



Ruthenium nanoparticles-modified reduced graphene prepared by a green method for high-performance supercapacitor application in neutral electrolyte[†]

Hagar K. Hassan,^a Nada F. Atta,^a Maher M. Hamed,^a Ahmed Galal^{*a} and Timo Jacob^{*bcd}

Here, we report a one-pot preparation method of a ruthenium-based reduced graphene oxide hybrid (Ru_{nano} -based RGO). The synthesis is based on a single step for the reduction of Ru^{3+} (RuCl_3) into Ru nanoparticles, and graphene oxide (GO) into reduced graphene oxide (RGO) without the use of any reducing agent. The structure of the hybrid material was ascertained using X-ray diffraction (XRD), X-ray photoelectron spectroscopy (XPS), Raman spectroscopy and high resolution transmission spectroscopy (HRTEM). The change in pH of the synthesis media resulted in different structures and morphologies. A mixture of Ru and RuO_2 (3 : 2) was obtained at pH 8.0, while only Ru-nanoparticles were obtained at pH 4.0. The supercapacitor performance of the hybrid was investigated in a neutral electrolyte, 1.0 M NaNO_3 , in a three-electrode setup. Ru_{nano} -based RGO prepared at pH 8 exhibited a maximum specific capacitance of 270 F g^{-1} (238.5 mF cm^{-2}), and maximum energy and power densities of 15 W h kg^{-1} and 76.4 kW kg^{-1} , respectively. The proposed supercapacitor has a stable potential window up to 1.1 V with very good cycling stability over 5000 cycles at 24 A g^{-1} .

Received 27th November 2016
Accepted 26th January 2017

DOI: 10.1039/c6ra27415c
rsc.li/rsc-advances

1. Introduction

Research and technology development for energy sources and storage alternatives are attracting considerable attention. This is because of the depletion of the energy supply from fossil fuels and increasing environmental pollution.¹ Another challenge is electricity storage with high power and energy densities.²

Electric double layer capacitors (EDLCs) are promising energy storage devices that store electric charges at the interface between the electrode surface and the electrolyte.³ Another class of capacitors is pseudocapacitors, in which the capacitance arises from a reversible reaction that takes place at the interface. Both EDLCs and pseudocapacitors belong to the class of supercapacitors or ultracapacitors due to their high specific capacitance compared to conventional capacitors. Moreover, supercapacitors provide several advantages over batteries such as high power density, long cycling stability, fast charging/discharging rate and lower environmental impact.⁴

Supercapacitors however have lower energy densities compared to batteries. Research work on the usage of energy storage devices is targeting the enhancement of the properties of supercapacitors for usage in many applications. Beside the low energy density of supercapacitors, pseudocapacitors have poor cycling stability.⁵ Construction of hybrid materials (for combination between EDLCs and pseudocapacitors) is one of the challenges to overcome the problem of both EDLCs and pseudocapacitors and to improve the properties of supercapacitors.⁶ This might allow for a maximum energy density at high power density without sacrificing the cycling stability.

In recent years, graphene attracted a significant interest as a promising electrode material for energy storage applications.⁷ Graphene-based supercapacitors showed high theoretical specific capacitance up to 550 F g^{-1} .⁸ Since restacking or aggregation of its sheets due to van der Waals forces limits reaching its theoretical specific capacitance value in real applications, modifying the graphene sheets in order to avoid their aggregation is of prime importance.

Transition metals and their oxides are interesting candidates to act as spacers between graphene sheets to prevent their aggregation and hence, to increase the specific capacitance of graphene-based supercapacitors.⁵ Here, RuO_2 is one of the most promising materials for supercapacitor applications due to its high specific capacitance from 1300 to 2200 F g^{-1} ,^{9,10} high rate capability, highly reversible faradic reaction and high thermal

^aDepartment of Chemistry, Faculty of Science, Cairo University, 12613 Giza, Egypt.
E-mail: galal@sci.cu.edu.eg

^bInstitute of Electrochemistry, Ulm University, 89081 Ulm, Germany. E-mail: timo.jacob@uni-ulm.de

^cHelmholtz-Institute-Ulm (HIU) [d], Helmholtzstr. 11, 89081 Ulm, Germany

^dKarlsruhe Institute of Technology (KIT), P.O. Box 3640, 76021 Karlsruhe, Germany

† Electronic supplementary information (ESI) available. See DOI: [10.1039/c6ra27415c](https://doi.org/10.1039/c6ra27415c)



stability.¹¹ However, Ru-based materials are usually suffering from their poor porosity, less cycling stability and rapid decrease of power density especially at high charge/discharge rates as well as the high price of Ru-based materials that limits their commercial use.⁵ The properties of Ru-based materials (*i.e.* Ru, RuO₂ and Ru-based perovskites) can be strongly altered by changing the method of their preparation or changing the material they are grown over. There are several works related to the preparation of RuO₂ over various carbon materials such as mesoporous carbon,¹² activated carbon,⁵ graphene^{10,13-17} or carbon nanotubes.¹⁸ However, extra improvement is still required because almost all of the previous RuO₂-based carbon supercapacitors are limited to operate in either highly acidic or basic electrolytes. This is in order to harvest the benefits of the high capacitance value of RuO₂ in spite of the fact that such electrolytes are not targets for the next generation of supercapacitors due to environmental issues. Additionally, some of the previous methods require long preparation times, sophisticated preparation methods, the use of hazardous materials or further annealing steps. Therefore, it is preferred to construct a supercapacitor based on Ru-based materials, using a neutral electrolyte as well as a preparation method based on a “green” and simple (*e.g.* one-step) method. This can be achieved by tuning the structure and/or morphology of Ru-based materials through changing its preparation method, so that for instance nanoparticulate materials can be obtained. Wang *et al.*⁹ reported the preparation of Ru-based materials with smaller particle size decreasing the diffusion resistance, facilitating the proton transport in the bulk of RuO₂ and increasing the specific capacitance as well as the storage efficiency.

In this work, we introduce the synthesis of Ru/RGO *via* a one-step, green and efficient method and its application as a supercapacitor electrode in a neutral electrolyte. The current method is based on the simultaneous reduction of a Ru precursor and GO in a single step and without using any reducing agent and under pH control. The mass percentage of the Ru precursor does not exceed 25% of the total weight of the material in order to minimize the cost of the capacitor. The proposed preparation method is accomplished without any annealing or calcination process. This is advantageous in reducing the energy consumption during the preparation step.

2. Experimental

2.1 One-step preparation of Ru-RGO

Graphene oxide was prepared by a modification of Hummer's method, following the same procedure as reported previously.¹⁹ Briefly, 5 g of high purity graphite was cured with H₂SO₄, P₂O₅ and K₂S₂O₈ in order to prepare the per-oxidized graphite. This is followed by stirring 5 g of dried per-oxidized graphite with 115 mL of concentrated H₂SO₄ in an ice bath for about 10 minutes, then 15 g of KMnO₄ was gradually added and stirring was continued for two more hours. The mixture acquired a bright yellow color after its dilution by water and treatment with H₂O₂. The bright yellow suspension was filtered and washed with 1 : 10 HCl solution, afterwards it was dried overnight at 80 °C in

an oven. Ru/RGO was prepared by adding 0.2 g of GO to 0.069 g of RuCl₃ and dissolved in distilled water followed by ultrasonication of the mixture for two hours. Ammonia solution was used to adjust the pH of the suspension to (2.0, 4.0 and 8.0), then it was placed in a conventional microwave that was operated at 720 W using 30 s cycles (20 s: on and 10 s: off) until the ignition process takes place; the total reaction time is 30 minutes. During the irradiation, the suspension becomes viscous with time and dehydrated. The resulting powder was ignited and a strong firing takes place, resulting in a black powder that indicates the successful reduction of RGO. For comparison, the chemically reduced graphene oxide (RGO) was also prepared following the procedure mentioned in ref. 20.

2.2 Structural, spectral and surface analysis

The prepared materials were characterized using X-ray diffraction (XRD), Raman spectroscopy, X-ray photoelectron spectroscopy (XPS), field-emission scanning electron microscope (FESEM) with energy-dispersive analysis by X-ray (EDX) and high-resolution transmission electron microscope (HR-TEM), the XRD was recorded by Panalytical X'Pert using Cu-K_α radiation ($\lambda = 1.540 \text{ \AA}$), the Raman spectra were obtained using Witec alpha 300R confocal Raman microscope with a 532 nm Nd:Yag laser. XPS measurements were performed with a Perkin Elmer 5300 XPS system with a non-monochromatised Mg-K_α X-ray source. Calibration was performed using the C 1s component (binding energy is 284.6 eV). An Mg-K_α X-ray was used with 300 W applied to the anode. The surface morphological analysis was performed using HR-TEM (Tecnai G20, FEI, Netherland, 200 kV, LaB6 Gun), FESEM and EDX (JEOL JSM-6360LA and Philips XL30).

2.3 Electrode fabrication and electrochemical measurements

An amount of 20 mg of the prepared Ru_{nan}-based RGO was sonicated with 1 mL of dimethyl formamide (DMF) for about 30 minutes in order to prepare a homogeneous suspension of 20 mg mL⁻¹ Ru-RGO suspension. By the same way we also prepared 20 mg mL⁻¹ of RGO (sonicated until a homogeneous suspension obtained). The electrodes used for the electrochemical capacitance measurements were prepared by adding 5 μ L of 2% Nafion solution on the surface of well-polished mirror-like Ni electrodes of diameter 0.6 cm (surface area 0.283 cm²) and connected with a copper rod for electrical conductivity. Then the Ni/Nafion electrode was dried at 80 °C followed by casting the active material (Ru_{nan}-based RGO or RGO) and left to dry in an oven in order to obtain Ni/N/Ru_{nan}-based RGO and Ni/N/RGO. The masses of the electroactive materials were measured as 0.25 mg for both Ru_{nan}-based RGO and RGO.

The electrochemical measurements were performed in a conventional three-electrode system using Ag/AgCl (4.0 M KCl) as reference electrode and a Pt wire as counter electrode. Cyclic voltammetry and galvanostatic charging/discharging experiments were performed using a Voltalab PGZ301 potentiostat. Electrochemical impedance spectroscopy (EIS) was performed using the same instrument at OCP from 100 mHz to 100 kHz with

an AC amplitude of 10 mV. Cyclic voltammetry curves were recorded from $E_i = 0.0$ V to $E_f = 0.7, 0.8, 0.9, 1.0$ and 1.1 V at scan rates ranging from 5.0 to 200 mV s $^{-1}$. Galvanostatic charging/discharging curves were recorded using chronopotentiometry at various current densities ranging from 1.0 A g $^{-1}$ to 26 A g $^{-1}$ and limiting the potential between 0.0 V to 0.9 V. The specific capacitance was calculated from cyclic voltammetry using eqn (1) and from galvanostatic charging-discharging data using eqn (2). Eqn (3) and (4) are used to calculate the energy and the power densities, respectively.

$$C_{sp} = \frac{\bar{I}}{\nu m \Delta V} \quad (1)$$

$$C_{sp} = \frac{I t_d}{m \Delta V} \quad (2)$$

$$E = \frac{1}{2} \frac{C_s \Delta V^2}{3.6} \quad (3)$$

$$P = \frac{E}{t_d} \times 3600 \quad (4)$$

\bar{I} is the integration of area enveloped inside the CVs, ΔV is the potential window in volts ($V_f - V_i$), ν is the scan rate, C_{sp} is the specific capacitance in F g $^{-1}$, m is the mass of electroactive material in grams, I is the current in ampere, and t_d is the discharge time in seconds. For measuring power density and energy density IR drop was excluded from potential value. The factors 3.6 and 3600 comes from the unit conversions of time and mass from g to kg and from s to h, respectively.

3. Results and discussion

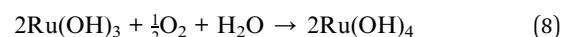
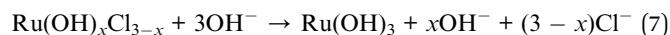
By changing the preparation method one is able to form particles of different morphologies, sizes and of different structures as well. This should additionally alter the electric and electrochemical properties of the final hybrid. Hence, in the current study we investigated the effect of microwave irradiation on the precursors GO and RuCl₃ in the absence of any reducing, stabilizing or precipitating agents. The microwave irradiation was used to provide a simultaneous reduction of both GO and Ru³⁺, and the structure of the obtained material was investigated. Microwave chemistry is the science that uses microwave irradiation to boost chemical reactions.²¹ There are two main heating mechanisms during the microwave irradiation; dipolar polarization and ionic conduction. Generally, microwave heats any polar molecules such as water or the conducting ions in a solvent or in a solid. Dipolar mechanism includes the re-orientation of water molecules under the effect of alternating electric field. The frictions and collisions between molecules during the reorientation as well as the rotation are responsible for the heating effect.²¹ In the case of ions, the ionic conduction is based on the movement of ions in the solution under the effect of the orientation of the electric field. And since the fluctuation of the electric field is constant, the direction of ions movement is constant through the solution. This will result in an increase in the temperature due to frictions and collisions.

So, the semiconducting or conducting samples heat up when their ions lose their electric energy that is generated during the irradiation. This loss in the electric energy is due to the electrical resistance of the materials.²² Microwave chemistry is therefore based on the efficient superheating of the material so that the product may acquire different properties. There is extensive work using microwave irradiation to reduce metal nanoparticles in the presence of a reducing agent such as hydrazine hydrate, urea, sodium borohydride, sugar, sodium citrate, citric acid, etc.²² The use of some of these compounds such as citric acid as a reducing agent is based upon the formation of a complex with metal salts. GO contains in its structure oxygenated functional groups that can form a complex with many metal salts. This complex facilitates the reduction of metal ions into metal nanoparticles as a result of the high density of electrons on the GO surface.

3.1 Surface, structural and spectral characterization

3.1.1 X-ray diffraction (XRD). X-ray diffraction is a very useful tool providing information on the structure of the crystalline material as well as the predominance of one phase over the other in case of the presence of different phases. XRD patterns of GO, RGO and Ru_{nan}-based RGO are shown in Fig. 1. As depicted in Fig. 1a, GO showed a strong diffraction peak at $2\theta = 8.78^\circ$ with a d -spacing of 10.05 Å, while RGO showed a broad hump at 26.5° corresponding to the (002) plane. On the other hand, Ru_{nan}-based RGO showed a weak diffraction peak at 26.6° with a d -spacing 3.34 Å related to the (002) plane, which gives an indication of the successful reduction of GO as depicted in Fig. 1b. The XRD pattern of Ru_{nan}-based RGO prepared in pH 8.0 showed the formation of Ru nanoparticles, as indicative by the diffraction peak at 43.97° corresponding to the (101) plane with a d -spacing of 2.058 Å. These values are in very good agreement with the reference card (04-003-0364). In the same time, RuO₂ is also formed as a secondary phase with the main diffraction pattern at 28.02° (d -spacing of 3.184 Å), according to reference card number (00-040-1290), with a relative intensity of 64.7%. This indicates that at pH 8 a mixture of Ru nanoparticles and RuO₂ was formed in 3 : 2 ratio.

It is worth mentioning that this is the first time Ru_{nan}-based RGO with Ru-rich composition was prepared without any reducing agent. The microwave irradiation provided enough energy to successfully reduce both GO and Ru³⁺. RuCl₃ dissolved in water forms hydroxyl-salts of Ru of the type [Ru(OH)_x, Cl_{3-x}] (see eqn (5)). In the presence of high concentrations of NH₄OH, [Ru(OH)_x, Cl_{3-x}] leads to the formation of Ru(OH)₃. Further oxidation results in the formation of the corresponding oxide according to eqn (5)–(9).¹³ The major fraction of ruthenium ions are bonded to the surface of graphene and consequently reduced on the RGO surface, forming Ru nanoparticles.



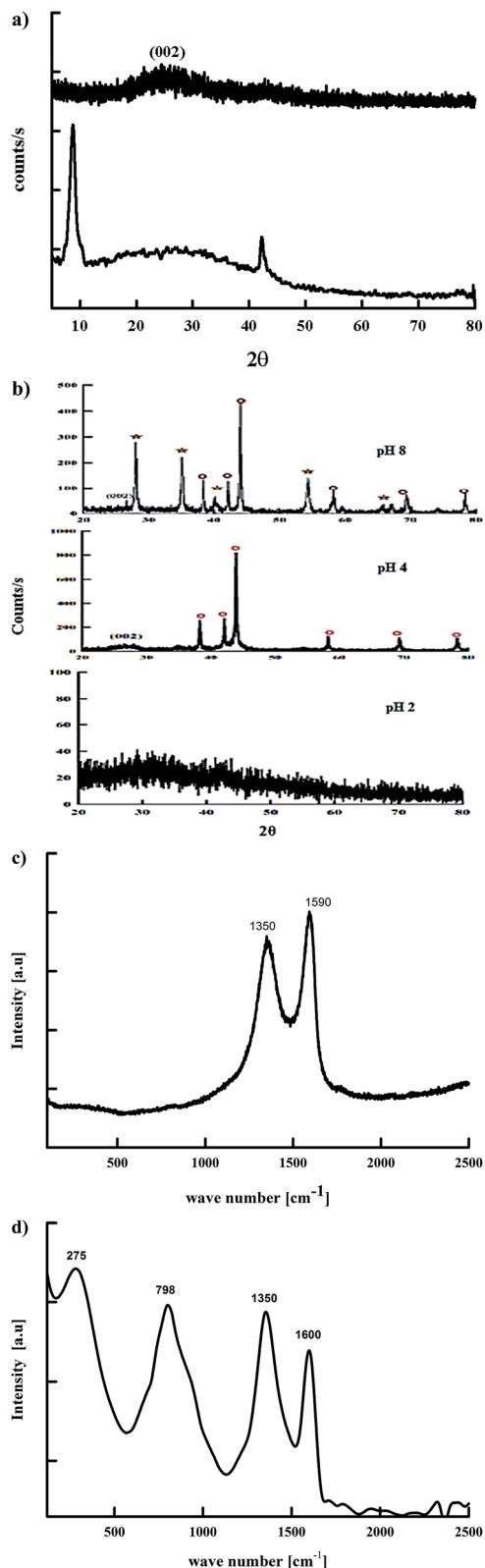
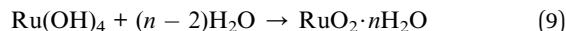


Fig. 1 XRD of (a) GO and RGO, (b) Ru_{nano} -based RGO (prepared at pH 2, 4 and 8) (circles are related to the phases of Ru and stars are related to RuO_2), (c) Raman spectra of GO and (d) Raman spectra of Ru_{nano} -based RGO (pH 8).

As the pH of the precursors strongly affects the prepared material, we also investigated the influence of pH on the initial solution of hydrated RuCl_3 and GO on the obtained composite structure. As shown in Fig. 1b the sample prepared at pH 4.0 contains solely Ru nanoparticles, where decorated RGO is the phase obtained without the formation of RuO_2 . Wang *et al.*⁹ reported that Ru nanoparticle formation is favored at pH values less than 4.9. At pH 4.0 Ru^{3+} ions are bound to GO sheets *via* complexation that helps reducing Ru^{3+} ions into Ru nanoparticles that are strongly attached to RGO sheets as a result of a high electron density on the GO surface. On the other hand, the sample prepared at pH 2.0 does not show any diffraction peaks, indicating an amorphous-like structure. This leads to the conclusion that pH 2.0 is not suitable for preparation of neither crystalline Ru nor RuO_2 . This is due to the strong acidic solution favoring “free” ionic formation and minimizing the possible reduction of Ru^{3+} ions to bind “*in situ*” onto RGO. However, the firing process that occurred during the microwave irradiation of the sample results in the reduction of GO into RGO as indicated by the absence of GO-related peaks. Hence, only samples prepared at pH 4.0 and 8.0 were electrochemically tested as supercapacitor electrodes.

3.1.2 Raman spectroscopy. Raman spectra of GO (see Fig. 1c) are characterized by two strong peaks at 1350 cm^{-1} corresponding to the D-band that arises from a breathing mode of K-point phonons of A_{1g} symmetry and the second peak at 1600 cm^{-1} , corresponding to the G-band that arises from the first scattering of E_{2g} phonon of sp^2 carbon atoms. These two bands are characteristic for RGO as well, while the difference in the spectra of GO and RGO is the intensity ratio of D-band to G-band. Increased D/G ratio indicates a successful reduction of GO into RGO. The spectra of Ru_{nano} -based RGO (pH 8) (Fig. 1d) shows the D-band of RGO at 1348 cm^{-1} and G-band at 1600 cm^{-1} , where the D/G ratio increases in case of Ru_{nano} -based RGO (pH 8) from 0.89 to 1.11 as indicator of the successful reduction of GO to RGO. The red shift of the D-band arises from the presence of Ru-based material as a spacer between RGO sheets. The Ru_{nano} -based RGO (pH 8) sample shows also two relatively broad peaks at 285 cm^{-1} and 798 cm^{-1} , corresponding to the symmetric A_{1g} vibration of $\text{Ru}-\text{O}$ and B_{2g} of RuO_2 , respectively.^{23,24}

3.1.3 X-ray photoelectron spectroscopy (XPS). The XPS analysis is a powerful tool to identify elements and the possible structure of a compound or a composite material. Fig. 2 shows the XPS spectra of Ru_{nano} -based RGO (pH = 8.0). The deconvoluted photoemission peaks show $\text{Ru}-3\text{p}_{1/2}$ and $3\text{p}_{3/2}$ at 484 eV and 461.8 eV, respectively, confirming the presence of Ru and RuO_2 nanoparticles in the composite.²⁵ Ru-3d peaks are observed at 284.9 and 279.9 eV, corresponding to $\text{Ru}-3\text{d}_{3/2}$ and $\text{Ru}-3\text{d}_{5/2}$ and related to RuO_x and Ru nanoparticles, respectively.^{26,27} It is worth to mention that the total content of RuO_2 and Ru in the sample is 5.2% (3.15% Ru nanoparticles and 2.08% RuO_2).

The de-convolution of C-1s showed four peaks at 287.6, 286.8, 285.7 and 284.6 eV that can be assigned to $\text{C}=\text{O}$, $\text{COO}/$

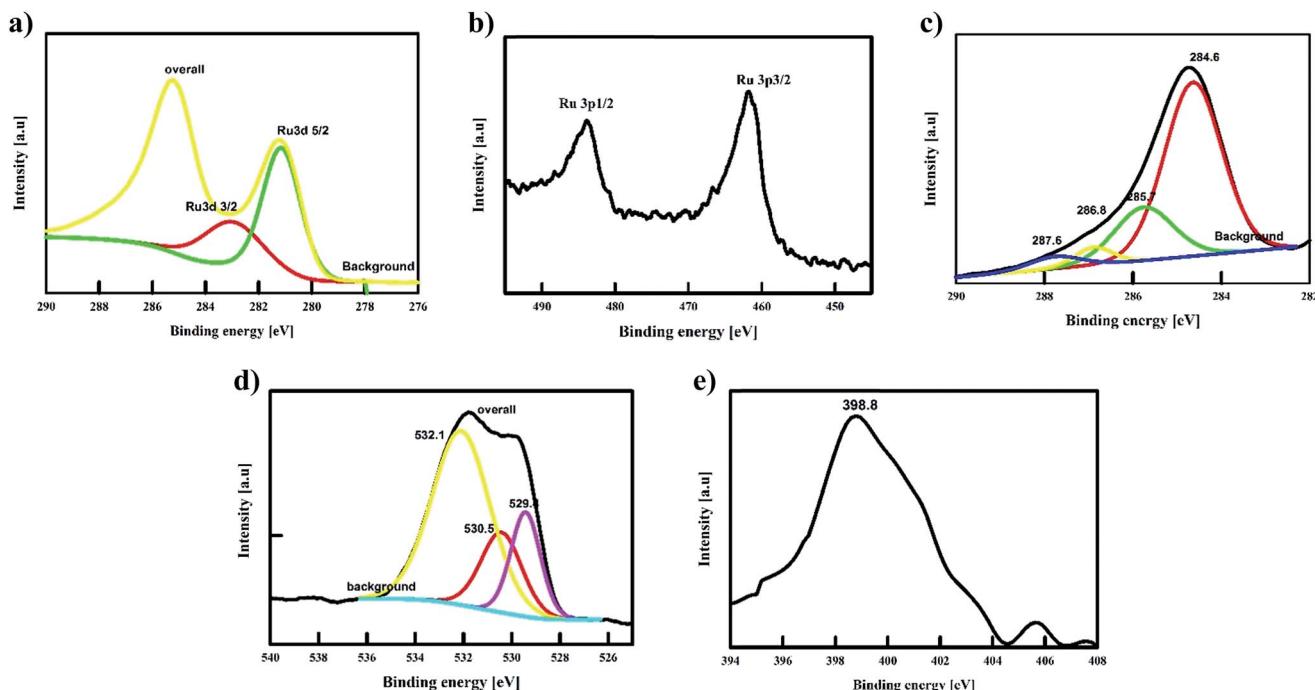


Fig. 2 XPS spectra of Ru_{nano}-based RGO (pH 8) shows the deconvoluted peaks of Ru 2p, Ru 3d, C 1s, O 1s and N 1s.

COH, C–O and C=C/C–H. The relative intensity of C–C/C=C compared to the oxygenated peaks as well as the corresponding peak area are used to evidence the successful reduction of GO into RGO. As depicted in the C-1s spectra of GO before its reaction (Fig. S1†), the C–C/C=C peak to the oxygenated peak ratio is 1 : 4. While in the case of Ru_{nano}-based RGO (pH = 8.0), C–C/C=C to the oxygenated peaks ratio is relatively higher compared to the case of GO as shown in Fig. 2c. In the latter case, the ratio is about 1.02 : 1.0 that confirms the successful reduction of GO into RGO. Moreover, the peaks of the O-1s spectra at 232.1 eV and 230.5 eV are assigned to the graphenic oxygen, while the peak at 229.4 eV corresponds to the oxygen of RuO₂. Additionally, the presence of nitrogen-doped graphene is confirmed by the N-1s peak at 398.78 eV that corresponds to the N–C bond, indicating the formation of N-doped graphene, which in turn plays a crucial role in separating graphene sheets and preventing their restacking.

3.1.4 TEM analysis. The TEM image of Ru_{nano}-based RGO (pH = 8.0) is shown in Fig. 3a, clearly indicating that the Ru-based materials (Ru and RuO₂) are uniformly distributed on the surface of RGO sheet. The presence of both Ru and RuO₂ has been confirmed from selected area electron diffraction (SEAD) pattern by calculating the *d*-spacing(s) and correlating them to those in the reference cards of both Ru and RuO₂. By the aid of the SEAD pattern we can calculate the average particle size of Ru and RuO₂ nanoparticles and differentiate between them. The SEAD pattern (Fig. 3c) of a selected area indicates that the particles of this area are mainly Ru with an averaged size of 1–3 nm. While by selecting another area (Fig. 3c), in which only RuO₂ has been investigated, and by the aid of the SEAD pattern (Fig. 3d), we found that some RuO₂ nanoparticles

form small clusters with a maximum size of 16 nm. The images further indicate that some Ru nanoparticles are embedded into the RGO sheets.

3.2 Electrochemical characterization

The electrochemical behavior of Ru-based graphene as supercapacitor electrode depends on the main phase within the structure.^{10,13–17} The main phase formed is RuO₂ with the possible formation of some Ru nanostructures as minor phase.^{10,13–17} As it was anticipated in this work, Ru_{nano} has the highest percentage in hybrid Ru_{nano}-based RGO when being prepared at pH 8.0 and is the only phase at pH 4.0.

Often H₂SO₄ and KOH are used as electrolytes in supercapacitors and pseudocapacitors measurements. Relatively high concentrations of these electrolytes are recommended in order to decrease the equivalent series resistance (ERS). While higher power densities are expected, the corrosive nature of the concentrated acid or base results in the deterioration of current collectors and their leakage will be detrimental, especially on decreasing the cycling stability of the capacitor. Limitation of the operating potential windows in acidic^{13,28} and basic^{11,29} electrolytes is another drawback.

RuO₂-based supercapacitors have been reported to operate in highly concentrated H₂SO₄^{30,31} and many researchers suggested that RuO₂-based supercapacitors are suitable only in such electrolytes.^{13,28} A high pseudocapacitance was obtained due to a surface reaction between Ru ions and protons (H⁺) in acidic media according to eqn (10):



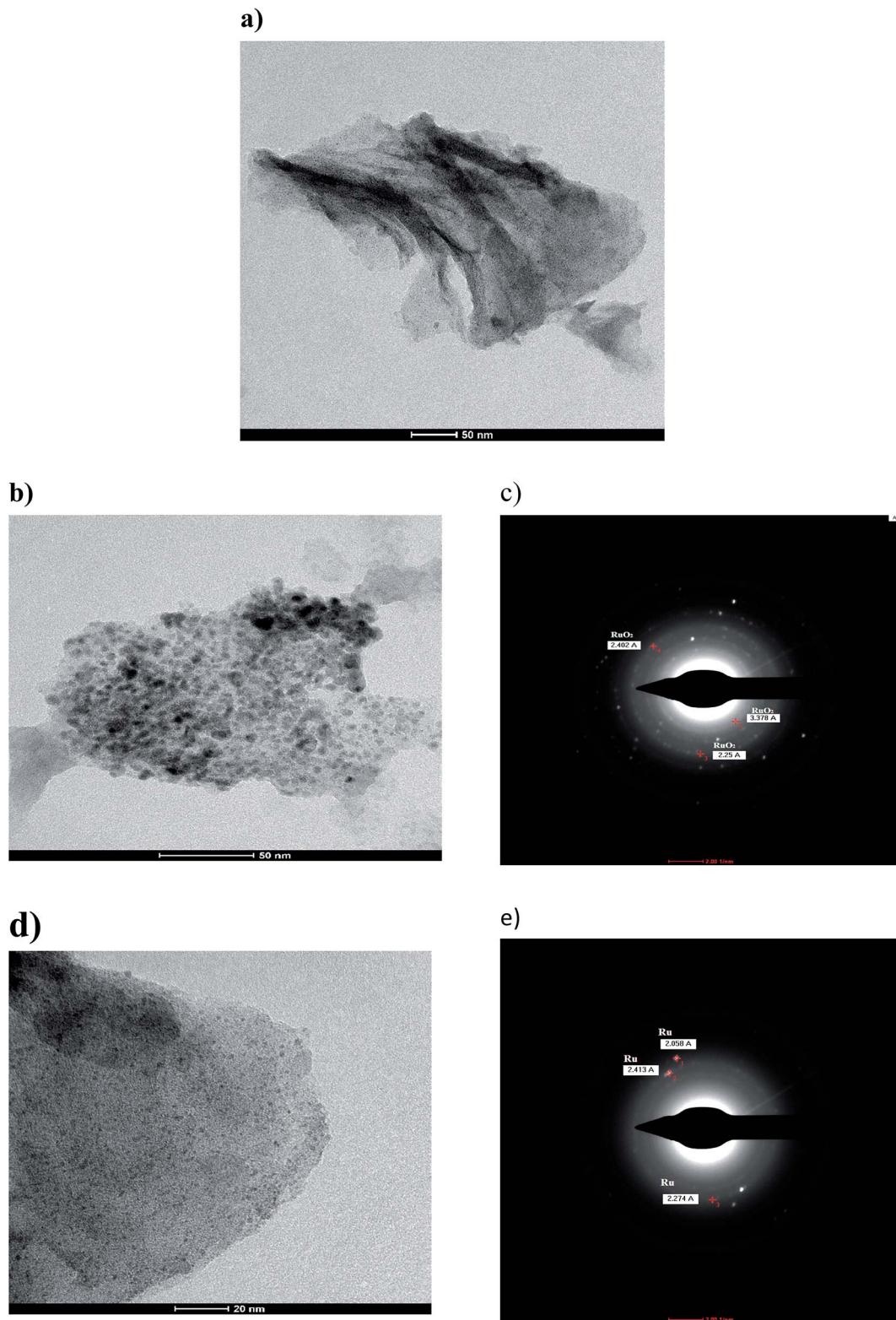


Fig. 3 (a) TEM of Ru_{nan}-based RGO (pH 8) (b) zoom in TEM of Ru_{nan}-based RGO (pH 8) showing Ru nanoparticles and (c) its corresponding SEAD pattern, (d) zoom in TEM of Ru_{nan}-based RGO (pH 8) showing RuO₂ nanoparticles and (e) its corresponding SEAD pattern.

Using a neutral electrolyte not only solves the corrosion problem but also allows the supercapacitor to operate at a wider potential window.³⁰

3.2.1 Cyclic voltammetry measurements. The electrochemical capacitance of the obtained Ru-based RGO (prepared at pH 4.0 and 8.0) as well as RGO has been tested by cyclic

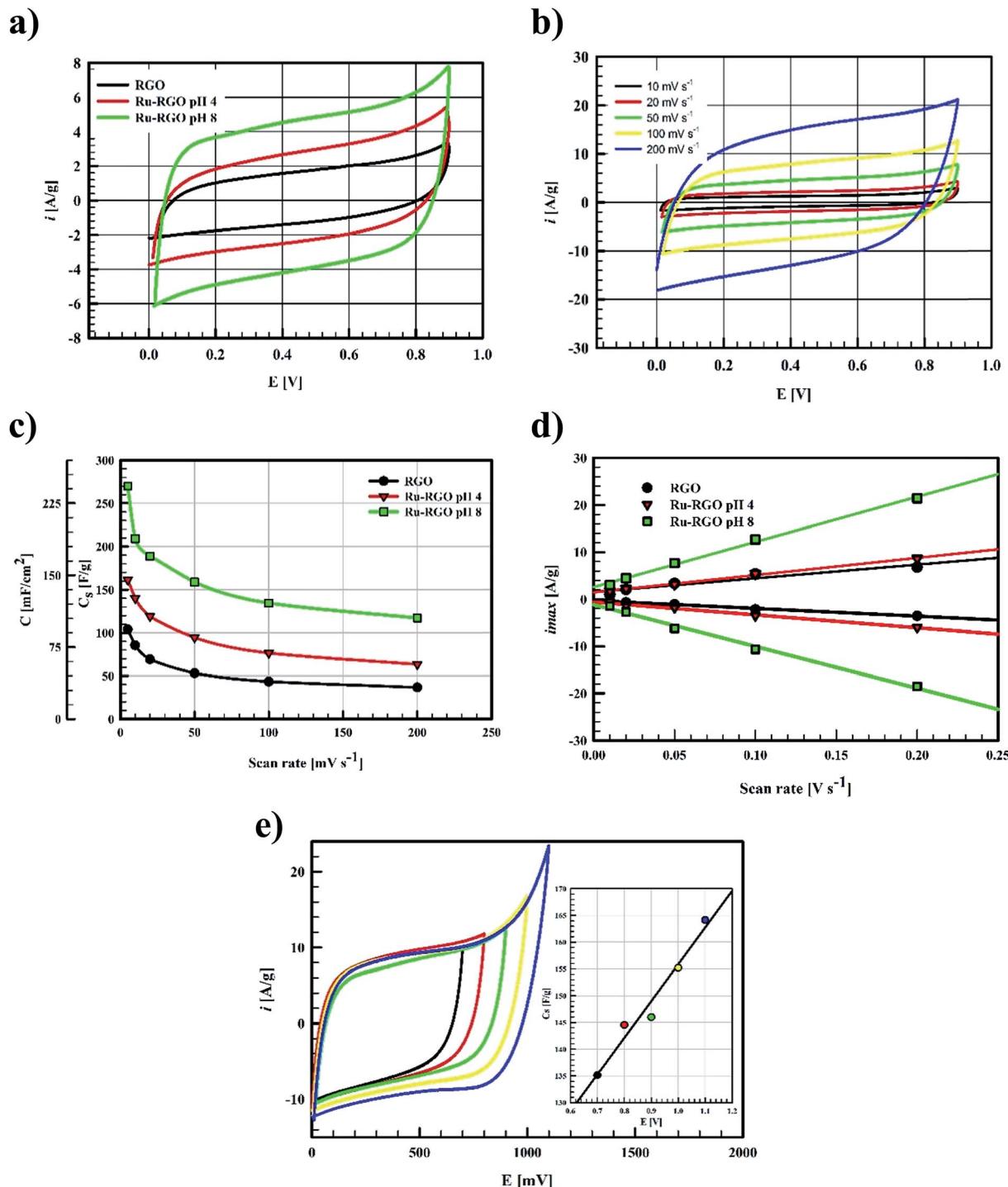


Fig. 4 (a) CV of RGO (black line), Ru_{nan}-based RGO (pH 4) (red line) and Ru_{nan}-based RGO (pH 8) (green line) measured at scan rate 50 mV s⁻¹ in 1 M NaNO₃, (b) CVs of Ru-RGO (pH 8) at various SR from 5 mV s⁻¹ to 200 mV s⁻¹, (c) effect of scan rate on specific capacitance of RGO (black line), Ru_{nan}-based RGO (pH 4) (red line) and Ru_{nan}-based RGO (pH 8) (green line), (d) the effect of scan rate on the maximum current density obtained at charging and discharging process and (e) CVs of Ru_{nan}-based RGO (pH 8) obtained at various upper potential and the relation between the upper potential and specific capacitance (inset).

voltammetry in 1.0 M NaNO₃ in the potential window from 0.0 to 0.9 V (see Fig. 4a). It is noticed that all the electrodes show a rectangular CV without appearing of any redox peaks indicating that the electric double layer capacitance shares the major contribution on the value of specific capacitance. Ru_{nan}-

based RGO (prepared at pH 8) provides the highest specific capacitance (C_{sp}), reaching 159 F g⁻¹ (140 mF cm⁻²) at a scan rate of 50 mV s⁻¹ and reaches 270 F g⁻¹ (238.5 mF cm⁻²) at a scan rate of 5 mV s⁻¹ compared to 161.5 F g⁻¹ (142.6 mF cm⁻²) and 104 F g⁻¹ (91.9 mF cm⁻²) at a scan rate of 5.0 mV s⁻¹ for

Ru_{nano} -based RGO (prepared at pH 4.0) and RGO, respectively. This indicates that inclusion of Ru nanoparticles acts as a spacer for RGO sheets and results in increasing the electrochemical capacitance of Ru_{nano} -based RGO (prepared at pH 4) over RGO. The presence of RuO_2 in Ru_{nano} -based RGO (prepared at pH 8.0) nearly doubles the value of the capacitance compared to Ru_{nano} -based RGO (prepared at pH 4.0) at 5.0 mV s^{-1} . In order to gain more information about the charge storage capabilities of the ruthenium-based RGO composites, cyclic voltammetry was carried out at various scan rates ranging from 5.0 to 200 mV s^{-1} as shown in Fig. 4b. It is clearly shown that as the scan rate increases the specific capacitance decreases as the diffusional component of the current starts to increase. The CVs still have a rectangular shape up to 100 mV s^{-1} and the distortion in the CV shape at 200 mV s^{-1} is very small, indicating a typical electric double layer capacitance (EDLC) behavior. Fig. 4c reveals the calculated specific capacitance under various scan rates of RGO, Ru_{nano} -based RGO (prepared at pH 4.0) and Ru_{nano} -based RGO (prepared at pH 8.0), where again Ru_{nano} -based RGO (pH 8.0) provides the highest specific capacitance. By looking at the relation between the maximum current obtained during charging and discharging (Fig. 4d), we can conclude that Ru_{nano} -based RGO (pH 8.0) shows very good symmetry compared to Ru_{nano} -based RGO (pH 4.0) and the calculated diffusion coefficient of Na^+ ions through RGO, Ru_{nano} -based RGO (prepared at pH 4.0) and Ru_{nano} -based RGO (prepared at pH 8.0) are 0.36×10^{-8} , 1.3×10^{-8} and $3 \times 10^{-8} \text{ cm}^2 \text{ s}^{-1}$, respectively. The effect of upper potential on the specific capacitance of Ru_{nano} -based RGO (pH 8.0) was investigated by CV and data are shown in Fig. 4e. The figure reveals that as the upper potential limit increases, the specific capacitance increases. This indicates the applicability of this supercapacitor to provide potentials up to 1.1 V in aqueous medium.

3.2.2 Galvanostatic charging/discharging measurements. Galvanostatic charging/discharging experiments were also conducted to evaluate the specific capacitance of Ru_{nano} -based RGO (prepared at pH 8.0) as shown in Fig. 5a. The specific capacitance of Ru_{nano} -based RGO (prepared at pH 8) at current density 2.4 A g^{-1} is 80 F g^{-1} (70.6 mF cm^{-2}) compared to 53.3 F g^{-1} (47 mF cm^{-2}) and 29.6 F g^{-1} (26.1 mF cm^{-2}) for Ru_{nano} -based RGO (prepared at pH 4.0) and RGO, respectively. It is also noticed that there is a slight change in the potential at the beginning of the discharge process, which reflects low ohmic resistance. The difference between the potential value at the onset of discharge, and the point at which a change in the slope of the discharge curve takes place is due to the *IR*-drop. As the current density increases, the specific capacitance decreases (see Fig. 5b); Ru_{nano} -based RGO (prepared at pH 8.0) shows a specific capacitance 40 F g^{-1} at high current density 20 A g^{-1} compared to 21.3 F g^{-1} and 13.3 F g^{-1} for Ru_{nano} -based RGO (prepared at pH 4.0) and RGO, respectively. Both Ru_{nano} -based RGO prepared at pH 8.0 and 4.0 show excellent cyclic stability even at a high current density of 24 A g^{-1} , as shown in Fig. 5c. However, Ru_{nano} -based RGO (prepared at pH 4) shows a continuous increase in the C_{sp} value upon cycling, reaching 158% capacitance retention after 5000 cycles. While, Ru_{nano} -based RGO (prepared at pH 8.0) showed an increase in the

specific capacitance in the first 2600 cycles to reach 108% of the initial specific capacitance then achieved specific capacitance retention 92.5% after successive cycling for 5000 cycles. It is worth to mention that upon storage the electrode in the electrolyte solution overnight after performing 5000 cycles, its performance enhanced again to reach 146% of the initial performance as shown in Fig. 5d. This enhanced performance was also stable for another 100 cycles as shown in the inset of Fig. 5d. Based on the aforementioned results, Ru_{nano} -based RGO (prepared at pH 8.0) has very good cycling stability with the ability of enhancing its initial performance upon its storage in the electrolyte solution.

Ragon plots in Fig. 5e are used to compare the performances of the current surfaces, showing the relation between power density in W kg^{-1} and the corresponding energy density in W h kg^{-1} . As shown in Fig. 5e, for a given power density the corresponding energy density is highest when using Ru_{nano} -based RGO (prepared at pH 8.0). This also shows an energy density of 7.7 W h kg^{-1} at 1766 W kg^{-1} , holding an energy density of 4.3 W h kg^{-1} at a higher power density (8624 W kg^{-1}). The energy density decreases at high power densities due to the fast voltage decay at high power densities.³¹

The current supercapacitor shows higher specific capacitance and cyclic stability compared to that reported by Antiohos *et al.*³² at 90% MnO-10% GO ($42.5 \pm 9.2 \text{ mF cm}^{-2}$) at scan rate 50 mV s^{-1} , in 1.0 M NaNO_3 . Li *et al.*⁶ reported 211.5 F g^{-1} at 2.0 mV s^{-1} in $1.0 \text{ M Na}_2\text{SO}_4$. Further, He *et al.* used Ru (20%)/mesoporous carbon and reported 269 F g^{-1} at 2.0 mV s^{-1} in 6.0 M KOH ,¹¹ while Xia *et al.* reported 238.6 F g^{-1} at scan rate 50 mV s^{-1} in 6.0 M KOH over RGO/CoMoO₄ material.³³ Lou *et al.*³⁴ synthesized Ru/porous activated carbon (derived from *Moringa Oleifera* fruit shells) composites using the bio-waste carbon source and ruthenium acetylacetone as the co-feeding metal precursors (Ru/MOC). Ru/MOC provided a specific capacitance of 291 F g^{-1} (the area of the electrode is 1.0 cm^2) at scan rate 10 mV s^{-1} in a potential range 0.0 to 800 mV in $1.0 \text{ M H}_2\text{SO}_4$ and showed 90% of its original C_s after only 2000 cycles. Finally, Takikawa and coworkers reported 153 F g^{-1} at 0.125 A g^{-1} using RuO_2 /activated carbon.¹² The values reported in this work for C_{sp} are comparable to those measured earlier,¹⁸ using RuO_2 (30.4%)/p-MWCNT in $1.0 \text{ M H}_2\text{SO}_4$ (231 F g^{-1} at a scan rate of 20 mV s^{-1}). Others reported the use of RGO/SnO₂-CNT, resulting in 218 F g^{-1} at scan rate 200 mV s^{-1} in 30 wt% KOH (81% of initial capacitance remains after 6000 cycles).³⁵ Although using RuO_2 -RGO as electrode material¹⁴ shows higher specific capacitance (that may be due to the use of $1.0 \text{ M H}_2\text{SO}_4$ as electrolyte), the electrode material used in the present study (Ru_{nano} -based RGO) shows higher cycling stability. And by taking into account the very small area of the used electrode (0.283 cm^2), its use in capacitor applications can be considered as a good candidate for a commercial, green and easily-prepared supercapacitor.

3.2.3 Electrochemical impedance spectroscopy. Nyquist plots of EIS measurements of RGO Ru_{nano} -based RGO (pH 4.0 and 8.0) in 1 M NaNO_3 are shown in Fig. 6a. RGO shows a semicircle in the high frequency regions and a vertical line (with the lowest slope) in the low frequency regions. The intersection of semicircle with x -axis at high frequency region is



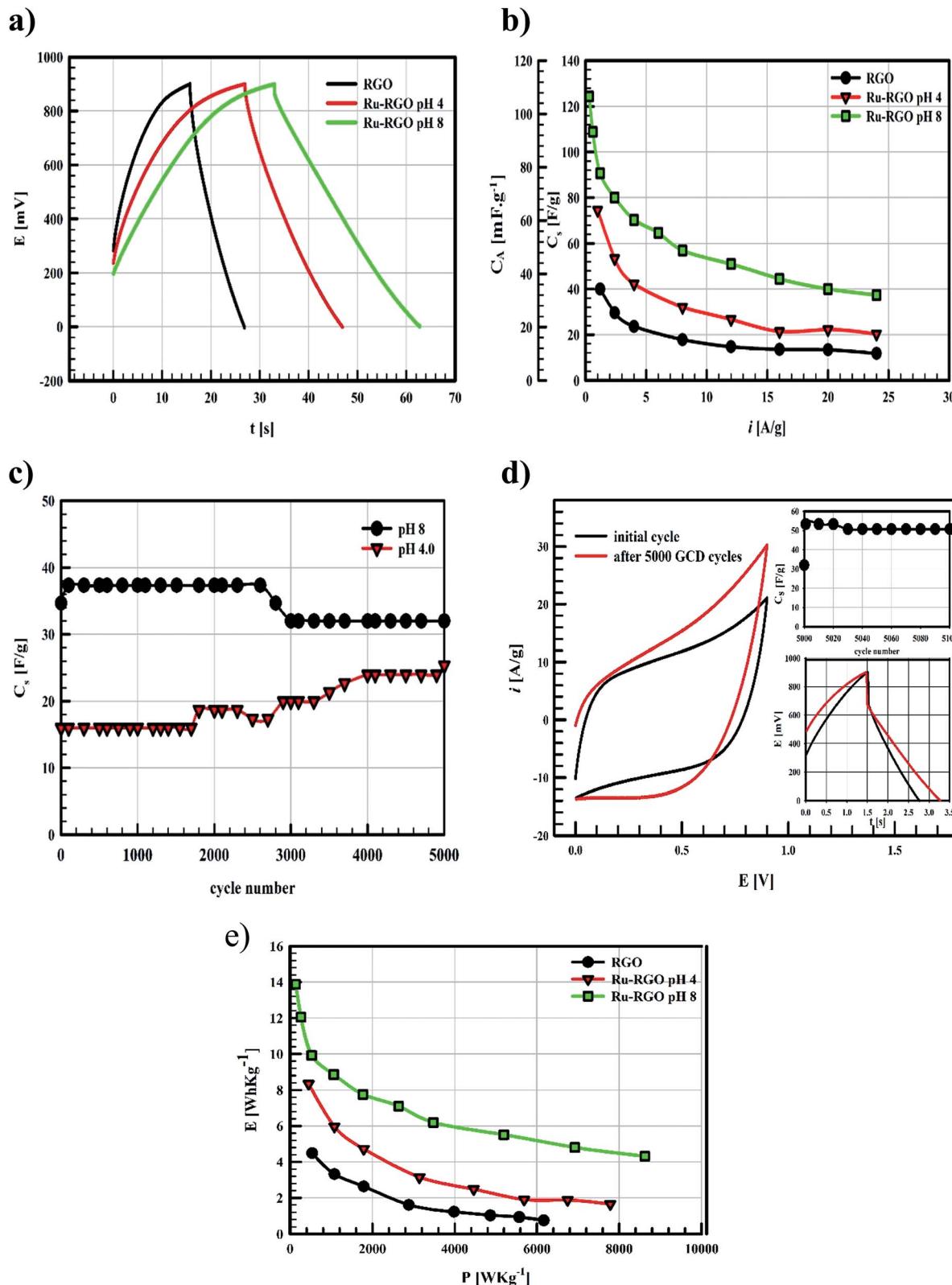


Fig. 5 (a) GCD curves of RGO (black line), Ru_{nanoo}-based RGO (pH 4) (red line) and Ru_{nanoo}-based RGO (pH 8) (green line) measured at current density 24 A g⁻¹ in 1.0 M NaNO₃, (b) the relation between the applied current density and the measured specific capacitance, (c) cycling stability of and Ru_{nanoo}-based RGO (pH 8 and 4) measured at current density 24 A g⁻¹ for 5000 cycles (d) the initially recorded CV compared to after performing 5000 GCD cycles at 24 A g⁻¹ followed by its storage overnight in the electrolyte solution. The inset is GCD at 24 A g⁻¹ and the measured specific capacitance recorded for another 100 cycles after the storage process. And (e) Ragone plot of RGO (black line), Ru_{nanoo}-based RGO (pH 4) (red line) and Ru_{nanoo}-based RGO (pH 8) (green line).



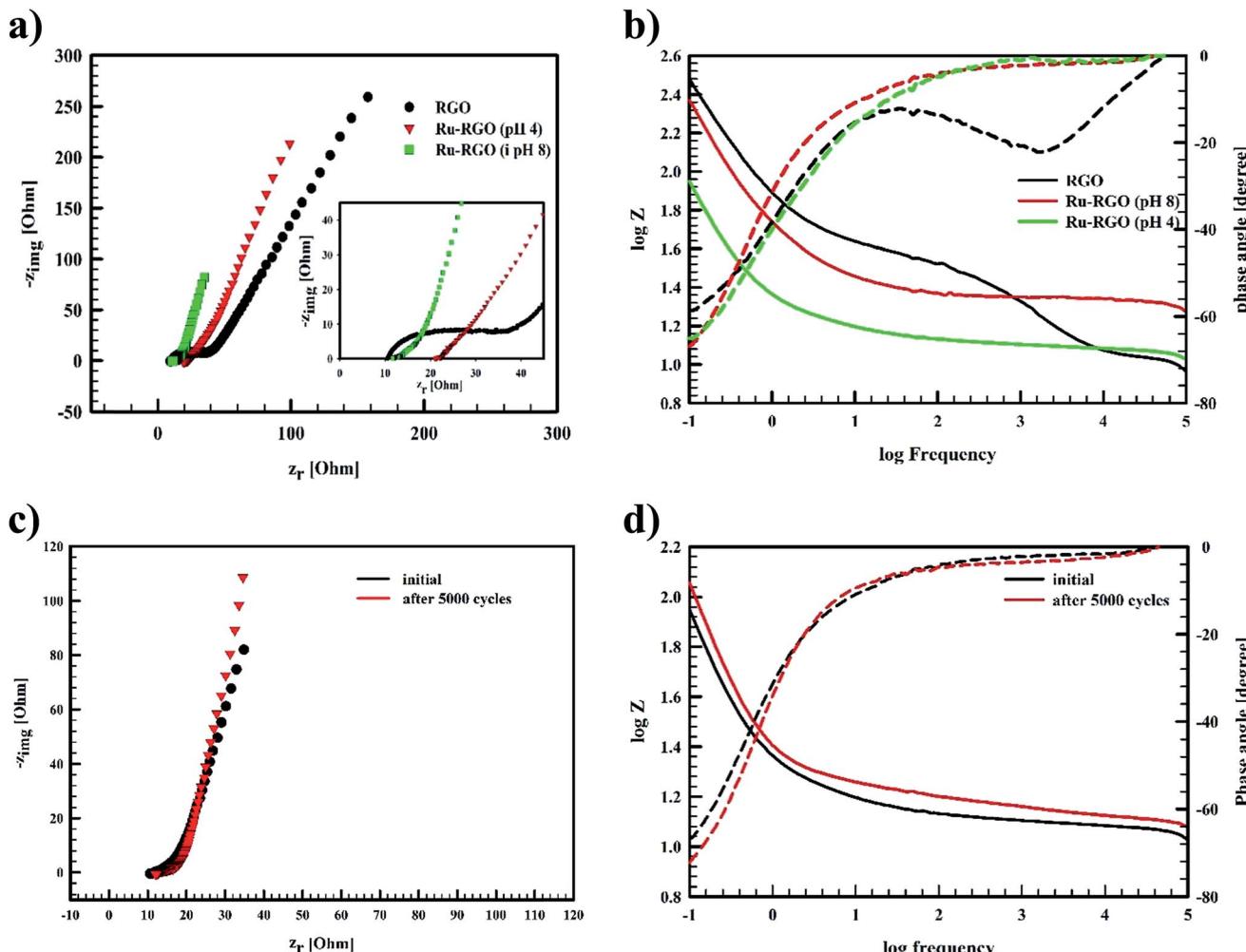


Fig. 6 (a) Nyquist plot of RGO (black line), Ru_{nan}-based RGO (pH 4) (red line) and Ru_{nan}-based RGO (pH 8) (green line), (b) Bode plot of RGO (black line), Ru_{nan}-based RGO (pH 4) (red line) and Ru_{nan}-based RGO (pH 8) (green line), (c) the initial Nyquist plot of Ru_{nan}-based RGO (pH 8) (black) and after 5000 cycles (red) at 24 A g⁻¹, and (d) the corresponding Bode plot.

the value of ESR and the diameter of the semicircle is the charge transfer resistance. Otherwise, both Ru_{nan}-based RGO prepared at pH 4.0 and 8.0 do not show a semicircle at high frequencies, rather a distorted semicircle with a very small diameter is formed that indicates very low charge transfer resistance compared to RGO. The vertical line in the low frequency region is related to the diffusion resistance of the electrolyte at the electrode interface.⁶ The supercapacitors approach ideal behavior as the line at low frequency region becomes more vertical with increasing slope. As shown in Fig. 6a Ru_{nan}-based RGO (prepared in pH 8.0) has a vertical line with relatively higher slopes compared to Ru_{nan}-based RGO (prepared at pH 4.0) and RGO, and tends to display better capacitive behavior. The calculated solution resistances for RGO, Ru_{nan}-based RGO (prepared at pH 4.0) and Ru_{nan}-based RGO (prepared at pH 8.0) are 10.4 Ω, 19.5 Ω and 10.6 Ω, respectively.

The Bode plot in Fig. 6b reveals that Ru_{nan}-based RGO (prepared at pH 8.0) has a phase angle φ of 70° at a frequency of 100 mHz, compared to 65° for Ru_{nan}-based RGO (prepared at pH 4) and 54° for RGO. The value of $\varphi = 70^\circ$ is close to the ideal capacitor at this frequency. Additionally, the time constants (τ)

for Ru_{nan}-based RGO (prepared at pH 8.0) can be calculated from the Bode plot, which is the time necessary to discharge the capacitor to 36.8% and it can be directly calculated by:

$$\tau = \frac{1}{2\pi f^*}, \quad (11)$$

where f^* is the characteristic frequency (frequency at phase angle 45°).

It is established that the lower the time constant, the faster the charge/discharge characteristics. The calculated value of τ for Ru_{nan}-based RGO (prepared at pH 8) is 0.357 s. This small value indicates fast charge/discharge properties of the current material. From the calculated time constant, the value of the maximum energy density can be deduced from the following relations:

$$\tau = \frac{E_{\text{max}}}{2P_{\text{max}}} \quad (12)$$

$$P_{\text{max}} = \frac{\Delta V^2}{4mR_{\text{ESR}}} \quad (13)$$

The calculated E_{\max} and P_{\max} of Ru_{nano} -based RGO (prepared at pH 8) are 15 W kg^{-1} and 76.4 kW kg^{-1} , respectively.

EIS of Ru_{nano} -based RGO (pH 8.0) was also measured after applying 5000 cycles at 24 A g^{-1} and the results were compared with the initial EIS measurement shown in Fig. 6c and d. We observe no significant change in the solution resistance after 5000 GCD cycles. However, the line at low frequency region became slightly more vertical after 5000 GCD cycles compared to the initial one. Moreover, there was no significant change in the phase angle as shown in the Bode plot (Fig. 6d).

4. Conclusion

Ru_{nano} -based material supported on RGO was successfully prepared by a one-pot, fast, green and pH-controlled synthesis method with the aid of microwave irradiation and in the absence of any other reducing or stabilizing agent. No calcination steps were necessary to obtain crystalline Ru-nanoparticles. The results show that Ru_{nano} -based RGO (prepared at pH 8.0) possesses the highest electrochemical capacitance compared to the one prepared at pH 4.0. Synthesis at pH 2.0 was not suitable to reduce Ru^{3+} ions, and consequently did not show a good capacitive behavior. All measurements were carried out in a neutral electrolyte (1.0 M NaNO_3). Ru_{nano} -based RGO (prepared at pH 8.0) showed a specific capacitance of 270 F g^{-1} at scan rate of 5.0 mV s^{-1} (238.5 mF cm^{-2} , surface area 0.283 cm^2). Ru_{nano} -RGO systems prepared at pH 8 reveal a very good cycle stability even after 5000 cycles, where it showed a C_{sp} retention 92.5% that increased to 146% when the electrode stored in the electrolyte solution overnight. Ru_{nano} -RGO (prepared at pH 8.0) showed a maximum power density of 76.4 kW kg^{-1} with a maximum energy density of 15.0 W h kg^{-1} .

Acknowledgements

The authors would like to acknowledge the partial financial support from Cairo University through the Vice President Office for Research Funds.

References

- Y. Cao, B. Lin, Y. Sun, H. Yang and X. Zhang, *J. Alloys Compd.*, 2015, **638**, 204–213.
- Z. B. Zhou, M. Benbouzid, J. F. Charpentier, F. Scuiller and T. H. Tang, *Renewable Sustainable Energy Rev.*, 2013, **18**, 390–400.
- J. W. Leea, J. M. Kob and J.-D. Kima, *Electrochim. Acta*, 2012, **85**, 459–466.
- P. Simon and Y. Gogotsi, *Nat. Mater.*, 2008, **7**, 845–854.
- N. Lin, J. Tian, Z. Shan, K. Chen and W. Liao, *Electrochim. Acta*, 2013, **99**, 219–224.
- Z. Li, J. Wanga, S. Liu, X. Liu and S. Yang, *J. Power Sources*, 2011, **196**, 8160–8165.
- C. G. Liu, Z. N. Yu, D. Neff, A. Zhamu and B. Z. Jang, *Nano Lett.*, 2010, **10**, 4863–4868.
- M. F. El-Kady, V. Strong, S. Dubin and R. B. Kaner, *Science*, 2012, **335**, 1326–1330.
- P. Wang, H. Liu, Y. Xu, Y. Chen, J. Yang and Q. Tan, *Electrochim. Acta*, 2016, **194**, 211–218.
- J. Y. Hwang, M. F. El-Kady, Y. Wang, L. Wang, Y. Shao, K. Marsh, J. M. Ko and R. B. Kaner, *Nano Energy*, 2015, **18**, 57–70.
- X. He, K. Xie, R. Li and M. Wu, *Mater. Lett.*, 2014, **115**, 96–99.
- X. J. He, Y. J. Geng, S. Oke, K. Higashi, M. Yamamoto and H. Takikawa, *Synth. Met.*, 2009, **159**, 7–12.
- X. Leng, J. Zou, X. Xiong and H. He, *J. Alloys Compd.*, 2015, **653**, 577–584.
- L. Na, J. Tian, Z. Shan, C. Kuan and W. Liao, *Electrochim. Acta*, 2013, **99**, 219–224.
- R. Rakhi, W. Chen, D. Cha and H. Alshareef, *J. Mater. Chem.*, 2011, **21**, 16197–16204.
- N. Soin, S. S. Roy, S. K. Mitra, T. Thundat and J. A. McLaughlin, *J. Mater. Chem.*, 2012, **22**, 14944–14950.
- L. Deng, J. Wang, G. Zhu, L. Kang, Z. Hao, Z. Lei, Z. Yang and Z.-H. Liu, *J. Power Sources*, 2014, **248**, 407–415.
- X. Liu, T. A. Huber, M. C. Kopac and P. G. Pickup, *Electrochim. Acta*, 2009, **54**, 7141–7147.
- N. I. Kovtyukhova, P. J. Ollivier, B. R. Martin, T. E. Mallouk, S. A. Chizhik, E. V. Buzaneva and A. D. Gorchinskiy, *Chem. Mater.*, 1999, **11**, 771–778.
- N. F. Atta, H. K. Hassan and A. Galal, *Anal. Bioanal. Chem.*, 2014, **406**, 6933–6942.
- M. J. Collins Jr, *Future Med. Chem.*, 2010, **2**, 151–155.
- Y.-J. Zhu and F. Chen, *Chem. Rev.*, 2014, **114**, 6462–6555.
- R. C. Rocha, M. G. Brown, C. H. Londergan, J. C. Salsman, C. P. Kubiak and A. P. Shreve, *J. Phys. Chem. A*, 2005, **109**, 9006–9012.
- L. M. M. Tejada, A. Muñoz, M. A. Centeno and J. A. Odriozola, *J. Raman Spectrosc.*, 2016, **47**, 189–197.
- R.-R. Bi, X.-L. Wu, F.-F. Cao, L.-Y. Jiang, Y.-G. Guo and L.-J. Wan, *J. Phys. Chem. C*, 2010, **114**, 2448–2451.
- J. Y. Shen, A. Adnot and S. Kaliaguine, *Appl. Surf. Sci.*, 1991, **51**, 47–60.
- J. C. Fuggle, T. E. Madey, M. Steinkilberg and D. Menzel, *Surf. Sci.*, 1975, **52**, 521–541.
- M. Y. Ho, P. S. Khiew, D. Isa, T. K. Tan, W. S. Chiu and C. H. Chia, *Nano*, 2014, **9**(6), 143002–1430027.
- K. Seung and Y. B.-H. Kim, *Electrochim. Acta*, 2015, **186**, 337–344.
- P. K. Shen, C.-Y. Wang, S. P. Jiang, X. Sun and J. Zhan, *Electrochemical Energy, advanced materials and technology*, CRC Press, 2016.
- P. Tamilarasan and S. Ramaprabhu, *Energy*, 2013, **51**, 374–381.
- D. Antiohos, K. Pingmuang, M. S. Romano, S. Beirne, T. Romeo, P. Aitchison, A. Minett, G. Wallace, S. Phanichphant and J. Chen, *Electrochim. Acta*, 2013, **101**, 99–108.
- X. Xia, W. Lei, Q. Hao, W. Wang and X. Wang, *Electrochim. Acta*, 2013, **99**, 253–261.
- B.-S. Lou, P. Veerakumar, S.-M. Chen, V. Veeramani, R. Madhu and S.-B. Liu, *Sci. Rep.*, 2016, **6**, 19949, DOI: 10.1038/srep19949. PMID: 26818461.
- R. B. Rakhi and H. N. Alshareef, *J. Power Sources*, 2011, **196**, 8858–8865.

