1. Introduction

Chemical–mechanical polishing/planarization (CMP) is a critical step in semiconductor circuits, which is employed after the metal deposition step to eliminate any topography on a wafer’s surface with the combination of chemical reactions and mechanical forces. Copper metallization is rapidly replacing the aluminum-based processes in the CMP process. Copper is high resistance to electromigration and low electrical compared to early used aluminum. This transition, however, poses new challenges to wastewater treatment. Copper CMP effluents can account for 30–40% of the water discharge in the semiconductor manufacturing and contain high concentrations of soluble copper and a complex mixture of organic constituents. Copper is toxic to aquatic organisms, which has been reported to cause toxic effects on algae and fish or amphibians at as low as 5–10 ppb. The dissolved copper in CMP wastewater is in the range of 5–100 mg L⁻¹. The 100 mg L⁻¹ concentration of copper in CMP wastewater has also been reported. It is thus evident that the copper concentration range of CMP wastewater is by far higher than copper effective concentration towards aquatic faunas and its concentration is strictly regulated in many countries both for discharges to municipal wastewater treatment facilities and for direct discharge to receiving waters. In China, the integrated circuit has been designated as a pillar industry in the 10th Five Year Economic plan since 2000. This has certainly driven large scale domestic and foreign semiconductor manufacturing plants in the past decade. High production volumes of semiconductor circuits bring forth a large amount of water consumption and increase the amount of wastewater discharged into the local environment. Therefore, an appropriate treatment of copper CMP wastewater is indispensable to meet stringent discharge regulations.

Various technologies have been employed for the removal of copper form copper CMP wastewater. Flocculation/clarification is a conventional treatment for the removal of copper ions. However, this process requires large tanks and generates an undesirable copper-rich sludge and secondary environmental pollution. Other methods such as those employing membranes, ion exchange, and electrochemical technologies have also been considered. Drawbacks associated with the abovementioned methods including the
addition of chemicals or high-cost limit their application. Accordingly, the need for the development of alternatives to copper CMP wastewater treatment has led to an increasing interest in sorption processes. Among various sorbents, bacterial biomass might be an efficient class of biosorbents for the removal of metal ions because of their high surface to volume ratios and presence of potentially active chemisorption sites. Both the living and non-living cells of bacteria can be used as biosorbents and have their own advantages and disadvantages. However, it is a controversial issue to choose living and non-living bacterial cells to remove heavy metals. Living cells might be preferential in the continuous treatment of heavy metal-bearing effluent and have been successful in the treatment of heavy metals in genuine industrial and municipal effluents. Biosorption and bioaccumulation often simultaneously participate in heavy metal removal by living bacterial cells. Hence, the use of heavy metal-resistant bacteria has gained extensive attention in the removal of heavy metals from aqueous solution because they can resist the toxicity of heavy metals even at high concentrations due to specific resistance systems, such as efflux and uptake mechanisms, and extracellular precipitation. Several studies have explored the copper removal from model copper CMP wastewater using metal tolerant bacteria from the genus Staphylococcus and Lactobacillus casei. However, bacterial mechanisms for the removal of various heavy metals are often highly specific for certain metals due to the heterogeneity of bacterial cell constituents. The high diversity and complex structure of bacteria imply that there are many ways for the metal to be bound by bacterial cells. Hence, there is much work that needs to be carried out in screening promising bacteria for their features and mechanisms for the bioremoval of heavy metal species.

The genus Cupriavidus is a Gram-negative β-proteobacteria that is well known for its resistance to heavy metals. Accordingly, Cupriavidus species is of great interest in the field of environmental biotechnology. Some Cupriavidus sp. such as C. metallidurans CH34, C. necator N-1, C. pinatubonensis JMP134, and C. taiwanensis have been reported to resist various heavy metals. In particular, C. metallidurans CH34 has been employed in the removal of heavy metals by biosorption and bioprecipitation in a bioreactor. To the best of our knowledge, the ability of C. gilardii in terms of removal of heavy metals has not been documented to date. In our previous study, we isolated a strain CR3 from a natural asphalt deposit. The whole genome of C. gilardii strain CR3 was grown in a Luria broth (LB) medium that contained 10.0 g tryptone, 5.0 g yeast extract, and 10.0 g NaCl per liter. Cells of C. gilardii CR3 were initially inoculated into sterile LB and placed on a shaker at 28 °C for 24 h. The live cells were harvested at an early stationary phase by centrifugation at 10 000g for 5 min. The cells were rinsed with phosphate buffer solution (PBS) three times and were subsequently stored in PBS at 4 °C. Cell concentrations in the suspension were determined by drying an aliquot onto a pre-weighed aluminum foil container to a constant weight at 60 °C.

Synthetic Cu-CMP wastewater used in this study was similar to that described by Mosier et al. (2015). Briefly, the stock synthetic Cu-CMP wastewater containing 5 mM Cu2+ was prepared with CuCl2·2H2O and deionized water. All chemicals used in this study were of analytical grade and were purchased from Sigma Aldrich Co., Ltd, USA. Biochemicals were obtained from Dingguo Biotechnology Co., Ltd., China.

2. Materials and methods

2.1 Preparation of biosorbent and synthetic Cu-CMP wastewater

The bacterium strain of C. gilardii CR3 used in this study was isolated from a natural asphalt deposit in our previous study. The GenBank accession number of C. gilardii CR3 is JX945578. C. gilardii CR3 was grown in a Luria broth (LB) medium that contained 10.0 g tryptone, 5.0 g yeast extract, and 10.0 g NaCl per liter. Cells of C. gilardii CR3 were initially inoculated into sterile LB and placed on a shaker at 28 °C for 24 h. The live cells were harvested at an early stationary phase by centrifugation at 10 000g for 5 min. The cells were rinsed with phosphate buffer solution (PBS) three times and were subsequently stored in PBS at 4 °C. Cell concentrations in the suspension were determined by drying an aliquot onto a pre-weighed aluminum foil container to a constant weight at 60 °C.

2.2 Scanning electron microscopy and energy dispersive X-ray spectroscopy

The concentrations of Cu2+ in the supernatant were determined via atomic absorption spectrophotometer (AA6300, Shimadzu, Japan). Scanning electron microscopy equipped with energy dispersive X-ray spectroscopy (SEM-EDX) was used to study the chemical characteristics of C. gilardii CR3 before and after Cu2+ binding. The biosorbent samples were fixed using 2.5% glutaraldehyde for 12 h and dehydrated in graded concentrations by alcohol. Samples were freeze-dried for 24 h and the dried cells were coated with a thin layer of gold and examined using SEM-EDX.

2.3 Fourier transform infrared spectroscopy

The possible involvement in the functional groups of cells of strain CR3 during the removal of Cu2+ from aqueous solution was elucidated using the Fourier transform infrared spectroscopy (FTIR) analysis. FTIR spectra of the cells loaded with and without Cu2+ were obtained in KBr discs using an FTIR
spectrometer (Perkin Elmer 2000). The FTIR samples were analyzed in the range of 4000–400 cm⁻¹ with a resolution of 4 cm⁻¹.

2.4 Zeta potential

The zeta potential for C. gilardii CR3 was measured before and after the Cu²⁺ binding using the Zeta Meter (Malvern, Nano ZS90). The concentration of Cu²⁺ was set at 64 mg L⁻¹ in the measurement of zeta potential versus pH of cells (2.0–7.0).

2.5 Batch experiments

All the batch experiments were carried out by mixing 1 g L⁻¹ biosorbent at ambient temperature (28 ± 2 °C) and the mixtures were agitated at the speed of 180 rpm for 24 h. Control experiments were carried out in the free-biosorbent in every test to investigate the removal of Cu²⁺, which might occur via chemical precipitation and sorption on to vessel walls.

The effects of pH, initial Cu²⁺ concentration of CMP wastewater, and contact time on biosorption and bioaccumulation capacity were studied. The effect of wastewater pH on the binding capacity of bacteria for Cu²⁺ was determined in the pH range from 2.0 to 7.0 with the intervals of 1.0 using 50 mL solutions containing 64 mg L⁻¹ of Cu²⁺. The Cu-CMP wastewater was adjusted to required pH values using 0.1 mol L⁻¹ HCl and 0.1 mol L⁻¹ NaOH. Biosorbent was added to Cu-CMP wastewater. Similarly, the effect of initial Cu²⁺ concentration on biosorption and bioaccumulation was conducted by changing the initial Cu²⁺ concentration to 6.4, 19.2, 32, 44.8, 64, 83.2, 96, 108.8, and 128 mg L⁻¹ (at optimum pH 5.0 for 24 h). In addition, the batch studies were carried out at room temperature (28 ± 2 °C) with 64 mg L⁻¹ of Cu²⁺ using 1 g L⁻¹ of biosorbent dosage at different time intervals between 5 and 400 min at pH 5.0. The Cu²⁺ binding capacity and removal efficiency of the tested bacterium were calculated according to the following equation:

\[ q_t = \frac{v(C_0 - C_t)}{m} \]  

(1)

Removal efficiency (%) = \( \frac{C_0 - C_e}{C_0} \times 100 \)  

(2)

where \( q_e \) is the amount of Cu²⁺ uptake at any given time in mg g⁻¹, \( v \) is the volume of Cu-CMP wastewater in the reaction system in L, \( m \) is the dosage of added biosorbents in g, and \( C_0 \) and \( C_e \) are the initial and equilibrium concentration of Cu²⁺ in the solution in mg L⁻¹, respectively.

2.6 Isotherms and kinetics

The experimental data were evaluated using a Langmuir isotherm and Freundlich isotherm equations. The linear form of the Langmuir isotherm and Freundlich isotherm equations can be expressed as follows:

\[ \frac{C_e}{q_e} = \frac{1}{q_m} + \frac{C_e}{q_m} \]  

(3)

\[ \ln q_e = \ln K_f + \frac{1}{n} \ln C_e \]  

(4)

where \( q_e \) is the amount of Cu²⁺ uptake at unit of sorbents in mg g⁻¹, \( C_e \) is the equilibrium concentration of Cu²⁺ in the solution in mg L⁻¹, \( q_m \) is the theoretical monolayer sorption capacity in mg g⁻¹, \( n \) and \( K_f \) are Freundlich equilibrium constants, and \( b \) is the Langmuir equilibrium constant.

To analyze the kinetics of Cu²⁺ using C. gilardii CR3, the linear form of the pseudo-second order kinetics and the intraparticle diffusion model were applied to the experimental data:

\[ \frac{t}{q_t} = \frac{1}{kq^2} + \frac{1}{q_t} \]  

(5)

\[ q_t = k_d t^n + d \]  

(6)

where \( k, k_d \), and \( d \) are the rate constant of kinetic sorption, \( q_t \) is the amount of Cu²⁺ uptake by unit of sorbents in mg g⁻¹, and \( q_t \) (mg g⁻¹) is the amount of Cu²⁺ uptake at any given time \( t \) (min).

2.7 Statistical analysis

All the experiments were performed in triplicate. The data is shown as mean values with the error bars representing the 95 percent confidence intervals around the average value.

Fig. 1 Scanning electron microscopy (SEM) images of Cupriavidus gilardii CR3 in the absence (a) and presence (b) for Cu²⁺ concentration of 64 mg L⁻¹ at pH 5.0 for 24 h.
3. Results and discussion

3.1 SEM-EDX analysis

Cell surface properties govern the metal–microorganism interactions to a large extent.\textsuperscript{31} The morphological characterization of \emph{C. gilardii} CR3 was studied before and after Cu\textsuperscript{2+} was loaded. After Cu\textsuperscript{2+} was loaded, the morphology of the bacterial cells did not show a greater variation in cell size and surface roughness. Moreover, the structural integrity of the cell was clear, without appearance of cell damage. Interestingly, reticulo-filamentous substances were obviously observed in the outer cells after Cu\textsuperscript{2+} was loaded (Fig. 1b), which were mostly like those secreted by \emph{C. gilardii} CR3 in the presence of ions from copper stress. This might reflect a self-protection of the strain CR3. Extracellular polymeric substances released by bacteria have been recommended as surface active agents and have an important role in binding heavy metals.\textsuperscript{32,33} The present results might be associated with the genetic characteristics of \emph{C. gilardii} CR3. Strain CR3 was isolated from a natural asphalt seep located at the Rancho La Brea Tar Pits (Los Angeles, USA), in which these natural asphalt seeps have existed for at least 40 000 year.\textsuperscript{34} The distinctive echo niche caused bacterium CR3 to develop intrinsic heavy metal resistance mechanisms including self-detoxification through ion efflux, metal-complexation and

Fig. 2  Energy dispersive X-ray spectroscopy (EDX) of \emph{Cupriavidus gilardii} CR3 in the absence (a) and presence (b) for Cu\textsuperscript{2+} concentration of 64 mg L\textsuperscript{−1} at pH 5.0 for 24 h.
metal-reduction, and a powerful DNA self-repair.\textsuperscript{27} Moreover, the EDX analysis confirmed that Cu\textsuperscript{2+} combined with the biomass of \textit{C. gilardii} CR3 (Fig. 2). There was no copper peak appearing on the cells before Cu\textsuperscript{2+} was loaded (Fig. 2a). In contrast, two additional signals of copper existed in the EDX profile after Cu\textsuperscript{2+} was loaded on the cells (Fig. 2b).

3.2 FTIR analysis

A number of sorption peaks that represented various functional groups of the cell surface were observed in the FTIR spectra (Fig. 3). The major sorption bands of \textit{C. gilardii} CR3 were as follows: 3307 cm\textsuperscript{-1}, corresponding to –OH and –NH stretching associated with protein; 2927 cm\textsuperscript{-1}, for asymmetric/symmetric stretching vibration of CH\textsubscript{2}; 1654, 1390, and 1072 cm\textsuperscript{-1}, for the C=O stretching of amide I, the symmetric stretching C=O peptide bond of carboxyl groups, and the stretching of PO\textsubscript{2}\textsuperscript{-} in phosphodiesters and phosphates, respectively; and 545 cm\textsuperscript{-1}, for phosphate or sulphate functional groups. In a comparison with the spectra of \textit{C. gilardii} CR3 prior to being treated with Cu\textsuperscript{2+}, a minute shift in the peak position was observed (<15 cm\textsuperscript{-1}) after Cu\textsuperscript{2+} binding. The band shifts were from 3307 cm\textsuperscript{-1} to 3309 cm\textsuperscript{-1} for –OH and –NH, from 1390 cm\textsuperscript{-1} to 1391 cm\textsuperscript{-1} for –COOH, and from 1072 cm\textsuperscript{-1} to 1069 cm\textsuperscript{-1} for PO\textsubscript{2}\textsuperscript{-} in the presence of Cu\textsuperscript{2+}. Functional groups such as carboxyl, hydroxyl, amino, and phosphate groups played a key role during the metal ion binding process.\textsuperscript{35,36} The present FTIR analysis also provided evidence that these functional groups were involved in the binding of copper for \textit{C. gilardii} CR3.

3.3 Zeta potential analysis

Bacterial surface charge is usually characterized by the zeta potential.\textsuperscript{37} To elucidate the mechanism that \textit{C. gilardii} CR3 sorbs Cu\textsuperscript{2+}, the zeta potentials of \textit{C. gilardii} CR3 were measured at different pH values with and without Cu\textsuperscript{2+} binding (Fig. 4). The isoelectric point of \textit{C. gilardii} CR3 was around 2.2. The zeta potential of cells changed from positive to negative once the pH exceeded 2.2, and the negative values of zeta potential increased with continuously increasing pH. The values of the zeta potential of the cell ranged from about +3 mV to about –15 mV. The variations of zeta potential in strain CR3 with pH revealed the electrostatic interaction between bacteria and Cu\textsuperscript{2+} in an aqueous system, which is consistent with the previous observations that most bacteria carry a net negative charge.\textsuperscript{35} Additionally, the obtained results indicated that the zeta potentials of the cell with

![Fig. 3](image-url)  
**Fig. 3** Fourier transform infrared spectroscopy (FTIR) of \textit{Cupriavidus gilardii} CR3 before and after binding the initial Cu\textsuperscript{2+} concentration of 64 mg L\textsuperscript{-1} at pH 5.0 for 24 h.

![Fig. 4](image-url)  
**Fig. 4** Zeta potential of \textit{Cupriavidus gilardii} CR3 before and after binding initial Cu\textsuperscript{2+} concentration of 64 mg L\textsuperscript{-1} at different pH values.
and without binding Cu\textsuperscript{2+} exhibited similar trends over the pH range from 2.0 to 7.0. The absolute values of the zeta potential of the cells after binding with Cu\textsuperscript{2+} were lower than those prior to binding at the pH range from 2.2 to 7.0.

### 3.4 Effect of pH on copper bioremoval

The pH value of the aqueous solution is an important variable in the binding of metals on the biosorbents. The influence of the initial pH value of the solution on the removal of Cu\textsuperscript{2+} was examined in the pH range from 2 to 7. The effects of pH on the binding of Cu\textsuperscript{2+} at various pH values (2.0–7.0) are shown in Fig. 5. The results showed that pH values significantly affect bacterial binding capacity to copper in the solution. The capacity of Cupriavidus gilardii CR3 and removal efficiency of Cu\textsuperscript{2+} increased with the elevation of the pH value. At the pH range of 2.0–5.0, the binding capacity sharply increased from 2.08 mg g\textsuperscript{-1} (removal 2% Cu\textsuperscript{2+}) to 18.33 mg g\textsuperscript{-1} (removal 27% Cu\textsuperscript{2+}). A 30% removal of Cu\textsuperscript{2+} occurred at pH between 6.0 and 7.0.

The pH value affects the solution chemistry of metal ions and the activity of the functional groups of the cell wall.\textsuperscript{38} Copper can be present in solution as three different species: Cu\textsuperscript{2+}, CuOH\textsuperscript{+}, and Cu(OH)\textsubscript{2}\.\textsuperscript{39} Under lower pH conditions, Cu\textsuperscript{2+} and CuOH\textsuperscript{+} were more favorable copper species. The competition of Cu\textsuperscript{2+} and protons in solution to the binding sites on the bacterium cell wall can explain the relationship between pH values and the uptake of metal ions.\textsuperscript{38} In this study, the binding capacities of copper were greatly influenced by the pH value of the solution. At pH 2.0, the lowest binding capacity was obtained. The reason for this is that H\textsubscript{3}O\textsuperscript{+} in the solution may more easily occupy the binding sites on the surface of Cupriavidus gilardii CR3 than copper. With an increase in pH (2.0–5.0), the amount of negative charges on the surface of the strain CR3 increased (Fig. 4), which led to the enhancement in binding capacity. When the pH value was over 5.3, the precipitation of Cu\textsuperscript{2+} hydroxide occurred. The formation of metal hydroxide and other metal–ligand complexes significantly reduced the amount of metal ions bound at higher pH.\textsuperscript{40} Therefore, the optimum pH condition was 5.0 in this study, and all the subsequent experiments were conducted at pH 5.0.

### 3.5 Effect of initial Cu\textsuperscript{2+} concentration on copper bioremoval

As depicted in Fig. 6, the Cu\textsuperscript{2+} binding capacity of Cupriavidus gilardii CR3 increased with the increase of initial Cu\textsuperscript{2+} concentration from 6.4 mg L\textsuperscript{-1} to 128 mg L\textsuperscript{-1}, and the removal efficiency inversely changed. The rapid removal of Cu\textsuperscript{2+} was observed under the lower Cu\textsuperscript{2+} condition (<64 mg L\textsuperscript{-1}). With an increase in Cu\textsuperscript{2+} concentration, the removal of Cu\textsuperscript{2+} became gradually slow. The lowest removal ratio was achieved at the highest initial Cu\textsuperscript{2+} concentration, where the equilibrium binding capacity for the cells was 22 mg g\textsuperscript{-1}. Therefore, the intersecting point of the two curves was the best ratio of the Cu\textsuperscript{2+} concentration and bacterial concentration.

### 3.6 Effect of contact time on copper bioremoval

The equilibrium time is of crucial importance because it is one of the parameters for economical wastewater treatment plant application. The effect of contact time on bioremoval of Cupriavidus gilardii CR3 at the initial concentration of 64 mg L\textsuperscript{-1} is shown in Fig. 7. The results revealed that the binding process as a function of contact time occurs in two steps. The first step involves a rapid metal uptake within the first 50 min of contact time, suggesting that this step was a rapid process. Within the first 50 minutes of binding, 76% of the total binding capacity was achieved. Rapid binding during the initial stage of the process was attributed to the larger surface area that was available for the binding of metal ions.\textsuperscript{41} The second step was characterized by the subsequent slower binding capacity of the metal. The exhaustion of binding sites during the process resulted in lower sorption rates during the second stage.\textsuperscript{42}

### 3.7 Isotherms study

Several sorption isotherms can be used to describe the equilibrium concentration of the sorbed species as a function of the fluid phase concentration. The Langmuir isotherm assumes monolayer sorption and the Freundlich isotherm is an empirical equation based on sorption for a heterogeneous surface.\textsuperscript{43,44} In the present study, the linearized Langmuir and Freundlich

![Fig. 5](image_url) Effect of pH on copper bioremoval of Cupriavidus gilardii CR3 with an initial Cu\textsuperscript{2+} concentration of 64 mg L\textsuperscript{-1} for 24 h.

![Fig. 6](image_url) Effect of initial Cu\textsuperscript{2+} concentration on copper bioremoval of Cupriavidus gilardii CR3 at pH 5.0 for 24 h.

![Fig. 7](image_url) Effect of contact time on copper bioremoval.
isotherms at various initial concentrations of copper ions were employed to determine the sorption isotherm data (Fig. 8). The correlation coefficient and other parameters obtained from Langmuir and Freundlich sorption isotherms for the sorption of Cu$^{2+}$ are shown in Table 1. The Langmuir model ($R^2 = 0.99$) exhibited a better fit to the data of copper than the Freundlich model ($R^2 = 0.49$) in the studied initial Cu$^{2+}$ concentration range. It could be inferred that the binding of the cells of C. gilardii CR3 to Cu$^{2+}$ is monolayer sorption. Most of bacterial binding for metals follow the Langmuir isotherm.\textsuperscript{45–47} According to Langmuir isotherm, the predicted maximum biosorption capacity for C. gilardii CR3 was 18.25 mg g\(^{-1}/C_0\), which is very close to the experimental binding capacity \(q_e\) value (18.33 mg g\(^{-1}\)). These results suggest that the binding of copper by strain CR3 principally occurred at cell sorption rather than accumulation because the maximum binding capacity calculated by the Langmuir isotherm did not consider the amount of metal uptake by bacteria.\textsuperscript{48} Although both biosorption and bioaccumulation are involved in living cells for the removal of various heavy metals from aqueous solution, biosorption usually accounts for the main part in the binding capacity of heavy metals.\textsuperscript{49–51} For instance, the cell surface of a Lactococcus lactis strain accounted for over 95\% of the amount of cadmium binding.\textsuperscript{52} Additionally, a comparison within Cupriavidus strains showed that the copper biosorption capacity of C. gilardii CR3 was almost the same as that of the C. taiwanensis TJ208 (19.0 mg g\(^{-1}\)), whereas lower than that of C. metallidurans CH34 (26.69 mg g\(^{-1}\)) under the same experiment conditions.\textsuperscript{53} Therefore, it is still efficient compared to other bacteria in terms of copper biosorption capacity: the copper biosorption capacity of Bacillus sp. is 16.25 mg g\(^{-1}/C_0\) (ref. 54) and Sphaerotilus natans is 5.4 mg g\(^{-1}\).\textsuperscript{55} Thus, C. gilardii CR3 can be an interesting alternative for the bioremoval of Cu$^{2+}$ from Cu-CMP wastewater.

### 3.8 Kinetics study

Kinetic studies are important for the practical application of binding process because they can provide insight into the reaction pathways and the mechanisms of the sorption reaction.\textsuperscript{56} To investigate the kinetic patterns of copper on C. gilardii CR3, the pseudo-second order kinetic model and intraparticle diffusion model were applied in this study (Fig. 9). The kinetics of copper binding by strain CR3 was well described by the pseudo-second order kinetic model ($R^2 = 0.99$) and intraparticle diffusion model ($R^2 = 0.98$) (Table 2). The experimental data fitting the pseudo-second order model demonstrated that the mechanism of the sorption process was controlled by chemical
sorption involved in exchange sorption of metal ions and a surface proton. Moreover, the fit line with an intraparticle diffusion model did not pass through the origin, suggesting that intraparticle diffusion occurred in the binding process, but it was not the sole rate-controlling step. Furthermore, the amount of negative charges on the surface of strain CR3 increased with an increase in pH value based on the measurement of the zeta potential (Fig. 4), which confirmed that the electrostatic interaction contributed to the binding process. Therefore, chemisorption and physisorption might be simultaneously involved in the binding process by means of ion exchange and electrostatic interactions between copper ions and functional groups in the cells of strain CR3.

4. Conclusions

The present study was carried out to use living cells of *C. gilardii* CR3 as a biosorbent for the removal Cu$^{2+}$ from Cu-CMP wastewater. The binding ability was pH-dependent, and the highest binding capacity of copper ions (18.33 mg g$^{-1}$) was achieved at pH 5.0. The biosorption equilibrium was well described by the Langmuir isotherm. Both the pseudo-second order kinetic model and the intraparticle diffusion model were in good agreement with all the experimental data. The obtained results may provide pre-stage for an initial evaluation of binding mechanisms and operational conditions for the continuous removal of Cu$^{2+}$ from CMP wastewater utilizing living cells of *C. gilardii* CR3 in a fixed-bed reactor.

Authors' contributions

XW and FW conceived and designed the experiments. YY and XW performed the experiments. MH and DZ analyzed the data. MH and MH contributed reagents, materials, and analysis tools. XW and YY wrote the paper. XW and FW revised the paper.

Ethics approval and consent to participate

Not applicable.

Conflict of interest

The authors declare that they have no competing interests.

Abbreviations

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<thead>
<tr>
<th>Abbreviation</th>
<th>Description</th>
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<tbody>
<tr>
<td>CMP</td>
<td>Chemical–mechanical polishing/planarization</td>
</tr>
<tr>
<td>LB</td>
<td>Luria broth</td>
</tr>
<tr>
<td>PBS</td>
<td>Phosphate buffer solution</td>
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<tr>
<td>SEM-EDX</td>
<td>Scanning electron microscopy equipped with energy dispersive X-ray spectroscopy</td>
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<tr>
<td>FTIR</td>
<td>Fourier transform infrared spectroscopy</td>
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