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Natural phosphorus-doped honeycomb carbon materials as oxygen reduction catalysts derived from *Pulsatilla chinensis* (Bunge) Regel

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Facile phosphorus (P)-doped carbon materials based on *Pulsatilla chinensis* (Bunge) Regel (PCR) were directly used in oxygen reduction reaction (ORR). The half-peak potential of PCR carbonized at 1000 °C (PCR-1000) was 30 mV lower than that of the commercial Pt/C (20%) catalysts. The PCR-1000 showed 87.0% of the initial current density after 20 000 s stable testing at the rotation of 1600 rpm in O₂-saturated 0.1 M KOH. This study provides a method for the preparation of P-doped ORR catalysts derived from biomass. In addition, the honeycomb structure carbon materials can also provide a carrier for other applications including composite catalysts in fuel cells and electrode materials in supercapacitors.

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Introduction

Fuel cells are considered as a renewable energy technology and have been widely researched in recent years for their merits of high energy density and zero emission.¹ However, the progress of fuel cells is limited due to the sluggishness of the ORR in the cathode.^{2,3} At present, platinum and its alloys are the most effective ORR catalysts. However, their applications are restricted because of their high-cost. Therefore, the development of inexpensive and alternative cathodic catalysts is desirable in the area of fuel cells. Based on these conditions, many researchers developed many metal-free catalysts. Nitrogen-doped carbon materials were extensively used for their high ORR activity in ORR since the first report of high ORR activity of vertically aligned N-doped carbon nanotubes by Dai and co-workers in 2009.⁴ Phosphorus has the same number of valence electron as nitrogen; in addition, the larger atomic size and lower electronegativity of P can introduce a defect-induced active surface and local charge density through its lone pair electrons in the 3p orbital, which should be facilitated by O₂ adsorption in the ORR.^{1,5} However, the development of P-doped carbon materials in ORR is sluggish compared with the quick development of nitrogen-doping. Recently, some phosphorus sources such as phosphate,⁶ triphenylphosphine,⁷⁻⁹ phosphoric acids,¹⁰⁻¹² and ionic liquid¹³ were used for P-doped carbon materials. P-doped graphite layers and P-doped graphitic carbon nanospheres all exhibited high electrocatalytic activity. However, their tedious synthesis process and low P-doped content restricted their industrial applications. Although different chemical methods were used for the preparation of P-doped catalysts, the

results were unsatisfactory for difficult doping and not retaining the P element at high temperature. Apart from this point, Strelko *et al.* researched the mechanism of the influence of P doping in a carbon matrix and, from a theoretical study of the band gap energy, found that the dopant content was related to the catalytic activity.¹⁴ We found that the content with better ORR catalytic effects in P-doped carbon materials was not larger than 3 at% in many previous reports. Based on these two reasons, we considered to realize P-doping from biomass.^{5-13,15} Biomass has many nutrients including different elements, especially P that is very important for biomass growth. P in the biomass is contained in various structures, which may contribute to increasing the diversity of P-doping. We considered that the natural P-doping in biomass may obtain the uniform structures and different phosphorus groups. At present, there are few reports on the P-doped catalysts from biomass including yeast cells⁵ and coconut shells¹⁵ as precursors.

Herein, we reported facile P-doped carbon materials used in ORR for the first time. The PCRs in this study were obtained from Xiao Hinggan Mountains in the northeast of China, which were distributed in hills, slopes or fields. The carbonization of the PCR biomass material is a low-cost, simple, and readily scalable approach for the synthesis of P-doped honeycomb carbon materials. In this study, the larger phosphorus content in the carbonization of the PCR biomass was 1.46 at%, which was compared with the contents obtained from chemical P-doping methods. Moreover, the carbonized PCR exhibited superior properties in the ORR. The half-peak potential of PCR-1000 was 30 mV lower than that of the commercial Pt/C (20%) catalyst, and 87.0% of the initial current density was maintained after 20 000 s of stable testing at the rotation of 1600 rpm in O₂-saturated 0.1 M KOH. These results demonstrated that the method derived from the P-doped biomass materials not only demonstrated superior properties in ORR but

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also showed the implementation of a novel idea for the difficult preparation of P-doped catalysts in fuel cells.

Experimental

PCR was obtained from Xiao Hinggan Mountains in the northeast of China (Fig. 1). The PCR biomass was dried at 100 °C for 12 h before heat treatment. The dried PCR was put in a tube furnace and then heated at 400 °C for 1 h. Subsequently, the obtained carbon materials were directly carbonized from 700 to 1000 °C for 1 h at the heating rate of 10 °C min⁻¹ in N₂ atmosphere with the flow rate of 100 ml min⁻¹. After carbonization, the samples were washed several times with distilled water and dried at 60 °C. PCRs carbonized at 700, 800, 900, and 1000 °C were denoted as PCR-700, PCR-800, PCR-900, and PCR-1000, respectively.

Scanning electron microscopy (SEM, HITACHI S-4800), Transmission electron microscope (TEM, JEM-2100HR), and X-ray diffraction (XRD, Rigaku D/Max 2500/PC) were used to characterize the structure and morphology. X-ray photo electron spectroscopy (XPS, ESCALAB 250) was used to determine the

composition. The electrochemical evaluation was carried out using an electrochemical workstation (RST5200F, Zhengzhou shiruisi Instrument Technology Co., Ltd. China). Cyclic voltammogram (CV) and rotating-disk electrode (RDE) polarization curves were obtained using a conventional three-electrode electrochemical system. Platinum wire (CHI115) and Ag/AgCl (sat.) (CHI111) electrodes were used as the counter and reference electrodes, respectively. A disk (4 mm in diameter, 0.1256 cm², Jiangsu Jiangfen Electroanalytical Instrument Co., Ltd.) was used as the working electrode to test the CV and RDE curves. Before measurements, the glassy carbon (GC) electrode was polished with gamma alumina powders (0.05 mm) until a mirror-like surface was obtained, and then washed twice with distilled water and dried in a vacuum. PCR catalysts (12.5 mg) were put in an ethanol solution (1 ml) and were treated by ultrasonication for 20 minutes. Then, 2 μl solution was placed on the surface of the GC electrode to achieve a content of 200 μg cm⁻². Finally, a drop of Nafion solution (5 wt%, DuPont) was placed onto the abovementioned surface for improving the adhesion of active materials and the GC electrode surface. The RDE polarization curves were measured by the reference electrode of Ag/AgCl (sat.) in a 0.1 M KOH solution at the scan rate of 10 mV s⁻¹ and rotation rates from 400 to 3600 rpm. CV curves were measured in the same range. O₂ was bubbled into the cell for 30 min to obtain O₂-saturated solutions before the measurements. To eliminate the influence of the background current, the CV and RDE curves were also tested in N₂-saturated solution. The current–time (*i*-*t*) curves were measured at the constant potential of -0.40 V in 0.1 M KOH solution for 20 000 s under continuous O₂ bubbling with a flow rate of 20 ml min⁻¹ at the fixed rotation rate of 1600 rpm.



Fig. 1 The schematic for *Pulsatilla chinensis* (Bunge) Regel.

Results and discussion

Two different structures were shown in Fig. 2. Honeycomb structures of PCR-700, PCR-800, and PCR-1000 carbon materials

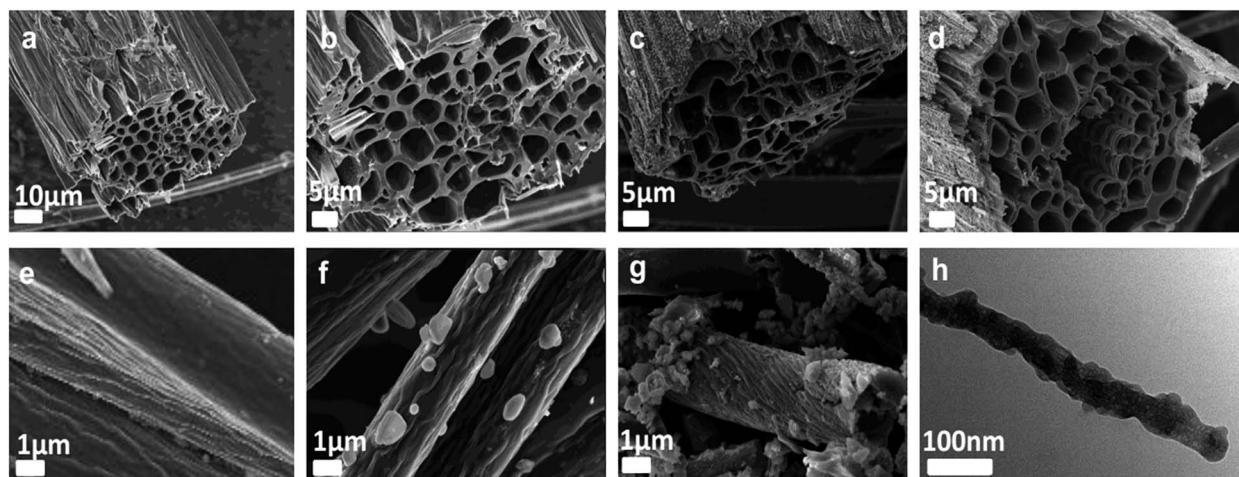


Fig. 2 SEM images: low magnification of PCR-700 (a), middle magnification of PCR-700 (b), middle magnification of PCR-800 (c), middle magnification of PCR-1000 (d), fiber structure of PCR-700 (e), fiber structure of PCR-800 (f), fiber structure of PCR-900 (g), TEM image: fiber structure of PCR-1000 (h).



were shown in Fig. 2a and b, and Fig. 2c and d, respectively. Fig. 2b was the enlargement of Fig. 2a, showing the honeycomb structure. With the increasing temperature, the surface of the honeycomb structures showed more particles, which can be reflected from the mapping images shown in Fig. 3, indicating some metal elements. Apart from the honeycomb structures, fiber structures were also observed (Fig. 2e–h), and the diameter of PCR carbon materials gradually decreased with the increasing carbonized temperature. As compared to the fibers of PCR-700 and PCR-800, PCR-900 showed more flexural folds and particles at the surface of the fibers (Fig. 2g). The diameter of honeycomb pores and the wall thicknesses are from 1 nm to 10 μm and from 500 nm to 2 μm , respectively (Fig. 2b–d). Apart from this, more particles on the PCR-1000 surface, as shown in Fig. 2h, were present in the TEM image, which can be seen clearly from its illustration corresponding to the position of the elemental mapping images of PCR-1000. Fig. 3 revealed the corresponding elemental compositions including C, O, P, Ca, and K, in which P and O are distributed in the interior of PCR-1000. As shown in Fig. 4, two obvious peaks of PCR carbon materials corresponded to the (002) and (100) reflections. Distance spaces of PCR-700, PCR-800, PCR-900, and PCR-1000 are 0.357, 0.362, 0.362, and 0.364 nm, respectively. With the increasing temperature, the degree of crystallinity of PCR carbon materials gradually decreased. The large peak width and absence of the peaks related to the interlayer order were due to the presence of defects.¹⁶ Usually, with the increasing temperature, the degree of crystallinity of carbon materials increased, but the PCR series carbon materials presented the opposite results, which can be understood by the high percentage of the edge orientation. Otherwise, the P content had a great influence on the structure of the PCR-1000, which adopted a more disordered structure corresponding to the results of XPS. Compared with other samples, PCR-1000 exhibited more defects, which facilitated the improvement of the conductivity and increased the catalytic sites. Raman spectra of PCR-700, PCR-800, PCR-900, and PCR-1000 are shown in Fig. 5. All spectra exhibited the well-known D-band and G-band peaks associated with the disordered and graphitized structures, respectively.^{17,18} The ratios of the D-band and G-band in terms of

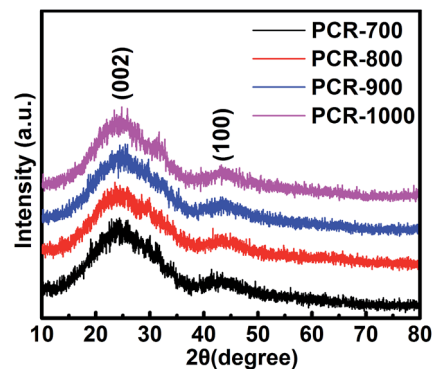


Fig. 4 XRD patterns of PCR-700, PCR-800, PCR-900, and PCR-1000.

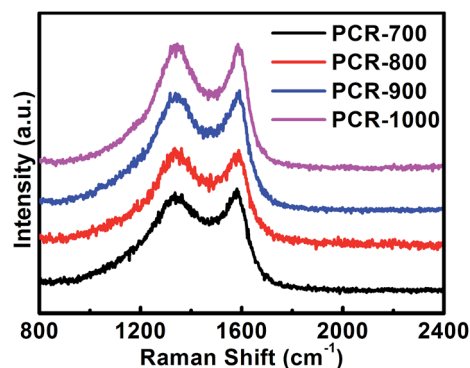


Fig. 5 Raman spectra of PCR-700, PCR-800, PCR-900, and PCR-1000.

intensity (I_D/I_G) for PCR-700, PCR-800, PCR-900, and PCR-1000 are 3.05, 3.11, 2.94, and 3.28, respectively. Generally, the I_D/I_G ratio decreased with the increasing carbonized temperature, suggesting the improvement of the crystallinity. The I_D/I_G gradually decreased until PCR-900, but PCR-1000 I_D/I_G reached the largest value in the Raman spectrum, suggesting more disorder structures, which was agreement with the XRD results.

The XPS survey revealed the presence of C, O, P, N, Ca, and K on the surface of PCR carbon materials in Fig. 6a. Table 1

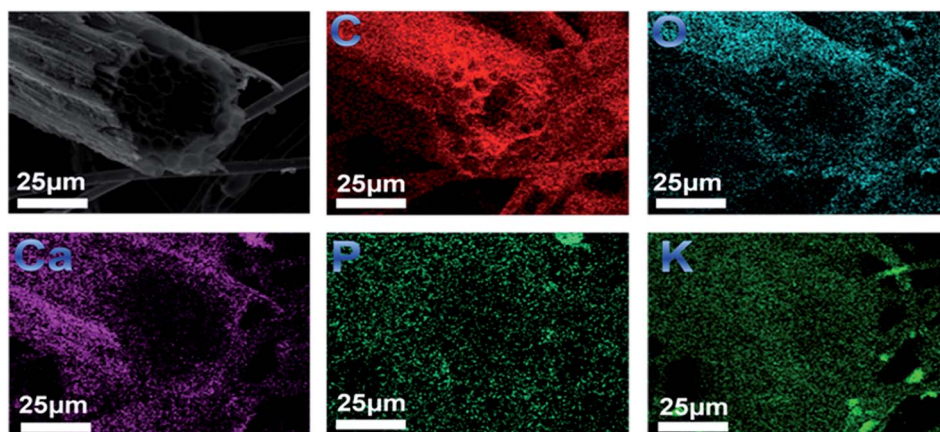


Fig. 3 Elemental mapping images of PCR-1000 including C, O, Ca, P, and K.



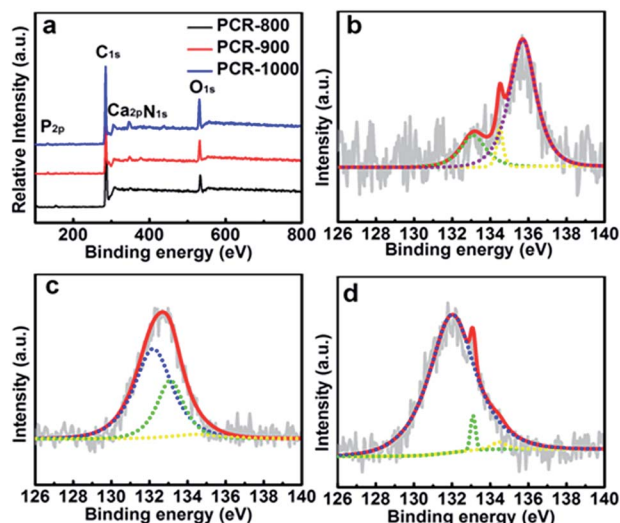


Fig. 6 (a) XPS survey of PCR-800, PCR-900, and PCR-1000. P_{2p} spectra of PCR-800 (b), PCR-900 (c), and PCR-1000 (d). Fitted lines, blue: C_3PO , green: $C-PO_3$, yellow: $C-O-PO_3$, and pink: $P-O$.

summarizes the atomic surface concentrations of different catalysts. The contents of O, P, and Ca all increased with the carbonized temperature, whereas those of C and N gradually decreased. Note that significant amount of phosphorus (from 0.34 to 1.46 at%) remained on the surface of the three catalysts. The thermal treatment increased the oxygen content from 8.95 to 14.45 at%.

The P_{2p} spectrum of phosphorus for PCR biomass carbon had been reported as C_3PO , $C-PO_3$, $C-O-PO_3$, and $P-O$ groups at the binding energy regions of 132.1, 133.1, 133.7–134.6, and 135.7 eV, respectively.^{19–21} A peak at the value around 134.2 eV was assigned to P groups bonded to a carbon site through an O atom ($C-O-PO_3$ and/or $(C-O)_3PO$).^{19,22–24} A second peak that appeared at the binding energy of about 133.1 eV indicated the existence of $C-P$ bonding as in $C-PO_3$.^{22–24} A small peak at

Table 1 Surface composition of different catalysts observed in the XPS spectra (at%)

Catalyst	Relative intensities of different elements (at%)					
	C	O	P	N	Ca	K
PCR-800	88.45	8.95	0.34	0.32	1.58	0.36
PCR-900	80.70	13.64	1.46	0.31	2.17	1.12
PCR-1000	79.95	14.45	1.26	0.27	3.48	0.59

Table 2 Relative intensities of XPS P_{2p} spectra

Catalyst	Relative intensities of different peaks (at%)			
	C_3PO	$C-PO_3$	$C-O-PO_3$	$P-O$
PCR-800	0	18.43	5.65	75.92
PCR-900	65.87	30.34	3.79	0
PCR-1000	96.12	2.39	1.49	0

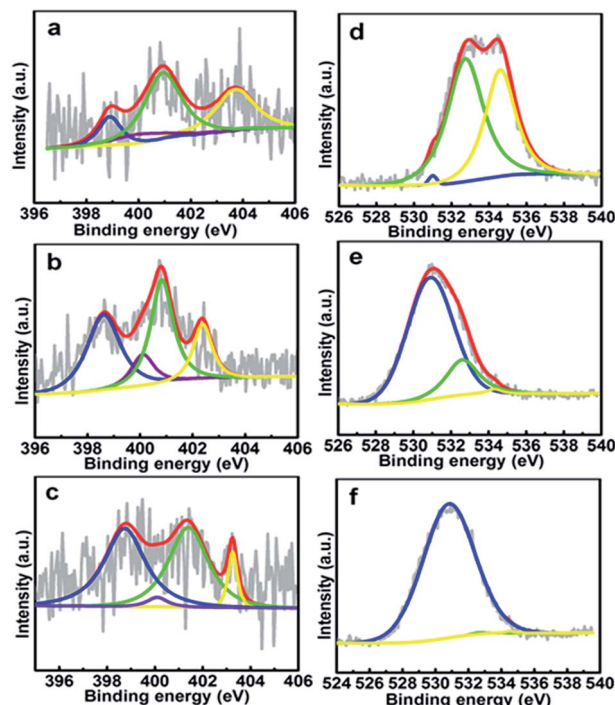


Fig. 7 N 1s spectra of PCR-800 (a), PCR-900 (b), and PCR-1000 (c), (blue: pyridinic-N, violet: pyrrolic-N, green: quaternary-N, yellow: pyridine-N-O), O 1s spectra of PCR-800 (d), PCR-900 (e), and PCR-1000 (f), blue: =O in carbonyl, carboxyl and phosphates, green: $-O-$, yellow: chemisorbed O and H_2O .

around 132.1 eV was ascribed to C_3PO groups and the peak at 135.7 eV was related to $P-O$ groups.^{19,22–24} Similar peaks have been obtained for phosphoric acid-activated carbons from kraft lignin.^{17,19} As shown in Table 2, the carbonization process caused a decomposition of the less stable $P-O$ groups and $C-O-PO_3$ groups, increasing the amount of $C-PO_3$ and C_3PO groups. Phosphorus directly bonded to the edge carbon atoms in

Table 3 Relative intensities of XPS N 1s spectra

Catalyst	Relative intensities of different peaks (at%)			
	Pyridinic-N	Pyrrolic-N	Quaternary-N	Pyridine-N-O
PCR-800	12.31	8.70	47.81	31.18
PCR-900	37.80	10.44	36.54	15.22
PCR-1000	46.42	2.58	43.82	7.18

Table 4 Relative intensities of XPS O 1s spectra

Catalyst	Relative intensities of different peaks (at%)		
	=O in carbonyl and carboxyl	$-O-$	Chemisorbed O and H_2O
PCR-800	0.85	57.18	41.97
PCR-900	80.89	18.10	1.01
PCR-1000	99.60	0.22	0.18



graphite *via* a C–P–O bond.²⁵ The results of P_{2p} spectra were in agreement with the defects analysis of XRD and Raman.

The XPS N_{1s} spectra of the three catalysts were divided into four groups. Fig. 7a–c correspond to PCR-800, PCR-900, and PCR-1000, respectively. The peaks were attributed to pyridinic-N at 398.7 ± 0.3 eV, pyrrolic-N at 400.3 ± 0.3 eV, quaternary-N at

401.4 ± 0.5 eV, and pyridinic-N–O at $402\text{--}405$ eV.²⁶ In Table 3, the variety of the four types is pyridinic-N and the proportion gradually increased from 12.31% to 46.42%, which takes more active sites in the ORR.

The XPS O_{1s} spectra of the three catalysts were divided into three groups. The peak at $530.6\text{--}530.9$ eV was attributed to an

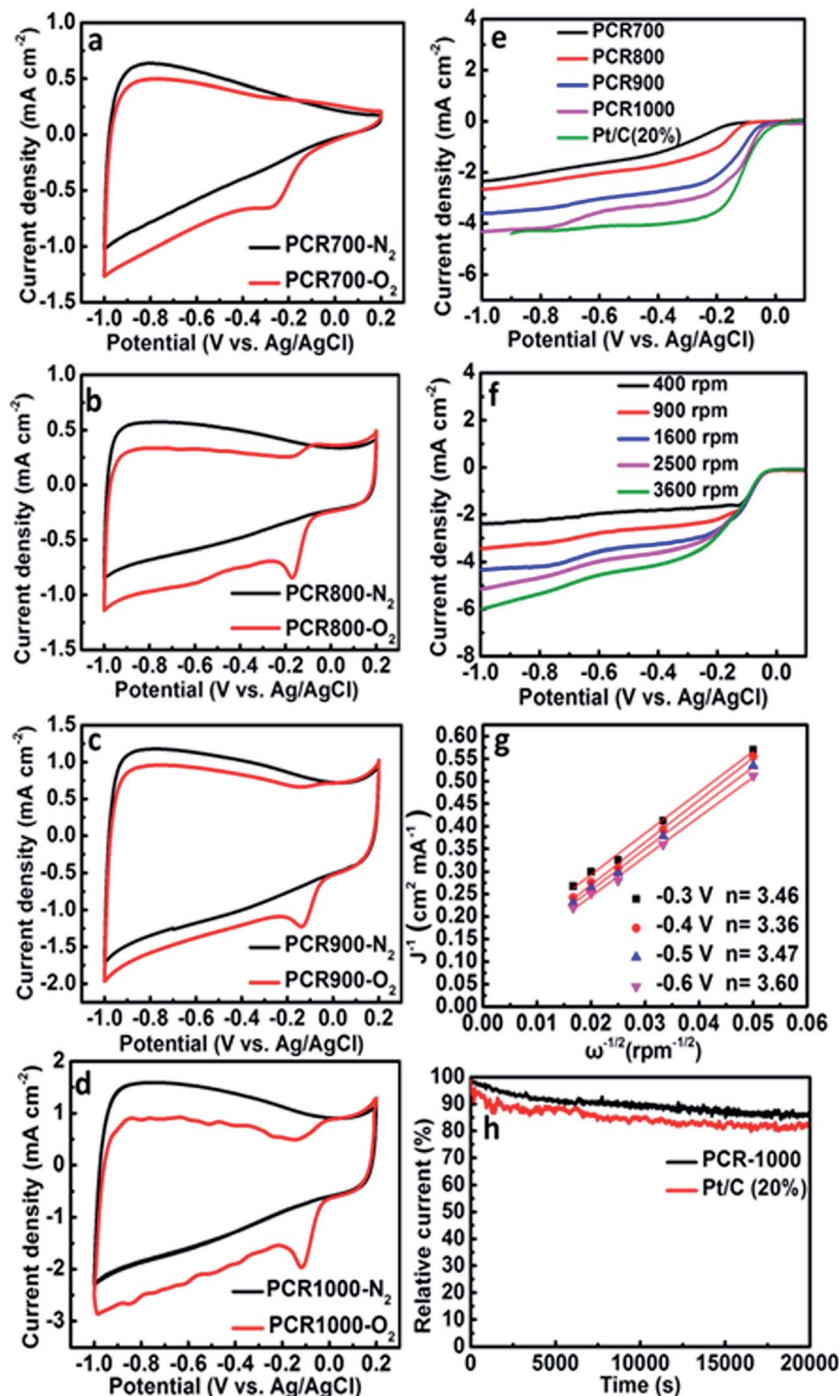


Fig. 8 Cyclic voltammograms (CV) curves of PCR-700 (a), PCR-800 (b), PCR-900 (c), and PCR-1000 (d) black line- N_2 , red line- O_2 , linear sweep voltammograms of PCR-700, PCR-800, PCR-900, PCR-1000, and Pt/C catalysts at 1600 rpm and 10 mV s^{-1} (e), polarization curves for ORR in an O_2 -saturated 0.1 M KOH solution on PCR-1000 electrodes with the rotation rates from 400 to 3600 rpm (f), Koutecky–Levich plots of the PCR-1000 electrode at different potentials and their corresponding electron transfer number (g), and current–time chronoamperometric response for PCR-1000 and Pt/C (20%) catalysts tested at $-0.40\text{ V vs. Ag/AgCl}$ with the rotation speed of 1600 rpm (h).



oxygen doubly bonded to carbon (C=O) and non-bridging oxygen in the phosphate group (P=O).^{27,28} The peak at 532.5–532.9 eV was assigned to the combined effects of singly bonded oxygen (–O–) in C–O and in C–O–P groups. The peak at 535.0–535.9 eV was ascribed to the chemisorbed oxygen and water. With the increasing temperature, the =O group in the carbonyl and the –O– group and chemisorbed O groups decreased (Table 4), which was in accordance with the analysis results of the P_{2p} spectrum.^{27,28}

Compared to PCR-700, PCR-800, and PCR-900, the CV curve of PCR-1000 exhibited the largest ORR potential (–0.120 V vs. Ag/AgCl) (Fig. 8d) in an O₂-saturated 0.1 M KOH solution. On comparing the RDE curves of different catalysts obtained at the scan rate of 10 mV s^{–1} and at the rotation speed of 1600 rpm, as shown in Fig. 8e, it was observed that PCR-700, PCR-800, and PCR-900 catalysts showed relatively low ORR activity. On the contrary, PCR-1000 showed better electrocatalytic activity with the half-peak potential of –0.137 V vs. Ag/AgCl, which is 30 mV lower than that of a commercial Pt/C (20%) catalyst, showing excellent potential in numerous P-doped carbon materials. The half-peak current density was 2.18 mA cm^{–2}, and the limited current densities were 4.36 mA cm^{–2} (Fig. 8e). RDE voltammograms of PCR-1000 catalyst at the rotation speeds from 400 to 3600 rpm in an O₂-saturated 0.1 M KOH solution are shown in Fig. 8f. Derived Koutecký–Levich (K–L) plots (Fig. 8g) show the good linearity and parallelism, indicating that the ORR process of PCR-1000 follows first-order kinetics in the selected potential range from –0.3 to –0.6 V.

The electron transfer numbers (*n*) during ORR and the kinetic limited current density (*J_k*) can be calculated from the following K–L equation:²⁹

$$1/J = 1/J_k + 1/(0.62nFC_0D_0^{2/3}\nu^{-1/6}\omega^{1/2}) \quad (1)$$

where *J* is the measured current density (mA cm^{–2}), *J_k* is the kinetic limited current density (mA cm^{–2}), ω is the angular velocity of the disk ($\omega = 2\pi N$, *N* is the linear rotation speed), *n* is the overall number of electrons transferred per oxygen molecule during the ORR, *D₀* is the diffusion coefficient (cm s^{–1}), *F* is the Faraday constant (*F* = 96 500 C mol^{–1}), *C₀* is the bulk concentration of O₂ (mol L^{–1}), ν is the kinematic viscosity of the electrolyte, *k* is the electron transfer rate constant, and the values of *C₀*, *D₀*, and ν for an O₂-saturated 0.1 M KOH solution are 1.14 × 10^{–6} mol cm^{–3}, 1.73 × 10^{–5} cm² s^{–1}, and 0.01 cm² s^{–1}, respectively. The values of the *n* and *J_k* can be estimated from the slope and intercept of the K–L plots, based on eqn (1). The values of the transferred electron number in the PCR-1000 catalyst are 3.46, 3.36, 3.47, and 3.60 e, which is involved in both a two and four electron transfer.⁴ The stability of PCR-1000 and Pt/C (20%) catalysts were tested at the constant potential of –0.4 V in an O₂-saturated 0.1 M KOH solution at 1600 rpm (Fig. 8h). After 20 000 s of testing, PCR-1000 maintained 87% of its initial current value, which was higher than 82.1% of the Pt/C (20%) catalyst, indicating better stability than the Pt/C catalyst in an O₂-saturated 0.1 M KOH solution.

Conclusion

A facile method was used for the fabrication of phosphorus-doped honeycomb carbon materials as oxygen reduction catalysts derived from *Pulsatilla chinensis* (Bunge) Regel. The largest content of P in P-doped PCR biomass was 1.46 at%, which corresponds to the value obtained *via* other chemical method. The pyridinic-N group gradually increased from 12.31% to 46.42%, which takes more active sites in PCR-1000. PCR-1000 showed a better half-peak potential (–0.137 V vs. Ag/AgCl), which was 30 mV lower than that of the commercial Pt/C (20%) catalyst, showing excellent reactive potential in many previous P-doped carbon materials. In the PCR-1000 catalyst, the half-peak current density was 2.18 mA cm^{–2}, and the limited current densities were 4.36 mA cm^{–2}. The values of the transferred electron number in PCR-1000 catalyst were 3.46, 3.36, 3.47, and 3.60 e including both two and four electron transfers. A value of 87.0% of the initial current density was maintained after 20 000 s stable testing in an O₂-saturated 0.1 M KOH solution with a rotation of 1600 rpm. The electrochemical testing results demonstrated that the method of P-doped materials derived from plants containing phosphorus exhibited superior properties in oxygen reduction, which implemented a novel idea for the preparation of difficult to prepare P-doped carbon material catalysts in fuel cells.

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Notes and references

- 1 L. Dai, Y. Xue, L. Qu, H. J. Choi and J. B. Baek, *Chem. Rev.*, 2015, **115**, 4823–4892.
- 2 M. K. Debe, *Nature*, 2012, **486**, 43–51.
- 3 J. Masa, W. Xia, M. Muhler and W. Schuhmann, *Angew. Chem., Int. Ed.*, 2015, **54**, 10102–10120.
- 4 K. Gong, F. Du, Z. Xia, M. Durstock and L. Dai, *Science*, 2009, **323**, 760–764.
- 5 X. Gong, S. Liu, C. Ouyang, P. Strasser and R. Yang, *ACS Catal*, 2015, **5**, 920–927.
- 6 X. Sun, J. Xu, Y. Ding, B. Zhang, Z. Feng and D. S. Su, *ChemSusChem*, 2015, **8**, 2872–2876.
- 7 C. Zhang, N. Mahmood, H. Yin, F. Liu and Y. Hou, *Adv. Mater.*, 2013, **25**, 4932–4937.
- 8 D.-S. Yang, D. Bhattacharjya, S. Inamdar, J. Park and J.-S. Yu, *J. Am. Chem. Soc.*, 2012, **134**, 16127–16130.
- 9 Z.-W. Liu, F. Peng, H.-J. Wang, H. Yu, W.-X. Zheng and J. Yang, *Angew. Chem.*, 2011, **123**, 3315–3319.



- 10 Z. Chen, K. Li and L. Pu, *Bioresour. Technol.*, 2014, **170**, 379–384.
- 11 J. Wu, Z. Yang, Q. Sun, X. Li, P. Strasser and R. Yang, *Electrochim. Acta*, 2014, **127**, 53–60.
- 12 R. Li, Z. Wei, X. Gou and W. Xu, *RSC Adv.*, 2013, **3**, 9978–9984.
- 13 Z. Chen, K. Li, P. Zhang, L. Pu, X. Zhang and Z. Fu, *Chem. Eng. J.*, 2015, **259**, 820–826.
- 14 V. V. Strelko, V. S. Kuts and P. A. Thrower, *Carbon*, 2000, **38**, 1499–1503.
- 15 M. Borghei, N. Laocharoen, E. Kibena-Poldsepp, L.-S. Johansson, J. Campbell, E. Kauppinen, K. Tammeveski and O. J. Rojas, *Appl. Catal., B*, 2017, **204**, 394–402.
- 16 L. Zhao, Y. Qiu, J. Yu, X. Deng, C. Dai and X. Bai, *Nanoscale*, 2013, **5**, 4902–4909.
- 17 F. Tuinstra and J. L. Koenig, *J. Chem. Phys.*, 1970, **53**, 1126–1130.
- 18 R. J. Nemanich and S. A. Solin, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1979, **20**, 392–401.
- 19 J. Bedia, J. M. Rosas, J. Rodríguez-Mirasol and T. Cordero, *Appl. Catal., B*, 2010, **94**(1), 8–18.
- 20 M. Cuisinier, N. Dupré, J. F. Martin, R. Kanno and D. Guyomard, *J. Power Sources*, 2013, **224**, 50–58.
- 21 X. Wu and L. R. Radovic, *Carbon*, 2006, **44**, 141–151.
- 22 A. M. Puziy, O. I. Poddubnaya and A. M. Ziatdinov, *Appl. Surf. Sci.*, 2006, **252**, 8036–8038.
- 23 J. M. Rosas, J. Bedia, J. Rodríguez-Mirasol and T. Cordero, *Ind. Eng. Chem. Res.*, 2008, **47**, 1288–1296.
- 24 Y.-J. Lee and L. R. Radovic, *Carbon*, 2003, **41**, 1987–1997.
- 25 S. G. Oh and N. M. Rodriguez, *J. Mater. Res.*, 1993, **8**, 2879–2888.
- 26 D. Long, J. Zhang, J. Yang, Z. Hu, G. Cheng, X. Liu, R. Zhang, L. Zhan, W. Qiao and L. Ling, *Carbon*, 2008, **46**, 1253–1269.
- 27 A. G. Kannan, N. R. Choudhury and N. K. Dutta, *Polymer*, 2007, **48**, 7078–7086.
- 28 P. Y. Shih, S. W. Yung and T. S. Chin, *J. Non-Cryst. Solids*, 1999, **244**, 211–222.
- 29 C. Jin, F. Lu, X. Cao, Z. Yang and R. Yang, *J. Mater. Chem. A*, 2013, **1**, 12170–12177.

