


 Cite this: *RSC Adv.*, 2017, 7, 7413

# Th(IV) recovery from aqueous waste *via* hollow fiber renewal liquid membrane (HFRLM) in recycling mode: modelling and experimental validation

 Sareh Ammari Allahyari,<sup>a</sup> Seyed Javad Ahmadi,<sup>b</sup> Abdolhamid Minuchehr<sup>a</sup> and Amir Charkhi<sup>\*b</sup>

In the present work, the experimental investigation and mathematical modelling of Th(IV) recovery has been carried out by using the recycling mode of the hollow fiber renewal liquid membrane (HFRLM) process. Experiments were done at various operational conditions: the concentrations of TBP (tri-butyl phosphate) diluted in kerosene as a liquid membrane (LM), the nitrate and thorium(IV) ions concentration in the feed and flow rates of the feed and strip phases in the membrane module. The results indicate that practically more than 96% of thorium can be removed from an aqueous phase containing 198 ppm Th(IV) and 2 M NO<sub>3</sub><sup>−</sup> by using the LM comprising of 30% (v/v) of TBP while the feed and stripping solutions were being circulated at the flow rates of 5.6 mL min<sup>−1</sup> within 7 hours. In addition, a rigorous mathematical model has been developed to predict the experimental data. For this purpose, one-dimensional time dependent mass conservation equations in the longitude direction of the shell and lumen sides and in the radial direction of liquid membrane are numerically solved. The reasonable agreement between the experimental and the predicted data suggests the validity of the developed model. Therefore, this model might be useful for the design and optimization of the recovery of Th(IV) in the HFRLM process in the recycling mode.

 Received 7th November 2016  
Accepted 14th January 2017

DOI: 10.1039/c6ra26463h

[www.rsc.org/advances](http://www.rsc.org/advances)

## 1 Introduction

Thorium(IV) is a radiotoxic element of the actinide group, which due to its specific physiochemical properties has been utilized in a variety of industrial applications.<sup>1</sup> In addition, it has been considered as the most important replaced element for uranium in nuclear fuels.<sup>2</sup> These applications can produce many toxic and radioactive wastes. Finding a more efficient method for the separation of thorium from these wastes is desired due to prevention of its release to the environment and its recovery as a valuable element.<sup>3</sup>

Recently, liquid membrane (LM) separation technology has been used as a suitable replacement for the traditional methods of solvent extraction. This technique is a cost effective and energy efficient method *via* which a high pre-concentration factor in the separation of various elements can be obtained.<sup>4</sup> Furthermore, different configurations of membrane contactors have been utilized to maximize the efficiency of the process with a long-term stability. These include non-dispersive solvent extraction (NDSX),<sup>5</sup> hollow fiber supported liquid membrane (HFSLM),<sup>6</sup> pseudo-emulsion-based hollow-fiber strip dispersion

(PEHFS),<sup>7,8</sup> hollow fiber contained liquid membrane (HFCLM)<sup>9</sup> and hollow fiber renewal liquid membrane (HFRLM).<sup>10</sup> Among these configurations, the HFRLM has been preferred due to its several advantages over the other methods. Because of the renewal effect of the liquid membrane layer developed at the inner surface of the fibers and the high membrane surface area, the mass transfer rate is higher than the conventional methods in the hollow fibers. In addition, the continuous replenishment of the liquid membrane embedded in the pores of hollow fibers causes the long-time stability of the liquid membrane. Moreover, due to lateral shear forces, no secondary pollution could be created by the emulsification. Finally, the low investment costs due to the low energy consumption and low carrier consumption as well as the easy scale-up of the process and the wide range of operation region are some other advantages of this process.<sup>11</sup> As a result, the HFRLM process would have potential applications, which include the removal of the organic pollutants, the aroma recovery,<sup>12,13</sup> the pharmaceuticals extraction,<sup>14</sup> the recovery of metallic compounds from industrial wastewaters and hydro-metallurgical processes.<sup>15–19</sup>

In order to provide an appropriate tool to consider some engineering issues like the detailed design, the estimation of cost, the scale up and optimization of the HFRLM process, it is necessary to develop a simple and precise mathematical model. Bringas *et al.* reported an overview on the different approaches

<sup>a</sup>Nuclear Engineering Department, Shahid Beheshti University, P.O. Box: 1983963113, Tehran, Iran

<sup>b</sup>Material and Nuclear Fuel School, Nuclear Science and Technology Research Institute, Tehran, Iran. E-mail: [acharkhi@aeoi.org.ir](mailto:acharkhi@aeoi.org.ir); Tel: +98-21-82063626



of the mathematical modeling of hollow fiber liquid membrane contactors.<sup>20</sup> Generally, two different approaches have been considered in the research: the diffusional-kinetic and mixed-kinetic approaches. In the diffusional-kinetic approach, the instantaneous chemical reactions (complexation and de-complexation) have been considered, therefore the global rate of the process depends only on the diffusional mass transport. In the later approach, the global rate of the process depends on both the diffusional mass transport and the reaction kinetic. In addition, the mathematical modelling of the fluid flow in a hollow fiber contactor has been categorized as a rigorous or simplified model. The simplified model is generally used for quick calculations by assuming the average properties and the concentrations on each side of the membrane, and by writing the macroscopic balances of the mass, momentum and energy. In the rigorous model and in the most accurate form, the equations of change and the interfacial chemical reactions might be considered in the differential form, and solution of this set of equations reveals useful information about the velocity, the temperature, the pressure and the concentration profiles. Zhongqi *et al.* introduced a simplified mathematical modeling of copper(II) transport through the HFRLM, which was in good agreement with the experimental results.<sup>11</sup> They assumed that the fast complexation of solute with the carrier had been carried out at the aqueous-organic interfaces. At another work,<sup>21</sup> the authors investigated the effects of pH on the mass transfer of copper in the HFRLM system by incorporating counter transport of hydrogen ions as well as the chelating reaction in the individual mass transfer resistances. The individual mass transfer resistance of the strip phase was neglected. Yang and coworkers have also mathematically modeled the transfer of CuSO<sub>4</sub> in the HFRLM process in different operating conditions such as the single-pass, the recycling and the cascade modes. By considering the non-ideal flow in the shell side, a new mass transfer correlation was proposed according to Peclet number. The experimental results proved the reliability and accuracy of the model.<sup>22</sup>

In the present work, recovery of thorium in HFRLM process in the recycling mode is mathematically modelled based on the rigorous diffusional-kinetic model. Due to the comparatively high recovery of the small values of the solute, the recycling mode is preferable to the one-through mode. Since there has been few attempts to model the HFRLM process in the recycling mode, in this study a novel, simple and precise mathematical model has been developed based on the mass conservation in the longitude direction of the shell, the lumen sides, and radial direction of the liquid membrane phase as well. Unlike the previous works, variation of the renewal liquid membrane layer concentration at the inner surface of the lumen side has been considered for the first time in the modelling of the HFRLM.<sup>22</sup> The dependent PDE equations have been solved numerically and the results have been validated with the experimental results. In addition, the effect of carrier, the nitrate ion and the initial thorium concentrations and the flow rates of the feed and stripe phases on the recovery of Th(IV) has been mathematically investigated.

## 1.1 Mathematical model

In this study a rigorous model has been developed in the longitude direction of the shell and lumen sides as well as the radial direction of the liquid membrane phase by considering the following assumptions in the hollow fiber renewal liquid membrane in the recycling mode (Fig. 1).

- (1) The operation is carried out at constant pressure and temperature conditions, so the physiochemical properties of fluids have remained unchanged.
- (2) The momentum and energy balances have been neglected.
- (3) Due to the instantaneous or fast kinetic of chemical reactions (complexation and de complexation) of Th(IV) with the extractant and the stripping solution, the diffusional-kinetic regime has been considered as a transport mechanism of Th(IV).
- (4) The mass conservation of Th(IV) has been considered in a one-dimensional and time dependent mode.
- (5) Th(IV) concentration profile in the radial direction of the shell and lumen sides has been neglected.
- (6) The extraction of Th(IV) from nitrate media has been considered as an equilibrium reaction.
- (7) Fluid flows through the lumen side in a laminar and fully developed regime.
- (8) Axial dispersion has been neglected and only the axial convection has been considered.
- (9) Mass transfer in the shell and tube sides has been approximated by the film and the renewal theory, respectively.
- (10) The fibers have uniform pore size and porosity and they have been modeled as a thin cylindrical pipe with a uniform diameter.
- (11) Due to pre-mixing of aqueous and organic phases, which causes a saturated feed in HFRLM, the instability of the liquid membrane phase due to the solubility-related loss would be eliminated.
- (12) Due to the rigorous agitation, the concentrations of Th(IV) in the feed and strip reservoir tanks are uniform.
- (13) Due to the rigorous agitation and the fast extraction reaction, the aqueous phase of the reservoir tank is assumed to be in equilibrium with the organic droplets.
- (14) Due to the fast extraction reaction and formation of fine droplets in aqueous phase, the aqueous phase inside the lumen side has been considered in an equilibrium state with the organic droplets.

Th(IV) mass balances of the aqueous feed and strip solutions as well as the liquid membrane phase are given in eqn (1)–(3), respectively. The variation of Th(IV) concentration inside the liquid membrane droplets dispersed in the feed phase, has also been considered in the eqn (1) by using the term of  $\left(\frac{A/O+1}{A/O+m}\right)$ . The importance of this term has been neglected at the only model developed up to now for the HFRLM process in the recycling mode.<sup>22</sup>

Furthermore, the Th(IV) mass balances in the reservoir tanks are represented in eqn (4) and (5).

$$\frac{\partial C_f}{\partial t} = -\frac{Q_f}{nA_f} \frac{\partial C_f}{\partial z} - \left(\frac{A/O+1}{A/O+m}\right) \frac{\varepsilon P_f}{A_f} N_{A_f} \quad (1)$$



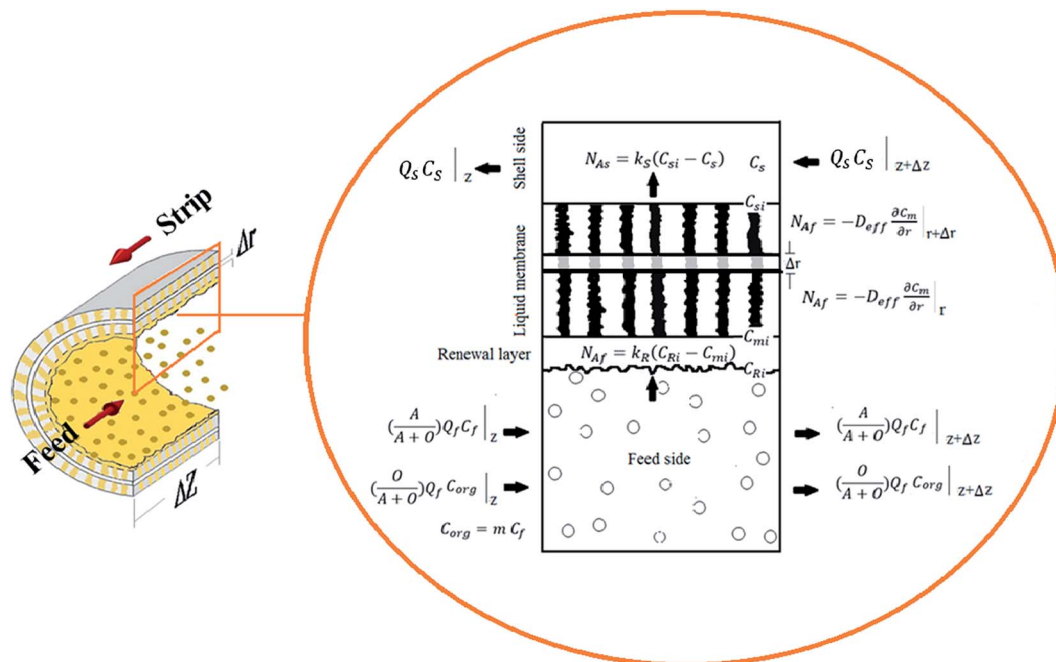


Fig. 1 Schematic diagram for determination of mass balance in single hollow fiber.

$$\frac{\partial C_s}{\partial t} = -\frac{Q_s}{A_s} \frac{\partial C_s}{\partial z} + n\epsilon \frac{P_s}{A_s} N_{A_s} \quad (2)$$

$$\frac{\partial C_m}{\partial t} = \frac{D_{eff}}{r} \frac{\partial}{\partial r} \left( r \frac{\partial C_m}{\partial r} \right) \quad (3)$$

$$V_{tf} \frac{dC_{ft}}{dt} = Q_f C_{fin} - Q_f C_{ft} \quad (4)$$

$$V_{ts} \frac{dC_{st}}{dt} = Q_s C_{sin} - Q_s C_{st} \quad (5)$$

where  $Q_s$ ,  $Q_f$ ,  $P_s$ ,  $P_f$ ,  $A_f$  and  $A_s$  are the volumetric flow rate in the shell and lumen sides, the wetted perimeter of each fiber in the tube and shell sides, the cross sectional area of the individual fiber and the strip phase passing through the shell side, respectively. Also  $n$ ,  $\epsilon$  and  $D_{eff}$  are the number and the porosity of the fibers and the effective diffusivity of Th(iv) in the liquid membrane phase, respectively. Furthermore,  $V_{tf}$ ,  $V_{ts}$ ,  $C_{ft}$  and  $C_{st}$  represent the volume and Th(iv) concentration of the feed and strip phases in the reservoir tanks, respectively. In eqn (1),  $A/O$  is aqueous to organic phase volume ratio. In addition,  $m$  is the distribution coefficient of the extraction process. Finally,  $C_s$ ,  $C_f$  and  $C_m$  are the concentration of Th(iv) in the strip, feed and liquid membrane phases, respectively.

In the above equations  $N_{A_f}$  and  $N_{A_s}$  are the feed and strip mass transfer fluxes of Th(iv). By considering the entry of the feed and strip phases counter-currently, the boundary and initial conditions of the eqn (1) and (2) are as the following:

$$t = 0, 0 < z < l, C_f = 0, C_s = 0 \quad (6)$$

$$t > 0, z = 0, C_f = C_{ft}, z = l, C_s = C_{st} \quad (7)$$

The initial and boundary conditions of eqn (3) are as the following:

$$t = 0, r_{in} < r < r_{out}, C_m = 0 \quad (8)$$

$$t > 0, r = r_{in}, N_{A_f} = -D_{eff} \frac{\partial C_m}{\partial z} \quad (9)$$

$$t > 0, r = r_{out}, N_{A_s} = -D_{eff} \frac{\partial C_m}{\partial z} \quad (10)$$

The initial conditions in the reservoir tanks are given as the following:

$$t = 0, C_{ft} = C_{ft0}, C_{st} = 0 \quad (11)$$

By considering the local mass transfer coefficients,  $N_{A_f}$  and  $N_{A_s}$  could be calculated as follows:

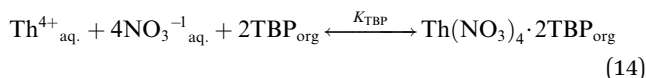
$$N_{A_f} = k_R (C_{Ri} - C_{mi}) \quad (12)$$

$$N_{A_s} = k_s (C_{si} - C_s) \quad (13)$$

where,  $k_s$  and  $k_R$  indicate the local mass transfer coefficients in the aqueous phase passing through the shell side and organic renewal layer, respectively. In addition,  $C_{Ri}$  and  $C_{mi}$  are the interfacial concentration at the inner surface of renewal layer and LM, respectively. Also,  $C_{si}$  is the interfacial concentration of Th(vi) at the interface surface of LM and strip phase, and  $C_{si}$  and  $C_{Ri}$  are assumed to be in the equilibrium with the liquid membrane and the feed phases at interface surface, respectively. These equilibrium concentrations have been calculated according to the distribution coefficient.



Thorium is extracted from the nitrate media by using TBP according to the following equilibrium reaction.<sup>23</sup>



where TBP and  $\text{Th}(\text{NO}_3)_4 \cdot 2\text{TBP}$  denote the free carrier molecule and the metal-carrier complex, respectively.  $\text{NO}_3^{-1}$  is the counter-ion, which transports from the feed to strip solution along with  $\text{Th}(\text{IV})$ .

The equilibrium constant ( $K_{\text{eq}}$ ) of the above reaction is expressed as:

$$K_{\text{eq}} = \frac{[\text{Th}(\text{NO}_3)_4 \cdot 2\text{TBP}]_{\text{org}}}{[\text{Th}^{4+}]_{\text{aq.}} [\text{NO}_3^{-1}]_{\text{aq.}}^4 [\text{TBP}]_{\text{org}}^2} \quad (15)$$

where  $[\text{Th}(\text{NO}_3)_4 \cdot 2\text{TBP}]_{\text{org}}$ ,  $[\text{TBP}]_{\text{org}}$ ,  $[\text{NO}_3^{-1}]_{\text{aq.}}$  indicate the solute-carrier complex at the feed-membrane interface and the free carrier and nitrate ion concentrations, respectively.

By considering  $m$  as the distribution coefficient of the extraction process, eqn (15) could be written as:

$$m = K_{\text{eq}} [\text{NO}_3^{-1}]_{\text{aq.}}^4 [\text{TBP}]_{\text{org}}^2 \quad (16)$$

The distribution coefficient of the back extraction process is the same as the above equation by using the nitrate concentration of strip phase. The equilibrium constant of eqn (14) is 0.6.<sup>24</sup> In our previous work,<sup>25</sup> the parameters of empirical correlations for  $k_{\text{R}}$  and  $k_{\text{s}}$ , mentioned in eqn (12) and (13), were modified and have been used in this study. In addition  $D_{\text{eff}}$  has been considered as equal to  $9.2 \times 10^{-11} \text{ m}^2 \text{ s}^{-1}$  due to the following equation:

$$D_{\text{eff}} = \frac{D_0 \varepsilon}{\tau} \quad (17)$$

where  $D_0$  is the diffusivity of thorium through the liquid membrane phase, which has been calculated by Wilke–Chang theory.<sup>26</sup>

The set of partial differential equations given in eqn (1)–(3) is converted to the coupled ordinary differential equations by spatially discretizing the equations.

Then, the obtained set of initial value ODEs is solved by using the 4-th order Runge–Kutta method.

The recovery efficiency  $R$  (%) of the system has been calculated by eqn (18).

$$R(\%) = \frac{C_{\text{st}}}{C_{\text{ft0}}} \times 100 \quad (18)$$

The average relative error (ARE), as given in the following equation, has been used in all the experiments to estimate the accuracy of the model.

$$\text{ARE}\% = \frac{1}{n} \sum_{j=1}^n \frac{|C_{\text{t}}^{\text{exp}} - C_{\text{t}}^{\text{cal}}|}{C_{\text{t}}^{\text{exp}}} \quad (19)$$

where,  $n$  is the number of the experimental data in various operating conditions and  $C_{\text{t}}^{\text{cal}}$  and  $C_{\text{t}}^{\text{exp}}$  are the calculated and experimental tank concentration, respectively.

## 2 Material and methods

### 2.1 Reagents

The aqueous phases have been prepared by using analytical grade reagents. The donor (feed) solution containing 99.16 to 396.65  $\text{mg L}^{-1}$  of  $\text{Th}(\text{IV})$  and 2  $\text{mol L}^{-1}$  of  $\text{NaNO}_3$  was prepared by dissolving appropriate amounts of thorium nitrate and sodium nitrate salts, which were provided by Merck company, in deionized water. The pH of the donor phase has been adjusted to 2 by adding an appropriate amount of diluted  $\text{HNO}_3$  and  $\text{NaOH}$  solutions and has been measured by a digital pH meter (model Metrohm 691). A 0.1  $\text{mol L}^{-1}$   $\text{HNO}_3$  aqueous solution is selected as an acceptor (stripping) phase.

The organic liquid membrane phase has been prepared by dissolving tributylphosphate (TBP) ( $\geq 99\%$ ) as a carrier in kerosene by the volume ratios of 15–45%. The commercial extractant and solvent of TBP and kerosene were purchased from Merck (Germany) and Sigma-Aldrich company (Germany), respectively.

### 2.2 Hollow fiber renewal liquid membrane module (HFRLM)

A hollow fiber module containing 50 polypropylene fibers, which were supplied from the Persian Polymer Company, fabricated and prepared for laboratory scale. The main characteristics of the fabricated module are given in Table 1. Furthermore, characterization of utilized hollow fibers by taking SEM images as well as measuring the porosity of membrane has been carried out.

The SEM images of the surface and cross section of hollow fibers were captured with a Scanning Electron Microscope (SEM: model of ZEISS EVO-18). The samples were broken in liquid nitrogen.

Due to measure the bulk porosity of the fibers, the sample hollow fibers were first weighed ( $w_1$ ). Then the weighted fibers were immersed in kerosene solution for 72 hours. By assumption of filling all the empty voids with the liquid membrane phase, the fully impregnated fibers were removed from the

Table 1 Characteristics of hollow fiber module

Shell characteristics	
Material	Glass
Length, (cm)	35
Internal diameter, $2R_i$ (cm)	0.6
Outer diameter (cm)	0.9
Fiber characteristics	
Material	PP (polypropylene)
Number of fibers in module	50
Effective length, $l$ (cm)	30
Internal diameter, $d_{\text{t}}^{\text{int}}$ ( $\mu\text{m}$ )	275
External diameter, $d_{\text{t}}^{\text{ext}}$ ( $\mu\text{m}$ )	375
Effective surface area of module, $A_{\text{eff}}$ ( $\text{m}^2$ )	$6.9 \times 10^{-3}$
Membrane porosity, $\varepsilon$	0.4
Membrane tortuosity, $\tau$	2





kerosene. Any excess kerosene on the outer surface of the fibers was wiped away. Next, the fibers were quickly weighed again. Therefore, the volume of empty void ( $V$ ) of hollow fiber wall can be calculated by the following relation:

$$V = \frac{1}{n} \left( \frac{w_2 - w_1}{\rho_K} \right) - \frac{\pi}{4} D_{in}^2 L \quad (20)$$

where,  $\rho_K$ ,  $D_{in}$ ,  $n$  and  $L$  are the kerosene density, inner diameter, number and length of fibers. The porosity of the hollow fibers membrane ( $\epsilon$ ) was estimated by the ratio of empty void ( $V$ ) to the total volume of the hollow fiber wall  $\left( \frac{\pi}{4} (D_{out}^2 - D_{in}^2) L \right)$ . An average value of the porosity which measured for three sets of fibers is about 40%.

In addition, a schema of HFRLM system in the recycling and counter current mode is represented in Fig. 2.

Thorium concentration in the aqueous phase is measured by inductively coupled plasma optical emission spectrometry (ICP-OES: model Liberty 220 Varian) analyzer.

Impregnating the pores of fibers with the liquid membrane (TBP diluted in kerosene) is carried out by passing the LM phase through the fibers in the flow rate of  $2.3 \text{ mL min}^{-1}$  for 30 min. After that, the stirred mixture of LM and the aqueous donor phase at the volume ratio of the aqueous to organic phase ( $A/O$ ) of 20 : 1 is pumped through the lumen side of the module while the aqueous acceptor phase (stripping solution) flows through the shell side in counter current mode. The outlet streams of both sides of the module are recycled back to the respective reservoir tanks with volumes of 250 mL. The flow rates of lumen and shell sides are carefully controlled in the range of  $2.3\text{--}10.6$  and  $5.6\text{--}10.6 \text{ mL min}^{-1}$ , respectively. In order to avoid mixing of the phases in the module, it is necessary to perform a positive pressure on the shell side. The pressure difference between the shell and lumen side is maintained at 0.3 bar. Samples of approximately 1 mL are taken from of the feed and strip

reservoir tanks at regular time intervals and  $\text{Th(IV)}$  concentration in the aqueous phases is measured by using ICP-OES.

### 3 Results and discussion

#### 3.1 SEM analysis of fibers

Scanning electron micrographs of hollow fiber are shown in Fig. 3(a)–(d). From the micrograph of Fig. 3(a), the outer and inner diameter of the fibers are measured about 375 and 275  $\mu\text{m}$ , respectively which is properly in agreement with the data of manufacture fibers (Table 1). From the micrograph of Fig. 3(b), the dimension of the pores is measured less than 200 nm. Fig. 3(c) and (d) illustrate a closer view of cross section of fibers which indicate the porosity of the utilized hollow fiber.

#### 3.2 Concentration distribution of thorium in the hollow fiber module

For a typical experiment run, the concentration distribution of thorium in the lumen, the liquid membrane and the shell sides of the module at different time intervals is illustrated in Fig. 4. The stirred mixture of the donor and the LM phase enters from one side of the module ( $z = 0$ ), whereas the strip solution enters from the opposite side ( $z = 1$ ). As the feed passes through the fibers, due to the transport of  $\text{Th(IV)}$  into the liquid membrane, the  $\text{Th(IV)}$  concentration is reduced along the fibers. At the shell side of the module, the back extraction of thorium occurs and consequently the thorium concentration at the outlet stream is higher than inlet stream. As shown in Fig. 4, by the time, the axial and radial  $\text{Th(IV)}$  concentration gradients in various parts of module (shell, HF, LM) are reduced.

Mass transfer rate of  $\text{Th(IV)}$  from the tube side to the shell side is shown in the Fig. 5 at various times. The results indicate that the rates are equal to each other during considered time

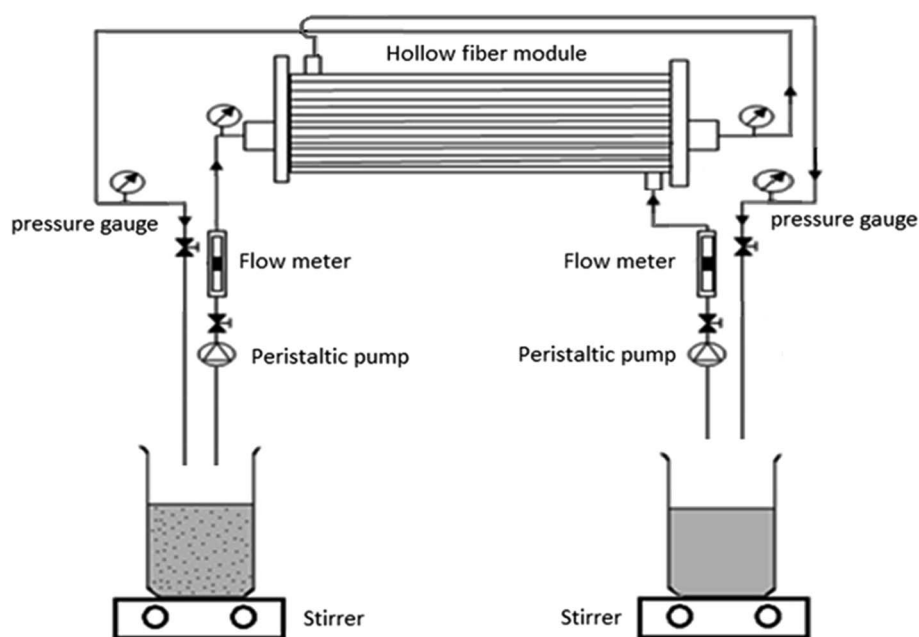


Fig. 2 A schema of HFRLM process in recycling and counter current mode.



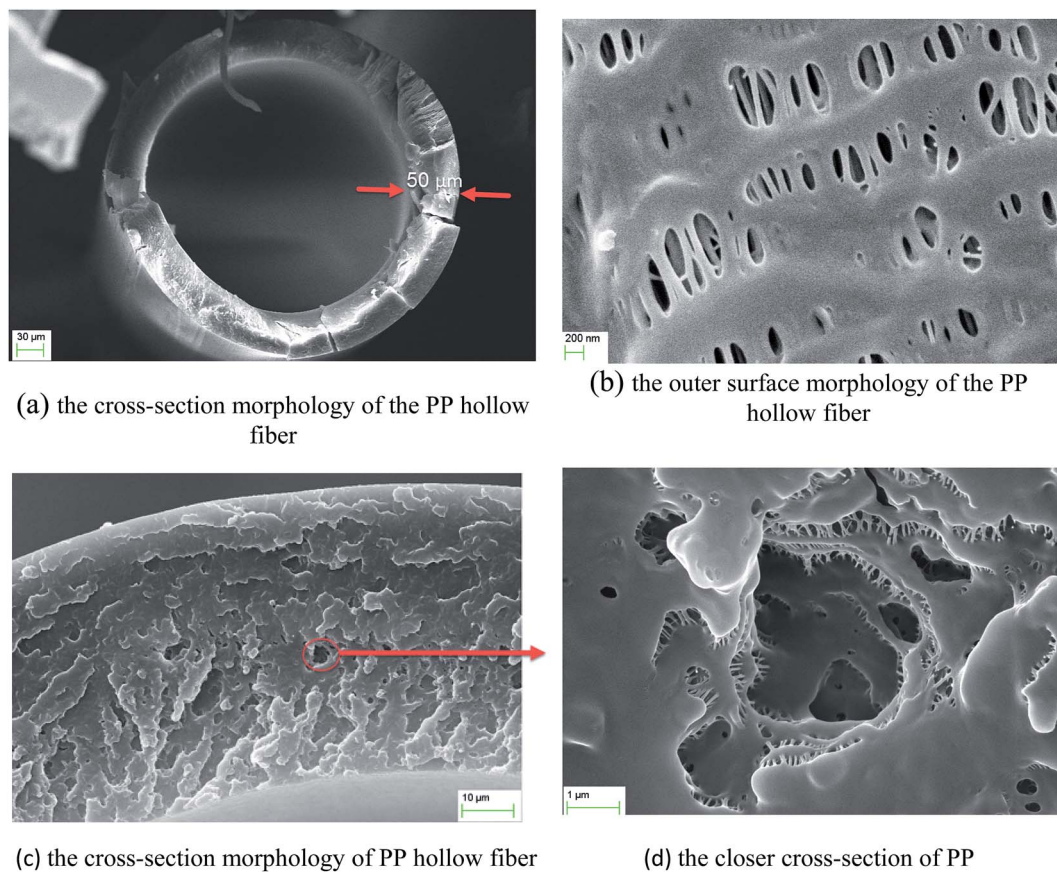


Fig. 3 SEM images of structure of PP hollow fiber.

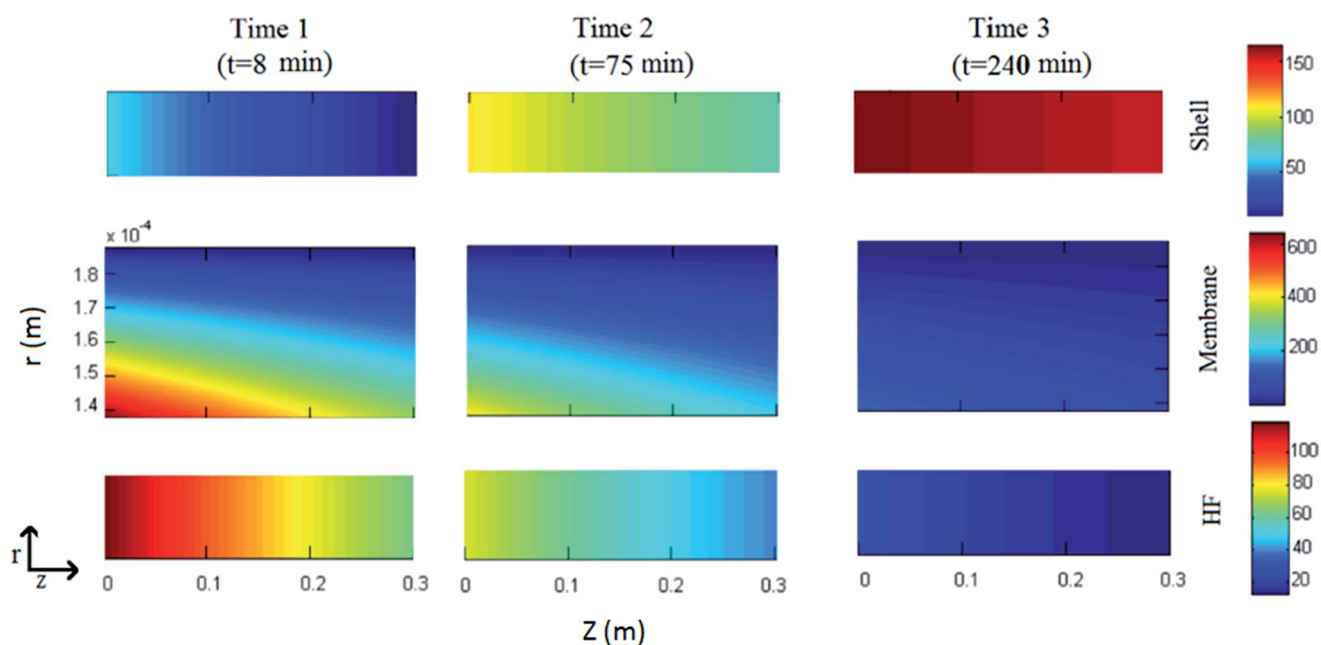


Fig. 4 Unsteady state concentration distribution of thorium in the module (feed phase:  $[Th^{4+}]_{init} = 198.32 \text{ mg L}^{-1}$ ,  $[NaNO_3] = 2 \text{ mol L}^{-1}$ , flow rate of tube and shell sides:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[HNO_3] = 0.1 \text{ mol L}^{-1}$ ,  $[TBP] = 30\% \text{ (v/v)}$ ).



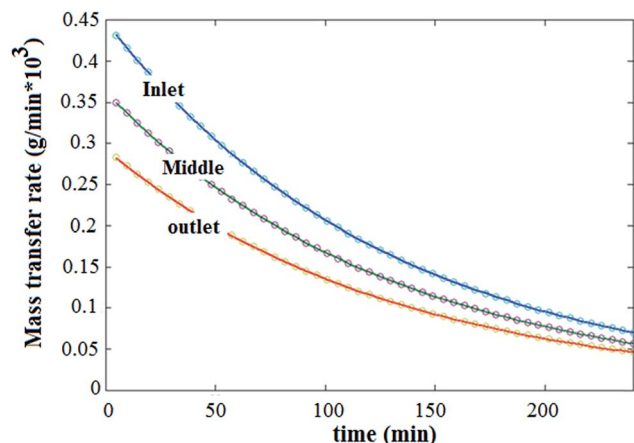


Fig. 5 Mass transfer rate of Th(IV) at various time, from feed side (solid lines) into the shell side (circles).

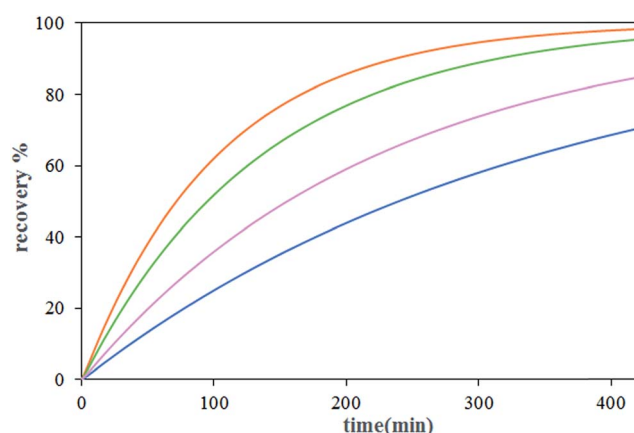


Fig. 6 The effect of carrier concentrations on thorium(IV) recovery efficiency (feed phase:  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$ , flow rate of tube and shell sides:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (—), 45% TBP (—), 30% TBP (—), 20% TBP (—), 15% TBP).

interval. Thus, over each time interval, the steady state condition in the liquid membrane is obtained. This finding is due to the low mass transfer resistance of LM in comparison with ones in feed and shell sides. Therefore, for reduction of the computational complexity and time, a pseudo steady state model for HFRLM process can be used.

### 3.3 Effect of carrier concentration on the mass transfer performance of HFRLM

The effect of carrier concentration on Th(IV) transport rate through HFRLM is shown in Fig. 6 and 7. Fig. 6 indicates that after 420 min an increase in the carrier concentration from 15 to 45% leads to the increase of recovery efficiency by 40% at the same feed and strip flow rates of  $5.6 \text{ mL min}^{-1}$ . This is related to the improvement of the uptake capacity of the liquid membrane on its interface to the feed side, while, the increase of the carrier concentration higher than 30% (v/v) TBP is less effective on the transport of Th(IV). As can be seen in this Fig. 7, the simulation results are in good agreement with the experimental data. According to eqn (16), the increase of the carrier concentration leads to an increase in the distribution coefficient of the Th(IV) extraction. Thus, by assuming instantaneously formation of solute-carrier complexes at the aqueous-organic interface, an increase in the concentration of Th(IV)-TBP complex at the renewal layer and in the renewal layer-LM interface leads to enhancement of driving force of Th(IV) transport through the HFRLM and consequently the mass transfer rate as well. Since the distribution coefficient of the back-extraction of Th(IV) is much less than that of the extraction one, an increase in the carrier concentration leads to the facilitation of Th(IV) transport through the HFRLM.

The variation of Th(IV) concentration in the reservoir tanks is also shown in Fig. 7. According to this figure, an increase in the concentration of TBP in kerosene from 15 to 30% (v/v) leads to an increase in the Th(IV) transfer rate so that recovery of 50% is achieved after 244 and 92 min, respectively. Moreover, simulation results illustrate that at the carrier concentration of 45%, the recovery of 50% is obtained at 72 min which indicates

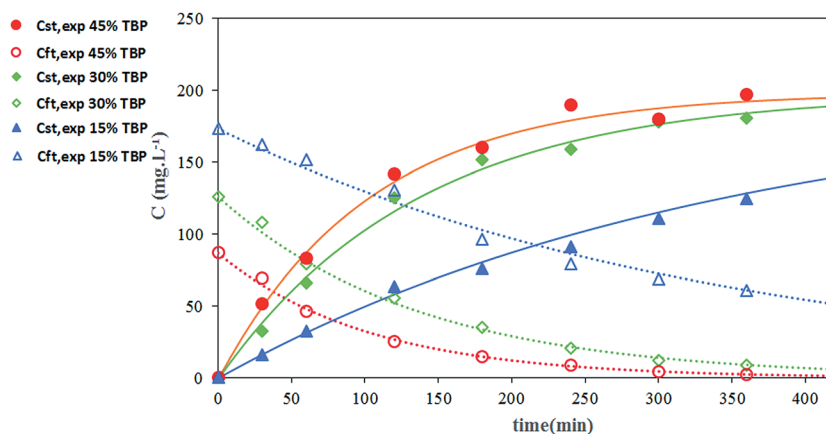


Fig. 7 The effect of carrier concentrations on thorium(IV) concentration in the reservoir tank (feed phase:  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$ , flow rates of tube and shell sides:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (—),  $C_{\text{st,cal}}$  at 45% TBP (—),  $C_{\text{st,cal}}$  at 30% TBP (—),  $C_{\text{st,cal}}$  at 15% TBP (—),  $C_{\text{st,exp}}$  at 45% TBP (—),  $C_{\text{st,exp}}$  at 30% TBP (—),  $C_{\text{st,exp}}$  at 15% TBP (—)).



that the increase of carrier concentration over than 30% has no significant effect on mass transfer rate. Furthermore, increasing the carrier concentration from 15 to 45% volume percent causes an increase in the distribution coefficient of the extraction process as well, which leads to a decrease in the initial solute concentrations of the donor phase, which is in equilibrium with liquid membrane droplets.

### 3.4 The influence of concentration of nitrate ion in feed

The effect of nitrate ion concentration in the feed phase is illustrated in Fig. 8 and 9 at various sodium nitrate concentrations of 1 M, 2 M and 3 M. According to eqn (16), an increase in the nitrate concentration of the feed phase causes the enhancement of the extraction distribution coefficient of Th(IV), which in turn enhances the extraction of metal ion from feed side and the transport of solute through the HFRLM as well. Furthermore Fig. 8 indicates that the increasing of the nitrate ion concentration of the feed phase from 1 M to 2 M causes a 3

times enhancement in thorium recovery, whereas a higher concentration has no significant effect on the recovery of Th(IV). As shown in Fig. 9, by increasing the nitrate ion concentration, the uphill effect takes place at a shorter time, which indicates the enhancement of the mass transfer rate. In addition, this figure shows that there is a good agreement between the calculated and experimental data.

### 3.5 Effect of initial concentration of Th(IV) in the feed phase

The transportation of Th(IV) via the HFRLM system at different initial concentration of Th(IV) in the feed (99.16, 198.32, 396.65 mg L<sup>-1</sup>) has been carefully studied. The variation of the recovery efficiency and Th(IV) concentration at the reservoir tanks are shown in Fig. 10 and 11, respectively. The presented results in Fig. 10 indicate that the recovery profile is almost the same at all investigated experimental conditions. The maximum recovery obtained at TBP concentration of 30% (v/v),

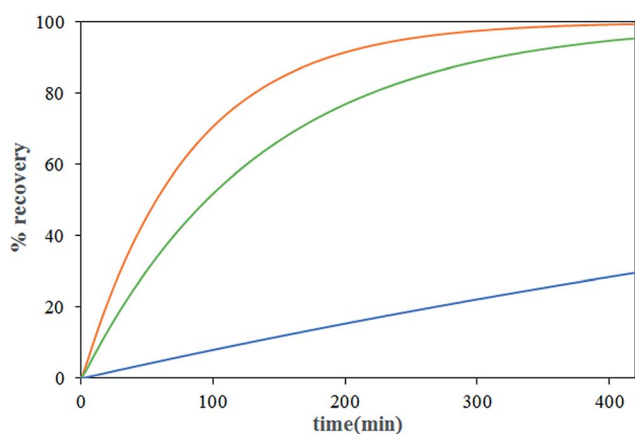


Fig. 8 The effect of nitrate ion concentrations of the feed on thorium(IV) recovery efficiency (feed phase:  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$ , flow rates:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (—),  $3 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $2 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $1 \text{ mol L}^{-1} \text{ NaNO}_3$ ).

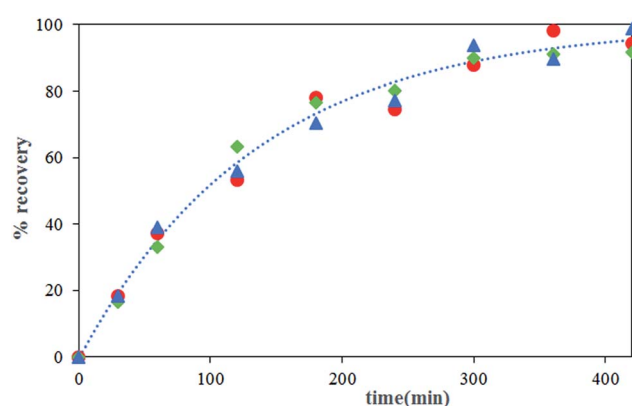


Fig. 10 The effect of initial concentrations of thorium(IV) ion on recovery efficiency ( $[\text{TBP}] = 30\% \text{ (v/v)}$ , flow rates of tube and shell sides:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (●),  $[\text{Th}^{4+}]_{\text{init}} = 396.65 \text{ mg L}^{-1}$  (◆),  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$  (▲),  $[\text{Th}^{4+}]_{\text{init}} = 99.16 \text{ mg L}^{-1}$  (●●●), simulated result).

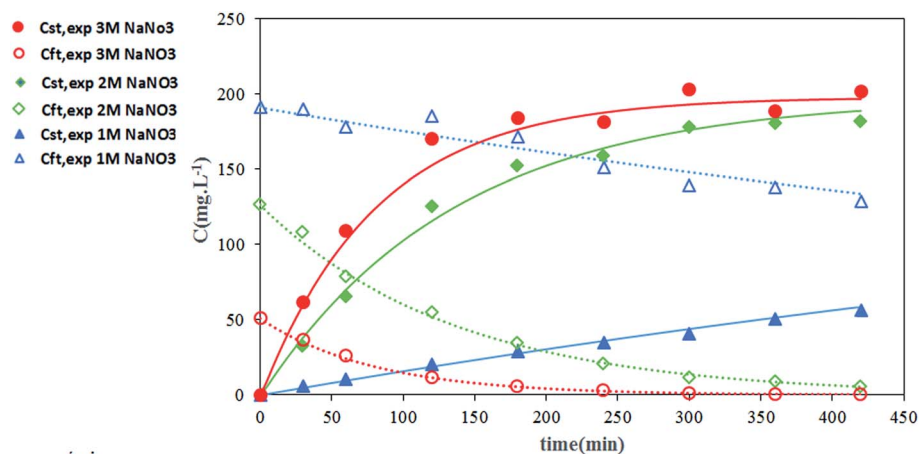
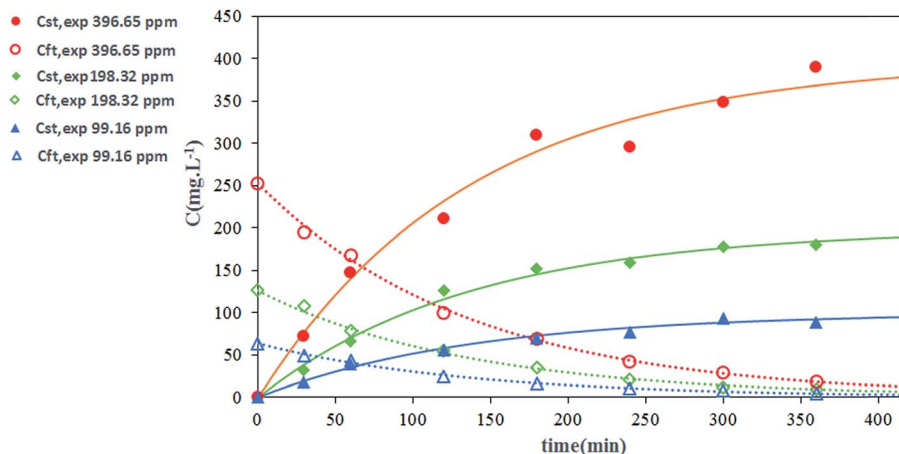


Fig. 9 The effect of nitrate ion concentrations of the feed on thorium(IV) concentration in the reservoir tank (feed phase:  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$ , flow rates:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$  (—),  $C_{\text{ft,cal}}$  at  $3 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $C_{\text{ft,cal}}$  at  $2 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $C_{\text{ft,cal}}$  at  $1 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $C_{\text{st,cal}}$  at  $3 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $C_{\text{st,cal}}$  at  $2 \text{ mol L}^{-1} \text{ NaNO}_3$  (—),  $C_{\text{st,cal}}$  at  $1 \text{ mol L}^{-1} \text{ NaNO}_3$ ).





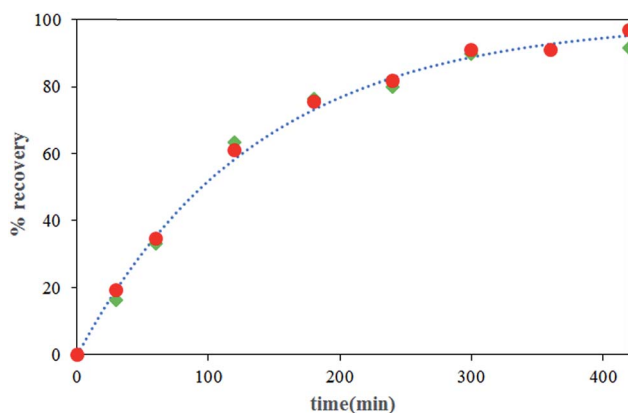


**Fig. 11** The effect of initial thorium(IV) concentration on the concentrations of Th(IV) at reservoir tank ([TBP] = 30% (v/v), feed and strip flow rates:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (•••••),  $C_{f,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 396.65 \text{ mg L}^{-1}$  (•••••),  $C_{f,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$  (•••••),  $C_{f,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 99.16 \text{ mg L}^{-1}$  (—),  $C_{s,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 396.65 \text{ mg L}^{-1}$  (—),  $C_{s,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$  (—),  $C_{s,\text{cal}}$  at  $[\text{Th}^{4+}]_{\text{init}} = 99.16 \text{ mg L}^{-1}$ ).

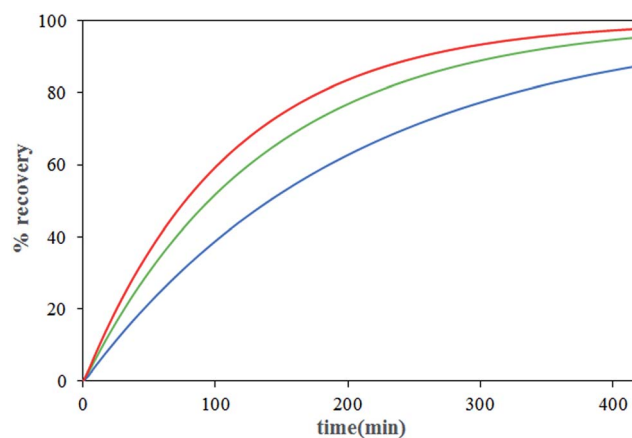
nitrate ion concentration of 2 M and 0.1 M in the feed and strip phases respectively, is 95.3% after 420 min. Since the free concentration of TBP is much higher than Th(IV)–TBP complex, the distribution coefficients of the extraction and back extraction at all investigated conditions are similar according to eqn (16), so the recovery efficiency of experiments is the same as well. In addition, at the diluted conditions, same as the condition of this study, the transport properties might be assumed constant. Thus, the uphill effect takes place at the same time. Moreover, increasing of Th(IV) concentration of the feed could shift the equilibrium equation towards the formation of a more Th(IV)–TBP complex at the feed–membrane interface. Thus, this causes the enhancement of the mass transfer driving force and more solute transports through the hollow fiber renewal system as shown in Fig. 11. Furthermore, the experimental data can prove this claim. This observation has been made by Vernekar *et al.*, who modeled Cobalt(II) transport through a HFSLM in the recycling mode.<sup>27</sup>

### 3.6 Effect of flow rate

The effects of flow rate of the strip and feed phases on thorium recovery are shown in Fig. 12 and 13, respectively. As shown in these figures, although the recovery of thorium has been increased by increasing the tube side flow rate, the increasing of strip phase flow rate has no significant effect on the recovery of thorium. This indicates that the mass transfer resistance in the tube side is higher than the shell side, which is in agreement with our previous study.<sup>25</sup> Enhancement of the tube side velocity causes a decrease in the renewal layer thickness and an increase in the renewal rate of organic layer at the inner surface of hollow fiber, which leads to an increase in the local mass transport coefficient. Furthermore, Fig. 14 shows the variation of Th(IV) concentration with time at the reservoir tanks. This figure shows that the calculated results are in good agreement with experimental data.



**Fig. 12** The effect of shell side flow rate on Th(IV) recovery efficiency ([TBP] = 30% (v/v), tube side flow rate:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (•),  $Q_s = 10.6 \text{ mL min}^{-1}$  (•),  $Q_s = 5.6 \text{ mL min}^{-1}$  (•), simulated result).



**Fig. 13** The effect of tube side flow rate on Th(IV) recovery efficiency ([TBP] = 30% (v/v), shell side flow rate:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (—),  $Q_f = 10.6 \text{ mL min}^{-1}$  (—),  $Q_f = 5.6 \text{ mL min}^{-1}$  (—),  $Q_f = 2.3 \text{ mL min}^{-1}$ ).

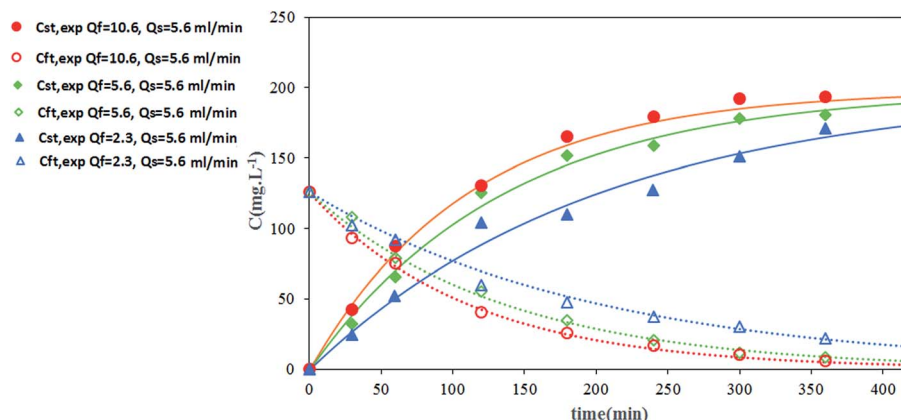


Fig. 14 The effect of tube side flow rate on Th(IV) concentration profiles in the reservoir tank (feed phase:  $[\text{Th}^{4+}]_{\text{init}} = 198.32 \text{ mg L}^{-1}$ , flow rates of strip phase:  $5.6 \text{ mL min}^{-1}$ , strip phase:  $[\text{HNO}_3] = 0.1 \text{ mol L}^{-1}$ , (—, +, +, +, +, +).  $C_{f,\text{cal}}$  at  $Q_f = 10.6 \text{ mL min}^{-1}$  (—, +, +, +, +, +),  $C_{f,\text{cal}}$  at  $Q_f = 5.6 \text{ mL min}^{-1}$  (—, +, +, +, +, +),  $C_{f,\text{cal}}$  at  $Q_f = 2.3 \text{ mL min}^{-1}$  (—, +, +, +, +, +),  $C_{s,\text{cal}}$  at  $Q_f = 11 \text{ mL min}^{-1}$  (—, +, +, +, +, +),  $C_{s,\text{cal}}$  at  $Q_f = 5.6 \text{ mL min}^{-1}$  (—, +, +, +, +, +),  $C_{s,\text{cal}}$  at  $Q_f = 2.3 \text{ mL min}^{-1}$  (—, +, +, +, +, +).

### 3.7 Accuracy of the mathematical model

Accuracy of developed model has been measured as the average relative error. The calculated error using by using eqn (19) is less than 10% and this shows the validity of the developed model.<sup>28,29</sup>

## 4 Conclusions

In this paper, a rigorous mathematical model has been presented to simulate thorium recovery from an acidic medium through the HFRLM in the recycle mode. The model's predictions have been validated by the experimental data. The important findings of this study are given as following:

(1) The mass transfer rate of Th(IV) from the feed to the LM phase is equal to that of the LM to the stripping phase during each time interval. Thus, considering the steady state condition in the liquid membrane might be a correct assumption, due to the low mass transfer resistance of the LM in comparison with the ones in the feed and shell sides. Therefore, for reduction of the computational complexity and time, a pseudo steady state model for recycling mode of the HFRLM process can be used.

(2) Experimental results show that increasing the carrier concentration, the tube side flow rate as well as nitrate ion concentration of the feed phase, could increase the recovery of solute. While, enhancement of the shell side flow rate and initial thorium concentration have no significant effect on recovery efficiency of thorium.

(3) Finally, the calculated data are found to be in good agreement with the experimental results, which indicates that the developed model might be a useful tool for the design and scale-up of the HFRLM process for recovery of thorium from acidic waste solution.

## Nomenclature

### Abbreviations

HFRLM	Hollow fiber renewal liquid membrane
HFSLM	Hollow fiber supported liquid membrane

LM  
TBP

Liquid membrane  
Tri-butyl phosphate

### Symbols

$A_{\text{eff}}$	Effective mass transfer area, $\text{m}^2$
$A$	Cross sectional area of passing fluid, $\text{m}^2$
$C$	Concentration, $\text{mol m}^{-3}$
$D$	Diameter, m
$D_{\text{eff}}$	Effective diffusivity, $\text{m}^2 \text{s}^{-1}$
$D_{\text{in}}$	Inner diameter of fiber, m
$D_{\text{out}}$	Outer diameter of fiber, m
$D_0$	Diffusivity of solute in liquid membrane phase, $\text{m}^2 \text{s}^{-1}$
$K$	Local mass transfer coefficient, $\text{m s}^{-1}$
$k_{\text{eq}}$	Equilibrium constant, $\text{kmol m}^{-3}$
$l$	Effective length of hollow fiber, m
$m$	Distribution coefficient of extraction process
$M$	Molecular mass of diluent, $\text{kg mol}^{-1}$
$N_A$	Mass transfer flux of solute, $\text{mol m}^2 \text{s}^{-1}$
$n$	Number of hollow fibers in module
$O$	Organic phase volume, $\text{m}^3$
$P$	Wetted perimeter, m
$Q_s$	Volumetric flow rate of the strip phase, $\text{m}^3 \text{s}^{-1}$
$Q_f$	Volumetric flow rate of the feed phase, $\text{m}^3 \text{s}^{-1}$
$R_i$	Internal radius of the shell side of the module, m
$R\%$	Recovery efficiency, %
$r_{\text{in}}, r_{\text{out}}$	Internal and external radii of hollow fiber, m
$r$	Radial direction, m
$t$	Time, s
$T$	Absolute temperature, K
$V_t$	Volume of the reservoir tank, $\text{m}^3$
$V_M$	Molar volume of extractant, $\text{m}^3 \text{mol}^{-1}$
$x$	Association parameter of diluent



## Paper

z Longitude direction, m

## Greek letters

$\varepsilon$  Porosity of hollow fiber membrane support  
 $\tau$  Tortuosity of hollow fiber membrane support  
 $\mu$  Viscosity of solvent (extractant + diluent), Pa s  
 $\rho_K$  Density of kerosene ( $\text{Kg m}^{-3}$ )

## Superscripts

Cal Calculated  
 exp Experimental  
 ext External  
 int Internal

## Subscripts

A Acceptor phase  
 aq. Aqueous phase  
 f Feed phase  
 fin Inlet feed phase to the reservoir tank  
 ft Feed phase in the reservoir tank  
 ft0 Initial feed phase in the reservoir tank  
 init Initial  
 m Membrane phase  
 mi Interfacial surface at the inner liquid membrane phase  
 org Organic phase  
 R Renewal layer  
 Ri Interfacial surface at the inner renewal layer  
 s Strip phase  
 si Interfacial surface of the liquid membrane and strip phases  
 sin Inlet strip phase to the reservoir tank  
 st Strip phase in the reservoir tank  
 t Reservoir tank

## References

- 1 M. Metaxas, V. Kasselouri-Rigopoulou, P. Galiatsatou, C. Konstantopoulou and D. Oikonomou, *J. Hazard. Mater.*, 2003, **97**, 71.
- 2 V. Dekoussar, G. R. Dyck, A. Galperin, C. Ganguly, M. Todosow and M. Yamawaki, *Thorium fuel cycles: potential benefits and challenges*, in, *International atomic energy agency (IAEA)*, Vienna, 2005.
- 3 K. Gupta, V. Manchanda, M. Subramanian and R. Singh, *Solvent Extr. Ion Exch.*, 2000, **18**, 273.
- 4 E. B. M. F. San Román, R. Ibanez and I. Ortiz, *J. Chem. Technol. Biotechnol.*, 2010, **85**, 2.
- 5 A. Kiani, R. Bhavé and K. Sirkar, *J. Membr. Sci.*, 1984, **20**, 125.
- 6 Q. Yang and N. M. Kocherginsky, *J. Membr. Sci.*, 2006, **286**, 301.
- 7 F. J. Alguacil, M. Alonso, F. Lopez, A. López-Delgado, I. Padilla and H. Tayibi, *Chem. Eng. J.*, 2010, **157**, 366.
- 8 K. Wieszczczycka, M. Regel-Rosocka, K. Staszak, A. Wojciechowska, M. T. A. Reis, M. R. C. Ismael, M. L. F. Gameiro and J. M. Carvalho, *Sep. Purif. Technol.*, 2015, **154**, 204.
- 9 X. P. Dai, Z. F. Yang, R. G. Luo and K. K. Sirkar, *J. Membr. Sci.*, 2000, **171**, 183.
- 10 Z. Ren, W. Zhang, H. Meng, J. Liu and S. Wang, *J. Membr. Sci.*, 2010, **365**, 260.
- 11 Z. Ren, W. Zhang, Y. Liu, Y. Dai and C. Cui, *Chem. Eng. Sci.*, 2007, **62**, 6090.
- 12 Z. Ren, W. Zhang, H. Li and W. Lin, *Chem. Eng. J.*, 2009, **146**, 220.
- 13 Z. Ren, X. Zhu, W. Liu, W. Sun, W. Zhang and J. Liu, *Chin. J. Chem. Eng.*, 2014, **22**, 1187.
- 14 Z. Ren, W. Zhang, Y. Lv and J. Li, *Biotechnol. Prog.*, 2009, **25**, 468.
- 15 J. S. Carletto, R. M. Luciano, G. C. Bedendo and E. Carasek, *Anal. Chim. Acta*, 2009, **638**, 45.
- 16 Z. Dingjie, J. Yue and L. Xiaolong, *Ind. Water Treat.*, 2009, **11**, 019.
- 17 W. Zhang, C. Cui, Z. Ren, Y. Dai and H. Meng, *Chem. Eng. J.*, 2010, **157**, 230.
- 18 A. Fortuny, M. T. Coll and A. M. Sastre, *Desalination*, 2014, **343**, 54.
- 19 C. Yun, R. Prasad and K. Sirkar, *Ind. Eng. Chem. Res.*, 1993, **32**, 1186.
- 20 M. F. San Román, E. Bringas, R. Ibanez and I. Ortiz, *J. Chem. Technol. Biotechnol.*, 2010, **85**, 2.
- 21 Z. Ren, W. Zhang, Y. Dai, Y. Yang and Z. Hao, *Ind. Eng. Chem. Res.*, 2008, **47**, 4256.
- 22 Z. Ren, Y. Yang, W. Zhang, J. Liu and H. Wang, *J. Membr. Sci.*, 2013, **439**, 28.
- 23 S. Michaud, M. Miguiditchian, G. Deblonde, B. Dinh, X. Hérès and G. Andreoletti, *Procedia Chem.*, 2012, **7**, 251.
- 24 M. Benedict, T. H. Pigford and H. W. Levi, *Nuclear Chemical Engineering*, McGraw Hill Book Company, 1981.
- 25 S. A. Allahyari, A. Minucmehr, S. J. Ahmadi and A. Charkhi, *J. Membr. Sci.*, 2016, **520**, 374.
- 26 C. R. Wilke and P. Chang, *AIChE J.*, 1955, 264.
- 27 P. V. Vernekar, Y. D. Jagdale, A. W. Patwardhan, A. V. Patwardhan, S. A. Ansari, P. K. Mohapatra and V. K. Manchanda, *Chem. Eng. Res. Des.*, 2013, **91**, 141.
- 28 P. Kandwal, S. Dixit, S. Mukhopadhyay and P. K. Mohapatra, *Chem. Eng. J.*, 2011, **174**(1), 110.
- 29 P. J. Harrington and G. W. Stevens, *J. Membr. Sci.*, 2001, **192**(1), 83.

