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Fabrication of SiO₂@silicalite-1 and its use as a catalyst support

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SiO₂@silicalite-1, using silica sol (pH = 9.47, SiO₂ ≈ 30 wt%) as the silica source, was directly synthesized in a eutectic mixture where silicalite-1 grains were formed in the three-dimensional net structure of silica gel, grown *in situ* by transforming amorphous SiO₂ into an MFI-type structure and coated with amorphous SiO₂. The alkalinity, template agent, and crystallization time strongly affect the physicochemical properties of SiO₂@silicalite-1. The physicochemical properties of these samples were characterized by scanning electron microscopy (SEM), X-ray diffraction (XRD), transmission electron microscopy (TEM), UV-vis diffuse reflectance spectra and nitrogen adsorption. The results show that the SiO₂@silicalite-1 is synthesized in a eutectic mixture and its physicochemical properties can be tuned by controlling the content of sodium hydroxide and tetrapropyl ammonium bromide (TPABr). A time-dependent study reveals that the formation process obeys an *in situ* epitaxial growth and phase transformation mechanism. Finally, SiO₂@silicalite-1 was used as the support to prepare TiO₂-loaded SiO₂@silicalite-1 (TiO₂@SiO₂@silicalite-1). After five loading procedures, it could load 0.44% TiO₂ nanoparticles, which is higher than the TiO₂ nanoparticle loading in traditional silicalite-1 (0.13%). When the catalysts were used as a catalyst for the degradation of rhodamine B (RhB) aqueous solution under UV light, the photocatalytic efficiency of TiO₂@SiO₂@silicalite-1 (89.2%) is higher than TiO₂@silicalite-1 (only 34.6%). The rate of degradation using TiO₂@SiO₂@silicalite-1 is 4.3 times faster than that using TiO₂@silicalite-1. Furthermore, SiO₂@TiO₂@silicalite-1 exhibits high stability of photocatalytic performance. After five repeated cycles, the photocatalytic efficiency of TiO₂@SiO₂@silicalite-1 is 88.09%, which reduces only by 1.1%.

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1. Introduction

Zeolites are microporous crystalline materials, which have high surface areas, excellent hydrothermal stability, controllable porosity and acid/basic/redox properties. Many zeolites have been discovered and used in industrial applications,^{1–5} such as NaX,⁶ MCM-41,⁷ 5A,⁸ zeolite Y,⁹ ZSM-5.¹⁰ They are composed of SiO₄ and AlO₄ tetrahedral structures and have been widely used as adsorbents,^{11–13} ion exchangers and catalysts in the chemical industry.^{14,15} Among these applications, zeolites were widely used as supporters in many fields. For example, Yang *et al.* employed SBA-15 as a support to load nanocrystalline titanium dioxide (TiO₂) in its pores, which showed much higher photo-degradation ability for methylene blue than commercial pure P25 TiO₂ nanoparticles.¹⁶ Panpranot *et al.* used MCM-41 as the support to develop MCM-41-supported CoRu catalysts for CO hydrogenation, in which Co and Ru were well loaded on the pores of MCM-41 and showed high activity.¹⁷

ZSM-5, a kind of crystalline microporous aluminosilicate with the MFI-type framework, has a strong and tunable acidity,

excellent shape selectivity and good hydrothermal stability, leading to their wide use as sorbents,^{18–20} catalysts, such as petrochemical processing for improving gasoline quality and producing important chemical materials.^{21–23} The properties of ZSM-5 zeolite are greatly influenced by the Al₂O₃/SiO₂ ratio and the morphology.²⁴ Recently, the high-silica ZSM-5 zeolite (silicalite-1) has received much attention because of its high shape selectivity for molecular sorption and diffusion, catalytic applications.^{25–27}

Generally, silicalite-1 is synthesized as powders. It consists of pure SiO₂ and has high surface areas. Accordingly, it is a potential to load some materials on silicalite-1, such as TiO₂ nanoparticles. In our previous research, the surface of silicalite-1 is smooth and the cavity of silicalite-1 is about 0.5–0.55 nm.²⁸ So it is difficult to load some materials on its surface firmly or in its pores due to its smooth surface and small cavity. What is more, when silicalite-1 is loaded with some materials in its pores, its surface area will decrease obviously. It can influence its catalytic activity significantly. In order to meet growing environmental challenges, it is important to explore new technologies to treat or synthesize silicalite-1 which can load a large amount of other materials and keep its large surface area.

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In this work, we demonstrate a sol-gel method for one-step synthesis of $\text{SiO}_2\text{@silicalite-1}$ which has high crystallinity and rough surface. We also explored the formation of silicalite-1. Silicalite-1 grains are formed by transforming amorphous SiO_2 into MFI-type structure. The surface of $\text{SiO}_2\text{@silicalite-1}$ is coated with amorphous SiO_2 nanoparticles and full with grooves. The amorphous SiO_2 and grooves can provide more spaces for some materials to load on the surface of silicalite-1. Finally, $\text{SiO}_2\text{@silicalite-1}$ was used as support to load TiO_2 nanoparticles. The catalytic activity of the catalyst was investigated by the photocatalytic degradation of RhB.

2. Experimental

2.1. Synthesis of $\text{SiO}_2\text{@silicalite-1}$

All chemicals were purchased from Sinopharm Chemical Reagent Co., Ltd (analytical grade) and used as received without further purification. The samples were synthesized by first adding X g (where X is 0.5, 1, 1.5, respectively) tetrapropyl ammonium bromide (TPABr) and then Y g (where Y is 0.1, 0.2, 0.3, respectively) sodium hydroxide to 30 ml silica sol (30 wt%, pH = 9.46) under vigorous stirring for 2 h at room temperature. After that, the gel mixture was transferred to the Teflon-lined stainless-steel autoclave and heated at 180 °C in the baking oven for different time (3 h, 12 h, 24 h). The schematic diagram of the synthetic procedure of core-shell $\text{SiO}_2\text{@silicalite-1}$ is shown in Fig. 1. The obtained gelatinous monoliths were slightly poured out and dried at 60 °C for 48 h. Finally, all samples were calcined at 550 °C for 6 h with the heating rate of 5 °C min⁻¹ in the air atmosphere to remove the templates. And then, the core-shell $\text{SiO}_2\text{@silicalite-1}$ was grinded into powders.

2.2. Synthesis of $\text{TiO}_2\text{@SiO}_2\text{@silicalite-1}$

$\text{SiO}_2\text{@silicalite-1}$ was used as support to load TiO_2 nanocrystalline by the sol-gel method. The precursor solution of

TiO_2 was prepared as follows: 11.2 ml tetrabutyl titanate (TBOT) and 2 ml acetic acid were dissolved in 33.8 ml absolute ethanol. The solution was stirred vigorously at room temperature for 20 min, followed by the addition of a mixture containing 4.5 ml distilled water, 10 ml ethanol and 0.5 ml nitric acid. Then, the solution was stirred continuously at room temperature to hydrolyze for 5 minutes to obtain a TiO_2 sol. After that, 5 g $\text{SiO}_2\text{@silicalite-1}$ powders was immersed into the above solution. And the solution kept stirring for 20 minutes at room temperature. Then, the TiO_2 -loaded $\text{SiO}_2\text{@silicalite-1}$ ($\text{TiO}_2\text{@SiO}_2\text{@silicalite-1}$) was filtrated and dried at drying oven and, subsequently, the process was repeated from immersion to drying for 5 times. Finally, $\text{TiO}_2\text{@SiO}_2\text{@silicalite-1}$ was calcined at 550 °C in air for 2 h, which resulted in the preparation of $\text{TiO}_2\text{@SiO}_2\text{@silicalite-1}$ catalyst. In order to draw a comparison, the traditional silicalite-1 was used as support to load TiO_2 at the same process to prepare the $\text{TiO}_2\text{@silicalite-1}$ catalyst.

2.3. Catalyst characterization

The surface morphology of all ground samples was observed using scanning electron microscopy (SEM) on Hitachi SU-70 microscope at acceleration voltage of 5 kV. The X-ray powder diffraction patterns were recorded on a Bruker-AxsD8 diffractometer using Cu-K α radiation (40 kV and 40 mA) in the angular range (2θ) from 5 to 60°. Transmission electron microscopy (TEM) images were taken using a TEM (JEOL Model JEM-2100) operated at an accelerating voltage of 100 kV using a 20 μm aperture. UV-vis diffuse reflectance spectra was recorded on a UV-101 PC scanning spectrophotometer. The nitrogen adsorption-desorption isotherm were recorded using TriStar 3020 (Micromeritics, USA) analyzer. The specific surface areas were calculated using the multiple-point Brunauer-Emmett-Teller (BET) method by the Barrett-Joyner-Halenda (BJH) model. The micropore size distribution was measured using the Horvath-Kawazoe (HK) analysis method.

2.4. Catalytic evaluation

In order to evaluate the catalytic activity of the catalyst, all samples were used to degrade the rhodamine B (RhB) solution in an enclosed box with side the volume of 25 × 25 × 25 cm³. A Philips lamp irradiating UV light with the power of 40 W was fixed on the top of the box. The distance between the UV light and the reactor is 26 cm. And a magnetic stirrer was located at the bottom of the box. Before UV light irradiation, the solution was stirred continuously in the dark for 30 min to ensure the adsorption-desorption equilibrium between RhB and the catalyst. Afterwards, the UV light was turned on and the solution was irradiated for 90 min. In order to determine the residual dye concentration in solution during this time, a few milliliters of the solution was withdrawn from the reactor and centrifuged to remove the photocatalyst before measuring. The fractional degradation efficiency (D_R) of RhB was calculated by the following equation:

$$D_R (\%) = \frac{(C_0 - C_t)}{C_0} \times 100$$

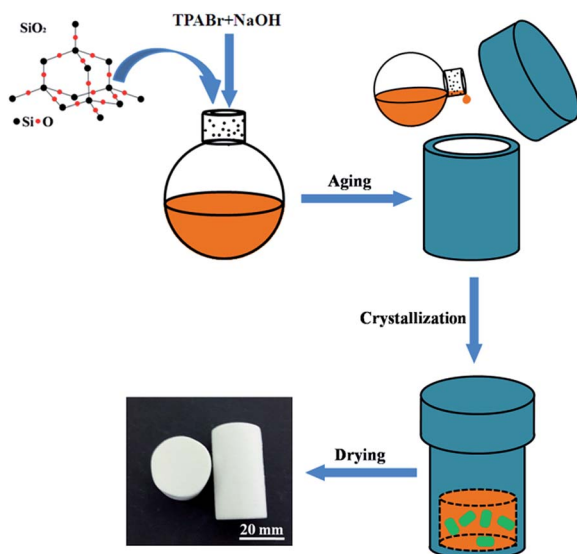


Fig. 1 Schematic diagram of the synthetic procedure and the synthesized $\text{SiO}_2\text{@silicalite-1}$.



where C_0 is the initial concentration of RhB and C_t is the concentration at definite interval of time, respectively.

3. Results and discussion

3.1. Preparation of SiO₂@silicalite-1 and its characterization

3.1.1. Effects of alkalinity. SiO₂@silicalite-1 was synthesized by sol-gel method. The effect of alkalinity on the morphology and crystallinity of SiO₂@silicalite-1 was investigated. The sol composition was composed of 30 ml silica sol, 1 g TPABr, X g NaOH (where X was 0.1, 0.2, and 0.3, respectively) and crystallized at 180 °C for 24 h. Fig. 2 shows the SEM images and XRD patterns of products prepared with different NaOH/SiO₂ mass ratios. From the XRD patterns of Fig. 2, it can be seen that three samples show distinct peaks at 7.98°, 8.82°, 13.24°, 13.94°, 14.80°, 15.53°, 15.89°, 17.79°, 20.37°, 23.18°, 24.02° and 24.46°, which are ascribable to (101), (020), (300), (012), (301), (202), (040), (031), (501), (151) and (303) reflections of silicalite-1. It indicates that silicalite-1 is formed in three samples.^{29,30} Fig. 2a–c are SEM images of three samples. It can be seen that all silicalite-1 grains exhibit the twinned structure. And the particle size of silicalite-1 has a close relative with the sodium hydroxide content. When the sample is synthesized with the NaOH/SiO₂ mass ratio of 0.1 : 9, it is composed of silica and silicalite-1 crystals. The particle size of silicalite-1 is about 40 μm (Fig. 2a). With the increase of the NaOH/SiO₂ mass ratio (from 0.1 : 9 to 0.3 : 9), the particle size of silicalite-1 particles decreases (from 40 μm to 25 μm). It indicates that more nuclei are formed at the higher sodium hydroxide content and relative smaller crystals are produced. It is in agreement with Watanabe's group experiments.³¹ Fegan's group has also reported that crystals formed at high alkalinity are smaller.³² Except silicalite-1 particles, there is also existing amorphous silica around silicalite-1 particles. Fig. 2 also shows the XRD patterns of the three samples. The intensity of diffraction peaks of silicalite-1 increases when the NaOH/SiO₂ previous work, we successfully synthesized the silicalite-1 with high crystallinity.²⁸ Using the synthesized silicalite-1 as comparison, the

crystallinity degree of SiO₂@silicalite-1 is 68.95%, 82.71%, 77.19%, respectively. It indicates that when the NaOH/SiO₂ mass ratio is 0.2 : 9, it is beneficial to synthesize the MFI-type structure. The result is in line with the observations reported in the literature.³³

3.1.2. Effects of tetrapropyl ammonium bromide (TPABr).

In this experiment, we investigated the effect of the TPABr content on synthesizing the SiO₂@silicalite-1. The products were synthesized at 180 °C for 24 h with a sol composition of 30 ml silica sol, 0.2 g NaOH, X g TPABr (X is 0.5, 1.0, and 1.5, respectively). Fig. 3 shows the SEM images and XRD patterns with different TPABr content. It can be noted that when the product is synthesized with 0.5 g TPABr, there are a lot of amorphous SiO₂ particles around silicalite-1 particles. The crystallinity degree is only 28.75%. Except the characteristic peaks of silicalite-1, the XRD patterns of the sample with 0.5 g TPABr also shows a wide range diffraction from 20 to 25°, which is corresponding to amorphous SiO₂. The average particles size of silicalite-1 as estimated by SEM is about 50 μm. When the TPABr content is 1.0 g, the average particles size of silicalite-1 is about 40 μm and the crystallinity degree is 82.71%. When the TPABr content increases further (1.5 g), the average particles size of silicalite-1 decreases continuously (20 μm). And the crystallinity degree of silicalite-1 is increased to 88.11%. TPABr content can promote the formation of silicalite-1 nuclei.

3.1.3. Effects of crystallization time. By controlling the amount of sodium hydroxide and TPABr in the synthesis gel, we obtained the suitable alkalinity and TPABr content on synthesizing the SiO₂@silicalite-1. In this experiment, the composition of the precursor solution was 30 ml silica sol, 0.2 g NaOH, 1.5 g TPABr. It was aging for 2 h at room temperature and crystallized at 180 °C for different time. The effect of crystallization time was examined. Fig. 4 shows the SEM images and XRD patterns of products crystallized with different crystallization time that fell within the 3–24 h range.

Fig. 4a is the SEM image of the product which is crystallized for 3 h. It can be seen that silicalite-1 particles are not formed. The XRD patterns just shows weak characteristic peaks of

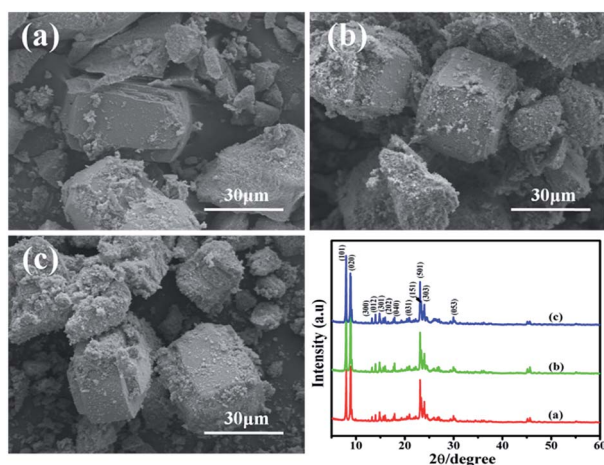


Fig. 2 SEM images of products prepared with NaOH/SiO₂ mass ratios of (a) 0.1 : 9, (b) 0.2 : 9, (c) 0.3 : 9 and XRD patterns.

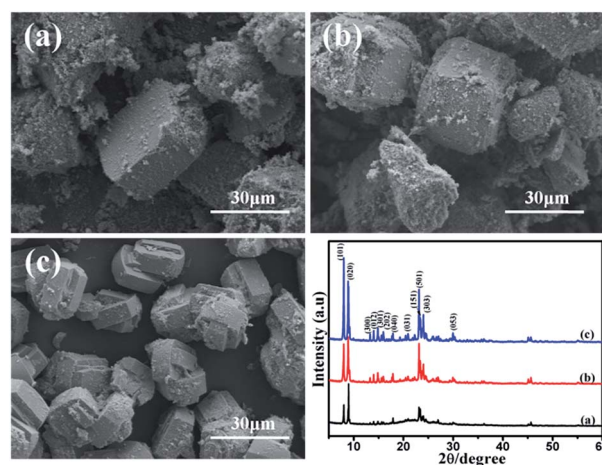


Fig. 3 SEM images of products prepared with different amount of TPABr (a) 0.5 g, (b) 1.0 g, (c) 1.5 g and XRD patterns.



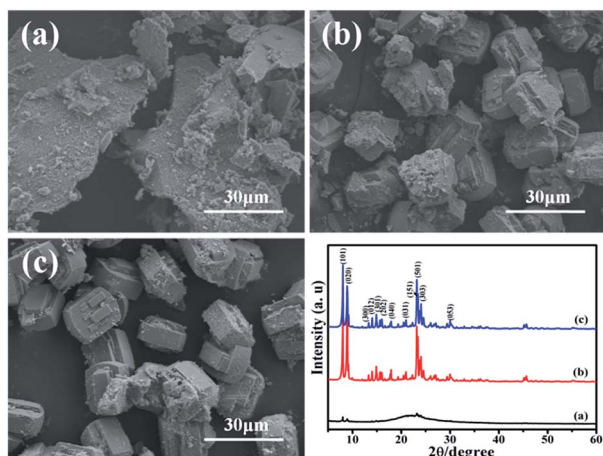


Fig. 4 SEM images of products prepared with a crystallization time of (a) 3 h, (b) 12 h, (c) 24 h and XRD patterns.

silicalite-1 and exhibits a wide range diffraction from 20 to 25°, which is corresponding to amorphous SiO₂. It can be inferred that SiO₂ colloidal particles just aggregate and form gelation with the three-dimensional net structure after 3 h of hydrothermal treatment.³⁴ The solvent and solute are immobilized in the three-dimensional net structure and loses its mobility. After being dried, water molecules are evaporated and SiO₂ colloidal particles are left and bond together. When the product is synthesized by crystallizing for 12 h, it shows the obvious characteristic peaks of silicalite-1. And silicalite-1 particles exhibit the twinned-lozenge shape and the average crystal size of the harvested particles is about 20 μm (Fig. 4b). The crystallinity degree is about 117.76%. Further extension of the crystallization time (after 24 h), the particle size shape and crystal size of silicalite-1 are almost the same with the sample crystallizing for 12 h (Fig. 4c). But the crystallinity degree decreases to 88.11%. It indicates that 12 h is the suitable time to form silicalite-1 zeolite.

3.2. Preparation of TiO₂@SiO₂@silicalite-1 and its characterization

In order to draw a comparison, the traditional silicalite-1 was synthesized according to our previous report.²⁸ In a typical process, 40 ml of silica sol (30 wt%, pH = 9.46) and 60 ml of distilled water were mixed together. Then 1.5 g of sodium hydroxide and 3 g of tetrapropylammonium bromide were added into above mixture. After the mixture was stirred at room temperature for 6 h, the obtained gel precursor was transformed into the Teflon-lined stainless-steel autoclave and heated at 180 °C for 24 h to obtain the silicalite-1. Fig. 5 illustrates the high magnification SEM images of silicalite-1 and SiO₂@silicalite-1. It can be seen that the surface of the traditional silicalite-1 is smooth (Fig. 5a). Fig. 5b is the high magnification SEM images of SiO₂@silicalite-1. It can be seen that the surface of SiO₂@silicalite-1 coats with nanoparticles. The size of the nanoparticles is about 50 nm. Except these nanoparticles, there are also existing lots of grooves.

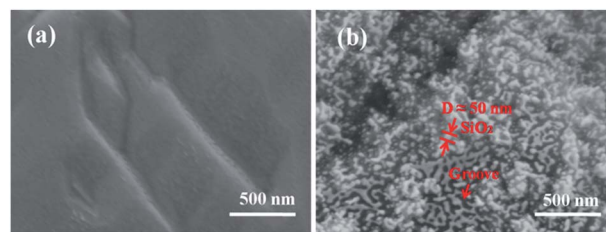


Fig. 5 High magnification SEM images of (a) silicalite-1 and (b) SiO₂@silicalite-1.

In order to investigate the physicochemical properties of the SiO₂@silicalite-1 (30 ml silica sol, 0.3 g NaOH, 1.5 g TPABr, crystallized for 12 h), BET techniques was employed. As references, silicalite-1 (15 μm) and SiO₂@silicalite-1 were characterized. The nitrogen sorption isotherms and pore size distributions are shown in Fig. 6. It can be seen that the nitrogen sorption isotherms of silicalite-1 and SiO₂@silicalite-1 are both type I, which shows a steep increase at low relative pressure ($P/P_0 < 0.1$) region corresponding to the presence of micropores in the samples.³⁵ Silicalite-1 and SiO₂@silicalite-1 almost keep the same micropore size distribution and have a very narrow micropore size distribution with a mean pore diameter of 0.5–0.55 nm, which is comparable to the pore diameter of the crystallographic five membered ring of silicalite-1.

The textural properties of the synthesized silicalite-1 and SiO₂@silicalite-1 are shown in Table 1. The particle size of silicalite-1 and SiO₂@silicalite-1 are 15 and 20 μm, respectively. The BET surface area and external surface area of silicalite-1 are 334 m² g^{−1} and 195 m² g^{−1}, respectively. When the sample is formed into SiO₂@silicalite-1, it keeps the high BET surface area (279 m² g^{−1}).

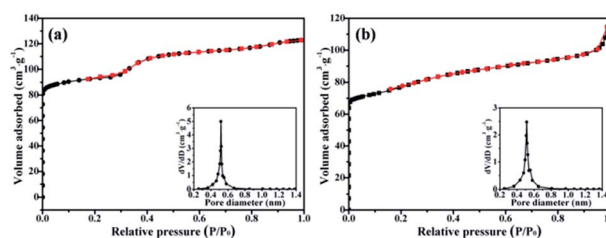


Fig. 6 Nitrogen sorption isotherms and pore size distribution of (a) silicalite-1 and (b) SiO₂@silicalite-1.

Table 1 Textural characteristic of silicalite-1 and SiO₂@silicalite-1

Catalyst	Particle size ^a (μm)	S _{BET} ^b (m ² g ^{−1})	S _{EXT} ^c (m ² g ^{−1})	V ^d (ml g ^{−1})
Silicalite-1	15	334	195	0.190
SiO ₂ @silicalite-1	20	279	161	0.177

^a Particle size: estimated based on SEM results. ^b S_{BET}: BET surface area.

^c S_{EXT}: external surface area. ^d V: micropore volume.



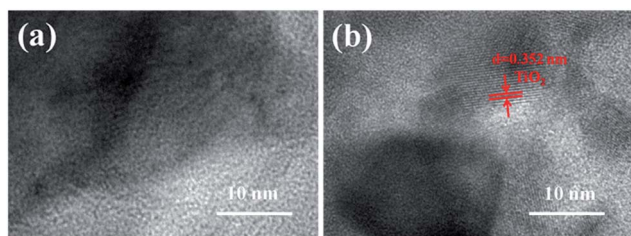


Fig. 7 TEM images of (a) SiO_2 @silicalite-1 and (b) TiO_2 @ SiO_2 @silicalite-1.

Fig. 7 shows the TEM images of SiO_2 @silicalite-1 and TiO_2 @ SiO_2 @silicalite-1. The SiO_2 @silicalite-1 was synthesized with the sol composition of 30 ml silica sol, 0.2 g NaOH, and 1.5 g TPABr. The high magnification SEM image of SiO_2 @silicalite-1 (Fig. 5a) shows that the surface of SiO_2 @silicalite-1 is coated with nanoparticles. From Fig. 7a, it can be seen that these nanoparticles are amorphous and aggregate together. It can be inferred that silicalite-1 grains are formed in the three-dimensional net structure of silica and grow *in situ* by transforming amorphous SiO_2 into MFI-type structure. Fig. 7b is the high magnification TEM image of TiO_2 @ SiO_2 @silicalite-1. It can be seen that there are existing diffractions with lattice constant of 0.352. It indicates that TiO_2 is well loaded in the pores of amorphous SiO_2 .

In the preparation of TiO_2 -loaded SiO_2 @silicalite-1 (TiO_2 @ SiO_2 @silicalite-1), we used the SiO_2 @silicalite-1 as support, which was synthesized with 30 ml silica sol, 0.2 g NaOH, 1.5 g TPABr and crystallized for 12 h at 180 °C. The synthetic procedure is given as above. Fig. 8 shows that diffuse reflectance UV-vis spectra of SiO_2 @silicalite-1, TiO_2 @silicalite-1, and TiO_2 @ SiO_2 @silicalite-1. It can be seen that SiO_2 @silicalite-1 just shows a gentle absorption band at about 250–280 nm. When silicalite-1 and SiO_2 @silicalite-1 are loaded with TiO_2 nanoparticles, they all exhibit a highly intense band at about 388 nm. According to the Kubelka–Munk radiative transfer model, the band gap (E_g) of TiO_2 @silicalite-1 and TiO_2 @ SiO_2 @silicalite-1

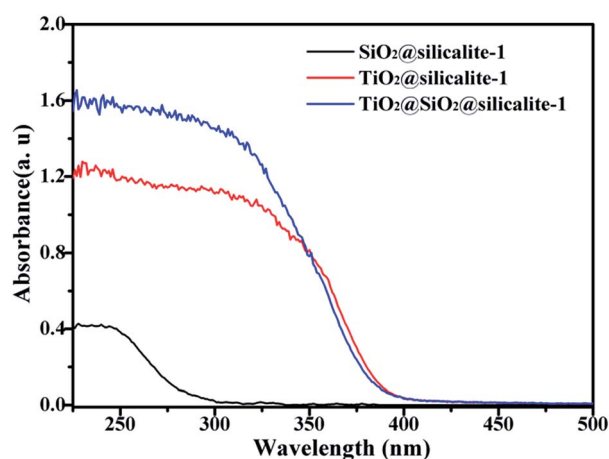


Fig. 8 UV-vis spectra of silicalite-1, TiO_2 @silicalite-1 and TiO_2 @ SiO_2 @silicalite-1.

are about 3.2 eV.³⁶ It indicates that TiO_2 nanoparticles are successfully loaded silicalite-1 and SiO_2 @silicalite-1. However, from Fig. 8, it can be found that the absorption peak of TiO_2 @ SiO_2 @silicalite-1 corresponding to TiO_2 is higher than TiO_2 @silicalite-1. It can be inferred that content of TiO_2 in TiO_2 @ SiO_2 @silicalite-1 may be higher than in TiO_2 @silicalite-1.

The catalytic performance of all samples was examined by degradation of RhB. Fig. 9a shows the results of RhB removal by the catalysts. In order to ensure the adsorption-desorption equilibrium between RhB and the catalyst, all samples were mixed with RhB solution and stirred in the dark. After 30 min of continuously stirring in the dark, 5.24% of RhB is absorbed by SiO_2 @silicalite-1. When SiO_2 @silicalite-1 is loaded with TiO_2 , it shows 4.75% absorption. The adsorptive property of the catalyst is reduced, which may be contributed to the loading of TiO_2 in the pores of SiO_2 @silicalite-1. The absorption of TiO_2 @silicalite-1 is 5.22%, which is higher than TiO_2 @ SiO_2 @silicalite-1. It may be the reason that the surface area of TiO_2 @silicalite-1 is higher than the surface of TiO_2 @ SiO_2 @silicalite-1. Then, all experiments are exposed under the UV light for 120 min. From Fig. 9a, we can see that 4.68% of RhB is degraded by itself under the UV light. The existing of SiO_2 @silicalite-1 powders cannot accelerate the degradation of RhB. When SiO_2 @silicalite-1 is loaded with TiO_2 nanoparticles, it can degrade 89.19% of RhB. It shows high photocatalytic activity. And the traditional silicalite-1 was used as support to load TiO_2 at the same process to prepare the TiO_2 @silicalite-1 catalyst. Only 34.58% of RhB is degraded by TiO_2 @silicalite-1 catalyst. The catalytic activity of TiO_2 @ SiO_2 @silicalite-1 is higher than TiO_2 @silicalite-1. It may be the reason that SiO_2 @silicalite-1 can load more TiO_2 than the traditional silicalite-1. The kinetics of all samples for degradation of RhB under UV light irradiation is illustrated in Fig. 9b. The photocatalysis degradation kinetic reaction can be described by pseudo-first-rate order kinetics, $\ln(C_0/C) = kt$, where k is a pseudo-first-rate kinetic constant and t is the irradiation time. The calculated k value of TiO_2 @silicalite-1 is 0.00384 min^{-1} . For SiO_2 @ TiO_2 @silicalite-1, the calculated k value is 0.0165 min^{-1} , which is 4.3 times the k value of TiO_2 @silicalite-1. The photocatalysis activity of the composite is improved greatly. Accordingly, the change of the photocatalytic activity of TiO_2 @ SiO_2 @silicalite-1 catalyst is probably due to the increase of TiO_2 loading in the catalyst. More TiO_2 nanoparticles can be loaded in the pores of amorphous SiO_2 and the grooves of silicalite-1.

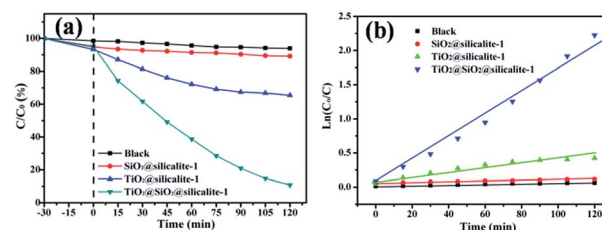


Fig. 9 Photocatalytic degradation (a) and kinetics (b) of blank, silica-1, TiO_2 @silicalite-1 and TiO_2 @ SiO_2 @silicalite-1 for degradation of RhB in the dark and under UV light irradiation with a catalyst content of 1 g l^{-1} , $C_0 = 5 \text{ mg l}^{-1}$.



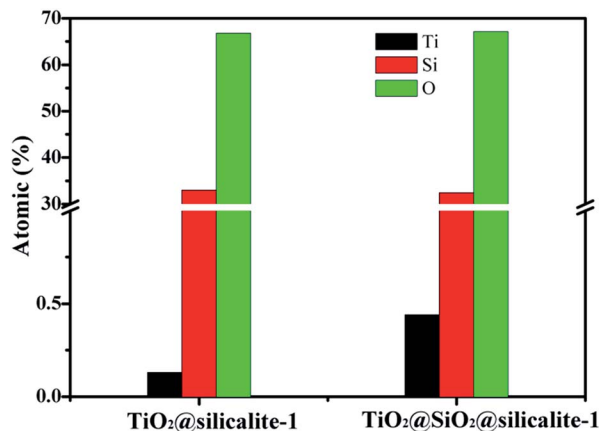


Fig. 10 The content of Ti, Si, O element in TiO₂@silicalite-1 and TiO₂@SiO₂@silicalite-1.

The energy dispersive spectrometer of TiO₂@silicalite-1 and TiO₂@SiO₂@silicalite-1 was measured by a Hitachi SU70 microscope and the element proportion was shown in Fig. 10. It can be seen that the atomic ratio of O/Si of TiO₂@silicalite-1 and TiO₂@SiO₂@silicalite-1 is about 2. After five times of loading, the Ti element in TiO₂@silicalite-1 is just 0.13%. However, the Ti element in TiO₂@SiO₂@silicalite-1 is 0.44%, which is almost 3.48 times the Ti content of TiO₂@silicalite-1.

More TiO₂ nanoparticles can be loaded in the pores of amorphous SiO₂ and grooves on silicalite-1 surface.

In order to investigate the stability of photocatalytic performance under UV light, TiO₂@SiO₂@silicalite-1 was used as catalyst to degrade RhB in five repeated cycles. The results are shown in Fig. 11. From Fig. 11, it can be seen that the photocatalytic performance of SiO₂@TiO₂@silicalite-1 exhibits effective photostability under the UV light. After five repeated cycles, the photocatalytic efficiency of TiO₂@SiO₂@silicalite-1 is 88.09%, which reduces only by 1.1%. It may be the reason that TiO₂ nanoparticles load in the pores of amorphous of SiO₂ and tightly adhere on the surface of silicalite-1.

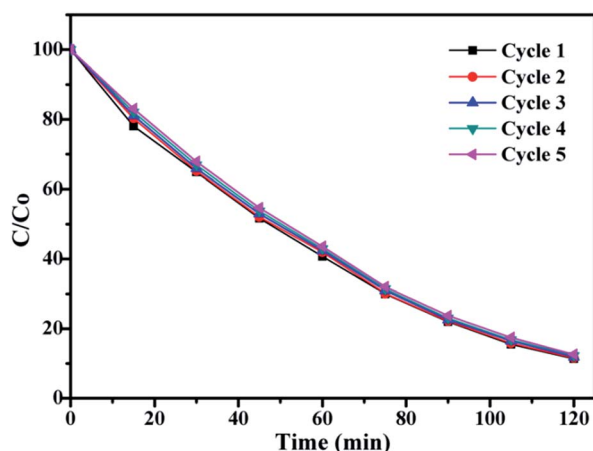


Fig. 11 Five photocatalytic degradation cycles of RhB using TiO₂@SiO₂@silicalite-1 under UV light.

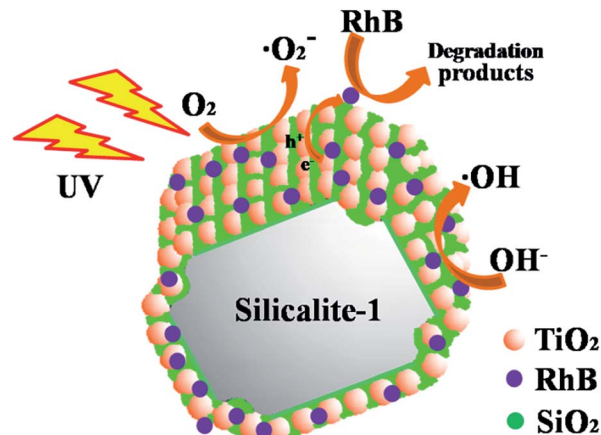


Fig. 12 The schematic diagram for the adsorption and photocatalytic degradation of RhB molecules on catalyst.

Fig. 12 shows the schematic diagram for the adsorption and photocatalytic degradation of RhB molecules on catalyst. As we know silicalite-1 zeolite has high adsorptive property. It can keep a high RhB concentration on its surface for its adsorption. After a period of contact, there is an adsorption-desorption equilibrium between the catalyst and the solution. When TiO₂ particles are loaded on SiO₂@silicalite-1 and irradiated by UV light, it can generate the hydroxyl radicals and valence band holes to oxidize RhB molecules into inorganic compounds. The adsorption of the zeolites and photocatalysis of the deposited TiO₂ have been combined, resulting in synergistic effects in improving the photocatalytic efficiency of TiO₂. When SiO₂@silicalite-1 monolith is loaded with TiO₂ particles, it can degrade RhB molecular effectively. From Fig. 9a, it can be seen that after 120 min of UV light irradiation, 89.19% of RhB is degraded. The catalyst shows high photocatalytic activity.

4. Conclusions

A series of SiO₂@silicalite-1 with different silicalite-1 particle size and crystallinity degree were synthesized according to the hydrothermal method using silica sol as Si source, TPABr as template. By adding an appropriate amount of tetrapropyl ammonium bromide and sodium hydroxide in the silica sol and controlling the crystallization time, the crystallinity degree of silicalite-1 monolith can be 117.76% comparing to the synthesized silicalite-1. SiO₂@silicalite-1 containing TiO₂ (TiO₂@SiO₂@silicalite-1) was also prepared by a sol-gel method. TiO₂ was well loaded on the pores of amorphous SiO₂ and the grooves of silicalite-1 surface. And it shows high catalytic performance for degradation of RhB and keeps high stability of photocatalytic performance. SiO₂@silicalite-1 is potential to be used as support to load same materials using in different catalysis.

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