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Sulfur doped graphene as a promising metal-free electrocatalyst for oxygen reduction reaction: a DFT-D study†

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As an efficient metal-free catalyst, graphene doped with heteroatoms is highly active in promoting electrochemical oxygen reduction reaction (ORR). The detailed kinetic and thermodynamic behaviors of the entire ORR process on sulfur doped monovacancy graphene (SG_v), as well as the original mechanism are investigated by the dispersion-corrected density function theory (DFT-D) calculations. It is found that the SG_v is rather stable and the sulfur dopant is probably the active center. There are two proposed ORR pathways by kinetic process: the dissociation of OOH and the hydrogenation of OOH with the rate-determining steps of 0.75 eV and 0.62 eV, respectively. And the Gibbs free energy diagram of the entire ORR indicates that the dissociation of OOH is precluded, because the process of reduction step of O into OH is endothermic, while the hydrogenation of HOOH is the most favorable pathway even at high potential of 0.86 V. Our DFT-D simulation suggests that the SG_v would be an efficient electrocatalyst for ORR.

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1. Introduction

The energy-conversion efficiency of low-temperature fuel cells (FCs) is mainly limited by the slow kinetics of the oxygen reduction reaction (ORR) on cathode.^{1,2} The most effective cathodes of the commercial FCs is Pt and its alloys.³ Given the high cost, low abundance and poor durability of the Pt-based catalysts in FCs, non-precious metals cathodes or metal-free carbon-based cathodes have sparked worldwide interest in very recent years owing to their low cost, environmental friendliness, outstanding activity and stability.^{4–9} The cheap and stable graphene-based materials doped with heteroatoms,¹⁰ such as nitrogen,^{11,12} phosphorus,^{13,14} and their mixtures,^{15–17} present high catalytic activity and selectivity for ORR. The detailed kinetic behaviors and mechanisms of the entire ORR process on the modified graphene, such as phosphorus doped graphene,^{18,19} nitrogen doped graphene,^{20,21} boron doped graphene,²² MnN₄ embedded graphene,²³ and FeN₄ embedded

graphene,^{24,25} have been revealed successfully through density functional theory (DFT) calculations.

As one of the candidate for the metal-free cathode materials, the sulfur doped graphene (SG) is also reported to present high electrocatalytic activity for ORR.^{26–31} By replacing one C atom with sulfur, there are two kind of sulfur doped graphene, graphitic S and thiophene S: the thiophene S was found mainly at the edges of graphene bonding to C,²⁸ and the graphitic S was found inside the lattice structure of the graphene.^{27,32} For ORR mechanism on the thiophene S doped graphene, Zhang *et al.*³³ using several sulfur doped graphene clusters models demonstrated that the sulfur-doped graphene clusters with sulfur or sulfur oxide locating at graphene edges show electrocatalytic activity for ORR, and the zigzag edge or the neighboring carbon atoms of doped sulfur oxide atoms are the catalytic active sites. Summarily, from their reaction energy barrier calculation, Zhang *et al.*³³ predicted that the (thiophene) sulfur-doped graphene could show ORR catalytic properties comparable to Pt.

On the other hand, as other important sulfur doping configuration, the graphitic sulfur doped graphene presents the great stability from our simulations presented below. Moreover, to the best of our knowledge, the theoretical studies on the electrocatalytic activity of the graphitic sulfur doped graphene is still lack, and the detailed kinetic behaviors, the catalytic active center and the catalytic mechanism for the ORR on the graphitic sulfur doped graphene remains unclear. Thus, in the current study, we have studied the entire ORR mechanisms on the graphitic sulfur doped graphene (SG_v) using the first-principles method. The adsorption properties for all possible

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ORR involved species and the activation energies for all possible elementary reactions of the entire ORR on the SG_V in acid environment. Furthermore, the most favorable pathway for ORR on SG_V is identified based on the calculation of the transition states. Our current results not only show that the ORR on SG_V proceeds more possibly *via* a direct hydrogenation of HOOH reaction pathway but also demonstrate that the SG_V would be an efficient metal-free electrocatalyst for ORR.

2. Theoretical methods

All of the spin-polarized calculations were performed within dispersion-corrected density functional theory (DFT-D) computations as implemented in Dmol³ code embedded in Materials Studio (Accelrys, SanDiego, CA). The generalized gradient approximation (GGA) with Perdew–Burke–Ernzerhof (PBE) functional was employed using the DFT semi-core pseudopotential³⁴ with long-range dispersion correction *via* Grimme approach³⁵ to describe exchange and correlation effects,³⁶ and the polarization p-function (DNP) as the basis set for the double numerical atomic orbital augmented was chosen. The convergence tolerances of the geometry optimization are set to 10⁻⁵ Ha (1 Ha = 27.21 eV) for the energy, 0.002 Ha Å⁻¹ for the maximum force, and 0.005 Å for the maximum displacement. The electronic SCF tolerance is set to 10⁻⁶ Ha. In order to achieve accurate electronic convergence, we apply a smearing of 0.005 Ha to the orbital occupation. The *k*-points grid is set as 5 × 5 × 1 and the *k*-points are generated automatically using the Monkhorst–Pack method³⁷ for the relaxation calculations. And denser meshes of 15 × 15 × 1 are used to calculate the density of states (DOS) for the electronic relaxation. The transition states are obtained by linear synchronous transit (LST)/quadratic synchronous transit (QST) method in Dmol³ code for ORR elemental steps. It is found that all the minima possess real frequencies, and the transition states have only one imaginary frequency.

Gibbs free energy of the ORR intermediates can be calculated with the approach developed by Nørskov *et al.*³⁸ The change in free energy for the elemental step is defined as $\Delta G = \Delta E + \Delta ZPE + T\Delta S + \Delta G_U + \Delta G_{PH} + \Delta G_{field}$, where ΔE is the reaction energy based on DFT-D calculations, ΔZPE is the zero point energy, T is the temperature and equal to 300 K, ΔS is the change in the entropy. ZPE and S of the ORR intermediates are calculated based on the vibrational frequencies. ΔG_U are the free energy contributions due to variations in the electrode potential U . $\Delta G_U = -neU$, where n is the number of electrons transferred and U is the electrode potential *vs.* the standard hydrogen electrode (SHE). In our system, ΔG_{PH} isn't discussed. ΔG_{field} is normally ignored due to its small value.³⁸ We set the reference potential to be that of the standard hydrogen electrode. The free energy of 1/2H₂ can be used to instead of that of (H⁺ + e⁻). The free energy of H₂O was calculated in the gas phase at 300 K and the free energy of O₂ was obtained from the reaction O₂ + 2H₂ = 2H₂O for which a free energy change is 4.92 eV.³⁹

The adsorption energy (E_{ad}) is defined as $E_{ad} = E_{adsorbate} + E_{support} - E_{adsorbate/support}$, where $E_{adsorbate}$, $E_{support}$ and $E_{adsorbate/support}$ are the total energies of the free adsorbate, the

corresponding support and the support with the adsorbate, respectively. With this definition, a positive value indicates an exothermic adsorption. We use the same periodic box dimensions and the same level of calculations to obtain all the energies for $E_{adsorbate}$, $E_{support}$, and $E_{adsorbate/support}$.

3. Results and discussion

3.1. Properties of SG_V

The SG_V is modeled by replacing one of the lattice carbon atoms with a S atom from a 4 × 4 graphene supercell with the vacuum layer of 15 Å. According to our careful tests, the size of supercell is enough to study the ORR process. In the optimized SG_V (as shown in Fig. 1a), the S atom protrudes out of the graphene plane forming three S–C1 bonds of 1.47 Å. The C1–C2 bond neighboring to the S dopant is 1.40 Å (see Fig. 1a). The formation energy (E_f) of the SG_V given by $E_f = E_{total}(m, n) - E_{total}(Gra) - m\mu_S + n\mu_C$. In the formula, $E_{total}(S)$ is the total energy of the supercell with the defect complex, and $E_{total}(Gra)$ is the total energy of the pristine graphene. The symbols μ_S and μ_C are the chemical potentials of S and C and m and n is the number of S and the substituted C, respectively. μ_C is calculated as the carbon atom in the graphene. For the chemical potential of S, which is tunable in the experiments, we set the range from H₂S to the free atomic sulfur. The E_f of the S doped graphene is from 3.68 to 0.45 eV. The increased μ_S is corresponding to the decreased E_f , indicating that S is able to dope into graphene at high S chemical potential. The E_f of S doped divacancy graphene with μ_S of the free atomic sulfur is 1.71 eV (higher 1.26 eV than SG_V), because the formation of divacancy graphene is more difficult. Therefore, SG_V is easier to form by thermodynamic standpoint.

To understand the introduced changes of the SG_V electronic structure upon sulfur doping, the calculated DOS of the SG_V and the pure graphene are presented in Fig. 1b. Different from that of pure graphene, upon sulfur doping, some sharp peaks emerges at the Fermi level of the SG_V. According to the partial density of states (PDOS) analysis (presented in Fig. 1c), the emerged peaks are mainly from the sulfur dopant and the its neighboring C atoms. As suggested in previous works, those kind of states near the Fermi level would facilitate to the charge transfer between the substrate and the adsorbed species,^{40,41} and the sulfur doping induced electronic states emerging near the Fermi level should play an important role in the enhanced chemical activity of graphene.

In addition to the novel catalytic activity of SG_V with electronic structure analysis, on the other hand, the stability of SG_V is a precondition for ORR. Fig. 2a presents the deformation charge density (DCD) map for the SG_V, which confirms the formation of the stable covalent bond between sulfur atom and its neighboring C atoms, in line with the recent experimental results.²⁷ According to the Mulliken atomic charges analysis, the atomic S is negatively charged by 0.09 $|e|$, the atomic C1 is negatively charged by 0.06 $|e|$ and the atomic C2 is positively charged by 0.06 $|e|$, indicating a negligible charge transfer between the S and graphene sheet, due to the similar electro-negativity of S and C.



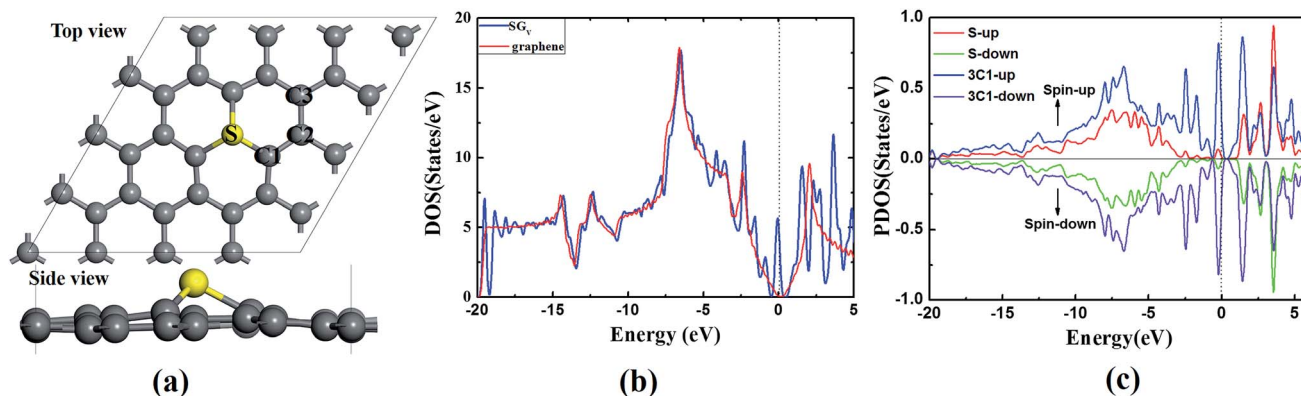


Fig. 1 (a) Top and side views of the most stable configurations of sulfur doped monovacancy graphene (SG_V), the carbon atoms near the sulfur atom are marked as "C1", "C2", and "C3", respectively. Hereafter, the gray and yellow spheres represent C and S atoms, respectively. (b) The density of states (DOS) for the SG_V and pure graphene. (c) The partial density of states (PDOS) for S and three "C1" atoms of SG_V. The Fermi level is indicated with a black dotted line.

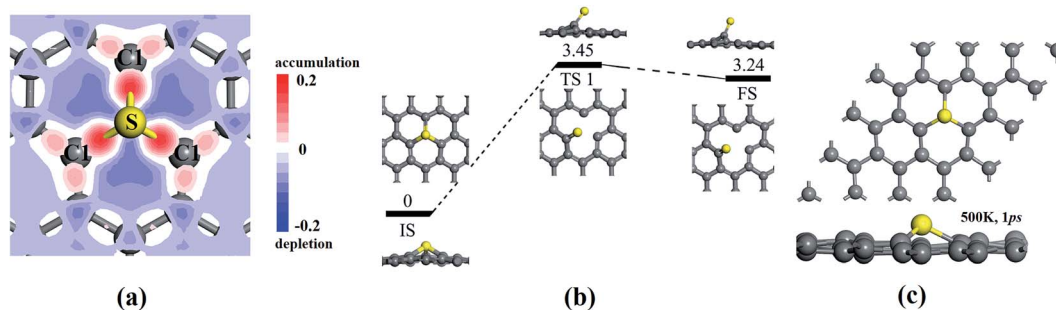


Fig. 2 (a) Deformation charge density (DCD) map in the SG_V plane, which displays the charge density overlap between the atomic sulfur and its neighboring atomic carbon. Blue, white and red represent charges depletion, non-transition and accumulation. (b) Optimized structures for the initial states (IS), the transition states (TS), and the final states (FS) (side view and top view) along the diffusion pathway of sulfur from the most stable configuration to the second most stable configuration. (c) The final SG_V structure from the molecular dynamics simulation at 500 K.

The stability of the SG_V model is further investigated from the mobility of the doping S and diffusion of the S dopant to its neighbor stable adsorption site. The various configurations of the S doped graphene or the adsorption of the S atom on the monovacancy-graphene support are presented in Fig. S1 in the ESI,[†] and it is turned out that SG_V presented above is the most stable configuration from the rather strong interaction between the S atom and the defect-graphene support. The S atom adsorbed on one of C atoms around the defect is the second most stable configuration (see Fig. S1[†] and 2b). The extremely big diffusion barrier (marked as "TS 1" in Fig. 2b) of 3.45 eV for the mobility of the S dopant from the most stable adsorption configuration (the selected SG_V model) to the second most stable adsorption configuration indicates the immobility of the atomic S anchored at the monovacancy-graphene and the great stability of the selected SG_V model. To further confirm the thermodynamical stable of SG_V, the first principles molecular dynamics calculations are performed in a period of 1000 fs at the temperature of 500 K. The final structure of SG_V from the molecular dynamics calculations is presented in Fig. 2c, and it is turned out that the C and S atoms are almost in the same plane just with slight distortion.

3.2. The adsorption of various ORR involved species on SG_V

The most stable configurations of the various ORR involved species on SG_V, including O₂, OOH, HOOH, O, H, OH, and H₂O, are presented in Fig. 3, and their corresponding adsorption properties are summarized in Table 1.

The adsorption of O₂ is a prerequisite of the ORR proceeding on the catalyst, thus we firstly investigate the adsorption characters of O₂. The most stable adsorption configuration of an O₂ molecule is on a carbon (marked as "C2") site with the E_{ad} of 0.25 eV, the O–O bond length of 1.24 Å and the O and C2 distance of 3.17 Å (see Fig. 3a). Given the weak interaction together with the negligible charge transfer between the O₂ molecule and the SG_V support, this adsorption configuration (presented in Fig. 3a) would be a physisorption configuration without the formation of the chemical bond between O₂ molecule and the SG_V support.

The co-adsorption of O₂ and H* (hereafter, "*" represents the reactant binding to the support) is an important intermediate product, where the O₂ (the O–O bond of 1.29 Å) and H* is respectively adsorbed on the hollow site and the C1 site forming a C–H bond of 1.13 Å (see Fig. 3b). The E_{ad} of O₂ and H* is 0.11 eV and 1.64 eV, respectively. The adsorbed O₂ is negatively charged by 0.32 |e|, and the H* is positively charged by 0.24 |e|.



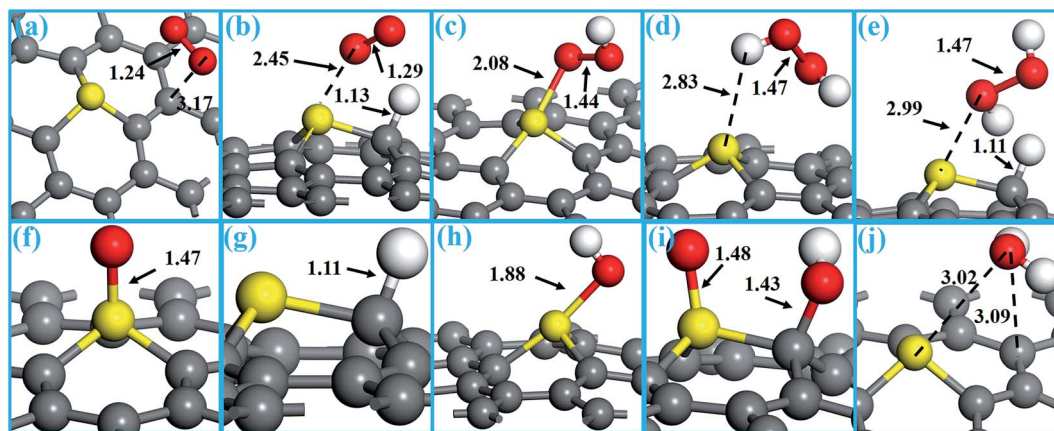


Fig. 3 The relaxed structures for the various ORR involved species adsorbed on SG_V : (a) O_2 , (b) $O_2 + H^*$, (c) OOH^* , (d) $HOOH$, (e) $HOOH + H^*$, (f) O^* , (g) H^* , (h) OH^* , (i) $O^* + OH^*$, (j) H_2O . Hereafter, red and white spheres represent O and H atoms, respectively.

It is found that the OOH species prefers to adsorb on sulfur dopant with the formation of the O-S bond of 2.08 Å (see Fig. 3c), and the O-O bond of the adsorbed OOH is 1.44 Å and the O-H bond is 0.98 Å. The E_{ad} of OOH^* is 0.66 eV and the OOH^* is negatively charged by 0.29 $|e|$. The $HOOH$ species weakly adsorbs on SG_V with the small E_{ad} of 0.34 eV and the negligible charge transfer (0.03 $|e|$), and the distance between H and S is 2.83 Å (see Fig. 3d).

As one of the intermediate products, the most stable co-adsorption configuration of the $HOOH$ and H^* is presented in Fig. 3e: the $HOOH$ stays above the sulfur dopant, and the H^* adsorbed on C1 forming a C-H bond of 1.11 Å. The E_{ad} of $HOOH$ and H^* is 0.49 eV and 1.94 eV, respectively. The adsorbed $HOOH$ is negatively charged by 0.03 $|e|$. The H^* is positively charged by 0.21 $|e|$.

As another ORR reactant, the atomic O is preferably adsorbed on S site with the rather big E_{ad} of 5.06 eV, and the length of the formed S-O bond is 1.47 Å (see Fig. 3d). The atomic O is

negatively charged by 0.35 $|e|$. The atomic H prefers to adsorb on C1 site (neighboring the S dopant) with an E_{ad} of 1.78 eV, and the formed C-H bond is 1.11 Å (see Fig. 3g). The atomic H is positively charged by 0.17 $|e|$, confirming that H^* could be taken as H^+ in the calculations.

As an important intermediate product, OH is preferably adsorbed on S site forming an S-O bond of 1.88 Å with an E_{ad} of 1.79 eV (see Fig. 3h). The OH^* is negatively charged by 0.27 $|e|$. The co-adsorption of O and OH is other important intermediate product, where the O and OH adsorb on the S site and the C1 site forming an S-O bond of 1.48 Å and a C-O bond of 1.43 Å, respectively (see Fig. 3i).

As the final product, H_2O is weakly adsorbed on the SG_V with the small adsorption energy of 0.35 eV (see Fig. 3j). The weak adsorption together with the long distance and the negligible charge transfer between H_2O and the SG_V support indicates that the formed H_2O would be easily released as the final product of ORR.

Table 1 The corresponding adsorption configurations, adsorption energies (E_{ad} in eV) and Mulliken charges (Δq in $|e|$) are summarized. Geometric and energetic parameters of reaction intermediates as identified in the stable state^a

Reaction intermediates	Configurations	Bond lengths (Å)	E_{ad}	Δq
O_2	Top-C2	d(O-O) = 1.24	0.25	-0.21
$O_2 + H^*$	Hollow	d(O-O) = 1.29	0.11	-0.32
		d(H-C) = 1.13	1.64	0.24
OOH^*	Hollow	d(O-O) = 1.44	0.66	-0.29
$HOOH$	Hollow	d(O-O) = 1.47	0.34	0.03
$HOOH + H^*$	Hollow	d(O-O) = 1.47	0.49	0.03
		d(H-C) = 1.11	1.94	0.21
O^*	Top-S	d(O-S) = 1.47	5.06	-0.35
H^*	Top-C1	d(H-C) = 1.11	1.78	0.17
OH^*	Top-S	d(O-S) = 1.88	1.79	-0.27
$O^* + OH^*$	Top-S	d(O-S) = 1.48	4.49	-0.35
		d(O-C) = 1.43	1.22	-0.09
H_2O	Hollow	d(O-H) = 0.97	0.35	0.04

^a *represented that the reactant binds to the SG_V surface.

3.3. Chemical behavior in various reaction pathways

As mentioned above, it is found that the rather stronger adsorption of H than the O_2 (1.78 eV vs. 0.25 eV), and the adsorption of H is greatly preferable than O_2 . Thus, the direct dissociation pathway of the adsorbed O_2 on SG_V support would not be further considered in here, and the H pre-adsorbed configuration would be focused. As shown in Fig. 4, it is found that the weakly adsorbed O_2 is easily hydrogenated by the pre-adsorbed H into OOH^* on SG_V : firstly, the $O_2 + H^*$ co-adsorbed configuration would form the metastable state (MS) of OOH^* with a reaction barrier of 0.39 eV and an exothermic reaction energy of 0.78 eV. Then the metastable adsorbed OOH^* convert into the most stable adsorbed OOH^* with a small reaction barrier of 0.18 eV and a slight exothermic reaction energy of 0.02 eV. The most stable adsorbed OOH^* is dissociated into the metastable co-adsorbed O + OH species *via* the activation barrier of 0.44 eV and an exothermic reaction energy of 1.17 eV. Finally, the formed OH^* would diffuse from S site to C1 site *via* the diffusion barrier of 0.75 eV and an exothermic



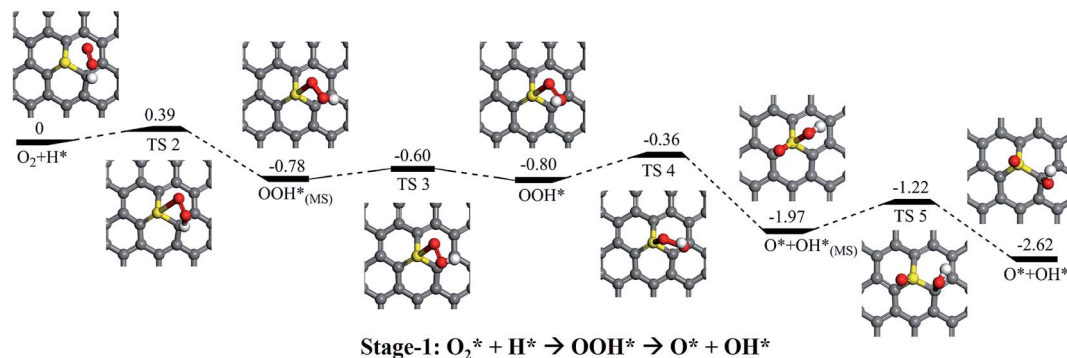


Fig. 4 Schematic potential energy profiles for the formation of OOH^* and the direct dissociation of OOH^* on SG_v (Stage-1). The calculated transition states (TS) are denoted as "TS 2", "TS 3", "TS 4" and "TS 5", respectively and the metastable state is denoted as MS.

reaction energy of 0.65 eV. The above process is marked as Stage-1.

In addition to the direct OOH^* dissociation pathway, the hydrogenation of OOH^* species is also considered, and the corresponding process is noted as "Stage-2" in Fig. 5. In Stage-2, upon the hydrogenation of the OOH^* species by the co-adsorbed H^* , the $HOOH$ is formed *via* a reaction barrier of 0.53 eV and an exothermic reaction energy of 1.23 eV. We find that the formed $HOOH$ is difficultly dissociated into two OH species due to the weak adsorption of $HOOH$ on SG_v . "Stage-3" in Fig. 5 is presented the hydrogenation of the $HOOH$ by the co-adsorbed H^* : the co-adsorbed $OH^* + H_2O$ is formed *via* a reaction barrier of 0.36 eV and an exothermic reaction energy of 2.70 eV. As mentioned above, the formed H_2O would easily release due to the weak interaction between H_2O and SG_v .

Following the formation of the $O^* + OH^*$ species from the dissociation of OOH^* , there are two possible pathways for their further hydrogenation: the O^* hydrogenation or the OH^* hydrogenation. On the one hand, we found that the hydrogenation of the OH would result in the formation of H_2O *via* a reaction barrier of 0.68 eV and an exothermic reaction energy of 3.28 eV, respectively (see Stage-4 in Fig. 6). As the final product of the ORR process, the H_2O molecule would be easily released due to the week interaction between H_2O and SG_v (with an adsorbed atomic O), which has a rather small E_{ad} of 0.39 eV.

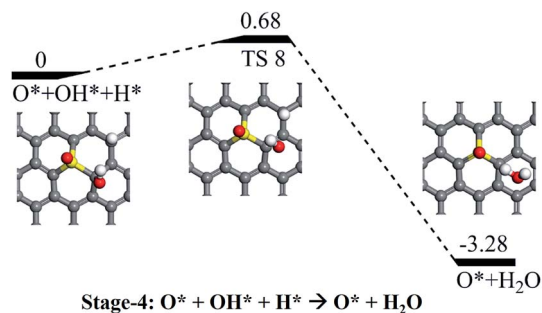


Fig. 6 Schematic potential energy profiles for the formation of the H_2O from the hydrogenation of the OH (Stage-4). The corresponding TS is denoted as "TS 8".

On the other hand, the hydrogenation of the O^* of the co-adsorbed $O^* + OH^*$ is difficult because of the rather high reaction barrier (2.77 eV), and the pathway is not further considered.

Following the release of the (first) formed H_2O molecule, the left atomic O staying on sulfur will be hydrogenated into the OH species *via* a reaction barrier of 0.40 eV and an exothermic reaction energy of 0.19 eV (see Stage-5 in Fig. 7). Again, the formed OH would be further hydrogenated into H_2O *via* the reaction barrier of 0.62 eV and an exothermic reaction energy of

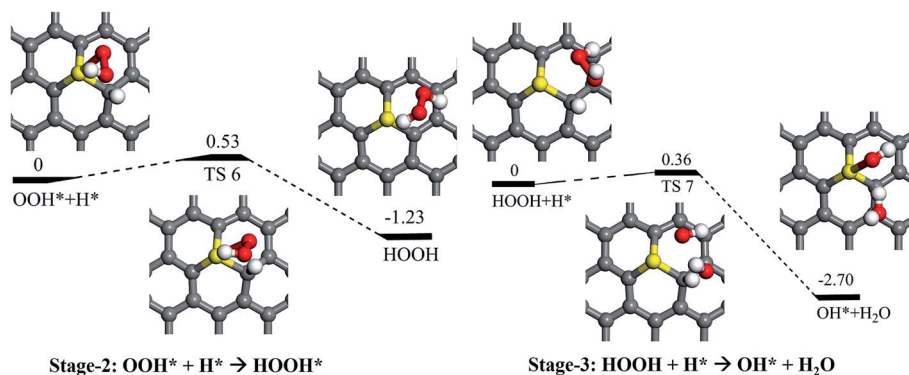


Fig. 5 Schematic potential energy profiles for the hydrogenation of OOH^* (Stage-2), and the hydrogenation of $HOOH$ (Stage-3). The corresponding TS are denoted as "TS 6" in Stage-2 and "TS 7" in Stage-3, respectively.



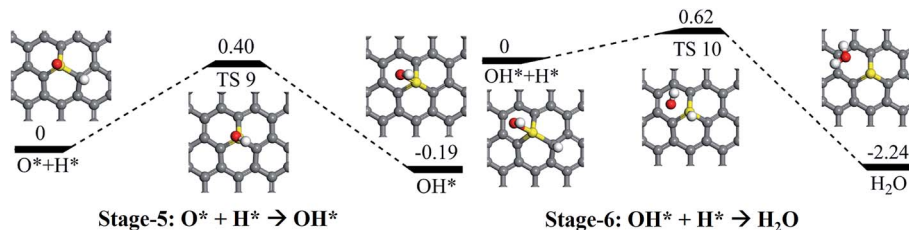


Fig. 7 Schematic potential energy profiles for the formation of the OH species from the hydrogenation of the atomic O (Stage-5) and the (second) H_2O formation from the hydrogenation of the OH (Stage-6). The corresponding TS are denoted as "TS 9" in Stage-5 and "TS 10" in Stage-6, respectively.

2.24 eV (see Stage-6 in Fig. 7). The formed H_2O would easily release due to the weak interaction between water and the support, resulting in the recovery of SG_V .

3.4. The selectivity of two pathways with kinetics and thermodynamics

According to the reaction stages presented above, two possible pathways are proposed for the entire ORR process on SG_V , which are summarized in Fig. 8a.

Pathway I: Stage-1 (to form $O + OH$) \rightarrow Stage-4 \rightarrow Stage-5 \rightarrow Stage-6

Pathway II: Stage-1 (to form OOH) \rightarrow Stage-2 \rightarrow Stage-3 \rightarrow Stage-6

In pathway I (black lines in Fig. 8), the pathway starts from the direct hydrogenation of the adsorbed O_2 , resulting in the formation of the OOH^* species. The form OOH^* is dissociated into $O^* + OH^*$. Upon a diffusion process, the formed O^* and

OH^* would be further hydrogenated into H_2O . The rate-determining step of pathway I is the diffusion of OH^* and the corresponding reaction barrier is 0.75 eV.

In pathway II (red lines in Fig. 8), the pathway starts from the co-adsorption of $O_2 + H^*$. The O_2 will proceed with three sequential hydrogenation reactions to form OOH^* , $HOOH$ and the final ORR product of released H_2O molecule, whereas the left OH^* is chemisorbed on sulfur. Finally, the OH^* is hydrogenated into the second H_2O . The rate-determining step of this pathway is the formation of the second H_2O with a reaction barrier of 0.62 eV.

According to the computed free energy diagram, all of the reduction steps of pathway I and pathway II are presented in Fig. 8b and c, respectively. In pathway I, all of the reduction steps except the O reduction are downhill at 0 potential, as shown in Fig. 8b. The reduction step of O into OH has a positive ΔG of 0.86 V at 0 potential, which is the thermodynamic rate-determining step. When the thermodynamic process is

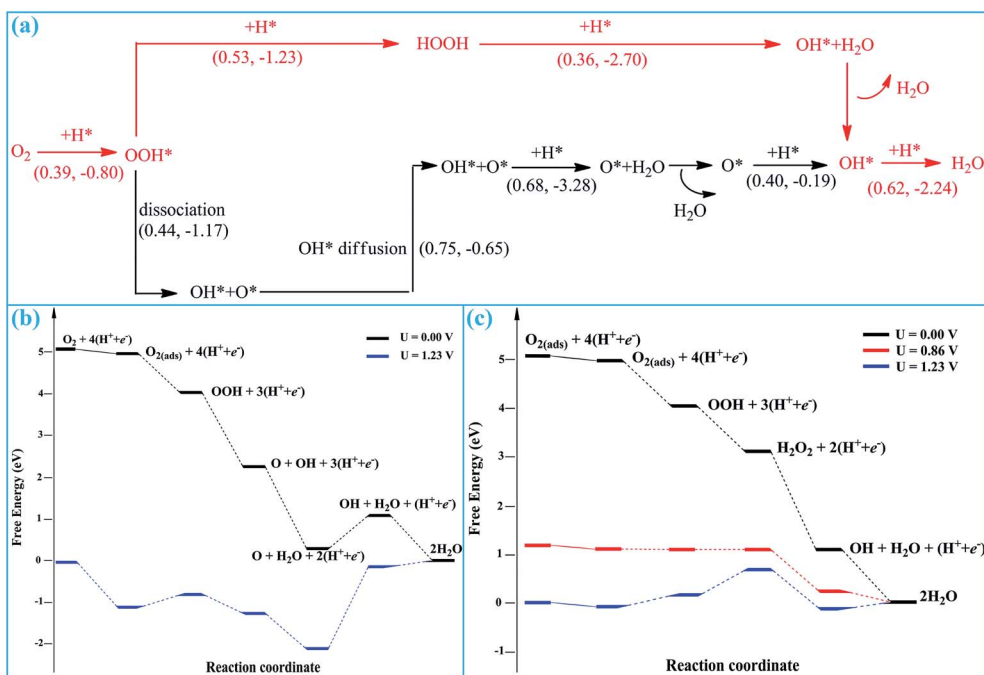


Fig. 8 (a) Proposed pathways for the entire ORR process on SG_V . The most favorable reaction pathway is expressed as red lines. The activation and reaction energies (in eV) are given in parentheses in the form of "(TS, ΔE)", accordingly. Free-energy diagram for ORR along (b) pathway I and (c) pathway II in on the SG_V , the subscript (ads) represents the adsorption on the SG_V .



continuously exothermic, it can be favorably applicable to ORR. Therefore, pathway I is precluded by the Gibbs free energy calculation, because the process of reduction step of O into OH is endothermic at 0 potential. However, in pathway II, all of the reduction steps are downhill at 0 potential, indicating that the entire ORR process successfully occurs by thermodynamic calculations. And even the ORR process occurs at the high potential of 0.86 V. Therefore, the pathway II is the most possible for ORR on SG_v.

4. Conclusions

DFT-D calculation is performed to investigate the detailed kinetic and thermodynamic behaviors of the entire ORR process on SG_v. It is found that the SG_v is rather stable due to the strong interaction between the sulfur and the defected-graphene. The O₂ molecule prefers to form OOH* species on SG_v with the co-adsorbed H*. There are two proposed pathways for the entire ORR process on SG_v, *i.e.*, the OOH* direct dissociation pathway and the OOH* hydrogenation pathway. The reaction barriers of the corresponding rate-determining steps of the two proposed pathways are 0.75 eV and 0.62 eV, respectively. According to the Gibbs free energy simulation at various potentials, the process of the OOH dissociation pathway is precluded, because the process of reduction step of O into OH is endothermic. While the process of the hydrogenation of HOOH species is effective at 0 potential, and it can occur at a high potential of 0.86 V. It is expected that the SG_v would be an efficient metal-free electro-catalyst for ORR.

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