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Highly efficient one-pot three-component Betti reaction in water using reverse zinc oxide micelles as a recoverable and reusable catalyst†

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An efficient synthesis of Betti bases *via* a one-pot three-component reaction of 2-naphthol, substituted aldehydes and anilines in aqueous media is reported. Through screening different catalysts, reverse zinc oxide nanomicelles show good activity and high selectivity. The catalyst retained activity after six cycles of reuse and were recoverable. The method has the advantages of atom economy, short reaction time, good yield, and easy work-up as well as being environmental friendly.

Introduction

The Betti reaction, which is a typical Mannich reaction, is one of the most important fundamental carbon-carbon bond-forming reactions in organic chemistry.¹ The Betti reaction provides 1-(α -aminoalkyl)-2-naphthols, the so-called Betti bases. The optical Betti base isomers are valuable ligands in asymmetric synthesis.^{2,3} Recently, the importance of Betti bases was well established in pharmaceutical chemistry because of their bioactivities, including anti-bacterial, antipain, antihypertensive and bradycardiac activities.⁴⁻⁶ The original Betti reaction is between an aldehyde, ammonia or urea and β -naphthol in an ethanolic solution of potassium hydroxide in two steps for a long time (9–36 h).^{7,8} However, the synthetic approach is plagued by a number of serious disadvantages and has limited applications. Therefore, numerous modern versions of Betti reaction have been developed to overcome the drawbacks of the classical method. In general, the preferable route is the use of a one-pot three-component strategy that allows for a wide range of structural variations by using quinolins,⁹ naphthols¹⁰ and alkyl amines to exchange ammonia.¹¹ However, these methods suffered from severe side reactions and longer reaction times with lower yields.

In terms of green chemistry, the development and application of multi-component coupling reactions in water is favourable, as they provide simple and rapid access to a large number of organic molecules *via* a tolerable method.^{12,13} But the Betti

reaction is difficult to be initiated in water because the activity of amine is inhibited by the strong hydrogen bond with water. In recent years, rare earth and lanthanide triflates,¹⁴ sulfanilic acid-functionalized silica-coated magnetite nanoparticles,¹⁵ surfactants¹⁶ or iminium salts^{17,18} as catalysts for direct-type Mannich reactions in water or under solvent-free conditions have been successfully applied.^{19,20} In most cases the catalysts are not used friendly and non-recoverable. To our knowledge, it is still a challenge for aromatic primary amines are employed as the substrates in water for their electron-withdrawing properties and steric hindrance. As a consequence, it is necessary to develop environmentally benign reactions and atom-economic catalytic processes for Betti reaction.

Nanocrystalline metal oxides find excellent application as active adsorbent for gas, for destruction of hazardous chemicals, and as catalyst in various organic reactions.²¹ Karmakar *et al.* reported an efficient method for the preparation of Betti base derivatives over nanocrystalline MgO in water.²² Unfortunately, the yield of aniline's product is lower than 20%. Moreover, nanocrystalline metal oxides usually undergo agglomeration when dispersed in solutions because of large specific surface area and high surface activity. This behavior has a major impact on their reactivity.²³

Under the inspiration of metallomicelles,^{24,25} which were utilized for their certain similarities with the natural enzymes, we designed reverse zinc oxide nanomicelles as functional catalyst to meet the dual requirements of the disperse stability and nano-size control. Reverse micelles have been depicted as passive nano-reactors that *via* their shapes template the growing crystalline nuclei into narrowly dispersed or even perfectly uniform nano-sized particles. Reverse micellar systems containing metal oxide could possibly afford high local reactant concentrations near the catalysts, helping enhance the catalytic rate.²⁶

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In this study, direct-type three-component Betti reaction using water as the only solvent at room temperature catalyzed by reverse ZnO nanomicelles is reported. Reverse ZnO nanomicelles are prepared starting from zinc acetate as a precursor, triethanolamine as the base and hexadecyl trimethyl ammonium bromide as the surfactant in cyclohexane. This route has the advantages of atom-economy, operational-simplicity, eco-friendliness and cost-efficiency.

Experimental section

Materials

Unless otherwise noted, all reagents and solvents were purchased from commercial suppliers and used without further purification.

Preparation of reverse zinc oxide nanomicelles. Hexadecyl trimethyl ammonium bromide (CTAB, 0.3644 g, 1 mmol) was dissolved in cyclohexane (10 mL). The suspension was heated at 80 °C for 1 h and cooled to room temperature under stirring. After being stirred overnight, triethanolamine (5 mL) was added dropwise under vigorous stirring until the milk white emulsion was formed. Zinc acetate ($\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$) (2.19 g, 1 mmol) was dissolved in distilled water (100 mL) whilst stirring to obtain the 1 mol L⁻¹ zinc solution. Zinc acetate solution (0.2 mL, 1 mol L⁻¹) was injected into the above mentioned emulsion slowly under moderate stirring until the mixture turn cleared and stirred for another 1 h. Then the solution was sonicated at 200 W for 15 min in an ice bath. The reacted solution kept clear, and the reverse ZnO micelles were formed in solution and stabilized by surfactant.

Synthesis of Betti base derivatives. The catalytic activity of the reverse ZnO nanomicelles was investigated by measuring their ability to synthesis Betti base derivatives. To a mixture of aromatic aldehydes (1 mmol), β -naphthol (1 mmol, 0.144 g) and aniline (1 mmol, 0.094 g) in 10 mL water in a round bottom flask, catalytic amount of reverse ZnO nanomicelles was added and stirred at room temperature for appropriate time. After completion of the reaction as monitored by thin layer chromatography (TLC), the aqueous layer was decanted to recycle the catalyst for another reaction. The crude product was filtered and purified by recrystallization from ethanol (95%). The novel products were characterized by ¹H NMR, ¹³C NMR, IR and HRMS.

Characterization methods. Dynamic light scattering (DLS) measurements were performed on a NICOMP 380ZLS zeta potential/particle size analyzer (PSS, USA) equipped with a He-Ne laser at a wavelength of 660 nm. The experimental data were analyzed by the CONTIN method, which is based on an inverse-Laplace transformation of the data and provides access to a size distribution histogram for the analyzed micellar solutions. The ζ -potential of ZnO nanomicelles was determined by laser Doppler anemometry using a NICOMP 380ZLS.

Transmission electron microscopy (TEM) was performed on JEOL JEM 1230 (Joel, Japan) at an acceleration voltage of 120 kV. Then a drop of the reverse ZnO nanomicelle sample was applied onto a carbon-coated copper TEM grid. After solvent evaporation ZnO micelles remained dispersed on the carbon film of the TEM grid.

X-ray diffractometer were employed on Rigaku D/max-III type instrument (Denki, Japan) using Cu K α (1.54 Å) radiation at a beam voltage 40 kV and a 30 mA beam current. The patterns were obtained by continuous scans in a 2θ range of 5–90° with a scan rate of 4 deg min⁻¹ and step 0.02 deg. XRD sample was obtained by centrifugation, washing, drying the micelles in vacuum at 70 °C in the vacuum drying oven.

Analytical thin layer chromatography (TLC) was performed using glass backed TLC extra hard layer pre-coated with silica gel (0.25 mm, 60 Å pore size) and visualized by exposure to ultraviolet light (254/366 nm) using a Camag UV Lamp (Switzerland).

Proton nuclear magnetic spectra (¹H NMR, 400 MHz) and carbon nuclear magnetic resonance spectra (¹³C NMR, 100 MHz) were recorded on an JNM-ECZ400s/L spectrometer (Joel, Japan) using CDCl₃ or DMSO-*d*₆ as a solvent and TMS as an internal standard. IR spectra were recorded on a Jasco FT-IR 4100 Series spectrophotometer (Perkin-Elmer, USA), ν max (cm⁻¹) are partially reported. HRMS analyses were carried out using a Bruker micrOTOF-Q instrument (Bruker, USA).

Results and discussion

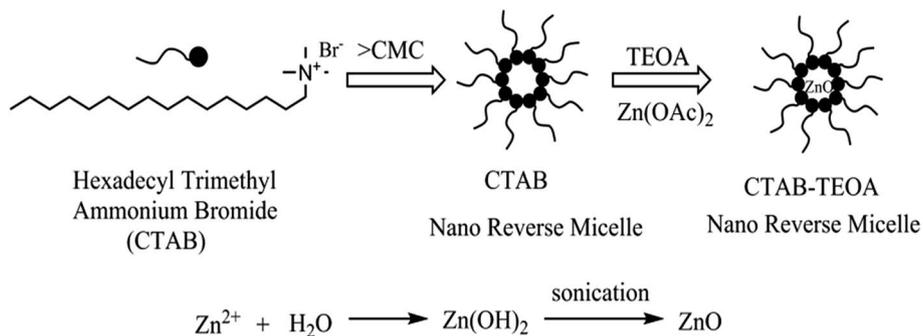
Preparation and characterization of reverse ZnO nanomicelles

Reverse ZnO nanomicelles were prepared according to the modified sol-gel based procedures²⁷ using zinc acetate dehydrate as precursor and cyclohexane as solvent. Reverse micelles were formed by the self-assembly of the surfactant in apolar solvents. Metal salts are solubilized within the interior of inverse micelles where were shield by the hydrophilic head groups of the surfactant from the oleic surroundings.²⁸ Ionic surfactants as amphiphilic molecules are often used to form reverse micelles for protein solubilization, such as anionic di-2-ethylhexyl sodium sulfosuccinate (AOT)²⁹ and cationic hexadecyl trimethyl ammonium bromide (CTAB).³⁰ In this case, CTAB was employed in the preparation of reverse micelles. Its relatively polar head group regions solubilize and confine added zinc salts and act as reaction cages when forced hydrolysis is initiated. The reaction sequence is sustained by material exchange between reverse micelles and final particle size is dependent on stabilization by the surfactant. In our study, triethanolamine was used as base to maintain pH value and sonication was applied to initiate the hydrolysis of zinc ions. The reverse ZnO nanomicelles were formed which were sterically stabilized in solution by surfactant (Scheme 1).

The particle size and distribution of reverse ZnO nanomicelles was measured with NICOMP 380ZLS and the morphology was characterized using transmission electron microscopy (TEM). The distribution of sphere diameters is shown in Fig. 1. The image revealed that the average diameter was 120 nm and the size distribution was 80%.

Fig. 2 illustrates the transmission electron microscopy (TEM) micrograph of the reverse ZnO nanomicelles. It can be observed that the particles are highly dispersed, sub-spherical and monodisperse on the TEM grid. The size range of them is approximate 200 nm. The particles displayed a homogeneous size distribution and the particle diameters ranged from 20 to 500 nm, as evidenced by Zata-diameter measurement.





Scheme 1 Preparation process of reverse ZnO nanomicelles.

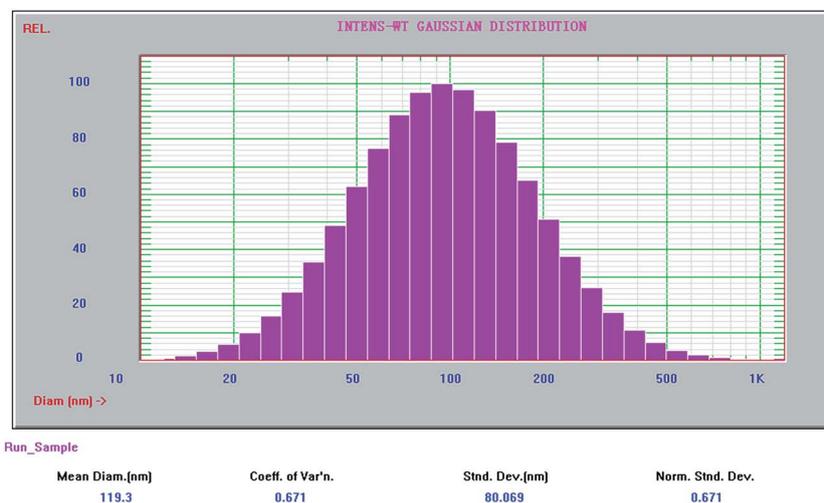


Fig. 1 Size distribution for reverse ZnO nanomicelles with average diameter of 120 nm.

The formation of ZnO nanomicelles was investigated by X-ray diffraction (XRD). The XRD pattern of ZnO nanomicelles after centrifugation, washing and drying was shown in Fig. 3.

The peaks appeared at 2θ value ranging from 31.73° , 34.38° , 36.22° , 47.50° , 56.56° , 62.81° , 66.34° , 67.91° , 69.03° , 72.6° and 76.90° values corresponded to pure ZnO. All XRD diffraction peaks of ZnO micelles are shown in a good agreement with hexagonal structure of zincite phase reported in JCPDS file card no. 05-0664. No peaks of impurity are observed, indicating that ZnO precursor has been completely transformed into nano ZnO. The result indicated that the sonication step is a key factor in avoiding the fast precipitation and achieving a pure zinc oxide.

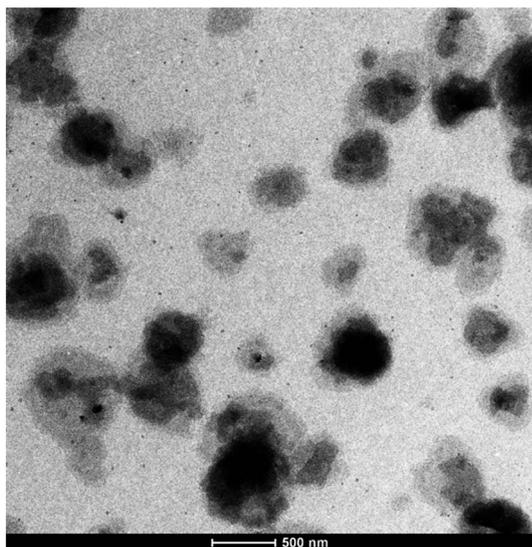


Fig. 2 TEM image of reverse ZnO nanomicelles.

Catalyst screening

Initially, the preparation of 4a from *m*-nitrobenzaldehyde (1 equiv.), naphthol (1.2 equiv.) and aniline (1 equiv.) in water was investigated as a model system to search for the effective and reusable catalytic system (Table 1).

It was found that the reaction failed in the absence of catalyst from room temperature to 90°C for 24 h. Conventional basic or acidic catalyst such as sodium hydroxide, triethylamine and acetic acid only afforded Schiff base 6a as the terminal products (Table 1, entry 1–4). Typical Lewis acid catalysts such as ZnCl_2 and FeCl_3 provided bis product 5a as a major product in low yield whereas, 4a was formed as a minor product (<10% yield)



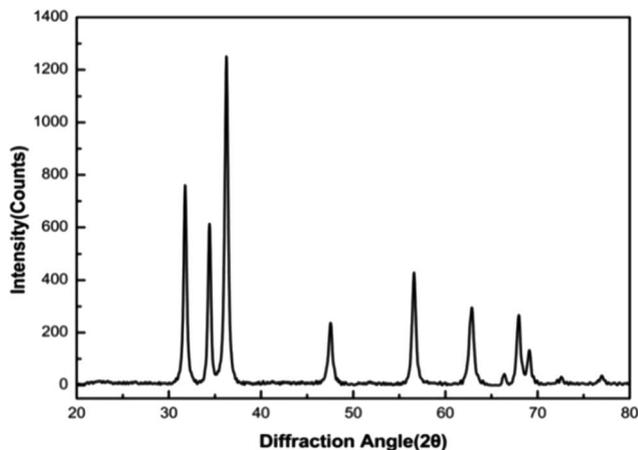


Fig. 3 XRD graph of ZnO nanomicelle particles.

(Table 1, entry 5 and 6). In contrast, same test reaction was carried out with different supported Lewis acid catalysts such as $\text{ZnCl}_2\text{-SiO}_2$, $\text{BF}_3\text{-SiO}_2$ and $\text{TiO}_2\text{-SiO}_2$, the yields of desired product 4a were increased from 23% to 42% (Table 1, entry 7–9).

To study the role of nanomicelles, the reaction was performed with surfactants and ZnO nanoparticles (ZnO NPs). Surprisingly, the reaction was not successful either with non-ionic surfactants (*i.e.* PEG200–PEG2000) or with ionic

surfactants (*i.e.* CTAB). Even though the mixture of ZnCl_2 and CTAB was added, there was no appreciable improvement in yield (Table 1, entry 10–13). Meanwhile, it was found that the smaller ZnO NPs (10 nm) can afford the desired product 4a, which is better than ZnO NPs (100 nm) (Table 1, entry 14 and 15). The best result showed that exposure of the reaction mixture to 10 mmol% reverse ZnO nanomicelles (100 nm) in 10 mL water at 70 °C for 30 min provided 4a with excellent yield (Table 1, entry 16). The larger size of nanomicelles (>500 nm) did not give more satisfied results (Table 1, entry 18). The normal ZnO nanomicelles catalyst is neither activity nor selectivity to Betti reaction. Obviously, Betti reaction is sensitive to the structure of the catalyst. All the results can be explained as follows. In the catalytic process, the catalyst consists of small metal oxide particles dispersed on an inert support which was the key components of the catalyst. The activity of the catalyst will generally depend on the size of the metal oxide particles (active sites), where a catalyst with small particles will give high activity due to the large number of atoms available on the metal particles surfaces.³¹

Solvent screening

To investigate the feasibility of the strategy, various solvents such as dichloromethane, tetrahydrofuran, acetonitrile, ethanol and dimethyl formamide were used to improve the yield of 4a

Table 1 Comparison of efficiency of various catalysts in one-pot synthesis of 4a

Entry	Catalyst ^a	T (°C)	Time (h)	4a yield ^b (%)	5a yield ^b (%)	6a yield ^b (%)
1	20% NaOH	90	8	—	60	—
2	40% NaOH	90	8	—	20	40
3	CH_3COOH	90	8	—	—	10
4	Et_3N	90	8	—	—	—
5	FeCl_3	90	8	4	65	—
6	ZnCl_2	90	8	7	61	—
7	$\text{ZnCl}_2\text{-SiO}_2$	90	8	23	47	—
8	$\text{BF}_3\text{-SiO}_2$	90	8	29	43	—
9	$\text{TiO}_2\text{-SiO}_2$	90	8	42	41	—
10	PEG-200	90	8	—	—	—
11	PEG-2000	90	8	—	—	—
12	CTAB	90	8	—	—	—
13	ZnCl_2 + CTAB	90	8	Trace	47	Trace
14	ZnO NPs (10 nm)	90	1	43	—	—
15	ZnO NPs (100 nm)	90	1	12	—	—
16	Reversed ZnO micelles (100 nm)	70	0.5	98	—	—
17	Reversed ZnO micelles (200 nm)	70	0.5	96	—	—
18	Reversed ZnO micelles (500 nm)	70	0.5	41	—	—
19	Normal ZnO micelles (100 nm)	70	8	27	19	—

^a Conditions: 1 (1.2 mmol), 2a (1 mmol), 3a (1 mmol), catalyst (3–6) (15 mol%), (7–11) (0.01 g mmol^{-1} substrate), (12–20) (10 mol%), solvent (water, 10 mL), 0.5–24 h. ^b Isolated yield.



Table 2 Effects of solvent on the synthesis of 4a^a

Entry	Solv. ^a	Temp. (°C)	Time (h)	4a yield ^b (%)	5a yield ^b (%)
1	CH ₂ Cl ₂	25	24	7	14
2	CH ₃ CN	60	24	12	26
3	THF	60	8	6	12
4	DMF	80	8	34	28
5	EtOH	60	8	46	22
6	Water	25	0.15	87	—
7	Water	70	1	96	—

^a Reactions were carried out in solvent 5.0 mL on a 1 mmol scale at the special temperature using 10 mol% reverse ZnO nanomicelles as catalyst. ^b Isolated yield.

and reduce the formation of 5a. In polar non-proton solvents such as CH₂Cl₂, THF and CH₃CN, both 4a and 5a were obtained in poor yield. However, in polar protic solvents like C₂H₅OH and DMF, 5a was obtained in higher yield in comparison to 4. It was found that ethanol was superior to other examined solvents in view of yield (Table 2, entry 5). However, with further consideration on the development of environment-friendly chemistry, we tried to conduct Betti reaction in water. To our delight, when the reaction was performed in water at room temperature, only product 4a was formed. With further studies by increasing reaction temperature, the desired product 4a could be obtained smoothly in 87% yield in a shortened reaction time 8 min (Table 2, entry 6). The results revealed that the reaction in aqueous media condition produced the product in shortest reaction time and highest yield (96%). Previously, reversible dehydration condensation in water is in principle a difficult transformation because the large excess of water pushes the equilibrium in favor of the hydrated compounds.³² To date, dehydration reactions in water have usually been achieved by employing the novel catalyst.^{33–38} In this case, it was assumed that water can favor the nucleophilic addition reaction between imine intermediate and β-naphthol for the high polarity.

Surfactant screening

To elucidate the role of nanomicelles, various non-ionic and ionic surfactants (5 mol%) based ZnO reverse micelles (5 mol%) were tested on the reaction of *m*-nitrobenzaldehyde, naphthol and aniline in water. The results are summarized in Table 3. The reaction was run in the presence of TX 10 at room temperature for 8 h, as a result, only very low yield of product 4a

Table 3 Effect of surfactant-based reverse micelle on synthesis of 4a^a

Entry	Surfactant	Catalyst	Time (h)	4a yield ^b (%)
1	TX10	ZnO	8	—
2	DBSA	ZnO	48	19
3	SDS	ZnO	48	23
4	CTAB	ZnO	0.15	96

^a Conditions: 1 (1.2 mmol), 2a (1 mmol), 3a (1 mmol), surfactant (1–4) (5 mol%), catalyst (ZnO NPs, 5 mol%), solvent (water, 10 mL), room temperature, 9 min to 48 h. ^b Isolated yield.

was obtained. In the presence of *p*-dodecylbenzenesulfonic acid (DBSA) and sodium dodecyl sulfate (SDS), which acts both Bronsted acid and anionic surfactant, only 19% and 23% of desired product 4a was gained. CTAB acts as cationic surfactant. CTAB based reverse ZnO nanomicelles proved to be the best choice. It was known that the reverse micelles are small molecular assemblies that encapsulate nanoscopic pools of water.³⁹ The results revealed that the interaction of the cationic surfactant head groups and their Br[−] counter ions in water pool is believed to play a major role in controlling the yields by promoting the dehydration condensation in water.

Stability of the reaction system

The effect of reaction temperature was carefully examined (Fig. 4). The reaction was smoothly promoted with the increasing of temperature from room temperature to 70 °C. After 9 minutes, the highest yield was obtained at 70 °C. However, the catalyst activity declined as the reaction temperature increased over 70 °C. Loss in surface area caused by the dissociation of the micelles could be the reason for the decreased activity. The effect of reaction time was also investigated (Fig. 5). It was found that the reaction was started 2 min later and maintained a stable state for a long period. The yield did not increase significantly from 8 min to 20 min.

Recyclability of catalyst

Additionally, recycle and reusability of the catalyst was examined (Table 4). After the completion of the reaction, the aqueous layer was decanted and the catalyst was reused for the next run under the same conditions. As shown in Table 4, the desired products yields remained essentially constant for the six successive cycles, reflecting high stability and reusability of the catalyst. This is consistent with the characterization results for this catalyst. Dynamic light scattering (DLS) measurements

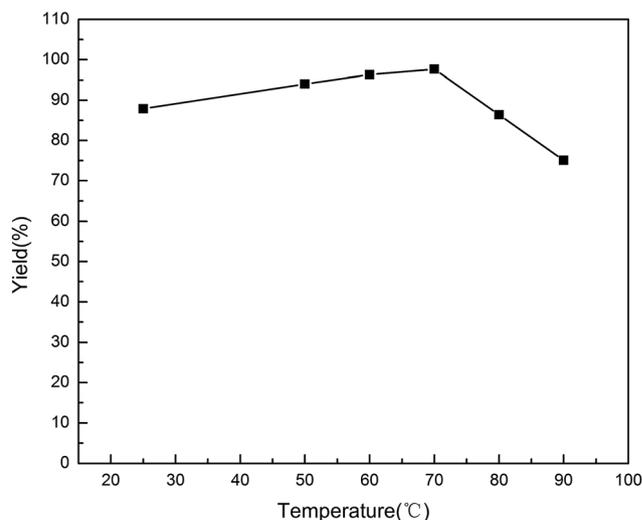


Fig. 4 Influence of reaction temperature on synthesis of 4a. Reactions were carried out on a 0.1 mmol scale; reverse ZnO nanomicelles 20 mol%.



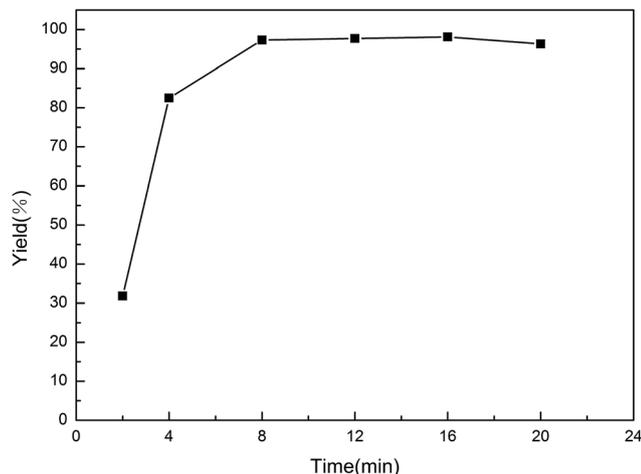


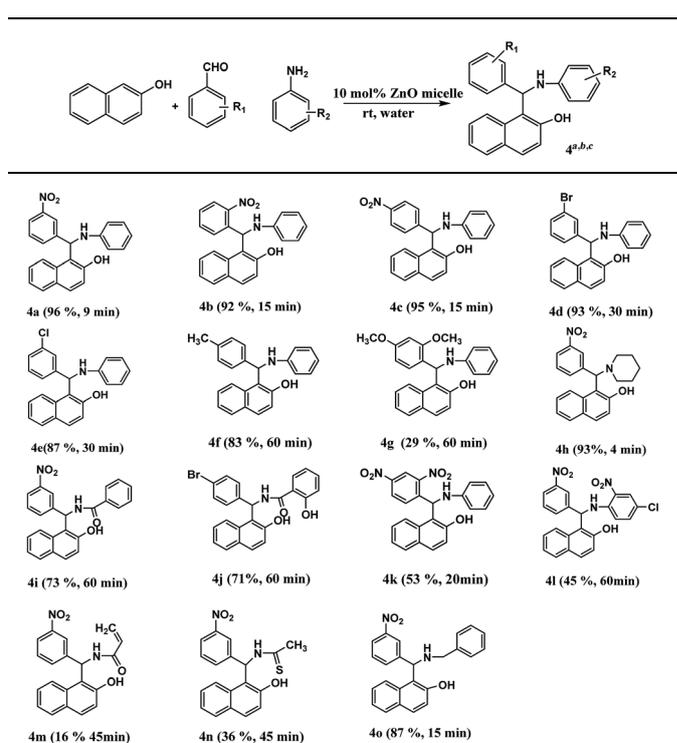
Fig. 5 Effect of reaction time on yield of 4a. Reactions were carried out on a 0.1 mmol scale; reverse ZnO nanomicelles 20 mol%.

results showed the particle diameters of ZnO nanomicelles ranged from 120 to 250 nm and displayed a homogeneous size distribution after the five successive runs. The increasing ζ -potential indicated the catalytic system tend to be stable gradually. No remarkable changes could be evidenced in the recovery of the catalytic system.

Comparison of the substrates

To explore the generality and scope of the protocol, the reaction of various aromatic aldehydes, amines and β -naphthol were extended to prepare a series of Betti base derivatives and the results are depicted in Table 5. It was found that both electron-rich and electron-deficient aromatic aldehydes could be converted to the desired Betti base derivatives in the presence of reverse ZnO nanomicelles (10 mol%) at 70 °C in 1 h in good yields (Table 5, 4a–4g). However, halogen or electron-withdrawing groups at *meta* position accelerated the reaction rather than electron-donating groups. Furthermore, the amine component of the three-component reaction was surveyed. To our delight, the reactions proceeded smoothly when aniline were used as substrate. Under the same reaction conditions, when piperidine was used as an amine partner, the desired product 4h was obtained in 93% yield in 4 minutes (Table 5, 4h). The results indicated that aliphatic amine and hydrazine were more active than aniline. In addition, a variety of Betti base derivatives were obtained in moderate yields when amide or

Table 5 Synthesized Betti base derivatives catalyzed by reverse ZnO nanomicelles



^a Reactions were conducted with naphthol (1.2 mmol), benzaldehyde (1 mmol), amine (1 mmol) and reverse ZnO nanomicelles (10 mol%) in water (10 mL) at room temperature. ^b The products were characterized by melting point, ¹H, ¹³C NMR, IR and mass spectroscopy. ^c Isolated yield.

thioacetamide were used as the substrates because the electron density of amino group was reduced by the electron-withdrawing conjugation effects of unsaturated bonds (Table 5, 4i, 4j, 4m and 4n).

Exploration of the possible mechanism

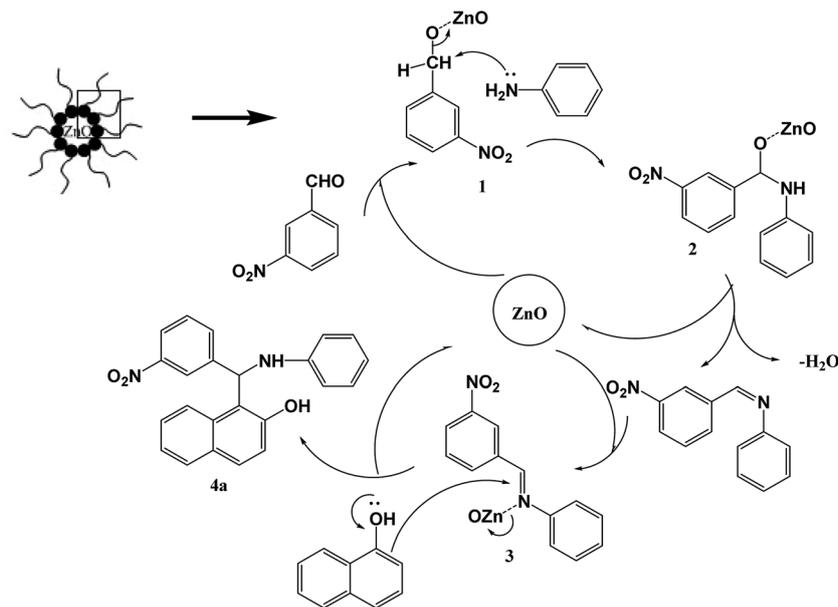
The mechanism of this multi-component reaction involves a Knoevenagel condensation/Michael addition cascade process.⁴⁰ A tentative mechanism for the zinc oxide-catalyzed Mannich is proposed in Scheme 2 in the revised manuscript. Firstly, ZnO coordinates to benzaldehyde and generates the zinc oxide-aldehyde intermediate **1**, which reacts with aniline to

Table 4 The reuse of reverse ZnO micelles in the synthesis of 4a^a

Cycle	Fresh	Reuse 1	Reuse 2	Reuse 3	Reuse 4	Reuse 5	Reuse 6
Yield ^b (%)	96	95	96	97	95	95	94
Diameter (nm)	116 ± 12	119 ± 13	142 ± 13	166 ± 15	186 ± 14	216 ± 17	251 ± 15
ζ (mV)	6.27 ± 0.52	7.04 ± 0.24	10.21 ± 1.03	13.02 ± 0.37	16.47 ± 0.62	20.14 ± 0.29	23.11 ± 0.74

^a Conditions: **1** (1.2 mmol), **2a** (1 mmol), **3a** (1 mmol), catalyst (reverse ZnO nanomicelles, 10 mol%), solvent (water, 10 mL), 70 °C, 1 h. ^b Isolated yield.





Scheme 2 Proposed mechanism for the synthesis of 4a.

generate the imine intermediate **2** after the dehydration and remove of zinc oxide. Subsequently, the imine intermediate is activated when ZnO binds to the nitrogen atom **3** which promotes the following nucleophilic addition of β -naphthol to give the desired product **4** and regenerates zinc oxide. Based on the results obtained, the formation of the imine intermediate is likely the rate-determining step. ZnO, which is a Lewis acid, provides efficient acidic sites to form coordinate bond between oxygen atom (or nitrogen atom) which can activate the formation of imine intermediate, followed by facilitating the nucleophilic addition. Reverse micelles played an important role in combining the substrates, stabilizing the imine intermediate and promoting the dehydration condensation under neutral conditions in water catalyzed by the cationic water pool encapsulated in reverse micelles. The enhanced surface area due to nano particle size is an added advantage for its reactivity. Water promotes the nucleophilic addition reaction between imine intermediate and β -naphthol for the high polarity. All these important factors are responsible for the high accessibility of the substrate molecules on the catalyst surface. Thus, the reaction occurred more easily in a micelles special with respect to its functioning as a nanoreactor.

Conclusions

In conclusion, we have developed an efficient synthesis of Betti bases *via* the reaction of naphthols, aldehydes with amines in aqueous media. Reverse ZnO nanomicelles effectively catalyzed Betti reaction in excellent yields without the formation of any other by-product. This method has the advantages of atom economy and environmentally benign. The results prove the crucial role of reverse ZnO micelles as

nanoreactors and will find more extensive applications in the field of green chemistry.

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References

- 1 H. Gao, J. Sun and C. G. Yan, *Chin. Chem. Lett.*, 2014, **26**, 353.
- 2 C. Cardellicchio, M. A. M. Capozzi and F. Naso, *Tetrahedron: Asymmetry*, 2010, **21**, 507.
- 3 A. Kumar, M. K. Gupta and M. Kumar, *ChemInform*, 2010, **41**, 1582.
- 4 A. Y. Shen, C. T. Tsai and C. L. Chen, *Eur. J. Med. Chem.*, 1999, **34**, 877.
- 5 I. Szatmári and F. Fülöp, *Curr. Org. Synth.*, 2004, **1**, 155.
- 6 H. R. Shaterian and H. Yarahmadi, *Tetrahedron Lett.*, 2008, **49**, 1297.
- 7 H. Gao, J. Sun and C. G. Yan, *Chin. Chem. Lett.*, 2014, **26**, 353.
- 8 A. Decottignies, C. Len and A. Fihri, *ChemSuschem*, 2010, **5**, 502.
- 9 C. Mukhopadhyay, S. Rana and R. J. Butcher, *ARKIVOC*, 2010, **12**, 291.
- 10 I. Szatmári, A. Hetényi, L. Lázár and F. Fülöp, *J. Heterocycl. Chem.*, 2004, **41**, 367.
- 11 H. Gao, J. Sun and C. G. Yan, *Chin. Chem. Lett.*, 2015, **26**, 353.



- 12 A. Decottignies, C. Len and A. Fihri, *Chemsuschem*, 2010, **5**, 502.
- 13 T. Maegawa, Y. Kitamura, S. Sako, T. Udzu, S. Ai and A. Tanaka, *Chemistry*, 2007, **13**, 5937.
- 14 S. Kobayashi, H. Ishitani, S. Komiyama, D. C. Oniciu and A. R. Katritzky, *Tetrahedron Lett.*, 1996, **37**, 3731.
- 15 H. Moghanian, A. Mobinikhaledi, A. G. Blackman and E. Sarough-Farahani, *RSC Adv.*, 2014, **4**, 28176.
- 16 A. Kumar, M. K. Gupta and M. Kumar, *Tetrahedron Lett.*, 2010, **51**, 1582.
- 17 J. Auge, N. Lubin-Germain and J. Uziel, *ChemInform*, 2007, **38**, 1739.
- 18 B. Karmakar and J. Banerji, *Tetrahedron Lett.*, 2011, **52**, 4957.
- 19 A. Shahrissa, R. Teimuri-Mofrad and M. Gholamhosseini-Nazari, *Mol. Diversity*, 2015, **19**, 1.
- 20 A. Rezaeifard, P. Farshid, M. Jafarpour and G. K. Moghaddam, *RSC Adv.*, 2014, **4**, 9189.
- 21 A. Kołodziejczak-Radzimska, E. Markiewicz and T. Jesionowski, *J. Nanomater.*, 2012, **20**, 5603.
- 22 B. Karmakar and J. Banerji, *Tetrahedron Lett.*, 2011, **52**, 4957.
- 23 A. Punnoose, K. Dodge, J. W. Rasmussen, J. Chess, D. Wingett and C. Anders, *ACS Sustainable Chem. Eng.*, 2014, **2**, 1666.
- 24 W. Jiang, B. Xu, J. Li, Q. Lin, X. Zeng and H. Chen, *Int. J. Chem. Kinet.*, 2007, **39**, 672–680.
- 25 F. Li, F. Feng, L. Yu and J. Xie, *J. Solution Chem.*, 2014, **43**, 1331–1343.
- 26 T. Dwars, E. Paetzold and G. Oehme, *Angew. Chem., Int. Ed.*, 2005, **44**, 7174.
- 27 J. Zhang, L. D. Sun, X. C. Jiang, C. S. Liao and C. H. Yan, *Cryst. Growth Des.*, 2003, **4**, 309.
- 28 A. Manino, J. P. Wilcoxon, A. P. Sylwester and J. S. Kawola, *American Chemical Society national meeting*, Denver, CO, United States, 1993.
- 29 J. G. Liu, J. M. Xing, C. Shen, C. L. Yang and H. Z. Liu, *Biochem. Eng. J.*, 2004, **21**, 273.
- 30 R. Shen, J. G. Liu, J. M. Xing and H. Z. Liu, *Chin. J. Process Eng.*, 2005, **5**, 255.
- 31 H. Abdizadeh, Y. Vahidshad, H. R. Baharvandi and M. A. Baseri, *International Conference on Ultrafine Grained and Nanostructured Materials*, 2012, **5**, 559.
- 32 C. J. Li and L. Chen, *Chem. Soc. Rev.*, 2006, **35**, 68.
- 33 T. Murase, Y. Nishijima and M. Fujita, *J. Am. Chem. Soc.*, 2012, **134**, 162.
- 34 Y. Mori, K. Kakumoto and K. Manabe, *ChemInform*, 2000, **41**, 3107.
- 35 K. Manabe, X. M. Sun and S. Kobayashi, *J. Am. Chem. Soc.*, 2001, **123**, 10101.
- 36 K. Suzuki, *Pure Appl. Chem.*, 1994, **66**, 1557.
- 37 P. Cotanda, A. Lu, J. P. Patterson, N. Petzetakis and R. K. O'Reilly, *Macromolecules*, 2012, **45**, 2377.
- 38 H. Firouzabadi, N. Iranpoor and A. Jafari, *Adv. Synth. Catal.*, 2005, **347**, 655.
- 39 I. R. Piletic, D. E. Moilanen, D. B. Spry, N. E. Levinger and M. D. Faye, *J. Phys. Chem. A*, 2006, **110**, 4985.
- 40 P. Heidari, R. Cheraghali and H. Veisi, *Appl. Organomet. Chem.*, 2016, **30**, 991.

