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Enhancement of CO₂ permeability of poly(vinyl ether)s having oxyethylene chains by the sequence control of crosslinking sites

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Sequential living cationic copolymerizations of vinyl ethers having oxyethylene side chains (MOEO₂VE and MOEO₃VE) with vinyl ether having a crosslinkable group (VEEM) provided the ABA-typed triblock copolymers poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM)s and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM)s. All triblock copolymers are sticky liquids at room temperature, and the *T*_gs were −77 to −73 °C. Heating the triblock copolymers afforded membranes by thermal crosslinking *via* methacrylate groups at the outer segments in the polymers. All the membranes showed high CO₂ permselectivity ($P_{\text{CO}_2}/P_{\text{N}_2} = 40\text{--}51$) due to the high CO₂ solubility selectivity ($S_{\text{CO}_2}/S_{\text{N}_2} = 44\text{--}61$). The CO₂ permeability of the triblock copolymers was higher than that of the random copolymers with the same composition ratios. This indicates that the inner segment (MOEO₂VE and MOEO₃VE) effectively enhanced the gas diffusivity in the polymer matrix because the crosslinking points are present only in the outer segments.

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Introduction

Gas separation using polymer membranes becomes important due to the low costs and energy consumption compared with traditional separation techniques, such as adsorption and absorption.^{1–3} For CO₂ separation from mixed gases, a variety of polymers containing oxyethylene chains were studied because oxyethylene chains have highly favorable interactions with CO₂ molecules and preferentially dissolve CO₂ molecules.⁴ For example, poly(ethylene oxide) exhibits high CO₂ permselectivity ($P_{\text{CO}_2}/P_{\text{N}_2} = 50$), but the CO₂ permeability is low ($P_{\text{CO}_2} = 12$ barrer).⁵ Gas molecules generally go through amorphous regions in polymer matrices. The high crystallinity of poly(ethylene oxide) interferes with the diffusion of gases in the membrane. To enhance the gas permeability, an introduction of short oxyethylene chains to the rigid polymers and a crosslinking between short oxyethylene chains are investigated.^{6–12} For instance, the crosslinked membranes prepared from low-molecular-weight poly(ethylene glycol) diacrylate show high CO₂ permeability and high CO₂/N₂ selectivity ($P_{\text{CO}_2} = 100\text{--}150$ barrer, $P_{\text{CO}_2}/P_{\text{N}_2} = 45\text{--}65$) [6–10]. The P_{CO_2} and $P_{\text{CO}_2}/P_{\text{N}_2}$ values of poly(ether imide)s bearing amorphous poly(ethylene oxide) are 100–150 barrer and 45–65, respectively.¹¹

We have been studying the membranes of poly(vinyl ether)s for CO₂ separation.^{13–17} Among them, poly(vinyl ether)s having oxyethylene chains are wholly amorphous and have high content of oxyethylene units, and therefore they exhibit high gas

permeability and high CO₂ permselectivity. However, the homopolymers of vinyl ethers having oxyethylene chains are sticky liquid, whose *T*_g's are −50 to −20 °C. The copolymerization with other vinyl ethers such as 2-admantyl vinyl ether and vinyloxy ethoxy ethyl methacrylate [VEEM] is needed for the preparation of polymer membranes.^{14–17} For example, the copolymers of vinyl ether having oxyethylene chain with VEEM are heated in bulk to afford the membranes of crosslinked polymers by the reaction of methacrylate side groups. We reported that the gas permeability increased as the composition rate of VEEM in the copolymer decreased because the lack of crosslinking point enhanced the flexibility of polymer chains.^{15,17} The crosslinking of polymer chains is necessary to obtain membranes, but it decreases the gas permeability.

In the present study, we synthesized the triblock copolymers, poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM), with various composition ratios (Scheme 1). The triblock copolymers have the crosslinkable groups only at the outer sides in the polymer chains, and therefore the sequence of poly(MOEO₂VE) or poly(MOEO₃VE) is predicted to make the polymer chains flexible compared to the random copolymer with the same composition. In consequence, the triblock copolymer may show high CO₂ permeability.

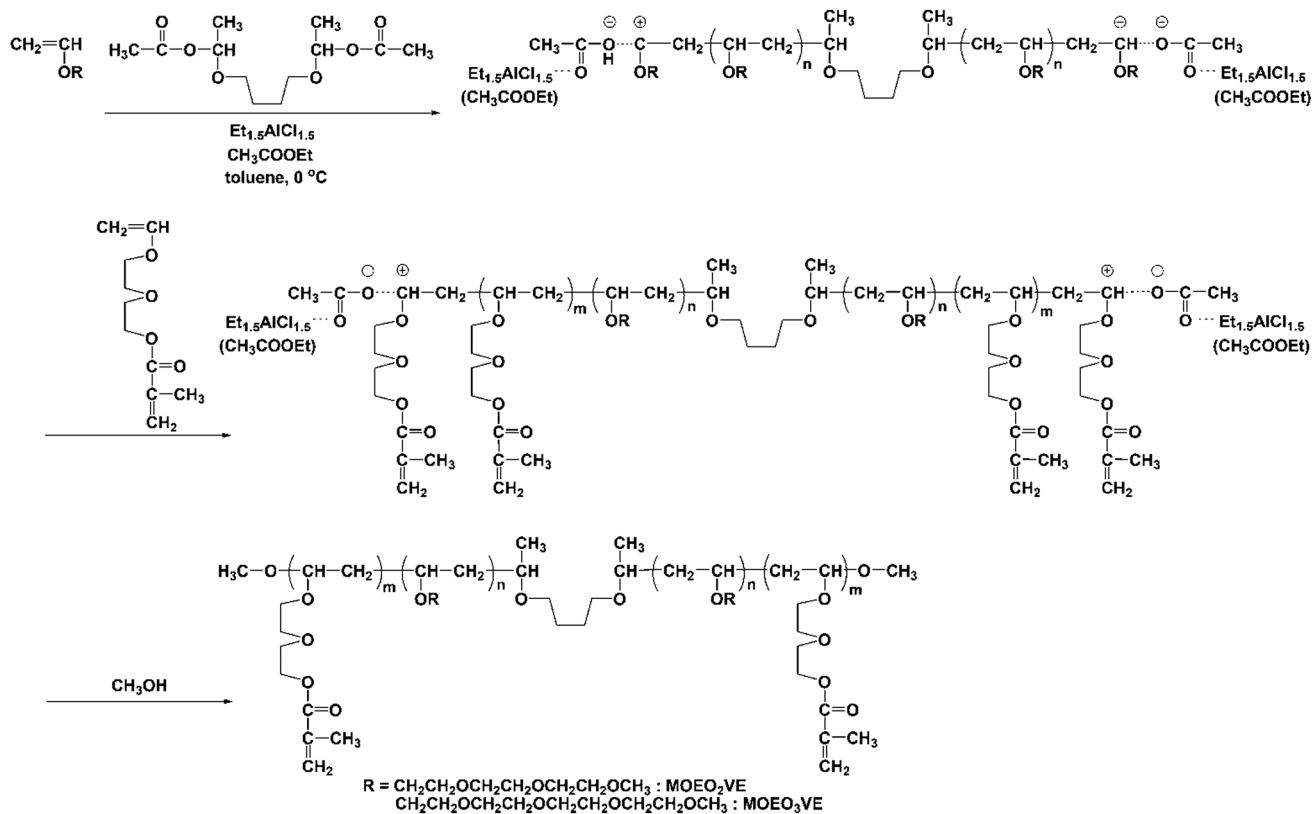
Experimental

Materials

Materials 2-[2-(2-methoxyethoxy)ethoxy]ethyl vinyl ether (MOEO₂-VE) was supplied by Maruzen Petrochemical Co. Ltd. 2-[2-[2-(2-methoxyethoxy)ethoxy]ethoxy]ethyl vinyl ether (MOEO₃VE) was synthesized by the reaction of 2-[2-(2-methoxyethoxy)ethoxy]ethyl

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Scheme 1 Synthesis of triblock copolymers, poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM).

methanesulfonate ($\text{CH}_3\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OCH}_2\text{CH}_2\text{OSO}_2\text{CH}_3$) with sodium salt of 2-hydroxyethyl vinyl ether ($\text{CH}_2=\text{CH}-\text{O}-\text{CH}_2\text{CH}_2-\text{OH}$), which was supplied by Nippon Carbide Industries Co. Inc. The detailed procedure was reported in our previous work.¹⁵ 2-(2-Vinylxyethoxy)ethyl methacrylate (VEEM) was supplied by Nippon Shokubai Co. Ltd. and distilled over *tert*-butylcatechol under reduced pressure. 1,4-Bis(1-acetoxyethoxy)butane [$\text{CH}_3\text{CH}(\text{OCOCH}_3)\text{OCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OCH}(\text{OCOCH}_3)\text{CH}_3$; BAEB] was synthesized by the treatment of 1,4-butanediol divinyl ether with acetic acid at 60 °C and purified by distillation under reduced pressure three times over calcium hydride. $\text{Et}_{1.5}\text{AlCl}_{1.5}$ (1.82 M solution in toluene) was commercially obtained from Sigma-Aldrich Co. Ltd. and used without further purification. Toluene for polymerization was distilled twice over calcium hydride.

Polymerization

The polymerizations were carried out under a dry nitrogen atmosphere in a glass flask equipped with a three-way stopcock, which was baked at *ca.* 400 °C on a heat gun just before the reaction. The toluene solution of cationogen (BAEB) and Lewis acid ($\text{Et}_{1.5}\text{AlCl}_{1.5}$) were prepared in the separate tubes. The solution of BAEB was added to a MOEO₂VE or MOEO₃VE solution containing internal standard (tetrahydronaphthalene) for gas chromatography. The solution of $\text{Et}_{1.5}\text{AlCl}_{1.5}$ was added to the mixture at 0 °C. MOEO₂VE or MOEO₃VE was first polymerized, and then VEEM (neat) was added to the reaction mixture under the same conditions. Polymerization was terminated by adding

excess amount of methanol (2.0 mL) containing a small amount of aqueous ammonia. Monomer conversion was determined from its residual concentration measured by gas chromatography. To work-up the polymer, the reaction mixture was diluted with CH_2Cl_2 and washed with 0.1 mol L⁻¹ hydrochloric acid and 10 wt% aqueous sodium chloride in this order. The organic solution was concentrated by rotary evaporator and the resultant polymer was isolated by precipitation into a large excess of hexane.

Fabrication of composite membrane

Composite membranes were prepared using a porous Teflon film (Advantech). The pore size, porosity, and thickness are 10 μm, 80%, and 50 μm, respectively. The porous Teflon film was put on a Petri dish, and the toluene-solution of uncrosslinked polymer obtained by the polymerization was poured into the dish. The solvent was slowly evaporated at room temperature under atmospheric pressure for a few days. After evaporation of the solvent, the Petri dish was placed in an oven at 110 °C for 15 min. This heating led the crosslinking between polymer side chains. The formed membrane of crosslinking polymer with Teflon film was cooled to room temperature and peeled from the dish. The membrane was dried at room temperature under reduced pressure for 24 h.

Measurements

The molecular weight distributions (MWDs) of polymers were measured by gel permeation chromatography (GPC) in



chloroform at 40 °C at 1.0 mL min⁻¹ flow rate on a Shimadzu LC-10AD chromatograph equipped with three polystyrene gel columns (Shodex K-807L, K-805L, and K-804L) and a Shimadzu RID-6A refractive index detector. The number-average molecular weight (M_n) and polydispersity ratio (M_w/M_n) were calculated from chromatograms based on a polystyrene calibration. ¹H NMR (500 MHz) spectra were recorded on Jeol ECX-500 instrument in CDCl₃ at room temperature. Scanning electron microscopy (SEM) was measured with HITACHI-HITEC S-3400N. Gas permeability coefficients were measured with a Tsukubarikaseiki K-315-N gas permeability apparatus at 25 °C under 1 atm upstream pressure using dry gases. The permeability coefficient (P) expressed in barrer unit (1 barrer = 10⁻¹⁰ cm³(STP) cm cm⁻² s⁻¹ cmHg⁻¹) was calculated from the slope of the steady-state line. The diffusivity coefficient (D) was determined by the time lag method using the following equation:

$$D = l^2/6\theta \quad (1)$$

here, l is the membrane thickness and θ is the time lag, which is given by the intercept of the asymptotic line of time-pressure curve to the time axis. The solubility coefficient (S) was calculated by using the following equation:

$$S = P/D. \quad (2)$$

Results and discussion

Polymerization

In our present study, the random copolymerization of MOEO₃VE with VEEM proceeded in the living manner using isobutyl vinyl ether-acetic acid adduct (CH₃CH(O^tBu)OCOCH₃) as an initiator, Et_{1.5}AlCl_{1.5} as a Lewis acid catalyst, and CH₃-COOEt as an added base in toluene at 0 °C.^{17,18} In the present study, MOEO₂VE was first polymerized using BAEB as a bifunctional initiator, Et_{1.5}AlCl_{1.5}, and CH₃COOEt in toluene at 0 °C to

synthesize the triblock copolymer. Then, VEEM was added to the reaction mixture and polymerized. Fig. 1 shows the GPC traces of the products obtained by the polymerization with the different monomer feed ratios. As shown in Fig. 1, all the MWD curves of the products showed the formation of the triblock copolymers, poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM), by the sequential polymerization described above. The peaks of the homopolymers shifted to the high-molecular-weight region by adding the second monomer, VEEM, with relatively narrow molecular weight distribution. No homopolymers peaks were observed in the MWD curves of the triblock copolymers, but other peaks appeared in higher-molecular-weight region ($M_n > 1 \times 10^6$) especially in the polymerization with high feed ratio of VEEM. These products seem to be formed by the intermolecular reaction through methacrylate units of VEEMs, but all the triblock copolymers were totally soluble in polymerization solvents. Fig. 2 shows the GPC curves of the products obtained by the sequential polymerizations of MOEO₃VE and VEEM at various monomer feed ratios. Similarly to the case of MOEO₂VE and VEEM, the sequential polymerizations produced the triblock copolymers, poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM), with relatively narrow molecular weight distributions. High-molecular-weight product formed by the intermolecular reaction was present in the polymerization with 2 : 1 monomer feed ratio. All the copolymers were soluble in toluene, chloroform, tetrahydrofuran, dichloromethane, *etc.* They did not dissolve in hexane and water.

Fig. 3 shows ¹H NMR spectra of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) obtained by the sequential polymerizations at the 2 : 1 monomer feed ratio. The peaks assignable to methacrylate group were observed at 6.12 and 5.58 ppm, and the intensity ratio of the signal at 6.12 ppm for methacrylate group to the signal at 4.27 ppm for the methylene adjacent to the ester oxygen was 1 to 2, which suggests that methacrylate groups of VEEM hardly react during cationic polymerization although the products obtained by

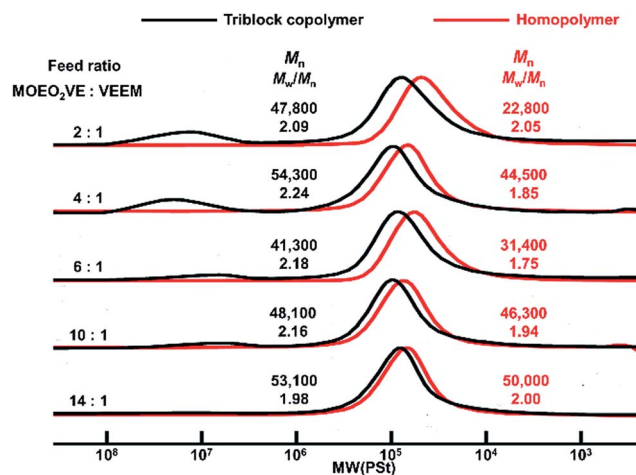


Fig. 1 GPC traces of the products: synthesis of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) with BAEB/Et_{1.5}AlCl_{1.5}/CH₃COOEt in toluene at 0 °C for 24 h: [monomer]_{0,total} = 1.2 M, [BAEB]₀ = 1.0 mM, [Et_{1.5}AlCl_{1.5}]₀ = 25 mM, [CH₃COOEt]₀ = 1.0 M.

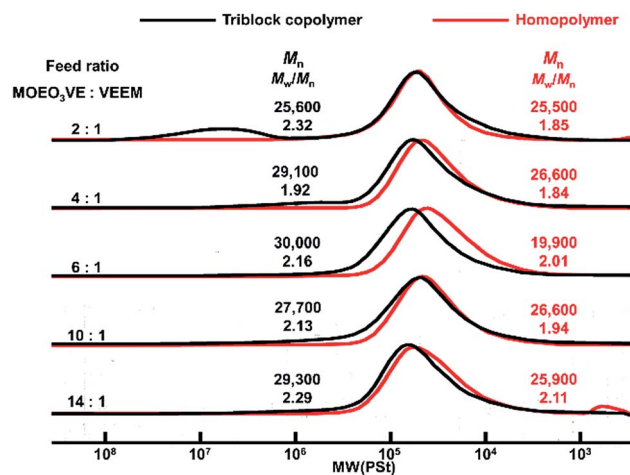


Fig. 2 GPC traces of the products: synthesis of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) with BAEB/Et_{1.5}AlCl_{1.5}/CH₃COOEt in toluene at 0 °C for 24 h: [monomer]_{0,total} = 1.2 M, [BAEB]₀ = 1.0 mM, [Et_{1.5}AlCl_{1.5}]₀ = 25 mM, [CH₃COOEt]₀ = 1.0 M.



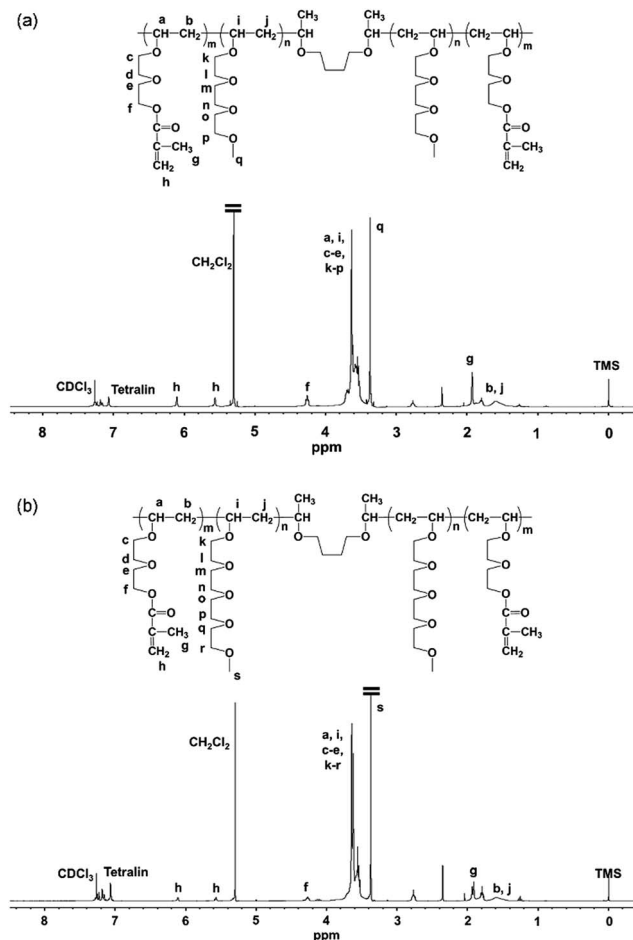


Fig. 3 ^1H NMR spectra of (a) poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (feed ratio 2 : 1) and (b) poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (feed ratio 2 : 1).

the intermolecular reaction were observed in GPC chromatograms. The copolymer compositions were determined by the peak intensity ratio of the signal for the methylene adjacent to the ester oxygen to the signals for the main chain methine and the pendant methylene and methyl group adjacent to the ether oxygen. For the poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM), the peak intensity of the signal for the methylene (f) was compared with the peak intensity of the signals for the main chain methine (a, i), pendant methylene (c-e, k-p), and methyl group (q). For the poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM), the peak intensity of the signal for the methylene (f) was compared with the peak intensity of the signals for the main chain methine (a, i), pendant methylene (c-e, k-r), and methyl group (s). The composition ratios calculated from ^1H NMR spectra corresponded well to the monomer feed ratios.

Thermal crosslinking and thermal property

Fig. 4(a), (c), (e) and (g) shows the IR spectra of the triblock copolymers obtained by the polymerization. Fig. 4(b), (d), (f) and (h) shows the IR spectra of the triblock copolymers heated at 110 °C in bulk. The absorptions at 1620 cm^{-1} assigned to the stretching of C=C double bonds and at 1710 cm^{-1} assigned to

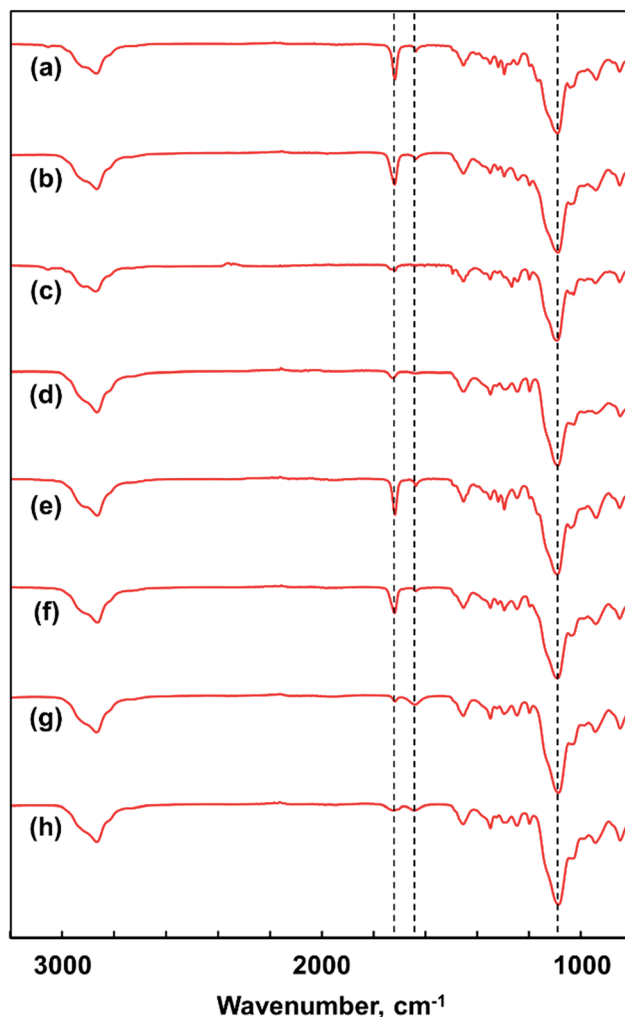


Fig. 4 IR spectra (a) before and (b) after heating for poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 2 : 1), (c) before and (d) after heating for poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 14 : 1), (e) before and (f) after heating for poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 2 : 1), (g) before and (h) after heating for poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 14 : 1).

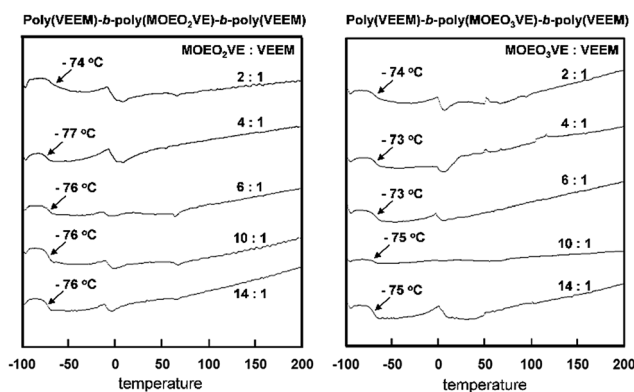


Fig. 5 DSC thermograms of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM)s and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM)s.



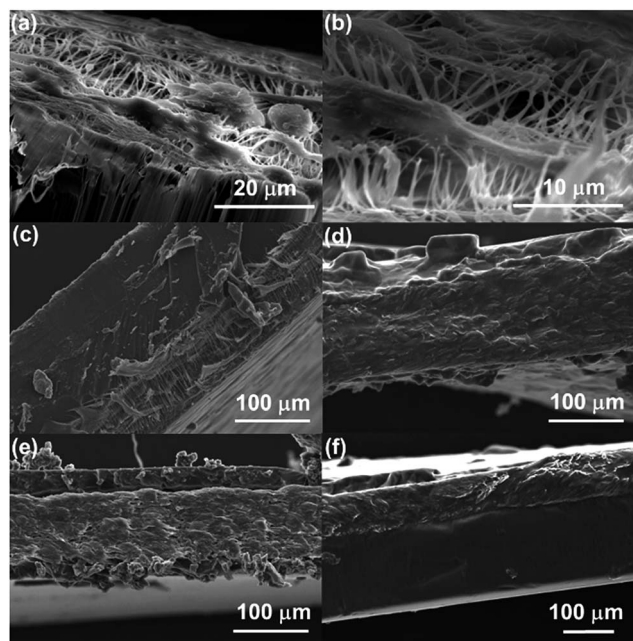


Fig. 6 SEM images of the cross-section of (a) porous PTFE film, (b) porous PTFE film, (c) poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 2 : 1), (d) poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 14 : 1), (e) poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 2 : 1), (f) poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 14 : 1).

the stretching of C=O carbonyl bonds were observed in all the spectra. However, the peak intensity of the C=C double bonds was decreased compared to the peak intensity of the absorption assigned to the stretching of C–O–C at 1080 cm⁻¹ after heating. The conversion of C=C double bonds was calculated from the

peak intensity at 1620 cm⁻¹ on the basis of the peak at 1080 cm⁻¹ attributed to C–O–C. The conversion of C=C double bonds calculated from the spectra (a) and (b) for poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) was approximately 50%. Those for the other triblock copolymers were in the range of 40–60%. The reaction between methacrylate groups in bulk did not proceed completely but reached around 50% conversion, whose result was the same as the case of the random copolymers, poly(MOEO₃VE)-*ran*-poly(VEEM), reported previously.¹⁷

Fig. 5 shows the DSC thermograms of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM). The prepared five poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM)s exhibited *T*_gs of –77 to –74 °C, which were very close to *T*_g of –73 °C of the homopolymer, poly(MOEO₂VE).¹⁶ The *T*_gs of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM)s were –75 to –73 °C, which were also similar to *T*_g of –72 °C of the homopolymer, poly(MOEO₃VE).¹⁶ In both types of triblock copolymers, the *T*_gs were independent on the composition ratio of MOEO₂VE/MOEO₃VE and VEEM. The low *T*_g is advantage for gas separation membrane because the flexible polymer chains enhances the gas diffusivity in the polymer matrix.

Gas permeability of crosslinked membranes

Fig. 6 shows SEM images of the cross-section of porous PTFE film (a and b) and the representative composite membranes (c–f). PTFE film had a filament structure and a lot of pores were observed in Fig. 6(a) and (b). On the other hand, no pores were observed in Fig. 6(c)–(f). This indicates that the copolymers penetrated into the pores of PTFE film and all the composite membranes were non-porous membranes. Since the pore size of PTFE film is as large as 10 μm, poly(vinyl ether) can infiltrate

Table 1 Gas permeability, diffusion, and solubility coefficients (*P*^a, *D*^b, *S*^c) of membranes at 25 °C

Composition	<i>P</i> _{N₂}	<i>P</i> _{O₂}	<i>P</i> _{CO₂}	<i>P</i> _{O₂} / <i>P</i> _{N₂}	<i>P</i> _{CO₂} / <i>P</i> _{N₂}	<i>D</i> _{N₂}	<i>D</i> _{CO₂}	<i>S</i> _{N₂}	<i>S</i> _{CO₂}	<i>D</i> _{CO₂} / <i>D</i> _{N₂}	<i>S</i> _{CO₂} / <i>S</i> _{N₂}
Poly(VEEM)-<i>b</i>-poly(MOEO₂VE)-<i>b</i>-poly(VEEM)											
2 : 1	4.5	13	230	2.9	51	8.4	7.3	0.53	31	0.87	58
4 : 1	7.5	18	310	2.4	41	14	11	0.54	30	0.79	56
6 : 1	7.6	18	340	2.3	45	12	11	0.63	32	0.91	51
10 : 1	8.4	20	370	2.4	45	14	12	0.59	33	0.85	56
14 : 1	9.3	22	420	2.4	45	16	14	0.60	31	0.88	52
Poly(VEEM)-<i>b</i>-poly(MOEO₃VE)-<i>b</i>-poly(VEEM)											
2 : 1	6.6	17	280	2.6	42	12	11	0.48	26	0.92	54
4 : 1	8.1	21	340	2.6	43	16	14	0.49	25	0.88	51
6 : 1	9.4	24	430	2.6	46	17	15	0.57	28	0.88	49
10 : 1	12	29	500	2.4	42	21	19	0.58	26	0.90	45
14 : 1	12	30	530	2.5	44	21	20	0.57	27	0.95	47
Poly(MOEO₃VE)-<i>ran</i>-poly(VEEM)											
2 : 1	3.3	8.9	140	2.7	42	11	7.4	0.30	19	0.85	51
4 : 1	4.9	12	220	2.4	45	10	8.7	0.49	25	0.93	55
6 : 1	6.7	17	300	2.5	45	18	13	0.37	23	0.86	55
10 : 1	7.9	19	360	2.4	46	21	15	0.38	24	0.88	52
14 : 1	9.2	23	420	2.5	46	18	16	0.51	26	0.89	57

^a In the unit of barrer (1 barrer = 1 × 10⁻¹⁰ cm³(STP) cm cm⁻² s⁻¹ cmHg⁻¹). ^b In the units of 1 × 10⁻⁷ cm² s⁻¹. ^c In the units of 1 × 10⁻³ cm³(STP) cm⁻³ cmHg⁻¹.



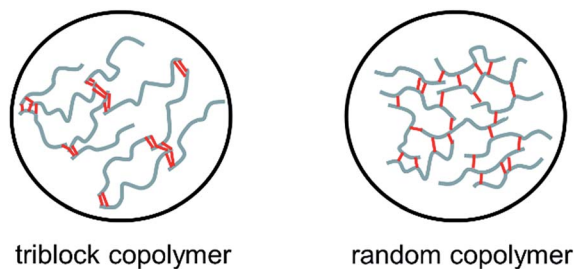


Fig. 7 Possible crosslinked structures of triblock-copolymer and random-copolymer.

into the pores. The composite membranes prepared in this study were smooth in surface and uniform in thickness.

The gas permeability of the composite membranes was examined at 25 °C. Table 1 summarizes the nitrogen, oxygen, carbon dioxide permeability coefficients (P_{N_2} , P_{O_2} , P_{CO_2}) and separation factors (P_{O_2}/P_{N_2} , P_{CO_2}/P_{N_2}), along with the data for random copolymers of poly(MOEO₃VE)-*ran*-poly(VEEM) for comparison. The gas permeability coefficients increased with the decreasing composition ratio of VEEM in the copolymers. For instance, the P_{CO_2} values of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) increased from 230 barrer to 420 barrer. The CO_2 values of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) increased from 280 barrer to 530 barrer. The low composition rates of VEEM means that the triblock copolymers had low crosslinking points. The lack of crosslinking points increases the polymer chain mobility. In the same composition ratio, the P_{CO_2} values of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) were higher than those of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM). From these results, it was found that the gas permeability increases with decreasing the composition ratio of VEEM and increasing the length of oxyethylene side chain. This tendency is the same as the case of random copolymers we reported in the previous study.^{15,17} Interestingly, the gas permeability of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) was higher than that of random copolymer, poly(MOEO₃VE)-*ran*-poly(VEEM), with the same

composition ratios, indicating that the monomer sequence in the copolymer is essential to the gas permeability. The random copolymers are crosslinked randomly in the main chains, whereas the triblock copolymers are crosslinked at outer segments in the main chains. The images of the crosslinked structures are illustrated in Fig. 7. The main chains of the triblock copolymers may become more flexible than the random copolymers, which would result in higher gas permeability. The O_2 separation factors (P_{O_2}/P_{N_2}) of the membranes were 2.3 to 2.9, which are ordinary values. On the other hand, the CO_2 permselectivity of all the copolymers was very high, and the separation factors (P_{CO_2}/P_{N_2}) were 41 to 51.

Gas permeability of non-porous polymer membranes can be divided into the diffusion and solution terms theoretically in the case of rubbery polymers. To inspect the gas permeability of the composite membranes in detail, the gas diffusion and solubility coefficients were estimated by time-lag method. Table 1 lists the diffusion and solubility coefficients (D and S) of the triblock and random copolymers. The S_{CO_2} values of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) and poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) were 25 to 33 cm³(STP) cm⁻³ cmHg⁻¹, and the S_{CO_2}/S_{N_2} values were as large as 45 to 58. The high solubility selectivity for CO_2 caused the high CO_2 permselectivity of the triblock copolymers. The D values increased with the decreasing composition ratio of VEEM in the same manner as the gas permeability coefficients. The D_{CO_2} value of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 14 : 1) was 14×10^{-7} cm² s⁻¹, which is twice as large as that of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM) (composition ratio 2 : 1). Similarly, the D value of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 14 : 1) was twice as large as that of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 2 : 1). The increment of gas permeability coefficient (P) caused by the decrease of VEEM ratio is originated mainly from the increase of gas diffusivity in the polymer matrix. The crosslinking points in polymer matrix restrict a motion of polymer chains, which lowers the gas diffusion in polymer matrix. For CO_2 permeation, poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-

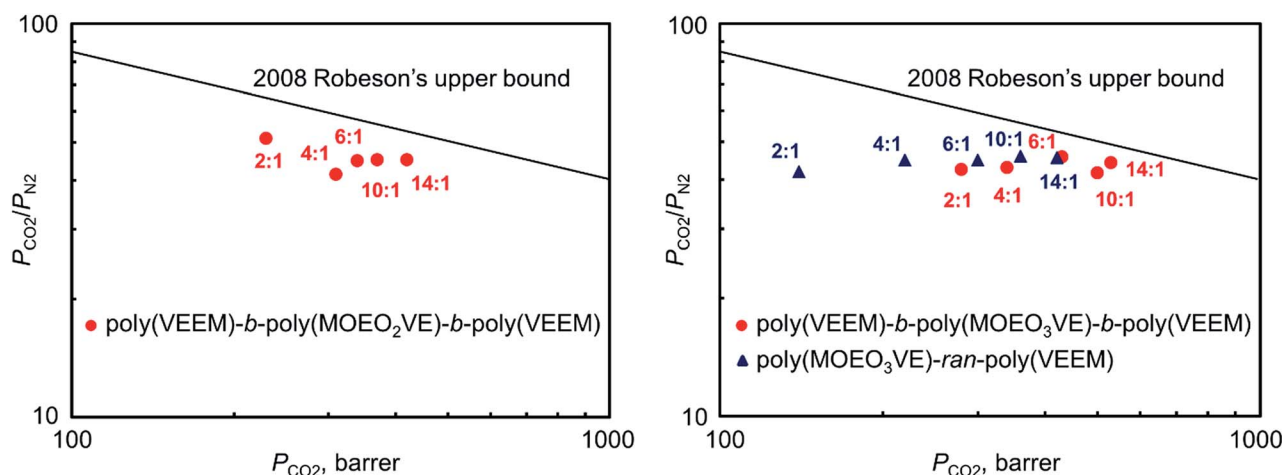


Fig. 8 Permeability–selectivity upper bound plots of poly(VEEM)-*b*-poly(MOEO₂VE)-*b*-poly(VEEM)s, poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM)s, and poly(MOEO₃VE)-*ran*-poly(VEEM)s for CO_2/N_2 gas pair.



poly(VEEM) showed slightly high diffusivity and solubility compared to the random copolymer, poly(MOEO₃VE)-*ran*-poly(VEEM). Therefore, the increment of flexibility of oxyethylene segment by block copolymerization enhanced not only the gas diffusivity but also the gas solubility in the polymer matrix.

The gas permeability–selectivity tradeoff plots for CO₂/N₂ gas pair of the present copolymers are shown in Fig. 8. The CO₂ permeability increased continuously with the decrease of the composition ratios of VEEM, and the CO₂ permselectivity of all the copolymers maintained high level around 50. The triblock copolymers obviously showed higher permeability than the random copolymers. Notably, the data of poly(VEEM)-*b*-poly(MOEO₃VE)-*b*-poly(VEEM) (composition ratio 14 : 1) was very close to the 2008 Robeson's upper bound.¹⁹

Conclusions

Highly CO₂-permeable and -permselective triblock copolymers were synthesized by the sequential cationic polymerizations of MOEO₂VE/MOEO₃VE and VEEM. The non-porous composite membranes with PTFE film were obtained by solution-cast followed by thermal crosslinking. The gas permeability increased with increasing the oxyethylene side chain length and decreasing the VEEM contents because the flexible polymer chains enhanced the gas diffusivity in polymer matrix. Noteworthy, the monomer sequence is dependent on the gas permeability. Triblock copolymers exhibited high gas permeability than the random copolymers even they have the same composition ratios. The triblock copolymers have the crosslinking only at the end segments, which makes the polymer chains more flexible than the random copolymers. These findings would contribute to the membrane field.

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