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Design and synthesis of dodecahedral carbon nanocages incorporated with Fe₃O₄†

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Novel dodecahedral carbon nanocages (NCs) modified with Fe₃O₄ nanoparticles (Fe₃O₄/C NCs) were constructed by utilizing a zeolitic imidazolate framework (ZIF-8) as the template and polydopamine (PDA) as the carbon source. The obtained nanocages not only display specific morphology with controllable feature structures, but also possess a large surface area and inner cavity. This unique structure of the nanocages and special synthesis approach make Fe₃O₄ NPs homogeneously distributed in the system, which further enriched the function of the carbon nanocages. Due to the integrated functional nanoparticles and large interior of such specific structures, the as-prepared Fe₃O₄/C NCs are further used as anode materials in lithium ion batteries (LIBs), which exhibit high activity in lithiation/delithiation cycling process and stable capacity retention. The nanocages we prepared may provide a train of thought to construct a series of hollow carbon nanostructures with functional materials incorporated.

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1. Introduction

Hollow nanocages have attracted great attention because of their unique properties such as low density, high surface area, shell permeability, chemical and physical stability, *etc.*^{1,2} They have successfully found applications in many fields ranging from delivery vehicles to batteries.^{3–5} Various nanocages have been prepared by template methods, supramolecular assembly, galvanic replacement, *etc.*^{6–8} However, there still exist some issues to be addressed, such as complicated synthesis, tedious post-treatment procedures, high cost, *etc.* In addition, nanocages with non-sphere shapes and regular interiors are still difficult to obtain. These unique nanocages are of particular interests due to their special physical properties and can serve as building blocks for assembly into hierarchical synthetic architectures. Moreover, they bring advantages that aren't able to be achieved through traditional round shapes driven by surface-tension.⁹ Therefore, it is significant to obtain nanocages with specific morphology and controllable properties.

To solve these problems, researchers have developed approaches to synthesize anisotropic nanocages with different structures. For instance, Lou's group successfully established a synthetic approach of inorganic nanocages *via* a templating method.^{5,9–12} In addition, we exploited a novel 'chelation

competition induced polymerization' (CCIP) method to prepare hollow polydopamine nanocontainers.¹³ These nanocages not only possess the inherent properties of materials, but also exhibit some new functions which are derived from the new hollow structure. Moreover, after annealing under inert gas, these polymer-based hollow structures can be further transformed into carbon nanocages.

Nowadays carbon materials are more and more popular due to their good electronic conductivity, high availability and cost efficiency and shown great potential in many fields like lithium-ion batteries (LIBs).^{14–16} In order to meet the ever-growing demands of market, anode materials are no longer limited to traditional graphite electrodes (theoretical capacity of 370 mA h g^{−1}), various carbon materials with different structures were prepared.^{17–19} Moreover, nanocomposites integrated carbon with other materials possessing higher theoretical capacity (such as metal oxide, Fe₃O₄, 924 mA h g^{−1}) were also investigated.^{18,20–22} Compared to the solid one, carbon nanostructures with hollow interior have shown much more advantages in LIBs since they will facilitate electrolyte transportation, lithium ion diffusion, provide more space to alleviate the strain arising from Li insertion/removal and enlarge the surface to volume ratio, resulting in faster reaction kinetics and higher specific capacity.^{20–22} For example, Xiong group has prepared a polyhedral carbon nanocages by carbonizing ZIF-8@ZIF-67 core-shell crystals. Such bi-metal embedded N-doped nanocages show an initial discharge capacity of 809 mA h g^{−1} and a capacity retention of 702 mA h g^{−1} after 400 cycles at a current density of 0.2 A g^{−1}.²³ Despite of the achieved progress, it is still very important for us to construct more novel functional nanostructures as anode materials for LIBs.

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In this work, a novel kind of dodecahedral carbon nanocages was successfully constructed. The method we developed exactly provided an effective way to prepare a unique carbon nanocage structure by using zeolitic imidazolate framework (ZIF-8) as template and polydopamine as carbon resource. The obtained carbon nanocages not only maintain the shape of ZIF-8 with highly uniform shells and tunable sizes, but also make it easy to introduce functional nanoparticles (*e.g.* Fe₃O₄ NPs in this work, named Fe₃O₄/C NCs) into the system. The carbon nanocage matrix has more space for the volume expansion of Fe₃O₄ NPs due to their large inner cavity and surface area, as well as provide a continue network for ion and electron transfer by embedding nanoparticles into the system. Moreover, N-doped is easily realized by choosing PDA as carbon resource, further improving the conductivity of the nanocages. The electrochemical performance as anode materials of LIBs was evaluated. The results show good cyclic performance and stability, indicating their promising application in energy fields.

2. Experimental section

2.1 Materials

Iron(III) acetylacetonate (Fe(acac)₃, 97%), dibenzyl ether (≥98%, GC), 1,2-hexadecanediol, oleic acid (≥99%, GC), oleylamine (OA) (≥98%), polyvinylpyrrolidone (*M_w* 40 000), zinc nitrate hexahydrate (Zn(NO₃)₂·6H₂O), 2-methylimidazole (99%) and dopamine hydrochloride were purchased from Sigma-Aldrich. Hexane, chloroform, ethanol were purchased from Beijing Chemical Works.

2.2 Preparation of OA-stabilized Fe₃O₄ nanoparticles

OA-stabilized Fe₃O₄ NPs were prepared following the procedure of Sun *et al.*²⁴ The as-prepared NPs were precipitated with ethanol, washed several times and centrifuged at 7000 rpm for 10 min. Then Fe₃O₄ NPs were re-dispersed in chloroform.

2.3 Preparation of Fe₃O₄/ZIF-8 nanostructures

Before the encapsulation procedure, OA-stabilized Fe₃O₄ NPs were firstly modified with surfactant PVP. In brief, Fe₃O₄ NPs were precipitated with ethanol, and centrifuged at 7000 rpm for 10 min. A solution of PVP (*M_w* = 40 000) in chloroform (20 mg mL⁻¹) was used to re-dispersed the NPs. Then the mixture was kept for 24 hours, the PVP-modified Fe₃O₄ NPs were precipitated with hexane and collected by centrifugation at 7000 rpm for 10 min. The NPs were washed with chloroform for several times to remove the excess free PVP. And then the sample was re-dispersed in methanol.

In the encapsulation procedure, methanol was used as solvent. 10 mL solution of 2-methylimidazole (25 mM) and 10 mL 25 mM Zn(NO₃)₂·6H₂O (containing 1 mL PVP-stabilized Fe₃O₄ NPs methanol solution) were mixed and then allowed to react at room temperature for 1 hour without stirring. The product was collected by centrifugation at 5000 rpm for 10 min, washed several times with methanol, and final dispersed in 3.0 mL methanol.

2.4 Preparation of Fe₃O₄/PDA NCs

3.0 mL Fe₃O₄/ZIF-8 methanol solution was mixed with 4.5 mL methanol and 1.5 mL dopamine (20 mM in methanol). The mixture was stirring and refluxing at 60 °C for 12 h, the color of which turned to black gradually. The product was collected by centrifugation at 6000 rpm for 10 min and washed with methanol for several times, and freeze-drying overnight.

2.5 Preparation of Fe₃O₄/C NCs

Fe₃O₄/PDA NCs powder was thermal annealed at 550 °C under N₂ flow with 3 °C min⁻¹ for 4 h.

2.6 Electrochemical measurements

The working electrodes were prepared by mixing the Fe₃O₄/C NCs, acetylene black, and poly(vinylidene fluoride) (PVDF) at a weight ratio of 80 : 10 : 10 in electrolyte of dimethyl carbonate (DMC) and ethylene carbonate (EC) (1 : 1 v/v), pasting the mixture on pure Cu foil and dried at 100 °C under vacuum for 10 h. The discharge-charge measurements were carried out at several different current densities between the cut off potentials of 0.01 and 3 V using Land CT2001A battery test system. Cyclic voltammogram (CV) measurements were performed on CHI750E electrochemical workstation.

2.7 Characterization

Transmission electron microscopy (TEM) images were collected on a Hitachi H-800 electron microscope operated at 200 kV with a CCD camera. High-resolution TEM images and mapping images were recorded using a JEM-2100F electron microscope at an acceleration voltage of 200 kV with a CCD camera. Scanning electron microscope (SEM) images were taken with a JEOL FESEM 6700F electron microscope with primary electron energy of 3 kV. The samples were sputtered with a thin layer of Pt prior to imaging. Inductively Coupled Plasma (ICP) results were conducted by OPTIMA 3300DV spectrometer. X-ray diffraction (XRD) analysis was performed by X-ray diffractometer, Model Rigaku Ru-200b, using a nickel-filtered Cu Kα radiation and the data were collected from 10° to 80°. X-ray photoelectron spectroscopic (XPS) analysis was done on an ESCALAB 250 spectrometer. Raman spectrum was recorded using a Horiba Jobin-Yvon micro Raman spectrometer, equipped with a microscope and a laser of 633 nm as the excitation source. N₂ adsorption-desorption measurements were conducted by Micromeritics ASAP 2010 instrument. The pore size distribution was calculated by the Barrett-Joyner-Halenda (BJH) method. And the specific surface area was calculated by using BJH and Brunauer-Emmett-Teller (BET) method. Hysteresis loop was tested by SQUID-VSM system. Magnetic measurements were performed using a SQUID magnetometer (SQUID-MPMS-XL) at 300 K by cycling the magnetic field between -30 and 30 kOe.

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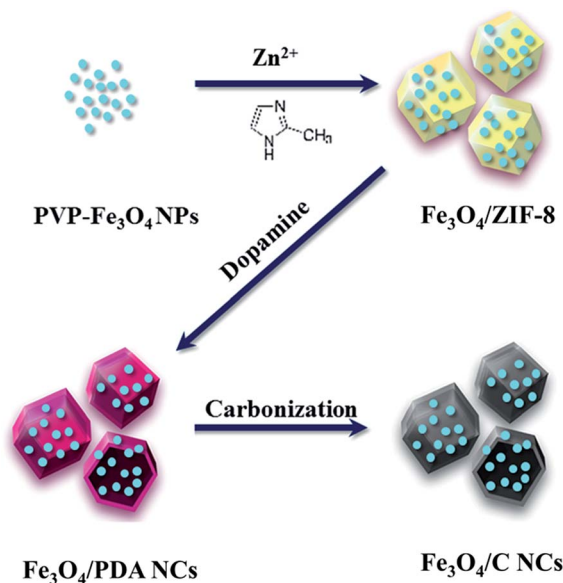


3. Results and discussion

3.1 Fabrication and regulation of Fe₃O₄/C NCs

The synthesis process of dodecahedral Fe₃O₄/C NCs is schematically illustrated in Scheme 1. Firstly, OA-stabilized Fe₃O₄ nanoparticles (NPs), obtained through hydrothermal method, were firstly washed and modified with polyvinylpyrrolidone (PVP).²⁴ The Fe₃O₄/ZIF-8 nanostructure could be easily obtained by adding PVP-Fe₃O₄ NPs into the mixture of Zn(NO₃)₂ and 2-methylimidazole.²⁵ Then, dopamine (DA) was added to react with the as-prepared Fe₃O₄/ZIF-8 and refluxed for 7 h. During this process, the Zn²⁺ in ZIF-8 was grabbed by DA *via* its catechol group, leading the disassembly of the template and releasing 2-methylimidazole which triggered the polymerization of DA.¹³ Finally, the nanostructure got hollowed and Fe₃O₄/polydopamine nanocages (Fe₃O₄/PDA NCs) formed. After dried in vacuum and annealed at 550 °C under N₂ flow (3 °C min⁻¹ for 4 h), the polymer shell of the PDA nanocages was transformed into a carbon layer and formed Fe₃O₄/C NCs. In our method, the hollow C nanocage was prepared without extra template removal, and the functionalization of the C nanocage, such as incorporating Fe₃O₄ NPs to the system is easily achieved through surfactant modification.

Electron microscopy measurements were conducted to monitor the product morphology changes in each preparation process. As shown in Fig. S1,† OA-stabilized Fe₃O₄ NPs were about 5 nm.²⁴ After forming Fe₃O₄/ZIF-8, the nanostructure was about 350 nm with many Fe₃O₄ NPs loading in and on (Fig. 1A). The dodecahedral shape of Fe₃O₄/ZIF-8 is very similar to that of pure ZIF-8, suggesting that incorporating functional nanoparticles have no obvious effects to the morphology of ZIF-8 nanostructure. With the reaction proceeding, the color of the solution turned gradually from brownish (Fe₃O₄/ZIF-8) to black, finally forming the hollow nanocages (Fig. 1B). The obtained Fe₃O₄/PDA NCs are *ca.* 364 nm with *ca.* 30 nm of PDA shell.



Scheme 1 Schematic illustration of the synthesis of Fe₃O₄/C NCs.

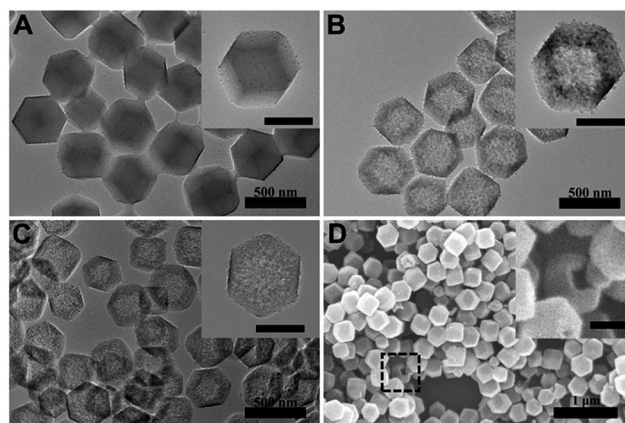


Fig. 1 TEM images of Fe₃O₄/ZIF-8 (A), Fe₃O₄/PDA NCs (B), Fe₃O₄/C NCs (C) and SEM image of Fe₃O₄/C NCs (D). Insets: TEM images of Fe₃O₄/ZIF-8, Fe₃O₄/PDA NCs, Fe₃O₄/C NCs and SEM image of a broken Fe₃O₄/C NCs with high magnification. The scale bar in the insets is 200 nm.

The hollowing process was monitored by TEM images in Fig. 2. With the time processing, the original solid Fe₃O₄/ZIF-8 was firstly split to 'core-shell' structure (Fig. 2B). Then the 'core' part of the structure was getting smaller and smaller, and totally disappeared (Fig. 2C and D). Fe₃O₄/C NCs were prepared by further annealing Fe₃O₄/PDA NCs under N₂ flow. As shown in Fig. 1C, the obtained NCs maintained the original shape of ZIF-8 nanostructures. SEM image also shows the appearance of Fe₃O₄/C NCs, which perfectly retained the dodecahedral morphology (Fig. 1D). The large hollow inner cavity observed from a typical broken structure also proves the formation of nanocages (inset in Fig. 1D). Optical photographs in Fig. S1B† shows that when placing a magnet, all the Fe₃O₄/C NCs the powder was attracted to the wall of glass vial, suggesting their good magnetism of the NCs. Moreover, the completely reversible *M* (*H*) hysteresis curve suggests the superparamagnetic structure of Fe₃O₄/C NCs (Fig. S1D†), which is consistent with Fe₃O₄ NPs we used (Fig. S1C†). Elemental mapping measurements were also conducted to see the distribution of Fe₃O₄ in

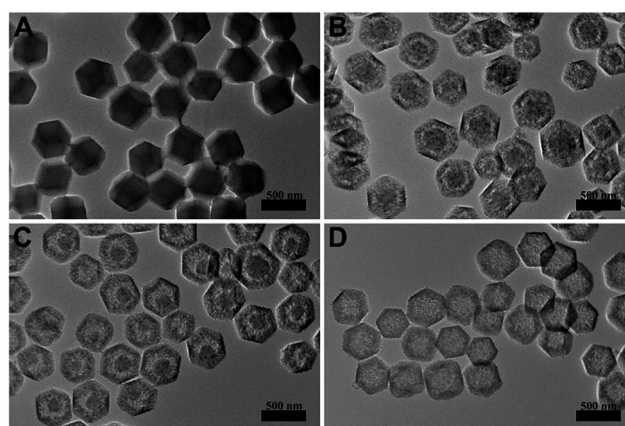


Fig. 2 TEM images of Fe₃O₄/PDA NCs with the reaction time of 0 h (A), 2 h (B), 4 h (C) and 7 h (D).



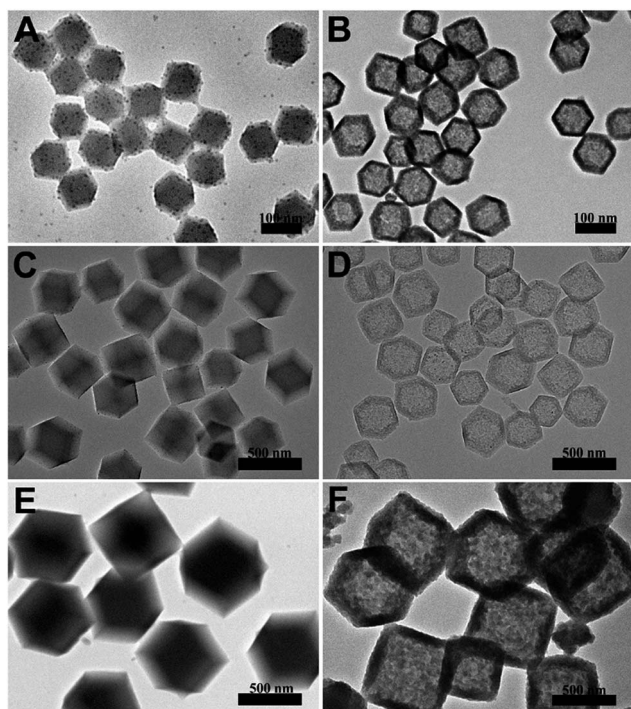


Fig. 3 TEM images of $\text{Fe}_3\text{O}_4/\text{ZIF-8}$ nanostructure with different sizes (97 nm A, 350 nm C and 574 nm E) and the corresponding $\text{Fe}_3\text{O}_4/\text{C}$ NCs (90 nm B, 332 nm D and 561 nm F).

both $\text{Fe}_3\text{O}_4/\text{PDA}$ NCs and $\text{Fe}_3\text{O}_4/\text{C}$ NCs. As shown in Fig. S2,† the element C presents a hollow dodecahedral shape and distributes throughout the structure (red), while Fe (green) major distributed on the backbone of the carbon nanocages. All these results prove that Fe_3O_4 NPs are retained in the final products even experiencing hollowing process. Inductively coupled plasma (ICP) results present the content of metallic elements in Table S2,† suggesting the loading capacity of Fe_3O_4 could reach 20.3% of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs. Series of $\text{Fe}_3\text{O}_4/\text{C}$ NCs were obtained by precisely controlling the reaction condition. For example, the size of the NCs could be regulated by changing that of ZIF-8 templates. As shown in Fig. 3, $\text{Fe}_3\text{O}_4/\text{C}$ NCs from down-to-100 nm (90 nm) to over 500 nm (561 nm) could be achieved. Moreover, the loading capacity of the Fe_3O_4 NPs could also be modulated (Fig. S3†) and the functional nanoparticles is replaceable. These controllable nanocages may meet the demands of application in variety of fields by choosing suitable sizes and functionalities in the future.

3.2 Characterization of $\text{Fe}_3\text{O}_4/\text{C}$ NCs

Raman spectroscopy is displayed to characterize the nature of $\text{Fe}_3\text{O}_4/\text{C}$ NCs (Fig. 4). Two major peaks at *ca.* 1336 and 1580 cm^{-1} are corresponding to the 'D' and 'G' bands of carbon materials, respectively.²² The intensity ratio of these two bands (I_D/I_G) reaches 1.47, suggesting that the presence of carbon in the NCs is in the form of amorphous structure. Besides, the peaks at *ca.* 310, 534, and 626 cm^{-1} are assigned to Fe_3O_4 NPs, meanwhile *ca.* 426, 559, and 1100 cm^{-1} belong to ZnO, which is

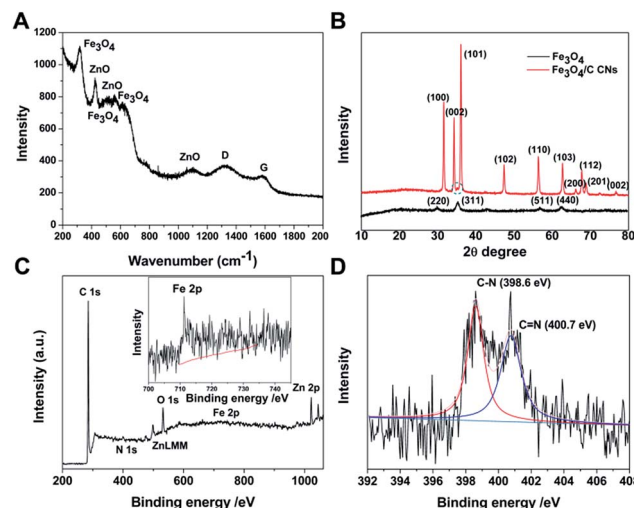


Fig. 4 Raman spectrum of $\text{Fe}_3\text{O}_4/\text{C}$ NCs (A), XRD patterns of Fe_3O_4 (black line) and $\text{Fe}_3\text{O}_4/\text{C}$ NCs (red line) (B), XPS spectrum (C) and N 1s spectrum (D) of $\text{Fe}_3\text{O}_4/\text{C}$ NCs.

further proved by the results of powder X-ray diffraction (XRD) measurements in Fig. 4B.^{26,27} $\text{Fe}_3\text{O}_4/\text{C}$ NCs exhibit strong wurtzite patterns of ZnO (red line, JCPDS card number 36-1451), which is formed during the annealing process, and the content is 37.4% (Table S2†).²⁸ HRTEM also prove the crystallinity of ZnO (Fig. S4†). Comparing to $\text{Fe}_3\text{O}_4/\text{PDA}$ NCs, no obvious structure changes were observed in $\text{Fe}_3\text{O}_4/\text{C}$ NCs, suggesting that the existence of ZnO has no effect on the structure of our system. It is worth noting that only a weak peak (*ca.* 35°) belongs to Fe_3O_4 (green dashed circle, JCPDS no. 65-3107) was observed in the pattern. This is due to the originally low intensity of pure Fe_3O_4 (black line), together with the shielding effect of the ZnO.²² A diffuse scattering peak observed at around 20° is attributed to the amorphous form of carbon nanocages, which is consisted with the Raman spectroscopy results (Fig. 4A). The chemical composition of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs was characterized by X-ray photoelectron spectroscopy (XPS). As shown in Fig. 4C, an overall XPS spectrum of $\text{Fe}_3\text{O}_4/\text{C}$ NCs indicates the presence of C, Fe, O, N and Zn elements in the NCs, and the peak at 711.2 eV can be assigned to the Fe 2p_{3/2}, referencing to the standard XPS spectrum for Fe_3O_4 .²⁹ In the high-resolution N 1s spectrum (Fig. 4D), two peaks observed at ~398.6 eV and 400.7 eV are identified to C–N and C=N, indicating N is doped in the structure of carbon nanocages. It is derived from the amine group in the PDA, which also serves as carbon source in our system, and the doping content is 2.8%, as shown in Table S1.†

Such doping is beneficial to improve the electrical conductivity and surface hydrophilicity, further increasing the potential application of $\text{Fe}_3\text{O}_4/\text{C}$ NCs in electrochemistry field. During the formation of $\text{Fe}_3\text{O}_4/\text{PDA}$ NCs, the disassembling process of ZIF-8 template demonstrates the permeability of PDA shell for the mass diffusion through the interstices (Fig. 2). After the carbonization process, the porous property of $\text{Fe}_3\text{O}_4/\text{C}$ NCs was investigated through N_2 adsorption–desorption isotherm (Fig. 5). The nanocages display a typical IV-type isotherm, indicating the presence of mesopores. An obvious H3-type



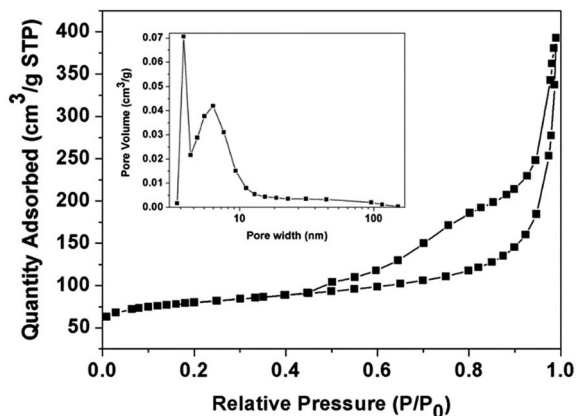


Fig. 5 Nitrogen adsorption-desorption isotherms of $\text{Fe}_3\text{O}_4/\text{C}$ NCs. Inset: BJH desorption dV/dw pore width of $\text{Fe}_3\text{O}_4/\text{C}$ NCs.

hysteresis loop is observed after P/P_0 reaching 0.5, due to the adsorption of N_2 in the inner cavities of $\text{Fe}_3\text{O}_4/\text{C}$ NCs.²² The BET surface area is $383.9 \text{ m}^2 \text{ g}^{-1}$, and the pore sizes of the nanocages are mainly focus on *ca.* 3.8 nm and 6.2 nm, deriving from the formation of nanocage and the gas release during the carbonization of the polymer shell. Such relatively high surface area will provide highly permeable for mass diffusion of electrolyte, improving the interaction between the nanocages and solution.

3.3 Lithium storage properties of $\text{Fe}_3\text{O}_4/\text{C}$ NCs

The Fe_3O_4 was chosen to incorporate in the nanocages because it has been regarded as one of promising anode materials for rechargeable LIBs. It possesses high theoretical capacity ($\sim 927 \text{ mA h g}^{-1}$), natural abundance and nontoxicity.^{26,30–32} Motivated by such intriguing structure, we evaluated the electrochemical properties of the obtained $\text{Fe}_3\text{O}_4/\text{C}$ NCs. Fig. 6A shows the representative cyclic voltammograms (CVs) of the sample between 10 mV and 3 V at a scan rate of 0.1 mV s^{-1} . A relatively weak reduction peak is observed around 0.1 V, which responses to the multistep reduction processes from ZnO to Zn^0 and the generation of Li-Zn alloy.^{33,34} The peak around 0.5 V in the first scan of the electrode is attributed to the formation of the solid electrolyte interphase (SEI) layer. The CV curves at the second and third cycles show a broad peak at 0.5–1.0 V, corresponding to the reduction of ZnO and Fe_3O_4 and the formation of amorphous Li_2O .³³ The reduction peaks at $\sim 0.9 \text{ V}$ in the first cycle and $\sim 0.83 \text{ V}$ at 2nd scan are ascribed to the reduction of Fe_3O_4 and the formation of SEI layer. There are three adjacent small oxidation peaks appearing below 1 V in the following anodic scan, which may be indicated as the multistep dealloying process of Li-Zn alloy (LiZn , Li_2Zn , LiZn_2 , and Li_2Zn_3).³³ In addition, the smooth and broad oxidation peak around 1.3–1.7 V may be ascribed to the decomposition of Li_2O and the oxidation of Fe^0 .²⁵ The little changes in these oxidation peaks indicate a good reversibility of the electrochemical reaction.

The charge-discharge voltage profiles of $\text{Fe}_3\text{O}_4/\text{C}$ NCs with different cycles at 100 mA g^{-1} in the voltage range of 0.01–3.0 V are shown in Fig. 6B. In the first cycle, $\text{Fe}_3\text{O}_4/\text{C}$ NCs exhibit an initial discharge capacity of 1370 mA h g^{-1} and a charge capacity of 789

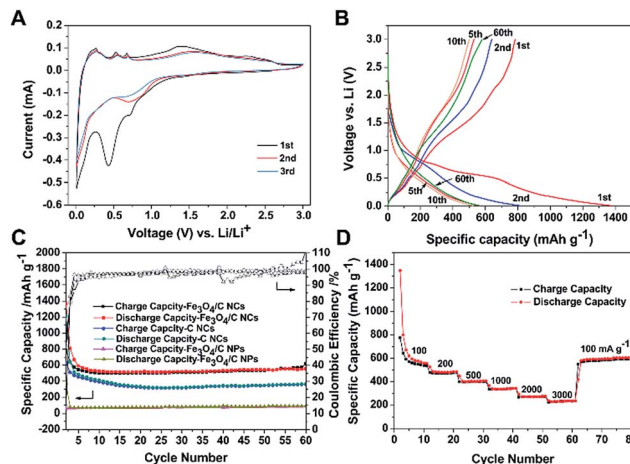


Fig. 6 Cyclic voltammograms of $\text{Fe}_3\text{O}_4/\text{C}$ NCs between 10 mV and 3 V at a scan rate of 0.1 mV s^{-1} (A), charge-discharge voltage profiles (B), cycling performance and the corresponding CE of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs compared to C NCs and solid $\text{Fe}_3\text{O}_4/\text{C}$ NPs at 100 mA g^{-1} (C) and rate capability of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs at various current rates (D).

mA h g^{-1} , with an initial coulombic efficiency (CE) of 57.6% based on the total mass of the sample. The capacity loss in the first cycle may be attributed to the irreversible processes such as the formation of SEI layer and amorphous Li_2O , trapped Li^+ in the inner cavity of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs and the decomposition of the electrolyte, *etc.*^{24,25} Fig. 6C shows the cycling performance and the corresponding CE of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs. The first charge and discharge capacity are $789\text{--}1370 \text{ mA h g}^{-1}$ and gradually remain at 584 mA h g^{-1} with the prolonging of cycling. Moreover, its CE after 4 cycles maintains at about 100%. The rate capability of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs was also characterized to evaluate their electrochemical performance. As shown in Fig. 6D, with the increasing of the current density (100, 200, 500, 1000, 2000, 3000 mA g^{-1}), the $\text{Fe}_3\text{O}_4/\text{C}$ NCs can still deliver reversible capacities of 574, 485, 405, 338, 269 and 235 mA h g^{-1} , respectively. Remarkably, when the current density turns back to 100 mA g^{-1} , a reversible capacity of 603 mA h g^{-1} can be obtained. To demonstrate the advantage of $\text{Fe}_3\text{O}_4/\text{C}$ NCs, the cycling performance of carbon nanocages and solid $\text{Fe}_3\text{O}_4/\text{C}$ nanoparticles ($\text{Fe}_3\text{O}_4/\text{C}$ NPs with the same composition with $\text{Fe}_3\text{O}_4/\text{C}$ NCs) is also investigated under the same conditions for comparison. As shown in Fig. 6C, a capacity of *ca.* 367 mA h g^{-1} is delivered by C NCs, while only *ca.* 98 mA h g^{-1} of the $\text{Fe}_3\text{O}_4/\text{C}$ NPs, showing much lower cycling performance. The comparison of the cycling performance shows the superiority of both incorporating Fe_3O_4 NPs and the specific hollow structure in $\text{Fe}_3\text{O}_4/\text{C}$ NCs. Moreover, our $\text{Fe}_3\text{O}_4/\text{C}$ NCs have shown obviously improved electrochemical performance compared to the C NCs (Fig. S5†). The status of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs electrode after cycling was characterized by SEM. As shown in Fig. S6A,† no obvious changes in morphology of $\text{Fe}_3\text{O}_4/\text{C}$ NCs after cycling, and the hollow interior observed in the typical broken particle prove the nanostructures were $\text{Fe}_3\text{O}_4/\text{C}$ NCs (Fig. S6B†), suggesting that our hollow nanocages can effectively relieve the volume expansion, stabilizing and preserving the integrity of $\text{Fe}_3\text{O}_4/\text{C}$ NCs during the charge-discharge process.



The good cycling performance of the $\text{Fe}_3\text{O}_4/\text{C}$ NCs is benefitted from their specific structure, which effectively integrated advantages of each component. Firstly, the mesoporous wall and large inner cavity of carbon nanocages greatly enlarge the surface sites, increase the electrochemical reaction area, shorten the diffusion pathway of Li ions and allow to buffer the volume change during the charge–discharge process. Secondly, the incorporated Fe_3O_4 NPs and ZnO formed in carbonization display positive contributions in the charge–discharge process. Moreover, Fe_3O_4 NPs are uniformly distributed in carbon nanocages through this method, which fully take their advantages of nanoscaled and large surface area. These synergistic effects of nanocages and Fe_3O_4 contribute to the desired performance of $\text{Fe}_3\text{O}_4/\text{C}$ NCs, indicating their promising application in LIBs.

4. Conclusions

In summary, a distinct carbon nanocages modifies with Fe_3O_4 NPs incorporated has been constructed successfully by using ZIF-8 as template and PDA as carbon source. The morphology of $\text{Fe}_3\text{O}_4/\text{C}$ NCs retained the dodecahedral shape of ZIF-8, and series of NCs can be prepared by precisely controlling the reaction conditions. Since the distinct properties of hollow carbon nanocages and functional Fe_3O_4 NPs, the lithium storage properties of the obtained $\text{Fe}_3\text{O}_4/\text{C}$ NCs are studied and display stable capacity retention and cycling performance even at high rates of 3000 mA g^{-1} in comparison with C NCs, suggesting their promising application in LIBs. The nanocages we fabricated may open up a new platform for preparation of various carbon-based hollow nanostructures with functionalities.

Acknowledgements

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Notes and references

- 1 X. Wang, J. Feng, Y. Bai, Q. Zhang and Y. Yin, *Chem. Rev.*, 2016, **116**, 10983.
- 2 X. Lou, L. A. Archer and Z. Yang, *Adv. Mater.*, 2008, **20**, 3987.
- 3 Z. Luo, X. Ding, Y. Hu, S. Wu, Y. Xiang, Y. Zeng, B. Zhang, H. Yan, H. Zhang, L. Zhu, J. Liu, J. Li, K. Cai and Y. Zhao, *ACS Nano*, 2013, **7**, 10271.
- 4 C. J. Ochs, H. Hong, G. K. Such, J. Cui, A. Postma and F. Caruso, *Chem. Mater.*, 2011, **23**, 3141.
- 5 Z. Wang and X. Lou, *Adv. Mater.*, 2012, **24**, 4124.
- 6 G. Zhang, H. B. Wu, T. Song, U. Paik and X. W. Lou, *Angew. Chem., Int. Ed.*, 2014, **53**, 12590.
- 7 A. O. Moughton and R. K. O'Reilly, *J. Am. Chem. Soc.*, 2008, **130**, 8714.
- 8 C. M. Cobley and Y. N. Xia, *Mater. Sci. Eng., R*, 2010, **70**, 44.
- 9 O. Shchepelina, V. Kozlovskaya, S. Singamaneni, E. Kharlampieva and V. V. Tsukruk, *J. Mater. Chem.*, 2010, **20**, 6587.
- 10 G. Zhang, H. Wu, T. Song, U. Paik and X. Lou, *Angew. Chem., Int. Ed.*, 2014, **53**, 12590.
- 11 L. Shen, L. Yu, X. Yu, X. Zhang and X. Lou, *Angew. Chem., Int. Ed.*, 2015, **54**, 1868.
- 12 X. Yu, L. Yu, L. Shen, X. Song, H. Chen and X. Lou, *Adv. Funct. Mater.*, 2014, **24**, 7440.
- 13 S. Xiang, D. Wang, K. Zhang, W. Liu, C. Wu, Q. Meng, H. Sun and B. Yang, *Chem. Commun.*, 2016, **52**, 10155–10158.
- 14 G. Jeong, Y.-U. Kim, H. Kim, Y.-J. Kim and H.-J. Sohn, *Energy Environ. Sci.*, 2011, **4**, 1986.
- 15 M. Nie, D. Chalasani, D. P. Abraham, A. Bose and B. L. Bose, *J. Phys. Chem. C*, 2013, **117**, 1257.
- 16 K. Xu and A. V. Cresce, *J. Mater. Chem.*, 2011, **21**, 9849.
- 17 B. Li, Y. Wang, H. Rong, Y. Wang, J. Liu, L. Xing, M. Xu and W. Li, *J. Mater. Chem. A*, 2013, **1**, 12954.
- 18 C. Ding, Y. Zeng, L. Cao, L. Zhao and Y. Zhang, *J. Mater. Chem. A*, 2016, **4**, 5898.
- 19 F. Zheng, Y. Yang and Q. Chen, *Nat. Commun.*, 2014, **5**, 5261.
- 20 Y. Yang, J. Li, D. Chen and J. Zhao, A facile electrophoretic deposition route to the $\text{Fe}_3\text{O}_4/\text{CNTs}/\text{rGO}$ composite electrode as a binder-free anode for lithium ion battery, *ACS Appl. Mater. Interfaces*, 2016, **8**, 26730–26739.
- 21 F. Maronia, S. Gabriellia, A. Palmieria, E. Marcantonio, F. Croceb and F. Nobili, High cycling stability of anodes for lithium-ion batteries based on Fe_3O_4 nanoparticles and poly(acrylic acid) binder, *J. Power Sources*, 2016, **332**, 79–87.
- 22 Q. Meng, F. Zhang, L. Wang, S. Xiang, S. Zhu, G. Zhang, K. Zhang and B. Yang, Facile fabrication of mesoporous N-doped $\text{Fe}_3\text{O}_4/\text{C}$ nanospheres as superior anodes for Li-ion batteries, *RSC Adv.*, 2014, **4**, 713–716.
- 23 M. Huang, K. Mi, J. Zhang, H. Liu, T. Yu, A. Yuan, Q. Kong and S. Xiong, *J. Mater. Chem. A*, 2017, **5**, 266.
- 24 S. Sun and H. Zeng, *J. Am. Chem. Soc.*, 2012, **124**, 8204.
- 25 G. Lu, S. Li, Z. Guo, O. K. Farha, B. G. Hauser, X. Qi, Y. Wang, X. Wang, S. Han, X. Liu, H. Zhang, Q. Zhang, X. Chen, J. Ma, S. C. J. Loo, W. D. Wei, Y. Yang, J. T. Hupp and F. W. Huo, *Nat. Chem.*, 2012, **4**, 310.
- 26 C. Ding, Y. Zeng, L. Cao, L. Zhao and Y. Zhang, *J. Mater. Chem. A*, 2016, **4**, 5898.
- 27 J. L. Santosa, T. R. Reinaa, S. Ivanovaa, M. A. Centenoa and J. A. Odriozola, *Appl. Catal., B*, 2017, **201**, 310.
- 28 P. Li, Z. Wei, T. Wu, Q. Peng and Y. Li, *J. Am. Chem. Soc.*, 2011, **133**, 5660.
- 29 Z. Li, M. Jaroniec, P. Papakonstantinou, J. M. Tobin, U. Vohrer, S. Kumar, G. Attard and J. D. Holmes, *Chem. Mater.*, 2007, **19**, 3349.
- 30 Z. Xiao, Y. Xia, Z. Ren, Z. Liu, G. Xu, C. Chao, X. Li, G. Shen and G. Han, *J. Mater. Chem.*, 2012, **22**, 20566.
- 31 S. K. Behera, *J. Power Sources*, 2011, **196**, 8669.
- 32 Y. Deng, Q. Zhang, Z. Shi, L. Han, F. Peng and G. Chen, *Electrochim. Acta*, 2012, **76**, 495.
- 33 Y. Wang, X. Jiang, L. Yang, N. Jia and Y. Ding, *ACS Appl. Mater. Interfaces*, 2014, **6**, 1525.
- 34 M. Ahmad, Y. Shi, A. Nisar, H. Sun, W. Shen, W. Miao and J. Zhu, *J. Mater. Chem.*, 2011, **21**, 7723.

