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Exploring the catalytic activity of Lewis-acidic uranyl complexes in the nucleophilic acyl substitution of acid anhydrides†

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The catalytic activities of several uranyl complexes, such as *N,N'*-disalicylidene-*o*-phenylenediaminato(ethanol)dioxouranium(vi) (UO₂(salophen)EtOH), bis(dibenzoylmethanato)(ethanol)dioxouranium(vi) (UO₂(dbm)₂EtOH), pentakis(*N,N*-dimethylformamide)dioxouranium(vi) ([UO₂(DMF)₅]²⁺), and tetrakis(triphenylphosphine oxide)dioxouranium(vi) ([UO₂(OPPh₃)₄]²⁺), were examined in the nucleophilic acyl substitution of acid anhydrides. Among them, [UO₂(OPPh₃)₄]²⁺ was the most efficient to give ethyl acetate and acetic acid from acetic anhydride (Ac₂O) and ethanol, and was resistant towards decomposition during the catalytic reaction. Several nucleophiles were also subjected to the catalytic acylation reaction using acetic and pivalic anhydride. Kinetic and spectroscopic studies suggested that [UO₂(OPPh₃)₄]²⁺ interacts with Ac₂O to form [UO₂(Ac₂O)(OPPh₃)₃]²⁺. Interaction of this actual catalyst with additional Ac₂O determines the rate of the overall nucleophilic acyl substitution reaction.

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Introduction

Although uranium is undoubtedly the most important element in nuclear engineering, the less abundant ²³⁵U isotope is primarily employed in the practical utilization of atomic power. Accordingly, this fissile isotope has to be enriched in the nuclear fuel process, particularly in light water reactors. As a result, a huge amount of depleted uranium is generated and stored.¹

Recently, there have been numerous studies towards the use of uranium compounds as catalysts in organic syntheses to develop a sophisticated use for depleted uranium that has already been refined.² In the former reports, three main methods appear to be present for organic syntheses: (i) catalysis using organouranium complexes,^{3–8} (ii) the activation of small molecules by low valent uranium complexes,⁸ and (iii) the activation of axial (namely, “-yl”) oxygen atoms in the [O=U=O]ⁿ⁺ moiety (*n* = 1, 2),^{9–14} thus far, the first research direction has succeeded in the coupling reactions of alkynes and the polymerization of lactones, but the latter two do not attain actual catalytic systems and remain as stoichiometric reactions. Furthermore, these reaction systems always require dry

anaerobic conditions because of the low stability of these (pre) catalysts towards oxygen and moisture.

Using the Lewis acidity of uranium should be another alternative for exploring its catalytic activity. Particularly, the uranyl ion ([O=U=O]²⁺, UO₂²⁺) is highly Lewis acidic and exhibits strong hardness according to Pearson's HSAB principle.¹⁵ Therefore, any undesired side reactions arising from the organometallic behaviour of the uranium centre can be avoided, although some organouranyl compounds have been known to be formed in exceptional circumstances.^{16–19} As a matter of fact, several examples of uranyl-catalysed reactions including the alcoholysis of esters,²⁰ Michael additions,^{21–26} Diels–Alder reactions²⁷ and polymerizations^{28–30} have been reported to date. Furthermore, its robust axial structure only allows a planar ligand arrangement in the equatorial coordination plane. Such a ligand set may provide some regio- or chemoselectivity to the reaction of interest.

Although the Lewis basicity of “-yl” oxygen atoms is not very high, they can interact with a strong Lewis acid such as B(C₆F₅)₃.^{31,32} Previously, Chen and co-workers studied the catalytic activity of the oxo ions of V(iv), Mo(vi), W(vi) and Cr(vi) in the nucleophilic acyl substitution of acid anhydrides.^{33,34} Among them, a high efficiency was recorded when using VO(OTf)₂ and MoO₂Cl₂. In their catalytic systems, these oxo ions are believed to be amphoteric and can efficiently activate acid anhydrides. This information strongly motivated us to examine the catalytic activity of the uranyl ion, although the Lewis basicity of its “-yl” oxygen atoms is rather weak.

In this study, several uranyl complexes were tested as Lewis acid catalysts in the nucleophilic acyl substitution to clarify

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† Electronic supplementary information (ESI) available: Progress and efficiency of acylation using uranyl(vi) catalysts, ¹H NMR spectra supporting the decomposition of the catalysts and side reactions, plots of the initial rate under several conditions, and the molar absorption spectra of the uranyl(vi) species. See DOI: 10.1039/c6ra27796a

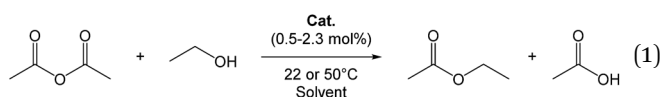


their requirements as a catalyst. Several nucleophiles were also subjected to the catalytic acylation reaction using acetic and pivalic anhydride. Furthermore, kinetic and spectroscopic studies were also performed to understand the catalytic mechanism in detail.

Results and discussion

Catalyst screening

At least one of the equatorial coordination sites of a uranyl complex must be offered for a substrate to be activated. Therefore, one or more leaving ligand(s) should be present in the uranyl catalyst. In this study, *N,N'*-disalicylidene-*o*-phenylenediaminato(ethanol)dioxouranium(vi) ($\text{UO}_2(\text{salophen})\text{EtOH}$, **1**), bis(dibenzoylmethanato)(ethanol)dioxouranium(vi) ($\text{UO}_2(\text{dbm})_2\text{EtOH}$, **2**), pentakis(*N,N*-dimethyl-formamide)dioxouranium(vi) ($[\text{UO}_2(\text{DMF})_5]^{2+}$, **3**), and tetrakis(triphenylphosphine oxide)dioxouranium(vi) ($[\text{UO}_2(\text{OPPh}_3)_4]^{2+}$, **4**) were employed as the catalyst candidates. Fig. 1 shows the schematic structures of these uranyl complexes. Acetic anhydride (Ac_2O) and ethanol (EtOH) were selected as substrates to test these complexes in the acyl substitution reaction (eqn (1)).[‡] The results obtained from the catalyst screening are summarized in Table 1.



First, complexes **1** and **2** were examined in CDCl_3 . Previously, our group reported that **1** and its related complexes tend to release a monodentate ligand, such as EtOH , to form a dimeric complex, $[\text{UO}_2(\text{salophen})]_2$, in chlorinated, non-coordinating solvents such as CHCl_3 and CH_2Cl_2 .³⁵ Therefore, we expected that a substrate will be efficiently activated through this substitution reaction with the initially coordinated EtOH molecule. Nevertheless, no progress of eqn (1) was observed at 22 °C in both systems. When the sample solutions were warmed to 50 °C, the desired reactions started, as shown in Fig. S1 (ESI[†]). To reach a yield >90%, it took 7 h and 4 h at 50 °C when using **1** and **2**, respectively. The obtained products were EtOAc and AcOH only, demonstrating the high selectivity of these systems. On the other hand, the ^1H NMR spectra revealed that both complexes underwent partial or full decomposition. This would be due to AcOH that was generated, as shown in eqn (1). Complex **1** appeared to be more resistant towards the decomposition with acetic acid because of its more stable tetradentate chelate, although new signals arising from several salophen^{2-} species including a free *N,N'*-disalicylidene-*o*-phenylenediamine ($\text{H}_2\text{salophen}$) were also detected, as seen in Fig. S2 (ESI[†]). In contrast, **2** was completely decomposed within 2 h as free dibenzoylmethane (Hdbm) was observed, as seen in Fig. S3 (ESI[†]). Through these experiments, it has been evidenced that the Lewis acidity of these uranyl

[‡] Control experiments in the absence of a uranyl complex did not result in any significant progress of the acylation reaction (e.g., 0.8% yield after 2 h in a CH_2Cl_2 solution of 0.5 M $\text{Ph}(\text{CH}_2)_2\text{OH}$ + 0.5 M Ac_2O).

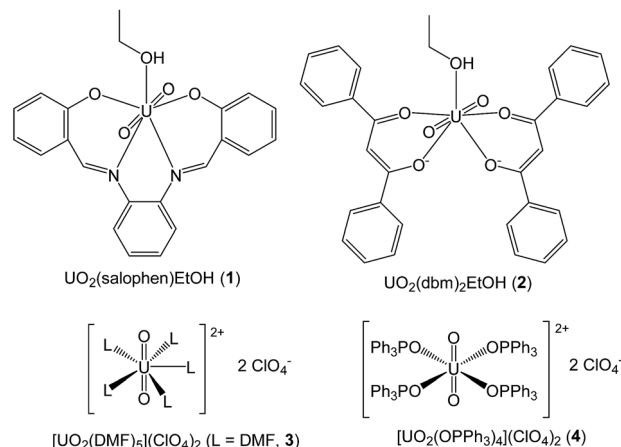


Fig. 1 A schematic structure of the uranyl(vi) complexes tested in this study (EtOH = ethanol).

complexes actually catalyses the acyl substitution reaction of Ac_2O . However, it is better to shift to other uranyl complexes because the mechanism and kinetics cannot be discussed in detail.

As another alternative for the catalyst in the acyl substitution reaction, complex **3** was investigated. Although the equatorial plane of the uranyl ion of this complex is fully occupied with 5-fold coordination offered by the DMF molecules, the exchange or substitution reaction of these monodentate ligands is expected to occur in high frequency because of the labile character of the uranyl ion.^{35–37} Therefore, one or both of the substrates in eqn (1) may enter the uranyl coordination sphere. The reaction progress at 22 °C in the presence of **3** (2.3 mol%) in CD_3CN is displayed in Fig. S4 (ESI[†]). On one hand, the substrates shown in eqn (1) are actually converted to the desired products in 87% yield after 6 h. On the other hand, the ^1H NMR spectrum in Fig. S5 (ESI[†]) indicates the occurrence of unknown products at 8.27 ppm (*cf.* CHO^- of DMF at 8.60 ppm). These substances cannot be a result of the reaction between Ac_2O and DMF. The peak integrals indicate that the concentrations of these undesired products are rather small when compared with that of the desired ester, whereas the uranyl complex (2.3 mol% in total) suffered from this undesired reaction. This means that several mixed ligand complexes other than **3** were also present in the reaction mixture. These *in situ* generated uranyl complexes may also exhibit some catalytic activity in the reaction shown in eqn

Table 1 A summary of the catalyst screening in the nucleophilic acyl substitution of acetic anhydride using ethanol^a

Catalyst (mol%)	Solvent	$T/^\circ\text{C}$	Time/h	Yield/%
1 (0.5)	CDCl_3	50	7	90
2 (0.5)	CDCl_3	50	4	90
3 (2.3)	CD_3CN	22	6	87
4 (1.3)	CD_2Cl_2	22	2	97

^a Reaction conditions: 0.50 M Ac_2O , 0.50 M EtOH and the uranyl complex were loaded in CD_2Cl_2 .



(1). The occurrence of unknown catalysts makes further discussion on the kinetics and mechanism complicated. Therefore, it is necessary to find a monodentate ligand with high stability towards acid anhydrides. Nevertheless, it should be emphasized that the higher catalytic activity of a uranyl complex was successfully attained in use of the solvate complex 3 as demonstrated by the lower reaction temperature.

Triphenylphosphine oxide (OPPh₃) is one of strong monodentate ligands for uranyl and could be durable enough to remain unchanged under the reaction conditions shown in eqn (1) because of its small potential to react with acid anhydrides and carboxylic acid. An isolable form of the uranyl-OPPh₃ complex is [UO₂(OPPh₃)₄]²⁺ (4), which is usually obtained as a salt of a poorly coordinating anion such as ClO₄⁻ and TfO⁻.^{38–40} The catalytic activity of 4 was examined in a similar manner to the abovementioned test systems. Fig. 2 shows the reaction progress in the presence of 4 (1.3 mol%) in CD₂Cl₂ at 22 °C. As a consequence, the products in eqn (1) were formed in 97% yield within 2 h. It is noteworthy that OPPh₃ remains unchanged in this system. In summary, the catalyst screening in Table 1 concluded that 4 is, in hand, the best uranyl complex to catalyse the reaction shown in eqn (1) from the viewpoints of activity and durability. Usually, the most popular coordination number of uranyl complexes in its equatorial plane is 5, as revealed by 1, 2, and 3 (Fig. 1). Even in the complexation of uranyl with OPPh₃, the pentacoordinate complex, [UO₂(OPPh₃)₅]²⁺ (5), tends to be formed below –40 °C (*vide infra*), where the thermal vibrations of the bulky OPPh₃ ligands decrease. When the solution is warmed, complex 5 tends to release OPPh₃ to give the 4-fold complex 4 because of the steric hindrance between the neighboring OPPh₃ molecules in the coordination sphere. As a result, some Lewis-acidic vacancies would be still left in the uranyl

equatorial plane of 4, which can be offered to the activation of a substrate in the current catalytic system.

Acyl substitution reactions catalyzed by complex 4

With the optimal catalyst 4, the catalytic acetylation and pivalation of several nucleophiles, including alcohols and thiophenol, were examined. The results are summarized in Table 2. In most cases except for ternary alcohols, the desired acylation products were obtained in good yield together with the carboxylic acids corresponding to the anhydride used. As a general trend, the acetylation was much faster than the corresponding pivalation, implying that the steric hindrance of an anhydride affects the reaction rate. When compared to entry 1, a much shorter reaction time was required for both the acetylation and pivalation of 2-phenylethyl alcohol (Ph(CH₂)₂OH, entry 2). This suggests that coordination of less sterically hindered nucleophiles, such as EtOH, inhibits the catalytic activity of the uranyl ion. As a matter of fact, the reaction rate decreased, when extra EtOH was added to the reaction (Fig. S6, ESI[†]). Such inhibition could also be the case in entry 3, where the carbonyl and hydroxyl groups of hydroxyacetone may interact with or chelate to the uranyl ion to prevent the desired acylation reaction. Nevertheless, this nucleophile is still amenable to acylation in quantitative yield, although a longer reaction time than Ph(CH₂)₂OH is required.

Entry 4 shows that *tert*-butyl alcohol is actually reactive in the current system even without any additional treatments such as raising the temperature and the addition of ternary amines. This point is distinguishable from the acylations catalyzed by vanadyl and molybdenyl ions, where this tertiary alcohol was completely inert towards the acylation reaction.^{33,34} The yield of *tert*-butyl acetate in entry 4 was 68% after 4 h, while Ac₂O already disappeared in the sample solution. This result was reproducible. Except for the expected products and leftover *tert*-butyl alcohol, the ¹H NMR spectrum of this sample (Fig. S7, ESI[†]) revealed the presence of isobutene [δ /ppm: 1.73 (t, J_{HH} =

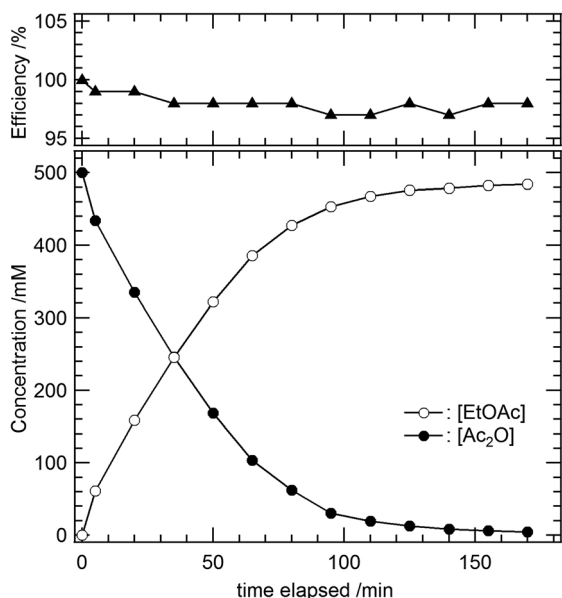


Fig. 2 The progress and efficiency of the acetylation of ethanol catalysed by [UO₂(OPPh₃)₄]²⁺ (4, 6.5 mM, 1.3 mol%) in CD₂Cl₂ at 22 °C. Initial conditions: [Ac₂O] = 0.50 M and [EtOH] = 0.50 M.

Table 2 [UO₂(OPPh₃)₄]²⁺-catalyzed acetylation (R' = CH₃) and pivalation (R' = C(CH₃)₃) of various nucleophiles^a

Entry	Nucleophile	Time/h	Yield ^b /%
1	C ₂ H ₅ OH	2 (24)	97 (90)
2	Ph(CH ₂) ₂ OH	0.75 (3)	95 (96)
3 ^c	H ₃ CC(=O)CH ₂ OH	2.5	>99
4	<i>tert</i> -BuOH	4	68
5	Ph ₃ COH	72	63
6	PhOH	2	95
7	PhSH	6	85

^a Reaction conditions: 0.50 M anhydride, 0.50 M nucleophile and 6.5 mM (1.3 mol%) complex 4 were loaded in CH₂Cl₂ or CD₂Cl₂ at 22 °C, unless stated otherwise. Data in the parentheses correspond to the pivalation reaction. ^b Determined by ¹H NMR peak integrals. ^c Catalyst: 2.0 mol%.



1.2 Hz, 6H, $\text{H}_2\text{C}=\text{C}(\text{CH}_3)_2$), 4.66 (septet, $J_{\text{HH}} = 1.2$ Hz, 2H, $\text{H}_2\text{C}=\text{C}(\text{CH}_3)_2$), which is the product resulting from the E_1 reaction of *tert*-butyl alcohol. Another product in this side reaction was water, which also consumes the anhydride to give the corresponding carboxylic acid. Trityl alcohol was also tested as another sterically hindered nucleophile (entry 5). As a result, the progress of the acetylation of this bulky nucleophile was actually observed, although the reaction was so sluggish that it took 3 days to reach 63% yield. This is also much different from the inertness of vanadyl and molybdenyl catalysts in the acylation of trityl alcohol.^{33,34} Such a difference may be related to the size of the metal center because the ionic radius of U^{6+} (0.73 Å at 6-coordinated) is much greater than those of V^{4+} (0.53 Å at 5-coordinated) and Mo^{6+} (0.59 Å at 6-coordinated).⁴¹

Phenol and thiophenol (entries 6 and 7) provide a good contrast of the difference in nucleophilic moieties. The nucleophilicity of PhS^- is widely known to be much greater than that of PhO^- . Nevertheless, phenol as a nucleophile exhibited a higher efficiency than thiophenol in terms of the reaction rate and yield. This difference appears to be related to the hardness of the nucleophilic atoms, although the details are still unclear.

Reaction kinetics and catalysis mechanism

Mechanistic insights into the nucleophilic acyl substitution reaction catalyzed by **4** were obtained by investigating how the concentrations of the substrates affected the reaction rate. Herein, the acetylation of $\text{Ph}(\text{CH}_2)_2\text{OH}$ (entry 2, Table 2) was employed as a model reaction. Fig. 3a shows the progress of this reaction in CD_2Cl_2 at 22 °C under different initial concentrations of Ac_2O ($[\text{Ac}_2\text{O}]_{\text{ini}}$). Upon increasing $[\text{Ac}_2\text{O}]_{\text{ini}}$, the reaction rate clearly increased. The initial rate (v_{ini}) estimated from Fig. 3a was plotted against $[\text{Ac}_2\text{O}]_{\text{ini}}$ as shown in Fig. S8 (ESI†). As a result, v_{ini} versus $[\text{Ac}_2\text{O}]_{\text{ini}}^2$ shows a linear relationship with a slope equal to $1.04 \times 10^{-3} \text{ M}^{-1} \text{ s}^{-1}$. We also examined the dependency of v_{ini} on the initial concentration of the nucleophile ($[\text{Ph}(\text{CH}_2)_2\text{OH}]_{\text{ini}}$). However, the progress of the reaction was almost the same regardless of the different concentrations of $[\text{Ph}(\text{CH}_2)_2\text{OH}]_{\text{ini}}$, as shown in Fig. S9 (ESI†) at least under the tested conditions ($[\text{Ph}(\text{CH}_2)_2\text{OH}]_{\text{ini}} = 0.25\text{--}1.00 \text{ M}$). This means that $\text{Ph}(\text{CH}_2)_2\text{OH}$ does not participate in the rate-determining step of this catalytic acyl substitution reaction.

When Ac_2O interacts with **4**, its OPPh_3 ligand(s) may dissociate from UO_2^{2+} . Therefore, we also studied the dependency of v_{ini} on the concentration of free OPPh_3 ($[\text{OPPh}_3]_{\text{free}}$). The results are displayed in Fig. 3b. With increasing $[\text{OPPh}_3]_{\text{free}}$, the reaction rate decreases. In Fig. S10 (ESI†), v_{ini} is proportional to $[\text{OPPh}_3]_{\text{free}}^{-1}$ with a slope equal to $6.54 \times 10^{-7} \text{ M}^2 \text{ s}^{-1}$.

In conclusion, the following is the rate equation for the acetylation of $\text{Ph}(\text{CH}_2)_2\text{OH}$ catalyzed by **4**.

$$v_{\text{ini}} = k[\text{Ac}_2\text{O}]^2[\text{OPPh}_3]^{-1}[\mathbf{4}] \quad (2)$$

The rate constant (k) of eqn (2) was estimated to be $4.02 \times 10^{-4} \text{ M}^{-1} \text{ s}^{-1}$ at 22 °C.

Further mechanistic information should be included in eqn (2) because multi-body collision is unlikely to occur. Thus, there should be an equilibrium preceding the rate-determining

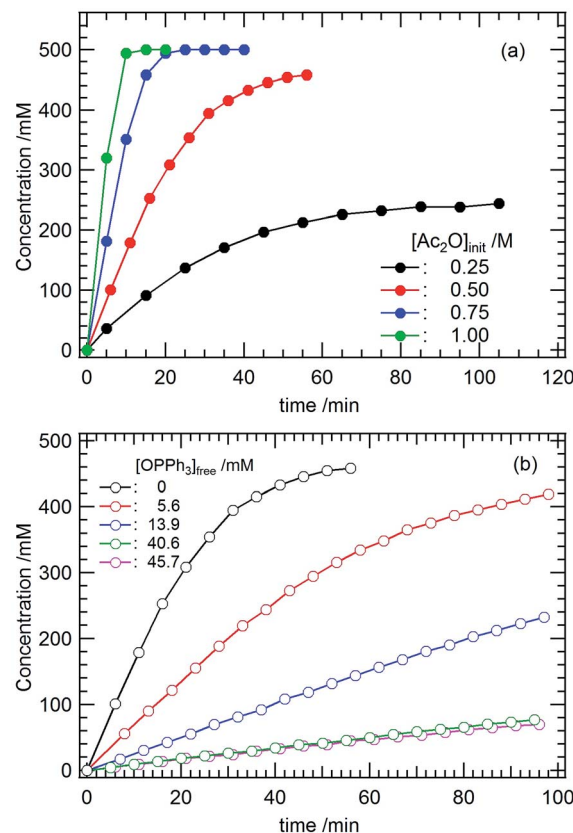
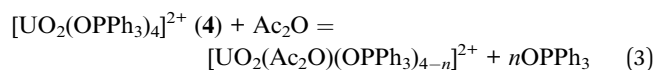


Fig. 3 The progress of the acetylation of $\text{Ph}(\text{CH}_2)_2\text{OH}$ catalysed by $[\text{UO}_2(\text{OPPh}_3)_4]^{2+}$ (**4**) in CD_2Cl_2 at 22 °C. (a) $[\text{Ac}_2\text{O}]$ dependency; initial condition: 0.50 M $\text{Ph}(\text{CH}_2)_2\text{OH}$, 0.25–1.00 M Ac_2O and 6.5 mM **4**, and (b) $[\text{OPPh}_3]$ dependency; initial condition: 0.50 M $\text{Ph}(\text{CH}_2)_2\text{OH}$, 0.50 M Ac_2O , 6.5 mM **4** and 0–45.7 mM OPPh_3 .

step. When Ac_2O was added to the CD_2Cl_2 solution containing **4**, the yellow color arising from the uranyl compound immediately diminished, implying ligand substitution between the coordinated OPPh_3 and Ac_2O . In contrast, any interaction of $\text{Ph}(\text{CH}_2)_2\text{OH}$ with **4** was unlikely to occur despite its nucleophilicity because v_{ini} is independent of $[\text{Ph}(\text{CH}_2)_2\text{OH}]_{\text{ini}}$, as shown in Fig. S9 (ESI†). This could be ascribed to steric hindrance arising from the phenyl groups. The complexation reaction between **4** and Ac_2O was studied in detail using a UV-vis absorption spectrophotometric titration study. The obtained spectral series recorded at different $[\text{Ac}_2\text{O}]$ is shown in Fig. 4. Upon increasing $[\text{Ac}_2\text{O}]$, a decrease in the intensity of the finely structured absorption bands arising from ligand-to-metal charge transfer (LMCT) in the UO_2^{2+} moiety was observed together with a slight red shift in the peak maxima. This observation indicates that OPPh_3 in **4** was substituted by Ac_2O as shown in eqn (3).



Although no isosbestic points were observed in Fig. 4, principal component analysis (PCA)^{42,43} has suggested that a total of



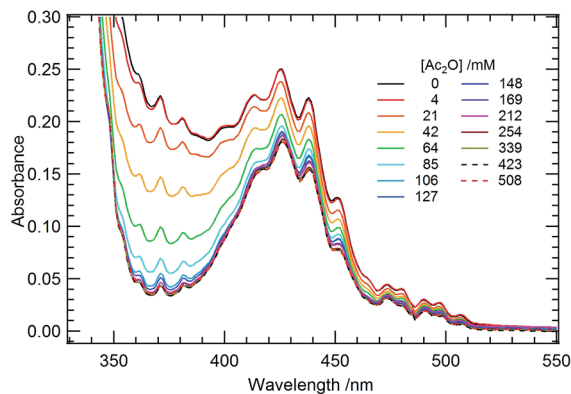
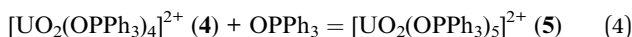


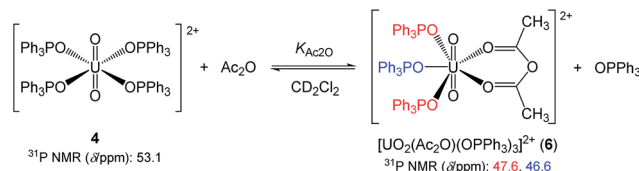
Fig. 4 UV-vis absorption spectra of a CH_2Cl_2 solution of $4(\text{ClO}_4)_2$ (10 mM) at 22 °C. $[\text{Ac}_2\text{O}]$ varied from 0 to 0.508 M.

2 components are involved in this spectral series. This result corroborates the occurrence of eqn (3) in the current system. In addition, the Job plot shown in Fig. S11 (ESI[†]) also suggests a 1 : 1 interaction between **4** and Ac_2O in eqn (3).

Although additional coordination to the fifth equatorial site of **4** is possible in a cooled solution (below -40 °C) as we have demonstrated the occurrence of $[\text{UO}_2(\text{OPPh}_3)_5]^{2+}$ (**5**) in the former,⁴⁰ there is not enough room for another ligand in the equatorial plane of **4** at 22 °C because of the thermal motion of the coordinated OPPh_3 molecules.



Therefore, the exclusion of OPPh_3 from the first coordination sphere of UO_2^{2+} should be convincing. For the equilibrium analysis of the data in Fig. 4, it is necessary to know the number of released OPPh_3 , n , in eqn (3). Fig. 5 displays the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum of the CD_2Cl_2 solution containing **4** (14.7 mM) and Ac_2O (0.40 M) at -70 °C. The sample had to be cooled to suppress the frequency of the chemical exchange of the ligands bound to UO_2^{2+} . In accordance with our former findings,⁴⁰ the ^{31}P signals at 53.1 ppm and 41.4 ppm were attributed to **4** and **5**, respectively. The new signals at 47.6 and 46.6 ppm appear upon the addition of Ac_2O . Upon increasing $[\text{Ac}_2\text{O}]$, the peak integrals of these signals increase, while the ratio between them remains constant at 2 : 1. Therefore, these signals should arise from the OPPh_3 ligands in the different environments of a single uranyl complex other than **4** and **5**. The most plausible species occurring in this test solution is $[\text{UO}_2(\text{Ac}_2\text{O})(\text{OPPh}_3)_3]^{2+}$ (**6**) as

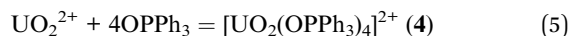


Scheme 1 The ligand substitution equilibrium between **4** and $[\text{UO}_2(\text{Ac}_2\text{O})(\text{OPPh}_3)_3]^{2+}$ (**6**) in CD_2Cl_2 at -70 °C.

shown in Scheme 1, where two of the OPPh_3 are chemically equivalent (red) and there is also another OPPh_3 in a different environment (blue). The coordination mode of Ac_2O may be bidentate in a similar manner to a β -diketonato ligand.

Note that no signals were observed at 29 ppm, where the ^{31}P signal of free OPPh_3 appeared. This implies that the concentration of free OPPh_3 was negligible. It is necessary to consider that OPPh_3 in the product part of Scheme 1 is also involved in eqn (4). In accordance with our previous report,⁴⁰ the logarithmic equilibrium constant of eqn (4) ($\log K_{45}$) at -70 °C is 3.33, which reveals that $[\text{OPPh}_3]_{\text{free}}$ is as small as 7×10^{-6} M under the current conditions. This is the reason why the free OPPh_3 was undetectable in Fig. 5.

The reaction in Scheme 1 was taken into account to analyze the UV-vis absorption spectral series shown in Fig. 4. Note that this treatment assumes that the equilibrium between **4** and **6** is also present in the current system at 22 °C. It was also necessary to consider the following complexation to form **4**.



A sufficiently large gross stability constant of **4** ($\beta_4 = [\mathbf{4}]/[\text{UO}_2^{2+}][\text{OPPh}_3]_{\text{free}}^4$, e.g., $\log \beta_4 = 20$) was assumed to simulate the predominant formation of **4** and practically negligible concentrations of free UO_2^{2+} . Formation of **5** can be ruled out at 22 °C because $\log K_{45}$ is as low as -1.31 . Analysis using HypSpec software⁴⁴ resulted in the logarithmic equilibrium constant of Scheme 1 ($\log K_{\text{Ac}_2\text{O}}$) equal to -0.513 ± 0.013 at 22 °C. The characteristic LMCT bands arising from the UO_2^{2+} moiety were assigned to both **4** and **6** as shown in Fig. S12 (ESI[†]), implying that the equilibrium analysis was successfully completed. It was also confirmed that the value of $K_{\text{Ac}_2\text{O}}$ was independent of the assumption of β_4 unless the occurrence of free UO_2^{2+} was significant in the simulation (where $\log \beta_4 < 12$).

Assuming that complex **6** is an actual catalyst in the acetylation of $\text{Ph}(\text{CH}_2)_2\text{OH}$, eqn (2) can be modified using $K_{\text{Ac}_2\text{O}} = [\mathbf{6}]/[\text{OPPh}_3]_{\text{free}}/[\mathbf{4}][\text{Ac}_2\text{O}]$ as follows:

$$v_{\text{ini}} = (k/K_{\text{Ac}_2\text{O}})[\text{Ac}_2\text{O}][\mathbf{6}] \quad (6)$$

The rate constant $k/K_{\text{Ac}_2\text{O}}$ is equal to $1.31 \times 10^{-3} \text{ M}^{-1} \text{ s}^{-1}$ at 22 °C. This rate equation indicates that the rate-determining step was the interaction between **6** and Ac_2O . Thus, $[\text{UO}_2(\text{Ac}_2\text{O})_2(\text{OPPh}_3)_3]^{2+}$ (**7**) will be formed as an intermediate to activate additional Ac_2O , followed by nucleophilic attack of $\text{Ph}(\text{CH}_2)_2\text{OH}$ to the electrophilic carbon of the carbonyl moiety in the activated Ac_2O . Taking into account the following points, (i) the stoichiometry of **7**, (ii) the absence of leaving OPPh_3 supported by eqn

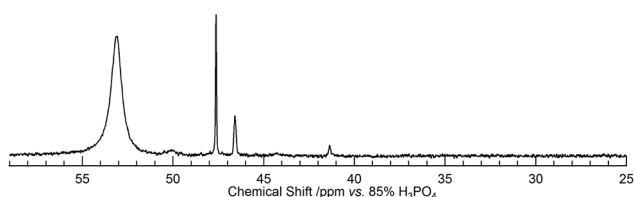
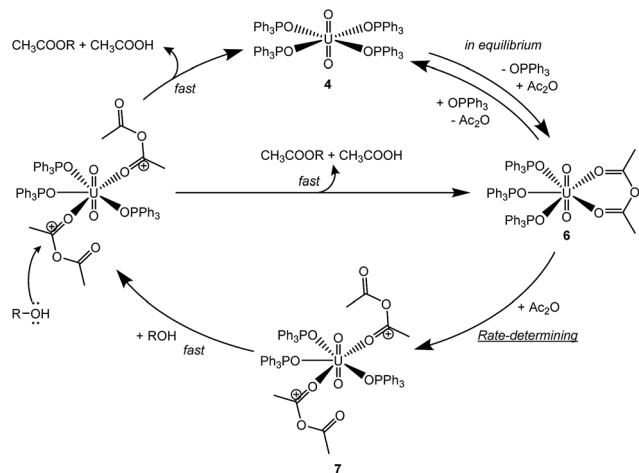


Fig. 5 $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum of CD_2Cl_2 solution dissolving **4** (14.7 mM) and Ac_2O (0.40 M) at -70 °C.





Scheme 2 The proposed cycle for the nucleophilic acyl substitution reaction of Ac_2O catalyzed by **4** ($\text{R} = \text{Ph}(\text{CH}_2)_2-$). The charges on the uranyl complexes are omitted for clarity.

(6) and (iii) the potential bidentate coordination manner of Ac_2O , UO_2^{2+} has to offer 7 coordination sites in its equatorial plane, being too crowded. Therefore, one or both of the coordinated Ac_2O molecules should be monodentate. In the latter case, the coordination number in the equatorial plane is 5, which is most commonly found in uranyl complexes. Although 6-fold uranyl complexes are also often observed, such a structure is allowed only by the chelating ligands bearing narrow bite angles such as CO_3^{2-} and NO_3^- .^{45,46} In summary, the proposed catalytic cycle of the nucleophilic acyl substitution is shown in Scheme 2. This catalysis mechanism is plausible, but still somewhat hypothetical. Our trials to crystallize **6** and uranyl complexes related to intermediate **7** have not been successful to date.

Lewis-acid catalysis is widely employed in various organic syntheses. A typical example is the series of triflate salts of several d-block metals and trivalent rare earths used in aldol reactions, where the Lewis-acid catalysts are much easier to handle than those traditionally employed, such as AlCl_3 and $\text{BF}_3 \cdot \text{Et}_2\text{O}$, because of their resistance towards hydrolysis.⁴⁷ The uranyl species also does not show a strong tendency towards hydrolysis and therefore, can be present stably in aqueous or aqueous/organic biphasic reaction systems. Furthermore, the dioxo structure of uranyl may result in the equatorial coordination shown in this study, which is much distinguishable from the known spherical metal ions as water-compatible Lewis-acid catalysts. We intend to expand the coordination chemistry of uranyl to explore its catalytic functions.

Conclusions

In this study, we performed screening of several uranyl complexes as Lewis-acid catalysts used in the nucleophilic acyl substitution reaction. As a result, all the uranyl species tested were found to more or less promote this reaction to afford an ester and carboxylic acid. The most active uranyl complex in hand was $[\text{UO}_2(\text{OPPh}_3)_4]^{2+}$ (**4**), which gives ethyl acetate in 97%

yield within 2 h at 22 °C in CD_2Cl_2 . Several nucleophiles were used in the acetylation and pivalation reactions using the corresponding acid anhydrides in the presence of **4**. In all the entries reported herein, the progress of these reactions was successfully enhanced by **4**. Even when using ternary alcohols such as *tert*-BuOH and trityl alcohol, **4** still shows its activity, although progress of the reaction in both systems is somewhat slow. According to the kinetic and mechanistic studies, **4** is in an equilibrium with $[\text{UO}_2(\text{Ac}_2\text{O})(\text{OPPh}_3)_3]^{2+}$ (**6**). Therefore, the former is a pre-catalyst and the latter is an actual catalyst in the current nucleophilic acyl substitution reaction.

Experimental

Materials

The uranyl complexes $\text{UO}_2(\text{salophen})\text{EtOH}$ (**1**),⁴⁸ $\text{UO}_2(\text{dbm})_2 \cdot \text{EtOH}$ (**2**),⁴⁹ $[\text{UO}_2(\text{DMF})_5](\text{ClO}_4)_2$ (**3**),⁵⁰ and $[\text{UO}_2(\text{OPPh}_3)_4](\text{ClO}_4)_2$ (**4**)⁴⁰ (Fig. 1) were prepared by the methods described elsewhere. All the chemicals used in this study were of reagent grade and used as received.

Kinetic and mechanistic studies using NMR spectroscopy

^1H and $^{31}\text{P}\{^1\text{H}\}$ NMR spectra of the sample solutions were recorded on a JEOL ECX-400 NMR spectrometer (^1H : 399.78 MHz, ^{31}P : 161.83 MHz). Tetramethylsilane (TMS) and 85% H_3PO_4 were employed as reference materials for the ^1H and ^{31}P chemical shifts, respectively. A typical procedure used to study the reaction kinetics of the catalytic acyl substitution reaction is described below. The uranyl catalyst was loaded in an NMR sample tube (5 mm O.D.) together with CD_2Cl_2 and acid anhydride. After optimization of the NMR shim, a nucleophile was injected to the NMR tube. The sample solution was vigorously shaken, followed by starting the kinetic experiment. The ^1H NMR spectrum was recorded every 5–30 min. The concentrations of the reactants and products were estimated from the peak integrals of the species occurring in the reaction mixture.

UV-vis titration experiments

A CH_2Cl_2 solution containing complex **4** (10 mM) was prepared. This solution was titrated with neat Ac_2O . At each titration step, the UV-vis absorption spectrum of the sample solution was recorded on a Agilent 8453 photodiode array spectrophotometer. The obtained spectral series was analyzed using HypSpec software.⁴⁴

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