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Preparation of protic ionic liquids containing cyclic oligosiloxane frameworks

T. Hirohara,^a T. Kai,^b J. Ohshita^b and Y. Kaneko^{*a}

Protic ionic liquids (PILs) containing cyclic oligosiloxanes with ammonium side-chain groups (**Am-CyS-IL** and **2Am-CyS-IL**) were successfully prepared by the hydrolytic condensation of 3-aminopropyldimethoxymethylsilane and 3-(2-aminoethylamino)propyldimethoxymethylsilane, respectively, in a water/methanol (1:19 v/v) mixed solution of bis(trifluoromethanesulfonyl)imide (HNTf₂). The differential scanning calorimetry (DSC) curves of **Am-CyS-IL** and **2Am-CyS-IL** exhibited baseline shifts attributable to glass transition temperatures (T_g) of -2 °C and 9 °C, respectively, and endothermic peaks at their melting temperatures (T_m) were not detected for both compounds, indicating their amorphous structures. In addition, fluidity was visually observed for each compound at ~ 35 °C and ~ 45 °C, respectively. On the basis of these results, we concluded that **Am-CyS-IL** and **2Am-CyS-IL** were PILs. To investigate the influence of the kind of siloxane framework on the nature of the ionic liquid, we prepared polyhedral oligomeric silsesquioxanes (POSSs) containing the same side-chain groups and counteranions (**Am-POSS** and **2Am-POSS**) from 3-aminopropyltrimethoxysilane and 3-(2-aminoethylamino)propyltrimethoxysilane, respectively, using an HNTf₂ catalyst. **Am-POSS** had a T_g of 15 °C and a T_m of 193 °C, and **2Am-POSS** had a T_g of 32 °C and a T_m of 208 °C, which were estimated by DSC analyses. Furthermore, fluidity was not observed below 150 °C for both compounds, indicating that these POSS compounds were not PILs. These results suggest that the cyclic oligosiloxane frameworks are an important factor for the preparation of siloxane-based PILs exhibiting fluidity at low temperature. In addition, the present PILs containing cyclic oligosiloxane frameworks exhibited relatively high decomposition temperatures (T_d : **Am-CyS-IL**: $T_{d5} = 351$ °C and $T_{d10} = 362$ °C, **2Am-CyS-IL**: $T_{d5} = 294$ °C and $T_{d10} = 309$ °C).

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Introduction

Ionic liquids (ILs) are generally defined as molten salts below 100 °C or 150 °C. ILs are classified as aprotic ILs (AILs) or protic ILs (PILs) depending on the absence or presence of protons. PILs are prepared by a simple neutralization (proton transfer) reaction between a Brønsted acid and a Brønsted base.^{1,2} Therefore, they have the advantages of being easier to prepare and cheaper compared with AILs. So far, PILs have been widely studied due to their remarkable potential as electrolytes for fuel cells,³ lithium-ion batteries⁴ and capacitors,⁵ as biomass extraction solvents,⁶ and for carbon dioxide captures.^{7,8} In particular, because PILs have relatively good thermal stability, they are expected to be applied as thermostable proton conductors (solid electrolytes) for fuel cells. However, the decomposition temperatures (T_d) of PILs on pyrolysis are generally lower than those of AILs because of proton transfer

from the cation back to the anion to reform the original acid and base neutral species.¹

So far, in the field of AILs, ILs containing siloxane frameworks have been reported.^{9–14} Siloxane compounds exhibit excellent thermal stability due to the high binding energy of their Si–O–Si linkages. An AIL containing polyhedral oligomeric silsesquioxane (POSS), a siloxane framework, was first prepared by Chujo, Tanaka, and co-worker.⁹ This POSS-AIL contained carboxylate anion side chains and imidazolium cations as counterions. Zhang *et al.* have also reported the preparation of an AIL containing POSS with imidazolium cations side chains and dodecyl sulfate anions as counterions.¹⁰ Meanwhile, we prepared a highly thermostable POSS-AIL containing imidazolium cationic side-chains and bis(trifluoromethanesulfonyl)imide (NTf₂) anions as counterions by a simple hydrolytic condensation method.¹¹ In addition, AILs containing randomly structured oligomeric silsesquioxanes with imidazolium¹¹ and quaternary ammonium¹² side-chain groups were also prepared. More recently, a room-temperature POSS-AIL containing two types of cationic side-chain groups, *i.e.* imidazolium and quaternary ammonium groups, and NTf₂ anions as counterions was prepared by a similar hydrolytic condensation method.¹³ Furthermore, as an example of AILs containing other types of

^aGraduate School of Science and Engineering, Kagoshima University, 1-21-40 Korimoto, Kagoshima 890-0065, Japan. E-mail: ykaneko@eng.kagoshima-u.ac.jp; Fax: +81 99 285 7794; Tel: +81 99 285 7794

^bGraduate School of Engineering, Hiroshima University, Higashi-Hiroshima, Hiroshima 739-8527, Japan



siloxane frameworks, we prepared imidazolium salt type AILs containing cyclic siloxanes.¹⁴ In these siloxane-based AILs prepared by us, the use of a superacid catalyst is essential for structural control and indicating IL nature. The T_d values of the aforementioned AILs containing siloxane frameworks were higher than those of AILs with side-chain structures of the corresponding siloxane-based AILs. Therefore, PILs containing siloxane frameworks are expected to exhibit higher thermal stability. However, to the best of our knowledge, the preparation of PILs containing siloxane frameworks has not yet been reported.

To successfully prepare such siloxane-based PILs, we referred to our previous study of the facile preparation of a single-structured cyclic tetrasiloxane containing ammonium groups by the hydrolytic condensation of 3-aminopropyldiethoxymethylsilane using aqueous superacid trifluoromethanesulfonic acid (HOTf) as a catalyst.¹⁵ However, this cyclic tetrasiloxane was not an IL, *i.e.* its melting temperature (T_m) is *ca.* 310 °C. In general, it has been reported that T_m and/or flow temperatures of ILs containing NTF₂ anions are lower than those containing OTf anions with the same cationic species.^{14,16}

In this study, when the hydrolytic condensation of dialkoxysilanes containing primary and/or secondary amino groups, such as 3-aminopropyldimethoxymethylsilane (APDMMS) and 3-(2-aminoethylamino)propyldimethoxymethylsilane (AEAPDMMS), were investigated using superacid HNTf₂ as a catalyst, we found that PILs containing cyclic oligosiloxane frameworks were successfully prepared. Here we report the preparation, characterizations and thermal properties of these PILs, and a comparison of the properties with POSSs containing the same side-chain groups and counterions.

Experimental

Materials

APDMMS, AEAPDMMS, 3-aminopropyltrimethoxysilane (APTMS) and 3-(2-aminoethylamino)propyltrimethoxysilane (AEAPTMS) were purchased from Tokyo Chemical Industry Co., Ltd. (Japan). Other reagents and solvents were purchased from Wako Pure Chemical Industries, Ltd. (Japan). All reagents and solvents were used without further purification.

Preparation of PILs containing cyclic oligosiloxane with one ammonium group per repeating unit (Am-CyS-IL)

A water/methanol (1 : 19 v/v) mixed solution of HNTf₂ (0.5 mol L⁻¹, 15 mL, 7.5 mmol) was added to APDMMS (purity: 97%, 0.505 g, 3.0 mmol) with stirring at room temperature. After the resulting solution was further stirred for 2 h, this solution was heated at ~60 °C in an open container until the solvent evaporated. Subsequently, the resulting viscous product was maintained at 100 °C for ~2 h, dissolved in methanol (1 mL) at room temperature, and then this methanol solution was poured into chloroform (30 mL). The chloroform-insoluble part was isolated by decantation, washed with chloroform (~30 mL × 3), and then dried by heating at 150 °C for ~10 h to yield 1.019 g of a viscous product (Am-CyS-IL; yield, ~85%; the ideal chemical formula of one repeating unit of the product

[CH₃SiO(CH₂)₃NH₃·(CF₃SO₂)₂N, FW = 398.4] was used for this determination). ¹H NMR (400 MHz, DMSO-*d*₆): δ 7.80–7.43 (3H, br, -NH₃), δ 2.83–2.64 (2H, br, -CH₂NH₃), δ 1.61–1.41 (2H, br, -SiCH₂CH₂CH₂NH₃), δ 0.63–0.37 (2H, br, -SiCH₂-), δ 0.19 to -0.03 (3H, br, -SiCH₃). ²⁹Si NMR (79.5 MHz, DMSO-*d*₆): δ -19.3 to -19.8 (D², cyclic tetramer), δ -21.6 to -21.9 (D², cyclic pentamer), δ -22.2 to -22.5 (D², cyclic hexamer).

Preparation of PILs containing cyclic oligosiloxane with two ammonium groups per repeating unit (2Am-CyS-IL)

A water/methanol (1 : 19 v/v) mixed solution of HNTf₂ (0.5 mol L⁻¹, 30 mL, 15 mmol) was added to AEAPDMMS (purity: 97%, 0.638 g, 3.0 mmol) with stirring at room temperature. The subsequent procedures were the same as those described above for the preparation of Am-CyS-IL, and 1.896 g of a viscous product (2Am-CyS-IL) was obtained (yield, ~87%; the ideal chemical formula of one repeating unit of the product [CH₃-SiO(CH₂)₃NH₂(CH₂)₂NH₃·2(CF₃SO₂)₂N, FW = 722.7] was used for this determination). ¹H NMR (400 MHz, DMSO-*d*₆): δ 8.56–8.14 (2H, br, -NH₂CH₂CH₂NH₃), δ 7.99–7.60 (3H, br, -NH₂-CH₂CH₂NH₃), δ 3.20–2.97 (4H, br, -NH₂CH₂CH₂NH₃), δ 2.97–2.75 (2H, br, -SiCH₂CH₂CH₂NH₂-), δ 1.69–1.38 (2H, br, -SiCH₂CH₂CH₂NH₂-), δ 0.70–0.36 (2H, br, -SiCH₂-), δ 0.27 to -0.01 (3H, br, -SiCH₃). ²⁹Si NMR (79.5 MHz, DMSO-*d*₆): δ -19.1 to -20.1 (D², cyclic tetramer), δ -21.3 to -21.8 (D², cyclic pentamer), δ -22.1 to -22.3 (D², cyclic hexamer).

Preparation of POSS with one ammonium group per repeating unit (Am-POSS)

An aqueous HNTf₂ solution (0.5 mol L⁻¹, 15 mL, 7.5 mmol) was added to APTMS (purity: 96%, 0.560 g, 3.0 mmol) with stirring at room temperature. The subsequent procedures were the same as those described above for the preparation of Am-CyS-IL, and 1.150 g of a solid substance (Am-POSS) was obtained (yield, ~98%; the ideal chemical formula of one repeating unit of the product [SiO_{1.5}(CH₂)₃NH₃·(CF₃SO₂)₂N, FW = 391.4] was used for this determination). ¹H NMR (400 MHz, DMSO-*d*₆): δ 7.70–7.46 (3H, br, -NH₃), δ 2.84–2.63 (2H, br, -CH₂NH₃), δ 1.63–1.40 (2H, br, -SiCH₂CH₂CH₂NH₃), δ 0.74–0.52 (2H, br, -SiCH₂-). ²⁹Si NMR (79.5 MHz, DMSO-*d*₆): δ -66.5 (T₈), δ -68.5 (T₁₀).

Preparation of POSS with two ammonium groups per repeating unit (2Am-POSS)

Aqueous HNTf₂ solution (0.5 mol L⁻¹, 30 mL, 15 mmol) was added to AEAPTMS (purity: 95%, 0.702 g, 3.0 mmol) with stirring at room temperature. The subsequent procedures were the same as those described above for the preparation of Am-CyS-IL, and 2.110 g of a solid substance (2Am-POSS) was obtained (yield, ~98%; the ideal chemical formula of one repeating unit of the product [SiO_{1.5}(CH₂)₃NH₂(CH₂)₂NH₃·2(CF₃SO₂)₂N, FW = 715.6] was used for this determination). ¹H NMR (400 MHz, DMSO-*d*₆): δ 8.48–8.23 (2H, br, -NH₂CH₂CH₂NH₃), δ 7.96–7.69 (3H, br, -NH₂CH₂CH₂NH₃), δ 3.20–2.91 (4H, br, -NH₂CH₂CH₂-NH₃), δ 2.91–2.68 (2H, br, -SiCH₂CH₂CH₂NH₂-), δ 1.69–1.33 (2H, br, -SiCH₂CH₂CH₂NH₂-), δ 0.78–0.40 (2H, br, -SiCH₂-). ²⁹Si NMR (79.5 MHz, DMSO-*d*₆): δ -67.3 (T₈), δ -69.4 (T₁₀).



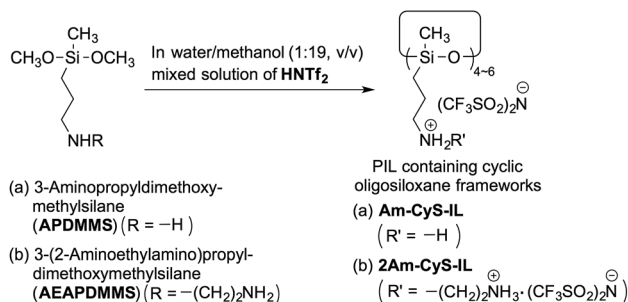
Measurements

^1H and ^{29}Si NMR spectra were recorded using a JEOL ECX-400 spectrometer (JEOL Ltd.). The atomic ratios of Si : S in the products were confirmed by energy-dispersive X-ray (EDX) analyses using an XL30 scanning electron microscope (FEI Co.). Matrix-assisted laser desorption ionization time-of-flight mass spectral (MALDI-TOF MS) analyses of the products were performed using a Shimadzu Voyager Biospectrometry Workstation Ver. 5.1 (SHIMADZU Co.) positive ion mode with 2,5-dihydroxybenzoic acid (DHB) as the matrix. Attenuated total reflectance infrared (ATR-IR) spectra were recorded using a JASCO FTIR-4200 spectrometer (JASCO Co.). Differential scanning calorimetry (DSC) analyses were performed using a DSC-60 Plus (SHIMADZU Co.). The sample was placed in an aluminium capsule and cooled to $-140\text{ }^\circ\text{C}$ at a rate of $20\text{ }^\circ\text{C min}^{-1}$ under a nitrogen flow (50 mL min^{-1}), and then heated from $-140\text{ }^\circ\text{C}$ to $250\text{ }^\circ\text{C}$ at the same rate. Glass transition temperatures (T_g) and T_m were determined as the onset of baseline shifts and as the tops of endothermic peaks, respectively, in the third set of curves (from $-140\text{ }^\circ\text{C}$ to $250\text{ }^\circ\text{C}$ at a rate of $20\text{ }^\circ\text{C min}^{-1}$) to eliminate the heat histories of the samples. X-ray diffraction (XRD) patterns were recorded at a scanning speed of $2\theta = 6.6^\circ\text{ min}^{-1}$ using an X'Pert Pro diffractometer (PANalytical) with Ni-filtered Cu $K\alpha$ radiation (0.15418 nm). Thermogravimetric analyses (TGA) were performed on an Exstar TG/DTA6200 (Seiko Instruments) at a heating rate of $10\text{ }^\circ\text{C min}^{-1}$ up to $1000\text{ }^\circ\text{C}$ under a nitrogen flow (250 mL min^{-1}). The decomposition temperatures were determined from the 5% (T_{d5}) and 10% (T_{d10}) weight losses.

Results and discussion

Preparation and characterizations of Am-CyS-IL and 2Am-CyS-IL

A PIL containing a cyclic oligosiloxane with one ammonium group per repeating unit, **Am-CyS-IL**, was prepared according to the following procedure (Scheme 1a): **APDMMS** was stirred in a water/methanol (1 : 19 v/v) mixed solution of HNTf_2 (0.5 mol L^{-1}) at room temperature for 2 h. Here, an excess amount of HNTf_2 was required (the feed molar ratio of $\text{HNTf}_2/\text{APDMMS}$ was 2.5). The resulting solution was heated at *ca.* $60\text{ }^\circ\text{C}$ in an open container until the solvent evaporated. The resulting



Scheme 1 Preparation of PILs containing cyclic oligosiloxane frameworks: (a) Am-CyS-IL and (b) 2Am-CyS-IL.

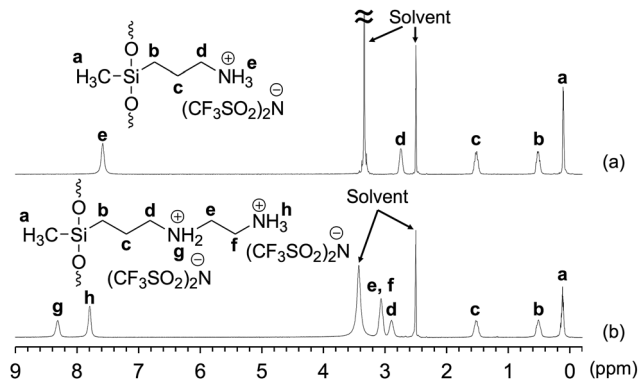


Fig. 1 ^1H NMR spectra of (a) Am-CyS-IL and (b) 2Am-CyS-IL in $\text{DMSO-}d_6$. Chemical shifts were referenced to DMSO (δ 2.5).

crude product was further heated at $100\text{ }^\circ\text{C}$ for 2 h, washed with chloroform, and then dried at $150\text{ }^\circ\text{C}$ for $\sim 10\text{ h}$ to obtain **Am-CyS-IL**. **Am-CyS-IL** was soluble in high-polarity solvents, such as water, methanol and acetone, but insoluble in low-polarity solvents, such as chloroform, toluene and hexane.

The ^1H NMR spectrum of **Am-CyS-IL** in $\text{DMSO-}d_6$ exhibited signals attributable to 3-aminopropyl and methyl groups, but signals due to the methoxy groups of **APDMMS** were not observed (Fig. 1a), indicating that the **APDMMS** reagent was not present in the product. The EDX pattern of **Am-CyS-IL** revealed that the Si : S atomic ratio was *ca.* 1.00 : 1.99, indicating that the molar ratio of the ammonium cation to the NTf_2^- anion in **Am-CyS-IL** was *ca.* 1 : 1 (Fig. 2a). In the MALDI-TOF MS analysis of **Am-CyS-IL**, several peaks were observed that corresponded to the masses of cyclic siloxane tetramer, pentamer and hexamer (Fig. 3a). The ^{29}Si NMR spectrum of **Am-CyS-IL** in $\text{DMSO-}d_6$ at $40\text{ }^\circ\text{C}$ showed three multiple signals in the D^2 region (-19.3 to -19.8 ppm for cyclic tetrasiloxane, -21.6 to -21.9 ppm for cyclic pentasiloxane, and -22.2 to -22.5 ppm for cyclic hexasiloxane) and no signals in the D^1 region at *ca.* -15 ppm (Fig. 4a). Furthermore, the aforementioned ^1H NMR spectrum contained multiple signals, a, attributable to methyl groups at 0.20 to -0.01 ppm (Fig. 1a). These results indicate that **Am-CyS-IL** was a mixture of cyclic tetra-, penta-, and hexasiloxanes, with some stereoisomers.

A PIL containing a cyclic oligosiloxane with two ammonium groups per repeating unit, **2Am-CyS-IL**, was prepared using almost the same procedure as that of **Am-CyS-IL** but using

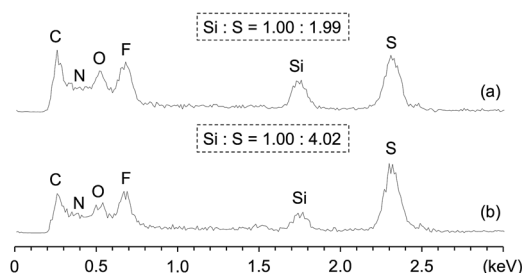


Fig. 2 EDX patterns of (a) Am-CyS-IL and (b) 2Am-CyS-IL.



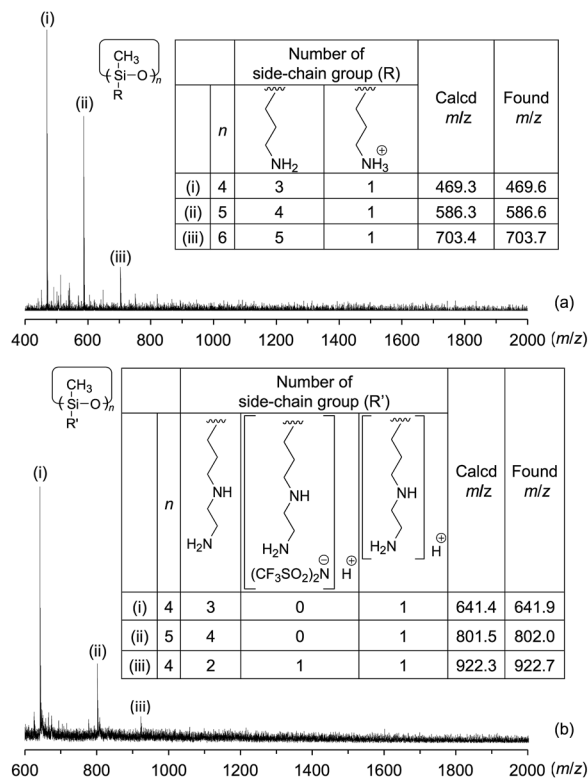


Fig. 3 MALDI-TOF MS of (a) Am-CyS-IL and (b) 2Am-CyS-IL.

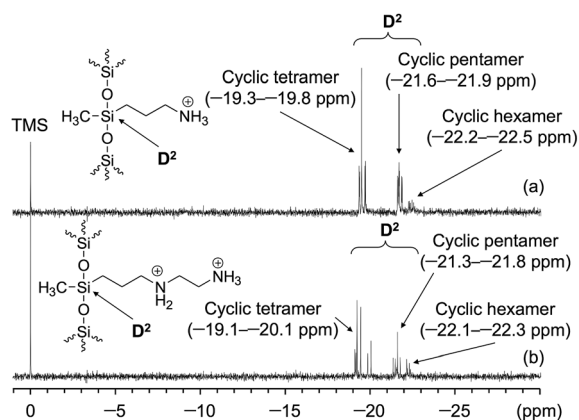


Fig. 4 ²⁹Si NMR spectra of (a) Am-CyS-IL and (b) 2Am-CyS-IL in DMSO-d₆ at 40 °C. Chemical shifts were referenced to tetramethylsilane (TMS; δ 0.0).

another organodialkoxysilane containing two amino groups, **AEAPDMMS**, as a starting material (Scheme 1b). **2Am-CyS-IL** was soluble in high-polarity solvents, such as water, methanol and acetone, but insoluble in low-polarity solvents, such as chloroform, toluene and hexane.

The ¹H NMR (Fig. 1b) and EDX (Fig. 2b) results of **2Am-CyS-IL** indicated that the **AEAPDMMS** reagent was not present in the product, and that the molar ratio of ammonium cations to NTF₂ anions in **2Am-CyS-IL** was ca. 1 : 1. The MALDI-TOF MS results indicated the existence of a mixture of cyclic siloxane tetramer

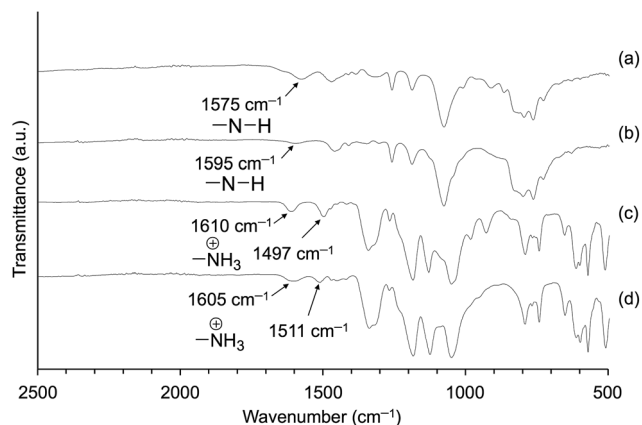


Fig. 5 ATR-IR spectra of (a) APDMMS, (b) AEAPDMMS, (c) Am-CyS-IL and (d) 2Am-CyS-IL.

and pentamer (Fig. 3b). Furthermore, the ²⁹Si NMR spectrum of **2Am-CyS-IL** showed three multiple signals in the D² region (-19.1 to -20.1 ppm for cyclic tetrasiloxane, -21.3 to -21.8 ppm for cyclic pentasiloxane, and -22.1 to -22.4 ppm for cyclic hexasiloxane) and no signals in the D¹ region at ca. -15 ppm (Fig. 4b), supporting the formation of cyclic siloxanes. In addition, these compounds had some stereoisomers, as demonstrated by the ¹H NMR spectrum with multiple signals, a, due to the methyl groups at 0.27 to -0.01 ppm (Fig. 1b) and the ²⁹Si NMR spectrum with three multiple signals, as described above (Fig. 4b).

To confirm the formation of ionic pairs of the resulting products (**Am-CyS-IL** and **2Am-CyS-IL**), we compared the ATR-IR spectra of these products and their starting materials (**APDMMS** and **AEAPDMMS**). The IR spectra of **APDMMS** and **AEAPDMMS** showed absorption peaks attributable to the bending vibration of amino groups at 1575 and 1595 cm⁻¹, respectively (Fig. 5a and b). Conversely, those of **Am-CyS-IL** and **2Am-CyS-IL** exhibited two typical absorption peaks attributable to the bending vibration of ammonium cations at 1610 and 1497 cm⁻¹ for **Am-CyS-IL** (Fig. 5c) and at 1605 and 1511 cm⁻¹ for **2Am-CyS-IL** (Fig. 5d). In addition, these spectra did not exhibit absorption peaks due to the bending vibration of amino groups. Therefore, we concluded that cationic species (ammonium cations) were included in **Am-CyS-IL** and **2Am-CyS-IL**. Meanwhile, because the NTF₂ anion is extremely stable, NTF₂ must be present as the anionic species in these compounds.

Thermal properties of Am-CyS-IL and 2Am-CyS-IL

The DSC analyses of **Am-CyS-IL** and **2Am-CyS-IL** were performed at a heating rate of 20 °C min⁻¹ under a nitrogen flow (50 mL min⁻¹). The baseline shifts assigned to T_g were observed at -2 °C for **Am-CyS-IL** (Fig. 6a) and at 9 °C for **2Am-CyS-IL** (Fig. 6b). Conversely, endothermic peaks due to T_m were not detected for both compounds (Fig. 6a and b), indicating that **Am-CyS-IL** and **2Am-CyS-IL** were amorphous compounds. In addition, sharp diffraction peaks were not observed in the XRD



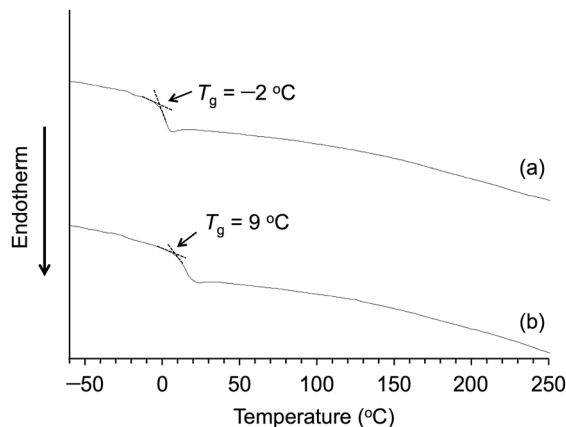


Fig. 6 DSC curves of (a) Am-CyS-IL and (b) 2Am-CyS-IL under a nitrogen flow.

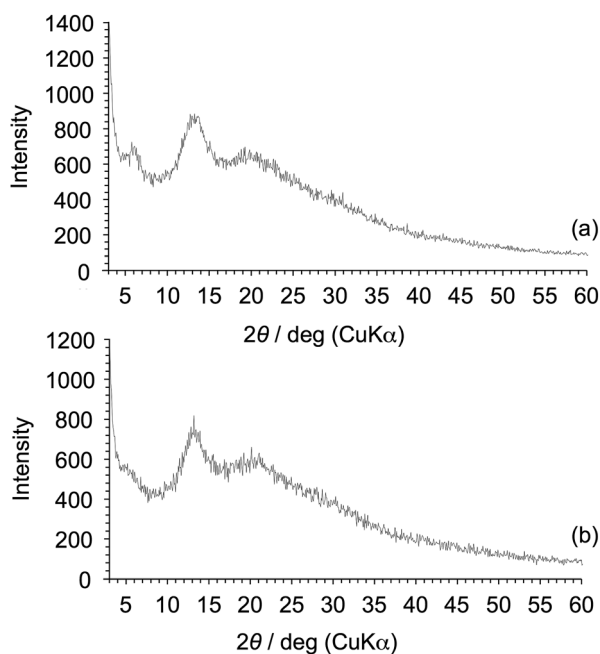


Fig. 7 XRD patterns of (a) Am-CyS-IL and (b) 2Am-CyS-IL. The amount of each product on the glass was ca. 1.0 mg cm^{-2} .

patterns of Am-CyS-IL and 2Am-CyS-IL (Fig. 7a and b), supporting their amorphous or poorly crystalline structures.

The flow temperatures of Am-CyS-IL and 2Am-CyS-IL were confirmed as follows. Samples in glass vessels were maintained in a horizontal position at $150 \text{ }^\circ\text{C}$ for 15 min, and then the vessels were cooled to room temperature, still in the horizontal position. Next, the vessels were maintained in a horizontal position at various temperatures (varied in $5 \text{ }^\circ\text{C}$ intervals) for 15 min, and then held for 15 min while tilting at each temperature. During this procedure, Am-CyS-IL and 2Am-CyS-IL showed obvious fluidity at $\sim 35 \text{ }^\circ\text{C}$ (Fig. 8a) and $\sim 45 \text{ }^\circ\text{C}$ (Fig. 8b), respectively. On the basis of these results, we concluded that Am-CyS-IL and 2Am-CyS-IL were PILs.

To investigate the amount of water in these PILs, the weight losses at $150 \text{ }^\circ\text{C}$ were evaluated by TGA (Fig. 9a and b).

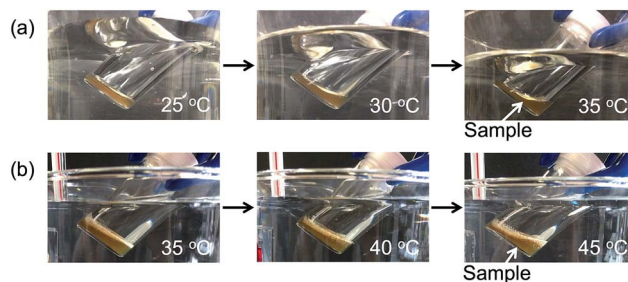


Fig. 8 Photographs of (a) Am-CyS-IL and (b) 2Am-CyS-IL at various temperatures.

Consequently, these values were 1.1% for Am-CyS-IL and 0.3% for 2Am-CyS-IL, indicating that these PILs contain a little moisture. However, because these samples were dried at $150 \text{ }^\circ\text{C}$ for 10 h after synthesis, it is presumed that they slightly absorbed moisture in the environment.

The thermal stabilities of Am-CyS-IL and 2Am-CyS-IL against pyrolysis were investigated by TGA at a heating rate of $10 \text{ }^\circ\text{C min}^{-1}$ up to $1000 \text{ }^\circ\text{C}$ under a nitrogen flow (250 mL min^{-1}). The T_{d5} and T_{d10} values were 351 and $362 \text{ }^\circ\text{C}$, respectively, for Am-CyS-IL (Fig. 9a). These values were higher than those of the propylamine-NTf₂ salt (281 and $295 \text{ }^\circ\text{C}$), having the side-chain structure of Am-CyS-IL, with a T_g of $-65 \text{ }^\circ\text{C}$ and a T_m of $28 \text{ }^\circ\text{C}$ (Fig. 9c). These results indicate that the thermal stability of Am-CyS-IL was enhanced by the incorporation of the cyclic oligosiloxane frameworks. The cyclic oligosiloxane frameworks could suppress molecular tumbling, resulting in the prevention of degradation.⁹ On the other hand, the T_{d5} and T_{d10} values of 2Am-CyS-IL were 294 and $309 \text{ }^\circ\text{C}$ (Fig. 9b), which are almost the same as those of the *N*-propylethylenediamine-NTf₂ salt (299 and $309 \text{ }^\circ\text{C}$), having the side-chain structure of 2Am-CyS-IL, with a T_g of $-29 \text{ }^\circ\text{C}$ and a T_m of $85 \text{ }^\circ\text{C}$ (Fig. 9d). Because the side-chain groups of 2Am-CyS-IL are long, the suppression of molecular tumbling would not work. The weights of the residues of Am-CyS-IL (ca. 12%; Fig. 9a) and 2Am-CyS-IL (ca. 10%; Fig. 9b) at $900 \text{ }^\circ\text{C}$ were almost the same as the theoretical SiO₂ yields (15% and 8%, respectively).

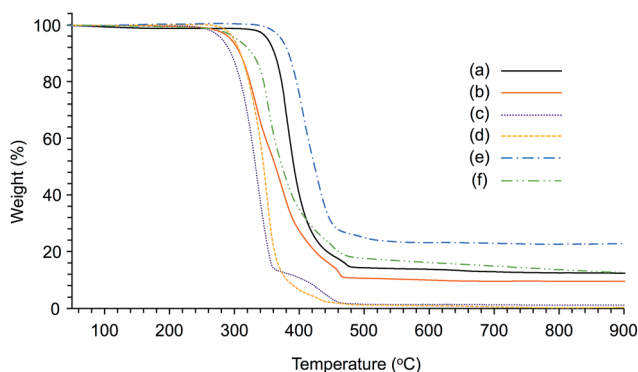


Fig. 9 TGA thermograms of (a) Am-CyS-IL, (b) 2Am-CyS-IL, (c) propylamine-NTf₂ salt, (d) *N*-propylethylenediamine-NTf₂ salt, (e) Am-POSS and (f) 2Am-POSS under a nitrogen flow.

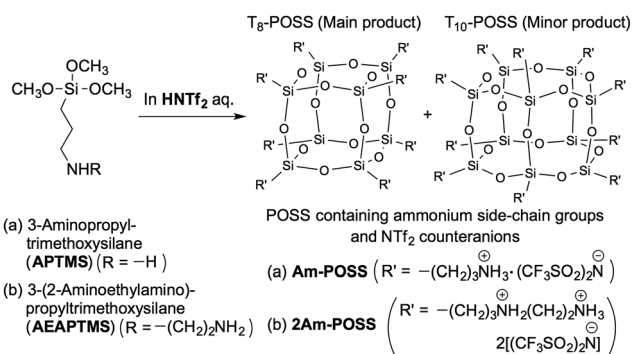


Influence of structures of siloxane frameworks on IL nature

To investigate the influence of the structures of the siloxane frameworks on the nature of the ILs, we prepared POSS compounds containing the same side-chain groups and counteranions (**Am-POSS** and **2Am-POSS**). In our previous studies, POSS compounds containing ammonium side-chain groups were prepared by the hydrolytic condensation of **APTMS**,¹⁷ **AEAPTMS**¹⁸ and an **APTMS/AEAPTMS** mixture¹⁸ using aqueous superacid HOTf. Therefore, to obtain POSS compounds containing the same counteranions as **Am-CyS-IL** and **2Am-CyS-IL**, the hydrolytic condensation of **APTMS** and **AEAPTMS** was performed using HNTf_2 as a catalyst, respectively (Scheme 2).

POSSs with one ammonium group (**Am-POSS**) and two ammonium groups (**2Am-POSS**) per repeating unit were prepared according to the following procedure (Scheme 2): **APTMS** and **AEAPTMS** were stirred in 0.5 mol L^{-1} aqueous HNTf_2 solutions at room temperature for 2 h, respectively. Here, excess amounts of HNTf_2 were required (feed molar ratio of $\text{HNTf}_2/\text{amino}$ group was 2.5). The subsequent procedures were the same as those described above for the preparation of **Am-CyS-IL** and **2Am-CyS-IL**.

Characterization of **Am-POSS** and **2Am-POSS** was performed using ^1H NMR (Fig. 10), EDX (Fig. 11) and ^{29}Si NMR (Fig. 12) measurements. Especially, the ^{29}Si NMR results



Scheme 2 Preparation of POSSs containing ammonium side-chain groups and NTf_2 counteranions: (a) **Am-POSS** and (b) **2Am-POSS**.

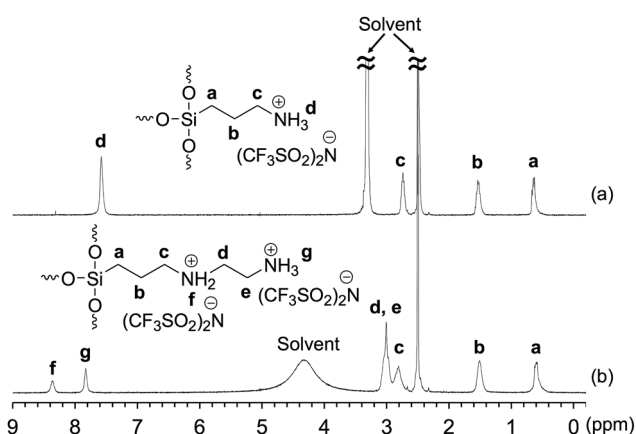


Fig. 10 ^1H NMR spectra of (a) **Am-POSS** and (b) **2Am-POSS** in $\text{DMSO}-d_6$. Chemical shifts were referenced to DMSO (δ 2.5).

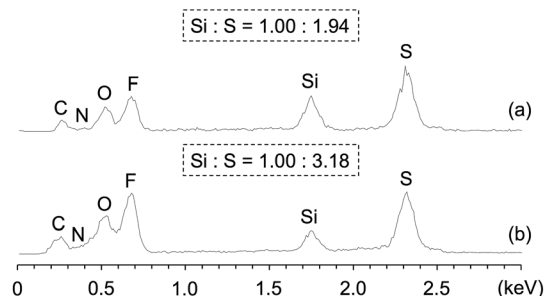


Fig. 11 EDX patterns of (a) **Am-POSS** and (b) **2Am-POSS**.

indicated that these compounds were mixtures of cage-like octasilsesquioxane (T_8 -POSS) and cage-like decasilsesquioxane (T_{10} -POSS) (Fig. 12). However, the molar ratio of ammonium cations to NTf_2 anions in **2Am-POSS** was *ca.* 1 : 0.79, *i.e.* a non-equimolar ratio (Fig. 11b). This is probably because it is difficult for HNTf_2 to approach the amino side-chain groups in **2Am-POSS** due to their bulky nature.

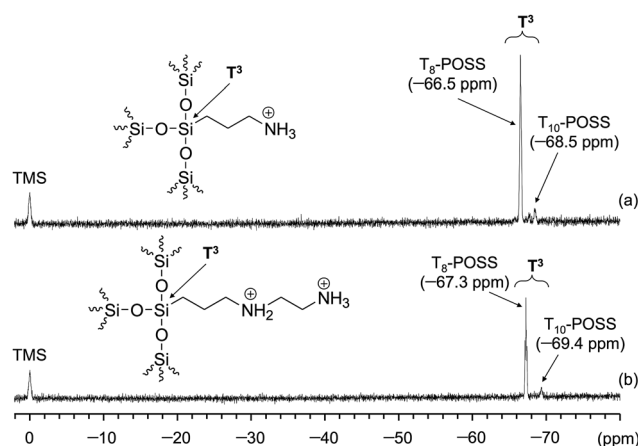


Fig. 12 ^{29}Si NMR spectra of (a) **Am-POSS** and (b) **2Am-POSS** in $\text{DMSO}-d_6$ at 40°C . Chemical shifts were referenced to TMS (δ 0.0).

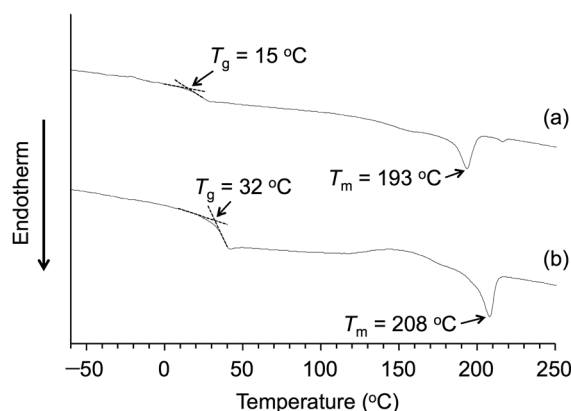


Fig. 13 DSC curves of (a) **Am-POSS** and (b) **2Am-POSS** under a nitrogen flow.



The DSC curve of **Am-POSS** showed a baseline shift due to T_g at 15 °C and an endothermic peak due to T_m at 193 °C (Fig. 13a). In addition, the XRD pattern of **Am-POSS** showed several sharp diffraction peaks (Fig. 14a), indicating that **Am-POSS** was a crystalline compound. Furthermore, this POSS compound did not exhibit fluidity at 150 °C (Fig. 15a), indicating that it was not an IL. On the other hand, the DSC curve of **2Am-POSS** exhibited a baseline shift due to T_g at 32 °C and an endothermic peak due to T_m at 208 °C (Fig. 13b). Additionally, the XRD pattern of **2Am-POSS** also supported the formation of a crystalline structure, *i.e.* several sharp diffraction peaks were detected (Fig. 14b). Furthermore, this compound did not show fluidity at 150 °C (Fig. 15b). Based on these results, we concluded that the cyclic oligosiloxane frameworks with flexible structures compared with POSS are important for the preparation of siloxane-based PILs exhibiting fluidity at low temperature.

Incidentally, **Am-POSS** and **2Am-POSS** exhibited high thermal stabilities in the TGA analyses. The T_{d5} and T_{d10} values

were 372 and 384 °C for **Am-POSS** (Fig. 9e) and 304 and 327 °C for **2Am-POSS** (Fig. 9f).

Conclusions

In this study, we found that PILs containing cyclic oligosiloxane frameworks, **Am-CyS-IL** and **2Am-CyS-IL**, could be successfully prepared by the hydrolytic condensation of **APDMMS** and **AEAPDMMS** using a water/methanol (1 : 19 v/v) mixed solution of HNTf_2 . The DSC curves of **Am-CyS-IL** and **2Am-CyS-IL** exhibited baseline shifts attributable to T_g at -2 °C and 9 °C, respectively, and endothermic peaks at T_m were not detected for both compounds, indicating their amorphous structures. In addition, fluidity was visually observed for each compound at ~35 °C and ~45 °C, respectively, indicating that **Am-CyS-IL** and **2Am-CyS-IL** were PILs. For comparison, POSSs containing the same side-chain groups and counteranions, **Am-POSS** and **2Am-POSS**, were prepared from **APTMS** and **AEAPTMS**, respectively, using an HNTf_2 catalyst. **Am-POSS** had T_g of 15 °C and T_m of 193 °C, and **2Am-POSS** had T_g of 32 °C and T_m of 208 °C. Furthermore, fluidity was not observed below 150 °C for either compound, indicating that these POSS compounds were not PILs. These results suggest that the cyclic oligosiloxane frameworks are an important factor for the preparation of siloxane-based PILs exhibiting fluidity at low temperature. The present PILs containing cyclic oligosiloxane frameworks exhibited relatively high thermal stabilities (**Am-CyS-IL**: T_{d5} = 351 °C and T_{d10} = 362 °C, **2Am-CyS-IL**: T_{d5} = 294 °C and T_{d10} = 309 °C).

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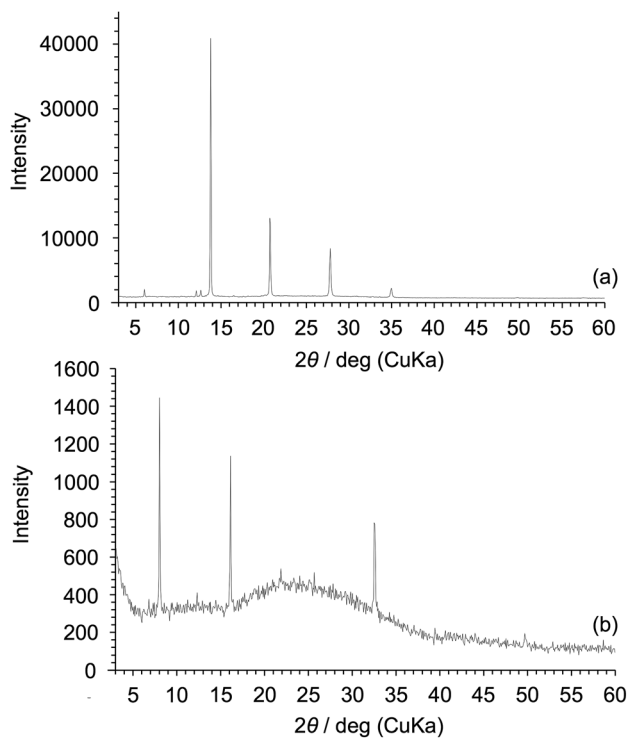


Fig. 14 XRD patterns of (a) **Am-POSS** and (b) **2Am-POSS**. The amount of each product on the glass was ca. 1.0 mg cm⁻².

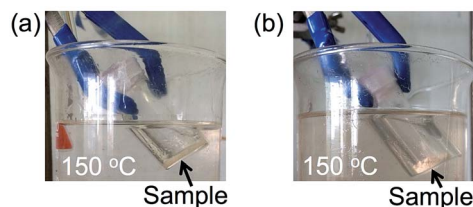


Fig. 15 Photographs of (a) **Am-POSS** and (b) **2Am-POSS** at 150 °C.



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