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TEMPO-catalyzed oxidative homocoupling route to 3,2'-biindolin-2-ones via an indolin-3-one intermediate†

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A combinative C2-selective arylation, and C3-selective carbonylation of free indole derivatives, by means of TEMPO catalysis and a silver oxidant under non-directing group conditions, was successful demonstrated. This new methodology is both atom and step efficient and is applicable to a broad scope of substrates, allowing the synthesis of a range of synthetically valuable 3,2'-biindolin-2-ones in moderate to excellent yields.

Introduction

Biindole scaffolds are important motifs in an array of natural products with diverse biological activities,^{1–3} exemplified by the TCDD (2,3,7,8-tetrachlorodibenzo-*p*-dioxin) antagonist bisindigotin **1** (Fig. 1) from the Chinese medicinal herb *Isatis indigotica*.⁴ This herbal plant has long been used as a folk medicine in China for treatment of viral diseases and diseases with an inflammatory nature.⁴ In addition, indirubin, a bis-indole scaffold and its derivatives, are present in some traditional Chinese medicines and inhibit cyclin-dependent kinases (CDKs).⁵ These medicines have historically been used to treat chronic diseases including leukemia.⁶

Accordingly, synthesis and functionalization of biindolyls have attracted much attention over decades.^{7–9} Recently, some progress on construction of the 2,3'-linked⁷ and 3,3'-linked⁸ biindolyl scaffolds was made, including the palladium- or copper-catalyzed intermolecular coupling reaction and iodine-induced dimerization of indoles. For example, Zhang and co-workers reported a mild and selective method for dimerization of indoles by palladium catalysis to give 2,3'-biindolyls in high yields at room temperature.^{9b} However, most of these procedures require expensive metal catalysts and high loading of metal oxidants. In addition, the regioselectivity control of C2 arylation can be quite challenging under non-directing group conditions. From the viewpoints of atom economy, cost efficiency and green chemistry, atmospheric oxygen is obviously superior to other reagents, and thus represents the quintessential oxidant.¹⁰ In the past decades, most efforts have been

directed to the development of transition metal-based catalysts. In contrast, much less attention has been paid to the development of non-metallic oxidation systems, largely ignoring their inherent advantages.^{10c} The radical TEMPO (2,2,6,6-tetramethylpiperidine 1-oxyl radical) and its derivatives are well-established catalysts for oxidation processes, and now used extensively in organic synthesis and industrial applications as a mild, safe, and economical alternative to heavy metal reagents as highly selective oxidation catalysts for the production of pharmaceuticals, flavors, fragrances, agrochemicals, and a variety of other specialty chemicals.¹¹ As part of our ongoing investigations on environmentally benign, selective, and controllable C–H bond functionalization, we studied the TEMPO-catalyzed oxidative homocoupling of indoles in air.¹² Herein, we report the first successful example of TEMPO-catalyzed oxidative homocoupling of indoles affording substituted C2–C3' bisindole-3-ones **9** (Scheme 1).

Results and discussion

Preliminary studies were carried out at 80 °C in an open tube, using commercially available unsubstituted indole (**5a**) in the presence of TEMPO and K₃PO₄. A variety of transition-metal catalysts, including FeCl₃·6H₂O, FeSO₄, ZrCl₄, ZnCl₂, CoCl₂, CuI, CuCl, and Ag₂CO₃, were screened. It was found that Ag₂CO₃ was the most efficient catalyst for this reaction, which gave the desired product **9a** in 51% yield (Table 1, entry 8); while other transition-metal catalysts did not undergo the conversion under the reaction conditions. The structure of **9a** was confirmed by X-ray crystallography (see ESI†). Among silver sources tested, including Ag₂CO₃, AgNO₃, AgOTf, and AgBF₄, were tested in DMF using TEMPO as the oxidant at 80 °C for 12 h, and Ag₂CO₃ was found to be the most effective catalyst (Table 1, entries 8–12). The amount of Ag₂CO₃ also has a large influence on the yield of **9a**. In the absence of Ag₂CO₃ catalyst, the desired product was not obtained under these conditions, whereas, at

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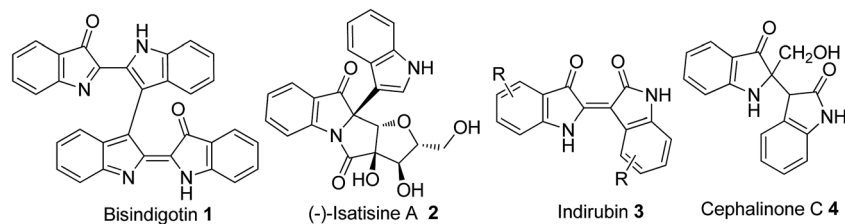


Fig. 1 Representative natural products with a 2-substituent indolin-3-one structural unit.

35 mol% of Ag_2CO_3 , the desired product is obtained in 57% yield, and at lower or higher loading, the yield decreases (Table 1, entries 12–16). It was also found that the use of TEMPO as an oxidant is critical to the reaction. In the absence of TEMPO, the homocoupling reaction did not proceed. However, the slightly more loading amounts of TEMPO result in a dramatic influence on the yield (Table 1, entries 17–20). For example, 81% yield of **9a** was obtained when using 0.15 mol% of TEMPO (Table 1, entry 19). Further assessment of the reaction solvents indicated that DMF was the optimal solvent, while other solvents gave lower yields or were ineffective (Table 1, entries 21–25).

Finally, we examined a series of bases (Table 1, entries 26–29). The observation revealed that NaOAc is slightly better than K_3PO_4 and other bases, such as LiOH, NaHCO_3 , and K_2CO_3 , are inferior to K_3PO_4 . After a great deal of screening on different parameters we found that the combinative C2-selective arylation and C3-selective carbonylation of indole by using TEMPO (15 mol%) in air as catalyst, Ag_2CO_3 (35 mol%) as an oxidant, and NaOAc (10 mol%) as base in DMF at 80 °C led to the highest efficiency (87% yield, Table 1, entry 26).

With a set of optimized conditions in hand, we next examined the indole scope of this TEMPO-catalyzed oxidative homocoupling reaction. As shown in Table 2, the reaction can tolerate a variety of functional groups at the 4, 5, 6, and 7 positions of indoles, such as F, Cl, Br, CH_3 , CH_3O , BnO, CO_2CH_3 , and CN, and the corresponding reactions proceeded smoothly to afford the desired products in moderate to excellent yields with high regioselectivity. The substituent effect on the indole ring was then investigated. The results have shown that electronegativities of substituents played a major role in

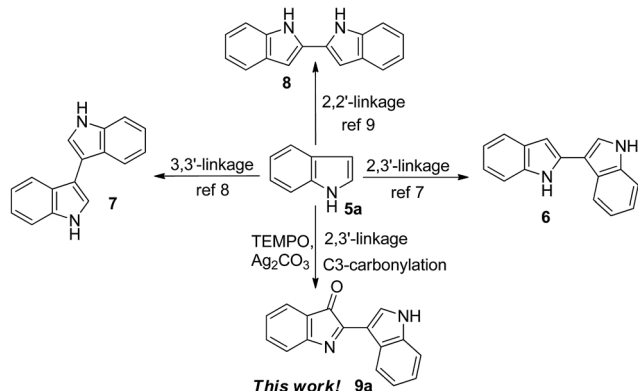
governing the reactivity of the substrates. Electron-donating substituents showed better results than electron-withdrawing substituents in this transformation. For example, 7-substituted indole derivatives with electron-donating substituents (CH_3 , OCH_3 , and OBn) afforded the desired **9l–n** in yields ranging from 76% to 90%, while 7-substituted indole derivatives with electron-withdrawing substituents (Cl and Br)

Table 1 Optimization of the reaction conditions^a

Entry	Catalyst (mol%)	TEMPO (mol%)	Solvent	Base (mol%)	Yield ^b (%)
1	$\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ (15)	10	DMF	K_3PO_4 (10)	0
2	FeSO_4 (15)	10	DMF	K_3PO_4 (10)	0
3	ZnCl_2 (15)	10	DMF	K_3PO_4 (10)	0
4	ZnCl_2 (15)	10	DMF	K_3PO_4 (10)	0
5	CoCl_4 (15)	10	DMF	K_3PO_4 (10)	0
6	CuI (15)	10	DMF	K_3PO_4 (10)	Trace
7	CuCl_2 (15)	10	DMF	K_3PO_4 (10)	Trace
8	Ag_2CO_3 (15)	10	DMF	K_3PO_4 (10)	51
9	AgNO_3 (15)	10	DMF	K_3PO_4 (10)	24
10	AgOTf (15)	10	DMF	K_3PO_4 (10)	35
11	AgBF_4 (15)	10	DMF	K_3PO_4 (10)	22
12	—	10	DMF	K_3PO_4 (10)	0
13	Ag_2CO_3 (5)	10	DMF	K_3PO_4 (10)	10
14	Ag_2CO_3 (25)	10	DMF	K_3PO_4 (10)	53
15	Ag_2CO_3 (35)	10	DMF	K_3PO_4 (10)	57
16	Ag_2CO_3 (45)	10	DMF	K_3PO_4 (10)	55
17	Ag_2CO_3 (35)	—	DMF	K_3PO_4 (10)	0
18	Ag_2CO_3 (35)	5	DMF	K_3PO_4 (10)	27
19	Ag_2CO_3 (35)	15	DMF	K_3PO_4 (10)	81
20	Ag_2CO_3 (35)	20	DMF	K_3PO_4 (10)	81
21	Ag_2CO_3 (35)	15	DMSO	K_3PO_4 (10)	67
22	Ag_2CO_3 (35)	15	Toluene	K_3PO_4 (10)	0
23	Ag_2CO_3 (35)	15	Pyridine	K_3PO_4 (10)	0
24	Ag_2CO_3 (35)	15	1,4-Dioxane	K_3PO_4 (10)	Trace
25	Ag_2CO_3 (35)	15	$\text{ClCH}_2\text{CH}_2\text{Cl}$	K_3PO_4 (10)	0
26	Ag_2CO_3 (35)	15	DMF	NaOAc (10)	87
27	Ag_2CO_3 (35)	15	DMF	LiOH (10)	0
28	Ag_2CO_3 (35)	15	DMF	NaHCO_3 (10)	30
29	Ag_2CO_3 (35)	15	DMF	K_2CO_3 (10)	48

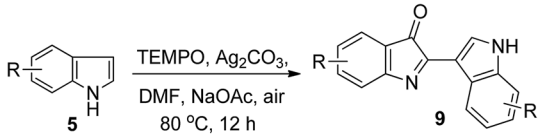
^a Condition: **5a** (0.3 mmol), solvent (1 mL), 80 °C, 12 h, under open air.

^b Isolated yields.



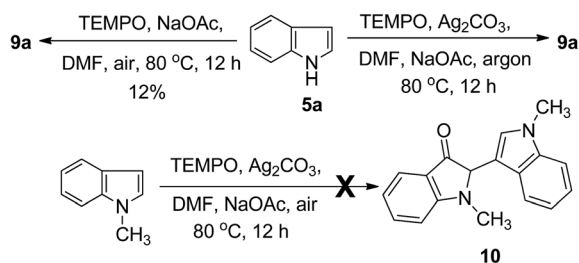
Scheme 1 Homocoupling of indole.



Table 2 Substrate scope^a


Entry	R	Product	Yield ^b [%]
1	H	9a	87
2	5-F	9b	66
3	5-Br	9c	75
4	5-CH ₃	9d	89
5	5-OCH ₃	9e	71
6	5-OBn	9f	78
7	6-F	9g	73
8	6-Cl	9h	64
9	6-CH ₃	9i	79
10	7-Cl	9j	61
11	7-Br	9k	72
12	7-CH ₃	9l	90
13	7-OCH ₃	9m	76
14	7-OBn	9n	82
15	4-CN	9o	38
16	4-CO ₂ CH ₃	9p	63

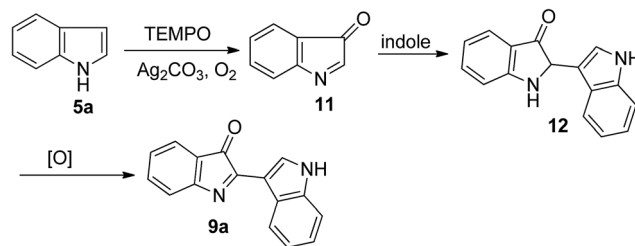
^a Reaction conditions: indole (0.3 mmol), TEMPO (15 mol%), Ag₂CO₃ (35 mol%), NaOAc (10 mol%), DMF (1 mL), 80 °C. ^b Isolated yields.



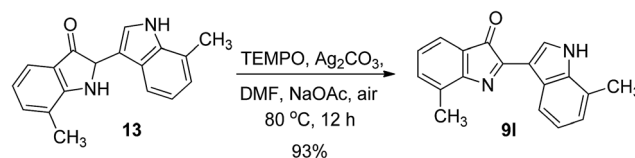
Scheme 2 Mechanistic studies.

provided the desired products in 61 and 72% yields, respectively. It is worth noting that substrate with a strong electron-withdrawing substituents at C4-position, such as CO₂CH₃ and CN, gave **9o** and **9p** in 63% and 38% yields, respectively. This is particularly important, since substrates with a strong electron-withdrawing group, such as a nitrile group, disfavored the homocoupling of indoles and there were few examples reported.^{7,8}

To gain some mechanistic insight into the process of this reaction, a series of control experiments were conducted (Scheme 2). Because the TEMPO-catalyzed oxidative homocoupling reaction was performed in air, the role of O₂ in this reaction was explored by conducting several control experiments. Under an O₂ atmosphere, the reaction yield was not increased, but a more rapid conversion of the starting material to the reaction product was observed by TLC detection compared to that performed under air conditions. However, only trace amount of the product was obtained under an argon atmosphere, even a long reaction time. These results indicated



Scheme 3 Plausible reaction pathway.



Scheme 4 Investigation of the reaction mechanism.

that O₂ is essential for the TEMPO-catalyzed transformation. In addition, in the control experiment of **5a** with TEMPO but without Ag₂CO₃, compound **9a** was obtained only in 12% yield. The result indicated that Ag₂CO₃ act as terminal oxidant. Under the optimized conditions, N-CH₃ indole was chosen as a substrate instead of indole. To our surprise, no conversion was observed, indicating that the substituents at the N1-position of the indole had a great influence on the reactivity.

Although the detailed mechanism remained unclear at the current stage, a plausible reaction pathway based on the basis of the results described above and relevant literature¹³ is outlined in Scheme 3. First, indole was oxidized slowly in the presence of TEMPO, Ag₂CO₃ and O₂ into indenone **11** which should be unstable and was never isolated.^{13a,d} Then, a rapid nucleophilic addition of another indole molecule on the C=N bond of this intermediate gave the intermediate **12**, which was oxidized rapidly to afford product **9a**.

To verify such a mechanistic scenario, we attempted to obtain the putative intermediate **12** or its derivatives. Fortunately, trace amounts of **13** can be determined in the reaction mixtures when 7-methyl-1*H*-indole was subjected to the standard reaction conditions, which showed us some clues on the reaction intermediate. In the next step, to obtain more information concerning the reaction pathway, we separated and collected the intermediate **13**, and subjected it to react with TEMPO under standard reaction conditions, and this gave product **9l** in 93% yield (Scheme 4). This result showed that **13** is the intermediate of the dimeric reaction.

Conclusions

In conclusion, we have developed a general and efficient method for the synthesis of 3,2'-biindolin-2-ones *via* a TEMPO-catalyzed oxidative homo dimerization in moderate to excellent yields with high regioselectivity. The advantages of this new method are broad substrate scope, operational simplicity, and high atom-economy. Moreover, the high halogen compatibility



of the process can provide a facile access to halo-substituted 3,2'-biindolin-2-ones.

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