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Synthesis and characterisation of cerium(IV)-incorporated hydrous iron(III) oxide as an adsorbent for fluoride removal from water†

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Surface-altered hydrous iron(III) oxide incorporating cerium(IV) (CIHFO) was prepared and characterised via modern analytical tools for applications in fluoride removal from groundwater. The material with a Fe : Ce ratio of 1.0 : 0.5 (mol : mol) calcined at 473 K shows $24.8 \pm 0.5 \text{ mg F}^- \text{ g}^{-1}$ adsorption capacity at pH 5.0–7.0 from a solution with a concentration of 15.0 mg L^{-1} ; the material was established to be microcrystalline ($\sim 5 \text{ nm}$) with a $140.711 \text{ m}^2 \text{ g}^{-1}$ surface area, irregular surface morphology and porous structure. The time-dependent fluoride adsorption capacities of CIHFO at 293, 303 and 313 K are well described by the pseudo-first order, pseudo-second order and Weber–Morris kinetic models, respectively. The adsorption reaction occurs via a film/boundary layer diffusion process. The very low Arrhenius activation energy ($E_a = 0.026 \text{ kJ mol}^{-1}$) indicates the high feasibility of fluoride adsorption over CIHFO. The equilibrium data fit better with the Freundlich and Redlich–Peterson ($g < 1.0$) isotherms than with the Langmuir isotherm, which suggests multilayer adsorption. The values of the Freundlich parameters, $n = 3.10, 4.47$ and 7.57 and $K_F = 8.58, 10.88$ and 11.25 at 293, 303 and 313 K, respectively, indicate high affinity for fluoride. Thermodynamic analysis of the reaction equilibrium shows that the reaction is highly exothermic ($\Delta H^0 = -25.924$ and $-36.279 \text{ kJ mol}^{-1}$ for $C_i = 25.0$ and 35.0 mg L^{-1}), whereas the negative ΔG^0 values indicate the spontaneous nature of the reaction. The fluoride adsorption over CIHFO occurs via ion-exchange that progresses to chemisorption. The presence of sulphate shows an adverse influence on fluoride adsorption by CIHFO, and the fluoride level of 2.4 g per L groundwater (9.05 mg F L^{-1}) can be reduced below the permissible value.

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1. Introduction

The generic term ‘fluorosis’ is closely associated with elevated levels of fluoride concentration in drinking water. A report called ‘Second thoughts about fluoride’, which was published in Scientific American in 2008, demonstrated the severity of fluoride poisoning.¹ Increased fluoride in groundwater is, therefore, recognised as a major threat to human health by the European Union (EU), World Health Organisation (WHO) and Environmental Protection Agency (EPA). Fluoride contamination is considered a top priority to ensure the quality of drinking water after microbial pollution.² The beneficial and detrimental effects of fluoride to human health are highly dependent upon

its concentration in drinking water and the volume of water consumed daily. The unusual bio-chemical behaviour of fluoride has necessitated restrictions on the use of groundwater for drinking without prior removal (treatment) of fluoride. At present, 32 countries worldwide report endemic levels of fluorosis.³ This problem has also been reported in India for the past few decades and is now considered a natural crisis. The supply of treated surface water to entire fluoride-prone areas is challenging, as it requires the investment of huge funds for technological shifts, manpower, operation and maintenance. Thus, it can be safely stated that access to safe drinking water is correlated with higher income levels, whereas poor rural populations are forced to consume more fluoride-contaminated water. The simplest and the cheapest way to ensure safe water is through cost-effective treatment technology. Defluoridation of excess fluoride (above 1.5 mg L^{-1}) from fluoride-rich groundwater is the only pragmatic option to overcome this problem. The adsorption technique is often considered to be the most efficient and commercially viable method for fluoride removal³ from drinking water compared to other conventional methods. The criteria to select a suitable adsorbent include adsorption capacity, cost of material, ease of operation, and the

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possibility of regeneration and reuse. Bimetallic and poly-metallic nanocrystallites have drawn attention in recent decades for their efficacy in this process. Incorporation of a metal ion inside the lattice of another metal oxide results in structural modification, lattice defects, and size gradation of the metallic product. This rearrangement increases the porosity, surface area, and number of adsorption sites with oxygen and hydroxyl groups, which in turn triggers dynamic adsorption. Iron(III) oxide is one of the most effective natural water purifiers; it is abundantly available in the earth's crust.⁴ Hydrous iron(III) oxide, however, shows relatively low adsorption affinity for hard fluoride ions in aqueous media.⁵ Numerous bimetal hydrous oxides of La(III), Ce(IV), Ti(IV) and Zr(IV) with Fe(III) or Al(III) have shown relatively high binding affinities with fluoride.^{4–18} Consequently, different researchers have prepared binary and ternary composites of metal oxides containing Fe(III) and have investigated their efficiencies in fluoride removal.^{4–6,16–31} As Ce(IV) is small in size and has a high positive charge (high ionic potential), it has a strong binding affinity with fluoride. Therefore, the surface of hydrous CeO₂ should have greater adsorption affinity for fluoride. Another unique feature of cerium oxide is its highly facile transformation between the +4 and +3 oxidation states.⁶ However, the use of hydrous CeO₂ in defluoridation technology would not be cost effective for rural populations of developing countries, such as India. Thus, it would be beneficial to alter the surface of hydrous Fe(III) oxide by incorporating different proportions of Ce(IV) oxide to upgrade the fluoride removal efficiency of pure Fe(III) oxide.

In this work, an attempt has been made to tune the synthesis of Ce(IV)-incorporated hydrous Fe(III) oxide (CIHFO) and apply it for fluoride adsorption from water at neutral pH. The optimised material has been characterized by modern analysis techniques, viz. powder X-ray diffraction (XRD), X-ray fluorescence (XRF), Brunauer–Emmett–Teller (BET) analysis, atomic force microscopy (AFM), transmission electron microscopy (TEM), scanning electron microscopy (SEM), optical microscopy (OM), Fourier transform infrared (FTIR) spectroscopy, thermogravimetry analysis (TGA) and differential thermal analysis (DTA). Finally, the adsorption of fluoride on the optimised material (Fe : Ce = 1.0 : 0.5) was further investigated systematically. Thus, this article reports the synthesis, characterisation and fluoride removal efficacy of this promising newly developed material in aqueous medium.

2. Materials and methods

2.1. Reagents

The chemicals used were ceric ammonium nitrate ((NH₄)₂–[Ce(NO₃)₆]), ferric chloride (FeCl₃), sodium fluoride (NaF), sodium hydroxide (NaOH), hydrochloric acid (HCl), nitric acid (HNO₃), liquid ammonia and ammonium carbonate [(NH₄)₂CO₃]. All chemicals and reagents used in this work were of guaranteed reagent (G.R.) grade (Merck India, Mumbai).

2.2. Fluoride solution

A standard stock solution of fluoride (1000 mg L^{–1}) was prepared by dissolving an appropriate amount (2.21 g) of

sodium fluoride in one litre of de-ionised water. The required concentrations of fluoride solution were achieved by appropriate dilution of the stock solution with fluoride-free double de-ionized water.

2.3. Fluoride analysis

The fluoride samples were analysed using an ion-selective electrode (Hach, USA). Here, 25 mL of filtrate was placed in a polythene (PE) beaker, and fluoride adjustment buffer powder (Hach) was dissolved to adjust the pH from 5.0 to 7.8. The calibrated ion-selective electrode was then dipped in a sample solution to determine the fluoride concentration.

2.4. Ce(IV)-incorporated hydrous Fe(III) oxide (CIHFO) preparation

The method of preparation of hydrous cerium(IV)-incorporated iron(III) oxide adopted here is different from previously described methods.^{6,18} Here, a batch of cerium incorporated hydrous Fe(III) oxide was synthesised by co-precipitation in an ambient laboratory environment. The concentration of ferric chloride (0.1 M) solution in 0.1 M HCl was kept constant, while the concentration of ammonium ceric nitrate in 0.1 M HNO₃ was varied. Both solutions were mixed together, and the pH of the mixture was increased to ~6.0 to 7.0 by dropwise addition of ammonia solution (mixture of 25% ammonia solution and 0.1 M ammonium carbonate solution). The brownish-yellow gel-like precipitate was developed and was aged for 24 hours. The precipitate was filtered through a 0.45 µm membrane filter and repetitively washed with deionised water to remove extraneous contaminants. The residue was dried at 75 °C (±5 °C) in an air oven. The hot mass immediately after drying was treated with cold water in order to break it into small particles. The air-dried particles were ground and sieved to collect the particles with sizes between 50 and 100 mesh (particle size: 140 to 290 µm), which were used to conduct the experiments.

2.5. Instruments

Powder X-ray diffraction (XRD), X-ray fluorescence (XRF), optical microscopy (OM), scanning electron microscopy (SEM), transmission electron microscopy (TEM), atomic force microscopy (AFM), Brunauer–Emmett–Teller (BET) nitrogen (vapour) adsorption–desorption and Fourier transform infrared (FTIR) spectroscopy were used to characterize the materials. The powder XRD patterns of the materials were recorded using an X-ray diffractometer (Bruker D8 Advance) equipped with Cu-Kα radiation ($\lambda = 1.5418 \text{ \AA}$) at a scanning speed of $0.4^\circ \text{ min}^{-1}$ between $2\theta = 10^\circ$ and 70° , operated at a 40 kV voltage and an applied potential current of 30 mA. Transmission electron microscopic (TECNAI G² 30 ST, FEI, Netherlands) images were recorded at different magnifications. The samples for TEM were suspended in 0.25% v/v vinyloc-E chloroform solution by sonication, and one drop of the suspension was casted onto a carbon film-coated copper grid and dried under an IR lamp. The grid was then mounted on a Double Tilt Specimen Holder to perform the TEM analysis. X-ray fluorescence (XRF-Panalytical Axios) of the optimised material was analysed by accurately



weighing 8.0 g of sample (~ 200 mesh), which was thoroughly mixed with 1.0 g of wax. This mixture was packed in an aluminium cup 40 mm in diameter and pressed at 200 KN. This was exposed to X-rays for analysis. Atomic force microscopic (AFM) images were taken with an Innova Atomic Force Microscope (Bruker AXS Pte Ltd) using optical beam deflection to monitor the displacement of a micro-fabricated silicon cantilever with a spring constant of 80 Nm^{-1} to visualize the topography of the mixed oxide surface. OM images were recorded using a Nikon Eclipse LV100 POL microscope. TG and DT were measured using a Perkin Elmer system; the variations in weight and heat energy during the phase transitions of the materials were measured in a variety of gaseous environments in the temperature range of 50°C to 1200°C . The surface areas of the materials were analyzed by the Brunauer–Emmett–Teller (BET) method by $\text{N}_{2(\text{vapour})}$ adsorption at 77 K using a high-speed surface analyzer (Quantachrome Nova4000e). The adsorption of $\text{N}_{2(\text{vapour})}$ at a temperature of 77 K results in an adsorption isotherm commonly referred to as the Brunauer–Emmett–Teller (BET) isotherm. Multiplying the number of gas molecules required to cover an adsorbent surface with a monolayer of adsorbed molecules with the cross-sectional area of an adsorbate material gives the approximate sample surface area. The pH_{PZC} values of the samples were analysed by the pH metric method used by Babić *et al.*³²

2.6. Experimental design

Batch adsorption is a systematic static process to evaluate the removal efficacy of an adsorbent. Here, a measured volume of fluoride solution (100 mL) with a desired initial concentration (C_i , mg L^{-1}) was placed in a 250 mL PE bottle along with a definite mass (0.050 g) of CIHFO (except for the dose variation experiments) and agitated at $250 (\pm 5)$ rpm inside a thermostatic shaker for 2 hours (except for the time-dependent experiments). The solid mass was immediately separated by filtration, and the filtrate was analysed to determine the residual concentration of fluoride (C_f , mg L^{-1}). As the material was prepared with the aim of treating groundwater, the batch experimental program was conducted in the neutral pH range, except for the pH influence experiments.

The composition of CIHFO was optimised by preparing samples with varying mole proportions of Ce in Fe. To optimise the drying temperature, the optimized CIHFO samples were separately dried for five hours at 100°C (373 K) to 600°C (873 K) with increments of 100°C ; these samples were employed for the adsorption experiments described above.

The effects of the solution pH on fluoride adsorption were investigated independently at pre-programmed pH_i values from 3.0 to 10.0 by unit increments, employing two different fluoride solutions (C_i : 10.0 and 15.0 mg L^{-1}). The pH value of each test solution was adjusted with 0.1 M HCl or 0.1 M NaOH solution.

Time-dependent fluoride adsorption experiments with CIHFO (reaction kinetics) were conducted with three separate concentrations of fluoride solution. The same experiment was also carried out at three different temperatures (293 K, 303 K and 313 K) with a fixed fluoride concentration (15.0 mg L^{-1}).

Equilibrium experiments were carried out using fluoride solutions with C_i values ranging from 10.0 to 60.0 mg L^{-1} . The scheduled experiments were conducted separately at 293 K, 303 K, and 313 K at a solution pH_i of 7.0, an agitation time of 2 hours and a speed of 250 ± 5 rpm. The fluoride concentration remaining at equilibrium (C_e , mg L^{-1}) was determined by the ion-selective electrode, and the equilibrium capacity (Q_e , mg g^{-1}) was calculated. The thermodynamic parameters (ΔS^0 , ΔH^0 , ΔG^0) were evaluated from the results of the equilibrium analysis.

The adsorption of fluoride over tuned CIHFO may be adversely influenced by the presence of ions occurring in groundwater, such as Ca^{2+} , Mg^{2+} , HPO_4^{2-} , SO_4^{2-} , HCO_3^- , NO_3^- and Cl^- . To evaluate this interference, batch absorber experiments were conducted at $\text{pH} \sim 7.0$ as described above using the fluoride solution with $C_i = 10.0 \text{ mg L}^{-1}$ in the presence of different concentrations of competing ions.

For the desorption experiments, 0.05 g of fluoride-enriched CIHFO ($16.4 \text{ mg F}^- \text{ g}^{-1}$) was agitated at $250 (\pm 5)$ rpm for one hour along with 50 mL of NaOH solution with various concentrations in the range from 0.1 to 4.0 M.

2.7. Adsorption capacity

Immediately after each experiment, the adsorbent particles were filtered through a $0.45 \mu\text{m}$ membrane filter. The residual fluoride level in each filtrate was analysed. The amount of fluoride adsorbed by CIHFO was calculated using the following mass balance equation (eqn (1)):

$$Q_e = (C_i - C_f) \times V/M \quad (1)$$

where Q_e is the adsorption capacity (mg g^{-1}), V is the volume of the solution (L), and M is the mass of the adsorbent (g) used to carry out the batch experiments. C_i and C_f were defined previously.

2.8. Field validation

To establish the applicability of CIHFO in groundwater treatment, validation of the removal efficacy of the optimised material with field water is an essential step. To achieve this goal, water samples were collected in thoroughly cleaned PE bottles from different hand pump-attached tube wells in the Kashipur block, the most adversely fluoride-affected region of Purulia district, West Bengal, India. The fluoride concentrations of the collected water samples were analysed in the laboratory by the ion selective electrode, and the water samples with maximum fluoride concentrations (Kharai village) were used for the field validation.

3. Results and discussion

3.1. Influence of cerium content in CIHFO on fluoride adsorption

Fig. 1A shows the changes in the fluoride adsorption capacity with changing mole ratio of Fe(III) to Ce(IV) in CIHFO from 1.0 : 0.0 to 0.0 : 1.0; the results showed an increase of the



fluoride adsorption amount with increasing cerium content from 1.0 : 0.0 to 1.0 : 0.5, followed by a decrease until the Fe : Ce mole proportion reached 1 : 1. Similar observations have been reported by others.^{7,10} Although the fluoride adsorption was highest for the pure ceria material (25.3 mg g⁻¹), the fluoride adsorption amount was appreciably high compared to pure iron oxide at a Fe to Ce ratio of 1.0 : 0.5. This may be beneficial in light of the cost of the material.

Fig. 1B shows that the adsorption amount/gram of calcined CIHFO decreased with increasing calcination temperature at and above 300 °C (573 K). The adsorbent calcined at 200 °C (473 K) showed the maximum fluoride adsorption amount (24.8 mg g⁻¹) from an aqueous solution of fluoride ($C_i = 15.0 \text{ mg L}^{-1}$). As the adsorbent material was prepared by a co-precipitation method, the as-prepared CIHFO contains hydrous oxides of iron and cerium; the hydroxyl groups on these metal oxide surfaces are the primary adsorption sites for fluoride adsorption. At the temperature of 473 K, the surface water molecules attached by hydrogen bonding are removed, and the surface hydroxyl groups become available to the solute for adsorption reactions, resulting in a higher adsorption value. These hydroxyl groups were deliberately eliminated from the surfaces of the metal oxides by increasing the calcination temperature (>473 K), which is similar to other reports.^{8,10,16} As a result, the adsorption capacity of the proposed adsorbent decreased sharply.

3.2. Characterisation of CIHFO

The CIHFO material with a Fe : Ce mole ratio of 1.0 : 0.5 calcined at 200 °C (473 K) showed the highest affinity for fluoride; it was characterised by powder XRD, TEM, OM, SEM-EDX, XRF for composition, FTIR, TG and DT analyses. The powder XRD pattern (Fig. S1A†) of the CIHFO (Fe : Ce = 1.0 : 0.5) sample shows the presence of microcrystalline particles and crystallites ~5 nm in size. Comparison of the 2θ (°) values for the XRD peak positions of the newly synthesized material CIHFO with literature reports¹⁸ of Fe₂O₃ and CeO₂ [23.03°, 28.91°, 32.78°, 48.29°, 57.11° for CIHFO; 24.21° (0 1 2), 33.25° (1 0 4), 35.72° (1 1 0), 40.97° (1 1 3), 49.6° (0 2 4), 54.22° (1 1 6), 57.78° (1 2 2/0 1 8),

62.61° (2 1 4), 64.17° (3 0 0), 72.18° (1 1 9), 75.67° (2 2 0) for Fe₂O₃ and 33.0° (2 0 0), 56.3° (3 1 1), 69.43° (4 0 0), 76.69° (3 3 1), 79.04° (4 2 0) for CeO₂] showed that two peaks of CIHFO very closely matched those of CeO₂ phase with cubic structures corresponding to (*h k l*) values of (1 1 1) and (2 2 0), while one peak was common to Ce₂O₃ phase with a hexagonal structure corresponding to the (*h k l*) value (2 0 1). However, only one XRD peak of CIHFO was found to be close to α -Fe₂O₃; this clearly indicates the incorporation of Fe³⁺ ion into the crystal structure of CeO₂. Furthermore, XRF analysis to determine the composition of the sample quantified the presence of 55.082% Fe₂O₃ and 33.626% CeO₂ with some other minor constituents as impurities, namely Al₂O₃ (0.059%), SiO₂ (0.456%), CaO (0.014%), MnO (0.395%), and Cl⁻ (6.97%). The established empirical formula of the hybrid oxide is Fe_{3.72}CeO_{7.59}. Fig. 2A–C shows TEM images at low and high resolution, respectively. Fig. 2A shows the presence of agglomerated particles, and B shows particles 5 nm in size which are microcrystalline in nature. The fingerprint-like appearance in the high resolution TEM image in Fig. 2C indicates the presence of internal crystallinity in the particles. The presence of the associated crystallite particles was also prominently observed in the optical microscopic images (Fig. 2D and E). The SEM images (Fig. 2F and G) show the irregular surface morphology of the material. The SEM-EDX spectrum (Fig. 2H) also supports the presence of higher iron content than cerium, in contrast to a material reported⁶ by another research group. This is in agreement with our intention when preparing the material. From the analytical data, the empirical formula of the hybrid oxide is established as Fe_{3.72}CeO_{7.59}.

The FTIR spectra of iron(III) oxide, cerium(IV) oxide and composition-optimised CIHFO are compared in Fig. 3A. It can be seen that the stretching and bending modes of the vibrations for O–H bonds are indicated by the bands at wave numbers of (ν) 3000 to 3500 cm⁻¹ and ~1630 cm⁻¹, respectively.^{10,13} In addition, the bands recorded in the spectra at ν = 1300 to 1400 cm⁻¹ are the stretching modes of M–O bonds and those around 750 cm⁻¹ are the bending modes of M–O bonds.³¹ The FTIR spectrum for HFO (black) showed bands for Fe–O at ν = 1318 cm⁻¹ and for HCO at ν = 1396 cm⁻¹, while that for CIHFO (red)

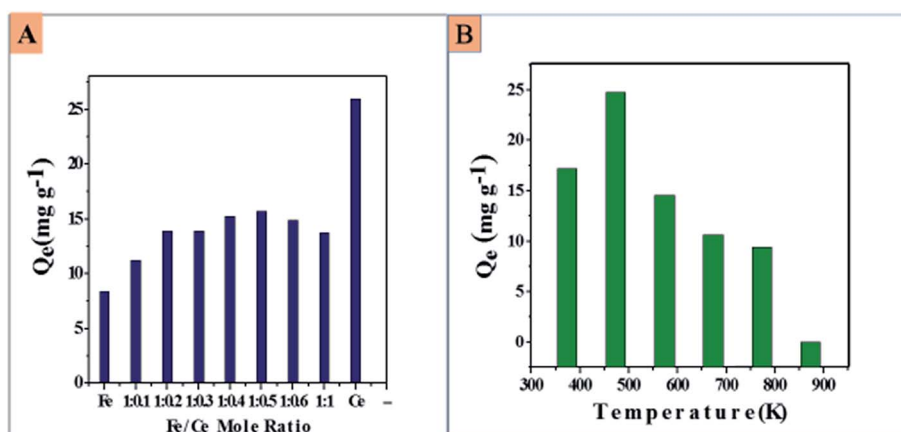


Fig. 1 Effects of variations of the (A) Fe/Ce mole ratio and (B) calcined temperature of the Fe : Ce (1.0 : 0.5) material on the fluoride adsorption capacity of CIHFO. Adsorbent dose: 0.05 g per 100 mL of fluoride solution, concentration: 15.0 mg L⁻¹, pH: 7.0 (±0.2), agitation time: 60 min.



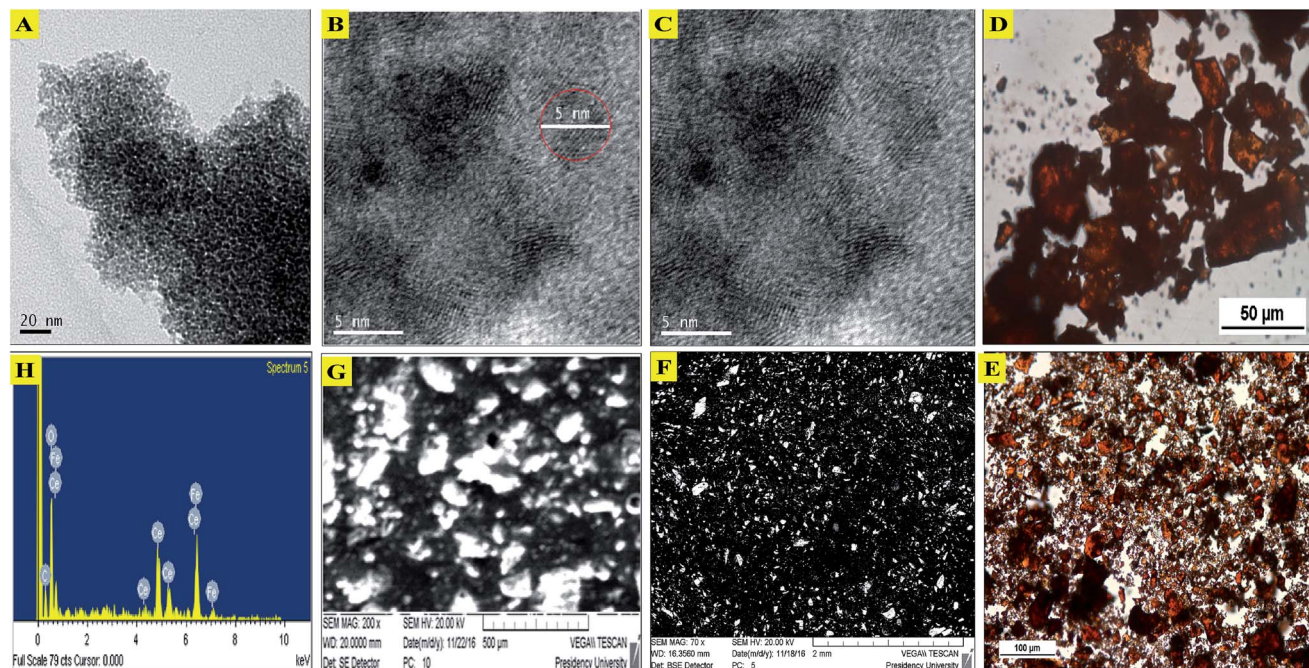


Fig. 2 Transmission electron microscopy images of CIHFO at different magnifications: (A) 20 nm, (B) 5 nm, and (C) 5 nm showing fingerprint-like impressions of internal crystallinity. Optical microscopy images of the material at (D) high resolution (50 μm scale) and (E) low resolution (100 μm scale). Scanning electron microscopy images of the material at (F) high magnification and (G) low magnification. (H) EDX spectra (a larger image is provided in the ESI (Fig. S5[†]) for clarity).

showed multiple band structures at $\nu = 1318$ and 1396 cm^{-1} ; this confirms the presence of Fe–O–Ce bonds in the material. The TG spectrum of CIHFO (Fig. 3B) recorded a maximum weight loss at around 230°C , and the percentage of this loss is about 25. This result is different from that of Yang *et al.*²¹ The weight loss was also confirmed from a sharp exothermic band in the DT spectrum (Fig. 3B). This result is presumably owing to the elimination of physically attached water molecules,

including surface-adsorbed hydrogen bonded and lattice water molecules.

Fig. 4 shows AFM images in 3D (image-A) and 2D (image-B), including the height histogram along the z-axis (C), the extracted profile (D), and the length measurement (E). It can be seen that the material has a regular surface morphology with high porosity, a pore volume of 21.4 \AA and a pore size distribution with a surface area of $140.711\text{ m}^2\text{ g}^{-1}$. Fig. S2[†]

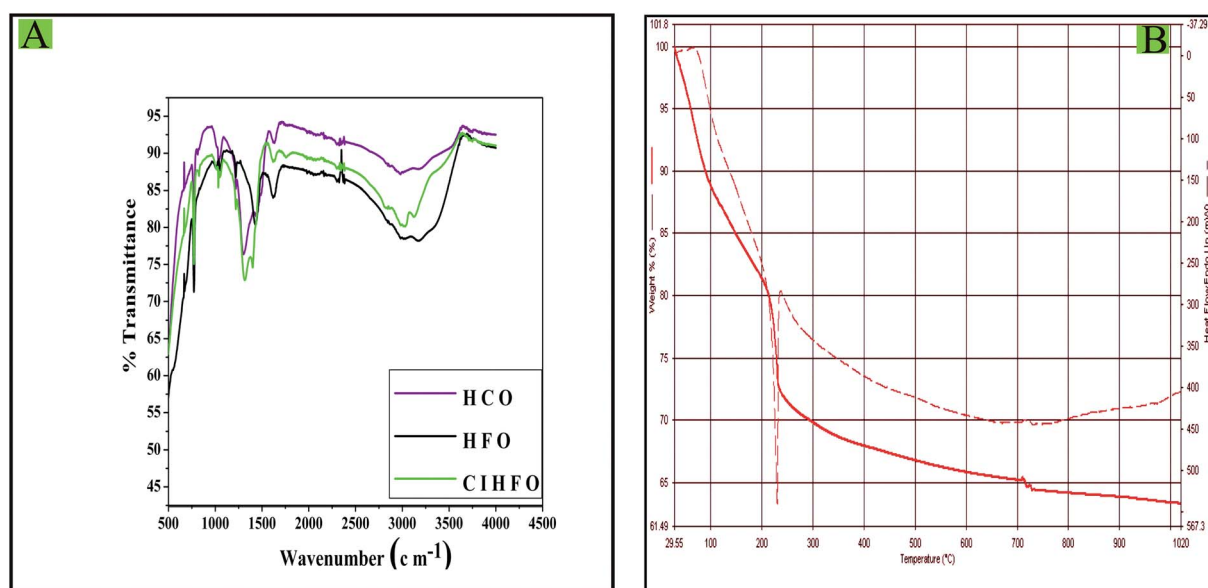


Fig. 3 (A) Fourier transform infrared spectra of HCO, HFO and CIHFO; (B) TGA and DTA spectra of CIHFO.



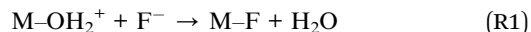
shows the plot of P/P_0 versus $1/[W(P_0/P) - 1]$; this plot shows a linear change, possibly due to the high porosity of the material.

The plot of ΔpH versus pH_i is shown in Fig. 5A; this plot optimizes the zero point surface charge pH (pH_{ZPC})³² of the material, where a transition of the surface charge occurs from positive to negative *via* the electrically neutral surface at pH 5.2.

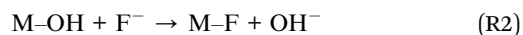
3.3. Effect of pH

Fluoride adsorption over an oxide surface is initially influenced by the pH of the solution (pH_i) and the pH_{ZPC} of the material. In general, anion adsorption is believed to be associated with exchange of hydroxyl ions (OH^-) or coulombic forces; hence, adsorption of fluoride over the surface of CIHFO occurs preferentially at a lower pH range. The influence of pH_i upon the fluoride adsorption process over the CIHFO surface is shown in Fig. 5B. It can be seen from the results that the fluoride adsorption capacity (Q_e , mg g^{-1}) decreased with increasing pH_i from 3.0 to 5.0 and remained almost the same up to pH_i 7.0. The Q_e value (mg g^{-1}) declined sharply with increasing pH_i from >7.0 to 10.0. A similar pH effect with a defluoridation process was also reported by Biswas *et al.*⁴ This indicates that the as-prepared material could be profitably used as a fluoride scavenger from water in the pH_i range from 5.0 to 7.0.

Fluoride adsorption over the prepared material (CIHFO) can be stipulated by a two-step ligand/anion exchange reaction, (R_1).



At strongly acidic pH ($\text{pH}_i = 3$), the high value of Q_e is due to the electrostatic/coulombic attraction forces of fluoride to the positively charged surface of the material. The decrease in Q_e with increasing pH_i from 3.0 to 5.0 is due to the fluctuation of the positive charge density of the solid surface. When $\text{pH} = \text{pH}_{\text{ZPC}}$ (5.2 ± 0.2), the solid oxide surfaces are electrically neutral; this is attributed to the equal distribution of positive and negative charges. Fluoride adsorption in these conditions can be simply explained by reaction (R_2).



When $\text{pH}_i > \text{pH}_{\text{ZPC}}$ (5.2 ± 0.2), the CIHFO surface becomes negative and imparts coulombic repulsion to the fluoride ions. Hence, the value of Q_e (mg g^{-1}) started to decline at $\text{pH}_i > 7.0$; this can be demonstrated by (R_3). This situation is in good agreement with the observed noticeable decrease of the equilibrium solution pH.

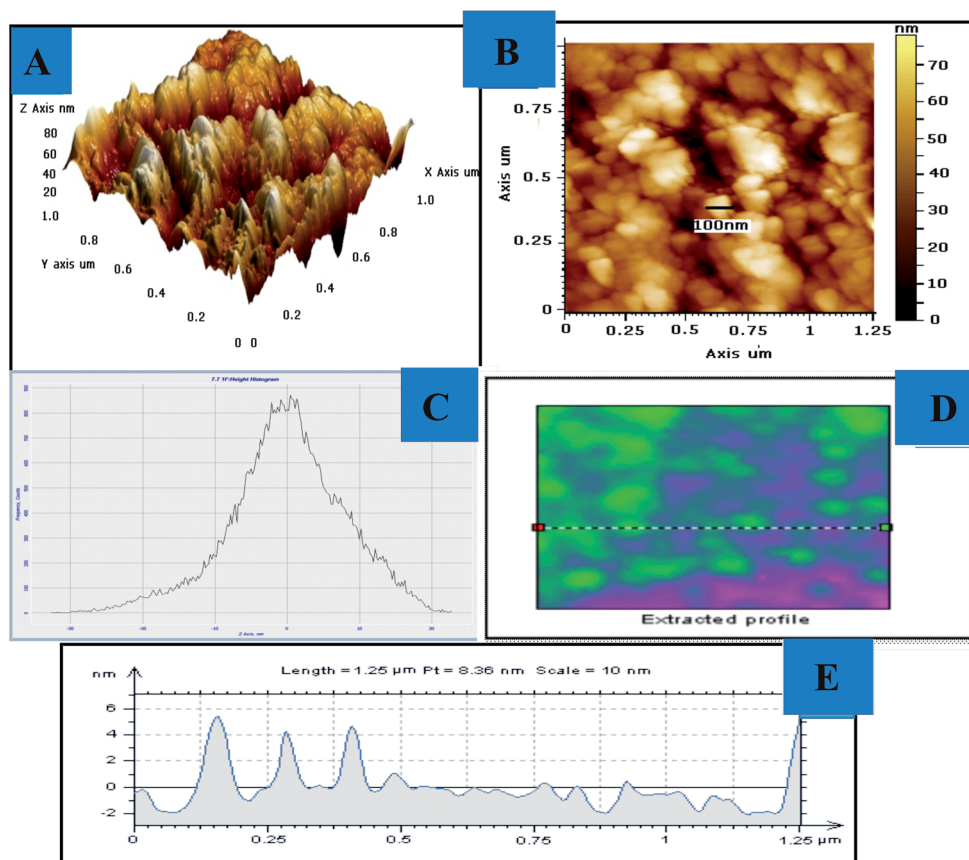
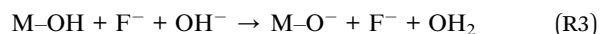


Fig. 4 Atomic force microscopy images of CIHFO: (A) 3-D view, (B) 2-D view, (C) height histogram along the Z-axis, and (D) extracted profile, (E) length measurement.



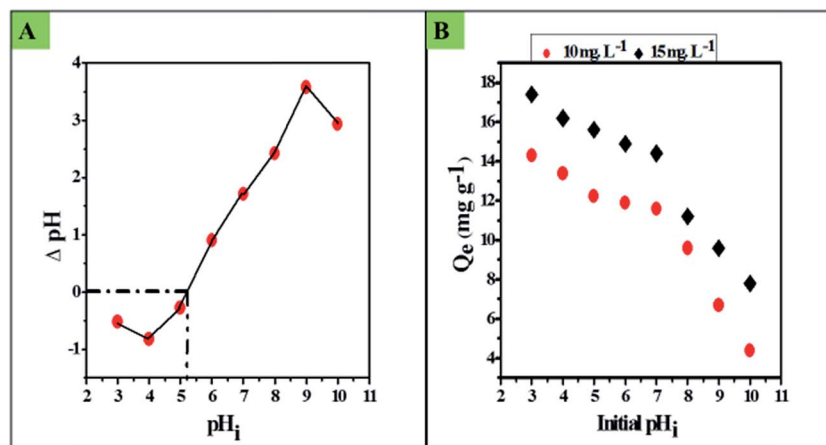


Fig. 5 (A) Plot of ΔpH versus pH_i for estimation of the pH_{ZPC} of CIHFO. (B) Effect of the initial solution pH on the adsorption of fluoride over the surface of CIHFO at 303 K with two different fluoride solution concentrations ($C_i = 10$ and 15 mg L^{-1}).

Hence, in a lower pH range, especially below pH_{ZPC} , the dominating positive charge over the solid surface triggers quick adsorption of the fluoride ions from the solution onto the solid surface due to coulombic forces; this elucidates the increase in the adsorption capacity. It is obvious that CIHFO may be an efficient scavenger of fluoride; this is in good agreement with results of other researchers,⁴ who also observed enhanced adsorption of fluoride with decreasing pH.

3.4. Effect of CIHFO dose

To determine the effects of the adsorbent dose on fluoride removal, the adsorbent dose was varied from 0.05 to 2.0 g per 50 mL fluoride solution with $C_i = 10.0$, 15.0 or 20.0 mg L^{-1} , respectively, for the adsorption reaction at 30°C (303 K) and at $\text{pH}_i \sim 7.0$. It was noticed that the fluoride removal capacity (Q_e , mg g^{-1}) decreased for a given C_i value with increasing adsorbent dose (Fig. 6). This phenomenon can be elucidated by the surface site heterogeneity model. As per this model, surface sites are composed of a series of spectra of binding energies. When the dose is applied in minute quantity, the availability of exposed surface sites increases the adsorption capacity (Q_e) value. This condition reversed as the adsorbent dose was increased, which is similar to other studies.^{4,15,23,25}

3.5. Kinetic modelling

The time-dependent adsorption capacity values, Q_t (mg g^{-1}), with increasing reaction time are presented in Fig. 7A–F. Here, A to C represent the Q_t data, respectively, for the adsorption of fluoride by CIHFO at reaction temperatures of 20°C (293), 30°C (303) and 40°C (313 K); (D) to (F) represent the Q_t data obtained for fluoride solutions with $C_i = 10.0$, 15.0 and 20.0 mg L^{-1} at 303 K and $\text{pH} \sim 7.0$. Inspection of the data revealed that more than 80% of the loaded fluoride transferred to the material surface from the solution in $t = 0$ to 40 minutes; the remaining 15% to 20% of the fluoride in solution required more time to transfer onto the solid surface. This later slow solute transfer may be due to coulombic and physical hindrance from the fluoride already

transferred to the solid surface from the solution as well as decreases of the available solute concentration and the adsorption sites of the adsorbent.^{15,23,25}

The kinetic data shown in Fig. 7 were analysed by the commonly used pseudo-first order (eqn (2)), pseudo-second order (eqn (3)), and intra-particle diffusion (eqn (4)) model equations;^{4,6,33} these well-known theoretical concepts are generally applied to understand the mechanism of fluoride adsorption on an adsorbent.

$$Q_t = Q_e[1 - e^{-k_1 t}] \quad (2)$$

$$Q_t = t \times k_2 \times Q_e^2 / [1 + (t \times k_2 \times Q_e)] \quad (3)$$

$$Q_t = K_{\text{id}} t^{0.5} + C \quad (4)$$

where Q_e (mg g^{-1}) and Q_t (mg g^{-1}) are the amounts of fluoride adsorbed by CIHFO at equilibrium at any time t (min); k_1 (min^{-1}), k_2 ($\text{g mg}^{-1} \text{ min}^{-1}$) and K_{id} represent the pseudo-first order, pseudo-second order and intra-particle diffusion rate constants, respectively; and C represents the boundary layer thickness. The non-linear plots of the kinetic data with pseudo-first order (dotted line) and pseudo-second order (bold line) model fits are also shown in Fig. 7A–F. The parameters of the kinetic model eqn (2) and (3) as estimated from the plots are demonstrated in Table 1. Considering the values of the regression coefficient (R^2) and the statistical error chi-square (χ^2), it can be observed that fluoride adsorption over the CIHFO surface occurred almost in accordance with both the pseudo-first and pseudo-second order models.^{29–31} The $Q_{t(\text{exp})}$ values were found to be very close to the modelled Q_t value from the pseudo-second order kinetics model at any particular time and temperature. It can be assumed from these results that the fluoride adsorption process can be well fitted with the pseudo-second order model; hence, the adsorption process proceeds by a chemisorption mechanism.^{11,31}

An increase of the rate of a reaction with temperature is associated with the availability of more fractions of adsorbent species with the threshold energy required for effective



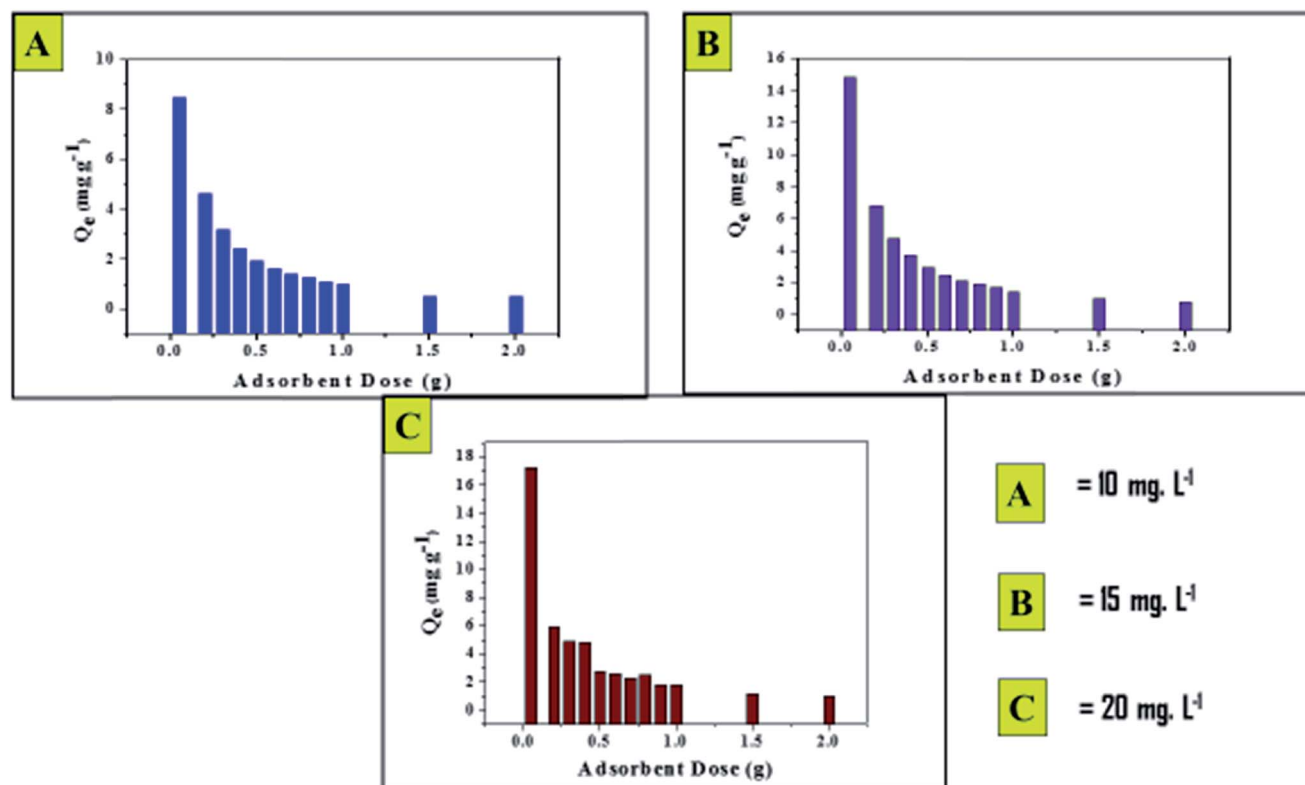


Fig. 6 Effects of adsorbent dose on fluoride removal at three different concentrations at pH 7.0 (± 2) and a temperature of 303 K (± 2 °C) with a contact time of 120 min.

molecular collision.³⁴ It was observed that the value of Q_t decreased with increasing reaction temperature. The decrease of the value of Q_t with increasing temperature indicates the exothermic nature of the adsorption reaction. In the case of increasing C_i , the rate change showed a similar result; however, the value of Q_t decreased with decreasing value of C_i . This is presumably due to the increased availability of solute loading sites per unit of the active sites.

As per the intra-particle diffusion model, the plot of Q_t against $t^{0.5}$ (Weber–Morris plot)³³ should be a straight line with a non-zero intercept value or a combination of two or three straight lines including the origin as a point. However, the present Q_t data acquired under our experimental conditions and at the investigated C_i of fluoride when plotted against $t^{0.5}$ demonstrated a straight line with a good R^2 value (Fig. 7G–I), indicating that fluoride adsorption over CIHFO occurred *via* a boundary layer (film) diffusion process.³³ The boundary layer (film) diffusion process was also indicated when the temperature-dependent Q_t data were plotted against $t^{0.5}$ (see ESI, Fig. S3†).

3.6. Equilibrium analysis

To determine the adsorption efficacy of a newly developed adsorbent, adsorption isotherms are basic requirements for the commercialisation of the adsorption system. Langmuir and Freundlich are two fundamental isotherm models that can conveniently depict the surface properties and adsorption

mechanism of such a system. Fig. 8A–C presents the equilibrium data for the adsorption of fluoride on CIHFO at pH ~ 7.0 and at temperatures of 293 (± 2) K, 303 (± 2) K, and 313 (± 2) K, respectively. We examined the adherence of the equilibrium data to the most widely accepted isotherm models, namely Langmuir³⁵ (bold line) (eqn (5)), Freundlich³⁶ (dotted line) (eqn (6)) and Redlich–Peterson³⁷ (inset) (eqn (7)) (Fig. 8A–C). The Langmuir model assumes that each molecule retains the same adsorption energy, there is no interaction between the molecules and all sites have equal affinities towards the adsorbate; hence, this isotherm denotes homogeneous adsorption. The Freundlich model is associated with non-ideal multilayer adsorption and heterogeneity of the adsorption sites. The non-linear expressions of the Langmuir, Freundlich and Redlich–Peterson eqn (5)–(7) can be set as

$$Q_e = q_m \times K_L \times C_e (1 + K_L \times C_e) \quad (5)$$

$$Q_e = K_F \times C_e^{1/n} \quad (6)$$

$$Q_e = \alpha C_e / 1 + \beta C_e^g \quad (7)$$

where C_e (mg L^{-1}) and Q_e (mg g^{-1}) are the equilibrium fluoride concentration and equilibrium adsorption capacity, respectively. The q_m value (mg g^{-1}) is the maximum monolayer adsorption capacity (mg g^{-1}), and K_L (L mg^{-1}) is the equilibrium constant related to the net enthalpy change of the reaction. K_F is the Freundlich adsorption capacity, and $1/n$ is an



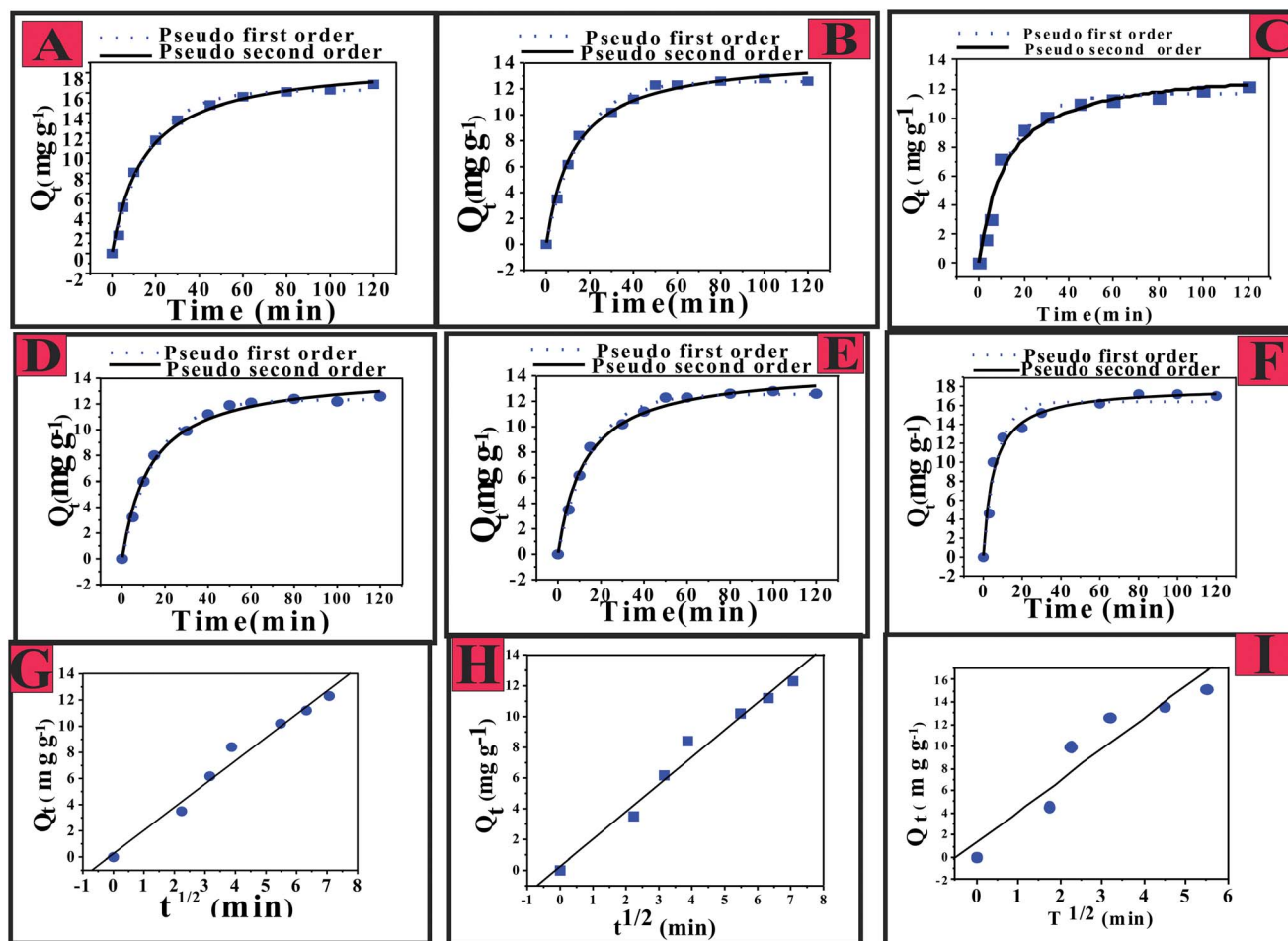


Fig. 7 The plots of fluoride adsorption capacity at any time and Q_t (mg g⁻¹) versus time, t (min) including non-linear pseudo-first order (---) and pseudo-second order (—) plots at three different temperatures, (A) 293 K, (B) 303 K and (C) 313 K, with an active fluoride concentration of 10 mg L⁻¹ (ionic strength = 1 M) at pH 7.0 and also with three fluoride solutions with different concentrations, (D) $C_i = 10.0$ mg L⁻¹, (E) $C_i = 15.0$ mg L⁻¹ and (F) $C_i = 20.0$ mg L⁻¹, at a temperature of 303 K (ionic strength = 1 M) and pH 7.0. The plots of Q_t (mg g⁻¹) versus $t^{1/2}$ (min) for three different fluoride concentrations, (G) 10.0 mg L⁻¹, (H) 15.0 mg L⁻¹ and (I) 20.0 mg L⁻¹, at a temperature of 303 K (ionic strength = 1 M) and pH 7.0.

Table 1 Estimated kinetic parameters for fluoride adsorption reactions with ClHFO at pH = 7.0 (±0.2) at 293, 303 and 313 K

Kinetics equation	Parameters	Fluoride concentration (mg L ⁻¹)			Temperature (K)		
		10	15	20	293	303	313
Pseudo-second order	k_2 (g mg ⁻¹ min ⁻¹)	0.005	0.007	0.010	0.003	0.005	0.006
	q_e (mg g ⁻¹)	14.40	16.37	17.98	19.20	14.61	13.48
	R^2	0.99	0.99	0.97	0.99	0.99	0.97
	χ^2	0.17	0.31	0.82	0.32	0.17	0.47
	χ^2	0.06	0.10	0.14	0.05	0.06	0.07
Pseudo-first order	k_1 (min ⁻¹)	0.06	0.10	0.14	0.05	0.06	0.07
	q_e (mg g ⁻¹)	12.34	14.70	16.3	16.28	12.56	11.63
	R^2	0.99	0.99	0.96	0.99	0.99	0.986
	χ^2	0.08	0.27	1.24	0.23	0.11	0.30
	χ^2	0.08	0.27	1.24	0.23	0.11	0.30
Intra-particle diffusion	K_{id} (g mg ⁻¹ min ^{0.5})	1.77	2.54	2.82	2.44	1.77	1.84
	R^2	0.96	0.94	0.88	0.96	0.96	0.92

empirical constant related to the adsorption intensity; α , β and g are the constants of the three-parameter Redlich–Peterson isotherm equation, which bears characteristics of both the

Langmuir and Freundlich isotherm models. (i) When the value of g moves to 1, eqn (7) converts to the Langmuir form; (ii) when g increases to below unity but is greater than zero, eqn (7)



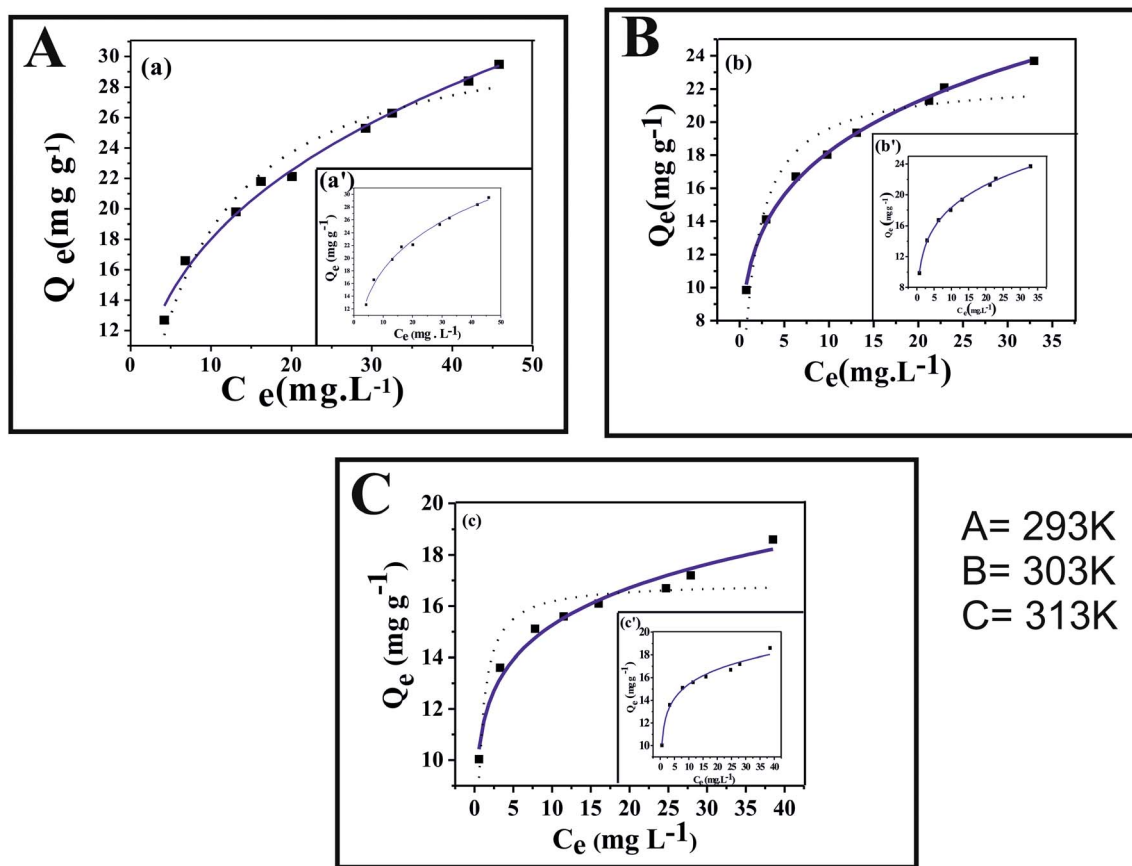


Fig. 8 (A) Adsorption isotherm plots and model fits: (a) Freundlich, Langmuir and (a') Redlich–Peterson (inset) at 293 K; (B) adsorption isotherm plots and model fits: (b) Freundlich, Langmuir and (b') Redlich–Peterson (inset) at 303 K; (C) adsorption isotherm plots and model fits: (c) Freundlich, Langmuir and (c') Redlich–Peterson fitting (inset) at 313 K.

converts to the Freundlich form; and (ii) when g moves to zero, eqn (7) converts to Henry's form.³⁷

The isotherm parameters generated from the plots of the equilibrium data on the origin using spreadsheet software are presented in Table 2. From the values of the regression coefficient (R^2) and the statistical error chi-square (χ^2) of non-linear analysis, it can be seen that the equilibrium data fit much better with the Freundlich model than with the Langmuir model. For further confirmation of the better Freundlich fit, the equilibrium data was also analysed with the Redlich–Peterson isotherm model (inset, Fig. 8); the as-estimated model parameters are also given in Table 2. Here, the value of g was found to be somewhat away from unity, confirming the greater conformity of the equilibrium data with the Freundlich isotherm.³⁷

The evaluated degree of the Freundlich coefficient ($1/n$) falls between 0 and 1, as the values of n obtained are 3.10, 4.47 and 7.57, respectively for the temperatures (293 K, 303 K and 313 K) applied in the equilibrium reactions. Again, the high value of n validated that the fluoride uptake by CIHFO followed the Freundlich isotherm more than the Langmuir isotherm.¹³ Thus, it can be said that the adsorption sites of this mixed metal oxide are heterogeneous, and the multilayer adsorption mechanism is suitable for the fluoride adsorption.

The Dubinin–Radushkevich (D–R) equation (eqn (8))^{38,39} was used to evaluate the free energy of the reaction; it can be

obtained from the equilibrium isotherm data to predict whether the nature of adsorption is physical, ion-exchange or chemical.

$$\ln Q_e = \ln Q_m - \beta \varepsilon^2 \quad (8)$$

where Q_e and Q_m are the equilibrium and maximum D–R adsorption capacities (mol kg^{-1}), respectively, whereas β signifies the activity coefficient that is related to the mean free energy ($\text{mol}^2 \text{kJ}^{-2}$) of the reaction. As eqn (8) represents the equation of a straight line, the plot of $\ln Q_e$ against ε^2 becomes a straight line, and the values of the slope and intercept of the plot (Fig. S4†) were used to calculate the values of β and Q_m (Table 2).

ε is the Polanyi potential, which can be extracted from eqn (9) below:

$$\varepsilon = RT \ln(1 + 1/C_e) \quad (9)$$

where R is the molar gas constant ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$) and T is the absolute temperature (K). The mean free energy change is basically the change of free energy required to transform the adsorbate molecule from infinity to the adsorbent surface.⁴⁰ The values of E_{DR} were calculated from eqn (10) and are incorporated in Table 2.



Table 2 Isotherm parameters of fluoride adsorption reactions with CIHFO at pH = 7.0 (± 0.2) at 293, 303 and 313 K

Isotherm model	Parameters	Temperature		
		293 K	303 K	313 K
Langmuir	θ (mg g ⁻¹)	32.62	22.55	16.91
	K (L mg ⁻¹)	0.13	0.66	2.18
	χ^2	1.29	2.88	1.20
	R^2	0.96	0.88	0.84
Freundlich	n	3.10	4.47	7.57
	K_F (mg g ⁻¹)	8.58	10.88	11.25
	R^2	0.99	0.99	0.98
	χ^2	0.31	0.05	0.15
Redlich–Peterson	α	16.25	156.29	170.47
	β	1.44	13.5	14.16
	g	0.74	0.79	0.88
	R^2	0.99	0.99	0.98
Dubinin–Radushkevich	χ^2	0.28	0.03	0.13
	Q_m (mol kg ⁻¹)	3.05	2.07	1.29
	β (mol ² kJ ⁻²) $\times 10^3$	3.14	2.08	1.21
	E_{DR} (kJ mol ⁻¹)	12.61	15.50	20.32
	R^2	0.60	0.96	0.98

$$E_{DR} = 1/(2\beta)^{1/2} \quad (10)$$

It can be seen that the values of Q_m decreased with increasing temperature, which supports the previous results obtained from equilibrium modelling. From the D–R equation and according to the value of E_{DR} , an adsorption process may be one of three types:^{38–40} (i) physisorption if $E_{DR} = 8.0$ kJ mol⁻¹, (ii) ion-exchange if $E_{DR} = 8.0$ to 16.0 kJ mol⁻¹ and (iii) chemisorption if $E_{DR} > 16.0$ kJ mol⁻¹. The gradual increase of the E_{DR} value with increasing temperature from 12.61 to 20.32 kJ mol⁻¹ indicates that the adsorption of fluoride proceeds from ion-exchange to chemisorption. At 303 K, the E_{DR} obtained in the present case is 15.50 kJ mol⁻¹, which is indicated to be at the borderline of the ion-exchange and chemisorption processes.^{8,38}

3.7. Thermodynamic parameters

Changes in thermodynamic parameters such as standard free energy (ΔG^0), standard enthalpy (ΔH^0) and standard entropy (ΔS^0) for this adsorption process were estimated using the following thermodynamic relations (eqn (11) and (12)),⁴⁰ assuming the activity coefficient to be unity at low solute concentrations.

$$\Delta G^0 = \Delta H^0 - T\Delta S^0 \quad (11)$$

and

$$\Delta G^0 = -2.303RT \log K_c \quad (12)$$

From eqn (11) and (12), eqn (13) can be derived:

$$\log K_c = \Delta S^0/2.303R - (\Delta H^0/2.303R)1/T \quad (13)$$

Every term in the context represented herewith has its usual significance. Eqn (14) is adapted from eqn (13) with a slight alteration, where the thermodynamic equilibrium constant K_c is substituted by q_e/C_e :

$$\log(q_e/C_e) = \Delta S^0/2.303R - (\Delta H^0/2.303R)1/T \quad (14)$$

ΔS^0 and ΔH^0 are constants within the studied temperature range; the values of both constants were computed from the intercept and slope of the straight line of the plot of $\log(q_e/C_e)$ versus $1/T$ (Fig. 9A), and ΔG^0 at various temperatures was assessed using eqn (12). The plots showed good linearity, as is evident from the high regression coefficient value ($R^2 = 0.98$), by taking the C_i of fluoride solution as 35.0 mg L⁻¹; the results are presented in Table 3. The ΔG^0 values for fluoride adsorption by CIHFO indicate the thermodynamic feasibility/spontaneity of the reactions. The negative values of ΔG^0 (Table 3) confirm that the adsorption of fluoride by CIHFO is spontaneous and can occur in a wide range of temperatures. The highly negative value of ΔH^0 (-36.279 kJ mol⁻¹) indicates that the adsorption process is also exothermic in nature.⁴¹ This was in agreement with the results of the isotherm analysis. The entropy change (ΔS^0) was found to be -86.537 J mol⁻¹ K⁻¹, indicating the irreversibility and stability of the adsorption process. This thermodynamics pattern has also been reported by other authors.^{11,14}

3.8. Activation energy

The Arrhenius equation is applicable to evaluate the activation energy of any kind of reaction and can be expressed as eqn (15):

$$k_2 = A \exp(-E_a/RT) \quad (15)$$

Here, k_2 is the rate constant (g mg⁻¹ min⁻¹), A (g mg⁻¹ min⁻¹) is a temperature independent factor and E_a (kJ mol⁻¹) is the activation energy of a particular reaction. R and T are well known and were mentioned earlier.

The logarithm of eqn (15) illustrates a linear relationship of the Arrhenius equation and can be drawn as below eqn (16):

$$\ln k_2 = \ln A + (-E_a/R)1/T \quad (16)$$

The plot of $\ln k_2$ against $1/T$ generates a straight line (Fig. 9B); the values of E_a and A were calculated from the slope and intercept of the plot and were found to be $E_a = 0.026$ kJ mol⁻¹ and $A = 4.95$ g mg⁻¹ min⁻¹. The significantly low value of the activation energy supports the high fluoride affinity of CIHFO and the feasibility of the reaction.^{8,41}

3.9. Effects of co-occurring ions on fluoride adsorption

Among commonly occurring groundwater anions, sulphate showed adverse effects on fluoride adsorption by CIHFO. The results of the adverse influence of sulphate on fluoride adsorption from the solution of $C_i = 10.0$ mg L⁻¹ are shown in Fig. 10A. It can be seen that as the sulphate concentration increased from 10.0 to 150 mg L⁻¹, the fluoride adsorption amount (9.5 mg g⁻¹) decreased to 3.3 mg g⁻¹, which is only ~30% of the initial value. However, the fluoride adsorption amount (9.5 mg g⁻¹) was



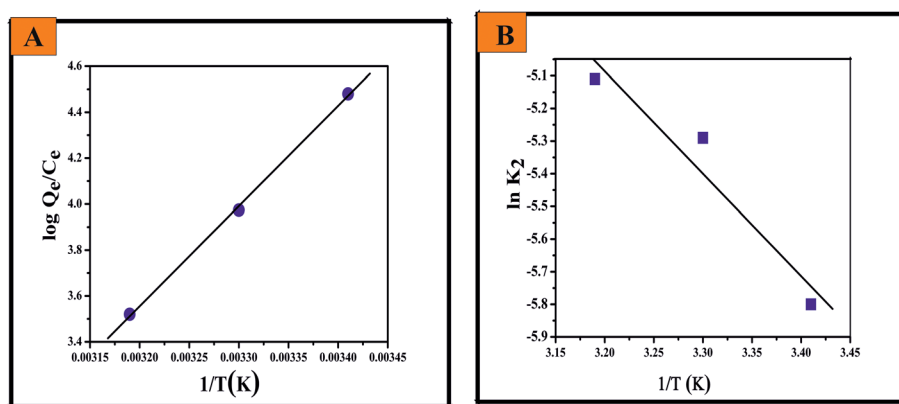


Fig. 9 The plots of (A) $\log Q_e/C_e$ versus $1/T$ (K) for the thermodynamic parameters and (B) $\ln K_2$ versus $1/T$ (K) for the activation energy of the fluoride adsorption reaction with CIHFO.

Table 3 Thermodynamic parameters for fluoride adsorption reactions with CIHFO at pH = 7.0 (± 0.2)

Fluoride concentration (mg L ⁻¹)	ΔH^0 (kJ mol ⁻¹)	ΔS^0 (J mol ⁻¹ K ⁻¹)	ΔG^0 (kJ mol ⁻¹) at different temperatures (K)		
			293	303	313
25	-25.924	-47.411	-11.941	-11.698	-10.971
35	-36.279	-86.537	-10.913	-10.011	-9.160

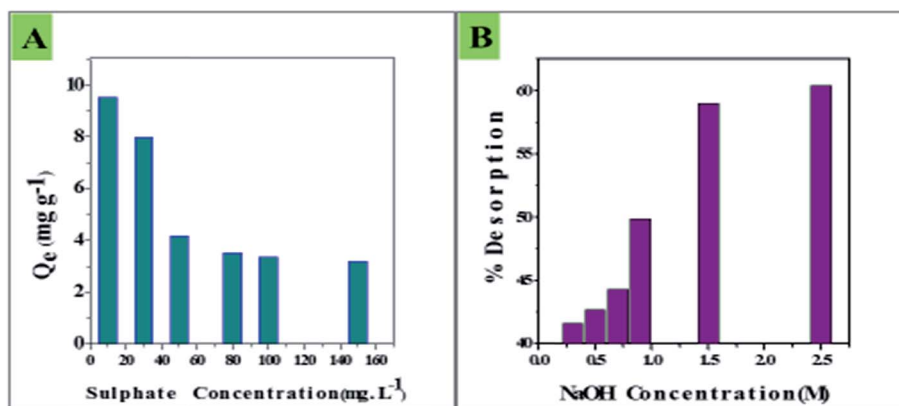


Fig. 10 Effects of (A) sulphate interference on fluoride adsorption by CIHFO and (B) molar concentration of NaOH on the desorption of fluoride from the CIHFO surface.

similar in the presence of other anions, such as phosphate, bicarbonate, nitrate and chloride. The sequence of the adverse influence of the tested ions on fluoride adsorption over CIHFO can be given as SO_4^{2-} , $\text{HCO}_3^- > \text{HPO}_4^{2-} > \text{NO}_3^- > \text{Cl}^-$. Some other studies^{12,20} support this interference trend; however, different interference trends have also been reported.^{6,8}

4. Desorption

From Fig. 10B, it is clear that with increasing basicity, the desorption percentage increased; a maximum of 61%

desorption occurred. A similar increasing trend of desorption percentage with increasing molar strength of base was reported by other researchers.^{8,11} The process is predicted to be borderline chemisorption and is exothermic in nature; therefore, it is obvious that the product is more stable than the reactant. Also, as the desorption tendency is moderate, the tendency of fluoride to leach into the environment also decreases. The fluoride-desorbed material was found to be recyclable up to a maximum of three cycles, with a drastic decrease of the fluoride absorption efficiency thereafter.



5. Fluoride removal from a natural water sample

Groundwater was collected from a hand pump-attached tube-well in Kharai village (Kashipur block) of Purulia district, West Bengal; the physico-chemical parameters of the groundwater were analysed, including pH (7.36), iron (total) (0.87 mg L^{-1}), fluoride (9.05 mg L^{-1}), calcium (8.05 mg L^{-1}), magnesium (1.99 mg L^{-1}), hardness (9.99 mg L^{-1}), TDS (1140 mg L^{-1}), salinity (517 mg L^{-1}), conductivity ($364 \mu\text{S}$) and turbidity (008 NTU). 500 mL of this water sample was stirred magnetically (500 rpm) for 2 hours with increasing CIHFO dosages from 0.1 g to 1.5 g. Immediately after the experiments, the samples were filtered and the fluoride levels were analyzed. It appeared that the fluoride level was reduced from 9.05 mg L^{-1} to below 1.0 mg L^{-1} when the dosage of CIHFO employed was 1.20 g; this is well below the WHO recommended guideline maximum concentration level. Thus, the experimentally approved CIHFO dosage is 2.40 g L^{-1} for the treatment of fluoride-rich (concentration: 9.05 mg L^{-1}) groundwater to reduce the fluoride level below 1.0 mg L^{-1} .

6. Mechanism of fluoride adsorption over CIHFO

Based on the measured pH_{zpc} of CIHFO and the pH effect on fluoride adsorption, it can be stated that under acidic pH conditions, more fluoride ions are adsorbed on the adsorbent surface due to electrostatic attractions between the negatively charged adsorbate and the positively charged adsorbent surface. As the energy of activation is low, the fluoride should be distributed very rapidly over the CIHFO surface using the minimum energy which is available from the mechanical agitation of the reaction mixture. The E_{DR} values (12.61 , 15.50 and $20.32 \text{ kJ mol}^{-1}$, respectively, at 293 , 303 and 313 K) estimated from modelling the equilibrium data with the D-R isotherm predict that the adsorption process is ion-exchange to borderline chemisorption. The obtained experimental results indicate that during the adsorption process, the univalent surface hydroxyl ions are exchanged by fluoride ions, leading to

the formation of a stable complex with a potential energy lower than that of the reactants (Fig. 11A). The large negative value of ΔH^0 and the lower tendency of fluoride desorption support this observation. Moreover, the FTIR spectrum of adsorbed CIHFO when compared with that of CHIFO (Fig. 11B) shows a significant reduction of the O–H stretching and bending band intensities, and the near disappearance¹³ of the bands at 1400 cm^{-1} and 750 cm^{-1} validates the proposed mechanism.

7. Conclusion

Surface-tuned iron(III) oxide with cerium(IV) oxide (CIHFO) is a highly porous, microcrystalline and amorphous material with a high surface area relative to that of iron(III) oxide. Microscopic images confirmed the presence of microcrystalline particles not only with nanodimensions but also with irregular surface morphologies. Exploration of this thermally stable material (473 K) for fluoride adsorption shows that the surface reaction occurs according to pseudo-first as well as pseudo-second order kinetics with a boundary layer (film) diffusion process. The very low Arrhenius energy of activation indicates the high affinity of CIHFO for fluoride. The better adherence of the equilibrium data to the Freundlich isotherm suggests the surface heterogeneity of the CIHFO multilayer adsorption phenomenon. The effects of tested ions on the fluoride adsorption reaction show no notable adverse influence, with the exception of sulphate. The FTIR analysis and other experimental results suggested that fluoride adsorption occurs through an ion-exchange mechanism which ultimately changes to chemisorption. Encouraging results were obtained when CIHFO was employed for high fluoride (9.05 mg L^{-1}) groundwater treatment for field validation.

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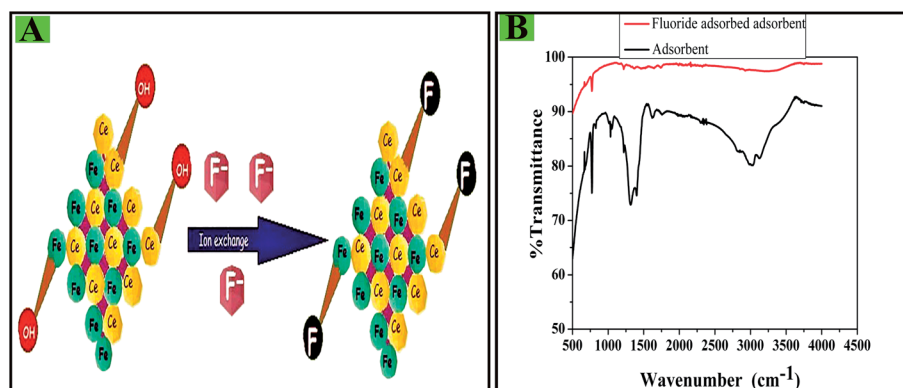


Fig. 11 (A) Schematic of the fluoride adsorption mechanism on the CIHFO surface. (B) Fourier transform infrared spectral data of the pure adsorbent and the fluoride-adsorbed adsorbent.



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