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Formation of uniform PbS quantum dots by a spin-assisted successive precipitation and anion exchange reaction process using PbX_2 ($X = \text{Br}, \text{I}$) and Na_2S precursors†

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We devised a straightforward spin-assisted successive precipitation and anion exchange reaction (spin-SPAER) process in order to deposit relatively uniform PbS quantum dots (QDs) on mesoporous TiO_2 (mp- TiO_2). For the spin-SPAER process, we used PbX_2 ($X = \text{I}, \text{Br},$ and Cl) precursors instead of a $\text{Pb}(\text{NO}_3)_2$ precursor and consequently deposited individual PbS QDs on mp- TiO_2 due to the suppressed overgrowth of PbS QDs, whereas the conventional spin-assisted successive ionic layer adsorption and reaction (spin-SILAR) process formed aggregated PbS QDs on the mp- TiO_2 surface due to continuous adsorption and reaction. In addition, the PbS QDs prepared by spin-SPAER showed better air stability than the PbS QDs prepared by spin-SILAR possibly due to the passivation by halogen elements such as I and Br. Accordingly, we could improve the overall power conversion efficiency of PbS QD-SSCs prepared by the spin-SPAER process using PbI_2 and PbBr_2 precursors to $\sim 26.7\%$ and $\sim 44.2\%$, respectively, compared to the PbS QD-SSCs prepared by spin-SILAR using the $\text{Pb}(\text{NO}_3)_2$ precursor.

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Metal chalcogenide, such as CdSe ,¹ CdTe ,² PbS ,³ PbSe ,⁴ Sb_2S_3 ,⁵ Sb_2Se_3 ,⁶ $\text{Sb}_2\text{S}_3-x\text{Se}_x$,⁷ HgTe ,⁸ ClSe_2 ,⁹ and $\text{ClTe}_{2-x}\text{Se}_x$,¹⁰ semiconductor- or quantum dot (QD)-sensitizers have been considered as promising candidates replacing the conventional Ru/organic dye sensitizer because of their unique properties such as high absorptivity, convenient bandgap tailoring, easy charge separation by a large dipole moment, and solution processability.¹¹

Among the metal chalcogenides, PbS QDs have been extensively studied because they are an ideal model light absorber due to their large Bohr radius, small bulk bandgap energy, strong absorptivity, wide absorption spectrum from visible to near infrared, and multiple exciton generation. PbS QD solar cells can be roughly classified as Schottky,¹² depleted heterojunction,¹³ or sensitized type,¹⁴ with respect to the device architecture. Unlike Schottky and depleted heterojunction type devices, the sensitized solar cells have unique device architecture of electron conductor, sensitizer, and hole conductor,

which can promptly transfer charge carriers into electron conductor and hole conductor. Occasionally the depleted heterojunction PbS QD solar cells exhibited higher device efficiency than the PbS QD-sensitized solar cells (QD-SSCs), the sensitized solar cells still have advantages to attain high efficiency with low cost in terms of device architecture and operating mechanism so that continuous studies are needed.

The PbS QD-SSCs are generally fabricated by depositing the PbS colloidal QDs (CQDs) on a mesoporous TiO_2 (mp- TiO_2) electron conductor or directly growing the PbS QDs on the mp- TiO_2 via successive ionic layer adsorption and reaction (SILAR) method. The PbS CQDs have uniform size distribution so that they are good for injecting the electrons from PbS CQDs into mp- TiO_2 because their conduction band edge can be uniformly positioned over the conduction band edge of mp- TiO_2 . However, their surfaces are passivated by insulating long alkyl chains in order to make uniform sized PbS CQDs via hot-injection method. Therefore, the insulating long alkyl ligands are replaced to short ligands such as 1,2-ethane dithiol (EDT), 3-mercaptopropionic acid (MPA), and halogens to improve the charge transport at mp- TiO_2 /PbS CQDs interface and PbS CQDs bulk. For instance, Im *et al.* reported 2.9% multiple-layered PbS QD-SSCs constructed by FTO/bl- TiO_2 (blocking TiO_2)/mp- TiO_2 /PbS CQDs/P3HT (poly-3-hexylthiophene)/PEDOT:PSS (poly(3,4-ethylenedioxythiophene) polystyrene sulfonate)/Au.¹⁵ Seo *et al.* reported 3.2% mesoporous PbS QD/ $\text{CH}_3\text{NH}_3\text{PbI}_3$ core/shell-SSCs.¹⁶ Kim *et al.* reported 3.9% PbS QD-SSCs with radial

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directional charge transport by using TiO_2 nanorod electron conductor instead of mp- TiO_2 .¹⁷ Very recently, Park *et al.* reported 8.07% mesoporous PbS embedding CuS QD-SSCs.¹⁸

On the other hand, the PbS QDs prepared by SILAR method did not have any insulating passivation layer on their surface due to direct grown on the mp- TiO_2 . However, the SILAR method required tedious repeated process cycle: adsorption of cations (anions) \rightarrow washing \rightarrow drying \rightarrow adsorption of anions (cations) \rightarrow washing \rightarrow drying. In addition, the size of PbS QDs deposited on mp- TiO_2 surface is not uniform because some tiny PbS nanoparticles are nucleated and grown on the mp- TiO_2 and at the same time the pre-deposited PbS nanoparticles are continuously overgrown by the repeated SILAR cycle, thereby deteriorating charge injection from PbS QDs into mp- TiO_2 . To reduce the tedious repeated process cycle, Joo *et al.* devised a spin-assisted SILAR (spin-SILAR) method for fabricating a liquid-junction CdS-sensitized ZnO nanorod photoelectrochemical cell.¹⁹ Im *et al.* reported the solid-state PbS-sensitized photovoltaic photodetectors and PbS-SSCs prepared by spin-SILAR.²⁰ By using spin-coating process instead of dipping process, the washing and drying process could be eliminated because the excess chemicals are automatically spun out and the film is dried during spin-coating process. However, it is still difficult to control the size uniformity of PbS QDs deposited on mp- TiO_2 *via* spin-SILAR. Therefore, here we tried to deposit more uniform PbS QDs on mp- TiO_2 *via* spin-assisted successive precipitation and anion exchange reaction (spin-SPAER) because the over growth of the pre-deposited PbS QDs on mp- TiO_2 can be reduced.

Fig. 1 is a schematic illustration of spin-SPAER process. For convention spin-SILAR, 5 mM $\text{Pb}(\text{NO}_3)_2/\text{H}_2\text{O}/\text{methanol}$ ionic solution was used so that the Pb^{2+} cations are adsorbed on mp- TiO_2 surface and then Pb^{2+} cations are reacted with S^{2-} anions to form PbS. During the repeated spin-SILAR process, the ions can be adsorbed on both mp- TiO_2 and pre-deposited PbS QDs. Accordingly, the size of PbS nanoparticles formed on mp- TiO_2 surface tends to be broadened as illustrated in upper right inset in Fig. 1 due to the formation of new PbS nuclei and the continuous overgrowth of pre-deposited PbS QDs. On the other

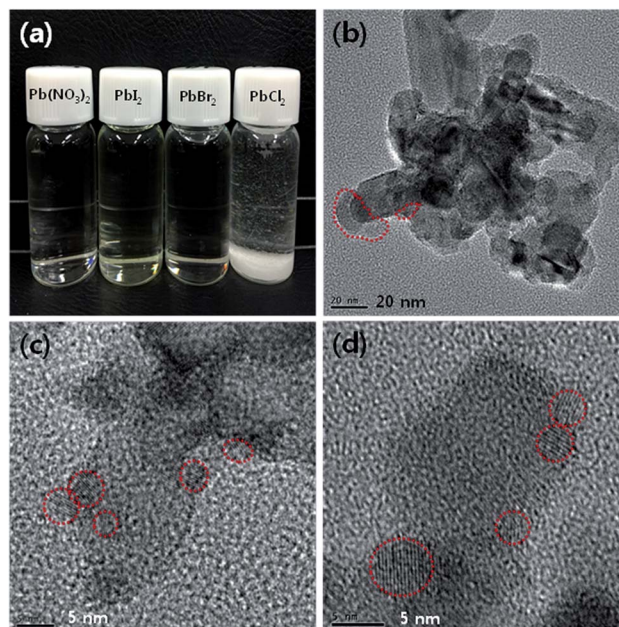


Fig. 2 (a) A photograph of 5 mM $\text{Pb}(\text{NO}_3)_2/\text{H}_2\text{O}/\text{methanol}$ ionic solution and PbX_2 ($X = \text{I}, \text{Br}, \text{and Cl}$)/DMF solutions and (b–d) TEM (transmission electron microscopy) images of PbS nanoparticles deposited on mp- TiO_2 , which are fabricated by 15 times repeated cycles of spin-SILAR process with 5 mM of $\text{Pb}(\text{NO}_3)_2/\text{H}_2\text{O}/\text{methanol}$ ionic solution (b) and spin-SPAER process with 5 mM of PbI_2/DMF (c), and PbBr_2/DMF solution (d), respectively. Red dotted region = PbS nanoparticles.

hands, for spin-SPAER method 5 mM PbX_2 ($X = \text{I}, \text{Br}, \text{and Cl}$)/DMF solution instead of 5 mM $\text{Pb}(\text{NO}_3)_2/\text{H}_2\text{O}/\text{methanol}$ ionic solution was spin-coated at 3000 rpm for 60 s. During the spin-coating process, the PbX_2 nanoparticles are precipitated on the surface of mp- TiO_2 (process I) as illustrated in Fig. 1. Then the PbX_2 is transformed into PbS by anion exchange reaction with 5 mM $\text{Na}_2\text{S}/\text{H}_2\text{O}/\text{methanol}$ solution during second stage spin-coating process (process II) at 3000 rpm for 60 s. To prevent the deposition/precipitation of PbX_2 on pre-formed PbS on mp- TiO_2 , 1 wt% 1,2 ethanedithiol (EDT)/ethanol solution was

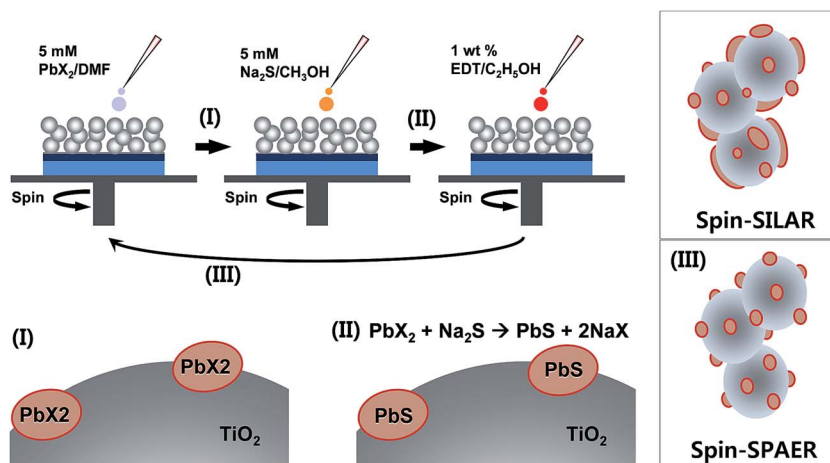


Fig. 1 Schematic illustration of spin-assisted successive precipitation and anion exchange reaction (spin-SPAER) process.



spin-coated at 3000 rpm for 60 s. By the repetition of the spin-SPAER process (process III), we expect that the individual PbS QDs will be formed on mp-TiO₂ as illustrated in lower right inset in Fig. 1 because the PbX₂ precipitates cannot adsorb on the pre-formed PbS QDs unlike to Pb²⁺ cations.

Fig. 2(a) is a photograph of 5 mM Pb(NO₃)₂/H₂O/methanol ionic solution and PbX₂ (X = I, Br, and Cl)/DMF solutions showing that Pb(NO₃)₂, PbI₂, PbBr₂ are well dissolved in H₂O/methanol solution and DMF, respectively, whereas the 5 mM PbCl₂ is precipitated in DMF due to poor solubility. The order of solubility of PbI₂, PbBr₂, and PbCl₂ toward DMF solvent was PbI₂ > PbBr₂ > PbCl₂. Therefore, we excluded the PbCl₂/DMF solution for further model experiments. Fig. 2(b)–(d) are TEM (transmission electron microscopy) images of PbS nanoparticles deposited on mp-TiO₂, which are fabricated by 15 times repeated cycles of spin-SPAER process with 5 mM of Pb(NO₃)₂/H₂O/methanol ionic solution, PbI₂/DMF, and PbBr₂/DMF solution, respectively. The PbS nanoparticles deposited on mp-TiO₂ from 5 mM of Pb(NO₃)₂/H₂O/methanol ionic solution had polydispersed size (see Fig. 2(b)). Most PbS nanoparticles had below 5 nm but they are aggregated on mp-TiO₂ and some PbS nanoparticles had over 5 nm in size. Accordingly the energetic electron injection from PbS nanoparticles into mp-

TiO₂ will be deteriorated by relatively large PbS nanoparticles because the conduction band edge of PbS nanoparticles with larger size become lower than the conduction band edge of mp-TiO₂.²¹ On the other hands, the PbS QDs prepared from 5 mM PbI₂/DMF and PbBr₂/DMF solution showed smaller PbS nanoparticles with 3–5 nm in size (see Fig. 2(c and d)) and over grown PbS nanoparticles were not detected as shown in Fig. S1.† Apparently, the size PbS QDs deposited on mp-TiO₂ from PbBr₂/DMF solution seems to be slightly smaller than that from PbI₂/DMF solution, which might be attributed to the quicker precipitation (nucleation) of PbBr₂ than PbI₂ during spin-coating process. It should be noted that the individual PbS QDs without over growth are formed on mp-TiO₂ *via* spin-SPAER process as we proposed.

Fig. 3(a) is a representative SEM (scanning electron microscopy) cross-sectional image of PbS QD-SSC of which the PbS QDs are prepared by 15 times repeated cycles of spin-SPAER process with 5 mM of PbBr₂/DMF solution. The PbS QD-SSC was constructed by FTO/bl-TiO₂ (~50 nm in thickness)/mp-TiO₂ (~600 nm in thickness)/PbS QDs/P3HT (~30 nm in thickness)/Au (~60 nm in thickness). An energy band diagram of PbS QD-SSC was shown in Fig. 3(b). Upon illumination of light, the PbS QDs generate electron-hole pairs and the

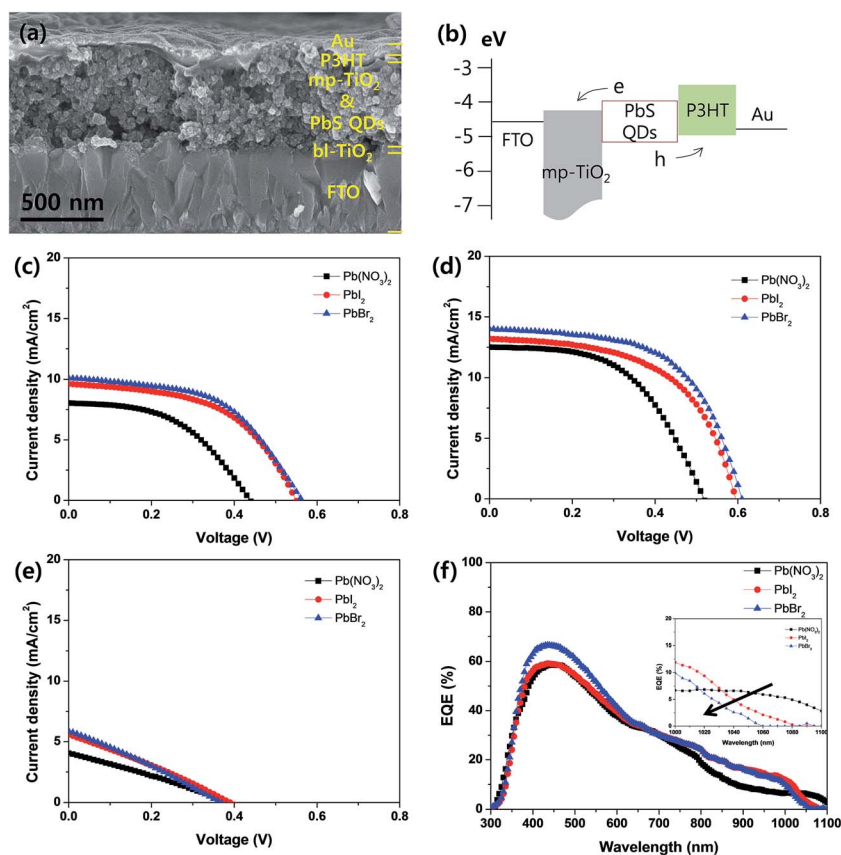


Fig. 3 (a) SEM (scanning electron microscopy) cross-sectional image of representative PbS QD-SSC composed to FTO/bl-TiO₂/mp-TiO₂/PbS QDs/P3HT/Au: PbS QDs are formed by 15 times repeated spin-SPAER process with 5 mM PbBr₂/DMF solution, (b) schematic energy band diagram, (c and d) current density–voltage (*J*–*V*) curves of PbS QD-SSCs with repeated coating cycles to (c) 10, (d) 15, and (e) 20 times of spin-SILAR (Pb(NO₃)₂) or spin-SPAER (PbI₂, PbBr₂) process, and (f) their corresponding EQE (external quantum efficiency) spectra: inset = zoom in at long wavelength region.



Table 1 Summary of photovoltaic properties of PbS QD-SSCs prepared by conventional spin-SILAR ($\text{Pb}(\text{NO}_3)_2$) and newly proposed spin-SPAER process (PbI_2 and PbBr_2)

Repeated cycles	Device	V_{oc} (V)	J_{sc} (mA cm^{-2})	FF (%)	η (%)
10	$\text{Pb}(\text{NO}_3)_2$	0.43	8.0	50	1.72
	PbI_2	0.54	9.6	53	2.75
	PbBr_2	0.56	10.1	52	2.94
15	$\text{Pb}(\text{NO}_3)_2$	0.51	12.5	54	3.44
	PbI_2	0.59	13.2	56	4.36
	PbBr_2	0.60	14.0	59	4.96
20	$\text{Pb}(\text{NO}_3)_2$	0.38	4.1	29	0.45
	PbI_2	0.39	5.6	28	0.61
	PbBr_2	0.36	5.9	28	0.59

electrons (holes) are promptly transfer/transported into mp-TiO₂ (P3HT), respectively. If the size of PbS QDs is fluctuated, the conduction band edge of PbS QDs is fluctuated so that the driving force of electron injection from PbS QDs into mp-TiO₂ is dependent on the fluctuation. Fig. 3(c–e) compared the current density–voltage (J – V) curves of PbS QD-SSCs prepared by conventional spin-SILAR and newly proposed spin-SPAER process. Their photovoltaic properties were summarized in Table 1. To optimize efficiency of mode devices and check the tendency of the effect of spin-SPAER process, we fabricated PbS QD-SSCs with different repeated cycles for the formation of PbS QDs. For the model experiments, we controlled the repeated cycles of spin-SILAR and spin-SPAER to 10 (Fig. 3(c)), 15 (Fig. 3(d)), and 20 times (Fig. 3(e)), respectively. Irrespective to the repeated cycles for the formation of PbS QDs, the PbS QD-SSCs prepared by spin-SPAER process showed better device performance than that prepared by spin-SILAR process. The improved performance of PbS QD-SSCs prepared by spin-SPAER

process is attributed to the formation of more uniform PbS-QDs on mp-TiO₂. Eventually the PbS QD-SSCs prepared by 15 times repeated cycles of spin-SILAR and spin-SPAER process exhibited the best device performance. The $\text{Pb}(\text{NO}_3)_2$ sample prepared by spin-SILAR process showed 0.51 V open-circuit voltage (V_{oc}), 12.5 mA cm^{-2} short-circuit current density (J_{sc}), 54% fill factor (FF), and 3.44% power conversion efficiency (η). The PbI_2 and PbBr_2 samples prepared by spin-SPAER process showed 0.59 V V_{oc} , 13.2 mA cm^{-2} J_{sc} , 56% FF, and 4.36% η and 0.60 V V_{oc} , 14.0 mA cm^{-2} J_{sc} , 59% FF, and 4.96% η at 1 Sun condition, respectively. Therefore, the best PbBr_2 sample had improvement of 18% V_{oc} , 12% J_{sc} , 9% FF, and 44% η compared to the $\text{Pb}(\text{NO}_3)_2$ device (see Table 1). The significant degradation of the performance in the PbS QD-SSCs with repeated coating cycles to 20 might be attributed to the formation of larger particles and the clogging of mesopores because the charge injection from large PbS sensitizer into mesoporous TiO₂ will be significantly deteriorated by the mismatch of energy band and the hole extraction efficiency will be also greatly deteriorated by the ineffective infiltration of P3HT hole transporting material with the mesopores in mesoporous TiO₂ film. Fig. 3(f) is EQE (external quantum efficiency) spectra of best devices (Fig. 3(d) samples). The calculated J_{sc} values from the integration of the EQE spectra of $\text{Pb}(\text{NO}_3)_2$, PbI_2 , and PbBr_2 sample was 12.1, 13.0, and 13.7 mA cm^{-2} , respectively so that these values were well matched to the measured J_{sc} values in J – V curves. The inset in Fig. 3(f) is the zoomed in EQE spectra around the band edge indicates that the onset wavelength of EQE spectrum is gradually blue-shifted in the order to $\text{Pb}(\text{NO}_3)_2$, PbI_2 , and PbBr_2 sample. This implies that the average size of PbS QDs deposited on mp-TiO₂ is gradually decreased in the order to $\text{Pb}(\text{NO}_3)_2$, PbI_2 , and PbBr_2 sample. This result is consistent with the TEM analysis in Fig. 2(b–d).

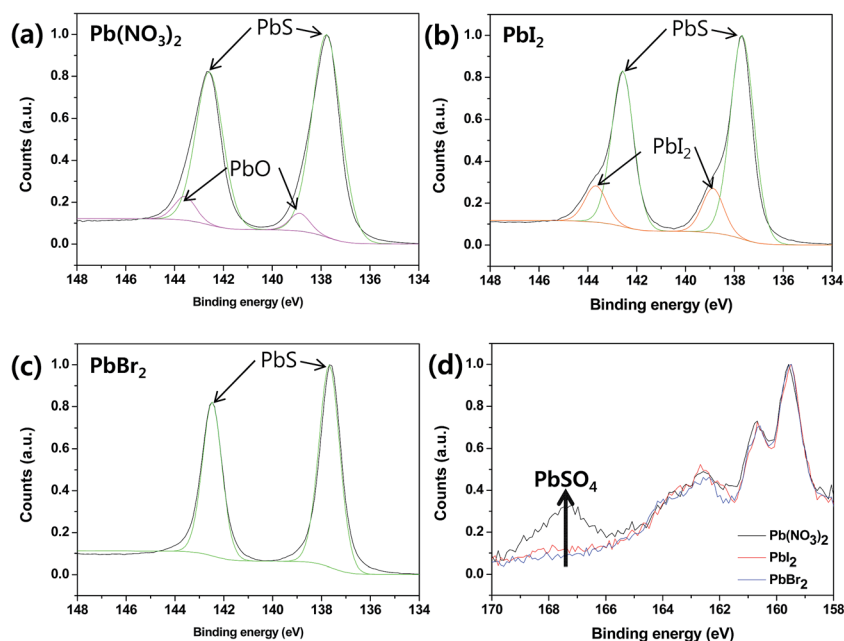


Fig. 4 XPS (X-ray photoelectron spectroscopy) spectra Pb_{2p} (a–c) and S_{2p} (d) peaks of PbS prepared by (a) $\text{Pb}(\text{NO}_3)_2$, (b) PbI_2 , and (c) PbBr_2 .



The electronic bandgap (E_g) of PbS QDs with a diameter (d) can be expressed by following equation.²²

$$E_g = 0.41 + 1/(0.025d^2 + 0.283d)$$

From above equation, the calculated average diameter of Pb(NO₃)₂, PbI₂, and PbBr₂ sample was over 3.7 nm, ~3.6 nm, and ~3.5 nm, respectively. The EQE is a product of light harvesting efficiency (η_{the}), charge separation efficiency (η_{cs}), and charge collection efficiency (η_{cc}). The absorption spectra of Pb(NO₃)₂, PbI₂, and PbBr₂ sample did not have significant difference (see Fig. S2†). Accordingly, the EQE improvement by newly designed spin-SPAER process was mainly ascribed to the improved charge separation efficiency (η_{cs}) and charge collection efficiency (η_{cc}).

In addition, the PbS QD-SSCs prepared by spin-SPAER process was much better air stability than the PbS QD-SSC prepared by spin-SILAR process. Therefore, we checked the XPS (X-ray photoelectron spectroscopy) spectra of PbS QDs prepared by 15 times repeated cycles of spin-SILAR (Pb(NO₃)₂ sample) and spin-SPAER process (PbI₂ and PbBr₂ sample) after 1 week storage in air in order to check the purity and stability of PbS QDs as shown in Fig. 4. The PbS QDs prepared by spin-SILAR process (Pb(NO₃)₂ sample) exhibited PbO peak from the analysis of Pb_{2p} peak analysis, whereas the PbS QDs prepared by spin-SPAER process (PbI₂ and PbBr₂ sample) did not show oxide peaks. In contrast, the PbS QDs prepared by spin-SPAER process using PbI₂ precursor exhibited PbI₂ peaks, whereas the PbS QDs prepared by using PbBr₂ precursor did not have PbBr₂ peaks. To distinguish the PbO peaks and PbX₂ (X = I, Br) peaks, we checked X_{3d} peaks as shown in Fig. S3† because the binding energy of PbO and PbX₂ in Pb_{2p} peaks is very close. Fig. S3† clearly confirms that the small quantity of PbI₂ was remained during spin-SPAER process, whereas the PbBr₂ impurity peaks were not detectable due to fully conversion into PbS QDs. The S_{2p} peaks in Fig. 4(d) confirm that the PbS QDs prepared by spin-SILAR process were oxidized to PbO and PbSO₄, whereas the PbS QDs prepared by spin-SPAER were not oxidized. From these results, we can conclude that more efficient and stable PbS QD-SSCs could be fabricated by newly proposed spin-SPAER process.

Conclusions

We could deposit relatively uniform PbS QDs *via* newly proposed spin-SPAER process whereas the conventional spin-SILAR process tends to form aggregated PbS QDs on mp-TiO₂ surface. The formation of individual PbS QDs on mp-TiO₂ by spin-SPAER might be attributed to the formation of nanoprecipitates of PbX₂ (X = I, Br) so that the continuous over growth of PbS nanoparticles was suppressed. The solubility of PbI₂ in DMF solvent was higher than the PbBr₂ so that the size of nuclei of PbI₂ formed during spin-coating process will be bigger than the that of PbBr₂. Accordingly, the larger PbS QDs was formed by PbI₂ precursor than the PbBr₂. The EQE spectra of PbS QD-SSCs prepared by Pb(NO₃)₂, PbI₂, and PbBr₂

precursor indicated that the size of PbS QDs was gradually increased in the order to PbBr₂ < PbI₂ < Pb(NO₃)₂. Therefore, the overall power conversion efficiency of PbS QD-SSCs prepared by spin-SPAER process using PbI₂ and PbBr₂ precursor could be enhanced to ~26.7% and ~44.2%, respectively, compared to the PbS QD-SSCs prepared by spin-SILAR using Pb(NO₃)₂ precursor. From the XPS analysis, we could find that the PbS QD-SSCs prepared by spin-SPAER have better stability than the device prepared spin-SILAR because the surface of Pb or S in PbS QDs prepared by spin-SILAR is more easily oxidized into PbO or PbSO₄ whereas the PbS QDs prepared by spin-SPAER are not easily oxidized possibly due to the passivation by I or Br element.

Acknowledgements

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