

CrossMark
click for updatesCite this: *RSC Adv.*, 2017, 7, 1026Received 21st October 2016
Accepted 28th November 2016

DOI: 10.1039/c6ra25594a

www.rsc.org/advances

Catalytic transfer hydrogenation of levulinic acid to γ -valerolactone over a bifunctional tin catalyst†

Shaodan Xu,^{*ab} Deqing Yu,^a Tao Ye^{ac} and Panpan Tian^a

The conversion of biomass-derived levulinic acid (LA) to γ -valerolactone (GVL) is important but extremely challenging due to the use of noble metal catalysts and high-pressure gaseous hydrogen. Here we describe the transfer hydrogenation of LA to GVL using 2-propanol as hydrogen donor over a Sn modified silica catalyst (SnO₂/SBA-15) which gives high LA conversion (85.1%) and GVL selectivity (95.2%). Key to this success is employing a simply synthesized Sn-modified silica catalyst (SnO₂/SBA-15) with both Lewis and Brønsted acidity, which could synergistically catalyze the tandem steps of hydrogen transfer and esterification, leading to successful transformation of LA to GVL.

Introduction

Biomass, which is regarded as a renewable feedstock, has attracted much attention for the production of fine chemicals and fuels.^{1–5} Many economically viable processes have been developed for the conversion of biomass.^{2,6} One efficient route is the fast pyrolysis and/or acidolysis of the cellulose and lignin feedstocks to various platform molecules, and then transformation of the platform molecules to the desired fuels and chemicals.^{7–12} However, the biomass platform molecules have high oxygen content and various functional groups, which makes the transformation route complex and nonselective.¹ Therefore, developing an efficient strategy to efficiently transform the biomass platform molecules is strongly desirable for producing high-quality fuels and chemicals.

Levulinic acid (LA), which can be easily produced from cellulosic feedstocks by acidolysis, has been recognized as one of the top-10 most promising platform molecules derived from biomass by the U.S. Department of Energy.¹³ One of the most important use of LA is to produce γ -valerolactone (GVL), which has been used as a high-quality solvent and fuel additive.^{14–20} Recently, many catalysts, such as homogeneous Pd complex,²¹ heterogeneous Ru/C,^{22,23} RuNi/C,²⁴ Au/ZrO₂,^{25,26} Ru/graphene,²⁷ have successfully catalysed the transformation of LA to GVL by a hydrogenation and subsequent dehydration steps.^{28–34} However, it should be noted that noble metals and gaseous hydrogen are necessary in most of these cases, which

make their further application difficult due to the pricy cost of noble metals and hazardous gaseous hydrogen. Therefore, efficient production of GVL from LA without using noble metal catalysts and gaseous hydrogen is still challenging. Recently, transfer hydrogenation of LA using the organic molecules (*e.g.* formic acid, 2-propanol) as hydrogen sources have been extensively studied, which is regarded to be potentially important for LA conversion. In the transfer hydrogenation of LA, the Zr- and Ni-based catalysts have attracted much attention because of their non-noble feature and high efficiency.^{35–38} In these cases, the metal atoms were attached into the zeolite or organic polymer framework to form Lewis acid sites, which benefit the transfer hydrogenation reaction. Considering the importance of LA conversion, developing more efficient catalysts for this process is still necessary.

In the present work, we reported a Sn modified silica catalyst (SnO₂/SBA-15), which is simply synthesized by grafting Sn sites on the surface of silica, could efficiently catalyse the transfer hydrogenation of LA to GVL using 2-propanol as hydrogen donor. For SnO₂/SBA-15 catalyst, the Sn sites are uniformly distributed on the SiO₂ support and exhibited both Lewis and Brønsted acidity. It is observed that the Lewis and Brønsted acidity are benefit for the tandem reactions of transfer hydrogenation and dehydration, respectively, giving GVL yield at as high as 81.0%. Importantly, the SnO₂/SBA-15 catalyst is very stable and recyclable, showing constant GVL yield in long reaction period in a fixed-bed reactor.

Experimental

Sample preparation

All chemicals were commercially available and directly used without further purification. The SBA-15 support was synthesized according to the method described in literature.³⁹

^aZhejiang Institute of Quality Inspection Science, Hangzhou 310018, China. E-mail: xushaodan6677@126.com

^bDepartment of Chemistry, Zhejiang University, Hangzhou 310027, China

^cMOE Key Laboratory of Macromolecular Synthesis and Functionalization, Department of Polymer Science and Engineering, Zhejiang University, Hangzhou 310027, China

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c6ra25594a



Synthesis of SnO₂/SBA-15 and HL-SnO₂/SBA-15 (high-loading SnO₂/SBA-15)

5.0 g of SBA-15 was dried in vacuum at 80 °C for 24 h, and then dispersed in 150 ml of *p*-xylene. After ultrasonic treatment for 1 h, a desired amount of dimethyldichlorostannane and 22 ml of triethylamine was added, and the mixture was continuously stirred for another 10 h at room temperature. After washing with *p*-xylene and ethanol, and calcining at 600 °C for 2 h in flowing oxygen (5% in N₂), the SBA-15 with grafted Sn species could be obtained. By adjusting the amount of dimethyldichlorostannane in the starting solution, the SnO₂/SBA-15 and HL-SnO₂/SBA-15 with different Si/Sn ratio could be obtained. The SnO₂/MCM-41 catalyst was synthesized in the same method as synthesizing SnO₂/SBA-15 except the use of MCM-41 as support instead of SBA-15.

Sample characterization

Nitrogen sorption isotherms were measured using a Micromeritics ASAP 2020M system. The samples were outgassed for 10 h at 200 °C before the measurements. The pore-size distribution was calculated using the Barrett–Joyner–Halenda (BJH) model. Titan ChemiSTEM and FEI Tecnai G2 F20 microscopes equipped with an EDAX detector were employed to acquire TEM, HAADF-STEM images, and energy dispersive X-ray spectrum (EDS) elemental maps. Powder X-ray diffraction patterns (XRD) were obtained with a Rigaku D/MAX 2550 diffractometer with CuK α radiation ($\lambda = 0.1542$ nm). The FTIR spectra of pyridine-adsorption were recorded on a Thermo-Fisher FTIR spectrometer. 15 mg of sample in a cell was used for the pyridine-adsorption IR test. The pyridine vapor was admitted into the cell and adsorption lasted for 1 h. Subsequently, the samples were heated up to 220 °C under vacuum and cooled down to RT for collecting the signal.

Catalytic tests

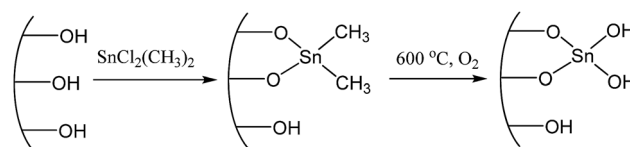
The reactions were performed in a 50 ml high-pressure autoclave with a magnetic stirrer. The substrate, catalyst, and solvent were mixed in the reactor, then pure N₂ was introduced to remove the air in autoclave and keep the N₂ pressure at 0.7–0.8 MPa. Then, the autoclave was heated to a given temperature. After the reaction, the product was taken out from the reaction system and analyzed by gas chromatography (flame ionization detector) with a flexible quartz capillary column (100 m in length) coated with FFAP. The recyclability of the catalyst was tested by separating the catalyst solids from the reaction system by centrifugation, washing with large quantity of methanol and water, drying at 60 °C, then the catalyst was reused in the next reaction. The reactions in the fixed-bed reactor were carried in a 10 ml quartz tube with an oven to control the temperature. The catalysts were made into 1–2 mm particles and mixed with quartz particles with similar sizes, then the mixture was localized in the middle of the tube. The liquid substrates were injected into the reactor by a pump, and N₂ flowed in the reactor during the whole reaction process.

Results and discussion

It is well known that ultra-small metal nanoclusters (*e.g.* Al, Zr, Sn) could create Brønsted acidic sites from the surface hydroxyl groups.⁴⁰ Therefore, it is possible to synthesize a metal catalyst with both Lewis and Brønsted acidity by loading uniformly dispersed metal species on an inert silica support. In this case, the mesoporous SBA-15 was employed as silica support because of the obvious advantage of the large surface area, which is expected to be favorable for the dispersion of acid sites as well as the diffusion of reactants. The Sn species are loaded on the SBA-15 support by a grafting method,^{40,41} where the dimethyldichlorostannane was grafted the mesoporous silica of SBA-15 by a condensation reaction with the silanol groups, followed by calcination to form hydroxyl group on Sn species (Scheme 1). The obtained SnO₂/SBA-15 has a Si/Sn ratio at 53, as determined by inductively coupled plasma optical emission spectrometry (ICP-OES) (Table 1).

Fig. 1A shows the small-angle X-ray diffraction (XRD) pattern of SnO₂/SBA-15. In the region at 0.6–2°, three peaks obviously appeared indexing to the (100), (110) and (200) reflections of the hexagonal mesoporous arrays.³⁹ Notably, the peaks assigning to SnO₂ crystals at ~34 and ~52° are completely absent in the wide-angle XRD pattern of SnO₂/SBA-15 (Fig. 1B), which might be due to the low metal loading. Fig. 1C shows the nitrogen sorption isotherm of SnO₂/SBA-15, which exhibit typical IV curve with a hysteresis loop, assigning to typical mesoporous structure. Additionally, the SnO₂/SBA-15 exhibits a surface area at 359 cm² g⁻¹ and uniform pore diameter distribution at 6.5 nm (Fig. 1D, Table 1), which are slightly smaller than those of the as-synthesized SBA-15 (Fig. S1†). This phenomenon might be due to the successful grafting of Sn species in the mesopores of SBA-15.

Electronic microscopy images give direct observation of the SnO₂/SBA-15. Fig. 2a shows the high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) by the Z-contrast, which exhibits high-quality mesoporous structure, in good agreement with the results of XRD pattern and N₂



Scheme 1 Synthesis of SnO₂/SBA-15 by grafting Sn species on the SBA-15 support.

Table 1 Structural parameters of various samples

Sample	Si/Sn ratio ^a	<i>S</i> _{BET} (m ² g ⁻¹)	Pore width (nm)
SBA-15	—	565	7.5
SnO ₂ /SBA-15	53	359	6.5
HL-SnO ₂ /SBA-15	13	320	6.3

^a By ICP analysis.



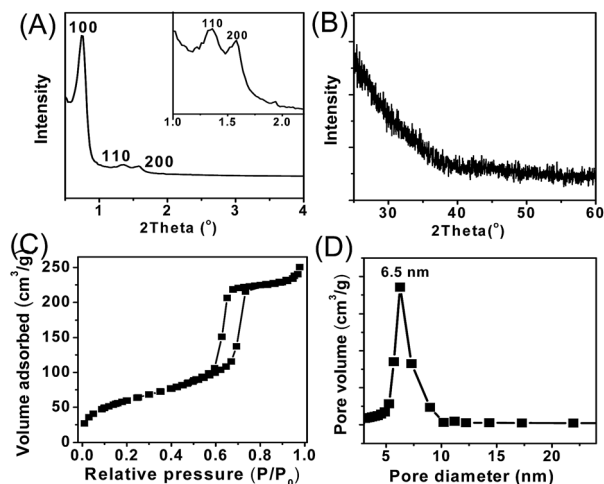


Fig. 1 (A) Small-angle and (B) wide angle XRD patterns, (C) N_2 sorption isotherms, and (D) mesopore size distribution of $SnO_2/SBA-15$.

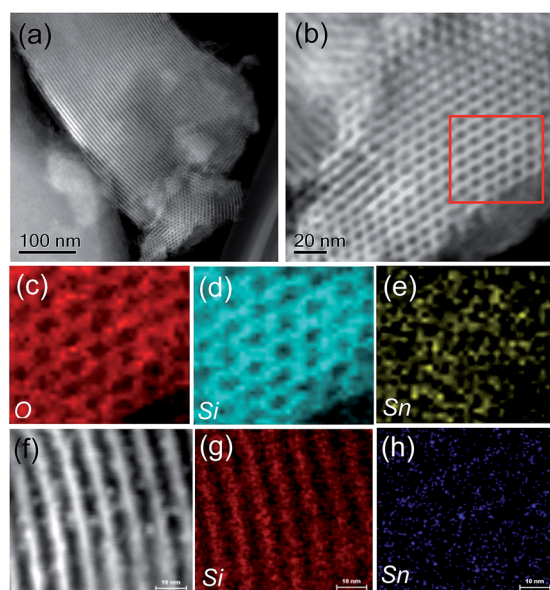


Fig. 2 (a and b) STEM images of $SnO_2/SBA-15$; (c) O, (d) Si and (e) Sn elemental maps of the red square in (b). (f) STEM image and (g) Si and (h) Sn EDS elemental maps of $HL-SnO_2/SBA-15$.

sorption isotherms. Notably, it is difficult to directly observe the Sn species in the high-resolution STEM image (Fig. 2b), which might be due to the uniform dispersion of Sn species on the surface of SBA-15. The energy dispersive X-ray (EDX) mapping analysis gives clear observation of the distribution of Sn element. As shown in Fig. 2c–e, the Sn element has similar distribution to Si and O elements. Furthermore, energy dispersive spectra of randomly selected regions on $SnO_2/SBA-15$ give similar signals (Fig. S2†). These results indicate the uniform dispersion of Sn on $SnO_2/SBA-15$ sample. For comparison, Fig. 2f–h show the electron microscopy characterization of $HL-SnO_2/SBA-15$. Clearly, many nanoparticles could be obviously observed in the mesopores of the SBA-15,

assigning to the aggregated Sn species because of high Sn loading.

The acidity of $SnO_2/SBA-15$ and $HL-SnO_2/SBA-15$ was studied by pyridine-adsorption IR spectra. As shown in Fig. 3, $SnO_2/SBA-15$ gave the obvious bands at 1453 and 1614 cm^{-1} , which are assigned to the coordinative bound of pyridine on Lewis acid sites.^{42–44} Interestingly, the spectrum of $SnO_2/SBA-15$ also gives a band at 1531 cm^{-1} , which is characteristic of pyridine adsorbed on the Brønsted acid sites. The band at 1491 cm^{-1} is from the contribution of pyridine adsorbed on both Lewis and Brønsted acid sites. These results indicate that $SnO_2/SBA-15$ has both Lewis and Brønsted acid sites, which should origin from the Sn atoms and silanol groups, respectively.^{31,45} In contrast, the $HL-SnO_2/SBA-15$ exhibited very weak band at 1531 cm^{-1} , indicating the sparsely Brønsted acid on $HL-SnO_2/SBA-15$, which might be due to the aggregation of Sn species could eliminate most of the Sn–OH groups. Furthermore, the band of pyridine on the Lewis acid sites of $HL-SnO_2/SBA-15$ appeared at 1446 cm^{-1} , with a red shift of 7 cm^{-1} from that of $SnO_2/SBA-15$ (1453 cm^{-1}). The red shift in IR band suggests the weaker pyridine–Lewis acid interaction on $HL-SnO_2/SBA-15$ than $SnO_2/SBA-15$. Similar phenomenon could also be observed in the band at 1491 cm^{-1} on $SnO_2/SBA-15$, which shifted to 1480 cm^{-1} on $HL-SnO_2/SBA-15$. These results indicate the significant advantages of $SnO_2/SBA-15$ in both Lewis and Brønsted acidity compared with $HL-SnO_2/SBA-15$.⁴⁶

The catalytic study starts from the conversion of LA to GVL in 2-propanol solvent. As summarized in Table 2, SBA-15 support without any Sn species is inactive for the formation of GVL (entry 1). Interestingly, when $SnO_2/SBA-15$ was used as catalyst, the conversion of LA could reach 85.1% with GVL selectivity at as high as 95.2% (entry 2). In contrast, the $HL-SnO_2/SiO_2$, which has the same individuals of Sn species and SiO_2 support to $SnO_2/SBA-15$ catalyst, exhibits much lower LA conversion at only 4.9% (entry 3). Additionally, we also tested the catalytic performances of various transition metal oxides with Lewis acidity, including Al_2O_3 , SnO_2 , ZrO_2 , which also give low LA conversion (1.9–14.4%) and GVL selectivity (<42%) (entries 4–6). These results indicate that the catalysts with only Lewis acidity is active for the transfer hydrogenation of LA to GVL, but $SnO_2/SBA-15$ with strong Lewis and Brønsted acidity is more efficient.

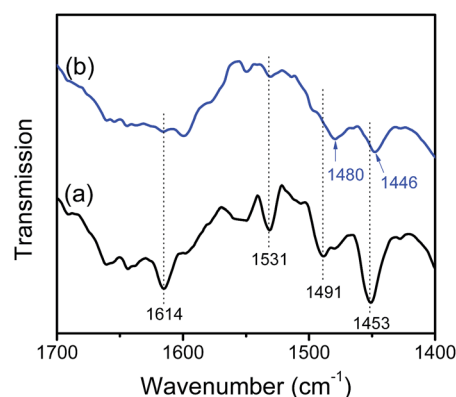


Fig. 3 Pyridine adsorption IR spectra of (a) $SnO_2/SBA-15$ and (b) $HL-SnO_2/SBA-15$.

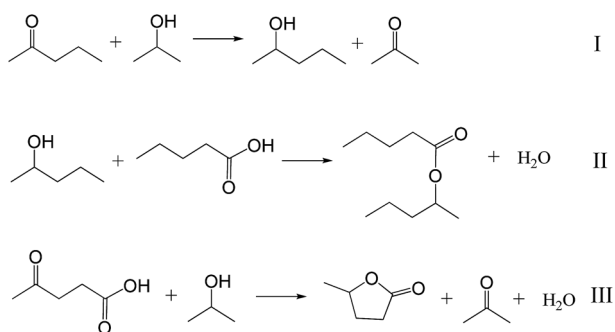


Table 2 Catalytic data in conversion of LA to GVL using 2-propanol over various catalysts^a

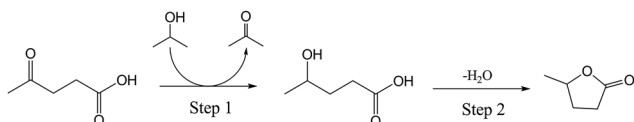
Entry	Catalyst	LA conv. (%)	GVL Sel. (%)
1	SBA-15	— ^b	— ^b
2	SnO ₂ /SBA-15	85.1	95.2 ^c
3	HL-SnO ₂ /SBA-15	4.9	96.5
4	Al ₂ O ₃	1.9	— ^b
5	SnO ₂	3.9	— ^b
6	ZrO ₂	14.4	42.1

^a Reaction conditions: 3 mmol of LA, 50 mmol of 2-propanol, 140 mg of catalyst, 110 °C, 8 h. ^b Undetectable. ^c The by-products are mainly the lipid from the esterification of acid with 2-propanol.

According to the well-known Meerwein–Ponndorf–Verley reduction using 2-propanol as hydrogen donor over Lewis acid sites,⁴⁷ it is reasonable to assume the conversion of LA to GVL should process the tandem steps of (I) transfer hydrogenation from 2-propanol to reduce the C=O to C–OH group on LA and (II) esterification of the C–OH and COOH groups to form GVL (Scheme 2). On the basis of the proposed reaction process, we discussed the reasons why SnO₂/SBA-15 shows better catalytic performances than the HL-SnO₂/SBA-15 by studying the reaction rate of each individual step, because the reaction rates in the beginning of the reactions at low conversion are sensitive to the activities of the catalysts. In order to understand the performances of the catalysts in each of the steps for transfer hydrogenation and esterification in LA-to-GVL reaction (Scheme 3), we employ different substrates to split the LA-to-GVL process into two individual reactions, as presented in Scheme 2, which are (I) transfer hydrogenation of C=O in 2-pentanone with 2-propanol to C–OH and (II) esterification of 2-



Scheme 2 The model reactions of (I) hydrogen transfer from 2-propanol to reduce the C=O group, (II) esterification of C–OH and COOH groups, and (III) transformation of LA to GVL using 2-pentanol as hydrogen donor.



Scheme 3 The two steps in LA-to-GVL process.

pentanol with *n*-pentanoic acid to ester. This method has been widely used to study the contribution of catalysts in the transformation with multiple steps,⁴⁸ which benefit the investigation on the catalytic activity of the catalyst in the individual steps, thus to understand the rate control step and the origin of high activity. Table 3 lists the reaction rates of the Sn-based catalysts in the individual reactions of (I) hydrogen transfer from 2-propanol to reduce the C=O group and (II) esterification of C–OH and COOH groups and the combined reaction of (III) transformation of LA to GVL using 2-propanol as hydrogen donor (Scheme 2). The reaction rates are obtained by the catalytic conversion at the beginning of each reaction within a short time of 20 min. It is worth emphasizing that the intramolecular esterification to form GVL in the LA-to-GVL process could occur even in the absence of Brønsted acid catalyst,^{37,38} here the intermolecular esterification (reaction II) was employed to slow down the reaction rate for clarifying the different activities of various catalysts. Clearly, the SBA-15 without Sn species failed to catalyze all the reactions (entry 1 in Table 3). Interestingly, SnO₂/SBA-15 is very active for both the hydrogen transfer (reaction I, Scheme 2) and esterification (reaction II, Scheme 2), giving reaction rate at 16.7 and 28.1 mmol g⁻¹ h⁻¹, respectively. In the conversion of LA to GVL (reaction III, Scheme 2), the rate of producing GVL over Sn/SiO₂ could reach as high as 12.1 mmol g⁻¹ h⁻¹ (entry 2 in Table 3). In contrast, the HL-SnO₂/SBA-15 which has only Lewis acidity, could solely catalyze the hydrogen transfer reaction and exhibit reaction rate at 3.1 mmol g⁻¹ h⁻¹ (entry 3 in Table 3). Considering hydrogen transfer is a control step in the reaction, the obviously higher activities of SnO₂/SBA-15 than HL-SnO₂/SBA-15 in hydrogen transfer step play important role to make it an efficient catalyst. As well, the high activity in the esterification step might also contribute to the good catalytic performances of SnO₂/SBA-15.

Importantly, the SnO₂/SBA-15 catalyst is reusable. After each reaction run, the catalyst can be easily separated from the reaction system and reused in the next run by simple washing and drying. As shown in Fig. 4, in the recycling process, a constant LA conversion (~85.0%) and GVL selectivity (~95.0%) could be obtained. Moreover, the performances of the fresh catalyst and used catalysts after the 6th run in the transfer hydrogenation reaction also have been tested in a shortened

Table 3 Reactions rates of various catalysts in the model reactions listed in Scheme 2

Entry	Catalyst	Reaction rate (<i>r</i> ₀ , mmol g ⁻¹ h ⁻¹)		
		Reaction I ^a	Reaction II ^b	Reaction III ^c
1	SBA-15	— ^d	<1.0	— ^d
2	SnO ₂ /SBA-15	16.7	28.1	12.1
3	HL-SnO ₂ /SBA-15	4.9	<1.0	<1.0

^a Reaction conditions: 3 mmol of 2-pentanone, 50 mmol of 2-propanol, 140 mg of catalyst, 110 °C, 30 min. ^b Reaction conditions: 5 mmol of 2-pentanol, 5 mmol of pentanoic acid, 4 g of toluene solvent, 100 mg of catalyst, 110 °C, 20 min. ^c Reaction conditions: 3 mmol of LA, 50 mmol of 2-propanol, 110 mg of catalyst, 140 °C, 20 min. ^d Undetectable.



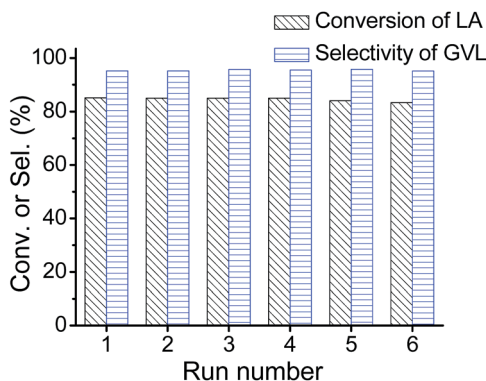


Fig. 4 Catalytic data of SnO₂/SBA-15 in the recyclability test. Reaction at 100 °C, reactor volume at 10 ml, 0.2 g of SnO₂/SBA-15, 2.0 wt% LA in 2-propanol solution, 0.2 ml min⁻¹ N₂.

reaction time, where similar catalytic performances were obtained, indicating the high stability of the Sn catalyst (entries 1, 2 in Table 4). By ICP analysis, the Sn leaching is negligible during the recycling process. The conversion of LA to GVL was also performed in a fixed-bed reactor using a mixture of LA and 2-propanol as feed (Fig. S3A†). In the catalytic tests, no obvious decrease in LA conversion and GVL selectivity was observed for a 50 h reaction (Fig. S3B†). On the other hand, the use of SBA-15 as supports has been compared with another typical silica support of MCM-41. The SnO₂/MCM-41 catalyst exhibits very comparable performance to that of SnO₂/SBA-15 (entry 3 in Table 4), suggesting that the silica support have little influence on the catalytic performances of Sn catalysts.

Table 4 Catalytic data in conversion of LA to GVL using 2-propanol over various catalysts in a shorten reaction time^a

Entry	Catalyst	LA conv. (%)	GVL Sel. (%)
1	Fresh SnO ₂ /SBA-15	59.5	91.0
2	Used SnO ₂ /SBA-15 ^b	56.2	93.7
3	SnO ₂ /MCM-41	63.4	88.1

^a Reaction conditions: 3 mmol of LA, 50 mmol of 2-propanol, 140 mg of catalyst, 110 °C, 4 h. ^b The catalyst after the 6 run in the recyclability test.

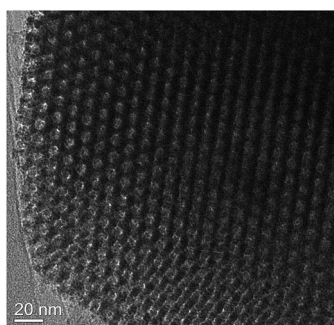


Fig. 5 TEM images of the randomly selected regions on the used SnO₂/SBA-15.

By TEM analysis (Fig. 5), the SnO₂/SBA-15 catalyst after a 50 h run still exhibited high-quality mesoporous structure. More importantly, no aggregation of Sn species was observed in the used SnO₂/SBA-15 catalyst, indicating the good stability of the catalysts. The high activity and selectivity combined with the good stability of SnO₂/SBA-15 catalyst make it potentially important for wide applications in the future.

Conclusions

In summary, we reported the transfer hydrogenation of LA to GVL using 2-propanol as hydrogen donor over a Sn modified silica catalyst (SnO₂/SBA-15) which gives high LA conversion (85.1%) and GVL selectivity (95.2%). The good catalytic performance is reasonably attributed to the strong Lewis acid sites on SnO₂/SBA-15, which could efficiently catalyze the individual reactions of hydrogen transfer, as well as the Brønsted acidity, which is benefit for the esterification step, thus leading to an efficient LA-to-GVL process without using noble metals and gaseous hydrogen. Although it is still difficult to conclude that the transfer hydrogenation excelled in LA conversion efficiency, because the LA concentration in transfer hydrogenation tests is usually lower than that in direct hydrogenation cases, the transfer hydrogenation strategy indeed offers new ideas in developing alternative pathways in LA conversion. It is expected that this work can be potentially used to develop efficient catalysts and processes for the conversion of other biomass platform molecules.

Acknowledgements

This work is supported by the Science and Technology Project of General Administration of Quality Supervision, Inspection and Quarantine of the People's Republic of China (AQSIQ, No. 2014QK203).

Notes and references

- 1 G. W. Huber, S. Iborra and A. Corma, *Chem. Rev.*, 2006, **106**, 4044.
- 2 L. Wang, H. Wang and F.-S. Xiao, *ChemSusChem*, 2014, **7**, 402.
- 3 D. Mohan, C. U. Pittman and P. H. Steele, *Energy Fuels*, 2006, **20**, 848.
- 4 D. M. Alonso, J. Q. Bond and J. A. Dumensic, *Green Chem.*, 2010, **12**, 1493.
- 5 B. Hu, K. Wang and M. M. Titirici, *Adv. Mater.*, 2010, **22**, 813.
- 6 F. Clippel, M. Dusselier and B. F. Sels, *J. Am. Chem. Soc.*, 2012, **134**, 10089.
- 7 Y. Roman-Leshkov, J. N. Chheda and J. A. Dumensic, *Science*, 2006, **312**, 1933.
- 8 A. V. Bridgwater, *Biomass Bioenergy*, 2012, **38**, 68.
- 9 P. Lenihan, A. Orozco and G. M. Walker, *Chem. Eng. J.*, 2010, **156**, 395.
- 10 T. Vom Stein, P. Grande and P. D. de Maria, *Green Chem.*, 2010, **12**, 1844.



- 11 S. Sukanuma, K. Nakajima and M. Hara, *J. Am. Chem. Soc.*, 2008, **130**, 12787.
- 12 S. Van de Vyver, L. Peng and B. F. Sels, *Green Chem.*, 2010, **12**, 1560.
- 13 M. Mascal and E. B. Nikitin, *Green Chem.*, 2010, **12**, 370.
- 14 W. R. H. Wright and R. Palkovits, *ChemSusChem*, 2012, **5**, 1657.
- 15 P. P. Upare, J.-M. Lee and J.-S. Chang, *J. Ind. Eng. Chem.*, 2011, **17**, 287.
- 16 D. J. Braden, C. A. Henao and J. A. Dumesic, *Green Chem.*, 2011, **13**, 1755.
- 17 H. A. Schuette and R. W. Thomas, *J. Am. Chem. Soc.*, 1930, **52**, 3010.
- 18 L. Deng, J. Li and Q.-X. Guo, *Angew. Chem., Int. Ed.*, 2009, **121**, 6651.
- 19 F. M. A. Geilen, B. Engendahl and W. Leitner, *Angew. Chem., Int. Ed.*, 2010, **122**, 5642.
- 20 F. M. A. Geilen, B. Engendahl and W. Lritner, *J. Am. Chem. Soc.*, 2011, **133**, 14349.
- 21 C. Ortiz-Cervantes, M. Flores-Alamo and J. J. Garcia, *ACS Catal.*, 2015, **5**, 1424.
- 22 L. E. Manzer, *Appl. Catal., A*, 2004, **272**, 249.
- 23 Z.-P. Yan, L. Lin and S. Liu, *Energy Fuels*, 2009, **23**, 3853.
- 24 Y. Yang, G. Gao and F. W. Li, *ACS Catal.*, 2014, **5**, 1419.
- 25 X.-L. Du, L. He and K.-N. Fan, *Angew. Chem., Int. Ed.*, 2011, **50**, 7815.
- 26 X.-L. Du, Q.-Y. Bi and K.-N. Fan, *ChemSusChem*, 2011, **4**, 1838.
- 27 J. J. Tan, J. L. Cui and Y. W. Li, *ACS Catal.*, 2015, **5**, 7379.
- 28 C. Xie, J. L. Song and B. X. Han, *ACS Sustainable Chem. Eng.*, 2016, **4**, 6231.
- 29 H. P. Winoto, B. S. Ahn and J. Jae, *J. Ind. Eng. Chem.*, 2016, **40**, 62.
- 30 A. H. Valekar, K.-H. Cho and Y. K. Hwang, *Green Chem.*, 2016, **18**, 4542.
- 31 M. J. Gilkey and B. J. Xu, *ACS Catal.*, 2016, **6**, 1420.
- 32 S. S. Enumula, V. R. B. Gurram and S. R. R. Kamaraju, *RSC Adv.*, 2016, **6**, 20230.
- 33 J. L. Song, L. Q. Wu and B. X. Han, *Green Chem.*, 2015, **17**, 1626.
- 34 W. R. H. Wright and R. Palkovits, *ChemSusChem*, 2012, **5**, 1657.
- 35 L. Bui, H. Luo and Y. Roman-Leshkov, *Angew. Chem., Int. Ed.*, 2013, **52**, 8022.
- 36 R. S. Assary, L. A. Curtiss and J. A. Dumesic, *ACS Catal.*, 2013, **3**, 2694.
- 37 J. L. Song, L. Q. Wu and B. X. Han, *Green Chem.*, 2015, **17**, 1626.
- 38 J. Yuan, S. S. Li and K.-N. Fan, *Energy Environ. Sci.*, 2013, **6**, 3308.
- 39 D. Y. Zhao, J. L. Feng and G. D. Stucky, *Science*, 1998, **279**, 548.
- 40 L. Wang, J. Zhang and F.-S. Xiao, *J. Mater. Chem. A*, 2014, **2**, 3725.
- 41 A. Corma, M. T. Navarro and M. Renz, *J. Catal.*, 2003, **219**, 242.
- 42 F. Babou, G. Coudurier and J. C. Vedrine, *J. Catal.*, 1995, **152**, 341.
- 43 M. I. Zaki, M. A. Hasan and L. Pasupulety, *Colloids Surf., A*, 2001, **190**, 261.
- 44 X. C. Wang, J. C. Yu and X. Z. Fu, *J. Photochem. Photobiol., A*, 2006, **179**, 339.
- 45 B. K. Chethana and S. H. Mushrif, *J. Catal.*, 2015, **323**, 158.
- 46 A. M. R. Galletti, C. Antonetti and M. Martinelli, *Green Chem.*, 2012, **14**, 688.
- 47 A. Corma, M. E. Domine and S. Valencia, *J. Am. Chem. Soc.*, 2002, **124**, 3194.
- 48 L. Wang, J. Zhang and F.-S. Xiao, *ACS Catal.*, 2015, **5**, 2727.

