Organic & Biomolecular Chemistry



PAPER View Article Online
View Journal | View Issue



Cite this: *Org. Biomol. Chem.*, 2017, **15**, 2397

Transition-metal-free synthesis of 3-(1-pyrrolidinyl)quinolines and 3-(1-pyrrolidinyl)quinoline 1-oxides *via* a one-pot reaction of 3-(1-pyrrolidinyl)crotonates with nitrobenzenes†

Robert Bujok,* Piotr Cmoch, Zbigniew Wróbel and Krzysztof Wojciechowski

A carbanion of tert-butyl 3-(1-pyrrolidinyl)crotonate adds to nitrobenzenes to form σ^H -adducts, which in the presence of pivaloyl chloride and triethylamine are converted into 3-(1-pyrrolidinyl)quinolines or 3-(1-pyrrolidinyl)quinoline 1-oxides depending on the nitrobenzene structure. This is the first methodology in which a quinoline ring is constructed from a substrate bearing a pyrrolidinyl ring. Starting from optically pure enamines, the method allows synthesis of the corresponding chiral products without racemisation.

Received 5th December 2016, Accepted 17th February 2017 DOI: 10.1039/c6ob02658c

rsc.li/obc

Introduction

Quinoline is a key building block in many naturally occurring and synthetically prepared compounds that have important practical applications. Thus, the synthesis of substituted quinolines is still of substantial interest. In recent years quinoline derivatives bearing a 1-pyrrolidinyl substituent at the 3-position have attracted attention due to their biological activity and potential applications as antimicrobials (antimalarial activity), hepatitis C virus polymerase inhibitors, antineoplastic (antitumor) agents, as well as compounds for osteoporosis treatment. Representative structures are shown in Fig. 1.

General methods for the synthesis of 3-(1-pyrrolidinyl)-quinolines consist of transition-metal-catalyzed replacement of halogen (mostly bromine) (Fig. 2, path **a**),^{4,7} reaction of 3-aminoquinoline with 4-chlorobutyryl chloride followed by reduction with LiAlH₄ (path **b**),³ ring-closing metathesis in 3-diallylaminoquinolines followed by hydrogenation (path **c**),³ and alkylation of 3-aminoquinoline with 1,4-dibromobutane (path **d**).³ One of the drawbacks of the recently developed methods for the synthesis of heterocycles, including quinolines, is the use of transition metal catalysts, which often requires arduous removal of the catalyst, particularly if the product has potential pharmaceutical applications.⁸ Thus new transition-metal-free methodologies are still in demand.

Institute of Organic Chemistry Polish Academy of Sciences, ul. Kasprzaka 44/52, 01-224 Warsaw, Poland. E-mail: robert.bujok@icho.edu.pl, krzysztof.wojciechowski@icho.edu.pl

†Electronic supplementary information (ESI) available. See DOI: 10.1039/c6ob02658c

Quinoline 1-oxides are usually obtained *via* direct oxidation of quinolines with peracids. Such oxidation will be ineffective if the quinoline derivative contains an amino group. Less widespread are methods in which the quinoline 1-oxides are formed as a result of interaction of the nitro group with adjacent substituents in *ortho*-substituted nitroarenes. To the best of our knowledge, the synthesis of 3-(1-pyrrolidinyl)quinoline-1-oxides is still unexplored.

During our studies on reactions of carbanions with nitroarenes we have developed numerous methods for the synthesis of heterocycles, especially indoles and quinolines. 10,11 Particularly effective for the synthesis of quinolines were methods in which the σ^{H} -adducts formed by addition of allyl carbanions to nitroarenes were subsequently treated with a silylating agent 12 such as trialkylchlorosilane or bis-trimethyl-silylacetamide (BSA) or with an acyl chloride, preferably pivaloyl chloride. 12c

Results and discussion

In this paper we present an efficient one-pot method for the synthesis of 3-(1-pyrrolidinyl)quinoline-2-carboxylic acid esters and the corresponding quinoline 1-oxides directly from nitrobenzenes. Unlike all the literature methods for the synthesis of 3-(1-pyrrolidinyl)quinoline, in this procedure a quinoline ring is constructed. Nitrobenzenes are cheap, commercially available starting materials, while quinoline derivatives are less abundant, so this synthesis is an interesting alternative to literature methodologies.

As a carbanion precursor we have chosen *tert*-butyl 3-(1-pyrrolidinyl)crotonate 1 prepared from *tert*-butyl acetoacetate and pyrrolidine in an excellent yield by simple mixing of

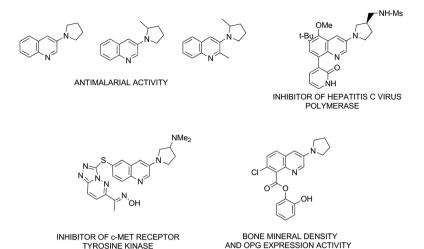


Fig. 1 Selected 3-(1-pyrrolidinyl)quinolines exhibiting biological activity.

Fig. 2 Synthetic strategies towards 3-(1-pyrrolidinyl)-quinolines.

equimolar amounts of the reagents without solvent (Scheme 1). ¹H NMR and nuclear Overhauser effect (1D-NOE) experiments proved the exclusive formation of the *E*-isomer.

As a model nitroarene for testing the reaction conditions we have chosen 2,4-dichloronitrobenzene, because earlier studies on vicarious nucleophilic substitution of hydrogen in nitroarenes revealed its relatively high electrophilicity. 13

In the initial experiment (Table 1) we used potassium tertbutoxide as a base for the generation of the carbanion and pivaloyl chloride to acylate the formed σ^H -adduct. The reaction was carried out in THF at −70 °C. Unfortunately, this attempt

Scheme 1 Synthesis of enamine 1.

Table 1 Effect of base on the yield of product 2a

Base	Yield ^a	
t-BuOK	0	
LDA	24	
n-BuLi	Traces of 2a (TLC)	
n-BuLi/NMP	12	
n-BuLi/TMEDA	18	
<i>n</i> -BuLi/HMPA	57	

^a The isolated yields are given.

was unsuccessful - no reaction with 2,4-dichloronitrobenzene occurred. Probably, potassium tert-butoxide is too weak a base to deprotonate the enamine 1. Indeed, when we used a stronger base, namely lithium diisopropylamide, the expected tert-butyl 6,8-dichloro-3-(1-pyrrolidinyl)quinoline-2-carboxylate 2a was obtained, but the yield was rather low (24%) and a large amount of by-products was formed. In the case of n-butyllithium only traces of the product 2a (TLC) were observed in the absence of any compounds, decreasing the degree of aggregation of the base.

n-Butyllithium exists as a tetramer¹⁴ in hexane/THF solution and HMPA is known to be the most effective reagent¹⁴ that breaks organo-lithium oligomers and increases the activity of the organolithium compound. Taking this into consideration, we carried out the reaction with *n*-butyllithium in

the presence of 1 equivalent of HMPA. In this case, treatment of the formed σ^H -adduct with triethylamine and pivaloyl chloride resulted in the formation of the ester 2a in 57% yield. Attempts to replace HMPA by other reagents (N-methylpyrrolidone (NMP) and TMEDA) were not successful – yields have not exceeded 20%. Because of the multistep nature of the investigated reaction, we have found the result obtained in the presence of HMPA to be satisfactory and applied these conditions 15 to screen the scope of the reaction (Table 2).

The reactions of 2,4-disubstituted nitrobenzenes with the enamine 1 led to the expected quinoline derivatives 2a-2e in 54-67% yields. As mentioned above, these are the overall yields of several steps.

Surprisingly, 4-substituted nitrobenzenes with no substituents at the 2-position react under analogous conditions to the enamine 1 to give *tert*-butyl 6-substituted 3-pyrrolidinyl-2-carboxylate-quinoline 1-oxides 3a–3h instead of the expected quinolines. The reaction is of general character – 4-substituted nitrobenzenes with electron-withdrawing (CF₃, Cl) and electron-donating groups (CH₃, *t*-Bu) give quinoline 1-oxides, usually in *ca.* 50% yield.

In the case of the 2-chloronitrobezene complex a mixture of compounds was formed and no attempt was made to separate them. The mixture was formed probably due to the attack of the carbanion on the 4-position of 2-chloronitrobezene.

To show the utility of our approach we have obtained carbanion precursors containing chiral pyrrolidines. Commercially available *N*-Boc-L-prolinol and *N*-Boc-D-prolinol were chosen for this purpose. Optically pure enamines (*S*)-1a and (*R*)-1a

Table 2 Synthesis of *tert*-butyl 3-(1-pyrrolidinyl)quinoline-2-carboxy-lates 2a-2e and *tert*-butyl 3-pyrrolidinyl-2-carboxylate-quinoline 1-oxides 3a-3h 15

Product	X	Y	Yield ^a	
			2	3
2a	Cl	Cl	57	_
2b	F	F	61	_
2c	Cl	OMe	67	_
2d	CH_3	Cl	54	_
2e	CH_3	F	55	_
3a	CF_3	H	_	38
3 b	Cl	H	_	50
3c	I	H	_	50
3 d	F	H	_	53
3e	SPh	H	_	50
3f	Ph	H	_	48
3g	CH_3	H	_	40
3h	t-Bu	Н	_	33

^a The isolated yields are given.

were prepared in very good overall yields (Scheme 5; for details see the ESI†). Some of the biologically active quinolines contain substituents at the 3-position of the pyrrolidine ring, so the corresponding enamines (S)-1b and (R)-1b were also obtained (Scheme 2).

Chiral enamines (*S*)-1a, (*R*)-1a, (*S*)-1b and (*R*)-1b react with nitrobenzenes in a similar manner to achiral enamine 1. Reaction with 2,4-disubstituted nitrobenzenes led to the corresponding non-racemic quinolines 4a–4d and reaction with 4-substituted nitrobenzenes led to non-racemic quinoline 1-oxides 5a–5c. The yields are somewhat lower than for achiral enamine 1, but the results are still acceptable considering the complex nature of the obtained structures (Table 3).

Scheme 2 Synthesis of optically pure enamines (5)-1a, (R)-1a, (5)-1b and (R)-1b (details are given in the ESI).

Table 3 Synthesis of optically pure tert-butyl 3-(1-pyrrolidinyl)-quino-line-2-carboxylates 4a-4d and tert-butyl 3-pyrrolidinyl-2-carboxylate-quinoline 1-oxides 5a-5c 15

					Yield ^a	
Product	X	Y	R	R'	4	5
4a	Cl	Cl	(S)-CH ₂ OBn	Н	37	
4b	Cl	OMe	(S)-CH ₂ OBn	Н	43	_
5a	I	Н	(S)-CH ₂ OBn	Н	_	35
5 b	\mathbf{F}	H	(S)-CH ₂ OBn	H	_	41
5c	F	H	(R)-CH ₂ OBn	Н	_	40
4c	Cl	OMe	Н	(S)-CH ₂ OBn	57	_
4d	F	F	H	(R)-CH ₂ OBn	40	_

^a The isolated yields are given.

To verify the optical purity of the non-racemic products, the R-enantiomer of quinoline 4b and the R-enantiomer of quinoline-1-oxide 5b were synthesized and both enantiomers of the examined compounds were converted to diastereomeric salts with R-Mosher's acid. Diagnostic signals of each salt formed from the corresponding enantiomer could be clearly recognized in ¹H NMR spectra. Only one diastereoisomer was detected in each case, which proved that all the examined compounds are optically pure (Fig. 3).

The surprising selectivity of the reactions (in all cases only quinoline or quinoline 1-oxide was obtained depending on the structure of the nitrobenzene; a mixture of these products was never observed) suggests that two different mechanisms operate. To gain an insight into the mechanisms, a few control experiments were carried out.

First of all, reactions of 2,4-dichloronitrobenzene and 4-chloronitrobenzene with enamine 1 under standard conditions but in the absence of pivaloyl chloride were carried out. In both cases no reaction was observed. These results indicate that σ^H -adducts formed from these nitrobenzenes do not undergo further reactions in the absence of t-BuCOCl and substrates are recovered after the hydrolysis of σ^H -adducts during aqueous work-up. So in the presence of pivaloyl chloride both σ^{H} -adducts are acylated and these intermediates are subsequently transformed into quinolines or quinoline 1-oxides depending on the character of the Y substituent (Scheme 3).

Acylated σ^H-adduct (I) probably undergoes elimination of the pivalate anion. The formed nitroso group enters an intramolecular Ehrlich-Sachs 12a,e reaction resulting in the formation of the corresponding 1-pivaloyloxy-dihydroquinoline derivative (II) (Scheme 4). When $Y \neq H$ due to steric interaction of the Y substituent and the bulky pivaloyloxy group, elimination of the pivalate anion takes place and quinoline 2a is

The formation of quinolone 1-oxides seems more difficult to explain. It could arise from nitro compound (III) (Scheme 5) via its intramolecular cyclization; 12c,16 however, experiments conducted in the absence of t-BuCOCl proved that the addition

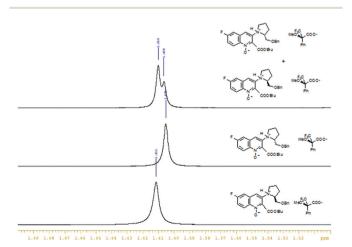


Fig. 3 Diagnostic signals of diastereoisomeric salts formed from 5b, 5c and R-Mosher's acid.

IN THE ABSENCE OF t-BuCOCI:

IN THE PRESENCE OF t-BuCOCI:

Scheme 3 Behaviour of σ^H -adducts in the absence of pivaloyl chloride and in the presence of the reagent.

Y=CI, F, OMe

Scheme 4 Plausible mechanism for the synthesis of guinoline 2a.

NO₂ COOt-Bu
$$t\text{-BuCOCl}$$

$$Et_3N, \text{THF, -70°C}$$

$$3a \overset{\circ}{\circ}$$

Attempted conversion of nitrocompound (III) into quinoline 1-oxide 3a under conditions employed for of synthesis of quinolines.

of the carbanion to nitroarene was reversible and the spontaneous oxidation of the σ^{H} -adduct did not occur (Scheme 3).

On the other hand the intermediate (III) might be formed from other than σ^H -adduct intermediates and then it could cyclize to quinoline 1-oxide 3a. To check this hypothesis an appropriate nitrocompound (III) was synthesized and exposed to reaction conditions (Scheme 5). Product 3a was not formed under these conditions - only the substrate and a moderate amount of polar by-products were observed. Thus quinoline 1-oxides are not formed *via* nitro-compound (III).

Based on these data, we suppose that 1-pivaloyloxy-dihydroquinoline derivative (II) is oxidized to 1-pivaloyloxy-quinoline

ox: eg p-chloronitrobenzene Nu: eg t-BuCOO-

Scheme 6 Plausible mechanism for synthesis of quinoline 1-oxide 3a.

(IV) by a nitrocompound (always used in excess) and then the pivaloyl group is attacked by a nucleophile (*e.g. t*-BuCOO⁻) leading to product **3a** (Scheme 6).

The *tert*-butoxycarbonyl group can be easily removed from compounds 2–3 by refluxing in 20% aqueous sulfuric acid. Hydrolysis and subsequent decarboxylation occurred, giving the expected products in good to excellent yields (81–97%) (Table 4).

The hydrolysis/decarboxylation domino reaction was also successfully performed for non-racemic quinoline 1-oxide 5a – the expected product 8a was obtained in 79% yield.

It is worth noting that treating a solution of *tert*-butyl esters **2a**, **2e** in dichloromethane with trifluoroacetic acid at room temperature resulted in the formation of the corresponding acids **9a**, **9e** in good yields of *ca.* 80% (Scheme 7). Elemental analysis indicated that compound **9a** was obtained as a free base (no fluorine present). In the case of compound **9e**, elemental analysis and ¹⁹F NMR spectroscopy suggest that the product contains "0.5 molecules" of trifluoroacetic acid.

Table 4 Hydrolysis/decarboxylation domino reaction of tert-butyl 3-(1-pyrrolidinyl)quinoline-2-carboxylates 2a-2d and tert-butyl 3-pyrrolidinyl-2-carboxylate-quinoline 1-oxides 3b-3h and $5a^{17}$

X	Y	R	Substrate	Product	Yield ^a
Cl	Cl	Н	2a	6a	93
F	F	H	2 b	6b	92
Cl	OMe	H	2c	6c	91
CH_3	Cl	H	2d	6d	93
Cl	H	H	3 b	7 b	81
I	H	H	3 c	7 c	87
F	H	H	3d	7d	97
Ph	H	H	3f	7 f	84
t-Bu	H	H	3h	7 h	91
I	H	(S)-CH ₂ OBn	5a	8a	79

^a The isolated yields are given.

CI CI N COOL-BU
$$CF_3COOH$$
 CI N COOH CI N COOH CI Sa_1 Sa_2 Sa_2 Sa_3 Sa_4 Sa_5 Sa_5

Scheme 7 Transformation of COOt-Bu group into COOH group (the isolated yields are given).¹⁸

Probably, 2 molecules of compound **9e** are coordinated by 1 molecule of trifluoroacetic acid.

Conclusions

We have developed a one-pot, transition-metal-free method for the synthesis of 3-(1-pyrrolidinyl)quinolines or 3-(1-pyrrolidinyl)-quinoline 1-oxides depending on the structure of the starting materials (2,4-disubstituted or 4-substituted nitrobenzenes). This is the first methodology in which a quinoline ring is constructed from a substrate bearing a pyrrolidinyl ring. The method allows direct synthesis of complicated structures from easily available enamines and commercial nitrobenzenes. Starting from optically pure enamines the method allows synthesis of the corresponding chiral products without racemisation.

Acknowledgements

This work was supported by National Scientific Center Grant No. 2012/07/B/ST5/00813. We would like to thank Dr Mikołaj Chromiński (also known as the Chemistry Wizard) for helpful discussions.

Notes and references

- (a) L. A. Mitscher, Chem. Rev., 2005, 105, 559; (b) M. Robert, J. Josef, K. Katarina and D. R. Richardson, Bioorg. Med. Chem., 2007, 15, 1280; (c) V. V. Kouznetsov, L. Y. V. Mendez and C. M. M. Gomez, Curr. Org. Chem., 2005, 9, 141; (d) K. Kaur, M. Jain, R. P. Reddy and R. Jain, Eur. J. Med. Chem., 2010, 45, 3245; (e) V. R. Solomon and H. Lee, Curr. Med. Chem., 2011, 18, 1488; (f) O. Afzal, S. Kumar, M. R. Haider, M. R. Ali, R. Kumar, M. Jaggi and S. Bawa, Eur. J. Med. Chem., 2015, 97, 871; (g) A. Marella, O. P. Tanwar, R. Saha, M. R. Ali, S. Srivastava, M. Akhter, M. Shaquiquzzaman and M. M. Alam, Saudi Pharm. J., 2013, 21, 1.
- 2 (a) M. Jose, P. Elena, S. Abdelouahid, C. Mariado and S. Elena, *Chem. Rev.*, 2009, **109**, 2652; (b) G. Liu, M. Yi,

- L. Liu, J. Wang and J. Wang, Chem. Commun., 2015, 51, 2911; (c) Y. Matsubara, S. Hirakawa, Y. Yamaguchi and Z. Yoshida, Angew. Chem., Int. Ed., 2011, 50, 7670; (d) G.-L. Gao, Y.-N. Niu, Z.-Y. Yan, H.-L. Wang, G.-W. Wang, A. Shaukat and Y.-M. Liang, J. Org. Chem., 2010, 75, 1305; (e) S. M. Prajapati, K. D. Patel, R. H. Vekariya, S. N. Panchal and H. D. Patel, RSC Adv., 2014, 4, 24463; (f) Y. Zhou, C. Wu, X. Dong and J. Qu, J. Org. Chem., 2016, 81, 5202-5208; (g) R. Bujok, Z. Wróbel and K. Wojciechowski, Tetrahedron Lett., 2016, 57, 1014-1018.
- Vandekerckhove, S. V. Herreweghe, J. Willems, B. Danneels, T. Desmet, C. de Kock, P. J. Smith, K. Chibaleand and M. D'hooghe, Eur. J. Med. Chem., 2015,
- 4 J. de Vicente Fidalgo, J. Li, R. C. Schoenfeld, F. X. Talamas and J. P. G. Taygerly, US 2010/311760A1, 2010.
- 5 A. G. Novartis, H. Deng, X. Fu, H. Guo, F. He, Y. Mi, X. Yan, H. Yu and J. Y. Zhang, WO 2012/107500A1, 2012.
- 6 J.-K. Kim, S.-W. Kim, S.-B. Shim, S.-Y. Ko, S. Chang, J.-S. Lee, D. P. Kang, H. C. Sung, J.-A. Park and Y. S. Yang, WO 2007/114672A1, Osotec Inc., 2007.
- 7 (a) E. F. DiMauro, J. L. Buchanan, A. Cheng, R. Emkey, S. A. Hitchcock, L. Huang, M. Y. Huang, B. Janosky, J. H. Lee, X. Li, M. W. Martin, S. A. Tomlinson, R. D. White, X. M. Zheng, V. F. Patel and R. T. Fremeau Jr., Bioorg. Med. Chem. Lett., 2008, 18, 4267; (b) F. X. Talamas, S. C. Abbot, S. Anand, K. A. Brameld, D. S. Carter, J. Chen, D. Davis, J. de Vicente, A. D. Fung, L. Gong, S. F. Harris, P. Inbar, S. S. Labadie, E. K. Lee, R. Lemoine, S. Le Pogam, V. Leveque, J. Li, J. McIntosh, I. Najera, J. Park, A. Railkar, S. Rajyaguru, M. Sangi, R. C. Schoenfeld, L. R. Staben, Y. Tan, J. P. Taygerly, A. G. Villasenor and P. E. Weller, J. Med. Chem., 2014, 57, 1914; (c) J. Cheng and M. L. Trudell, Org. Lett., 2001, 3, 1371.
- 8 C. E. Garrett and K. Prasad, Adv. Synth. Catal., 2004, 346, 889.
- 9 P. N. Preston and G. Tennant, Chem. Rev., 1972, 72, 627.
- 10 (a) M. Makosza and K. Wojciechowski, Heterocycles, 2001, 54, 445; (b) M. Makosza and K. Wojciechowski, Chem. Rev., 2004, 104, 2631; (c) M. Makosza and K. Wojciechowski, Heterocycles, 2014, 88, 75.
- 11 (a) M. Makosza and K. Wojciechowski, Top. Heterocycl. Chem., 2014, 37, 51; (b) M. Makosza and K. Wojciechowski, Chem. Heterocycl. Compd., 2015, 51, 210.

- 12 (a) K. Anczkiewicz, M. Królikiewicz, Z. Wróbel and Tetrahedron, Wojciechowski, 2015, 71, 3924: (b) M. Nowacki and K. Wojciechowski, RSC Adv., 2015, 5, 94296; (c) M. Bobin, A. Kwast and Z. Wróbel, Tetrahedron, 2007, 63, 11048; (d) Z. Wrobel, Synlett, 2004, 1929; (e) Z. Wróbel, Eur. J. Org. Chem., 2000, 521; (f) R. Bujok, A. Kwast, P. Cmoch and Z. Wróbel, Tetrahedron, 2010, 66, 698.
- 13 (a) S. Błażej and M. Makosza, Chem. Eur. J., 2008, 14, 11113; (b) F. Seelinger, S. Błażej, S. Bernhardt, M. Makosza and H. Mayr, Chem. - Eur. J., 2008, 14, 6108.
- 14 J. Clayden, Organolithium: Selectivity for Synthesis, Elsevier Science, 2002, pp. 2-7.
- 15 To a solution of the enamine 1 (0.317 g, 1.5 mmol) and HMPA (0.35 mL, 0.357 g, 2.0 mmol) in dry THF (10 mL) at -70 °C under argon, 2.5 M BuLi in hexane (0.67 mL, 1.67 mmol) was added in portions in 4 min and the solution was stirred at -70 °C for 30 min. A solution of the nitroarene (3.0 mmol) in THF (1.7 mL) was added and the resultant mixture was stirred for 10 min at -70 °C. Et₃N (1.0 mL, 7.2 mmol) was added and then pivalovl chloride (0.85 mL, 6.9 mmol) was added in portions (7 min). After the addition was complete, the solution was stirred at -70 °C for 2 h. The cooling bath was removed, water (10 mL) was added and the mixture was stirred for 5 min, then extracted with AcOEt (30 mL), dried and evaporated. The products were purified by column chromatography (SiO₂, hexane/AcOEt).
- 16 (a) Z. Wróbel, Tetrahedron Lett., 1997, 38, 4913; (b) Z. Wróbel, A. Kwast and M. Makosza, Synthesis, 1993, 31; (c) Z. Wróbel and M. Makosza, Tetrahedron, 1993, 49, 5315.
- 17 A suspension of compounds 2-5 (0.2-0.4 mmol) in 20% H₂SO₄ (2-4 mL) was heated to reflux for 3 h. The mixture was cooled to RT and diluted with water (15 mL) and solid K_2CO_3 was added in portions to achieve pH = 10. The mixture was extracted with AcOEt (2 × 25 mL), and the combined extracts were dried and evaporated. The crude products were washed with pentane or Et₂O. The obtained products were pure according to ¹H NMR.
- 18 To a solution of compound 2a (0.358 mmol) in CH₂Cl₂ (3.8 mL), trifluoroacetic acid (1.0 mL) was added and the solution was stirred at RT for 1 h. The solvent was evaporated and the crude product was washed with Et₂O. The obtained product was pure according to 1H NMR.