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Cation dependent charge transport in linear dicarboxylate based isotypical 1D coordination polymers†

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Two new mixed-ligand one-dimensional coordination polymers (1D CPs) [Cd(adc)(4-phpy)₂(H₂O)₂], (**1**) and [Zn(adc)(4-phpy)₂(H₂O)₂], (**2**) (H₂adc = acetylenedicarboxylic acid and 4-phpy = 4-phenylpyridine) have been synthesized and well characterized by elemental analysis, infrared spectroscopy, single crystal X-ray diffraction, powder X-ray diffraction (PXRD) and thermogravimetric analysis (TGA). Both compounds **1** and **2** are isostructural and fabricate 3D supramolecular networks by the combination of hydrogen bonding and C–H... π interactions. Interestingly, these two materials exhibit electrical conductivity and reveal Schottky barrier diode behavior. To shed light on the charge transport mechanism of the compounds, the mobility, transit time, diffusion length and density of states at a quasi Fermi level have been derived. The analysis indicates that compound **1** has higher mobility ($9.15 \times 10^{-7} \text{ m}^2 \text{ V}^{-1} \text{ s}^{-1}$) and diffusion length (1.116 μm) in comparison to compound **2** (mobility and diffusion length are $5.44 \times 10^{-7} \text{ m}^2 \text{ V}^{-1} \text{ s}^{-1}$ and 1.050 μm respectively). Compound **1**, with the larger cation and shorter H-bonding distance, shows higher electrical conductivity, which is 2.55 times greater than compound **2**.

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Introduction

Coordination polymers are inorganic–organic solid state materials containing metal ion centers or metal clusters linked by organic ligands extending in an array forming crystalline and highly ordered structures with high thermal stability.^{1–7} The design and synthesis of metal–organic coordination polymers or metal–organic frameworks (MOFs) have recently attracted immense interest, owing to their potential applications in gas storage,^{8–10} separation,^{11–14} magnetism,^{15–19} drug delivery,^{20–22} catalysis^{23–25} and sensing applications.^{26–28} Although it is a great challenge, by judicious choice of various ligands differing in lengths and functionalities, numerous MOFs with desired structures and properties have been obtained. However, utilization of MOFs in the fabrication of electronic devices is still at the age of infancy, because these materials exhibit low

electronic conductivity. Low conductivity of MOFs is a direct consequence of the engineering of these materials as they are typically constructed from hard metal ions linked by redox-inactive organic ligands, ensuing a poor conjugation pathway for charge transport.²⁹ Nevertheless, in the past few years, new approaches have been recognized for the construction of MOFs that exhibit high electrical conductivity.^{30–32}

A successful approach in the construction of thermodynamically stable coordination polymers is to utilize dicarboxylate ligands with rigid backbone that are capable of binding metal ions.^{33,34} These ligands would be good linear spacer and have been widely used due to their diverse coordination modes and bridging ability with metal ions. Recently, our group used linear dicarboxylate ligands to successfully synthesize a series of coordination polymers bonded through halogen–halogen interactions for the uptake of CO₂ gas.³⁵ Of various dicarboxylate ligands, utilization of acetylenedicarboxylate ([−]O₂CCCCO₂[−]) seems to be scarce in the literature.³⁶ Herein, we report two acetylenedicarboxylate based one-dimensional coordination polymers (1D CPs), [Cd(adc)(4-phpy)₂(H₂O)₂], (**1**) and [Zn(adc)(4-phpy)₂(H₂O)₂], (**2**) (H₂adc = acetylenedicarboxylic acid and 4-phpy = 4-phenylpyridine) assembled through weak interactions to form three-dimensional (3D) aggregates. Interestingly, these materials exhibit electrical conductivity of the order $10^{-7} \text{ S cm}^{-1}$ which is in the range of semiconductor and reveal Schottky barrier diode behavior. However, conductivity and photosensitivity of **1** are enhanced in both dark and under

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illumination because of the lesser band gap and increased charge transport due to the photon excitation. The mobility, diffusion length, density of states (DOS) at quasi Fermi level indicates the better charge transport for the device fabricated with **1**. In these materials, as the metal carboxylate chains are not expected to transport charge, the increase in conductivity of **1** is associated with increase in size of the ionic radii of the Cd^{2+} compared to Zn^{2+} .^{29,37} Besides, ionic radii and conductivity have been correlated with the simultaneous decrease in H-bonding distance among the adjacent layers in **1**. Indeed, the H-bonding distance decreases from 2.72 Å in **2** to 2.68 Å in **1**. Therefore, larger size of cation and shorter H-bonding distance among the layers lead to the higher electrical conductivity in compound **1**. Electrical conductivity of **1** is 2.55 times greater than **2**. Mobility of compound **1** is also enhanced by 68% compared to **2**.

Experimental section

Materials and physical method

The all chemicals purchased were reagent grade and were used without further purification. Elemental analysis (carbon, hydrogen and nitrogen) was performed on a Perkin-Elmer 240C elemental analyzer. Infrared spectrum in KBr (4500–500 cm^{-1}) was recorded using a Perkin-Elmer FT-IR spectrum RX1 spectrometer. Thermogravimetric analyses were recorded on a Perkin-Elmer Pyris Diamond TG/DTA in the temperature range 30–600 °C under a nitrogen atmosphere. The powder XRD data was collected on a Bruker D8 Advance X-ray diffractometer using Cu K α radiation ($\lambda = 1.548$ Å) generated at 40 kV and 40 mA. The PXRD spectrum was recorded in a 2θ range of 5–50. Solid state electronic absorption spectra were recorded on a PerkinElmer UV/VIS Spectrophotometer (LAMBDA 35). Solid state emission spectra were recorded on a HORIBA Jobin Yvon (Fluoromax-4) fluorescence spectrophotometer. The optical characterization was studied with the help of Shimadzu 2401 PC UV-vis spectrophotometer, in the range of 250–800 nm. Electrochemical impedance spectroscopy (EIS) study is done by Agilent 4295A LCR instrument.

Synthesis of compounds **1** and **2**

Synthesis of compound 1. A solution of 4-ppy (31 mg, 0.2 mmol) in MeOH (2 mL) was slowly and carefully layered to a solution of $\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ (62 mg, 0.2 mmol), in H_2O (2 mL) using 2 mL 1 : 1 (=v/v) buffer solution of MeOH and H_2O followed by layering of H_2adc (23 mg, 0.2 mmol) neutralized with Et_3N (0.027 mL, 0.2 mmol) in 2 mL EtOH. The colorless block crystals of $[\text{Cd}(\text{adc})(4\text{-ppy})_2(\text{H}_2\text{O})_2]$, **1** were obtained after three days (69 mg, yield 60%). Elemental analysis (%) calcd for $\text{C}_{26}\text{H}_{22}\text{CdN}_2\text{O}_6$: C 54.65, H 3.85, N 4.90; found: C 54.95, H 3.67, N 4.54. IR (KBr pellet, cm^{-1}): 1592 $\nu_{\text{as}}(\text{COO}^-)$, 1317 $\nu_{\text{sys}}(\text{COO}^-)$.

Compound 2. It was synthesized by a similar procedure as adopted for **1** except using $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (60 mg, 0.2 mmol). Colorless block shaped crystals $[\text{Zn}(\text{adc})(4\text{-ppy})_2(\text{H}_2\text{O})_2]$, **2** were obtained after few days. (68 mg, yield 65%). Elemental analysis (%) calcd for $\text{C}_{26}\text{H}_{22}\text{ZnN}_2\text{O}_6$: 59.40, H 4.20, N 5.35; found: C

59.57, H 4.15, N 5.56. IR (KBr pellet, cm^{-1}): 1590 $\nu_{\text{as}}(\text{COO}^-)$, 1322 $\nu_{\text{sys}}(\text{COO}^-)$.

General X-ray crystallography

Single crystal of the compound **1** and **2** having suitable dimensions, was used for data collection using a Bruker SMART APEX II diffractometer equipped with graphite-monochromated Mo K α radiation ($\lambda = 0.71073$ Å). The molecular structure was solved using the SHELX-97 package.³⁸ Non-hydrogen atoms were refined with anisotropic thermal parameters. Hydrogen atoms were placed in their geometrically idealized positions and constrained to ride on their parent atoms. The crystallographic data for **1** and **2** are summarized in Table S1† and selected bond lengths and bond angles are given in Tables S2 and S3.†

Device fabrication and characterization

To fabricate the Schottky device, at first, indium tin oxide (ITO) coated glass substrate was cleaned in isopropanol and ultrasonicated in a bath for 15 min. Then it was further cleaned in acetone and de-ionized water sequentially. This cleaned ITO coated glass was dried in vacuum chamber. At the same time well dispersed solutions of the sample **1** and **2** in the dimethyl sulfoxide (DMSO) medium were prepared and spin coated onto the dried ITO glass at 400 rpm for 1 min with the help of SCU 2700 spin coating unit. This spin coating step was repeated for 4 times. After drying in vacuum, the film thickness was measured as 10 μm by surface profiler. At the last step of fabrication, aluminum (Al) electrodes were deposited onto the film by a Vacuum Coating Unit 12A4D of HINDHIVAC under pressure 10^{-6} Torr. The area of the Al electrodes was maintained as $7.065 \times 10^{-6} \text{ m}^2$ by the shadow mask. Current–voltage measurements of the fabricated device were done by a Keithley 2400 source meter by two-probe technique under dark and illumination at room temperature.

Electrical characterization

To study the current–voltage (I – V) characteristics, two devices of Al/compound **1**/ITO (Device-a) and Al/compound **2**/ITO (Device-b) were fabricated. A bias voltage varying from –6 V to +6 V was applied to the devices at room temperature under dark and light condition. The structure of Device-a (or Device-b) is shown on Fig. S1† including the direction of light.

Results and discussion

Structural descriptions of $[\text{Cd}(\text{adc})(4\text{-ppy})_2(\text{H}_2\text{O})_2]$, (**1**) and $[\text{Zn}(\text{adc})(4\text{-ppy})_2(\text{H}_2\text{O})_2]$, (**2**)

Single crystal X-ray crystallography revealed that both **1** and **2** are isotypical. However, compound **1** crystallizes in the triclinic space group $P\bar{1}$ with $Z = 2$. Asymmetric unit in **1** contains distorted octahedral Cd(II) centre ligated by two O atoms from two adc anions in monodentate fashion (Cd–O, 2.289(3)–2.292(3) Å) and two N atoms from two 4-ppy ligands (Cd–N, 2.339(4)–2.343(5) Å) in the equatorial plane, and by O atoms from two aqua ligands at the axial sites (Cd–O, 2.301(3)–2.306(3) Å)



(Fig. 1a). The connectivity of the neighboring carboxylate oxygen atoms with Cd(II) centers results in a 1-D coordination polymer. The aqua ligands makes a strong intermolecular hydrogen bond with coordinated O atom of bridging adc ligands with the O...O separation of 2.68–2.71 Å (Fig. 1b). In addition, these 1D chains are self-assembled through weak C–H... π interactions along the 4-phpy ligands with edge-to-face distances are in the range of 3.306–3.375 Å (Fig. 1c) in *ac* plane. Thus, 1D zigzag chains aggregate through hydrogen-bonding and C–H... π interactions to form a 2D architecture. Moreover, these 1D chains undergo further edge-to-face weak C–H... π interactions (3.188–3.506 Å) among the 4-phpy ligands forming 2D network in *bc* plane (Fig. 1d). These cooperative hydrogen bonding and C–H... π interactions (Tables S4 and S5, ESI†) fabricate a 3D supramolecular arrangement by stacking the 1D chains together (Fig. S2, ESI†).

On the other hand, compound **2** crystallizes in the monoclinic space group *C2/c* with *Z* = 4. Each Zn(II) center in **2** adopts a distorted octahedral geometry and is bonded to two O atoms from two adc anions in monodentate fashion (Zn–O, 2.131(3) Å) and two N atoms from two 4-phpy ligands (Zn–N, 2.139(2)–2.150(3) Å) in the equatorial plane, and by O atoms from two aqua ligands at the axial sites (Zn–O, 2.104(2) Å) (Fig. 4). The bond lengths are comparable to the reported distances in the literatures.³⁹ Overall topology of the compound **2** is isotypical with **1**. However, H-bonded O...O separation is 2.72 Å in **2**,

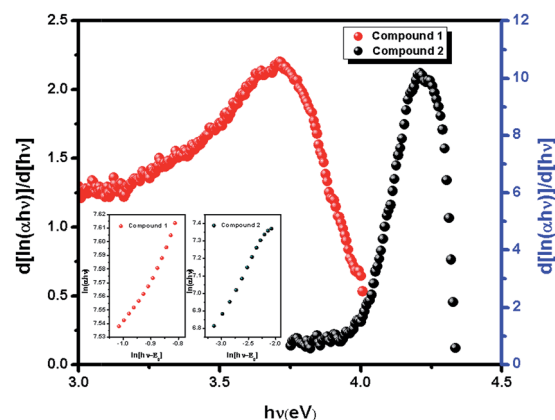


Fig. 2 The plot of $d[\ln(\alpha h\nu)]/d[h\nu]$ versus $h\nu$. The plot of $\ln(\alpha h\nu)$ versus $\ln(h\nu - E_g)$ [inset].

which is little higher than that of **1**. Unlike **1**, the edge-to-face C–H... π distances are symmetrical in both sides of the ring and span in the range of 3.169–3.409 Å. This symmetry has been removed in compound **1**.

TGA and PXRD analysis

To check the thermal stability of the compounds, thermogravimetric analyses (TGA) were performed with the powdered

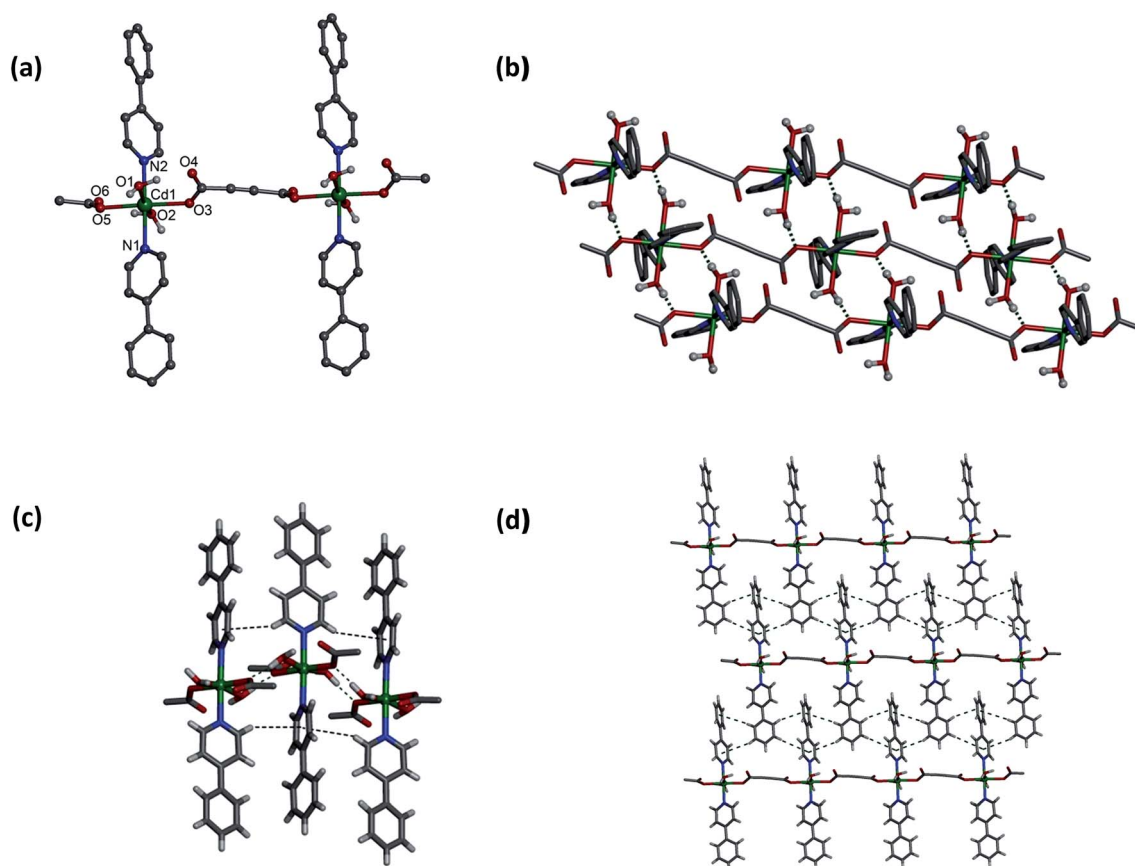


Fig. 1 (a) A portion of 1D coordination polymer of **1**. (b) View of 2D network formed by hydrogen bonding. (c) C–H... π interactions in **1** viewed along *b*-axis. (d) 2D network formed by C–H... π interactions by 1D chains viewed along *a*-axis.



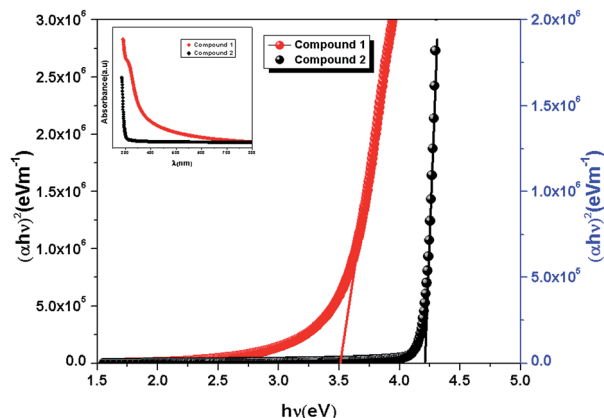


Fig. 3 Band gap determination of compound 1 and 2 using Tauc's equation. The absorbance data of 1 and 2 [inset].

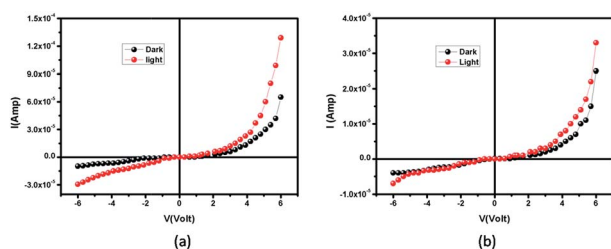


Fig. 4 I - V plot under dark and illumination for (a) compound 1 and (b) 2.

sample within the temperature range 30 °C to 600 °C under N_2 atmosphere. The results of TGA experiments indicate that the compound 1 is stable upto 100 °C whereas compound 2 is thermally stable upto 140 °C (Fig. S3†). Powder X-ray diffraction (PXRD) has been carried out at room temperature. All of the major peaks of PXRD patterns of as-synthesized 1 and 2 match quite well with those simulated from single crystal data indicating phase purity of the bulk (Fig. S4 and S5†).

Optical characterization

To measure the optical band gap corresponding to the excitation of electron from valance band to conduction band, the optical absorbance spectra of the synthesized materials have been recorded. The optical band gap of the compounds 1 and 2 are calculated using the Tauc's equation.⁴⁰

$$(\alpha h\nu) = c(h\nu - E_g)^n \quad (1)$$

where α is the absorption coefficient, c is a constant and E_g is the band gap. The value of n may vary for different types of transition. The eqn (1) may be rewritten as:⁴¹

$$\frac{d[\ln(\alpha h\nu)]}{d[h\nu]} = \frac{n}{h\nu - E_g} \quad (2)$$

In the plot of $d[\ln(\alpha h\nu)]/d[h\nu]$ versus $h\nu$ [Fig. 2], a discontinuity is observed. The discontinuity at a particular energy gives

the approximate value of band gap. Using the value of approximate band gap (E_g), $\ln[(\alpha h\nu)]$ versus $\ln[h\nu - E_g]$ is plotted [inset of Fig. 2]. The slope of the curve provides the value of n . Here the slope (~ 0.5) suggests that the transitions are directly allowed for both the materials. The more accurate band gap of the compounds were calculated by extrapolating the linear portion of $(\alpha h\nu)^2$ versus $h\nu$ [Fig. 3]. The value of optical direct band gap of 1 and 2 are evaluated as 3.51 eV and 4.21 eV. Inset of Fig. 3 shows the optical absorbance versus wavelength for the compounds 1 and 2.

Electrical characterization

The current-voltage (I - V) characteristics of compounds 1 and 2 were recorded under dark and illuminated conditions (Fig. 4). In both the conditions, 1 shows higher magnitude of current. At room temperature, the conductivity of 1 under dark and illuminated conditions are calculated as $1.53 \times 10^{-7} \text{ S cm}^{-1}$ and $3.04 \times 10^{-7} \text{ S cm}^{-1}$ respectively, whereas the values of these parameters of 2 are $5.90 \times 10^{-8} \text{ S cm}^{-1}$ and $7.78 \times 10^{-8} \text{ S cm}^{-1}$ respectively. Interestingly, it has been observed that both the conductivity and photosensitivity are enhanced for 1 in both the conditions. Here, 1 exhibits the photosensitivity of the order of 1.99, which is $\sim 51\%$ higher than the value, 1.32 obtained in 2.

To get better realization of the charge transport mechanism in the device, the thermionic emission (TE) theory is adopted. According to TE theory the current of a diode can be expressed as:⁴²

$$I = I_0 \left[\exp\left(\frac{qV}{nkT}\right) - 1 \right] \quad (3)$$

where

$$I_0 = AA^*T^2 \exp\left(-\frac{q\phi_B}{kT}\right) \quad (4)$$

here, I_0 , q , k , T , A , A^* , η and ϕ_B stand for the reverse saturation current, electronic charge, Boltzmann constant, temperature in Kelvin, effective diode area ($7.065 \times 10^{-6} \text{ m}^2$), effective Richardson constant ($1.20 \times 10^6 \text{ A m}^{-2} \text{ K}^{-2}$), ideality factor and the effective barrier height at zero bias respectively.

The I - V characteristics of 1 and 2 are presented in Fig. 5. The diverse conduction mechanisms of the devices are obtained from the different regions of the graph. Here Fig. 5 exhibits three different regions under forward bias. The slope of the region-I is less than 1 which is due to ohmic conduction mechanism ($I \propto V$) in the low voltage regime. In the region-II, with the slope 1.92–1.94, the conduction mechanism is governed by SCLC mechanism dominated by discrete trapping level ($I \propto V^2$). The region-III, with the slope 3.17–3.78, is influenced by SCLC mechanism ($I \propto V^n$) due to exponential distribution of traps.

The reverse saturation current is determined from the straight line intercept of $\ln I$ - V plot (inset of Fig. 5). The zero bias barrier height is calculated using the following expression.

$$\phi_B = \frac{kT}{q} \ln\left(\frac{AA^*T^2}{I_0}\right) \quad (5)$$



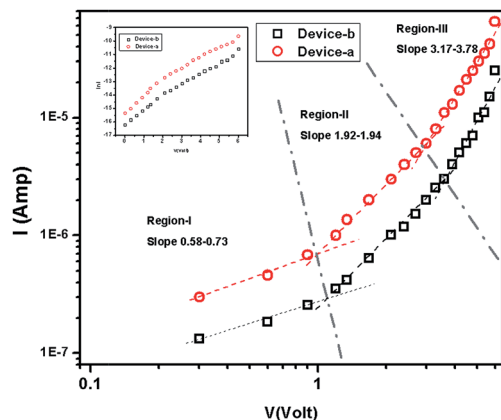


Fig. 5 Logarithmic plot of I - V characteristics in dark condition for 1 and 2. $\ln I$ versus V plot [inset].

The barrier heights of 1 and 2 have been obtained as 0.72 eV and 0.75 eV respectively. Significantly, the lower barrier height of 1 indicates the lower turn-on voltage and higher image force barrier lowering. So the compound 1 is expected to be better candidate for the fast switching device.⁴³

To get the better insight of the charge transport mechanism in the SCLC region (region-II), the effective carrier mobility is estimated from I versus V^2 plot (Fig. 6) using the Mott-Gurney equation.⁴⁴

$$I = \frac{9\mu_{\text{eff}}\epsilon_0\epsilon_r A}{8} \left(\frac{V^2}{d^3} \right) \quad (6)$$

where, ϵ_0 is the free space permittivity, ϵ_r is the dielectric constant of the compounds (10.26 and 5.99 for 1 and 2 respectively), μ_{eff} is the effective mobility of electron and d is the thickness (~ 10 μm) of the thin film. The effective area is assumed as 7.065×10^{-6} m^2 . The transit time (τ) of the charge carrier is estimated with the help of following equation [inset of Fig. 6].⁴⁵

$$\tau = \frac{9\epsilon_0\epsilon_r A}{8d} \left(\frac{V}{I} \right) \quad (7)$$

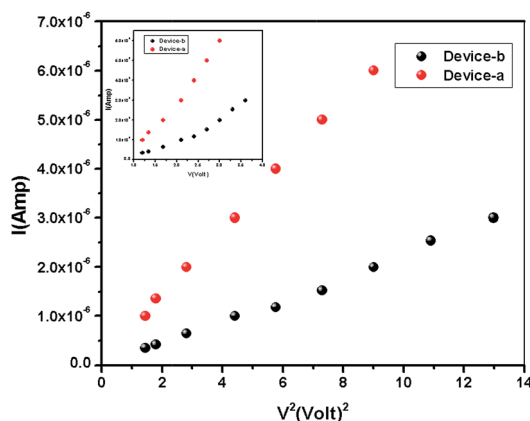


Fig. 6 I - V^2 plot for 1 and 2 in the SCLC region (region-II). I - V plot for determination of the transient time from the SCLC region-II [inset].

To calculate the diffusion length of charge carrier, Einstein-Smoluchowski equation is used.⁴⁵

$$\mu_{\text{eff}} = \frac{qD}{kT} \quad (8)$$

and

$$L_D = \sqrt{2D\tau} \quad (9)$$

where, D is the diffusion co-efficient. The carrier concentration (N_c) near the junction of the devices 1 and 2 is estimated by using the following equation,

$$N = \frac{\sigma}{q\mu_{\text{eff}}} \quad (10)$$

The values of carrier mobility, transient time and diffusion length are summarized in Table 1 with other parameters. It has been observed that the charge carrier mobility of 1 is enhanced up to 68% compared to 2, whereas the diffusion length is increased by 6%. Though the transient time is decreased for 1, but the $\mu\tau$ product is enhanced for 1. The charge carrier for 1 is also enhanced up to 54% compared to 2. The credibility of the compounds depends not only on the charge carrier mobility but also on DOS (density of states) near the Fermi level. The DOS near the Fermi level is determined from the SCLC region of I - V characteristic using den Boer method.⁴⁶

$$N(E_F) = \frac{2\epsilon_0\epsilon_r(V_2 - V_1)}{eL^2\Delta E_F} \quad (11)$$

where L is the thickness of the film, ΔE_F is the shift in the quasi-Fermi level, V_1 and V_2 different voltages applied to the diode in the SCLC region (region-II).

$$\Delta E_F = KT \ln \left(\frac{I_2 V_1}{I_1 V_2} \right) \quad (12)$$

where, I_2 and I_1 are the current values measured at voltages V_2 and V_1 respectively.

Beyond the turn-on voltage, the current increases exponentially with forward bias voltage due to space charge limited conduction (SCLC). Higher number of DOS at Fermi level generally indicates the occurrence of more number of trapping states. More trapping states directly influence the performance of device by degrading its quality. Here the DOS at Fermi level is decreased for 1. Therefore, from all the measured values listed in Table 1, it is very clear that the device fabricated by compound 1 is better than 2. The performance indicating parameters like, mobility, diffusion length, DOS are improved for 1.

Electrochemical impedance spectroscopy

Electrochemical Impedance spectroscopy (EIS) study is done in the regime of frequency 100 Hz to 1 MHz under 0.5 V bias voltage. Fig. 7 illustrates the nyquist plot of the compounds 1 and 2. It clearly indicates that the semicircle for 1 possesses lesser area than 2. This smaller semicircle depicts the bulk resistance (intercept of semicircle on the Z' axis) of 1 which is lower than 2. This lower bulk resistance for compound 1



Table 1 Comparison of I - V characteristic parameters of 1 and 2

Sample	Barrier height	μ_{eff} ($\text{m}^2 \text{V}^{-1} \text{s}^{-1}$)	τ (s)	$(\mu\tau)$ ($\text{m}^2 \text{V}^{-1}$)	DOS ($\text{V}^{-1} \text{m}^{-3}$)	L_D (μm)	Turn on voltage	SCLC conductivity (S cm^{-1})	N_c (m^{-3})
1	0.72	9.15×10^{-7}	2.63×10^{-5}	2.406×10^{-11}	2.576×10^{39}	1.116	0.9746	3.885×10^{-8}	10.45×10^{19}
2	0.75	5.44×10^{-7}	3.92×10^{-5}	2.132×10^{-11}	3.525×10^{39}	1.050	1.082	1.522×10^{-8}	6.78×10^{19}

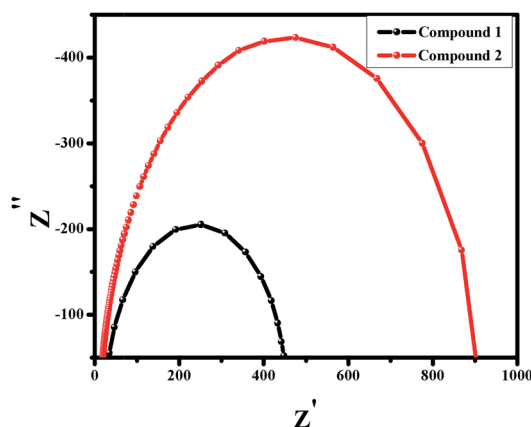


Fig. 7 Nyquist plot for compound 1 and 2.

indicates better possibility of charge transfer and reduction of chances of charge recombination.

Conclusions

In summary, we have synthesized two 1D coordination polymers from linear dicarboxylate adc by varying metal ion from Cd(II) to Zn(II). The 1D chains herein stack together by the combination of H-bonding and C-H \cdots π interactions to form 3D supramolecular structures. These two materials exhibit electrical conductivity and reveal Schottky barrier diode behavior. However, compound 1 shows enhanced photosensitivity, higher mobility, diffusion length and lower DOS at Fermi levels compared to compound 2. Improved charge transport has also been reflected in the device fabricated with 1. Therefore, it appears that compound 1 shows better device applicability compared to 2.

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