Fine silver sulfide–platinum nanocomposites supported on carbon substrates for the methanol oxidation reaction†

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Nanocomposites composed of silver sulfide (Ag₂S) and platinum (Pt) metal are promising for electrocatalysis. However, they suffer from low concentration synthesis or larger sizes due to limitations in the current synthetic approaches. Herein, we report an organic synthesis to address these deficiencies for the preparation of Ag₂S–Pt nanocomposites. This strategy starts with a dodecylamine (DDA)-based phase transfer of Ag⁺ ions from the aqueous phase to toluene. The transferred Ag⁺ ions are then loaded on the carbon substrates, followed by reaction with elemental sulfur to form carbon-supported Ag₂S nanocrystals, which are used as seeds for the growth of a Pt metal shell, resulting in the formation of the Ag₂S–Pt nanocomposites on the same carbon substrates. The as-prepared carbon-supported Ag₂S–Pt nanocomposites have a core–shell construction with an overall fine size of ca. 5 nm. In comparison with the commercial Pt/C catalysts from Johnson Matthey, fine Ag₂S–Pt nanocomposites supported on the carbon substrates exhibit a superior specific activity and durability for the methanol oxidation reaction under acidic conditions due to the strong electronic coupling effect between the Ag₂S and Pt domains.

1. Introduction

Nanocomposites composed of silver sulfide (Ag₂S) and platinum (Pt) have been demonstrated in our previous studies as potential candidates in electrocatalysis. The electron transfer from Ag₂S to Pt in the composite material due to the alignment of energy levels leads to a substantial increase in the electron density around the Pt domain, which inhibits the chemisorption of CO and promotes the methanol oxidation reaction (MOR). The selectivity of the Ag₂S–Pt nanocomposites for the MOR could even be used to construct a membraneless direct methanol fuel cell (DMFC) by integrating it with a cathode selective catalyst. The strategies used to prepare Ag₂S–Pt nanocomposites were usually carried out via a solution-based approach either by depositing Pt metal on the Ag₂S seeds or by a structural conversion process. Unfortunately, the shortcomings of the current synthetic routes are apparent. For the seed-mediated growth, BSPP (bis(p-sulfonatophenyl)phenyl phosphane) was used to direct the synthesis of the Ag₂S seeds in the aqueous phase, followed by the reduction of Pt precursors using sodium citrate to form Ag₂S–Pt heterogeneous nanocomposites. However, the expensive nature of BSPP and the aqueous phase entailed that the synthesis could only be performed at very low concentrations. For the structural conversion, core–shell Ag–Pt nanoparticles were used as precursors to react with the elemental sulfur for the production of nanocomposites consisting of Ag₂S and hollow Pt nanostuctures. However, the high content of Ag in the core–shell Ag–Pt precursors makes the formed Ag₂S–hollow Pt nanocomposites have a very low surface area exposed to the electrochemical reactions.

In this study, we report an organic synthesis to address the abovementioned deficiencies in the standard synthesis of Ag₂S–Pt nanocomposites. The novelty of this work lies in its synthetic approach, which is based on a combination of the phase transfer of Ag⁺ ions with the inhibition of Ag₂S growth on the carbon substrates to generate Ag₂S–Pt nanocomposites with fine sizes. This strategy starts with a dodecylamine (DDA)-based phase transfer of Ag⁺ ions from the aqueous phase to a non-polar organic solvent such as toluene. The transferred Ag⁺ ions are then loaded on the carbon substrates, followed by reaction with elemental sulfur to form carbon-supported Ag₂S nanocrystals (Ag₂S/C). Finally, the Ag₂S nanocrystals are used as seeds for the deposition of Pt metal, resulting in the formation of the Ag₂S–Pt nanocomposites on the same carbon substrates.

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of $\text{Ag}_2\text{S}$–Pt nanocomposites on the carbon substrates. The microscopic analyses indicate that the as-prepared carbon-supported $\text{Ag}_2\text{S}$–Pt nanocomposites ($\text{Ag}_2\text{S}$–PtNCs/C) are fine in size and have a core–shell construction. We demonstrated that the $\text{Ag}_2\text{S}$–PtNCs/C display a superior specific activity and durability in the MOR under acidic conditions. The overall fine sizes and thin Pt shell thickness, which are essential for a high specific surface area and for making the use of electronic coupling effect between the $\text{Ag}_2\text{S}$ and Pt domains, may account for the enhanced catalytic performance of the $\text{Ag}_2\text{S}$–PtNCs/C in the MOR. This concept may shed some light on the production of cost-effective and efficient electrocatalysts on a large scale.

2. Experimental

2.1 Reagents and material

Potassium tetrachloroplatinate(II) ($\text{K}_2\text{PtCl}_4$, 98%), silver nitrate ($\text{AgNO}_3$, 99%), acetic acid ($\text{CH}_3\text{COOH}$, 98%), aqueous $\text{HClO}_4$ solution (70%, ACS reagent), and Nafion 117 solution (5% in a mixture of lower aliphatic alcohols and water) from Aladdin Reagents; ethanol (99.5%), methanol (99%), and toluene (99.5%) from Beijing Chemical Works; elemental sulfur powder (S, chemical grade) from Xilong Chemical Co. Ltd.; oleylamine (70%, technical grade) from Sigma-Aldrich; dodecylamine (DDA, chemical grade) from Sinopharm Chemical Co. Ltd.; and Vulcan XC-72 carbon powder (XC-72C) with a BET surface area of 820 m$^2$/g and a diameter of ca. 250 nm from Cabot were used as received. All glassware and Teflon-coated magnetic stirring bars were cleaned with aqua regia, followed by copious rinsing with deionized water before drying in an oven.

2.2 Synthesis of the fine carbon-supported $\text{Ag}_2\text{S}$ nanocrystals ($\text{Ag}_2\text{S}/\text{C}$)

A DDA-based phase transfer method was used to synthesize the carbon-supported $\text{Ag}_2\text{S}$ nanocrystals with fine sizes. Typically, 50 mL of aqueous $\text{AgNO}_3$ solution (2 mM) was mixed with 50 mL of ethanol containing 1 mL of DDA for the formation of DDA–$\text{Ag}^+$ complexes. After 3 min of stirring, 50 mL of toluene was added and stirred for another 1 min. Subsequently, the upper toluene layer was collected, and a measured amount of XC-72C with a BET surface area of ca. 250 m$^2$/g and an average particle size of ca. 40 nm from Cabot, was used as received. All glassware and Teflon-coated magnetic stirring bars were cleaned with aqua regia, followed by copious rinsing with deionized water before drying in an oven.

2.3 Synthesis of fine carbon-supported $\text{Ag}_2\text{S}$–Pt nanocomposites ($\text{Ag}_2\text{S}$–PtNCs/C)

For the preparation of $\text{Ag}_2\text{S}$–PtNCs/C, the as-prepared $\text{Ag}_2\text{S}/\text{C}$ (10 wt% $\text{Ag}_2\text{S}$ on the carbon support) was redispersed in a 20 mL of oleylamine in a three-necked flask fitted with a condenser and a stirring bar. The solution was heated to 110 °C, and then 0.05 mmol of $\text{K}_2\text{PtCl}_4$ was added to the mixture under vigorous stirring. The resulting reaction system was kept at 110 °C for 2 h so as to allow the reduction of Pt$^{2+}$ ions by oleylamine on the surface of the fine $\text{Ag}_2\text{S}$ nanocrystals. The as-obtained $\text{Ag}_2\text{S}$–PtNCs/C were collected by centrifugation, washed thrice with methanol, and then dried in vacuum at room temperature for further characterizations and electrochemical measurements.
surface areas (ECSAs) of Pt. The catalyst performance in the MOR was also evaluated by cyclic voltammetry in an electrolyte (1 M methanol in 0.1 M perchloric acid) within a potential window of −0.2 V to 1 V at a scanning rate of 20 mV s⁻¹. The current densities for the tested catalysts were normalized by their ECSAs to obtain the specific activities.

3. Results and discussion

In this study, we aimed at making use of the substrates to inhibit the growth of Ag₂S nanocrystals for the preparation of Ag₂S–Pt nanocomposites with fine sizes. The overall strategy for the synthesis of Ag₂S–Pt nanocomposites with fine sizes on the surface of carbon substrates is summarized in Fig. 1. It was confirmed using ICP-AES (inductively coupled plasma atomic emission spectroscopy) analysis that the loading efficiency of Ag⁺ ions on the XC-72C carbon support from the organic medium (ca. 100%) is much higher than that from the aqueous phase (ca. 35%); therefore, before loading on the surface of the carbon substrates, Ag⁺ ions were first transferred to toluene via a process mediated by ethanol and dodecylamine. Then, upon mixing with elemental sulfur in toluene, Ag⁺ ions were converted into the Ag₂S nanocrystals on the surface of the carbon support. The as-prepared carbon-supported Ag₂S nanocrystals were subsequently used as seeds for the growth of a Pt metal shell in oleylamine, as indicated by the last step in Fig. 1.

The successful synthesis of Ag₂S nanocrystals on the surface of carbon substrates was confirmed by the powder X-ray diffraction (XRD) pattern, as shown in Fig. S1b in the ESI†. This XRD image for the Ag₂S/C sample with a Ag₂S/C mass ratio of 10% demonstrates that the as-prepared Ag₂S nanocrystals have a monoclinic phase (ESI Fig. S1a† for the Ag₂S reference with JCPDS card no. 140072). Fig. 2 shows the TEM images of the Ag₂S nanocrystals supported on carbon with a Ag₂S/C mass ratio of 20% (Fig. 2a), 10% (Fig. 2c), and 5% (Fig. 2e), which are spherical, monodispersed, and uniformly dispersed on the carbon substrates. The boundaries between the Ag₂S particles and the carbon substrates could be clearly discerned from the brightness contrast. The HRTEM images (inset in each TEM image) illustrate the lattice planes in these Ag₂S/C nanocrystals, confirming their high crystallinity. Moreover, as indicated by the corresponding particle size distributions (Fig. 2b, d and f), the average sizes for the carbon-supported Ag₂S nanocrystals are 4.85 nm with a standard deviation of 0.10 nm for the nanocrystals with a Ag₂S/C mass ratio of 20%, 4.03 nm with a standard deviation of 0.12 nm for the nanocrystals with a Ag₂S/C mass ratio of 10%, and 3.69 nm with a standard deviation of 0.14 nm for the nanocrystals with a Ag₂S/C mass ratio of 5%. As expected, the average diameter of the Ag₂S nanocrystals on the surface of carbon substrates is much smaller than that obtained in the aqueous phase by using BSPP–Ag⁺ complexes as precursors (7.2 nm) or that obtained from the thiol-stabilized Ag clusters in an organic medium (16 nm). We chose Ag₂S/C nanocrystals with a Ag₂S/C mass ratio of 10% as seeds for the preparation of carbon-supported Ag₂S–Pt nanocomposites (Ag₂S–PtNCS/C). As indicated in the TEM image (Fig. 3a) of the as-prepared Ag₂S–PtNCS/C, the average diameter of the Ag₂S nanocrystals on the surface of carbon substrates is increased from 4.03 nm to 5.04 nm after the deposition of Pt. This increase in the particle size might be attributable to the formation of the Pt shell with a thickness of 0.5 nm, which corresponds to 1–2 layers of Pt. Since the semiconductor nanocrystals have a catalytic capability for the reduction of metal ion precursors (in this case, Ag₂S and Pt ions), the isolated Pt deposition on the XC-72 carbon
substrate is less favourable. In the TEM images, the nanocomposites constructed with semiconductors and noble metals usually have apparent brightness contrasts between the semiconductor and metal domains due to the difference in the electron densities. In this study, although we did not succeed to observe the brightness contrasts between the Ag$_2$S and Pt domains in the Ag$_2$S–Pt nanocomposites due to their fine particle sizes and the influence of the carbon background, a successful synthesis of the Ag$_2$S–Pt nanocomposites as well as their structural features could be inferred by the HRTEM image and the EDX-based line scanning and mapping analyses in the STEM mode. Fig. 3b shows the HRTEM image of a single Ag$_2$S–Pt composite particle on the carbon substrates, in which the lattice difference between the peripheral and central regions indicates that the growth of Pt uniformly occurred on the Ag$_2$S seed, leading to the formation of Ag$_2$S–Pt nanocomposites with a core–shell construction. The distributions of Pt, Ag, and S elements in an arbitrary single particle, which were analyzed by EDX under the high-angle annular dark-field scanning TEM (HAADF-STEM) mode, were also used to confirm the core–shell structure of the Ag$_2$S–Pt nanocomposites. Nanoscale elemental mapping images (Fig. 3c–g) revealed that Ag and S in the composite particles are concentrated in the core region, whereas Pt is distributed throughout the entire particle, clearly suggesting the formation of a Ag$_2$S–Pt core–shell structure. Mapping analyses are fairly consistent with the element profiles of two arbitrary single Ag$_2$S–Pt particles (Fig. 3h and i), which also support that the Ag$_2$S–Pt nanocomposites formed on the carbon substrates by seed-mediated growth have a core–shell structure. The XRD pattern of the as-prepared Ag$_2$S–PtNCs/C was analyzed, as shown in Fig. S1c in the ESI.† Note that although the HRTEM image and STEM-based scanning and mapping analyses in Fig. 3 collectively illustrate the formation of core–shell Ag$_2$S–Pt nanocomposites on the surface of carbon substrates, the diffraction peaks from Pt metal are hardly detectable (as evinced by Fig. S1d in the ESI† for the Pt reference...
with JCPDS card no. 882343). The diffraction peaks in Fig. S1c (in the ESI†) also show the feature of a monoclinic Ag2S phase. Hence, it could be supposed that the interplanar spacing of Pt changes to match the lattice distances in Ag2S for its epitaxial growth on the surface of Ag2S seeds, analogous to that occurred on the carbon substrates, the partially nucleate on the existing Pt dots instead of continuous Pt shells are decorated on the multiple sites of the surface of the 7.2 nm Ag2S nanocrystals.\(^{\text{3}}\)

We examined the performance of Ag2S–Pt nanocomposites for catalyzing the methanol oxidation reaction (MOR) at room temperature and benchmarked it against that of the commercial Pt/C catalysts from Johnson Matthey (JM). As shown in Fig. 4a, cyclic voltammetry was used to determine the electrochemically active surface areas (ECSAs) of the Ag2S–Pt nanocomposites synthesized in the aqueous phase, in which the Pt dots instead of continuous Pt shells.

**Fig. 4a** Cyclic voltammograms of the Ag2S–Pt nanocomposites synthesized in the aqueous phase, in which the Pt dots instead of continuous Pt shells are decorated on the multiple sites of the surface of the 7.2 nm Ag2S nanocrystals.\(^{\text{2}}\)

The electronic coupling \(e_{f/b}\) between Pt and Ag2S core in the Ag2S–Pt nanocomposites is most likely due to the presence of residual impurities, e.g. residual DDA, amides originated from the oxidation of oleylamine ligands,\(^{25,26}\) and excess sulfur, adsorbed on the surface and their slightly larger sizes compared to those of Pt nanoparticles (ca. 3 nm) in the commercial catalysts.

**Fig. 4b** shows the cyclic voltammograms obtained for the MOR within a potential window of 0–1 V at a scan rate of 20 mV s\(^{-1}\), in which the current densities are normalized to the ECSA of Pt. As indicated, Ag2S–Pt nanocomposites display a superior activity in the MOR due to the electronic coupling effect between the Ag2S and Pt domains.\(^{2,6,7}\) In detail, the peak current density associated with the methanol oxidation in the forward scan, which is an important parameter while evaluating the catalytic activity,\(^{22,23}\) is 1.45 mA cm\(^{-2}\), which is much higher than that of the commercial Pt/C catalysts (0.99 mA cm\(^{-2}\)). It should be emphasized that the ECSA-based specific activities indicate the intrinsic catalytic performance of Pt in different chemical environments. If the calculations were based on the geometric area (0.196 cm\(^2\)) of the glassy carbon electrode, the current densities in the forward scan in the MOR for the Ag2S–Pt nanocomposites and commercial Pt/C catalysts are comparable, as shown in Fig. S4 in the ESI.\(^{\text{†}}\) Moreover, the onset potential in the MOR (crosspoint between the E-axis and the linear segment of the cyclic voltammograms) is ca. 0.29 V for Ag2S–Pt nanocomposites, which is lower than that of the commercial Pt/C catalyst (ca. 0.32 V), indicating that the oxidation of methanol is easier on Ag2S–Pt nanocomposites than that on the commercial Pt/C catalysts. In particular, the higher \(I_f/I_b\) (in which \(I_f\) and \(I_b\) are the forward and backward current densities, respectively) ratio as compared to that of the commercial Pt/C catalyst (1.45 vs. 0.61) indicates the higher durability of Ag2S–Pt nanocomposites for the methanol oxidation,\(^{24,25}\) which could be confirmed by the chronoamperometric (CA) tests. As shown in Fig. S2 in the ESI,\(^{\text{†}}\) the chronoamperograms show that the current density of Ag2S–Pt nanocomposites is higher than that of the commercial Pt/C for the entire time course, verifying that Ag2S–Pt nanocomposites exhibit a better electrocatalytic performance in the MOR.

The electronic coupling effect could be qualitatively illustrated by analyzing the binding energies of the Pt 4f region in the Ag2S–Pt nanocomposites and commercial Pt/C catalysts. As displayed in Fig. S3 in the ESI,\(^{\text{†}}\) while comparing with the binding energies of Pt 4f for commercial Pt/C catalysts, a shift to lower values is observed for Ag2S–Pt nanocomposites, indicating that there is an electron transfer or drift from the Ag2S core to the Pt shell in the nanocomposites. Analogous to the electron donation from Sn to Pt in the bimetallic Pt–Sn system and in other composite systems combining dielectric property with thermal conductivity,\(^{26–28}\) the electron transfer/drift from Ag2S to Pt can lead to a substantial increase in the electron density around the Pt atoms, which is favorable for weakening the chemisorption of CO-like intermediates, produced during methanol oxidation, on the Pt atoms, hence promoting the MOR by preventing the Pt surface from CO poisoning.

**4. Conclusions**

In summary, we demonstrated an organic-based synthesis of the carbon-supported Ag2S–Pt nanocomposites with fine sizes...
by combining the phase transfer of Ag⁺ ions with oleylamine reduction of the Pt metal precursors. The Ag₂S–Pt nanocomposites prepared on the surface of carbon substrates had a core–shell structure, with the thickness of the Pt shell being ca. 0.5 nm, corresponding to the 1–2 Pt atomic layers. The fine carbon-supported Ag₂S–Pt nanocomposites exhibited a superior activity and durability in the methanol oxidation reaction under acidic conditions due to the strong electronic coupling between the Ag₂S and Pt domains. By optimizing the overall particle size and the Pt shell thickness of the nanocomposites, we may further improve the catalytic activity of the carbon-supported Ag₂S–Pt nanocomposites in the methanol oxidation. In addition, the strategy used in this work might be extended for the generation of other semiconductor, e.g. CdS and PbS, or composite nanosystems, e.g. CdS–Pt and PbS–Pt, with fine sizes for electrocatalytic applications.

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Notes and references

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