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# Polymerizable ionic liquids and polymeric ionic liquids: facile synthesis of ionic liquids containing ethylene oxide repeating unit *via* methanesulfonate and their electrochemical properties†

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Polymeric ionic liquids have shown great potential in numerous application fields, and an easy access to polymerizable monomer is the key to the polymerized ionic liquids. The polymerizable monomer is usually synthesized *via* the quaternization of a corresponding bromide, which is usually prepared from bromination of the corresponding alcohol but with a low yield. In this study, a polymerizable monomer was prepared by quaternization of a mesylate, derived from the corresponding alcohol, followed by an ion exchange reaction. The polymerizable ILs M1–M4 showed ionic conductivity as high as  $10^{-3}$  S cm<sup>-1</sup>, and the oxygen effect was observed with the lithium salt. Polymeric ionic liquids P1 and P2 were prepared *via* radical polymerization on M1 and M2, and the ionic conductivity of polymeric ionic liquid P2 was as high as  $2.2 \times 10^{-4}$  S cm<sup>-1</sup> with electrochemical stability exceeding 7 V (vs. Li/Li<sup>+</sup>).

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## Introduction

With the fast development of electric vehicles and mobile electric devices, rechargeable batteries with high performance are highly required. Lithium ion batteries (LIBs) have been improved rapidly and possess high energy density and excellent cyclic stability. However, several safety problems have been encountered and numerous accidents have occurred due to LIBs. To solve the safety issue of LIBs, electrolytes are the main components that could be improved, which function as the charge carriers between electrodes. Traditionally, organic liquid electrolytes, such as carbonate ester, were widely used,<sup>1</sup> which have drawbacks such as volatility, leakage, flammability, and toxicity. Thus, the synthesis of solvent-free electrolytes has attracted the attention from materials scientists.

Solid polymer electrolyte (SPE)<sup>2,3</sup> is a popular concept used by people to reconstruct old electronic device systems. Polyethylene oxide (PEO)-based SPE has particularly been paid attention to for application in LIBs. Pure PEO has poor ionic

conductivity, which hardly meets the requirement for LIBs, and the addition of carbonate ester makes the leakage of carbonate ester a possibility. Ionic liquid has numerous advantages:<sup>4</sup> low vapor pressure, non-flammability, and good ion conduction, and these properties make it a better candidate for application in LIBs.

Solid or quasi-solid electrolytes,<sup>5</sup> such as polymer electrolytes, usually have good mechanical stability and simple processing. Among them, polymeric ionic liquids (PILs)<sup>6–8</sup> are the most promising solid electrolyte candidates, since they combine the advantages of ionic liquids and polymers. Plenty of pioneering studies have been directed at obtaining high ionic conductivity PILs.<sup>3,9–26</sup> Since the first PIL synthesized by the Ohno group,<sup>27</sup> several attempts have been made to investigate the tether effect<sup>28,29</sup> between the charge center and backbone on the ionic conductivity improvement of the PILs. The introduction of oxygen atom in the spacer was found to decrease the ionic conductivity of PILs<sup>30,31</sup> since the oxygen trapped the mobile cation was different from that in the traditional SPE. The Vidal group researched the nitrogen element in the spacer and an ionic conductivity as high as  $6.5 \times 10^{-4}$  S cm<sup>-1</sup> at 20 °C was achieved.<sup>32</sup> From the reported study, it was strongly suggested that the tether effect played an important role in determining the ionic conductivity. However, a systematic study on the tether effect as well as the heteroatom effect on ionic conductivity of ILs as well as PILs was not available.

Moreover, an easy access to polymerizable ILs is highly required in order to accelerate the development of high performance PILs suitable for battery applications. In this sense, the synthetic method towards ionic liquid monomers is

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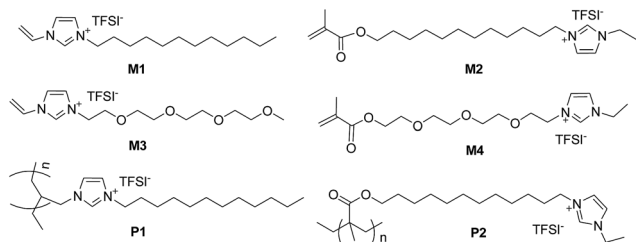


Fig. 1 Structure of some polymerizable ionic liquids and polymeric ionic liquids in this study.

very important as well as is study on the structure–property relationship of ILs and PILs. In this study, the design and synthesis of polymerizable ionic liquid monomer M1 to M4 as well as PILs P1 and P2 is presented (Fig. 1) towards developing a highly efficient synthetic procedure. In addition, the ionic conductivity of the synthesized ILs and PILs was investigated.

## Results and discussion

### Synthesis and characterization of polymerizable ionic liquid monomer M1 to M4

Our study initiated with the synthesis of polymerizable ionic liquid M1. According to the reported procedure, we tried the quaternization reaction between bromide 1 and *N*-vinylimidazole 2 in organic solvents, such as methanol and chloroform, and the imidazolium bromide was obtained as a precipitate from acetonitrile. In order to make the synthetic procedure greener and with higher economic efficiency, the solvent-free quaternization reaction was carried out at 80 °C. The reaction was effective and gave the target product in 96% isolated yield, which was transferred into polymerizable ionic liquid M1 by exchange with LiTFSI in 90% isolated yield (Scheme 1).

The chemical structure of M1 was determined by examining its <sup>1</sup>H-NMR spectrum (Fig. 2), which was very similar to that of its corresponding bromide (Fig. S1†).

The structure was further confirmed by high-resolution mass spectroscopy (Fig. 3 and 4).

With M1 prepared, we moved to the synthesis of M2 using dibromide 3 (Scheme 2).

The commercially available dibromide 3 was mono-substituted by methacrylic acid in the presence of Cs<sub>2</sub>CO<sub>3</sub>. It was found that the monoester reaction outcome strongly depended on the reaction conditions. Using excess dibromide and adding acid into the dibromide solution, monoester 4 was isolated in 45% yield after flash column chromatography. Bromide 4 was transferred into polymerizable ionic liquid M2 in



Scheme 1 Synthesis of ionic liquid M1 via alkylbromide 1.

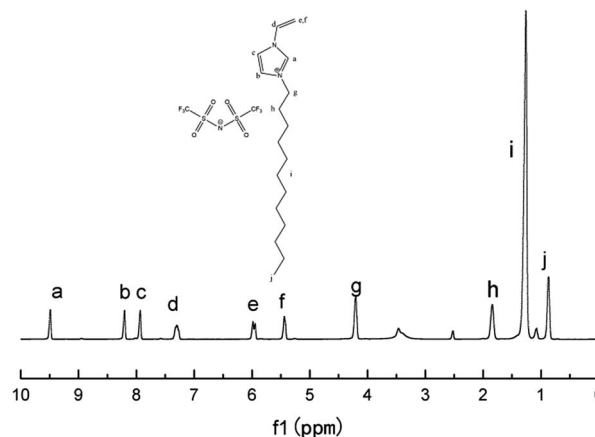


Fig. 2 <sup>1</sup>H-NMR spectrum of ionic liquid monomer M1 in DMSO-D<sub>6</sub>.

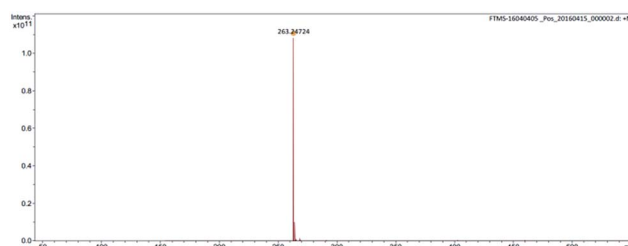


Fig. 3 Mass spectrum of cation of ionic liquid monomer M1.

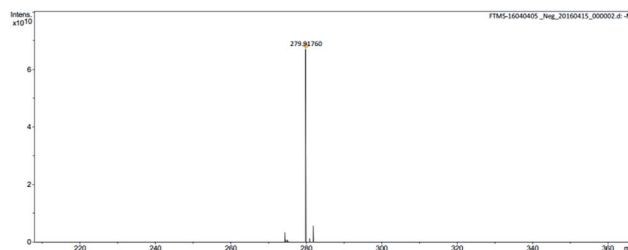
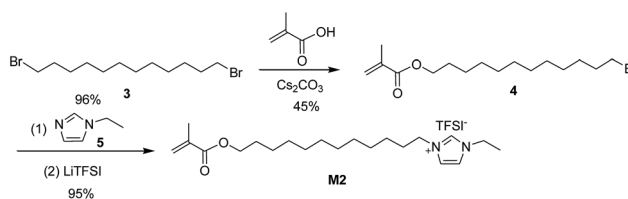


Fig. 4 Mass spectrum of anion of ionic liquid monomer M1.



Scheme 2 Synthesis of ionic liquid M2 via alkylbromide 3.

95% yield by treating 4 with *N*-ethylimidazole, 5 followed by ion exchange with LiTFSI.

The chemical structure of M2 was determined by examining its <sup>1</sup>H-NMR spectrum (Fig. 5), which was very similar to that of its corresponding bromide (Fig. S2†). This was further confirmed *via* high-resolution mass spectroscopy (Fig. 6).



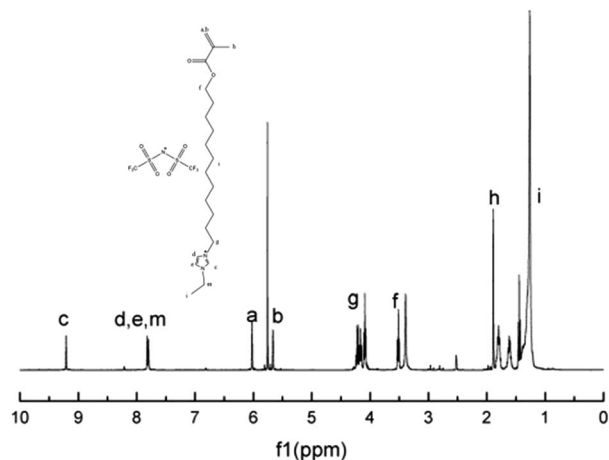


Fig. 5  $^1\text{H-NMR}$  spectrum of ionic liquid monomer M2.

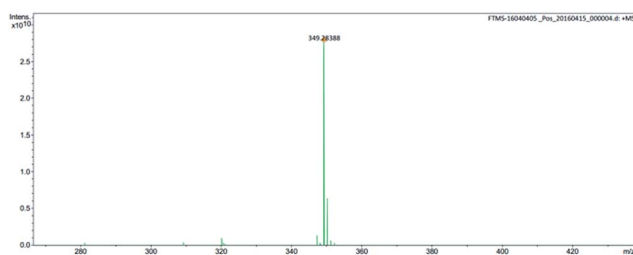


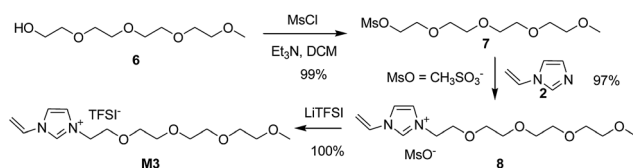
Fig. 6 Mass spectrum of anion of ionic liquid monomer M2.

With the experience of quaternization reaction and ion exchange reaction obtained from the synthesis of M1 and M2, we began with the synthesis of M3 and M4 following the same procedure. However, the bromination of alcohol 6 did not work using either HBr or  $\text{PBr}_3$ . As an alternative, alcohol 6 was transformed into mesylate 7 (Fig. S3†) in 99% isolated yield by treating 6 with methanesulfonyl chloride (MsCl) in the presence of  $\text{Et}_3\text{N}$ . To our delight, the quaternization reaction between mesylate 7 and *N*-vinylimidazole 2 worked nicely to give imidazolium mesylate 8 in 97% yield, which was transferred into M3 in 100% yield (Scheme 3).

The chemical structure of M3 was determined by examining its  $^1\text{H-NMR}$  spectrum (Fig. 7).

It was further confirmed by high-resolution mass spectroscopy of M3 (Fig. S6†).

Encouraged by success on the synthesis of M3, we moved to the synthesis of M4 starting from diol 9. Diol 9 was transformed into dimesylate 10 (Fig. S4†) in 99% isolated yield by treating 9 with methanesulfonyl chloride (MsCl) in the presence of  $\text{Et}_3\text{N}$ .



Scheme 3 Synthesis of ionic liquid M3 via mesylate.

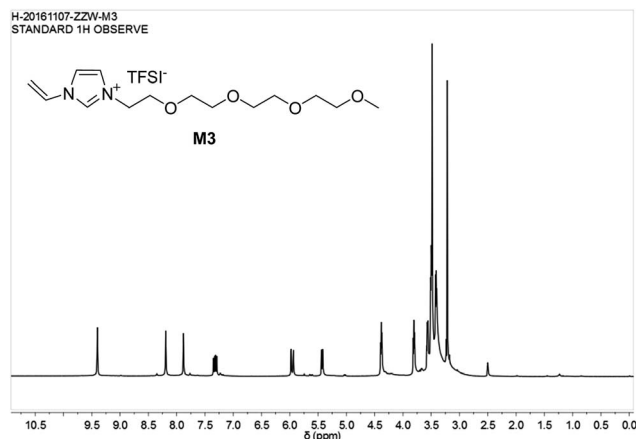


Fig. 7  $^1\text{H-NMR}$  spectrum of ionic liquid monomer M3.

The monosubstitution of dimesylate 10 was achieved using methacrylic acid in basic conditions to give monoester 11 (Fig. S5†) in 45% isolated yield. Finally, ionic liquid M4 was obtained in 94% yield by the quaternization reaction between mesylate 11 and *N*-ethylimidazole 5, followed by ion exchange reaction with LiTFSI (Scheme 4).

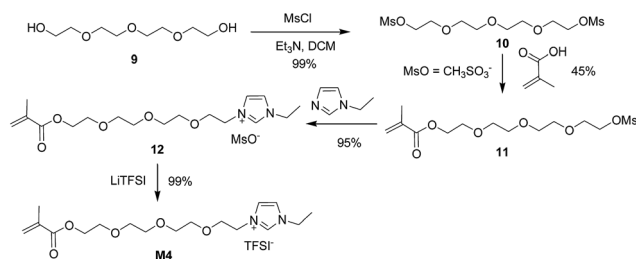
The mesylate anion of IL 12 was transferred into  $\text{TFSI}^-$  by exchange with LiTFSI in water, and M4 was obtained. This was confirmed *via* the  $^1\text{H-NMR}$  spectrum of M4 (Fig. 8).

### Synthesis of polymeric ionic liquid P1 and P2

With polymerizable ionic liquid monomer M1 and M2 in hand, we conducted the radical polymerization reaction. M1 was first polymerized in the presence of AIBN in DMF with heating to give P1 as a rubber-like solid; its  $^1\text{H-NMR}$  spectrum showed the complete disappearance of vinyl group between 5–6 ppm (Fig. 9).

Similarly, M2 was polymerized to give P2 with AIBN as the initiator. The structure of P2 was characterized by examining its  $^1\text{H-NMR}$  spectrum (Fig. 10).

Molecular weight determinations were attempted using GPC, but unfortunately no peak could be observed despite several attempts. As an alternative, we carried out the molecular weight measurement *via*  $^1\text{H-NMR}$  observations according to known methods,<sup>33</sup> which gave molecular weights of  $16\,588\text{ g mol}^{-1}$  for P1 and  $15\,782\text{ g mol}^{-1}$  for P2.



Scheme 4 Synthesis of ionic liquid M4 via mesylate.



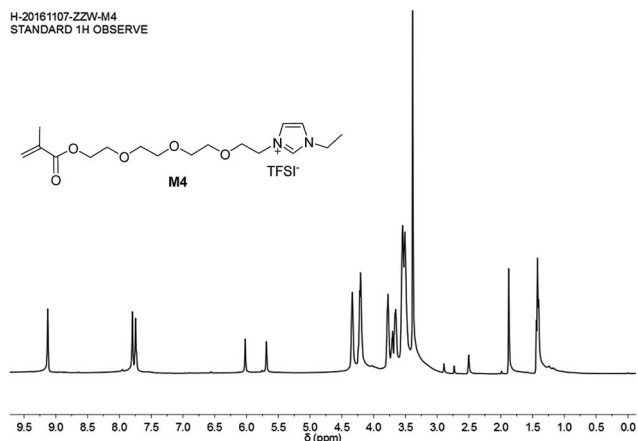


Fig. 8  $^1\text{H-NMR}$  spectrum of ionic liquid monomer M4.

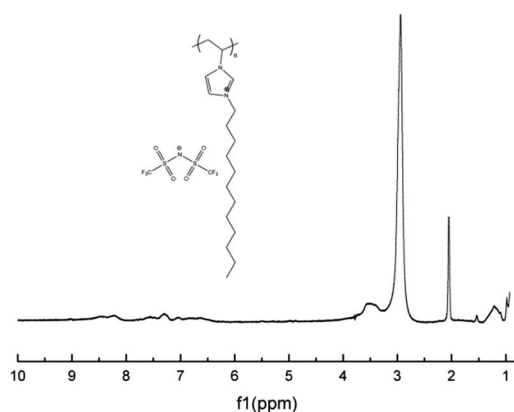


Fig. 9  $^1\text{H-NMR}$  spectrum of polymeric ionic liquid P1.

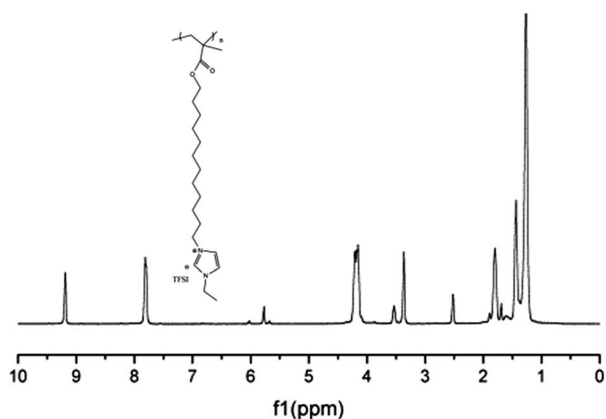


Fig. 10  $^1\text{H-NMR}$  spectrum of polymeric ionic liquid P2.

### Ionic conductivity of ionic liquid monomer M1 to M4 and polymer P1 and P2

The ionic conductivity of pure ionic liquids M1 to M4 was measured to be  $1.36 \times 10^{-3} \text{ S cm}^{-1}$ ,  $1.3 \times 10^{-3} \text{ S cm}^{-1}$ ,  $1.79 \times 10^{-3} \text{ S cm}^{-1}$  and  $1.89 \times 10^{-3} \text{ S cm}^{-1}$ , respectively, at  $30^\circ\text{C}$ .

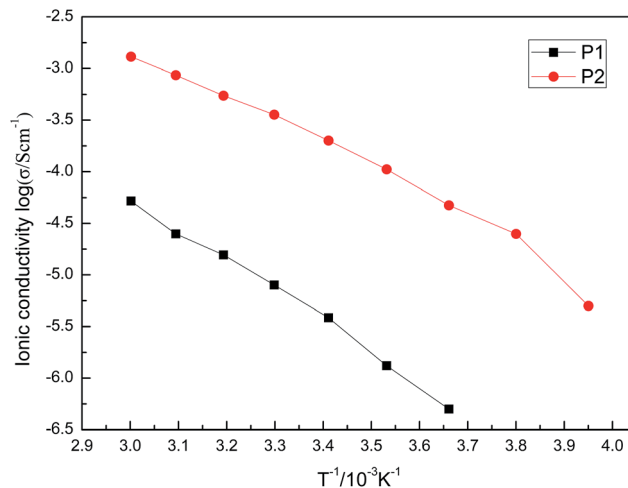


Fig. 11 Temperature dependence of polymer P1 and P2.

The ionic conductivity of P1 and P2 was dependent on the temperature, as shown in Fig. 11.

It was shown that P1 and P2 have higher ionic conductivity at higher temperatures, which arises from the movement improvement of the polymer chains and free volume. It also showed a nearly linear relationship between  $\log \sigma$  and  $T$ . The activation energies of P1 and P2 were calculated to be  $58.58 \text{ kJ mol}^{-1}$  and  $48.73 \text{ kJ mol}^{-1}$ , respectively. The ionic conductivity of P1 and P2 was measured to be  $1.1 \times 10^{-6} \text{ S cm}^{-1}$  and  $2.2 \times 10^{-4} \text{ S cm}^{-1}$ , respectively, at  $30^\circ\text{C}$ . Ionic conductivity of P1 decreased to one thousandth of M1 after polymerization of M1, which was observed for most of the PILs. However, P2 decreased to one tenth of M2 after M2 polymerization. This 100-fold difference of ionic conductivity definitely comes from the structural differences between P1 and P2. The long alkyl chain between the polymer backbone and charge centre keeps the TFSI<sup>-</sup> ions moving more freely compared with the case charge centre attached to the polymer backbone. Essentially, the large difference of ionic conductivity also corresponds to the big difference of  $T_g$  between P1 and P2, at  $-22^\circ\text{C}$  vs.  $-67^\circ\text{C}$ .

### Lithium salt concentration dependence of M1 to M4 and P1 to P2

In order to evaluate the potential of ILs and PILs in LIBs, the interaction between ILs (PILs) and lithium salt was investigated.

As shown in Fig. 12, with the increase of the ILs M1–M4 content, the ionic conductivity of the LiTFSI and ILs composites went up. Pure LiTFSI had ionic conductivity that was too low to be measured. When 94 wt% M1 was added into LiTFSI, the ionic conductivity increased to  $8 \times 10^{-4} \text{ S cm}^{-1}$  at  $30^\circ\text{C}$ ; similarly,  $9.6 \times 10^{-4} \text{ S cm}^{-1}$  for M2,  $1.17 \times 10^{-3} \text{ S cm}^{-1}$  for M3, and  $1.43 \times 10^{-3} \text{ S cm}^{-1}$  for M4 was observed. Clearly, oxygen-containing ionic liquids M3 and M4 showed higher ionic conductivity than all-carbon chain ionic liquids M1 and M2. This may be the result of dissociation ability of the ethylene oxide chain with the lithium salt. M2 had higher ionic conductivity than M1 and M4



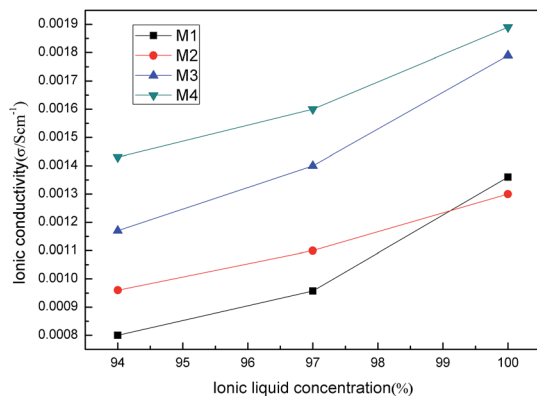


Fig. 12 IL helps lithium salt ionic conductivity at 30 °C.

higher than M3 under the same conditions. This may be caused by the longer chain of M2 and M4 than that of M1 and M3 as well as the plasticizing effect of the ester moiety.

In the case of PILs, P2 was investigated towards LiTFSI due to its high ionic conductivity. When LiTFSI was introduced into P2 during the polymerization process, the ionic conductivity of the composite decreased. The addition of LiTFSI to P2 may make a dimer of TFSI<sup>-</sup> formed, which will have bigger ion volume corresponding to the poor mobility, and thus provides lower ionic conductivity than pure P2. Higher Li-salt concentration resulted in higher dimer concentration in the system and lower ionic conductivity (Fig. 13).

### Electrochemical stability

For further application of P2 as solid electrolyte, it is very important to know its electrochemical stability. Generally, the point corresponding to the irreversible change is defined to be the decomposition voltage, which, in other words, means the stability boundary or so-called window.

As shown in Fig. 14, non-oxidizable electrochemical stability of P2 exceeded 7 V (*vs.* Li/Li<sup>+</sup>), whereas anti-reductive electrochemical stability was about 2.5 V. This demonstrated high potential for applications in high-voltage systems.

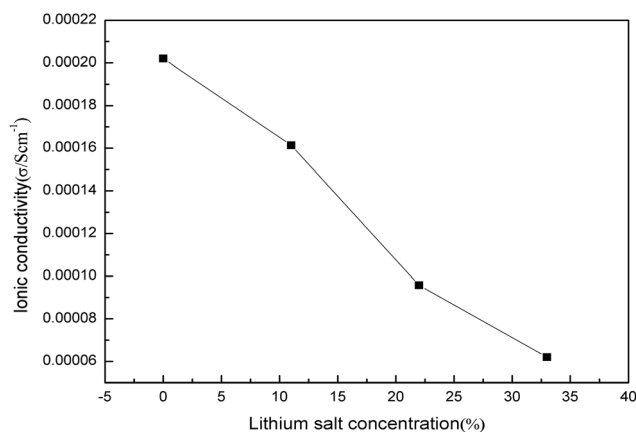


Fig. 13 Lithium salt concentration dependence of P2.

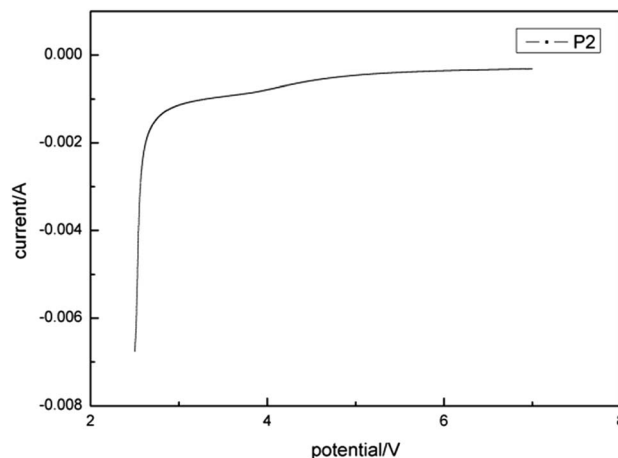


Fig. 14 Electrochemical stability of P2.

### Thermal stability

An ideal solid electrolyte for lithium batteries will have high thermal stability. The decomposition temperature of P1 and P2 was measured by TGA (Fig. S7†).

P1 lost 5% weight at 369 °C, whereas P2 at 234 °C, but P2 has higher maximum decomposed temperature, and decomposed more completely than P1.

## Experimental

### Materials

1-Vinyl-1*H*-imidazole, 1-bromoundecane, azodiisobutyronitrile (AIBN), 1,12-dibromo dodecane, *N*-ethylimidazole, 2,5,8,11-tetraoxatridecan-13-ol, methanesulfonyl chloride, 2,2'-((oxybis(ethane-2,1-diyl))bis(oxy))diethanol, cesium carbonate, methanesulfonyl chloride, lithiumbis((trifluoromethyl)sulfonyl)-amide (LiTFSI) and methanol were purchased from Tianjin Baiheng technology Co., Ltd. Dichloromethane, trichloromethane, and ethylacetate were purchased from Beijing Chemical Works. Anhydrous ether, and *N,N*-dimethylformamide were purchased from Beijing Tong Guang Fine Chemical Co., Ltd. All reagents were used as received.

### Characterisation method

The <sup>1</sup>H-NMR spectrum of sample dissolved in either CDCl<sub>3</sub> or DMSO-D<sub>6</sub> was recorded on a Varian Mercury plus 400 M nuclear magnetic resonance spectrometer. The molecular weight of ionic liquid monomers was obtained through the Fourier transform ion cyclotron resonance mass spectrometer, a Bruker APEX IV. The ionic conductivity was measured as follows:

$$\sigma = dl/(R \times A)$$

where, *d* is the film thickness, and *R* is the resistance value, which is obtained from the AC impedance measurement. *A* is the area of the film in contact with the blocking electrode. In the AC impedance spectrum, the value of *R* was obtained from *Z'* value when *-Z''* reached minimum. In this study, the structure



of the substrate for the AC impedance measurement was ss|PIL|ss (ss means stainless steel). Electrochemical stability was measured *via* linear sweep voltammetry or cyclic voltammetry. After the establishment of Pt/PIL/Pt, the cyclic voltammetry test was carried out with the scan speed of 0.1 mV s<sup>-1</sup> and voltage range between 2 V and 7 V.

### Synthetic procedures

**Synthesis of M1.** To 1-bromoundecane (1 ml, 4.17 mmol) in methanol (30 ml), 1-vinyl-1*H*-imidazole (0.4 ml, 4.17 mmol) was added at room temperature. The mixture was vigorously stirred at 100 °C (oil bath) for 10 h, and the solvents were removed by distillation. The residue was washed using ethyl acetate to give imidazolium bromide as a white precipitate (1.37 g, 4.0 mmol, 96%), which was pure enough based on its <sup>1</sup>H-NMR spectrum to be used directly for the next step.

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 0.85 (s, 3H), 1.23 (m, 18H), 1.82 (s, 2H), 4.20 (s, 2H), 5.42 (d, *J* = 6.3 Hz, 1H), 5.97 (d, *J* = 15.5 Hz, 1H), 7.32 (m, 1H), 7.96 (s, 1H), 8.23 (s, 1H), 9.59 (d, *J* = 7.2 Hz, 1H).

To a suspension of imidazolium bromide (2 g, 5.84 mmol) in deionized water (30 ml), LiTFSI (1.68 g, 5.84 mmol) in one portion was added, and the mixture was shaken vigorously in a separation funnel, followed by addition of dichloromethane into the abovementioned mixture. After being shaken 3–4 times, the lower layer was collected. The organics were dried with anhydrous sodium sulfate, filtered, and concentrated to give a light yellow oil (2.85 g, 5.256 mmol, 90%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 0.85 (s, 3H), 1.23 (m, 18H), 1.82 (s, 2H), 4.20 (s, 2H), 5.42 (d, *J* = 6.3 Hz, 1H), 5.97 (d, *J* = 15.5 Hz, 1H), 7.32 (m, 1H), 7.96 (s, 1H), 8.23 (s, 1H), 9.59 (d, *J* = 7.2 Hz, 1H).

**Synthesis of bromide 4.** To a solution of 1,12-dibromododecane (1 ml, 3.95 mmol, 2 equiv.) in DMF (5 ml) was added methacrylic acid (0.17 ml, 1.975 mmol, 1 eq.), cesium carbonate (0.643 g, 1.975 mmol, 1 equiv.) and hydroquinone (1 wt% of methacrylic acid). The resulting mixture was stirred at 32 °C for 48 h to give a clear solution. The solvents were removed with reduced pressure and the resulting residue was purified *via* flash column chromatography to give the title compound as a yellow oil (0.59 g, 1.77 mmol, 45%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 1.33 (m, 20H), 1.95 (s, 3H), 3.58 (t, *J*<sub>1</sub> = 6.72 Hz, *J*<sub>2</sub> = 6.71 Hz, 2H), 4.15 (t, *J*<sub>1</sub> = 6.5 Hz, *J*<sub>2</sub> = 6.6 Hz, 2H), 5.73 (d, *J* = 1.40 Hz, 1H), 6.08 (s, 1H).

**Synthesis of M2.** To bromide 4, (1 g, 2.99 mmol) 1-ethyl-1*H*-imidazole (0.34 ml, 2.99 mmol) was added, and the resulting mixture was stirred at 90 °C to give the imidazolium bromide as a white precipitate (0.985 g, 2.87 mmol, 96%), which was pure enough for the next step.

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 1.23 (m, 23H), 1.86 (s, 3H), 3.50 (s, 2H), 4.19 (dd, *J*<sub>1</sub> = 6.6 Hz, *J*<sub>2</sub> = 13.9 Hz, 2H), 5.66 (s, 1H), 5.99 (s, 1H), 7.85 (s, 3H), 9.34 (s, 1H).

To a suspension of the above imidazolium bromide (1 g, 2.912 mmol) in deionized water (30 ml) was added LiTFSI (0.842 g, 2.912 mmol) in one portion, and the mixture was shaken vigorously in a separation funnel followed by addition of

dichloromethane into the above mixture. After being shaken 3–4 times, the down layer was collected. The organics were dried with anhydrous sodium sulfate, filtered, and concentrated to give a light yellow oil (1.50 g, 2.76 mmol, 95%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 1.23 (m, 23H), 1.86 (s, 3H), 3.50 (s, 2H), 4.19 (dd, *J*<sub>1</sub> = 6.6 Hz, *J*<sub>2</sub> = 13.9 Hz, 2H), 5.66 (s, 1H), 5.99 (s, 1H), 7.85 (s, 3H), 9.34 (s, 1H).

**Synthesis of M3.** To a solution of alcohol 6 (1 ml, 5.11 mmol) in dichloromethane (30 ml) was added triethylamine (2.14 ml) and methanesulfonyl chloride (0.585 g, 5.11 mmol) dropwise at 0 °C. The resulting mixture was stirred at 0 °C for 45 min and then at room temperature for 10 h. The reaction was then quenched by adding HCl (0.1 M, 20 ml) and extracted with DCM (2 × 30 ml). The combined organics were dried over anhydrous sodium sulfate, filtered and then concentrated under reduced pressure to give the mesylate 7 as an oil (1.448 g, 5.059 mmol, 99%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 3.03 (s, 3H), 3.33 (s, 3H), 3.50 (m, 2H), 3.61 (m, 10H), 3.74 (m, 2H), 4.34 (m, 2H).

To the mesylate 7 (1 g, 3.492 mmol) was added 1-vinyl-1*H*-imidazole (0.34 ml, 3.492 mmol), the resulting mixture was stirred at 100 °C for 10 h. The resulting reaction mixture was washed with ethyl acetate to give the imidazolium mesylate 8 as an oil (1.288 g, 3.387 mmol, 97%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 3.00 (s, 3H), 3.55 (s, 3H), 3.75 (m, 2H), 3.85 (m, 10H), 4.11 (t, *J*<sub>1</sub> = 4.3 Hz, *J*<sub>2</sub> = 4.5 Hz, 2H), 4.80 (t, *J*<sub>1</sub> = 4.2 Hz, *J*<sub>2</sub> = 4.4 Hz, 2H), 5.60 (dd, *J*<sub>1</sub> = 8.6 Hz, *J*<sub>2</sub> = 2.5 Hz, 1H), 6.10 (dd, *J*<sub>1</sub> = 15.6 Hz, *J*<sub>2</sub> = 2.5 Hz, 1H), 7.5 (m, 1H), 7.9 (s, 1H), 8.0 (s, 1H), 10.3 (s, 1H).

To a suspension of mesylate 8 (1 g, 2.630 mmol) in deionized water (30 ml) was added LiTFSI (0.755 g, 2.630 mmol, 1 eq.), and the mixture was shaken vigorously in a separation funnel followed by dichloromethane addition into the above mixture. After being shaken 3–4 times, the bottom layer was collected. The organics were dried with anhydrous sodium sulfate, filtered, and concentrated to give M3 as a light yellow oil (1.487 g, 2.630 mmol, 100%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 3.24 (s, 3H), 3.44 (dd, *J*<sub>1</sub> = 5.7 Hz, *J*<sub>2</sub> = 3.4 Hz, 2H), 3.52 (dd, *J*<sub>1</sub> = 6.7 Hz, *J*<sub>2</sub> = 4.1 Hz, 10H), 3.83 (m, 2H), 4.4 (m, 2H), 5.45 (dd, *J*<sub>1</sub> = 8.7 Hz, *J*<sub>2</sub> = 2.3 Hz, 1H), 6.0 (dd, *J*<sub>1</sub> = 15.6 Hz, *J*<sub>2</sub> = 2.4 Hz, 1H), 7.3 (dd, *J*<sub>1</sub> = 15.6 Hz, *J*<sub>2</sub> = 8.7 Hz, 1H), 7.9 (s, 1H), 8.2 (s, 1H), 9.4 (s, 1H).

**Synthesis of M4.** To a solution of diol 9 (1 ml, 5.81 mmol) in dichloromethane (30 ml) was added triethylamine (2.43 ml) and methanesulfonyl chloride (1.33 g, 11.62 mmol) dropwise at 0 °C. The resulting reaction mixture was stirred at 0 °C for 45 min, and then at room temperature for 10 h. The reaction mixture was diluted with dichloromethane (30 ml) and washed with HCl (0.1 M, 30 ml). The organic layer was washed with pure water and then dried with anhydrous sodium sulfate. After elimination of solvents under reduced pressure, the dimesylate 10 was obtained as an oil (2.01 g, 5.75 mmol, 99%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 3.07 (d, *J* = 0.8 Hz, 3H), 3.65 (m, 8H), 3.76 (m, 4H), 4.37 (dd, *J*<sub>1</sub> = 3.70 Hz, *J*<sub>2</sub> = 4.70 Hz, 4H).

To a solution of dimesylate 10 (1 g, 2.853 mmol, 2 equiv.) in DMF (30 ml) was added methacrylic acid (0.13 ml, 1.427 mmol, 1 equiv.), cesium carbonate (0.465 g, 1.427 mmol, 1 equiv.) and hydroquinone (1 wt% of methacrylic acid). The resulting reaction



mixture was stirred at 32 °C for 48 h to give a clear solution. The solvents were removed under reduced pressure, and the residue was purified by flash column chromatography to give the title monoester 11 as a yellow oil (0.437 g, 1.283 mmol, yield 45%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 1.95 (m, 3H), 3.08 (d, 3H), 3.65 (m, 6H), 3.76 (m, 4H), 4.30 (m, 2H), 4.38 (m, 4H), 5.58 (m, 1H), 6.13 (dd,  $J_1 = 0.93$  Hz,  $J_2 = 1.44$  Hz, 1H).

To mesylate 11 (1 g, 2.936 mmol) was added 1-ethyl-1H-imidazole (0.33 ml, 2.936 mmol), and the resulting mixture was stirred at 100 °C for 10 h. The resulting reaction mixture was washed with ethyl acetate to give imidazolium mesylate 12 as an oil (1.217 g, 2.789 mmol, 95%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 1.57 (m, 3H), 1.94 (s, 3H), 2.75 (s, 3H), 3.64 (m, 12H), 4.32 (m, 4H), 4.52 (m, 2H), 5.59 (s, 1H), 6.11 (s, 1H), 7.34 (s, 1H), 7.65 (s, 1H), 9.77 (s, 1H).

To a suspension of imidazolium mesylate 12 (1 g, 2.291 mmol) in deionized water (30 ml) was added LiTFSI (0.658 g, 2.291 mmol 1 eq.), and the mixture was shaken vigorously in a separation funnel, followed by addition of dichloromethane into the abovementioned mixture. After being shaken 3–4 times, the bottom layer was collected. The organics were dried with anhydrous sodium sulfate, filtered, and concentrated to give M3 as a light yellow oil (1.351 g, 2.176 mmol, 95%).

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 1.44 (t,  $J_1 = 7.2$  Hz,  $J_2 = 6.2$  Hz, 3H), 1.89 (s, 3H), 3.63 (m, 12H), 4.23 (d,  $J = 5.9$  Hz, 4H), 4.35 (m, 2H), 5.70 (s, 1H), 6.04 (s, 1H), 7.76 (s, 1H), 7.81 (s, 1H), 9.14 (s, 1H).

**Synthesis of P1.** A mixture of M1 (1 g, 1.84 mmol) and AIBN (10 mg, 1 wt%) was heated at 90 °C under nitrogen for 1 h to give P1 as a rubber-like solid.

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ 0.80 (s, 3H), 1.23 (m, 18H), 1.82 (s, 2H), 4.20 (s, 2H), 7.32 (m, 1H), 7.96 (s, 1H), 8.23 (s, 1H), 9.59 (d,  $J = 7.2$  Hz, 1H).

**Synthesis of P2.** A mixture of M2 (1 g, 1.59 mmol) and AIBN (10 mg, 1 wt%) was heated at 90 °C under nitrogen for 1 h to give P2 as a rubber-like solid.

<sup>1</sup>H NMR (400 MHz, DMSO-D<sub>6</sub>): δ = 1.24 (m, 23H), 1.87 (s, 3H), 3.51 (s, 2H), 4.19 (dd,  $J_1 = 6.6$  Hz,  $J_2 = 13.9$  Hz, 2H), 7.79 (s, 3H), 9.17 (s, 1H).

## Conclusion

Ionic liquid monomer M1 to M4 and polymeric ionic liquid P1 to P2 were designed and synthesized in high yield. It is worth noting that the mesylate of ethylene oxide-containing alcohol could be synthesized in perfect yield and it could be transferred into imidazolium mesylate with high efficiency. M1 to M4 had ionic conductivity around  $1\text{--}2 \times 10^{-3}$  S cm<sup>-1</sup>. P2 showed ionic conductivity of  $2.2 \times 10^{-4}$  S cm<sup>-1</sup>, which was 200 times that of P1 due to the long alkyl chain tether separated polymer backbone and charge centre. Further systematic study of electrochemical properties of M1 to M4 and their corresponding PILs is ongoing and will be reported in due course.

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