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# Assessment of certain ionic liquids for separation of binary mixtures based on gamma infinity data measurements†

M. Karpińska,<sup>a</sup> M. Wlazło,<sup>a</sup> D. Ramjugernath,<sup>b</sup> P. Naidoo<sup>b</sup> and U. Domańska<sup>\*ab</sup>

The effect of interactions between organic solvents or water on the interfacial and bulk properties of 1-benzyl-3-methylimidazolium dicyanamide, [BzMIM][DCA] and 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [BzMIM][NTf<sub>2</sub>] were determined *via* measurement of activity coefficients  $\gamma_{13}^{\infty}$  at infinite dilution for 64 solutes. The data were obtained using the gas–liquid chromatography technique. Measurements were undertaken at six temperatures, in 10 K intervals, in the range of 318.15 to 368.15 K. The solutes studied included both non-polar and polar compounds, such as alkanes, alkenes, and alkynes, as well as aromatic hydrocarbons, alcohols, water, ethers, ketones, acetonitrile, pyridine, 1-nitropropane, thiophene, and esters. Density,  $\rho$ , and viscosity  $\eta$ , measurements for a range of temperatures,  $T$  for the chosen ionic liquids (ILs), [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] were also undertaken at pressure,  $p = 101$  kPa. The gas–liquid partition coefficients,  $K_L$  at infinite dilution, and the fundamental thermodynamic functions, partial molar excess Gibbs energy, enthalpy and entropy at infinite dilution were calculated from the experimental data measurements. The values of selectivity and capacity for three separation cases, *viz.* hexane/hex-1-ene, cyclohexane/cyclohexene, and ethylbenzene/styrene were calculated from  $\gamma_{13}^{\infty}$  values and compared to literature for imidazolium-based or dicyanamide-based, or bis((trifluoromethyl)sulfonyl)imide-based ionic liquids (ILs). The results from the study indicate that [BzMIM][DCA] has large selectivity values for all three of the separation cases studied.

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## Introduction

Ionic liquids (ILs) are nonvolatile salts with some rather unusual properties, such as high thermal stability, non-flammability, and high solvation properties. They have been widely studied, and have been proposed as potential replacement solvents for conventional organic solvents<sup>1–5</sup> in separation processes. Examples of studies considering ILs for separation include extraction of aromatic hydrocarbons from alkanes,<sup>5</sup> desulphurization and denitrification of diesel oil,<sup>6–8</sup> and extraction of butan-1-ol from water.<sup>9</sup>

Among the many potential separation applications, one of the most relevant is the employment of ILs as extractants for separation of alkenes from alkanes. Alkenes are particularly

useful as substrates for the production of polymers and petrochemicals. The conventional separation technique for the alkane/alkene system is by distillation. Distillation processes can be extremely energy-intensive and technologically complex. Other separation methods, such as membranes technology or adsorption processes generally achieve a poor separation selectivity. Over the last decade, the adsorption of alkenes onto carbon, or zeolite A, or zeolite imidazolate framework ZIF-7, or metal–organic frameworks (MOFs) instant zeolites have been proposed as feasible, economical, and green alternatives to distillation.<sup>10–12</sup>

Preliminary information about suitable solvents for separation can be obtained from activity coefficients measurements at infinite dilution ( $\gamma_{13}^{\infty}$ ) by the gas–liquid chromatography (GLC) technique. Solvent suitability can also be determined from liquid–liquid phase equilibrium measurements in ternary systems. The information obtained from infinite dilution activity coefficients and liquid–liquid equilibrium measurement is generally the first step in the engineering design process to determine the most suitable extraction and separation processes. The use of predictive models, such as Mo UNI-FAC, PC SAFT, or COSMO RS are to date not accurate and reliable enough for the design of separation processes and restricted to the description of simple systems only. Thus, the importance and need for experimental data. In this study we

<sup>a</sup>Department of Physical Chemistry, Faculty of Chemistry, Warsaw University of Technology, Noakowskiego 3, 00-664 Warsaw, Poland. E-mail: ula@ch.pw.edu.pl; Fax: +48 22 6282741; Tel: +48 22 6213115

<sup>b</sup>Thermodynamic Research Unit, School of Chemical Engineering, University of KwaZulu-Natal, Howard College Campus, King George V Avenue, Durban 4001, South Africa

† Electronic supplementary information (ESI) available: The sources and mass fraction purities of materials. Mean column pressure,  $\bar{p}$ , inlet column pressure,  $p_i$ , outlet column pressure,  $p_o$  and standard state of solutes at given temperatures, at standard state. Plots of  $\ln(\gamma_{13}^{\infty})$  versus  $1/T$  for the organic solute in ILs. See DOI: 10.1039/c6ra25208g



propose to estimate the solvent separation ability for some ILs by an analysis of their selectivities and capacities, calculated from  $\gamma_{13}^{\infty}$  values.

The present work is focused on new attractive ILs, namely – 1-benzyl-3-methylimidazolium dicyanamide, [BzMIM][DCA] and 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [BzMIM][NTf<sub>2</sub>]. In recent years ILs have been widely promoted as solvents for alkane/alkene separations.<sup>13–16</sup> The study of alkane/alkene separations using ILs reported in the literature deal mostly with short chain hydrocarbons. The imidazolium-based ILs were proposed for the propane/prop-1-ene<sup>13</sup> and the hexane/hex-1-ene separations.<sup>14</sup> Usually, an increase of the substituent size leads to a decrease of the separation factor and to an increase in the capacity.<sup>13,14</sup>

Recently,  $\gamma_{13}^{\infty}$  data were measured by us for 1-butyl-3-methylimidazolium dicyanamide, [BMIM][DCA],<sup>15</sup> and 1-allyl-3-methylimidazolium dicyanamide, [AMIM][DCA].<sup>16</sup> These measurements revealed a very promising selectivity and capacity for the hexane/hex-1-ene, cyclohexane/cyclohexene, and ethylbenzene/styrene separation processes.

The comparison of different ILs, applied for the alkane/alkene separation may be performed from the data reported in the review of Domańska *et al.*<sup>15</sup> Interestingly, high selectivity for the hexane/hex-1-ene separation were obtained for the ILs containing CN– groups substituted on the cation or anion. High selectivities whereas observed for 1-ethyl-3-methylimidazolium dicyanamide, [EMIM][DCA],<sup>17</sup> 1-(3-cyano-propyl)-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [CN–C<sub>3</sub>MIM][NTf<sub>2</sub>],<sup>18</sup> and 1-(3-cyano-propyl)2,4-dimethyl-imidazolium dicyanamide, [CN–C<sub>3</sub>MIM][DCA].<sup>18</sup> The ILs with dicyanamide, or tricyanomethanide anions were also found to be very good for extraction of thiophene, or pyridine from aliphatic hydrocarbons.<sup>6,8,19</sup> As an example, [AMIM][NTf<sub>2</sub>] has a selectivity ( $S_{12}^{\infty} =$

$\gamma_{13}^{\infty}/\gamma_{23}^{\infty} = 1.94$ ),<sup>20</sup> and [AMIM][DCA] has a selectivity ( $S_{12}^{\infty} = \gamma_{13}^{\infty}/\gamma_{23}^{\infty} = 2.51$ )<sup>16</sup> at  $T = 328.15$  K in the hexane/hex-1-ene separation process. The bis((trifluoromethyl)sulfonyl)imide, [NTf<sub>2</sub>]-based ILs are appropriate alternatives for apolar molecular liquids and are widely proposed as new entrainers for many processes.<sup>21</sup>

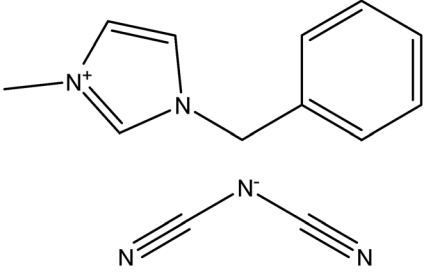
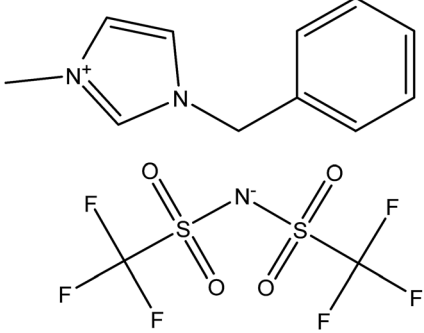
In this work we continue our investigations on the measurements of  $\gamma_{13}^{\infty}$  and an analysis of the obtained

**Table 2** Density,  $\rho$  and viscosity,  $\eta$  as a function of temperature,  $T$  for investigated ILs [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] at pressure  $p = 101$  kPa<sup>a</sup>

$T/K$	[BzMIM][DCA]		[BzMIM][NTf <sub>2</sub> ]	
	$\rho/g\text{ cm}^{-3}$	$\eta/\text{mPa s}$	$\rho/g\text{ cm}^{-3}$	$\eta/\text{mPa s}$
298.15	1.15814	102	1.49098 <sup>b</sup>	133 <sup>c</sup>
303.15	1.15480	—	1.48609	—
308.15	1.15147	56.1	1.48124	74.9
313.15	1.14816	—	1.47644	—
318.15	1.14488	34.5	1.47167	46.2
323.15	1.14163	—	1.46693 <sup>d</sup>	—
328.15	1.13839	23.2	1.46222	30.9
333.15	1.13517	—	1.45756	—
338.15	1.13183	16.5	1.45295	21.9
343.15	1.12869	—	1.44837	—
348.15	1.12528	12.3	1.44381	16.2
353.15	1.12201	—	1.43917	—
358.15	1.11876	9.56	1.43470	12.4
363.15	1.11544	—	1.43013	—
368.15	1.11211	7.61	1.42554	9.74

<sup>a</sup> Standard uncertainties  $u$  are  $u(T) = \pm 0.1$  K,  $u(\rho) = \pm 1.1 \times 10^{-3}$  g cm<sup>-3</sup>,  $u(\eta) = \pm 0.15\%$ ,  $u(p) = \pm 1$  kPa. <sup>b</sup>  $\rho/g\text{ cm}^{-3} = 1.491$  in ref. 23; 1.489 in ref. 24; 1.504 in ref. 25; 1.42 in ref. 26; at  $T = 293.15$  K; 1.151 at  $T = 293.15$  K in ref. 27; 1.491 (room temperature, 25 °C) in ref. 28; 1.441 at  $T = 295$  K in ref. 29. <sup>c</sup>  $\eta/\text{mPa s} = 113$  in ref. 23, 125 in ref. 24; 61 in ref. 26; 190 at  $T = 293.15$  K; 160 at  $T = 293.2$  K in ref. 27; 150.8 (room temperature, 25 °C) in ref. 28; 40 (room temperature) in ref. 22. <sup>d</sup>  $\rho/g\text{ cm}^{-3} = 1.458$  in ref. 23.

**Table 1** The name, abbreviation, structure, midpoint glass transition temperature,  $T_g$  and change of heat capacity at the glass transition temperature,  $\Delta C_{p(g)}$  of investigated ionic liquids at pressure  $p = 101$  kPa<sup>a</sup>

Name, abbreviation	1-Benzyl-3-methylimidazolium dicyanamide, [BzMIM][DCA]	1-Benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [BzMIM][NTf <sub>2</sub> ]
Structure		
$M/g\text{ mol}^{-1}$	239.28	453.38
$T_g/K$	217.0 <sup>b</sup>	216.9 <sup>c,d</sup>
$\Delta C_{p(g)}/J\text{ mol}^{-1}\text{ K}^{-1}$	169.9 <sup>b</sup>	199.5 <sup>c,d</sup>

<sup>a</sup> Standard uncertainties  $u$  are as follows;  $u(T_g) = \pm 0.1$  K;  $u(\Delta C_{p(g)}) = \pm 5$  J mol<sup>-1</sup> K<sup>-1</sup>,  $u(p) = \pm 1$  kPa. <sup>b</sup> A midpoint  $T_g/K = 195.4$  with  $\Delta C_{p(g)}/J\text{ mol}^{-1}\text{ K}^{-1} = 214$  in ref. 22. <sup>c</sup> A midpoint  $T_g/K = 217$  with  $\Delta C_{p(g)}/J\text{ mol}^{-1}\text{ K}^{-1} = 86.1$  in ref. 23. <sup>d</sup> A midpoint  $T_g/K = 215.8$  in ref. 24;  $T_g/K = 215.8$  K in ref. 25;  $T_g/K = 210.8$  K with jump  $\Delta C_{p(g)}/J\text{ mol}^{-1}\text{ K}^{-1} = 197$  in ref. 22.



**Table 3** The experimental activity coefficients at infinite dilution  $\gamma_{13}^{\infty}$  for the solutes in ionic liquid [BzMIM][DCA] at different temperatures for the hypothetical liquid at zero pressure<sup>a</sup>

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentane	71.4	65.0	59.5	54.7	50.5	
Hexane	105	95.4	87.8	81.0	75.1	70.2
3-Methylpentane	95.0	86.1	78.6	72.2	66.5	61.7
2,2-Dimethylbutane	115	101	88.8	78.9	70.4	
Heptane	159	144	132	121	111	103
Octane	245	221	201	184	169	156
2,2,4-Trimethylpentane	216	193	176	160	147	135
Nonane	372	332	299	271	247	227
Decane	575	509	455	409	369	336
Cyclopentane	22.2	21.0	20.0	19.0	18.2	17.4
Cyclohexane	36.0	33.5	31.4	29.6	27.9	26.4
Methylcyclohexane	58.7	54.2	50.3	46.9	43.9	41.3
Cycloheptane	44.1	41.0	38.4	36.0	34.0	32.2
Cyclooctane	58.3	53.9	50.2	46.9	44.0	41.5
Pent-1-ene	25.8	24.7	23.7	22.9	22.1	21.4
Hex-1-ene	40.1	38.1	36.4	34.7	33.3	32.1
Cyclohexene	13.9	13.6	13.3	13.0	12.8	12.5
Hept-1-ene	62.1	58.8	56.0	53.3	51.0	49.0
Oct-1-ene	98.3	92.2	87.0	82.5	78.5	74.9
Dec-1-ene	238	220	205	191	179	168
Pent-1-yne	5.20	5.36	5.50	5.65	5.79	5.93
Hex-1-yne	7.97	8.12	8.28	8.40	8.56	8.69
Hept-1-yne	12.7	12.8	12.9	13.0	13.1	13.2
Oct-1-yne	20.3	20.2	20.2	20.1	20.1	20.0
Benzene	2.36	2.44	2.53	2.61	2.70	2.78
Toluene	3.91	4.05	4.19	4.31	4.44	4.56
Ethylbenzene	6.56	6.71	6.85	6.97	7.10	7.22
<i>o</i> -Xylene	5.35	5.51	5.65	5.80	5.94	6.06
<i>m</i> -Xylene	6.77	6.95	7.12	7.27	7.42	7.57
<i>p</i> -Xylene	6.60	6.76	6.93	7.08	7.25	7.39
<i>n</i> -Propylbenzene	10.9	11.0	11.2	11.3	11.4	11.6
Iso-propylbenzene	10.6	10.9	11.1	11.3	11.5	11.7
Styrene	2.97	3.11	3.23	3.36	3.49	3.62
$\alpha$ -Methylstyrene	4.77	5.03	5.29	5.54	5.80	6.06
Thiophene	1.37	1.44	1.51	1.58	1.64	1.71
Pyridine	0.989	1.03	1.08	1.12	1.16	1.21
Methanol	0.503	0.511	0.518	0.525	0.533	0.539
Ethanol	0.874	0.872	0.87	0.869	0.868	0.867
Propan-1-ol	1.17	1.16	1.16	1.15	1.15	1.15
Propan-2-ol	1.37	1.35	1.33	1.31	1.29	1.27
Butan-1-ol	1.68	1.66	1.64	1.62	1.60	1.59
Butan-2-ol	1.76	1.75	1.74	1.73	1.72	1.71
2-Methyl-1-propanol	1.72	1.68	1.65	1.62	1.60	1.57
<i>tert</i> -Butanol	1.81	1.80	1.78	1.77	1.76	1.75
1-Pentanol	2.26	2.24	2.23	2.21	2.20	2.19
Water	0.378	0.387	0.396	0.404	0.412	0.42
Methyl acetate	1.85	1.94	2.03	2.11	2.19	2.28
Methyl propanoate	2.77	2.90	3.03	3.14	3.26	3.38
Methyl butanoate	4.40	4.55	4.68	4.81	4.94	5.07
Ethyl acetate	3.11	3.22	3.34	3.45	3.56	3.67
Tetrahydrofuran	1.71	1.78	1.86	1.93	2.00	2.08
1,4-Dioxane	1.09	1.16	1.23	1.30	1.37	1.45
<i>tert</i> -Butyl methyl ether	11.4	11.5	11.7	11.8	12.0	12.1
<i>tert</i> -Butyl ethyl ether	30.9	30.6	30.5	30.3	30.1	30.0
<i>tert</i> -Amyl methyl ether	16.9	17.0	17.1	17.2	17.3	17.3
Diethyl ether	9.88	9.94	9.99	10.0	10.1	10.1
Di- <i>n</i> -propyl ether	32.3	31.6	31.0	30.4	29.9	29.4
Di- <i>iso</i> -propyl ether	39.0	37.8	36.8	35.9	35.1	34.3
Di- <i>n</i> -butyl ether	81.2	77.2	73.8	70.8	68.0	65.5
Acetone	1.12	1.17	1.23	1.28	1.33	1.39



Table 3 (Contd.)

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentan-2-one	2.50	2.59	2.70	2.79	2.89	2.98
Pentan-3-one	2.52	2.62	2.73	2.83	2.93	3.02
Butanal	1.96	2.04	2.13	2.21	2.29	2.36
Acetonitrile	0.879	0.903	0.927	0.952	0.974	0.996
1-Nitropropane	1.53	1.58	1.63	1.68	1.73	1.77

<sup>a</sup> Standard uncertainties  $u$  are  $u(\gamma_{13}^{\infty}) = 3\%$ ,  $u(T) = 0.02$  K.

selectivities and capacities for hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene separation processes. This work proposes to study a particular type of ILs, 1-benzyl-3-methylimidazolium dicyanamide, [BzMIM][DCA] and 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [BzMIM][NTf<sub>2</sub>] as a function of temperature at ambient pressure to compare the results with data previously measured by us for [AMIM][DCA] and [AMIM][NTf<sub>2</sub>]. According to many sources in literature, the high interaction of aromatic compounds, here of ethylbenzene with the IL is caused by  $\pi$ - $\pi$  interactions between the cation with the benzyl group and aromatics. Because of this, we have selected cations with the benzyl group. In this paper we report the activity coefficients  $\gamma_{13}^{\infty}$  at infinite dilution, the gas-liquid partition coefficients  $K_L$  and thermodynamic functions at infinite dilution for all measured solutes. The thermodynamic properties obtained are analyzed with regard to intermolecular interactions. A physico-chemical characterization of the ILs was also performed by measuring their thermal properties with DSC, as well as densities and viscosities as a function of temperature.

## Materials and methods

### Materials

The ILs were purchased from Io-Li-Tec (Ionic Liquids Technologies, GmbH, Heilbronn, Germany), 1-benzyl-3-methylimidazolium dicyanamide, purity > 0.98 mass fraction; CAS: 958445-60-8, and 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, purity > 0.99 mass fraction; CAS: 433337-24-7. The names, abbreviation of names, structures, molar masses, glass transition temperatures ( $T_g$ ) and heat capacity, ( $C_{p(g)}$ ) values at glass transition are listed in Table 1. The different solutes, purchased from Aldrich or Fluka, had purities better than 0.99 mass fraction and were used without further purification due to the fact that the GLC technique separates any impurities on the column. The specification and purity is shown in Table 1S in the ESI.†

### Water content

The water content of the solvents was analyzed by the Karl-Fischer titration technique (method TitroLine KF). The sample of IL, or solvent was dissolved in methanol and titrated in steps of 0.0025 cm<sup>3</sup>. The uncertainty on the water content was  $u(w.c.) = 10$  ppm for the 3 cm<sup>3</sup> sample of IL injected. The water

content in ILs was 200 ppm and 320 ppm for [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>], respectively.

### Differential scanning calorimetry, DSC

The basic thermal characteristics of ILs, *i.e.* temperature of glass transition and heat capacity at glass transition temperature ( $T_g$ ,  $C_{p(g)}$ ) were measured using a differential scanning microcalorimetry technique (DSC). The applied scan rate was 5 K min<sup>-1</sup>, with power and recorder sensitivities of 16 mJ s<sup>-1</sup> and 5 mV, respectively. The apparatus (DSC 1 STAR<sup>®</sup>System from Mettler Toledo with Liquid Nitrogen Cooling System) was calibrated with a 0.999999 mol fraction purity indium sample. The repeatability of the glass transition temperature value was  $\pm 0.1$  K. The uncertainty of the glass temperature was  $u(T_g) = 0.1$  K, and that of the heat capacity at glass transition  $u(C_{p(g)}) = 5$  J mol<sup>-1</sup> K<sup>-1</sup>. The DSC diagrams of the ILs are shown as Fig. 1Sa and b in the ESI† and the final results together with the literature data are listed in Table 1.<sup>22-25</sup>

### Density measurements

The density of the ILs was measured using an Anton Paar GmbH 4500 vibrating-tube densimeter (Graz, Austria), thermostated over a temperature range of (298.15–368.15) K. Two integrated Pt-100 platinum thermometers provided good precision in temperature control internally ( $T \pm 0.01$  K). The densimeter has an automatic correction for the viscosity of the sample. The apparatus is precise to within  $1 \times 10^{-5}$  g cm<sup>-3</sup>, and the uncertainty of the measurements was estimated to be  $u(\rho) = \pm 1.1 \times 10^{-3}$  g cm<sup>-3</sup>. The densimeter's calibration was performed at atmospheric pressure using doubly distilled and degassed water (PURE LAB Option Q Elga Water System), specially purified benzene (CHEMIPAN, Poland 0.999), and dried air. The densities of ILs are listed in Table 2.<sup>23-29</sup>

### Viscosity measurements

Viscosity measurements were carried out on an Anton Paar GmbH AMVn (Graz, Austria) programmable rheometer, with a nominal uncertainty of  $\pm 0.15\%$  and reproducibility < 0.05% for viscosities from 7.6 mPa s to 140 mPa s. Temperature was controlled internally with a precision of  $\pm 0.01$  K in a range from (298.15 to 368.15) K. The diameter of the capillary was 1.8 mm and 3 mm for viscosities from (2.5 to 70) mPa s, and (20 to



**Table 4** The experimental activity coefficients at infinite dilution  $\gamma_{13}^{\infty}$  for the solutes in ionic liquid [BzMIM][NTf<sub>2</sub>] at different temperatures for the hypothetical liquid at zero pressure<sup>a</sup>

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentane	13.5	12.9	12.3	11.8	11.4	11.0
Hexane	18.4	17.6	16.8	16.1	15.5	14.9
3-Methylpentane	16.1	15.4	14.9	14.4	13.9	13.5
2,2-Dimethylbutane	15.8	15.1	14.6	14.1	13.6	13.2
Heptane	26.2	24.8	23.5	22.4	21.4	20.4
Octane	37.9	35.3	33.0	31.0	29.2	27.5
2,2,4-Trimethylpentane	26.1	24.9	23.8	22.8	21.9	21.1
Nonane	53.2	49.3	45.9	42.9	40.2	37.8
Decane	77.3	70.7	65.0	60.1	55.7	51.9
Cyclopentane	6.82	6.61	6.43	6.28	6.11	5.98
Cyclohexane	9.85	9.39	9.01	8.64	8.33	8.01
Methylcyclohexane	13.6	13.0	12.5	12.0	11.6	11.2
Cycloheptane	12.7	12.1	11.6	11.1	10.7	10.3
Cyclooctane	16.5	15.6	14.8	14.0	13.3	12.7
Pent-1-ene	6.66	6.61	6.56	6.52	6.47	6.43
Hex-1-ene	9.53	9.27	9.05	8.85	8.65	8.49
Cyclohexene	5.40	5.31	5.25	5.18	5.11	5.05
Hept-1-ene	13.2	12.9	12.6	12.4	12.1	11.9
Oct-1-ene	19.3	18.6	17.9	17.2	16.7	16.2
Dec-1-ene	38.8	36.8	35.1	33.5	32.1	30.8
Pent-1-yne	2.38	2.46	2.53	2.60	2.67	2.74
Hex-1-yne	3.32	3.40	3.47	3.55	3.62	3.68
Hept-1-yne	4.69	4.76	4.83	4.90	4.96	5.02
Oct-1-yne	6.64	6.65	6.7	6.73	6.75	6.77
Benzene	0.999	1.05	1.10	1.14	1.19	1.23
Toluene	1.49	1.55	1.60	1.66	1.72	1.77
Ethylbenzene	2.19	2.26	2.33	2.40	2.48	2.54
<i>o</i> -Xylene	1.98	2.04	2.10	2.16	2.22	2.27
<i>m</i> -Xylene	2.20	2.28	2.37	2.45	2.53	2.61
<i>p</i> -Xylene	2.21	2.30	2.38	2.46	2.54	2.62
<i>n</i> -Propylbenzene	3.17	3.26	3.35	3.43	3.52	3.59
Iso-propylbenzene	3.11	3.19	3.27	3.35	3.43	3.51
Styrene	1.26	1.32	1.39	1.46	1.52	1.59
$\alpha$ -Methylstyrene	1.88	1.98	2.10	2.20	2.31	2.42
Thiophene	0.860	0.900	0.938	0.974	1.01	1.05
Pyridine	0.487	0.512	0.537	0.561	0.586	0.610
Methanol	1.34	1.28	1.22	1.16	1.12	1.07
Ethanol	1.71	1.61	1.52	1.45	1.37	1.31
Propan-1-ol	2.08	1.96	1.85	1.76	1.68	1.60
Propan-2-ol	1.96	1.82	1.69	1.58	1.49	1.40
Butan-1-ol	2.69	2.50	2.33	2.19	2.06	1.95
Butan-2-ol	2.30	2.15	2.03	1.91	1.81	1.72
2-Methyl-1-propanol	2.64	2.44	2.25	2.09	1.95	1.83
<i>tert</i> -Butanol	1.91	1.81	1.73	1.65	1.59	1.53
1-Pentanol	3.18	2.97	2.78	2.62	2.47	2.33
Water	3.27	2.97	2.72	2.50	2.32	2.15
Methyl acetate	0.603	0.638	0.673	0.707	0.742	0.776
Methyl propanoate	0.768	0.814	0.863	0.909	0.955	1.00
Methyl butanoate	1.09	1.14	1.19	1.24	1.29	1.34
Ethyl acetate	0.788	0.834	0.881	0.926	0.974	1.02
Tetrahydrofuran	0.610	0.648	0.684	0.721	0.758	0.794
1,4-Dioxane	0.484	0.523	0.562	0.602	0.644	0.683
<i>tert</i> -Butyl methyl ether	2.30	2.41	2.52	2.62	2.72	2.83
<i>tert</i> -Butyl ethyl ether	4.97	5.11	5.26	5.40	5.54	5.67
<i>tert</i> -Amyl methyl ether	3.33	3.44	3.53	3.63	3.73	3.82
Diethyl ether	2.30	2.39	2.47	2.56	2.64	2.72
Di- <i>n</i> -propyl ether	5.93	5.99	6.04	6.08	6.13	6.19
Di- <i>iso</i> -propyl ether	5.68	5.79	5.88	5.98	6.08	6.18
Di- <i>n</i> -butyl ether	12.5	12.2	11.9	11.7	11.5	11.3
Acetone	0.386	0.410	0.434	0.459	0.483	0.507





Table 4 (Contd.)

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentan-2-one	0.652	0.693	0.734	0.773	0.814	0.855
Pentan-3-one	0.651	0.697	0.741	0.788	0.833	0.879
Butanal	0.675	0.711	0.747	0.782	0.818	0.852
Acetonitrile	0.436	0.455	0.472	0.490	0.507	0.523
1-Nitropropane	0.669	0.692	0.714	0.736	0.758	0.777

<sup>a</sup> Standard uncertainties  $u$  are  $u(\gamma_{13}^{\infty}) = 3\%$ ,  $u(T) = 0.02$  K.

230) mPa s, respectively. The diameter of the balls were 1.5 mm and 2.5 mm for the above viscosity ranges. The data obtained for the ILs are listed in Table 2.<sup>22–24,26–28</sup>

### Apparatus and experimental procedure

Experiments were performed using a Perkin Elmer Clarus 500 gas chromatograph equipped with a thermal conductivity detector (TCD). The data were collected and processed using the TotalChrom Workstation software. The column preparation and the packing method used in this work have been described in detail in our previous work.<sup>19,20</sup> Glass columns of length 1 m, with a 4 mm internal diameter were used. The solid support Chromosorb W/AW-DCMS 100/120 mesh was supplied by Sigma-Aldrich. Coating of the solid support material with the IL was performed by dispersing a certain portion of the IL in methanol, followed by evaporation of the solvent using a rotary evaporator. The masses of the stationary phase and of the solid support were weighed with a precision  $\pm 0.0001$  g, achieving an uncertainty in the IL loading on the column in the order of  $2 \times 10^{-4}$  mmol. The solvent loading on the column for 1-benzyl-3-methylimidazolium dicyanamide was 45.1% and 49.7% mass percent, and for 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide it was 45.0% and 50.7% mass percent. The large loading prevents possible residual adsorption of solute onto the column packing. Care was taken to ensure that the methanol was completely evaporated from the IL-coated solid support prior to column fabrication. Prior to each experiment, the column was conditioned by blowing hot carrier gas through it at a high flow rate ( $\sim 2.0$  cm<sup>3</sup> s<sup>-1</sup>) at 370 K for about 8 h. The pressure drop ( $P_1 - P_o$ ) was varied between 60 and 80 kPa depending on the flow rate of the carrier gas. The inlet pressure,  $P_i$ , was measured by a pressure gauge installed on the gas chromatograph with an uncertainty of  $\pm 0.1$  kPa and the outlet pressure,  $P_o$ , was measured using an Agilent Precision Gas Flow Meter having an uncertainty of  $\pm 0.07$  kPa. The mean column pressure,  $\bar{p}$ , inlet column pressure,  $P_i$ , outlet column pressure,  $P_o$  and standard state of solutes at given temperatures and st. state are listed in Table 2S in the ESI.†

The carrier gas used was helium. The flow rate of carrier gas was determined using an Agilent precision gas flow meter which was placed at the outlet after the detector and had an uncertainty of  $\pm 0.1$  ml min<sup>-1</sup>. The flow rate was set for a series of runs and was allowed to stabilize for at least 15 min before

any  $\gamma_{13}^{\infty}$  determinations were made. Solute injections ranged from 0.01 to 0.3  $\mu$ l and can be considered to be at “infinite dilution” on the column.

Temperature-dependent experiments were carried out in 10 K steps from (318.15 to 368.15) K. The temperature of the column was maintained constant to within  $\pm 0.02$  K. At a given temperature, each experiment was repeated two to three times to establish reproducibility. Retention times were generally reproducible to within  $10^{-3}$  to  $10^{-2}$  min depending upon the temperature and the individual solute. At each temperature, values of the dead time,  $t_G$ , equivalent to the retention time of a completely non-retained component were also measured. While our GC was equipped with a TCD detector, air was used as a non-retainable component. The estimated overall error in  $\gamma_{13}^{\infty}$  was less than 3%, taking into account the possible errors in determining the column loading, the retention times, and solute vapor pressure. The resultant activity coefficient values as a function of temperature are summarized in Tables 3 and 4 for [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>], respectively.

### Theoretical basis

The equations developed by Everett<sup>30</sup> and Cruickshank *et al.*<sup>31</sup> were used in this work to calculate  $\gamma_{13}^{\infty}$  for solutes in ILs:

$$\ln \gamma_{13}^{\infty} = \ln \left( \frac{n_3 RT}{V_N P_1^*} \right) - \frac{P_1^* (B_{11} - V_1^*)}{RT} + \frac{P_o J_2^3 (2B_{12} - V_1^{\infty})}{RT} \quad (1)$$

In this expression,  $n_3$  is the number of moles of solvent on the column packing,  $R$  is the universal gas constant,  $T$  is the column temperature,  $V_N$  denotes the net retention volume of the solute,  $P_1^*$  is the saturated vapor pressure of the solute at temperature  $T$ ,  $B_{11}$  is the second virial coefficient of pure solute,  $V_1^*$  is the molar volume of the solute,  $P_o$  is the outlet pressure,  $P_o J_2^3$  is the mean column pressure,  $B_{12}$  (where the subscript “2” refers to the carrier gas, in this case helium) is the mixed second virial coefficient of the solute and carrier gas, and  $V_1^{\infty}$  is the partial molar volume of the solute at infinite dilution in the solvent. The thermophysical properties required in developing the activity coefficients at infinite dilution were calculated using equations and constants known from literature.<sup>32</sup> The values of  $B_{12}$  were calculated using the Tsonopolous equation.<sup>33</sup> The pressure correction term,  $J_2^3$ , is given by:



Table 5 The experimental gas–liquid partition coefficients  $K_L$  for the solutes in ionic liquid [BzMIM][DCA] at different temperatures

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentane	1.38	1.16	0.989	0.855	0.749	
Hexane	2.76	2.19	1.77	1.45	1.22	1.03
3-Methylpentane	2.51	2.03	1.67	1.39	1.18	1.02
2,2-Dimethylbutane	1.33	1.15	1.01	0.892	0.802	
Heptane	5.29	3.99	3.07	2.41	1.94	1.58
Octane	9.85	7.07	5.18	3.92	3.02	2.37
2,2,4-Trimethylpentane	3.73	2.88	2.26	1.82	1.49	1.23
Nonane	18.4	12.6	8.89	6.45	4.80	3.65
Decane	33.8	22.1	15.0	10.5	7.53	5.52
Cyclopentane	6.70	5.29	4.24	3.46	2.87	2.41
Cyclohexane	12.0	9.12	7.10	5.63	4.54	3.72
Methylcyclohexane	14.4	10.8	8.26	6.45	5.14	4.16
Cycloheptane	37.8	27.1	19.9	15.0	11.5	9.05
Cyclooctane	97.1	66.3	46.4	33.4	24.6	18.5
Pent-1-ene	3.16	2.55	2.08	1.73	1.46	1.25
Hex-1-ene	6.10	4.65	3.61	2.88	2.33	1.91
Cyclohexene	33.3	24.2	17.9	13.6	10.5	8.33
Hept-1-ene	11.2	8.19	6.11	4.69	3.66	2.91
Oct-1-ene	20.4	14.2	10.2	7.45	5.60	4.30
Dec-1-ene	64.4	41.5	27.5	18.8	13.3	9.53
Pent-1-yne	21.4	15.6	11.7	8.92	6.95	5.50
Hex-1-yne	40.2	28.2	20.3	15.0	11.3	8.65
Hept-1-yne	71.6	48.2	33.4	23.7	17.3	12.9
Oct-1-yne	126	81.3	54.1	37.2	26.2	19.0
Benzene	183	125	86.9	62.4	45.7	34.2
Toluene	331	215	144	99.1	70.1	50.8
Ethylbenzene	525	329	213	143	98.3	69.4
<i>o</i> -Xylene	896	548	349	228	154	107
<i>m</i> -Xylene	575	357	230	153	104	73.2
<i>p</i> -Xylene	568	354	228	152	104	72.8
<i>n</i> -Propylbenzene	790	480	301	196	131	90.1
Iso-propylbenzene	628	379	238	155	104	71.4
Styrene	1710	1023	637	409	272	185
$\alpha$ -Methylstyrene	2347	1358	818	509	327	215
Thiophene	369	244	167	117	84.5	62.3
Pyridine	1691	1058	685	458	315	222
Methanol	579	380	258	179	128	93.4
Ethanol	636	407	269	183	128	91.5
Propan-1-ol	1181	719	455	298	201	140
Propan-2-ol	512	322	211	142	99.5	71.4
Butan-1-ol	2241	1298	785	495	323	218
Butan-2-ol	889	533	333	217	146	102
2-Methyl-1-propanol	1422	836	515	329	219	150
<i>tert</i> -Butanol	391	243	157	106	73.9	53.2
1-Pentanol	4132	2290	1328	804	506	330
Water	3504	2139	1351	880	592	406
Methyl acetate	108	74.2	52.3	37.9	28.1	21.3
Methyl propanoate	165	108	73.8	51.8	37.3	27.5
Methyl butanoate	250	160	106	72.5	51.0	36.7
Ethyl acetate	136	90.4	62.0	43.8	31.7	23.6
Tetrahydrofuran	155	107	75.1	54.4	40.3	30.4
1,4-Dioxane	907	576	376	253	175	124
<i>tert</i> -Butyl methyl ether	16.1	11.7	8.71	6.64	5.17	4.10
<i>tert</i> -Butyl ethyl ether	10.9	7.80	5.74	4.34	3.34	2.63
<i>tert</i> -Amyl methyl ether	31.8	22.1	15.8	11.6	8.70	6.67
Diethyl ether	9.36	7.05	5.44	4.28	3.43	2.80
Di- <i>n</i> -propyl ether	19.5	13.7	9.95	7.38	5.60	4.34
Di- <i>iso</i> -propyl ether	7.51	5.55	4.20	3.24	2.55	2.05
Di- <i>n</i> -butyl ether	60.2	39.2	26.4	18.2	13.0	9.45
Acetone	173	119	84.7	61.7	45.8	34.8



Table 5 (Contd.)

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentan-2-one	412	266	177	121	85.1	61.3
Pentan-3-one	400	259	173	118	83.2	59.9
Butanal	188	128	89.1	63.9	46.8	35.2
Acetonitrile	536	369	260	188	139	105
1-Nitropropane	2102	1293	823	542	367	255

$$J_2^3 = \frac{2 (P_i/P_o)^3 - 1}{3 (P_i/P_o)^2 - 1} \quad (2)$$

The net retention volume of the solute,  $V_N$ , is given by:

$$V_N = (J_2^3)^{-1} U_o (t_R - t_G) \quad (3)$$

where  $t_R$  and  $t_G$  are the retention times for the solute and an unretained gas, respectively, and  $U_o$  is the column outlet flow rate, taken from the apparatus.

While the activity coefficients at infinite dilution are determined as a function of temperature,  $\ln \gamma_{13}^\infty$  can be split to its respective partial molar excess thermodynamic functions:

$$\ln \gamma_{13}^\infty = \frac{\Delta H_1^{E,\infty}}{RT} - \frac{\Delta S_1^{E,\infty}}{R} \quad (4)$$

Assuming that the temperature dependence follows a linear van't Hoff plot:

$$\ln \gamma_{13}^\infty = a/T + b \quad (5)$$

the partial molar excess enthalpy,  $\Delta H_1^{E,\infty} = Ra$ , at infinite dilution can be obtained from the slope.

The gas-liquid partition coefficient  $K_L = (c_1^g/c_1^l)$  for a solute partitioning between a carrier gas and the IL was calculated from the solute retention according to the following equation

$$\ln(K_L) = \ln\left(\frac{V_N \rho_3}{m_3}\right) - \frac{P_o J_2^3 (2B_{12} - V_1^\infty)}{RT} \quad (6)$$

In which  $\rho_3$  is the density of the IL and  $m_3$  is the mass of the IL and  $V_1^\infty$  is the partial molar volume of the solute at infinite dilution.

## Results and discussion

### Thermal and physico-chemical characteristic

The DSC diagrams for [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] are shown in the ESI.† Numerical data are included in Table 1. The temperatures of the glass phase transition ( $T_g$ , midpoint) are very close to each other and are  $217.0 \pm 0.1$  K ([BzMIM][DCA]) and  $216.9 \text{ K} \pm 0.1$  ([BzMIM][NTf<sub>2</sub>]). The literature values are 195.4 K (ref. 22) for [BzMIM][DCA] and 217 K (ref. 23), 215.8 K (ref. 24), 215.8 K (ref. 25) and 210.8 K (ref. 22) for [BzMIM][NTf<sub>2</sub>].

The values of  $\Delta C_{p(g)}$  are in a range of  $169.9 \pm 5 \text{ J mol}^{-1} \text{ K}^{-1}$  ([BzMIM][DCA]) to  $199.5 \pm 5 \text{ J mol}^{-1} \text{ K}^{-1}$  [BzMIM][NTf<sub>2</sub>]. It differs from values in ref. 31 ( $\Delta C_{p(g)}/\text{J mol}^{-1} \text{ K}^{-1} = 86.1$ ) and is very close to ref. 30 ( $\Delta C_{p(g)}/\text{J mol}^{-1} \text{ K}^{-1} = 197$ ) for [BzMIM][NTf<sub>2</sub>].

Density and viscosity of [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] were measured in this work as a function of temperature and are listed in Table 2 together with information from literature for the same ILs. The density measurements were carried out in the temperature range from (298.15 to 368.15) K at atmospheric pressure. The value obtained at  $T = 298.15$  K is  $1.15814 \text{ g cm}^{-3}$  for [BzMIM][DCA] and  $1.49098 \text{ g cm}^{-3}$  for [BzMIM][NTf<sub>2</sub>]. The latter IL was measured in many laboratories, mostly at  $T = 298.15$  K (ref. 23–26 and 28) and  $T = 293.15$  K.<sup>27</sup> The density of [BzMIM][NTf<sub>2</sub>] in literature differs slightly from our data at higher temperatures.<sup>23</sup> The mean percent deviation for six temperatures in a range from 298.15 K to 323.15 K between our values and published in ref. 23 is  $-0.32\%$ , in ref. 24 (one point) is  $-0.13\%$  and in ref. 26 (one point) it is  $-4.76\%$ . The other literature data are at different temperatures, thus it is difficult to compare the results. The viscosity at  $T = 298.15$  K is  $102 \text{ mPa s}$  for [BzMIM][DCA] and  $133 \text{ mPa s}$  for [BzMIM][NTf<sub>2</sub>], which differ from the available literature values for [BzMIM][NTf<sub>2</sub>]<sup>23,24</sup> and is lower than those in other literature.<sup>34,36</sup> The mean percent deviation for two temperatures 298.15 K and 308.15 K between our values and published in ref. 23 is  $+17.5\%$ , in ref. 24 (one point) is  $+6\%$  and in ref. 26 (one point) it is  $+54\%$ . Undoubtedly, the purity of the compound in different publications is responsible for such a differences.

The average values of  $\gamma_{13}^\infty$  for the measured solutes in [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>], determined at six temperatures in a range from (318.15 to 368.15) K at atmospheric pressure are listed in Tables 3 and 4, respectively. Most of the virial coefficient values  $B_{11}$  and  $B_{12}$  and critical parameters for solutes used in this work were presented in our previous study.<sup>34</sup> An analysis of the values of  $\gamma_{13}^\infty$  shows large differences in possible interaction between a solute and the IL at infinite dilution. The data for [BzMIM][NTf<sub>2</sub>] shows almost five times lower values of  $\gamma_{13}^\infty$  for the non-polar solutes in comparison with [BzMIM][DCA], which means much larger interaction between the [NTf<sub>2</sub>]<sup>-</sup> anion with hydrocarbons than those of the [DCA]<sup>-</sup> anion. Aromatic hydrocarbons show much smaller differences; the values of  $\gamma_{13}^\infty$  for benzene, toluene and others are only two times smaller for the [NTf<sub>2</sub>]<sup>-</sup> anion. For polar substances, such as alcohols, the interaction between solute and the IL is similar





Table 6 The experimental gas–liquid partition coefficients  $K_L$  for the solutes in ionic liquid [BzMIM][NTf<sub>2</sub>] at different temperatures

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentane	4.95	3.96	3.23	2.67	2.25	1.91
Hexane	10.6	8.05	6.25	4.95	4.00	3.28
3-Methylpentane	10.1	7.66	5.96	4.73	3.83	3.15
2,2-Dimethylbutane	6.59	5.18	4.15	3.39	2.81	2.36
Heptane	21.8	15.8	11.7	8.85	6.85	5.41
Octane	43.1	29.9	21.5	15.7	11.8	9.09
2,2,4-Trimethylpentane	20.9	15.2	11.3	8.64	6.73	5.34
Nonane	87.2	57.5	39.3	27.6	20.0	14.8
Decane	170	108	71.0	48.2	33.7	24.2
Cyclopentane	14.8	11.4	8.91	7.10	5.77	4.75
Cyclohexane	29.7	22.1	16.8	13.0	10.3	8.31
Methylcyclohexane	42.3	30.5	22.6	17.1	13.2	10.4
Cycloheptane	88.8	62.1	44.7	32.9	24.9	19.1
Cyclooctane	233	156	107	75.8	55.1	40.9
Pent-1-ene	8.31	6.45	5.10	4.12	3.38	2.82
Hex-1-ene	17.4	13.0	9.85	7.65	6.07	4.89
Cyclohexene	58.3	41.8	30.8	23.2	17.8	14.0
Hept-1-ene	35.8	25.3	18.4	13.7	10.4	8.10
Oct-1-ene	70.5	47.9	33.5	24.2	17.8	13.5
Dec-1-ene	268	168	109	72.7	49.9	35.2
Pent-1-yne	31.7	23.1	17.2	13.1	10.2	8.06
Hex-1-yne	65.5	45.7	32.8	24.0	18.0	13.8
Hept-1-yne	131	87.8	60.4	42.7	31.0	23.0
Oct-1-yne	261	167	111	75.4	52.9	38.0
Benzene	294	197	136	96.5	70.2	52.2
Toluene	590	381	254	174	123	88.3
Ethylbenzene	1067	661	424	281	191	133
<i>o</i> -Xylene	1645	1003	634	415	279	193
<i>m</i> -Xylene	1204	739	469	307	207	144
<i>p</i> -Xylene	1150	706	450	296	200	139
<i>n</i> -Propylbenzene	1845	1102	682	438	288	196
Iso-propylbenzene	1455	877	548	354	236	161
Styrene	2741	1627	1003	640	421	285
$\alpha$ -Methylstyrene	4036	2336	1398	867	555	365
Thiophene	398	265	182	128	92.9	68.7
Pyridine	2330	1451	935	620	424	297
Methanol	147	103	74.4	54.8	41.3	31.8
Ethanol	221	149	104	74.3	54.6	40.8
Propan-1-ol	450	289	193	132	93.5	68.1
Propan-2-ol	242	162	112	79.7	58.2	43.9
Butan-1-ol	949	584	374	248	170	120
Butan-2-ol	463	294	194	133	94.1	68.5
2-Methyl-1-propanol	626	391	256	173	121	87.3
<i>tert</i> -Butanol	251	163	110	76.9	55.4	41.2
1-Pentanol	1989	1172	722	461	306	209
Water	275	189	133	96.3	71.1	53.8
Methyl acetate	226	153	106	76.5	56.1	42.2
Methyl propanoate	404	262	175	121	86.1	62.5
Methyl butanoate	687	434	283	191	132	94.2
Ethyl acetate	363	237	159	111	78.6	57.4
Tetrahydrofuran	295	199	138	98.7	72.1	53.8
1,4-Dioxane	1383	865	558	371	253	177
<i>tert</i> -Butyl methyl ether	54.2	38.0	27.4	20.3	15.4	11.9
<i>tert</i> -Butyl ethyl ether	45.8	31.7	22.5	16.5	12.3	9.40
<i>tert</i> -Amyl methyl ether	109	74.1	51.8	37.2	27.3	20.5
Diethyl ether	27.3	19.9	14.9	11.4	8.87	7.04
Di- <i>n</i> -propyl ether	72.0	49.1	34.5	25.0	18.5	14.0
Di- <i>iso</i> -propyl ether	34.9	24.6	17.8	13.2	9.96	7.67
Di- <i>n</i> -butyl ether	266	168	110	74.6	51.9	37.0
Acetone	340	232	162	117	85.7	64.3
Pentan-2-one	1072	675	440	296	204	145



Table 6 (Contd.)

Solute	T/K					
	318.15	328.15	338.15	348.15	358.15	368.15
Pentan-3-one	1049	661	430	288	198	139
Butanal	371	249	172	122	88.8	66.1
Acetonitrile	732	497	346	247	181	135
1-Nitropropane	3265	2007	1276	838	566	394

and for water it is much lower for the  $[\text{NTf}_2]^-$  anion. This confirms the common information about hydrophobicity of  $[\text{NTf}_2]^-$ -based ILs which is the most important for consideration for the three separation cases, *viz.* hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene. The values for the mentioned solutes at  $T = 328.15$  K: for hexane  $\gamma_{13}^\infty = 95.4$  or  $\gamma_{13}^\infty = 17.6$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively; for hex-1-ene  $\gamma_{13}^\infty = 38.1$  or  $\gamma_{13}^\infty = 9.27$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively; for cyclohexane  $\gamma_{13}^\infty = 33.5$  or  $\gamma_{13}^\infty = 9.39$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively; for cyclohexene  $\gamma_{13}^\infty = 13.6$  or  $\gamma_{13}^\infty = 5.31$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively; for ethylbenzene  $\gamma_{13}^\infty = 6.71$  or  $\gamma_{13}^\infty = 2.26$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively; and for styrene  $\gamma_{13}^\infty = 3.11$  or  $\gamma_{13}^\infty = 1.32$  in  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$ , respectively. For all these solutes, the interaction with the IL is almost three times larger for  $[\text{BzMIM}][\text{NTf}_2]$ . Thus, the selectivities for separation may be expected to not be very different from each other for the two ILs.

The global analysis drawn from Tables 3 and 4 shows that the  $\gamma_{13}^\infty$  values increase with the alkyl chain length in a series of alkanes, cycloalkanes, alkenes, alkynes, aromatic hydrocarbons (increasing radicals), alcohols, ethers, and ketones, which is an indication of the decrease of interactions between the solute and the IL at infinite dilution. Furthermore, despite the lower number of alkyl chains, the higher measured retention time is an evident observation of the decrease of the  $\gamma_{13}^\infty$  values, and of an increase in the interactions between the IL and solute. The lowest values of  $\gamma_{13}^\infty$  at all temperatures (the strongest interaction) in  $[\text{BzMIM}][\text{DCA}]$  are observed for water ( $\gamma_{13}^\infty = 0.387$ ), methanol ( $\gamma_{13}^\infty = 0.511$ ), ethanol ( $\gamma_{13}^\infty = 0.872$ ), acetonitrile ( $\gamma_{13}^\infty = 0.903$ ), and pyridine ( $\gamma_{13}^\infty = 1.03$ ) at  $T = 328.15$  K. The lowest values of  $\gamma_{13}^\infty$  at all temperatures (the largest interaction) in  $[\text{BzMIM}][\text{NTf}_2]$  are observed for acetone ( $\gamma_{13}^\infty = 0.410$ ), acetonitrile ( $\gamma_{13}^\infty = 0.455$ ), 1,4-dioxane ( $\gamma_{13}^\infty = 0.523$ ), tetrahydrofuran ( $\gamma_{13}^\infty = 0.648$ ), 1-nitropropane ( $\gamma_{13}^\infty = 0.692$ ), and pentan-2-one ( $\gamma_{13}^\infty = 0.693$ ) at  $T = 328.15$  K. As can be seen, the strongest interactions for the two different anions is observed for different solutes. Only acetonitrile is common for the two ILs, with the value two times larger in  $[\text{BzMIM}][\text{DCA}]$  than that in  $[\text{BzMIM}][\text{NTf}_2]$ .

The largest values of  $\gamma_{13}^\infty$  provide the initial information about the lower interactions. For  $[\text{BzMIM}][\text{DCA}]$  it may be observed for decane ( $\gamma_{13}^\infty = 509$ ), dec-1-ene ( $\gamma_{13}^\infty = 220$ ) and di-*n*-butyl ether ( $\gamma_{13}^\infty = 77.2$ ) at  $T = 328.15$  K. The largest values of  $\gamma_{13}^\infty$  for  $[\text{BzMIM}][\text{NTf}_2]$  are observed for decane ( $\gamma_{13}^\infty = 70.7$ ) and

dec-1-ene ( $\gamma_{13}^\infty = 36.8$ ) at  $T = 328.15$  K. These latter values are appreciably lower than those for  $[\text{BzMIM}][\text{DCA}]$ .

An interesting feature worth noting is a comparison to two ILs,  $[\text{AMIM}][\text{DCA}]^{16}$  and  $[\text{AMIM}][\text{NTf}_2]^{20}$  which show the same relations in the interactions for the measured solutes. Strong interactions for  $[\text{AMIM}][\text{DCA}]$  were observed with water and methanol ( $\gamma_{13}^\infty = 0.341$ ,  $\gamma_{13}^\infty = 0.505$  at  $T = 328.15$  K), all alcohols as well as with thiophene, pyridine, 1,4-dioxane, acetone, acetonitrile, and 1-nitropropane with  $\gamma_{13}^\infty < 2$  at  $T = 328.15$  K.<sup>16</sup> The  $\gamma_{13}^\infty$  data for the  $[\text{NTf}_2]^-$ -based ILs show much smaller interactions with water and alcohols. Nonetheless, the lower values of  $\gamma_{13}^\infty$  for polar solutes such as benzene, thiophene, pyridine, and 1-nitropropane for the  $[\text{NTf}_2]^-$ -based ILs suggest the high potential for the extraction of these compounds from alkanes, which is important in petrochemical processes, *e.g.* desulphurization and denitrification of fuels. The hydrocarbons, such as alkanes, alkenes and alkynes reveal much lower values of  $\gamma_{13}^\infty$  in  $[\text{AMIM}][\text{NTf}_2]^{20}$  than those observed earlier for  $[\text{AMIM}][\text{DCA}]^{16}$  ( $\gamma_{13}^\infty < 100$  at  $T = 328.15$  K). The  $\gamma_{13}^\infty$  values in  $[\text{AMIM}][\text{NTf}_2]$  were  $\gamma_{13}^\infty = 20.8$  and  $\gamma_{13}^\infty = 10.7$  at  $T = 328.15$  K for hexane and hex-1-ene, respectively.

The experimental  $\gamma_{13}^\infty$  values exhibit diverse temperature dependences corresponding on both endothermic and exothermic effects accompanying the interactions of solutes with the IL. The  $\gamma_{13}^\infty$  values decrease with an increasing temperature for alkanes, alkenes, cycloalkanes and cycloalkenes (see Fig. 2S and 3S in ESI<sup>†</sup>). The opposite influence of temperature is presented for alkynes (see Fig. 3S in ESI<sup>†</sup>) and for aromatic hydrocarbons (see Fig. 4S in ESI<sup>†</sup>). Only for the case of alcohols and water in  $[\text{BzMIM}][\text{NTf}_2]$  do the  $\gamma_{13}^\infty$  values decrease with an increasing temperature (see Fig. 5Sb in ESI<sup>†</sup>). For alcohols in  $[\text{BzMIM}][\text{DCA}]$  (see Fig. 5Sa in ESI<sup>†</sup>) and most of the ethers (see Fig. 6S in ESI<sup>†</sup>), esters, tetrahydrofuran, 1,4-dioxane, thiophene (see Fig. 7S in ESI<sup>†</sup>) and ketones, as well as 1-nitropropane (see Fig. 8S in ESI<sup>†</sup>) there is an increase in  $\gamma_{13}^\infty$  values with an increasing temperature. The diagrams mentioned above show the plot of the natural logarithm of the  $\gamma_{13}^\infty$  as a function of the inverse absolute temperature for all investigated solutes.

The gas-liquid partition coefficient,  $K_L$ , calculated from eqn (6) is an important property of the IL. This property shows the suitability of the IL for particular application in extraction. The data for  $[\text{BzMIM}][\text{DCA}]$  and  $[\text{BzMIM}][\text{NTf}_2]$  are listed in Tables 5 and 6. By inspection of these tables one can see that the lowest values are observed for alkanes ( $K_L = 2.19$  at  $T = 328.15$  K in  $[\text{BzMIM}][\text{DCA}]$ ), cycloalkanes, cycloalkenes, alkenes, alkynes



Table 7 Limiting partial molar excess Gibbs energies,  $\Delta G_1^{E,\infty}$  enthalpies  $\Delta H_1^{E,\infty}$  and entropies  $T_{\text{ref}}\Delta S_1^{E,\infty}$  for the solutes in investigated ionic liquids [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] at the reference temperature  $T_{\text{ref}} = 328.15$  K

Solute	[BzMIM][DCA]			[BzMIM][NTf <sub>2</sub> ]		
	$\Delta G_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta H_1^{E,\infty}/\text{kJ mol}^{-1}$	$T_{\text{ref}}\Delta S_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta G_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta H_1^{E,\infty}/\text{kJ mol}^{-1}$	$T_{\text{ref}}\Delta S_1^{E,\infty}/\text{kJ mol}^{-1}$
Pentane	11.4	8.20	-3.19	6.98	4.03	-2.95
Hexane	12.4	7.77	-4.67	7.82	4.15	-3.67
3-Methylpentane	12.2	8.41	-3.75	7.46	3.39	-4.07
2,2-Dimethylbutane	12.6	11.69	-0.90	7.41	3.50	-3.90
Heptane	13.6	8.40	-5.16	8.76	4.80	-3.96
Octane	14.7	8.74	-5.98	9.72	6.21	-3.52
2,2,4-Trimethylpentane	14.4	9.10	-5.26	8.77	4.14	-4.63
Nonane	15.8	9.67	-6.17	10.6	6.63	-4.00
Decane	17.0	10.45	-6.55	11.6	7.77	-3.85
Cyclopentane	8.31	4.75	-3.56	5.15	2.55	-2.60
Cyclohexane	9.58	5.99	-3.59	6.11	3.98	-2.13
Methylcyclohexane	10.9	6.87	-4.02	7.00	3.82	-3.17
Cycloheptane	10.1	6.12	-4.01	6.80	4.11	-2.69
Cyclooctane	10.9	6.61	-4.27	7.50	5.07	-2.43
Pent-1-ene	8.75	3.63	-5.12	5.15	0.70	-4.45
Hex-1-ene	9.93	4.36	-5.57	6.08	2.25	-3.83
Cyclohexene	7.12	2.03	-5.10	4.56	1.29	-3.27
Hept-1-ene	11.1	4.61	-6.50	6.98	2.03	-4.95
Oct-1-ene	12.3	5.28	-7.07	7.98	3.50	-4.47
Dec-1-ene	14.7	6.79	-7.93	9.84	4.47	-5.37
Pent-1-yne	4.58	-2.55	-7.13	2.46	-2.72	-5.18
Hex-1-yne	5.71	-1.69	-7.40	3.34	-2.04	-5.38
Hept-1-yne	6.96	-0.85	-7.81	4.26	-1.32	-5.58
Oct-1-yne	8.20	0.22	-7.98	5.17	-0.41	-5.58
Benzene	2.43	-3.21	-5.64	0.13	-4.13	-4.26
Toluene	3.82	-2.96	-6.78	1.20	-3.42	-4.61
Ethylbenzene	5.19	-1.85	-7.04	2.22	-2.91	-5.13
<i>o</i> -Xylene	4.66	-2.44	-7.10	1.95	-2.73	-4.67
<i>m</i> -Xylene	5.29	-2.16	-7.45	2.25	-3.39	-5.64
<i>p</i> -Xylene	5.21	-2.23	-7.44	2.27	-3.32	-5.60
<i>n</i> -Propylbenzene	6.54	-1.13	-7.68	3.22	-2.45	-5.67
Iso-propylbenzene	6.52	-1.93	-8.44	3.16	-2.37	-5.54
Styrene	3.10	-3.80	-6.90	0.76	-4.53	-5.29
$\alpha$ -Methylstyrene	4.41	-4.65	-9.06	1.86	-4.93	-6.79
Thiophene	0.99	-4.33	-5.32	-0.29	-3.82	-3.53
Pyridine	0.081	-3.87	-3.96	-1.83	-4.39	-2.56
Methanol	-1.83	-1.36	0.48	0.67	4.37	3.70
Ethanol	-0.37	0.14	0.52	1.30	5.14	3.84
Propan-1-ol	0.40	0.40	0.00	1.84	5.14	3.30
Propan-2-ol	0.82	1.42	0.60	1.63	6.57	4.94
Butan-1-ol	1.38	1.06	-0.32	2.50	6.27	3.77
Butan-2-ol	1.53	0.55	-0.97	2.09	5.61	3.52
2-Methyl-1-propanol	1.42	1.71	0.29	2.43	7.20	4.76
<i>tert</i> -Butanol	1.60	0.62	-0.99	1.62	4.33	2.71
1-Pentanol	2.20	0.59	-1.61	2.97	6.04	3.07
Water	-2.59	-2.06	0.53	2.97	8.16	5.19
Methyl acetate	1.81	-3.98	-5.79	-1.23	-4.91	-3.68
Methyl propanoate	2.90	-3.82	-6.72	-0.56	-5.20	-4.64
Methyl butanoate	4.13	-2.74	-6.87	0.36	-4.03	-4.39
Ethyl acetate	3.19	-3.26	-6.45	-0.50	-5.03	-4.53
Tetrahydrofuran	1.57	-3.78	-5.36	-1.18	-5.13	-3.95
1,4-Dioxane	0.40	-5.56	-5.96	-1.77	-6.72	-4.95
<i>tert</i> -Butyl methyl ether	6.66	-1.21	-7.87	2.40	-4.03	-6.43
<i>tert</i> -Butyl ethyl ether	9.33	0.57	-8.76	4.45	-2.59	-7.04
<i>tert</i> -Amyl methyl ether	7.73	-0.53	-8.26	3.37	-2.67	-6.04
Diethyl ether	6.27	-0.49	-6.75	2.38	-3.30	-5.68
Di- <i>n</i> -propyl ether	9.42	1.78	-7.65	4.88	-0.80	-5.68
Di- <i>iso</i> -propyl ether	9.91	2.49	-7.42	4.79	-1.62	-6.41
Di- <i>n</i> -butyl ether	11.9	4.17	-7.69	6.82	1.91	-4.91
Acetone	0.43	-4.15	-4.58	-2.43	-5.33	-2.89



Table 7 (Contd.)

Solute	[BzMIM][DCA]			[BzMIM][NTf <sub>2</sub> ]		
	$\Delta G_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta H_1^{E,\infty}/\text{kJ mol}^{-1}$	$T_{\text{ref}}\Delta S_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta G_1^{E,\infty}/\text{kJ mol}^{-1}$	$\Delta H_1^{E,\infty}/\text{kJ mol}^{-1}$	$T_{\text{ref}}\Delta S_1^{E,\infty}/\text{kJ mol}^{-1}$
Pentan-2-one	2.60	-3.45	-6.04	-1.00	-5.25	-4.25
Pentan-3-one	2.63	-3.59	-6.22	-0.98	-5.86	-4.87
Butanal	1.95	-3.63	-5.57	-0.93	-4.53	-3.60
Acetonitrile	-0.28	-2.45	-2.17	-2.15	-3.53	-1.39
1-Nitropropane	1.25	-2.84	-4.09	-1.00	-2.93	-1.92

and ethers, especially for diethyl ether in [BzMIM][DCA] ( $K_L = 7.05$  at  $T = 328.15$  K). The large values of  $K_L$  are observed for 1-pentanol ( $K_L = 2290$  at  $T = 328.15$  K in [BzMIM][DCA]), water ( $K_L = 2139$  at  $T = 328.15$  K in [BzMIM][DCA]) and  $\alpha$ -methylstyrene ( $K_L = 1358$  at  $T = 328.15$  K in [BzMIM][DCA] and  $K_L = 2336$  at  $T = 328.15$  K in [BzMIM][NTf<sub>2</sub>]) as well as for 1-nitropropane ( $K_L = 2007$  at  $T = 328.15$  K in [BzMIM][NTf<sub>2</sub>]). These values are much lower than those observed for [AMIM][DCA]<sup>16</sup> and similar to those observed for [AMIM][NTf<sub>2</sub>].<sup>20</sup> The high  $K_L$  value corresponds to a large affinity of the solute to the liquid phase. The  $K_L$  value increases with a decrease of temperature and with an increase of the alkane chain length for alkanes, alkenes, alkynes, cycloalkanes, alcohols, esters and ethers. The  $K_L$  values increase with an increase of the radicals in the aromatic compounds.

Table 7 lists the partial molar excess Gibbs energies,  $\Delta G_1^{E,\infty}$ , at infinite dilution, the molar excess enthalpies,  $\Delta H_1^{E,\infty}$ , at infinite dilution, and the partial molar excess entropies at infinite dilution,  $T_{\text{ref}}\Delta S_1^{E,\infty}$ , for all the solutes studied in [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>], respectively, at a reference temperature  $T = 328.15$  K. These thermodynamic functions specify the interaction between solute and the IL and are an important consideration in determining the suitability of the IL for extraction. The  $\Delta G_1^{E,\infty}$  was calculated from the temperature dependence of  $\gamma_{13}^\infty$ . The values of  $\Delta G_1^{E,\infty}$  are positive for all solutes except methanol, ethanol, water and acetonitrile in [BzMIM][DCA], as well as thiophene, pyridine, esters, tetrahydrofuran, 1,4-dioxane, ketones, acetonitrile and 1-nitropropane in [BzMIM][NTf<sub>2</sub>]. This is similar to the earlier measured data for [AMIM][DCA]<sup>16</sup> and [AMIM][NTf<sub>2</sub>].<sup>20</sup> The infinite dilution activity coefficient values are lower than one for these solutes, *i.e.*  $\gamma_{13}^\infty < 1$ , which corresponds to the (IL + solute) binary system with negative deviations from Raoult's law. The largest negative value of  $\Delta G_1^{E,\infty}$  was observed for water in [BzMIM][DCA] ( $-2.59$  kJ mol<sup>-1</sup>) and for acetone in [BzMIM][NTf<sub>2</sub>] ( $-2.43$  kJ mol<sup>-1</sup>). For the remaining solutes, including hexane ( $\Delta G_1^{E,\infty} = 12.48$  kJ mol<sup>-1</sup> in [BzMIM][DCA]) and hex-1-ene ( $\Delta G_1^{E,\infty} = 9.93$  kJ mol<sup>-1</sup> in [BzMIM][DCA]), positive deviations from ideality are obtained. The large positive values are observed for long-chain alkanes in [BzMIM][DCA], which indicates that the solute-IL interaction is not strong and limited miscibility is possible in the liquid phase.

The partial excess molar enthalpies at infinite dilution,  $\Delta H_1^{E,\infty}$ , determined from the Gibbs-Helmholtz equation exhibit negative values for alkynes, aromatic hydrocarbons, esters some

ethers, ketones and 1-nitropropane for both ILs. For these solutes relatively strong energetic solute-solvent interactions are observed. Additionally, there are negative  $\Delta H_1^{E,\infty}$  values for [BzMIM][DCA] with water and methanol. As expected, the [NTf<sub>2</sub>]<sup>-</sup>, the hydrophobic anion of the IL used in this work, exhibits possible  $\pi$ - $\pi$ , or  $n$ - $\pi$  interactions with most of the solutes, which leads to negative values of  $\Delta H_1^{E,\infty}$ . For the other solutes, no particular affinity between the IL and solute molecule is expected. Thus for aliphatic hydrocarbons, the endothermic interaction, resulting from the energetic weakness of their interaction with the IL is observed.

The partial excess molar entropies at infinite dilution,  $T_{\text{ref}}\Delta S_1^{E,\infty}$ , are small and negative for all solutes studied, excluding alcohols and water in [BzMIM][NTf<sub>2</sub>], and excluding a few alcohols and water in [BzMIM][DCA]. The solution of the majority of the solutes in both ILs is accompanied by entropy losses, which may suggest that the solute molecule arranges itself in the IL structure. The most positive value of the  $T_{\text{ref}}\Delta S_1^{E,\infty}$  term was observed for water,  $0.53$  kJ mol<sup>-1</sup> and  $5.19$  kJ mol<sup>-1</sup> for [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>], respectively, whilst for the previously measured [AMIM][DCA]<sup>16</sup> it was  $11.79$  kJ mol<sup>-1</sup>.

### Separation of hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene

To analyze the performance of the [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] ILs as extraction solvents for the chosen separation processes, the characteristic parameters for the separation, the selectivity ( $S_{12}^\infty = \gamma_1^\infty/\gamma_2^\infty$ ) and the capacity ( $k_2^\infty = 1/\gamma_2^\infty$ ) were calculated from the experimental activity coefficients at infinite dilution values. The results are listed in Table 8 for both ILs for the three separation cases: hexane (1)/hex-1-ene (2), cyclohexane (1)/cyclohexene (2), and ethylbenzene (1)/styrene (2) at  $T = 328.15$  K, along with literature data for ILs with the same anions, [DCA]<sup>-</sup> and [NTf<sub>2</sub>]<sup>-</sup>. To study the effect of the anion and cation structure on the extractive properties, the selectivity of 65 ILs have been compared to [BzMIM][DCA] and [BzMIM][NTf<sub>2</sub>] at the same temperature. In addition, to evaluate the potential of the ILs to be used at industrial scale, a comparison of capacity has also been performed.

Even though the experimental data present large discrepancies among different literature sources, one can conclude that the selectivities for the hexane (1)/hex-1-ene (2) separation, obtained with benzyl-based ILs is quite high with the best value for the imidazolium-based IL, [AMIM][DCA] ( $S_{12}^\infty = 2.51$ ).<sup>16</sup> The



**Table 8** Selectivities,  $S_{12}^{\infty}$ , and capacities,  $k_2^{\infty}$  for hexane/hexane, cyclohexane/cyclohexene and ethylbenzene/styrene separation problems for [DCA]<sup>−</sup> and [NTf<sub>2</sub>]<sup>−</sup> based ionic liquids at  $T = 328.15$  K

Ionic liquid abbreviation	$S_{12}^{\infty}$			$k_{12}^{\infty}$			Ref.
	Hexane/hexene	Cyclohexane/cyclohexene	Ethylbenzene/styrene	Hexene	Cyclohexene	Styrene	
[BzMIM][DCA]	2.50	2.46	2.16	0.03	0.07	0.32	This work
[BzMIM][NTf <sub>2</sub> ]	1.90	1.77	1.71	0.11	0.19	0.76	This work
[EMIM][DCA]	2.46	—	—	0.02	—	—	17
[AMIM][DCA]	2.51	2.51	2.24	0.02	0.05	0.25	16
[BMIM][DCA]	2.29	2.19	2.14	0.04	0.09	0.44	15
[BMPY][DCA]	2.34	2.32	2.31	0.05	0.12	0.72	35
[CPMIM][DCA]	2.30	—	—	0.02	—	—	18
[CPMMIM][DCA]	2.16	—	—	0.04	—	—	18
[N-C <sub>3</sub> OHMIM][DCA]	2.79	—	—	0.01	—	—	36
[MMIM][NTf <sub>2</sub> ]	2.10	—	—	0.07	—	—	37
[EMIM][NTf <sub>2</sub> ]	2.02	1.81	—	0.09	0.14	—	37
[EMIM][NTf <sub>2</sub> ]	3.02	—	—	0.13	—	—	38
[EMIM][NTf <sub>2</sub> ]	—	1.86	1.68	—	1.86	1.68	39
[EMMIM][NTf <sub>2</sub> ]	—	1.86	1.83	—	0.14	0.66	39
[AMIM][NTf <sub>2</sub> ]	1.94	1.79	1.75	0.09	0.15	0.63	20
[BMIM][NTf <sub>2</sub> ]	1.80	—	—	0.15	—	—	37
[BMIM][NTf <sub>2</sub> ]	2.13	—	—	0.08	—	—	18
[BMIM][NTf <sub>2</sub> ]	1.82	1.65	—	0.15	0.22	—	40
[HMIM][NTf <sub>2</sub> ]	1.67	—	—	0.23	—	—	41
[HMIM][NTf <sub>2</sub> ]	1.72	1.57	—	0.24	0.33	—	42
[HMIM][NTf <sub>2</sub> ]	1.53	1.56	—	0.23	0.64	—	43
[OMIM][NTf <sub>2</sub> ]	1.53	3.51	—	0.31	1.52	—	43
[D <sub>2</sub> MIM][NTf <sub>2</sub> ]	1.31	—	—	0.72	—	—	44
[C <sub>6</sub> H <sub>13</sub> OCH <sub>2</sub> MIM][NTf <sub>2</sub> ]	1.63	—	—	0.22	—	—	45
[(C <sub>6</sub> H <sub>13</sub> OCH <sub>2</sub> ) <sub>2</sub> IM][NTf <sub>2</sub> ]	1.41	—	—	0.47	—	—	45
[C <sub>2</sub> OHMIM][NTf <sub>2</sub> ]	1.99	—	—	0.05	—	—	36
[COC <sub>2</sub> MIM][NTf <sub>2</sub> ]	1.89	—	—	0.10	—	—	36
[(CO) <sub>2</sub> IM][NTf <sub>2</sub> ]	2.13	—	—	0.05	—	—	36
[CPMIM][NTf <sub>2</sub> ]	2.41	—	—	0.02	—	—	18
[CPMMIM][NTf <sub>2</sub> ]	2.16	—	—	0.04	—	—	19
[COC <sub>2</sub> MMOR][NTf <sub>2</sub> ]	2.13	1.97	1.82	0.07	0.12	0.57	46
[N-C <sub>3</sub> OHMMOR][NTf <sub>2</sub> ]	2.37	2.58	2.07	0.02	0.05	0.28	47
[PMPPI][NTf <sub>2</sub> ]	1.99	—	—	0.11	—	—	48
[BMPPI][NTf <sub>2</sub> ]	1.75	—	—	0.17	—	—	49
[PeMPIP][NTf <sub>2</sub> ] <sup>a</sup>	1.73	1.60	—	0.19	0.28	—	50
[HMPPI][NTf <sub>2</sub> ]	1.63	1.56	—	0.23	0.34	—	50
[COC <sub>2</sub> MPIP][NTf <sub>2</sub> ]	1.90	1.77	1.75	0.12	0.19	1.75	51
[BCN <sup>4</sup> PY][NTf <sub>2</sub> ]	2.13	2.07	1.85	0.07	0.14	0.69	52
[1,4BMPY][NTf <sub>2</sub> ]	1.93	—	—	0.15	—	—	53
[EPY][NTf <sub>2</sub> ]	2.20	1.95	—	0.08	0.13	—	54
[EPY][NTf <sub>2</sub> ]	1.08	—	—	0.14	—	—	55
[BPY][NTf <sub>2</sub> ]	1.37	—	—	0.16	—	—	55
[PePY][NTf <sub>2</sub> ] <sup>a</sup>	1.46	—	—	0.23	—	—	55
[N-C <sub>3</sub> OHpy][NTf <sub>2</sub> ]	2.15	2.04	—	0.05	0.10	—	56
[PMPYR][NTf <sub>2</sub> ]	2.00	—	—	0.13	—	—	21
[BMPYR][NTf <sub>2</sub> ]	1.72	—	—	0.15	—	—	43
[BMPYR][NTf <sub>2</sub> ]	1.86	—	—	0.14	—	—	21
[PeMPYR][NTf <sub>2</sub> ] <sup>a</sup>	1.86	—	—	0.17	—	—	21
[HMPYR][NTf <sub>2</sub> ]	1.53	—	—	0.22	—	—	57
[HMPYR][NTf <sub>2</sub> ]	1.65	—	—	0.22	—	—	58
[OMPYR][NTf <sub>2</sub> ]	1.52	—	—	0.31	—	—	57
[OMPYR][NTf <sub>2</sub> ]	1.58	—	—	0.33	—	—	58
[DMPYR][NTf <sub>2</sub> ]	1.46	—	—	0.38	—	—	58
[COC <sub>2</sub> MPYR][NTf <sub>2</sub> ]	1.91	1.78	1.7	0.11	0.17	1.70	59
[C <sub>8</sub> iQuin][NTf <sub>2</sub> ]	1.28	—	—	0.32	—	—	60
[Et <sub>3</sub> S][NTf <sub>2</sub> ]	2.09	—	—	0.09	—	—	61
[N <sub>1,1,1,2</sub> OH][NTf <sub>2</sub> ]	2.22	2.03	0.97	0.04	0.06	0.24	62
[N <sub>1,1,1,4</sub> ][NTf <sub>2</sub> ]	1.84	0.57	—	0.11	0.16	—	63
[N <sub>1,1,1,6</sub> ][NTf <sub>2</sub> ]	0.88	—	—	0.84	—	—	17
[N <sub>1,1,1,8</sub> ][NTf <sub>2</sub> ]	1.54	—	—	0.26	—	—	64





Table 8 (Contd.)

Ionic liquid abbreviation	$S_{12}^{\infty}$			$k_{12}^{\infty}$			Ref.
	Hexane/hexene	Cyclohexane/cyclohexene	Ethylbenzene/styrene	Hexene	Cyclohexene	Styrene	
[N <sub>1,1,1,10</sub> ][NTf <sub>2</sub> ]	1.50	—	—	0.36	—	—	64
[N <sub>1,4,4,4</sub> ][NTf <sub>2</sub> ]	1.55	—	—	0.23	—	—	64
[N <sub>1,8,8,8</sub> ][NTf <sub>2</sub> ]	1.27	—	—	1.16	—	—	65
[N <sub>2,2,2,8</sub> ][NTf <sub>2</sub> ]	1.54	1.48	1.65	0.28	0.40	1.39	66
[N <sub>8,8,8,8</sub> ][NTf <sub>2</sub> ]	1.23	—	—	0.95	—	—	64
[P <sub>6,6,6,1,4</sub> ][NTf <sub>2</sub> ]	1.51	—	—	2.04	—	—	67
[P <sub>6,6,6,1,4</sub> ][NTf <sub>2</sub> ]	1.21	—	—	1.09	—	—	68
[P <sub>6,6,6,1,4</sub> ][NTf <sub>2</sub> ]	1.19	—	—	0.98	—	—	69

<sup>a</sup> Pe-pentyl.

selectivities for [BzMIM][DCA] are much larger than those for [BzMIM][NTf<sub>2</sub>] for all separation cases. However, the capacity is two to three times larger for [BzMIM][NTf<sub>2</sub>] in comparison with [BzMIM][DCA]. The largest values of selectivity for the hexane (1)/hex-1-ene (2) separation are for [DCA]<sup>−</sup>-based ILs at  $T = 328.15$  K which are: 1-(3-hydroxypropyl)methylimidazolium dicyanamide, [N-C<sub>3</sub>OHMIM][DCA] ( $S_{12}^{\infty} = 2.79$ )<sup>36</sup> > [AMIM][DCA] ( $S_{12}^{\infty} = 2.51$ )<sup>16</sup> > [BzMIM][DCA] ( $S_{12}^{\infty} = 2.50$ ). The largest values of selectivity for the cyclohexane (1)/cyclohexene (2) separation are for [DCA]<sup>−</sup>-based ILs at  $T = 328.15$  K which are: [AMIM][DCA] ( $S_{12}^{\infty} = 2.51$ )<sup>16</sup> > [BzMIM][DCA] ( $S_{12}^{\infty} = 2.46$ ). The largest values of selectivity for the ethylbenzene (1)/styrene (2) separation are for [DCA]<sup>−</sup>-based ILs at  $T = 328.15$  K which are: 1-butyl-4-methylpyridinium dicyanamide, [BMPY][DCA] ( $S_{12}^{\infty} = 2.31$ )<sup>35</sup> > [AMIM][DCA] ( $S_{12}^{\infty} = 2.24$ )<sup>16</sup> > [BzMIM][DCA] ( $S_{12}^{\infty} = 2.16$ ). The largest values of selectivity for the hexane (1)/hex-1-ene (2) separation for [NTf<sub>2</sub>]<sup>−</sup>-based ILs at  $T = 328.15$  K are: 1-ethyl-3-methyl imidazolium bis((trifluoromethyl)sulfonyl)imide, [EMIM][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 3.02$ )<sup>38</sup> > [CPMIM][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.41$ )<sup>18</sup> > 1-(3-hydroxypropyl)methylmorpholinium bis((trifluoromethyl)sulfonyl)imide, [N-C<sub>3</sub>OHMOR][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.37$ )<sup>47</sup> > [N<sub>1,1,1,2</sub>OH][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.22$ )<sup>62</sup> > [EPY][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.20$ ).<sup>54</sup> The largest values of selectivity for the cyclohexane (1)/cyclohexene (2) separation for [NTf<sub>2</sub>]<sup>−</sup>-based ILs at  $T = 328.15$  K are: 1-octyl-3-methyl imidazolium bis((trifluoromethyl)sulfonyl)imide, [OMIM][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 3.51$ )<sup>43</sup> > [N-C<sub>3</sub>OHMOR][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.58$ ).<sup>47</sup> The largest values of selectivity for the ethylbenzene (1)/styrene (2) separation for [NTf<sub>2</sub>]<sup>−</sup>-based ILs at  $T = 328.15$  K are: [N-C<sub>3</sub>OHMOR][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 2.07$ )<sup>47</sup> > [BCN<sup>4</sup>PY][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 1.85$ )<sup>52</sup> > [EMMIM][NTf<sub>2</sub>] ( $S_{12}^{\infty} = 1.83$ ).<sup>39</sup>

Summing up, for all three separation cases, the [DCA]<sup>−</sup>-based ILs reveal larger selectivities than those for [NTf<sub>2</sub>]<sup>−</sup>-based ILs (with the exception of [EMIM][NTf<sub>2</sub>],<sup>38</sup> which is probably an experimental error in comparison with other results of the same IL). This fact reveals that the interaction of unsaturated hydrocarbons and aromatic hydrocarbons with [DCA]<sup>−</sup>-based ILs is larger than that with aliphatic hydrocarbons. Unfortunately, the high selectivity of the above-mentioned ILs are accompanied with low capacity. Among the 67 ILs listed in Table 8, the capacity is in a range ( $k_2^{\infty}$  is from 0.02 to 2.04) for hex-1-ene, ( $k_2^{\infty}$  is from 0.05 to 1.86) for cyclohexene, and ( $k_2^{\infty}$  is

from 0.24 to 1.75) for styrene at  $T = 328.15$  K. The highest capacity for the investigated ILs is exhibited by [BzMIM][NTf<sub>2</sub>] for hexene ( $k_2^{\infty} = 0.11$ ) and for styrene ( $k_2^{\infty} = 0.76$ ). The introduction of a benzyl group in the imidazolium cation has slightly increased the capacity, especially for styrene, with the selectivity on the same level in comparison to [AMIM][DCA]<sup>16</sup> and with slightly lower selectivity in comparison with [AMIM][NTf<sub>2</sub>].<sup>20</sup> The replacement of the ethyl group by the allyl group at the cation for [NTf<sub>2</sub>]<sup>−</sup>-based ILs has caused the opposite effect on the selectivity in the ethylbenzene (1)/styrene (2) separation, but increases the capacity for cyclohexene ( $k_2^{\infty} = 1.86$ )<sup>39</sup> and for styrene ( $k_2^{\infty} = 1.68$ ).<sup>39</sup>

## Conclusions

In this work, we have studied the potential use of 1-benzyl-3-methylimidazolium dicyanamide, [BzMIM][DCA] and 1-benzyl-3-methylimidazolium bis((trifluoromethyl)sulfonyl)imide, [BzMIM][NTf<sub>2</sub>] ILs in three processes for separation, *viz.* hexane/hex-1-ene, cyclohexane/cyclohexene, and ethylbenzene/styrene. For that purpose, experimental measurements were performed for the activity coefficients at infinite dilution for 64 solutes in both ILs at six temperatures ranging between (318.15 and 368.15) K. In addition, thermal properties, densities, and viscosities of both ILs were also measured as a function of temperature. The interactions of various types of organic solutes and water with two ionic liquids at infinite dilution were discussed and shown with regard to the activity coefficients, the gas-liquid-partition coefficients and the thermodynamic functions. Using the reported experimental data, along with other data from literature, the impact of the ILs cation and anion in the three chosen separation problems was analyzed. The reported results show high values of selectivity and low values of capacity for [BzMIM][DCA] and the opposite trend for [BzMIM][NTf<sub>2</sub>] ionic liquid as a separation agent. However, [BzMIM][DCA] has exhibited adequate extraction properties for all separation processes with high values of selectivities which are similar to [AMIM][DCA]<sup>16</sup> measured by us earlier. Moreover, densities and viscosities of [BzMIM][DCA] were lower than those of [BzMIM][NTf<sub>2</sub>] and slightly higher than those for [AMIM][DCA].<sup>16,27</sup> Therefore, in this context, the conclusion must be



made that the presence of the benzyl group in the imidazolium cation has not considerably increased the selectivity in both ILs in comparison with the allyl group, but has increased the capacity, especially for styrene.

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## References

- N. V. Plechkova and K. S. Seddon, *Chem. Soc. Rev.*, 2008, **37**, 123–150.
- J. P. Hallet and T. Welton, *Chem. Rev.*, 2011, **111**, 3508–3576.
- U. Domańska, *Ionic Liquids in Chemical Analysis, General Review of Ionic Liquids and Their Properties*, CRC Press, Taylor & Francis Group, Abingdon, UK, 2008, ch. 1.
- A. Cole, *Ionic Liquids in Separation Technology*, Elsevier Science and Technology, Research and Markets, 2014, <http://www.researchandmarkets.com/reports/2899469/ionic-liquids-in-separation-technology.htm>.
- G. W. Meindersma, A. R. Hansmeier and A. B. de Haan, *Ind. Eng. Chem. Res.*, 2010, **49**, 7530–7540.
- A. R. Hansmeier, G. W. Meindersma and A. B. de Haan, *Green Chem.*, 2011, **13**, 1907–1913.
- K. Kędra-Królik, F. Mutelet and J.-N. Joubert, *Ind. Eng. Chem. Res.*, 2011, **50**, 2296–2306.
- U. Domańska and M. Wlazło, *Fuel*, 2014, **134**, 114–125.
- A. Nann, C. Held and G. Sadowski, *Ind. Eng. Chem. Res.*, 2013, **52**, 18472–18481.
- M. Bülow, C. J. Guo, D. Shen, F. R. Fitch, A. I. Shirley and V. A. Malik, *Separation of alkenes and alkanes, US Pat.*, 6200366, 2001, [http://www.google.com/patents/US6200366?utm\\_source=gb-gplus-sharePatent](http://www.google.com/patents/US6200366?utm_source=gb-gplus-sharePatent).
- J. Bergh, C. Gücüyener, E. A. Pidko, E. J. M. Hensen, J. Gascon and F. Kapteijn, *Chem.–Eur. J.*, 2011, **17**, 8832–8840.
- D. Peralta, G. Chaplais, A. Simon-Masseron, K. Barthelet, C. Chizallet, A. A. Quoineaud and G. D. Pirngruber, *J. Am. Chem. Soc.*, 2012, **134**, 8115–8126.
- V. Mokrushin, D. Assenbaum, N. Paape, D. Gerhard, L. Mokrushina, P. Wasserscheid, W. Arlt, H. Kistenmacher, S. Neuendorf and V. Göke, *Chem. Eng. Technol.*, 2010, **33**, 63–73.
- Z. Lei, W. Arlt and P. Wasserscheid, *Fluid Phase Equilib.*, 2006, **241**, 290–299.
- U. Domańska, M. Wlazło and M. Karpińska, *Fluid Phase Equilib.*, 2016, **417**, 50–61.
- M. Wlazło, J. Gawkowska and U. Domańska, *Ind. Eng. Chem. Res.*, 2016, **55**, 5054–5056.
- F. Mutelet, A. L. Revelli, J.-N. Jaubert, L. M. Sprunger, W. E. Acree Jr and G. A. Baker, *J. Chem. Eng. Data*, 2010, **55**, 234–242.
- J. Zhang, Q. Zhang, B. Qiao and Y. Deng, *J. Chem. Eng. Data*, 2007, **52**, 2277–2283.
- U. Domańska, E. V. Lukoshko and M. Królikowski, *J. Chem. Thermodyn.*, 2013, **61**, 126–131.
- M. Wlazło, M. Karpińska and U. Domańska, *J. Chem. Thermodyn.*, 2016, **102**, 39–47.
- F. Mutelet, E.-S. R. E. Hassan, T. W. Stephens, W. E. Acree Jr and G. A. Baker, *J. Chem. Eng. Data*, 2013, **58**, 2210–2218.
- P. S. Kulkarni, L. C. Branco, J. G. Crespo, M. C. Nunes, A. Raymundo and C. A. M. Afonso, *Chem.–Eur. J.*, 2007, **13**, 8478–8488.
- S. V. Dzyuba and R. A. Bartsch, *ChemPhysChem*, 2002, **3**, 161–166.
- T. Mandai, A. Matsumura, M. Imanari and K. Nishikawa, *J. Phys. Chem. B*, 2012, **116**, 2090–2095.
- H. Shirota, H. Matsuzaki, S. Ramati and J. F. Wishart, *J. Phys. Chem. B*, 2015, **119**, 9173–9187.
- S. M. Mahurin, T. Dai, J. S. Yeary, H. Luo and S. Dai, *Ind. Eng. Chem. Res.*, 2011, **50**, 14061–14069.
- M. Larriba, P. Navarro, M. Gonzalez-Miquel, S. Omar, J. Palomar, J. García and F. Rodríguez, *Chem. Eng. Res. Des.*, 2016, **109**, 561–572.
- B. Gabri, A. Sander, M. C. Bubalo and D. Macut, *Sci. World J.*, 2013, **2013**, 512953, DOI: 10.1155/2013/512953.
- C. A. Ohlin, P. J. Dyson and G. Laurency, *Chem. Commun.*, 2004, 1070–1071.
- D. H. Everett, *Trans. Faraday Soc.*, 1965, **61**, 1637–1645.
- A. J. B. Cruickshank, B. W. Gainey, C. P. Hicks, T. M. Letcher, R. W. Moody and C. L. Young, *Trans. Faraday Soc.*, 1969, **65**, 1014–1031.
- Design institute for physical properties, sponsored by AIChE (2005; 2008; 2009; 2010). Dippr project 801-full version. Design institute for physical property research/aiche. Online version available at: [http://knovel.com/web/portal/browse/display?\\_ext\\_knovel\\_display\\_bookid=1187&verticalid=0](http://knovel.com/web/portal/browse/display?_ext_knovel_display_bookid=1187&verticalid=0).
- B. E. Poling and J. M. Prausnitz, *Properties of Gases and Liquids*, McGraw-Hill Publishing, 2001, <http://lib.myilibrary.com?ID=91317>.
- U. Domańska, M. Królikowska, W. E. Acree Jr and G. A. Baker, *J. Chem. Thermodyn.*, 2011, **43**, 1050–1057.
- M. Królikowski and M. Królikowska, *J. Chem. Thermodyn.*, 2014, **68**, 138–144.
- A.-L. Revelli, F. Mutelet, J.-N. Jaubert, M. Garcia-Martinez, L. M. Sprunger, W. E. Acree Jr and G. A. Baker, *J. Chem. Eng. Data*, 2010, **55**, 2434–2443.
- M. Krummen, P. Wasserscheid and J. Gmehling, *J. Chem. Eng. Data*, 2002, **47**, 1411–1417.
- N. Deenadajalu, T. M. Letcher and P. Reddy, *J. Chem. Eng. Data*, 2005, **50**, 105–108.
- A. Heintz, D. V. Kulikov and S. P. Verevkin, *J. Chem. Eng. Data*, 2002, **47**, 894–899.
- A. Heintz, L. M. Casás, I. A. Nesterov, V. N. Emel'yanenko and S. P. Verevkin, *J. Chem. Eng. Data*, 2005, **50**, 1510–1514.



- 41 T. M. Letcher, A. Marciniak, M. Marciniak and U. Domańska, *J. Chem. Thermodyn.*, 2005, **37**, 1327–1331.
- 42 A. Heintz and S. P. Verevkin, *J. Chem. Eng. Data*, 2006, **51**, 434–437.
- 43 R. Kato and J. Gmehling, *J. Chem. Thermodyn.*, 2005, **37**, 603–619.
- 44 J.-C. Moise, F. Mutelet, J. N. Jaubert, L. M. Grubbs, W. E. Acree Jr and G. A. Baker, *J. Chem. Eng. Data*, 2011, **56**, 3106–3114.
- 45 U. Domańska and A. Marciniak, *Fluid Phase Equilib.*, 2009, **286**, 154–161.
- 46 A. Marciniak and M. Wlazło, *J. Chem. Thermodyn.*, 2012, **47**, 382–388.
- 47 M. Wlazło, A. Marciniak, M. Zawadzki and B. Dudkiewicz, *J. Chem. Thermodyn.*, 2015, **86**, 154–161.
- 48 U. Domańska and K. Padaszyński, *J. Chem. Thermodyn.*, 2010, **42**, 1361–1366.
- 49 K. Padaszyński and U. Domańska, *J. Phys. Chem. B*, 2011, **115**, 8207–8215.
- 50 K. Padaszyński and U. Domańska, *J. Chem. Thermodyn.*, 2013, **60**, 169–173.
- 51 A. Marciniak and M. Wlazło, *J. Chem. Thermodyn.*, 2012, **49**, 137–145.
- 52 M. Wlazło, M. Karpińska and U. Domańska, *J. Chem. Thermodyn.*, 2016, **97**, 253–260.
- 53 U. Domańska and A. Marciniak, *J. Chem. Thermodyn.*, 2009, **41**, 1350–1355.
- 54 R. Kato and J. Gmehling, *Fluid Phase Equilib.*, 2004, **226**, 37–44.
- 55 P.-F. Yan, Q.-S. Liu, M. Yang, X.-M. Liu, Z.-C. Tan and U. Welz-Biermann, *J. Chem. Thermodyn.*, 2010, **42**, 1415–1422.
- 56 A. Marciniak, *J. Chem. Thermodyn.*, 2011, **43**, 1446–1452.
- 57 S. Nebig, V. Liebert and J. Gmehling, *Fluid Phase Equilib.*, 2009, **277**, 61–67.
- 58 W. E. Acree Jr, G. A. Baker, A.-L. Revelli, J.-Ch. Moise and F. Mutelet, *J. Chem. Eng. Data*, 2012, **57**, 3510–3518.
- 59 A. Marciniak and M. Wlazło, *J. Chem. Thermodyn.*, 2012, **54**, 90–96.
- 60 U. Domańska, M. Zawadzki, M. Królikowska, M. M. Tshibangu, D. Ramjugernath and T. M. Letcher, *J. Chem. Thermodyn.*, 2011, **43**, 499–504.
- 61 U. Domańska and A. Marciniak, *J. Chem. Thermodyn.*, 2009, **41**, 754–758.
- 62 U. Domańska, P. Papis and J. Szydłowski, *J. Chem. Thermodyn.*, 2014, **77**, 63–70.
- 63 A. Heintz, T. V. Vassitsova, J. Safarov, E. Bich and S. P. Verevkin, *J. Chem. Eng. Data*, 2006, **51**, 648–655.
- 64 W. E. Acree Jr, G. A. Baker, F. Mutelet and J.-C. Moise, *J. Chem. Eng. Data*, 2011, **56**, 3688–3697.
- 65 N. V. Gwala, N. Deenadayalu, K. Tumba and D. Ramjugernath, *J. Chem. Thermodyn.*, 2010, **42**, 256–261.
- 66 M. Wlazło and U. Domańska, *Sep. Purif. Technol.*, 2016, **162**, 162–170.
- 67 T. Banerjee and A. Khanna, *J. Chem. Eng. Data*, 2006, **51**, 2170–2177.
- 68 K. Tumba, T. M. Letcher, P. Naidoo and D. Ramjugernath, *J. Chem. Thermodyn.*, 2013, **65**, 159–167.
- 69 T. M. Letcher, D. Ramjugernath, M. Laskowska, M. Królikowski, P. Naidoo and U. Domańska, *J. Chem. Eng. Data*, 2008, **53**, 2044–2049.

