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Modulating Cu⁺ distribution on the surface of Ce-doped CuO composite oxides for SO₂-resistant NH₃-selective catalytic reduction of NO

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A series of (CeO_x)_mCuO composite oxides with different Ce/Cu molar ratios (*m*) were prepared by co-precipitation for ammonia-selective catalytic reduction of NO (NH₃-SCR-NO) in the absence and presence of SO₂. Through X-ray diffraction (XRD), X-ray photoelectron spectroscopy (XPS), H₂-temperature programmed reduction (H₂-TPR), and NH₃-temperature programmed desorption (NH₃-TPD) characterization, it was demonstrated that the prepared samples were featured by distinguished Cu⁺ distributions on the surface of the Ce-doped CuO composite structure, indicating a special interaction or synergistic effect present in the samples. The modulated distributions of Cu⁺ achieved by adjusting the Ce/Cu molar ratio were found to be crucial for optimizing the redox ability and acid feature among the samples. The outstanding SO₂-resistant NH₃-SCR-NO performance (NO conversion close to 100% in temperature window of 250–350 °C under 200 ppm SO₂-presence) was observed on (CeO_x)_{0.25}CuO which possessing the highest Cu⁺ distribution, indicating the significant dependence of Cu⁺ distribution in Cu-based catalysts on enhancing catalytic performance for NH₃-SCR-NO. These results also suggested that the distinguished synergistic effect among binary component Cu-based catalysts should be considered as an important factor to govern the performance of NH₃-SCR-NO_x, which could be inferred as important clues to innovate more efficient catalysts for NH₃-SCR-NO_x, as well as other heterogeneous catalysis applications based on Cu-catalysts.

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1. Introduction

Nitric oxides (NO_x) are a type of major gaseous pollutant that cause many harmful effects on both the environment and human health and are inevitably produced from fossil fuel combustion processes of power plants and vehicles.¹ For abatement of their emissions to the atmosphere, in view of current technique, the most efficient way is converting them to harmless N₂ and H₂O at the beginning of emissions by NH₃ through selective catalytic reduction (NH₃-SCR-NO_x) process over heterogeneous catalysts.^{2,3} V₂O₅-WO₃-TiO₂ has been used as a commercial catalyst for NH₃-SCR-NO_x, but it has drawbacks such as a narrow working temperature window, use of toxic vanadium species, and low N₂ selectivity at high temperatures,^{4–6} which motivated a new wave of innovation towards more efficient and environmentally friendly replacement catalysts for NH₃-SCR-NO_x. Among various attempts, copper-based catalysts supported on oxides or zeolites have recently emerged as attractive candidates due to their environmentally benign characteristics and high SCR performance in a wide temperature range.^{7–10} Co, Fe, Mn, and Ce have been doped into

the phase or co-deposited with Cu on supports, such as zeolites, TiO₂, and SiO₂, to fabricate composite Cu-based NH₃-SCR catalysts. Addition of these elements was found to be beneficial for creating charge imbalance, vacancies or unsaturated chemical bonds on the catalysts surface, and practical contribution to the catalytic improvements for NH₃-SCR-NO_x.^{11–15} These positive influences originated *via* an additive introducing method were profoundly defined as synergistic effect among different components. However, essential knowledge of this effect among current Cu-based NH₃-SCR-NO_x catalysts is still controversial.

In general, the so-called synergistic effect was always discussed as an interaction between the two components of the composite samples. Since most of the present Cu-based NH₃-SCR-NO_x catalysts can be roughly sorted as ternary-component composites composed of Cu species, assistant additives, and a supporter, there were at least three different interactions, Cu-supporter, Cu-additive, and additive-support, those could be involved in the composite structures, which led to inconsistent views on the origin and function of the synergistic effect. A popular view was mainly focused on the Cu-additive interaction, and additive-induced geometric modifications or electronic interactions were believed to be responsible for the observed enhancement in the catalytic performance. For instance, Dou

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*et al.*⁷ demonstrated that Ce doping improved the redox properties of Cu–Ce/ZSM-5 catalyst by enhancing the composition of Cu species in high valence and mobility of lattice oxygen compared to those of the Cu/ZSM-5 catalyst. Similar descriptions were also found on supported Cu–Fe or Cu–Mn catalysts.^{11,16,17} However, the interactions between active component and supporter were always ignored in these studies, and the situation was quite doubtful for current Cu-based NH₃-SCR-NO_x catalysts since this interaction was usually present as a significant factor to resolve the catalytic behaviors of supported metal or metal oxides catalysts.^{18–20} This is especially true in view of the low loadings of both Cu species and additives on supporters. In contrast, it is not a surprise that some more complicated synergistic effects such as the dual redox cycles among Cu–Ce–Ti triplet complex on Cu–Ce/TiO₂ catalysts were supposed to be the key factor for enhancing the performance of NH₃-SCR-NO.²¹ To clarify these arguments, a basic work could be decreasing the complexity of the catalysts such as *via* employing binary component samples as catalysts. To the best of our knowledge, these works, especially in directly constructing and systematically modulating the composite structure of binary component Cu-based NH₃-SCR-NO_x catalysts, have been seldom reported.²² In addition, one of the main drawbacks of current Cu-based NH₃-SCR-NO_x catalysts is their poor resistance to SO₂,^{23,24} and to date, research has been focused on simple composition catalysts, whereas binary component samples for SO₂-resistant NH₃-SCR-NO_x are still scarce.

In this study, a series of Cu–Ce composite oxides with systematic change in Cu/Ce molar ratio was prepared by a common co-precipitation method and was aimed at NH₃-SCR of NO in the absence and presence of SO₂. Compared to the absolute existence of Cu²⁺ on Cu–Ce-supporter ternary-component samples reported in literatures, the distinguished Cu⁺ distributions in the surface of Ce-doped CuO composite structure suggested that a kind of special interaction or synergistic effect existed on our samples. By properly modulating the distribution of Cu⁺ in these Cu–Ce binary composite oxides with a controlled Cu/Ce molar ratio and thermal processing temperature, superior SO₂-resistant activity for NH₃-SCR of NO was achieved. These results suggested that the distinguished synergistic effect among binary component Cu-based catalysts compared to popular ternary-component samples should be noticed as a type of significant factor to govern the performance for NH₃-SCR-NO_x, which could be inferred as important clues to innovate more efficient catalysts for NH₃-SCR-NO_x, as well as other heterogeneous catalysis applications upon Cu catalysts.

2. Experimental

2.1. Preparation of samples

Cu–Ce composite oxides with different Cu/Ce molar ratios were prepared *via* the co-precipitation of Cu(CH₃COO)₂ and Ce(NO₃)₃ with Na₂CO₃ as a precipitant. Simply, desired amounts of Cu(CH₃COO)₂·H₂O and Ce(NO₃)₃·6H₂O were dissolved in deionized water and then 0.1 M Na₂CO₃ aqueous

solution was added to the abovementioned mixture under vigorous stirring to adjust the pH of the system to 9. The solution was stirred at room temperature for 1 h and subsequently the solid powders were filtered out and washed with deionized water several times. After being dried at 110 °C in air for 5 h, the powders were calcined in a muffle furnace at 350 °C for 3 h to obtain catalyst samples. The compositions of all the samples were measured by the ICP-OES technique, and according to the measurements and the calcination temperature, the three Cu–Ce composite oxides obtained were labelled as (CeO_x)_{0.125}CuO-350, (CeO_x)_{0.25}CuO-350, and (CeO_x)_{0.50}CuO-350 in sequence. For reference, a pure CuO-350 sample was synthesized according to a similar procedure except for the absence of Ce(NO₃)₃. The calcination temperature at 350 °C was chosen based on the following considerations: (a) according to our experience, the specific surface area or dispersion of Cu-based composite oxide would be significantly lowered at the calcination temperature beyond 400 °C, and 350 °C was sufficient to obtain phase-stabilized Cu-based composite oxides; (b) current advanced Cu-based catalysts could achieve valuable high conversion for NH₃-SCR-NO_x below 350 °C;^{19,25} therefore, this temperature was pre-set as a working criterion for estimating the samples prepared in this study. However, according to the estimation of a reviewer on possible high temperature for using, the dried powders of (CeO_x)_{0.25}CuO were also calcined at 400 °C and 450 °C for further investigation, and they were designated as (CeO_x)_{0.25}-CuO-400 and (CeO_x)_{0.25}CuO-450.

2.2. Characterization

The actual compositions of the samples were quantified *via* inductively coupled plasma-optical emission spectroscopy (ICP-OES) using an Agilent Technologies 5100 ICP-OES.

X-ray powder diffraction experiments (XRD) were carried out *via* a Persee XD-3 X-ray diffractometer at the scan rate of 1° min^{−1} in the angle range of 20–80° using Cu Kα radiation.

Chemical states of the sample surface were investigated by X-ray photoelectron spectroscopy (XPS) using an Axis Ultra DLD electron spectrometer. The spectra were obtained at ambient temperature under ultra-high vacuum. During data processing of the XPS spectra, the binding energy was calibrated with reference to the C 1s peak of the contaminant carbon at 284.5 eV.

H₂-temperature programmed reduction (H₂-TPR) experiments were performed using a quartz U-type reactor connected to a thermal conductivity detector (TCD) with an Ar–H₂ mixture (10% of H₂ by volume, 30 ml min^{−1}) as the reductant. Prior to the reduction, the sample (50 mg) was pre-treated in a high purity N₂ stream at 300 °C for 1 h and then cooled down to room temperature. TPR measurements were carried out from 50 °C to the target temperature at the rate of 10 °C min^{−1}.

NH₃-temperature-programmed-desorption experiments (NH₃-TPD) were performed *via* a Micromeritics AutoChem II 2920 apparatus. Before analysis, about 50 mg of the sample was pre-treated at 120 °C in a helium stream for 1 h and then cooled



down to room temperature. After absorption of NH_3 (at room temperature) by a pulsed flow of 0.5% NH_3 in helium at the flow rate of 30 ml min^{-1} for 1 h, the system was purged with helium flow at 30 ml min^{-1} for 1 h to remove physically adsorbed NH_3 and then heated from room temperature to 800°C at the heating rate of $10^\circ\text{C min}^{-1}$ to monitor the desorption signals.

2.3. Catalytic tests

The catalytic activities of the prepared catalysts for NH_3 -SCR in excess oxygen were investigated using a fixed-bed quartz reactor and 180 mg of the catalyst (20–30 mesh) was used. The reactant gas was composed of 1000 ppm NO, 1000 ppm NH_3 , 5 vol% O_2 , 200 ppm SO_2 (when used), and balanced N_2 . The gas hourly space velocity (GHSV) was $40\,000 \text{ h}^{-1}$. The concentrations of NO and NO_2 remaining in the product were analysed by an on-line chemiluminescence-based NO- NO_x analyzer (Model 42i-HL, Thermo Scientific). For the concentration of the gases at steady state, the NO conversion was calculated as follows:

$$\text{NO conversion} = (1 - [\text{NO}]_{\text{outlet}}/[\text{NO}]_{\text{inlet}}) \times 100\%$$

3. Results and discussion

The XRD patterns of $(\text{CeO}_x)_m\text{CuO}$ (with Ce/Cu molar ratio m at 0.125, 0.25, and 0.5) and CuO samples are presented in Fig. 1. The diffraction peaks (Fig. 1a) observed at the 2θ values of 35.5° , 38.6° , 48.8° , 52.5° , 58.4° , and 61.6° for pure CuO-350 reference sample could be assigned to the crystalline features of monoclinic CuO (JCPDS 45-0937). In contrast, for the $(\text{CeO}_x)_m\text{CuO}$ -350 samples, with the increasing Ce composition (Fig. 1b–d), obvious and gradual decreases in the intensity for the

diffraction peaks of CuO phase were observed. In addition, there were no notable signals correlated to CeO_x diffractions that could be resolved with these samples; however, clear shifts up to 2θ of 35.7° and 38.9° for two main diffraction peaks compared to those of the pure CuO sample were observed in the $(\text{CeO}_x)_m\text{CuO}$ samples. Similar observations have also been reported on Cu–Ce/ZSM catalysts, in which the phenomenon was attributed to the effect of Ce doping into the crystalline of copper oxide.²³ Thus, the absence of CeO_x diffraction peaks and diffraction shifts compared to the XRD pattern of pure CuO on prepared $(\text{CeO}_x)_m\text{CuO}$ samples clearly indicated the doping of Ce species into CuO crystalline for these samples, which could be a hint of the presence of interaction between the components of our samples. According to the estimation of a reviewer, $(\text{CeO}_x)_{0.25}\text{CuO}$ -400 and $(\text{CeO}_x)_{0.25}\text{CuO}$ -450 (calcined at 400°C or 450°C) were also tested and the results are shown in Fig. 1e and f. Similar diffraction patterns with close peak positions and comparable intensities were found for the three $(\text{CeO}_x)_{0.25}\text{CuO}$ samples, which manifested phase stability of our samples under the thermal condition of 350 – 450°C .

Surface features of the samples were investigated by XPS characterization, as shown in Fig. 2. In contrast to pure CuO, the spectral feature at 880 – 920 eV observed for the $(\text{CeO}_x)_m\text{CuO}$ samples, as shown in Fig. 2A, corresponds to Ce 3d. These Ce 3d spectra were fitted into ten binding energy (BE) peaks, labelled as v-kind, originated from $3d_{5/2}$ photoemissions, and the associated $3d_{3/2}$ emissions are labelled as u-kind. The BE peaks marked as v (882.5 eV), v' (888.9 eV), and v'' (898.3 eV) originated from Ce(IV) $3d_{5/2}$, whereas the BE peaks labelled as u (901.2 eV), u' (907.5 eV), and u'' (916.8 eV) originated from Ce(IV) $3d_{3/2}$. The v' (881.5 eV)–v'' (885.1 eV) and u' (898.8 eV)–u'' (903.4 eV), refer to $3d_{5/2}$ and $3d_{3/2}$ of Ce(III), respectively, suggesting the co-presence of Ce(III) and Ce(IV) on $(\text{CeO}_x)_m\text{CuO}$ samples.^{21,26,27} For a CuO reference sample, BEs of 933.7 eV and 953.7 eV (Fig. 2B) for the two main peaks are close to the standard values of Cu $2p_{3/2}$ and Cu $2p_{1/2}$ of CuO, respectively,^{28,29} which were accompanied by two shakeup satellite peaks in the range of 938.2 – 946.3 eV and a isolated satellite peak around 962.1 eV . $(\text{CeO}_x)_m\text{CuO}$ samples showed similar peaks, except for the obvious BE shifts of peaks towards lower BE compared to CuO. With careful deconvolution measurements, besides Cu $2p_{3/2}$ and Cu $2p_{1/2}$ peaks of CuO, additional peaks located around 933.4 eV and 953.3 eV were also resolved for the $(\text{CeO}_x)_m\text{CuO}$ samples. These peaks, with lower BE compared to those of CuO, were close to the BE value of Cu^+ in the composite oxides reported in the literature,^{26,27} suggesting the existence of Cu^+ on the $(\text{CeO}_x)_m\text{CuO}$ samples. Among the species on the surface, Ce(III) and Cu^+ could be more attractive since they are believed to be helpful for creating charge imbalance or promoting active species mobility to facilitate the reaction;^{21,29} thus, the contents or distributions of these two species were worthy of further investigation.

Based on the abovementioned peak-fitting measurements, a literature method was applied to estimate the relative concentrations (C) of Ce(III) by the calculated peak area of Ce species according to the following equations:^{30–33}

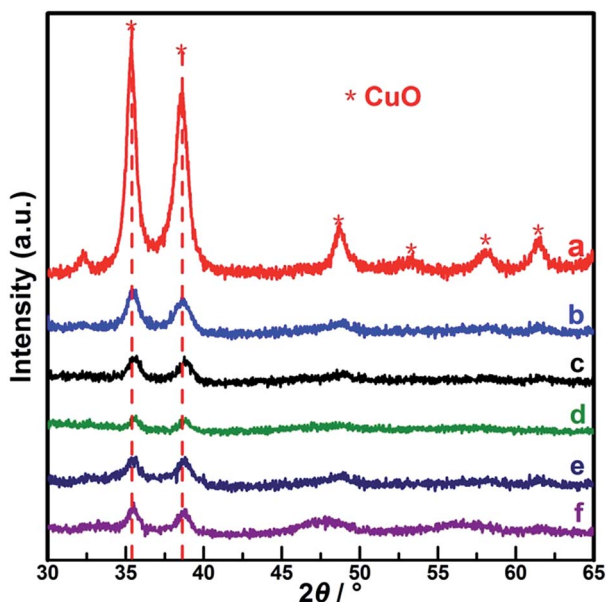


Fig. 1 XRD patterns of samples. (a) CuO, (b) $(\text{CeO}_x)_{0.125}\text{CuO}$ -350, (c) $(\text{CeO}_x)_{0.25}\text{CuO}$ -350, (d) $(\text{CeO}_x)_{0.50}\text{CuO}$ -350, (e) $(\text{CeO}_x)_{0.25}\text{CuO}$ -400, and (f) $(\text{CeO}_x)_{0.25}\text{CuO}$ -450.



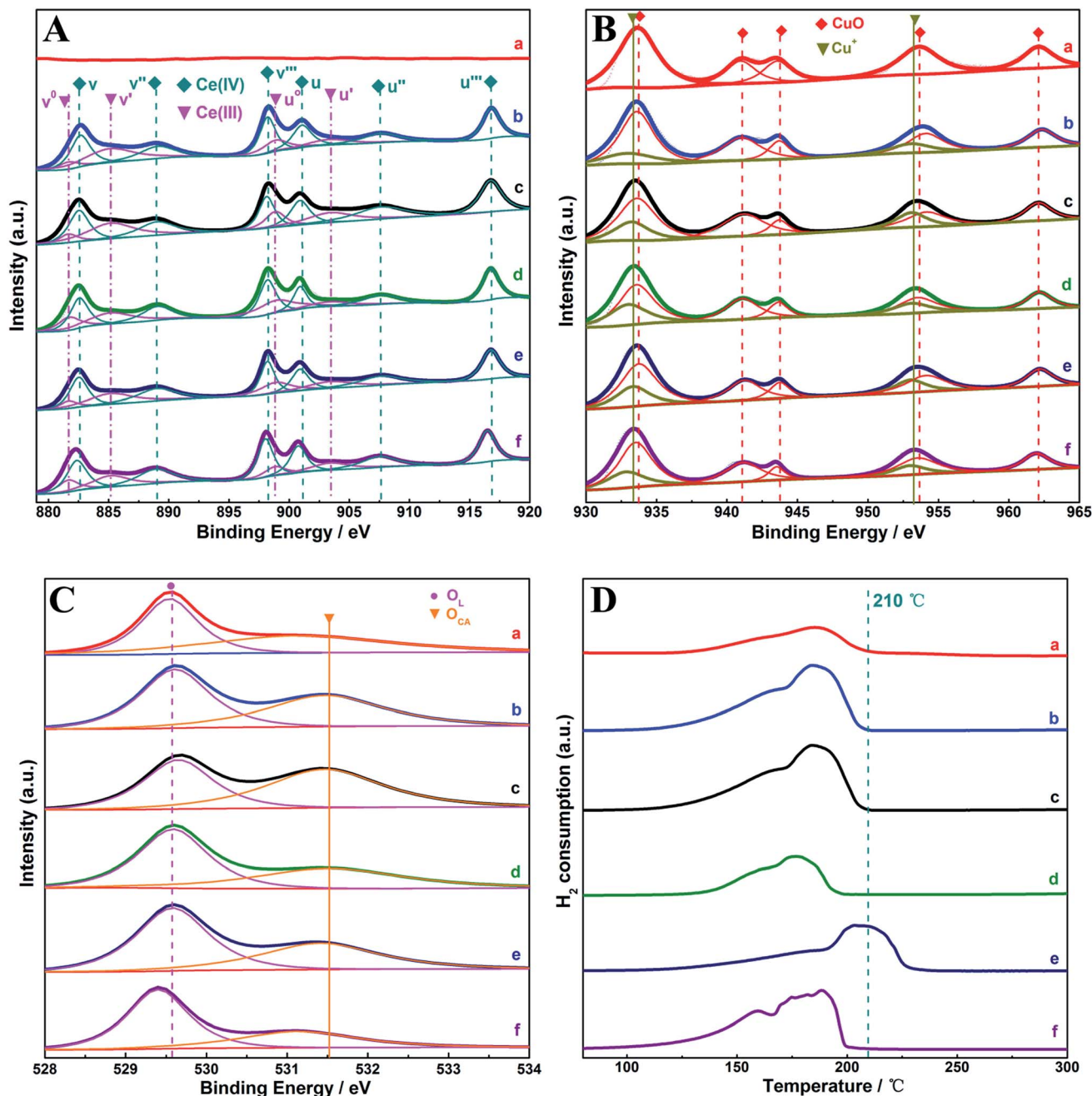


Fig. 2 XPS spectra and H₂-TPR profiles of the samples. (A) Ce 3d spectra; (B) Cu 2p spectra (C) O 1s spectra. (D) H₂-TPR profiles (a) CuO, (b) (CeO_x)_{0.125}CuO-350, (c) (CeO_x)_{0.25}CuO-350, (d) (CeO_x)_{0.50}CuO-350, (e) (CeO_x)_{0.25}CuO-400, (f) (CeO_x)_{0.25}CuO-450.

$$A_{\text{Ce(III)}} = A_{\text{v}_0} + A_{\text{u}_0} + A_{\text{v}'} + A_{\text{u}'}$$

$$A_{\text{Ce(IV)}} = A_{\text{v}} + A_{\text{u}} + A_{\text{v}''} + A_{\text{u}''} + A_{\text{v}'''} + A_{\text{u}'''}$$

$$C_{\text{Ce(III)}} = A_{\text{Ce(III)}} / (A_{\text{Ce(IV)}} + A_{\text{Ce(III)}})$$

Similarly, the contents of Cu⁺ in the (CeO_x)_m-CuO samples were also roughly estimated by the ratio between the sum of the peak area correlated to Cu⁺ and total area of the peaks around Cu 2p_{3/2} and Cu 2p_{1/2}. The corresponding species distributions of samples are listed in Table 1. In general, the increasing Ce

composition among the (CeO_x)_mCuO-350 samples ($m = 0.125, 0.25$, and 0.50) resulted in gradually decreased Ce(III) content accompanied by the highest distribution of Cu⁺ achieved on (CeO_x)_{0.25}CuO-350, suggesting that controlling the amount of Ce doped into CuO was crucial to confine the presence and distribution of Cu⁺ on the surface of the samples. These interactive changes among active components were always accepted as a hint or consequence of the presence of a synergistic effect in the composite samples.^{26,27} Note that the synergistic effect in the present samples was distinguished with that of the reported Cu-based NH₃-SCR-NO_x catalysts (especially with ternary-

Table 1 Surface atom concentrations (derived from XPS) and H₂ consumptions (derived from H₂-TPR) of the samples

Samples	Surface composition (at%)					Con.H ₂ ^a	
	Ce (Ce(III)/Ce)	Cu (Cu ⁺ /Cu)	O (O _{CA} /O)				
CuO-350	0	0	39.1	0	60.9	34.1	1
(CeO _x) _{0.125} CuO-350	8.5	41.5	24.7	18.8	66.8	49.4	2.2
(CeO _x) _{0.25} CuO-350	9.8	40.1	21.2	27.2	69.0	57.8	2.6
(CeO _x) _{0.50} CuO-350	15.0	27.7	18.9	23.4	66.1	41.7	1.4
(CeO _x) _{0.25} CuO-400	13.1	30.8	22.3	22.5	64.6	42.9	1.5
(CeO _x) _{0.25} CuO-450	12.4	30.5	24.8	19.3	62.8	35.3	1.7

^a The relative consumptions of H₂ (Con.H₂) on the samples were calibrated by the ratio between Con.H₂ peak area of the samples and that of the reference CuO-350 sample.

component samples) since one of the main features of the synergistic effect in these catalysts was embodied by the presence of isolated Cu²⁺ (with higher BE compared to that of CuO).^{27,29} In addition, the synergistic effect in composite structures would also influence the distribution of other active species in the samples such as active oxygen species required for redox reactions. The oxygen spectra of the prepared samples are shown in Fig. 2C, where lattice oxygen O_L (528.0–531.5 eV) and chemisorbed oxygen O_{CA} (529.0–533.7 eV) were fitted for all the samples,^{34–36} and the corresponding distributions are listed in Table 1. It is clear that higher concentrations of chemisorbed oxygen were distributed on (CeO_x)_mCuO-350 samples compared to those on CuO-350, and (CeO_x)_{0.25}CuO-350 showed the highest concentration of chemisorbed oxygen. Considering that chemisorbed oxygen was one of the key active species for most redox reactions, the result further manifested that doping Ce into CuO in a proper way to modulate the distribution of Cu⁺ on the surface of the present samples could be beneficial to optimize their surface state for the reactions. For (CeO_x)_{0.25}CuO samples obtained employing different temperature, with the increase of processing temperature, both the contents of Ce(III) and Cu⁺ were lowered, and the trend was accompanied by a decrease in the concentration of O_{CA}, suggesting that the higher calcination temperature would negatively influence the surface distribution of the active species in the present Cu–Ce composite oxides.

The influence of the surface distribution of active species on the redox properties of the samples was investigated by H₂-TPR, and the results are shown in Fig. 2D. The main H₂ consumption peaks are located in temperature range of 100–210 °C for all the samples except for (CeO_x)_{0.25}CuO-400 (extended to 230 °C), which was quite low compared to the temperature of reduced bulk CuO or Ce species reported on Cu-based composites.^{21,27,37–39} Thus, these features could be excluded for our samples. Generally, the H₂ consumption peaks below 260 °C for the reported Cu–Ce composite samples are attributed to the reduction of well-dispersed surface Cu species or active Cu species, such as Cu²⁺, on Cu–Ce interface.^{21,27,38} Obviously, these species were difficult to be estimated for the current samples according to previous XPS measurements. To illuminate the

origin of H₂-TPR features for our samples, the relative consumptions of H₂ (Con.H₂) for the samples were calibrated by the ratio between Con.H₂ peak area of samples and that of reference CuO-350 sample, and the results are listed in Table 1. From the data shown in the Table, even assuming that O_{CA} was very sensitive to consume H₂, the change of amounts of O_{CA} among the samples was not sufficient to solely afford the corresponding alternations in the H₂-consumption. Interestingly, the Con.H₂ values among the samples were roughly governed by the sum of the contents of Cu⁺ and O_{CA}, *i.e.*, the higher the sum of the contents, the higher the H₂-consumption; this dependence directly indicated that Cu⁺ was one of the significant species to react with H₂ on the present samples. The point could be demonstrated by the comparison between the (CeO_x)_{0.25}CuO-450 and reference CuO-350 samples. Although higher calcination temperature caused some deviation for (CeO_x)_{0.25}CuO-450 in contrast to the abovementioned dependence, with the baseline that O_{CA} content on (CeO_x)_{0.25}CuO-450 and CuO-350 were close to each other, the much higher H₂-consumption for (CeO_x)_{0.25}CuO-450 is directly linked to the additional Cu⁺ content on the sample compared to that on CuO-350. Therefore, it was reasonable to propose that the redox properties of the present Cu–Ce composite oxide were directly connected with the existence and distribution of Cu⁺ on the surface of the present samples.

Considering that the surface acid property of the catalyst was crucial for NH₃-SCR-NO_x, NH₃-TPD tests were performed for our samples, and the results are shown in Fig. 3. For CuO-350, the broad and weak desorption peak observed at 150–350 °C could be attributed to the feature of weak Brønsted acid sites on the CuO surface; in contrast, the desorption peaks of NH₃ in 350–450 °C were exhibited for the three (CeO_x)_mCuO-350 samples,

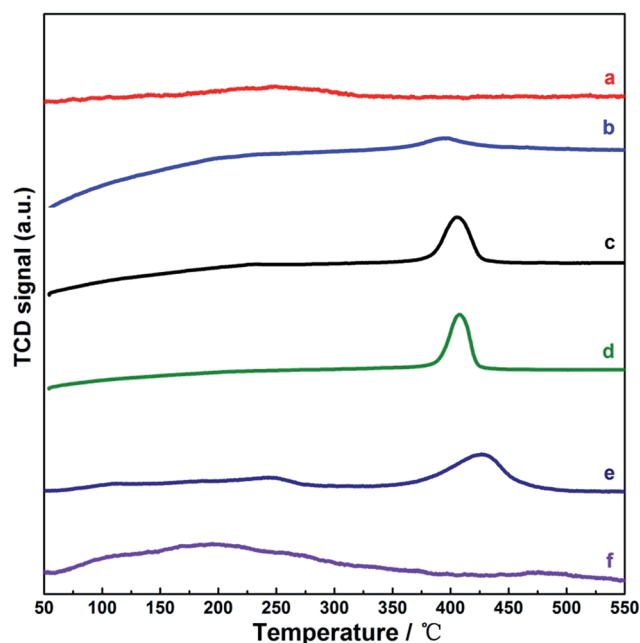


Fig. 3 NH₃-TPD profiles for the samples. (a) CuO, (b) (CeO_x)_{0.125}CuO-350, (c) (CeO_x)_{0.25}CuO-350, (d) (CeO_x)_{0.50}CuO-350, (e) (CeO_x)_{0.25}CuO-400, and (f) (CeO_x)_{0.25}CuO-450.



which were related to the structural Brønsted acid sites inferred to moderate and strong acidity or Lewis acid sites.^{25,40–42} According to the peak areas for NH_3 desorption, a five-fold change in the desorption (or adsorption) amount of NH_3 emerged by increasing the Ce composition with m ranging from 0.125 to 0.25 for the $(\text{CeO}_x)_m\text{CuO}$ -350 samples was observed; however, only a 1.2-fold increase in the desorption amount of NH_3 was observed on increasing m from 0.25 to 0.50. These results suggested that a 0.25 Ce/Cu molar ratio was the key composition to obtain the desired distribution of the acid sites in the present samples. Note that the sums of surface concentration of Ce(III) and Cu^+ on $(\text{CeO}_x)_{0.25}\text{CuO}$ -350 and $(\text{CeO}_x)_{0.50}\text{CuO}$ -350 were close to each other, but were higher than those of $(\text{CeO}_x)_{0.125}\text{CuO}$ -350 (Table 1). In view of the Ce(III) species or isolated transition metal ions in the composite oxides generally accepted as the resource to produce acid surface (Lewis acid sites),^{26,29} the observed differences of the acid amount in these $(\text{CeO}_x)_m\text{CuO}$ -350 samples could be attributed to the different distribution of the acid source species such as Ce(III) species and Cu^+ on the samples. Moreover, for the $(\text{CeO}_x)_{0.25}\text{CuO}$ samples prepared at different calcination temperatures (Fig. 3c, e and f), the sum of the Ce(III) species and Cu^+ concentration in the samples decreased with the increasing calcination temperature, which was accompanied by the lowered acid amount (desorption amount of NH_3) and weakened acidity (NH_3 desorption at low temperature). This connection further confirmed the proportional dependence of the sum distribution of Ce(III) species and Cu^+ on the surface acid properties among our samples. These results also suggested that the surfaces of the $(\text{CeO}_x)_m\text{CuO}$ -350 and $(\text{CeO}_x)_{0.25}\text{CuO}$ -400 samples were dominated by Lewis acid sites.

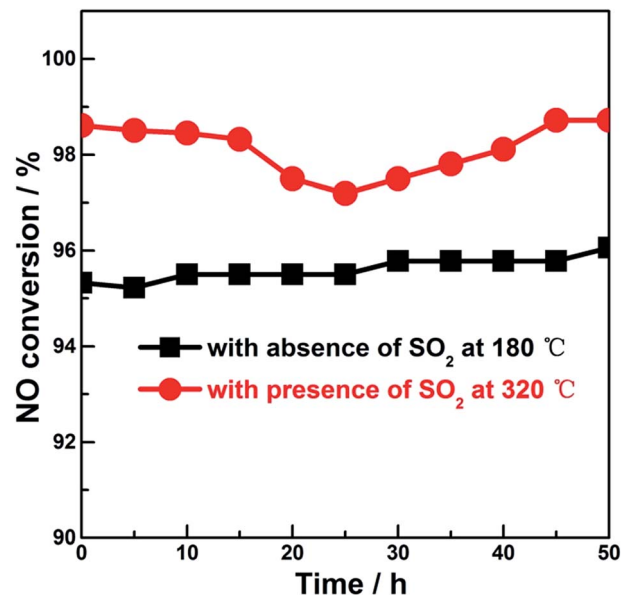


Fig. 5 NO conversion as a function of time over $(\text{CeO}_x)_{0.25}\text{CuO}$ -350.

The catalytic performances of the samples were investigated for NH_3 -SCR-NO in the absence and presence of SO_2 , which is shown in Fig. 4A and B. For reaction in the absence of SO_2 (Fig. 4A), in contrast to the low NO conversion (not more than 70%) over CuO -350, all the Ce-introduced samples could achieve NO conversion beyond 80% within the temperature window of 130–230 °C. Among the $(\text{CeO}_x)_m\text{CuO}$ -350 samples, the best performance (the highest NO conversion and the broadest temperature window) was observed for $(\text{CeO}_x)_{0.25}\text{CuO}$ -

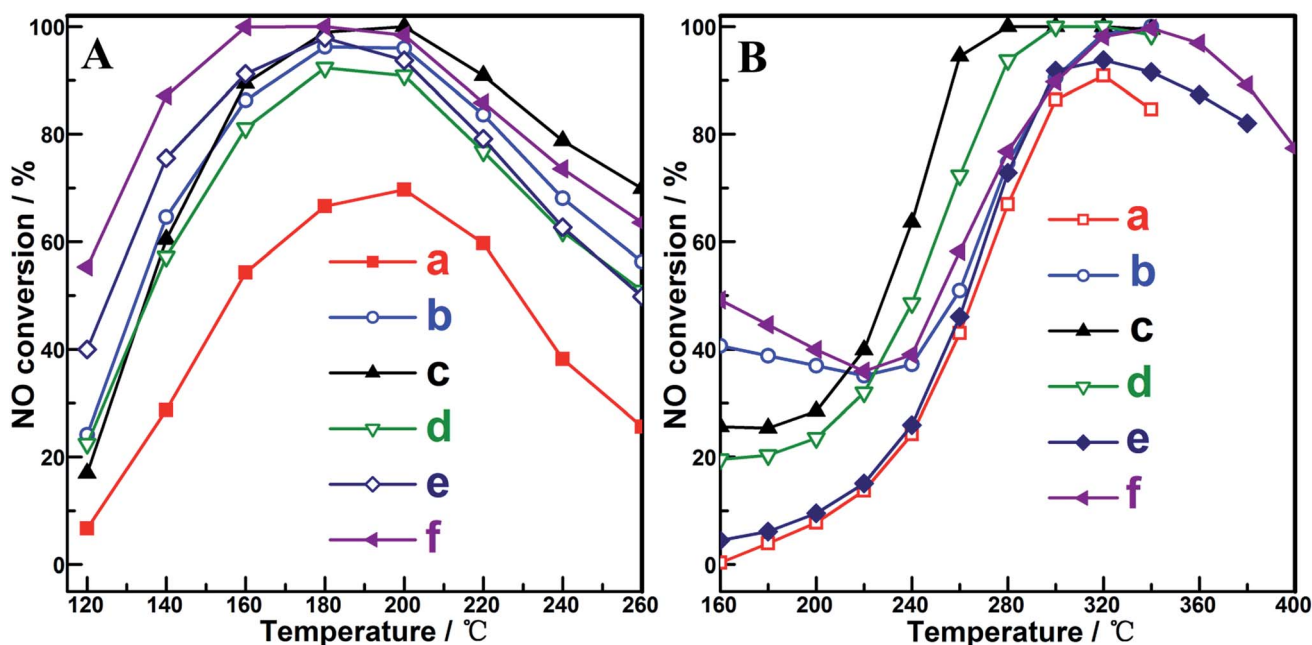


Fig. 4 Dependence of NO conversion on reaction temperature with (A) SO_2 -absence and (B) SO_2 -presence. Reaction conditions: $[\text{NO}] = [\text{NH}_3] = 1000 \text{ ppm}$, $[\text{O}_2] = 5\%$, $[\text{SO}_2] = 200 \text{ ppm}$ (when used), $\text{GHSV} = 40\,000 \text{ h}^{-1}$. (a) CuO , (b) $(\text{CeO}_x)_{0.125}\text{CuO}$ -350, (c) $(\text{CeO}_x)_{0.25}\text{CuO}$ -350, (d) $(\text{CeO}_x)_{0.50}\text{CuO}$ -350, (e) $(\text{CeO}_x)_{0.25}\text{CuO}$ -400, and (f) $(\text{CeO}_x)_{0.25}\text{CuO}$ -450.



Table 2 Preparation details of various Cu-based NH₃-SCR-NO catalysts and corresponding catalytic measurements in the presence of SO₂

Catalysts	Preparation method	Features of samples	Reaction condition	Temperature window ^a	Final conversion of NO ^b	References
Ce-Fe/WMH	Two-step impregnation	Enrichment of Ce ³⁺ on catalyst surface	1000 ppm NO, 1000 ppm NH ₃ , 100, 500, and 1000 ppm (when used) SO ₂ and 3 vol% O ₂ balanced with N ₂ , GHSV at 10 000 h ⁻¹	250 °C	92%	46
Ni-Ce-La composite oxide	Mulsifier-free emulsion polymerization	More lattice defects	1000 ppm NO, 1000 ppm NH ₃ , 500 ppm SO ₂ and 5 vol% O ₂ balanced with N ₂ , GHSV at 20 000 h ⁻¹	300 °C	90%	47
SnO ₂ -MnO _x -CeO ₂	Co-precipitation	Crystallinity loss	1000 ppm NO, 1000 ppm NH ₃ , 100 ppm SO ₂ , 2 vol% O ₂ balanced with N ₂ , GHSV at 35 000 h ⁻¹	220 °C	94%	48
MeO _x -MnO _x /SAPO-34 (Me = Pr or Ce)	Solvent dispersion	Rich in oxygen vacancies	800 ppm NO, 800 ppm NH ₃ , 100 ppm SO ₂ and 5 vol% O ₂ balanced with Ar, GHSV at 40 000 h ⁻¹	200 °C	60.5%	49
Ce-Cu composite oxides	Co-precipitation	Cu ⁺ on the surface of Ce-doped CuO	1000 ppm NO, 1000 ppm NH ₃ , 200 ppm SO ₂ and 5 vol% O ₂ balanced with N ₂ , GHSV at 40 000 h ⁻¹	250–350 °C	98%	This work

^a The temperature range for NO conversion beyond 80%. ^b The highest NO conversion.

350, which could be attributed to the outstanding distribution of the active species such as Ce(III), Cu⁺, and O_{CA}, and the corresponding optimized redox and surface acid properties of the sample. For the (CeO_x)_{0.25}CuO samples obtained at 400 °C and 450 °C, the higher NO conversion on these compared to that on (CeO_x)_{0.25}CuO-350 could be correlated to the more obvious feature of the weak acid sites observed in the low temperature range of 80–300 °C on these samples (seen in the NH₃-TPD profiles of Fig. 3e and f) since it was found that weak acid sites on the samples were sufficient to accomplish NH₃-SCR-NO at low reaction temperature.^{19,43} Considering that SO₂ was always present in normal NH₃-SCR-NO_x processes and a main challenge for current Cu-based catalysts is their poor SO₂ resistance, more practical and meaningful performance of NH₃-SCR-NO_x Cu-based catalysts should be focused on the catalytic behaviour in the presence of SO₂. As shown in Fig. 4B, meaningful conversion of NO (beyond 80%) in the presence of SO₂ on the samples was all shifted to the high temperature range (beyond 240 °C) compared to that in the absence of SO₂, and the best SO₂-resistant NH₃-SCR-NO performance with NO conversion close to 100% in the temperature window of 250–350 °C was obtained on (CeO_x)_{0.25}CuO-350 among the samples. The surface acid feature of catalyst was found to be the decisive factor that governed the performance of NH₃-SCR-NO_x. In our study, the acid amount on (CeO_x)_{0.50}CuO-350 was higher than that on (CeO_x)_{0.25}CuO-350; however, the SO₂ resistance and NO converting performance between two samples followed an opposite order, suggesting that the acid feature was not the only significant factor affecting the SO₂-resistant NH₃-SCR-NO. In contrast, the order of SO₂-resistant performance among our samples approximately followed their Cu⁺ concentration sequence. The outstanding SO₂-resistant performance of (CeO_x)_{0.25}CuO-350 could be correlated to its highest surface distribution of Cu⁺ among the samples. The correlation is in agreement with the recent findings on SO₂-removal on Cu-containing composites, in which the existence of

Cu⁺ in a notable amount was demonstrated to be crucial for the chemical removal of SO₂ due to its positive role in transferring charge or active oxygen species through the Cu(I)–Cu(II) redox cycle.^{44,45} The steady-state performance of (CeO_x)_{0.25}CuO-350 was further investigated under SO₂-absence and SO₂-presence conditions, as shown in Fig. 5. For both cases, the slight fluctuations within 2% were obtained on the sample, exhibiting the desired application stability of the present samples. In view of estimating the influence of the gas hourly space velocity (GHSV) on SO₂-resistant NH₃-SCR-NO performance from a reviewer, some typical studies were surveyed and are listed in Table 2. Although the GHSV employed in this study was in the middle rank compared with literature works, the higher concentration of SO₂ in the reactant gas could ensure the kinetic estimation for the present samples. The comparable high NO conversion in the presence of higher concentration of SO₂ obtained in the present (CeO_x)_mCuO samples suggested that these samples are a type of efficient and notable SO₂-resistant NH₃-SCR-NO_x catalyst, especially in view of the distinguished Cu⁺ distributions on the samples.

4. Conclusions

The synergistic effect involved in the present Cu–Ce composite structures, prepared *via* a common co-precipitation method, were found to be pronounced for producing distinguished Cu⁺ species on the surface of Ce-doped CuO. The distribution of Cu⁺ could be modulated by controlling the amount of Ce introduced, which was functional to optimize the redox and surface acid features of the samples, leading to outstanding SO₂-resistant NH₃-SCR-NO performance. These findings could serve for understanding the catalytic behaviour of Cu–Ce composite catalysts as well as innovating more efficient Cu-based catalysts for NH₃-SCR-NO_x and other heterogeneous catalysis applications upon Cu catalysts.



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