

Cite this: *RSC Adv.*, 2017, 7, 9975Received 12th January 2017
Accepted 30th January 2017

DOI: 10.1039/c7ra00506g

rsc.li/rsc-advances

CASSCF and CASMP2 study on the photoisomerization mechanisms of [tris(trialkylsilyl)silyl]cycloctetrasilene and related cyclobutene molecules†

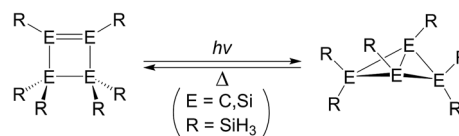
Ming-Der Su^{*ab}

The mechanisms for photochemical isomerization reactions are studied theoretically using the model systems, cycloctetrasilene and cyclobutene, both of which feature one four-membered-ring and one double bond, using the CAS(6,6)/6-311G(d) and CAS-MP2-(6,6)/6-311++G(3df,3pd) methods. The mechanisms, which are governed by the conical intersection concept, play a crucial role in these photorearrangement reactions, and explain the reaction pathways. The model computations demonstrate that the preferred reaction path for the photoisomerization reactions for both cycloctetrasilene and cyclobutene is as follows: four-membered-ring reactant → Franck–Condon region → conical intersection → butterfly-like photoproduct. These results allow a number of predictions to be made.

I. Introduction

For at least three decades, the chemistry of disilene, which features a silicon–silicon double bond, has been the subject of extensive study because of its novel structure and bonding properties.^{1,2} However, only very few photochemical reactions that involve the disilene systems have been studied experimentally.³

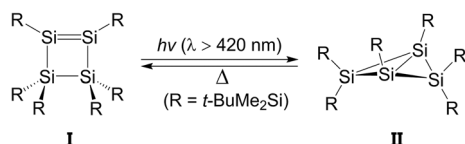
About 15 years ago, Kira and co-workers reported a distinctive photochemical reaction that rearranges hexakis(*tert*-butyldimethylsilyl)cycloctetrasilene (I) to its corresponding isomer, hexakis(*tert*-butyldimethylsilyl)-bicyclo[1,1,0]tetrasilane (II), as shown in Scheme 1.⁴ After heating, (II) is isomerized back to the initial reactant (I). However, to the best of the author's knowledge, there has been no theoretical study of the mechanism for this photo-isomerization reaction for cycloctetrasilene (I).^{4c} This may be because the computational methods for studying



Scheme 2

photochemistry were not so mature at that time. The author uses a simplified cycloctetrasilene (**Si-Rea-S₀**) as a model to study its photochemical mechanism for the isomerization reaction to bicyclo[1,1,0]tetrasilane (**Si-Pro-S₀**) using modern theoretical methods, *i.e.*, CASSCF (the complete-active-space SCF)⁵ and the multi-reference Møller–Plesset (CAS-MP2) algorithm.⁶ For the purposes of comparison, a similar mechanistic study is also made of a four-carbon system, **C-Rea-S₀**, as represented in Scheme 2.⁷

Conical intersection (CI) points⁸ can occur on the potential surfaces, from which the excited state decays to the ground state surface at the same multiplicity (usually from singlet to singlet). Therefore, the CI mechanism is used in this work to interpret the reaction trajectories for the photochemistry of both **Si-Pro-S₀** and **C-Pro-S₀** molecules. This paper shows that the CIs⁸ play a decisive role in the photo-rearrangement reactions for cycloctetrasilene and cyclobutene molecules, which have a four-membered-ring.



Scheme 1

^aDepartment of Applied Chemistry, National Chiayi University, Chiayi 60004, Taiwan.
E-mail: midesu@mail.nyu.edu.tw

^bDepartment of Medicinal and Applied Chemistry, Kaohsiung Medical University, Kaohsiung 80708, Taiwan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ra00506g

II. Theoretical methods

The CASSCF computations are executed using the MCSCF⁹ and CAS-MP2 (ref. 9) programs in GAUSSIAN 09.¹⁰ The active space



that is used to describe the photo-rearrangements for the four-membered-ring molecules (Scheme 2) that are studied in this work include six electrons in six orbitals, *i.e.*, $1\pi(\text{E}=\text{E}) + 1\pi^*(\text{E}=\text{E}) + 2\sigma(\text{E}-\text{E}) + 2\sigma^*(\text{E}-\text{E})$ orbitals. The state-averaged CASSCF(6,6) method was adopted to verify geometry on the intersection space (*i.e.*, root = 2). Moreover, the excited state gradient was obtained by using the CSASSCF method. The harmonic frequencies are computed analytically at the CAS(6,6)/6-311G(d) level of theory and are used to characterize either minima or transition states. Unless otherwise stated, the relative energies that are shown in this study are those that are ascertained at the CAS-MP2(6,6)/6-311++G(3df,3pd) level, using the CAS(6,6)/6-311G(d) geometry. These are designated CAS-MP2 and CASSCF, respectively. The Cartesian coordinates and the energies that are computed for the various points are available as ESI.†

III. General considerations

According to the available experimental results,⁴ there is no indication that a triplet state participates in these photoisomerization reactions. The CAS-MP2 computational results also show that the relative energies (kcal mol⁻¹) for the Franck-Condon (FC) regions are as follows: **Si-Rea-S₀** (0.0) < **Si-FC-S₁** (68.8) < **Si-FC-T₁** (93.6) and **C-Rea-S₀** (0.0) < **C-FC-S₁** (116.8) < **C-FC-T₁** (121.8).¹¹ From these theoretical results, it is concluded that triplet states do not play any role in the photochemical reactions for these cyclic disilene molecular systems, so this study concentrates on singlet surfaces for both **Si-Rea-S₀** and **C-Rea-S₀** reactants.¹²

In order to determine possible excited-state reaction paths that lead to CIs, a molecular orbital representation of the relationship between the likely reaction pathway and intersection is schematically illustrated in Scheme 3 (E = C and Si),¹³ which is a foundation for these theoretical studies. Upon irradiation with light, a singlet four-membered-ring molecule (**1**) is moved up to a singlet excited species (**2**), succeeding an intramolecular double bond (E=E) rotation. Then, as shown in (**3**), this excited molecule **1** can be considered as a composition of four kinds of carbenes (>E), existing four spin-up σ orbitals and four spin-

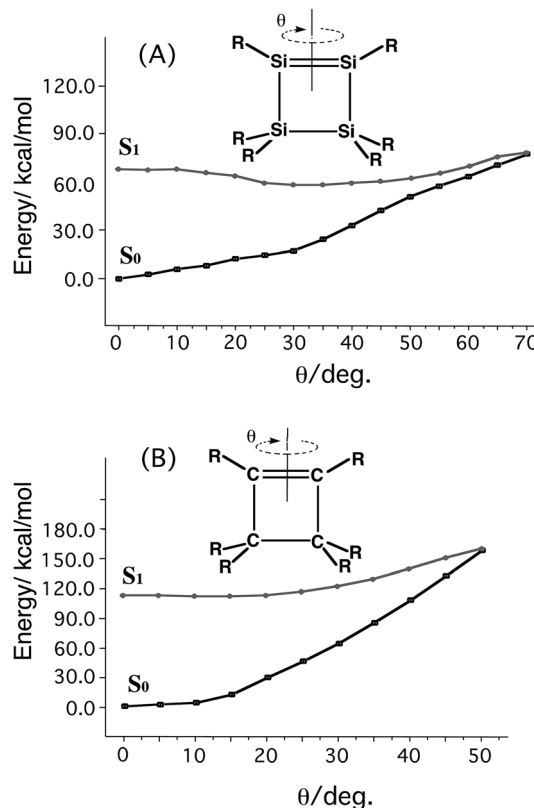


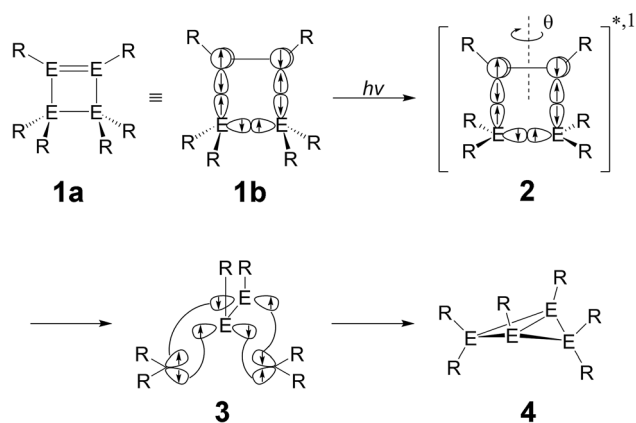
Fig. 1 The scans for finding a CI pathway using an E=E (E = C and Si) rotation in the reactants (A) Si-Rea-S₀ and (B) C-Rea-S₀; R = SiH₃ for the angle, θ , optimized for the S_0 and S_1 states at the CAS(6,6)/6-311G(d) level of theory. See the text and ESI.†

down σ orbitals.¹³ Consequently, through the conical intersection, a two-three-membered-ring photoproduct (**4**) can be obtained.

Using the molecular orbital analysis that is given in Scheme 3, since the narrow energy gap between S_0 and S_1 energy surfaces strongly implies the existence of a CI point, **Si-Rea-S₀** and **C-Rea-S₀** are used as examples to determine whether there is a low-lying S_1/S_0 surface crossing. Fig. 1 shows the potential energy profiles for the rotational angle, θ , as shown in Scheme 3.¹⁴ Fig. 1(A) and (B) show that the reaction pathways all terminate at a very similar S_1-S_0 energy gap at around 75° and 60°, respectively. These energy profile results suggest that a CI point occurs in the intramolecular rotation for the four-membered-ring geometries. In the next section, these findings are used to explain the mechanisms for the photochemical isomerization reactions of **Si-Rea-S₀** and **C-Rea-S₀**.

IV. Results and discussion

The discussion in Section III shows that there should exist a reaction route for photochemical isomerization reactions on the singlet excited energy surfaces of **Si-Rea-S₀** that contains a conical intersection point. Fig. 2 shows the energies for the key points relative to the energy of the reactant, **Si-Rea-S₀**. The optimized geometries for these important points on the



Scheme 3



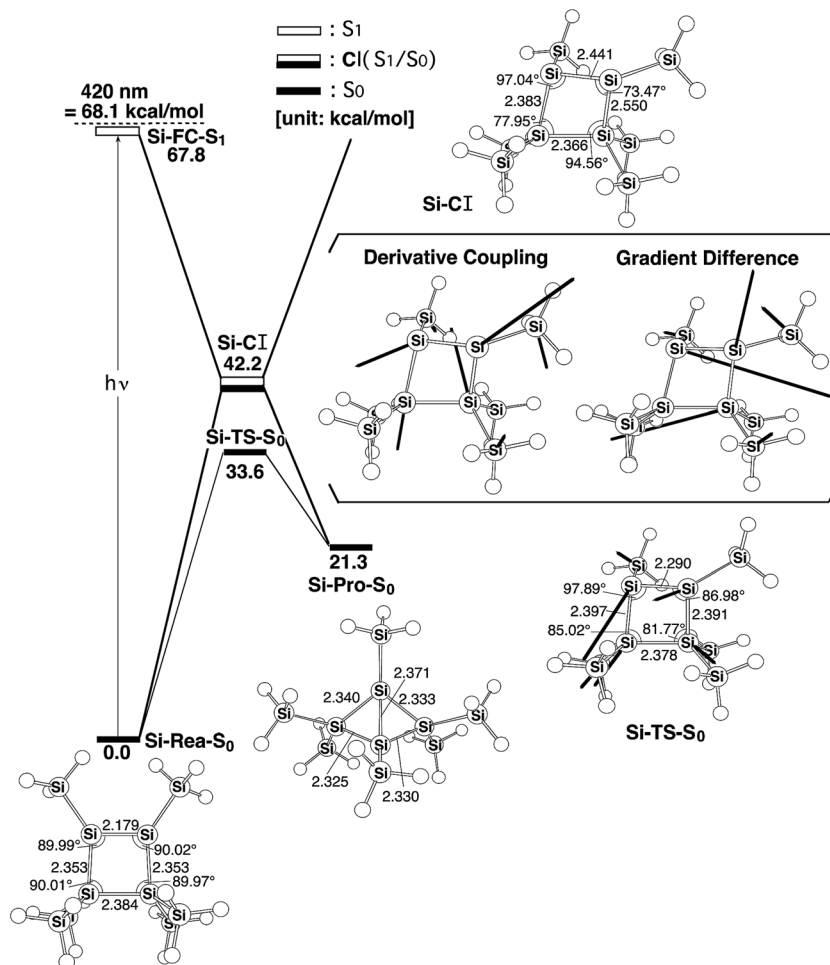


Fig. 2 The energy surfaces for the photochemical isomerization reaction for **Si-Rea-S₀**. The abbreviations, FC, TS and CI, denote Franck-Condon, transition state and conical intersection, respectively. The relative energies are obtained at the CAS-MP2-(6,6)/6-311++G(df,pd)//CAS(6,6)/6-311G(d) level of theory. The geometrical parameters for the CASSCF-optimized structures for the stationary points are also shown. Hydrogen atoms are omitted for clarity. The bold arrows show the main atomic motions in the **Si-TS-S₀** eigenvector. The derivative coupling and the gradient difference vectors – those which increase the degeneracy – are computed using CASSCF at the conical intersection, **Si-Cl**. For more information see the text.

mechanistic pathways in Fig. 2 are also shown in the same figure.

A comparison with the CASSCF-optimized geometry for the silicon reactant (**Si-Rea-S₀**), as shown in Fig. 2, shows that the theoretical values for the bond lengths and the bond angles in these four-membered-ring species are in good agreement with the experimental values.⁴ The bond lengths and the bond angles are within 0.07 Å and 5.9°, respectively. Therefore, CASSCF and CAS-MP2 computations should be satisfactory for a theoretical study of the photochemical and thermal reactions.

Initially, cyclotetrasilene (**Si-Rea-S₀**) with a four-silicon-membered-ring is irradiated by visible light and is promoted to the Franck-Condon region at the first excited singlet state (**Si-FC-S₁**), which is shown on the left side of Fig. 2. Subsequently, this species relaxes to the conical intersection point (**Si-Cl**), from which it can go to either the final photoproduct (**Si-Pro-S₀**) or the initial reactant (**Si-Rea-S₀**). These CAS-MP2 computations show that the relative energies for **Si-FC-S₁**, **Si-Cl**, and **Si-Pro-S₀**

with respect to **Si-Rea-S₀** are 68, 42, and 21 kcal mol^{−1}, respectively. The theoretical evidence strongly suggests that when cyclotetrasilene (**Si-Rea-S₀**) is irradiated by visible light, it easily undergoes a Si=Si double bond rotation to produce either the reactant (**Si-Rea-S₀**) or the corresponding bicyclo[1,1,0]tetrasilane (**Si-Pro-S₀**). The photochemical isomerization mechanism for **Si-Rea-S₀** that is shown in Fig. 2 is represented as follows:



The thermal (dark) reaction on the ground singlet state energy surface that is shown in Fig. 2 is also studied. Fig. 2 shows that the **Si-TS-S₀** point combines **Si-Rea-S₀** and **Si-Pro-S₀** on the ground singlet surface and that its energy lies 15 kcal mol^{−1} below the energy of the **Si-Cl** point. It is worthy of note that the CAS-MP2 calculations give barrier heights for **Si-Rea-S₀** → **Si-TS-S₀** and **Si-Pro-S₀** → **Si-TS-S₀** that are approximately 34



The photochemical conversion of **C-Rea-S₀** to **C-Pro-S₀** and its thermal reversion are studied. The computational results are all listed in Fig. 3. Similarly to the model for cyclotetrasilene

Cyclobutene (photo): **C-Rea-S₀** + $h\nu$ → **C-FC-S₁** → **C-CI** → **C-Pro-S₀**

However, the CAS-MP2 results demonstrate that much higher energy (117 kcal mol⁻¹) of **C-FC-S₁** than that of the irradiating light (420 nm = 68 kcal mol⁻¹) makes this photo-rearrangement reaction energetically unfeasible,¹⁵ so the energy of photo-irradiation is insufficient to generate the butterfly-like product, **C-Pro-S₀** *via* radiation-less decay (*i.e.*, the

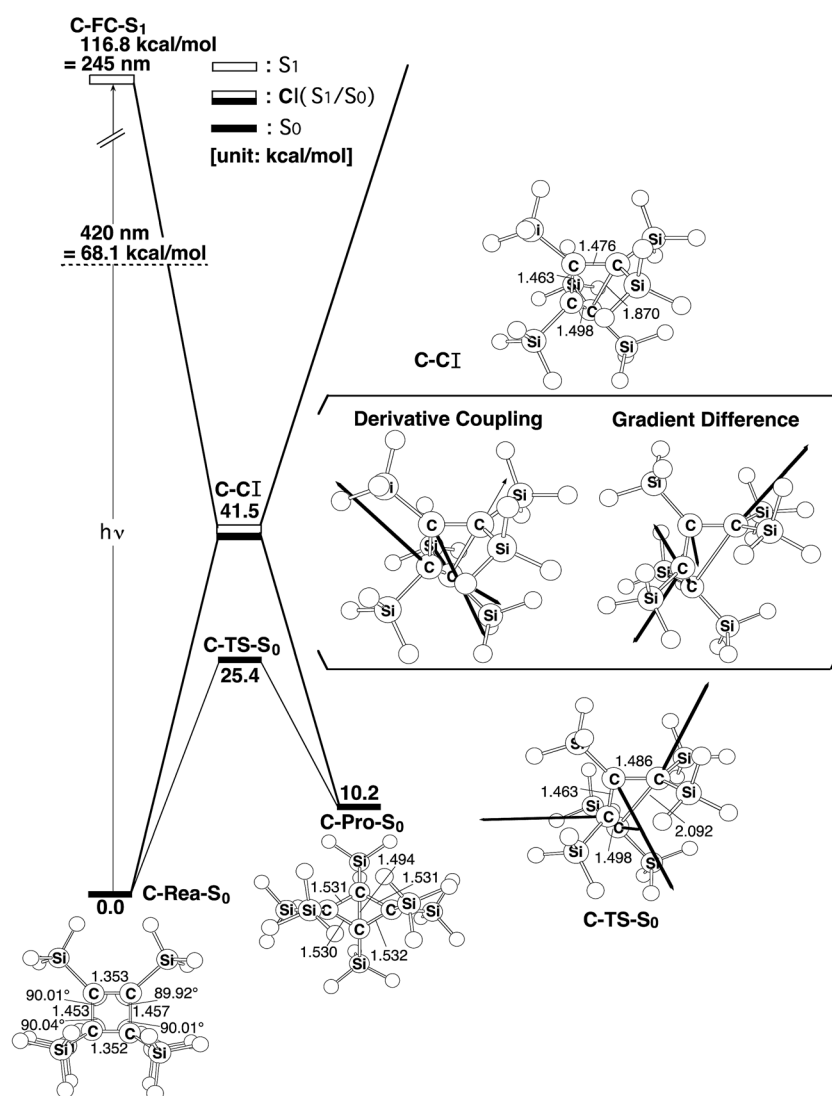


Fig. 3 The energy surfaces for the photochemical isomerization reaction for **C-Rea-S₀**. The abbreviations, FC, TS and CI, denote Franck-Condon, transition state and conical intersection, respectively. The relative energies are obtained at the CAS-MP2-(6,6)/6-311++G(df,pd)//CAS(6,6)/6-311G(d) level of theory. The geometrical parameters for the CASSCF-optimized structures for the stationary points are also shown. Hydrogen atoms are omitted for clarity. The bold arrows show the main atomic motions in the **C-TS-S₀** eigenvector. The derivative coupling and the gradient difference vectors – those that increase the degeneracy – are computed using CASSCF at the conical intersection, **C-CI**. For more information, see the text.

CI channel). Consequently, no photo-isomerization reaction should be observed in these cyclobutene systems for irradiation with visible light. However, the thermal reversion (**C-Pro-S₀** → **C-Rea-S₀**) is predicted to be easily achieved, since the energy that is required to surmount the reverse barrier (**C-Pro-S₀** → **C-TS-S₀**) is calculated to be 10 kcal mol^{−1} less than that which is required to surmount the forward barrier (**C-Rea-S₀** → **C-TS-S₀**), using the CAS-MP2 calculations. To the best of the author's knowledge, no experimental reports of the photochemical reactions of cyclobutenes have been published so the CAS-MP2 computational results that are shown in Fig. 3 remain a prediction.

V. Conclusion

The photochemical mechanisms for the isomerization reactions for **Si-Rea-S₀** and **C-Rea-S₀**, which have a skeleton that has one four-membered-ring and one double bond, are studied using the CASSCF and CAS-MP2 levels of theory. This is the first theoretical study of the reaction trajectory and the first theoretical estimation of the activation energy and the reaction enthalpy for both the singlet excited state and the singlet ground state. This study interprets the photochemical behavior of these four-membered-ring species that contain group 14 elements using a simple conical intersection mechanism. The available experimental findings for cyclotetrasilene are verified⁴ and the photochemical isomerization reactions for analogous cyclobutene systems are predicted.

Acknowledgements

The author is grateful to the National Center for High-Performance Computing of Taiwan for generous amounts of computing time, and the Ministry of Science and Technology of Taiwan for the financial support. In particular, the author also wishes to thank Professor Michael A. Robb, Dr S. Wilsey, Dr Michael J. Bearpark, (University of London, UK) and Professor Massimo Olivucci (Università degli Studi di Siena, Italy), for their encouragement and support during his stay in London. Special thanks are also due to reviewers 1 and 2 for very help suggestions and comments.

References

- 1 R. West, M. J. Fink and J. Michl, *Science*, 1981, **214**, 1343.
- 2 For reviews on disilenes, see: (a) R. West, *Pure Appl. Chem.*, 1984, **56**, 163; (b) G. Raabe and J. Michl, *Chem. Rev.*, 1985, **85**, 419; (c) A. H. Cowley and N. C. Norman, *Prog. Inorg. Chem.*, 1986, **34**, 1; (d) R. West, *Angew. Chem., Int. Ed. Engl.*, 1987, **26**, 1201; (e) G. Raabe and J. Michl, in *The Chemistry of Organic Silicon Compound*, ed. S. Patai and Z. Rappoport, Wiley, New York, 1989, part 2, ch. 17; (f) T. Tsumuraya, S. A. Batcheller and S. Masamune, *Angew. Chem., Int. Ed. Engl.*, 1991, **30**, 902; (g) R. S. Grev, *Adv. Organomet. Chem.*, 1991, **33**, 125; (h) M. Weidenbruch, *Coord. Chem. Rev.*, 1994, **130**, 275; (i) R. Okazaki and R. West, *Adv. Organomet. Chem.*, 1996, **39**, 231; (j) T. Iwamoto, H. Sakurai and M. Kira, *Bull. Chem. Soc. Jpn.*, 1998, **71**, 2741; (k) P. P. Power, *J. Chem. Soc., Dalton Trans.*, 1998, 2939; (l) P. P. Power, *Chem. Rev.*, 1999, **99**, 3463; (m) M. Kira and T. Iwamoto, *J. Organomet. Chem.*, 2000, **610**, 236; (n) P. P. Gaspar and R. West, in *The Chemistry of Organic Silicon Compounds*, ed. Z. Rappoport and Y. Apeloig, Wiley, New York, 2nd edn, 1999, part 3, pp. 2463–2568; (o) M. Haaf, T. A. Schmedake and R. West, *Acc. Chem. Res.*, 2000, **33**, 704; (p) B. Gehrhus and M. F. Lappert, *J. Organomet. Chem.*, 2001, **607–608**, 209; (q) N. J. Hill and R. West, *J. Organomet. Chem.*, 2004, **689**, 4165; (r) M. Kira, *J. Organomet. Chem.*, 2004, **689**, 4475; (s) O. Kuhl, *Coord. Chem. Rev.*, 2004, **248**, 411; (t) R. W. Alder, M. E. Blake, L. Chaker, J. N. Harvey, F. Paolini and J. Schutz, *Angew. Chem., Int. Ed.*, 2004, **43**, 5896; (u) B. Gehrhus, P. B. Hitchcock, R. Pongtavornpinyo and L. Zhang, *J. Chem. Soc., Dalton Trans.*, 2006, **15**, 1847.
- 3 (a) M. J. Michalczyk, R. West and J. Michl, *Organometallics*, 1985, **4**, 826; (b) B. D. Shepherd, D. R. Powell and R. West, *Organometallics*, 1989, **8**, 2664.
- 4 (a) M. Kira, T. Iwamoto and C. Kabuto, *J. Am. Chem. Soc.*, 1996, **118**, 10303; (b) T. Iwamoto, M. Tamura, C. Kabuto and M. Kira, *Organometallics*, 2003, **22**, 2342; (c) T. Iwamoto and M. Kira, *Chem. Lett.*, 1998, 277.
- 5 For instance, see: F. Bernardi, M. Olivucci and M. A. Robb, *J. Photochem. Photobiol., A*, 1997, **105**, 365.
- 6 J. J. W. McDouall, K. Peasley and M. A. Robb, *Chem. Phys. Lett.*, 1988, **148**, 183.
- 7 For the sake of simplicity, the SiH₃ substituent groups, instead of *t*-Bu₂MeSi groups, have been used in the present study. Of course, if one uses the above bulky group in the present theoretical study, the relative energies of CI and TS may change. Nevertheless, the basic concepts in these studies are still the same.
- 8 For reviews, see: (a) F. Bernardi, M. Olivucci and M. A. Robb, *Isr. J. Chem.*, 1993, 265; (b) M. Klessinger, *Angew. Chem., Int. Ed. Engl.*, 1995, **34**, 549; (c) F. Bernardi, M. Olivucci and M. A. Robb, *Chem. Soc. Rev.*, 1996, 321; (d) M. Klessinger, *Pure Appl. Chem.*, 1997, **69**, 773; (e) M. Klessinger and J. Michl, in *Excited States and Photochemistry of Organic Molecules*, VCH Publishers, New York, 1995.
- 9 M. J. Bearpark, M. A. Robb and H. B. Schlegel, *Chem. Phys. Lett.*, 1994, **223**, 269.
- 10 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci and G. A. Petersson, *et al.*, Gaussian, Inc., Wallingford CT, 2013.
- 11 Recently, Matsuo, Tamao, and co-workers experimentally reported an interestingly substituted disilene molecule, (Z)-1,2-di(1-phenyl)disilene, whose first excited state is theoretically estimated to be singlet. See: M. Kobayashi, N. Hayakawa, T. Matsuo, B. Li, T. Fukunaga, D. Hashizume, H. Fueno, K. Tanaka and K. Tamao, *J. Am. Chem. Soc.*, 2016, **138**, 758.
- 12 For discussion of excited singlet and triplet states of disilene, for instance, see a comprehensive study:



- G. A. Oyedepo and A. K. Wilson, *J. Phys. Chem. A*, 2010, **114**, 8806.
- 13 The MO representation of silylene at the singlet excited state can be found in a recent work. See: M.-D. Su, *Chem.-Eur. J.*, 2014, **20**, 9419.
- 14 The fixed geometrical parameters in (A) **Si-Rea** and (B) **C-Rea** are given in ESI.†
- 15 The reason that the Franck–Condon energy for **Si-Rea-S₀** (67.8 kcal mol^{−1}) is much less than that for **C-Rea-S₀** (116.8

kcal mol^{−1}), as already shown in Fig. 2 and 3 is easily understood with reference to the simple molecular systems, *i.e.*, ethylene (H₂C=CH₂) and silene (H₂Si=SiH₂). The energy gap for the HOMO (π) and the LUMO (π*) for ethylene is much larger than that for silene. Silene also has a higher HOMO (π) and a lower LUMO (π*) than ethylene. For details, see ref. 2.

