



Cite this: *New J. Chem.*, 2017, 41, 735

Vanadium and zinc complexes of 5-cyanopicolinate and pyrazine derivatives: synthesis, structural elucidation and *in vitro* insulino-mimetic activity study†

Tanja Koleša-Dobravc,^a Keiichi Maejima,^b Yutaka Yoshikawa,^c Anton Meden,^a Hiroyuki Yasui^b and Franc Perdih^{*a}

Vanadium(v) compounds with 5-cyanopicolinate acid (HpicCN), pyrazine-2-carboxylic acid (Hprz) and 3-aminopyrazine-2-carboxylic acid (HprzNH₂) and zinc compounds with HpicCN in the presence of 4-aminopyridine (4apy), pyridine (py) and 1,10-phenanthroline (phen) have been synthesized and characterized. The crystal structures of NH₄[VO₂(picCN)₂] (**3**), NH₄[VO₂(prz)₂] (**4**), NH₄[VO₂(przNH₂)₂·H₂O (**5**·H₂O), [Zn(picCN)₂(H₂O)₂] (**6**), [Zn(picCN)₂(4apy)₂]·C₇H₈ (**7**·C₇H₈), [Zn(picCN)₂(4apy)] (**8**), [Zn(picCN)₂(py)₂] (**9**) and [Zn(picCN)₂(phen)]·C₇H₈·2MeOH (**10**·C₇H₈·2MeOH) were determined by X-ray crystallography. The spatial arrangements of all the vanadium(v) complexes are similar, having carboxylate oxygen atoms in a mutual *trans* orientation. In the zinc bis(5-cyanopicolinate) complexes three different arrangements were found with *trans* (**6**) and a *cis* (**7**, **9** and **10**) octahedral and square-pyramidal (**8**) geometries. The insulino-mimetic activity of selected VO(IV), VO₂(V) and Zn(II) complexes was studied by *in vitro* inhibition of the free fatty acid (FFA) release from isolated rat adipocytes treated with epinephrine. All metal complexes showed insulino-mimetic activity and among them the VO(IV)–prz complex **2** was found to have higher insulino-mimetic activity than the positive control. The other vanadium compounds have activities similar to VOSO₄. The Zn complexes also exhibited some insulino-mimetic activity. Introduction of the N-donor 4apy to the zinc picCN complex **8** significantly increased inhibition of FFA release compared to **6**.

Received (in Victoria, Australia)
20th September 2016,
Accepted 6th December 2016

DOI: 10.1039/c6nj02961b

www.rsc.org/njc

Introduction

One of the most rapidly emerging diseases in the 21st century is thought to be diabetes mellitus (DM). The worldwide number of patients suffering from DM was estimated to be 382 million in 2013 representing ~8.3% of the adult population and is expected to increase to 592 million by 2035,¹ thus creating a significant health and economic burden. The prevalence of DM has been associated with ageing populations and the modern lifestyle, *i.e.* dietary changes, obesity, increasing urbanization,

reduced physical activity, and increased stress.^{1,2} DM is a condition characterized by abnormal glucose levels with a tendency for hyperglycaemia, due to a relative or absolute deficiency of insulin. DM is classified as type 1, type 2, or gestational diabetes. Type 2 represents the majority of diabetic patients (~90%) and is caused by insulin resistance or degradation of secreted insulin. It can be treated by dietary control, regular exercise and several chemotherapeutics or insulin. The most common treatment for diabetes is insulin. To some extent also oral hypoglycaemic drugs are used, however, they are often associated with undesirable side effects or a drop in efficacy is observed after prolonged use.³ Thus, the development of novel orally active compounds as an alternative to insulin injections as well as for the existing antidiabetics is still required and among the suitable candidates are zinc and vanadium coordination compounds.

In general, zinc and vanadium possess distinctly different biological roles in living organisms. Zinc is one of the most important trace elements in biological systems and is essential for the growth and development of microorganisms, plants and animals, and plays an important role in hundreds of metallo-enzymes and in thousands of metalloproteins.⁴ In humans, zinc is the second most abundant trace element with an estimated

^a Faculty of Chemistry and Chemical Technology, University of Ljubljana, Večna pot 113, P. O. Box 537, SI-1000 Ljubljana, Slovenia.
E-mail: franc.perdih@fkkt.uni-lj.si

^b Department of Analytical and Bioinorganic Chemistry, Division of Analytical and Physical Chemistry, Kyoto Pharmaceutical University, 5 Nakauchi-cho, Misasagi, Yamashina-ku, Kyoto 607-8414, Japan

^c Department of Health, Sports, and Nutrition, Faculty of Health and Welfare, Kobe Women's University, 4-7-2 Minatojima-nakamachi, Chuo-ku, Kobe, 650-0046, Japan

† Electronic supplementary information (ESI) available. CCDC 1480528–1480535 (3–10). For ESI and crystallographic data in CIF or other electronic formats see DOI: 10.1039/c6nj02961b



total quantity of 2.3 g, and only iron is more abundant than zinc. As an indispensable cation, zinc is a major regulatory ion in the metabolism of cells and presents many beneficial effects to human health. It has become apparent that zinc deficiency in humans is widely prevalent and significant;^{3–5} about 30% of the world's population is zinc deficient.⁶ Especially vulnerable are children under 5 years of age in developing countries and, for example, treatment of children suffering from deadly diarrhea with zinc significantly reduced child mortality in many third world countries.⁶ It has also been found that zinc is present in insulin, coordinated by three nitrogen atoms from histidine moieties and three water molecules.⁷ Contrary to zinc, functional vanadium compounds are scarcely present in living systems. Apart from some enzymes, such as vanadium-nitrogenase in nitrogen fixing bacteria and haloperoxidases, mainly present in marine algae, vanadium has been found in the form of amavadin, the storage proteins vanabins, and in the specialized blood cells, vanadocytes.⁸

It has been established that zinc complexes possess diverse biological activity, such as anticancer,⁹ antioxidant,¹⁰ antibacterial and antimicrobial activity,¹¹ as well as having effects in the treatment of Alzheimer's disease.¹² The biological functions of vanadium compounds, e.g. as antitumor, antiparasitic (Chagas disease), and osteogenic agents, have also been described.^{8,13} Additionally, zinc and vanadium compounds have attracted increasing attention also due to their antidiabetic effects. A class of very promising complexes consists of neutral Zn and V species with bidentate anionic organic ligands, for example maltolato and ethylmaltolato complexes are more effective in lowering the glucose concentration in blood serum than the parent zinc or vanadium inorganic salts.¹⁴ Additionally, picolinate complexes were also found to be strong inhibitors of fatty acid mobilization and are active in the treatment of STZ-induced diabetic rats.^{15,16} Even better insulin-enhancing effects were observed for 6-methylpicolinate complexes.^{16,17} These initial successes stimulated subsequent tests of several zinc and vanadium complexes formed with picolinate derivatives, such as halopicolinate, 6-ethylpicolinate and 3-hydroxypicolinate ligands,^{3,18} and also species formed with monoesters and amides of 2,5-dipicolinic acid.¹⁹ Recent efforts have been oriented towards the synthesis of new compounds with high activity and low toxicity incorporating a variety of organic ligands including antioxidants²⁰ affording different coordination environments. For example, complexes formed with allixin and *S*-allixin-*N*-methyl as well as with pyrrolidine-*N*-dithiocarbamate ligands have been demonstrated to be particularly active.²¹ Consequently, the field of antidiabetic zinc and vanadium compounds is gaining importance and novel compounds have been prepared and tested during the past few years.²² The recent development of noninvasive methods for analyzing and visualizing the biodistribution of complexes gave a fresh impetus in metallodrug discovery for the treatment of diabetes since it has been established that high antidiabetic activity correlates with a long retention time of complexes in the blood.²³ Although clinical trials on an ethylmaltolato vanadium complex as an antidiabetic drug were abandoned due to side effects affecting the kidneys of the patients it is currently being studied as a therapeutic for the

prevention, stoppage, and reparation of secondary tissue injuries. The main obstacle with the ethylmaltolato vanadium complex was the high dose necessary to achieve the therapeutic effect.^{14f} In order to overcome this challenge Rehder^{8b} has recently pointed out that ligands containing pyrone, pyridinone or pyridine-carboxylate moieties might be the best candidates with sufficiently long physiological half-lives to recombine with metal ions in order to overcome the problem of retention of some vanadium in the body and thus to decrease the potential toxicity.²⁴

In this respect we have revived interest in picolinate ligands and close analogues²⁵ in order to establish their structural characteristics and bioactivities. Herein we report eight vanadium and zinc compounds formed with 5-cyanopicolinic acid (HpicCN), pyrazine-2-carboxylic acid (Hprz) and 3-aminopyrazine-2-carboxylic acid (HprzNH₂). Structural characterization of all eight complexes was carried out by X-ray diffraction analysis. The insulin-mimetic activity of selected complexes was studied by *in vitro* inhibition of FFA release from isolated rat adipocytes treated with epinephrine.

Experimental section

Materials and physical measurements

Reagents and solvents were obtained from commercial suppliers as reagent grade and were used without further purification. Oxidovanadium(IV), VO(IV), compounds [VO(picCN)₂(H₂O)]·2H₂O (1·2H₂O) and [VO(prz)₂(H₂O)] (2) were synthesized from VOSO₄·5H₂O and HpicCN or Hprz as described previously.²⁵ Compound 4 was prepared by a modified synthetic procedure.²⁶

Infrared (IR) spectra (4000–600 cm^{−1}) of the samples were recorded using a Perkin-Elmer Spectrum 100, equipped with a Specac Golden Gate Diamond ATR as a solid sample support. Elemental (C, H, N) analyses were obtained using a Perkin-Elmer 2400 Series II CHNS/O Elemental Analyser. ¹H NMR spectra were recorded with Bruker Avance III 500 and Bruker Avance DPX 300 nuclear magnetic resonance spectrometers with DMSO and D₂O as solvents and TMS as an internal reference.

Synthesis

Ammonium bis(5-cyanopicolinate-κ²N,O)dioxidovanadate(V) (3). Ammonium metavanadate (58 mg, 0.50 mmol) was suspended in water (5 mL) and a solution of HpicCN (124 mg, 1.00 mmol) in methanol (3 mL) was added. The resulting solution was stirred under reflux for 15–30 min or until everything had dissolved. The solution obtained after hot filtration was stored at room temperature. After a few days of slow evaporation of the solvents, light yellow crystals had grown from the solution. Yield: 106 mg, 54%.

Elemental analysis (C₁₄H₁₀N₅O₆V) theoretical: C, 42.55; H, 2.55; N, 17.72. Found: C, 42.60; H, 2.52; N, 17.67%.

IR (ATR, cm^{−1}): 3162br, 3078br, 2847w, 2242w, 1680s, 1652s, 1601m, 1477m, 1424w, 1383m, 1348m, 1330s, 1283m, 1245m, 1152m, 1035m, 897s, 868s, 803m, 788w, 686m, 661m.

Ammonium bis(pyrazine-2-carboxylato-κ²N,O)dioxidovanadate(V) (4). Ammonium metavanadate (58 mg, 0.50 mmol) was suspended in water (5 mL) and a solution of Hprz (124 mg, 1.00 mmol) in methanol (3 mL) was added. The resulting solution was stirred



under reflux for 15 min or until everything had dissolved. The solution obtained after hot filtration was stored at room temperature. After a few days of slow evaporation of the solvents, yellowish brown crystals had grown from the solution. Yield: 92 mg, 53%.

Elemental analysis ($C_{10}H_{10}N_5O_6V$) theoretical: C, 34.60; H, 2.90; N, 20.17. Found: C, 34.54; H, 2.69; N, 19.94%.

IR (ATR, cm^{-1}): 3018br, 2878w, 1674m, 1632m, 1592w, 1470w, 1420m, 1350m, 1321s, 1283w, 1161m, 1042s, 896s, 870m, 861s, 786s, 744m, 647w.

Ammonium bis(3-aminopyrazine-2-carboxylato- κ^2N,O)dioxido-vanadate(v) hydrate (5·H₂O). Ammonium metavanadate (58 mg, 0.50 mmol) was suspended in water (5 mL) and a solution of HprzNH₂ (139 mg, 1.00 mmol) in methanol (3 mL) was added. The resulting solution was stirred under reflux for 15 min or until everything had dissolved. The solution obtained after hot filtration was stored at room temperature. After a few days of slow evaporation of the solvents, yellow crystals had grown from the solution. Yield: 123 mg, 62%.

Elemental analysis ($C_{10}H_{14}N_7O_7V$) theoretical: C, 30.39; H, 3.57; N, 24.81. Found: C, 30.40; H, 3.45; N, 24.6%.

IR (ATR, cm^{-1}): 3423m, 3172w, 2987m, 2875m, 1653s, 1597m, 1555w, 1439m, 1359s, 1321m, 1197m, 1145s, 907m, 875s, 840m, 817m, 731w, 696w.

trans-Di(aqua- κO)bis(5-cyanopicolinato- κ^2N,O)zinc(II) (6). Zinc acetate dihydrate (59 mg, 0.25 mmol) was dissolved in water (2 mL) and a hot solution of HpicCN (74 mg, 0.50 mmol) in methanol (5 mL) was added. After addition a white precipitate formed in the mixture. The resulting suspension was stirred at 60 °C for an additional 20 min and was then cooled to room temperature. The white precipitate was filtered, washed with methanol and dried in air. Yield: 98 mg, 99%.

Elemental analysis ($C_{14}H_{10}N_4O_6Zn$) theoretical: C, 42.50; H, 2.55; N, 14.16. Found: C, 42.63; H, 2.42; N, 14.10%.

IR (ATR, cm^{-1}): 3543w, 3066w, 2799w, 2239w, 1624s, 1597m, 1560m, 1390m, 1367s, 1292w, 1253w, 1038m, 1026m, 890w, 861w, 814w, 787m, 687s, 669w.

cis-Di(4-aminopyridine- κN)bis(5-cyanopicolinato- κ^2N,O)zinc(II)-toluene (1/1) (7·C₇H₈). Compound **6** (20 mg, 0.050 mmol) was suspended in toluene (5 mL) and then solid 4-aminopyridine (10 mg, 0.10 mmol) was added. The mixture was stirred for 10 min at 70 °C, and then hot filtered and stored at room temperature. After a few days of slow evaporation of the solvents, colorless crystals had grown from the solution. Elemental analysis confirmed the content of 0.8 equivalents of toluene. Yield: 24 mg, 77%.

Elemental analysis ($C_{29.6}H_{24.4}N_3O_8Zn$) theoretical: C, 57.19; H, 3.96; N, 18.03. Found: C, 57.21; H, 3.83; N, 18.14%.

IR (ATR, cm^{-1}): 3334w, 3172w, 2234w, 1610s, 1560m, 1519m, 1363s, 1286m, 1248w, 1210m, 1032w, 1012s, 853m, 831w, 833m, 781m, 729w, 691m, 660w.

(4-Aminopyridine- κN)bis(5-cyanopicolinato- κ^2N,O)zinc(II) (8). Compound **6** (99 mg, 0.25 mmol) was suspended in methanol (5 mL) and then solid 4-aminopyridine (48 mg, 0.50 mmol) was added. The mixture was stirred and heated under reflux for 30 min, and then hot filtered and stored at room temperature. After a few days of slow evaporation of the solvents, colorless crystals had grown from the solution. Yield: 87 mg, 38%.

Elemental analysis ($C_{19}H_{12}N_6O_4Zn$) theoretical: C, 50.30; H, 2.67; N, 18.52. Found: C, 49.92; H, 2.35; N, 18.23%.

IR (ATR, cm^{-1}): 3353m, 3198m, 2241w, 1655m, 1644m, 1645s, 1567m, 1525m, 1388w, 1354s, 1252m, 1216m, 1060w, 1038m, 1023s, 858m, 832m, 808m, 785m, 690s, 670m.

cis-Bis(5-cyanopicolinato- κ^2N,O)di(pyridine- κN)zinc(II) (1/1) (9). Compound **6** (20 mg, 0.050 mmol) was dissolved in pyridine (5 mL, 62 mmol). The resulting solution was stirred for 10 min at 70 °C, and then hot filtered and stored at room temperature. After a few days of slow evaporation of the solvents, colorless crystals had grown from the solution. The crystals decomposed when exposed to air. Elemental analysis shows that the pyridine moiety leaves the complex and crystal lattice out of solution, while it is replaced by water molecules. Yield: 15 mg, 58%.

Elemental analysis ($C_{14}H_{10}N_4O_6Zn$) theoretical for C, 42.50; H, 2.55; N, 14.16. Found: C, 42.48; H, 2.80; N, 14.71%.

IR (ATR, cm^{-1}): 3231br, 2233m, 1666s, 1602m, 1565m, 1391m, 1369s, 1253m, 1043m, 929w, 875m, 807m, 782m, 684s, 668m.

Bis(5-cyanopicolinato- κ^2N,O)(1,10-phenanthroline- κ^2N,N')-zinc(II)-toluene-methanol (1/1/2) (10·C₇H₈·2MeOH). Compound **6** (20 mg, 0.050 mmol) was suspended in toluene (5 mL) and then solid 1,10-phenanthroline (9 mg, 0.05 mmol) was added. The resulting mixture was stirred for 10 min at 70 °C and then hot filtered. Methanol (5 mL) was added and the solution was stored at room temperature. After a few days of slow evaporation of the solvents, colorless crystals had grown from the solution. The crystals decomposed on air. Elemental analysis showed that an uncoordinated toluene moiety and methanol leave the crystal lattice and one equivalent of water is absorbed when crystals are exposed to air moisture. Yield: 16 mg, 57%.

Elemental analysis ($C_{26}H_{16}N_6O_5Zn$) theoretical: C, 55.98; H, 2.89; N, 15.07. Found: C, 55.80; H, 2.76; N, 14.68%.

IR (ATR, cm^{-1}): 3422m, 3027w, 2240w, 1654s, 1597m, 1564m, 1520w, 1476w, 1430m, 1344s, 1295w, 1249m, 1211w, 1103w, 1038m, 847m, 811w, 780m, 725m, 695m, 664m.

X-ray structure determination

Crystals suitable for X-ray analysis were obtained from compound solutions or filtered reaction mixtures at room temperature by slow evaporation of the solvents over a few days. Single-crystal X-ray diffraction data were collected on either a Nonius Kappa CCD Diffractometer or an Agilent Technologies SuperNova Dual diffractometer with an Atlas detector using monochromated Mo-K α radiation ($\lambda = 0.71073 \text{ \AA}$) or Cu-K α radiation ($\lambda = 1.54184 \text{ \AA}$) at room temperature (3, 4, 7, 9) or at 150 K (5, 6, 8, 10). The data were processed using DENZO or CrysAlis Pro.²⁷ The structures were solved by direct methods using the program SHELXS-97 or SIR97 and refined on F^2 using full-matrix least-squares procedures (SHELXL-97 or SHELXL2014).²⁸ All non-hydrogen atoms were refined anisotropically. The hydrogen atoms on aromatic rings and NH₂ groups were treated as riding atoms in geometrically idealized positions. Hydrogen atoms attached to water oxygen atoms or to the nitrogen atom of ammonium cations were readily located from difference Fourier maps and refined by fixing the bond lengths and isotropic temperature factors as $U_{iso}(H) = 1.5U_{eq}(O,N)$. Two of the four methanol moieties in the



structure of $10\text{-C}_7\text{H}_8\cdot 2\text{MeOH}$ are disordered over two positions with a refined ratio of 0.47:0.53, and the C–O bond lengths were fixed on both disordered methanol solvate molecules. Crystallographic data are listed in Table S1 (ESI†).

In vitro insulin-like activity

The procedure for the *in vitro* test was adopted from the literature.²⁹ Male Wistar rats were sacrificed under anesthesia with ether. The adipose tissues were removed, chopped with scissors and digested with collagenase for 60 min at 37 °C in Krebs Ringer bicarbonate buffer (120 mM NaCl, 1.27 mM CaCl_2 , 1.2 mM MgSO_4 , 4.75 mM KCl, 1.2 mM KH_2PO_4 , 24 mM NaHCO_3 ; pH 7.4) containing 2% BSA (bovine serum albumin). The obtained adipocytes were then separated from the undigested tissues by filtration through a nylon mesh and washed three times. The metal complexes were dissolved in ultrapure water at 500 μM (final conc.), glucose solution (final conc.; 5 mM) was added to the isolated adipocytes, and the resulting suspension was incubated at 37 °C for 30 min. Finally, epinephrine solution (final conc.; 10 μM) was added to the reaction mixtures, and the resulting solutions were incubated at 37 °C for 180 min. The reactions were stopped by soaking in ice water, and the mixtures were centrifuged at 3000 rpm for 10 min. FFA (free fatty acids) levels in the outer solution of the cells were determined using an FFA kit (NEFA C-test WAKO, Wako Pure Chemicals).

All animal experiments in the present study were approved by the Experimental Animal Research of Kyoto Pharmaceutical University (KPU) and were performed according to the Guidelines for Animal Experimentation of KPU.

Results and discussion

Synthesis and IR spectroscopy

Oxidovanadium(IV), $\text{VO}(\text{IV})$, compounds **1** and **2** were synthesized from $\text{VOSO}_4\cdot 5\text{H}_2\text{O}$ and HpicCN or Hprz as described previously.²⁵ Dioxidovanadate(V), $\text{VO}_2(\text{V})$, compounds were prepared from ammonium metavanadate and the corresponding picolinic acid analogues (HpicCN (**3**), Hprz (**4**) or HprzNH_2 (**5**)) in a 1 : 2 ratio in water. IR spectroscopy of **3–5** shows the presence of N–H stretching vibrations of ammonium cations in the range between 3172 and 2847 cm^{-1} in all three compounds and stretching bands of VO_2^+ cations were observed between 907 and 840 cm^{-1} , while the characteristic absorption band for the cyano group in **3** was observed at 2242 cm^{-1} .

Zinc bis(5-cyanopicolinato) complexes were prepared in two steps. In the first step a white powder of zinc bis(5-cyanopicolinato) compound (**6**) was prepared in a water/methanol mixture from zinc acetate dihydrate and HpicCN acid in a ratio of 1 : 2. In the second step **6** was suspended in an organic solvent and diluted after the addition of neutral N-donor ligands 4apy (**7**– C_7H_8 , **8**) or phen (**10**– $\text{C}_7\text{H}_8\cdot 2\text{MeOH}$). Crystals of **9** were obtained by dissolving **6** in pure pyridine. 4Apy was selected to raise the hydrogen-bond formation potential through side amino groups. In addition, 4apy is already an FDA approved drug used in the treatment of multiple sclerosis as a blocker of

potassium channels.³⁰ Another two N-donors, pyridine and 1,10-phenanthroline, were used due to their structural similarity with 4apy with the aim of comparing their binding modes. The outcome of the reactions with 4apy and py is strongly dependent on the solvent. When the synthesis of 4apy complexes was carried out in chloroform or a chloroform/toluene mixture two 4apy molecules were coordinated to zinc. On the contrary, in methanol only one 4apy coordinates to the zinc, even though the reactant ratio was 1 : 2. Reaction with py in methanol yielded crystals of diaqua complex **6** after exposing the solution to the air for a few days, and the pyridine complex **9** could be isolated only from pure pyridine. However, when crystals **9** were dried in air, the complex decomposed and transformed into the compound **6**. IR spectra of **6** and **7–10** show characteristic absorption bands of the cyano group between 2241 and 2233 cm^{-1} and antisymmetric, $\nu_{\text{asym}}(\text{CO}_2)$, and symmetric, $\nu_{\text{sym}}(\text{CO}_2)$, stretching vibrations of the carboxylato group in the range 1666–1610 cm^{-1} and 1369–1344 cm^{-1} , respectively, in all five compounds. The difference Δ , a useful parameter for revealing the possible coordination mode of the carboxylato ligands, has values greater than 200 cm^{-1} for each of these complexes indicating a monodentate coordination of the carboxylato group.³¹ In the spectra of **7** and **8** an additional two bands in the range between 3353 and 3172 cm^{-1} are present corresponding to the N–H stretching vibrations of amino groups.

Structure of the complexes

Crystal structures of $\text{NH}_4[\text{VO}_2(\text{picCN})_2]$ (3**), $\text{NH}_4[\text{VO}_2(\text{prz})_2]$ (**4**) and $\text{NH}_4[\text{VO}_2(\text{przNH}_2)_2]\cdot \text{H}_2\text{O}$ (**5**· H_2O).** Crystals of compounds **3–5** were grown from the reaction mixtures after a few days of slow evaporation of the solvents. Diagrams of the structures of complexes **3–5** are shown in Fig. 1–3, and selected bond distances and angles are summarized in Table 1. The compounds crystallize in the monoclinic space group $P2_1/c$ (**3** and **5**· H_2O) and in the triclinic space group $P\bar{1}$ (**4**). The asymmetric units contain a complex dioxidovanadate(V) anion and an ammonium cation (**3** and **4**), and in addition, in **5**· H_2O one lattice water molecule is also present.

The vanadium atom in **3** is six-coordinated having an octahedral geometry with a *cis* arrangement of the oxido oxygen

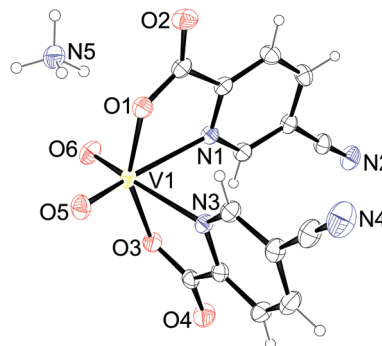


Fig. 1 Asymmetric unit of $\text{NH}_4[\text{VO}_2(\text{picCN})_2]$ (**3**). Thermal ellipsoids are shown with a 30% probability level at 293 K.



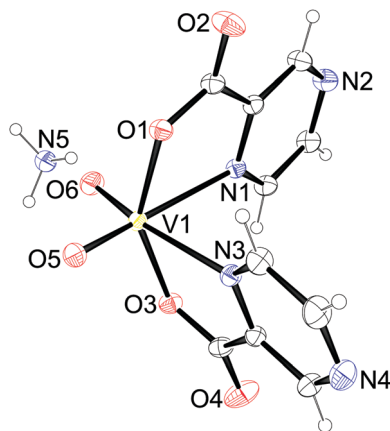


Fig. 2 Asymmetric unit of $\text{NH}_4[\text{VO}_2(\text{prz})_2]$ (**4**). Thermal ellipsoids are shown with a 30% probability level at 293 K.

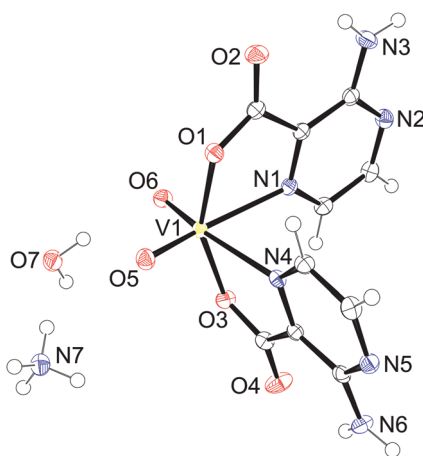


Fig. 3 Asymmetric unit of $\text{NH}_4[\text{VO}_2(\text{przNH}_2)_2] \cdot \text{H}_2\text{O}$ (**5**· H_2O). Thermal ellipsoids are shown with a 50% probability level at 150 K.

atoms of the VO_2^+ group and two bidentate 5-cyanopyridine-2-carboxylate (picCN) ligands coordinated with picolinato nitrogen atoms *trans* to the oxido oxygen atoms [$\text{V}-\text{N} = 2.3428(14)$ and $2.3275(15)$ Å] and carboxylate oxygen atoms in a mutual *trans* orientation [$\text{V}-\text{O} = 1.9824(12)$ and $1.9824(12)$ Å]. The octahedral geometry displays a distortion with $\text{O5}-\text{V1}-\text{N1}$ and $\text{O6}-\text{V1}-\text{N3}$ angles of $160.75(6)^\circ$ and $162.91(6)^\circ$, respectively, and a $\text{O1}-\text{V1}-\text{O3}$ angle of $149.58(4)^\circ$.

Ammonium cations and complex anions in the crystal structure of **3** are connected into infinite double layers parallel to the *ac* plane through several hydrogen bonds (Fig. S1, ESI†). Each ammonium cation is as a hydrogen bond donor involved in six different hydrogen bonds, connected to four adjacent complex anions. Hydrogen bond acceptors in $[\text{VO}_2(\text{picCN})_2]^-$ are carbonyl, carboxyl and oxido oxygen atoms. Additional weak $\text{C}-\text{H} \cdots \text{O}$ interactions are formed between the adjacent anions, but only within the same hydrogen-bonded layers. Hydrogen bonds and other weak intermolecular interactions in **3** are listed in Table S2 (ESI†).

The vanadium atom in **4** and **5** is also six-coordinated having an octahedral geometry similar to **3** with a *cis* arrangement of the oxido oxygen atoms of the VO_2^+ group and two bidentate

Table 1 Selected bond distances (Å) and angles ($^\circ$) for $\text{NH}_4[\text{VO}_2(\text{picCN})_2]$ (**3**), $\text{NH}_4[\text{VO}_2(\text{prz})_2]$ (**4**) and $\text{NH}_4[\text{VO}_2(\text{przNH}_2)_2] \cdot \text{H}_2\text{O}$ (**5**· H_2O)

Distance (Å)	3	4	5·H ₂ O
V1–O1	1.9824(13)	2.0100(10)	1.9831(12)
V1–O3	1.9824(12)	1.9619(10)	1.9624(12)
V1–O5	1.6236(13)	1.6327(10)	1.6255(12)
V1–O6	1.6216(14)	1.6243(10)	1.6528(13)
V1–N1	2.3434(14)	2.3021(12)	2.2914(14)
V1–N3/4 ^a	2.3275(15)	2.3075(12)	2.2969(15)
Angle ($^\circ$)	3	4	5·H ₂ O
O1–V1–O3	153.78(5)	149.59(4)	154.47(5)
O1–V1–O5	94.27(6)	94.43(5)	93.92(6)
O1–V1–O6	102.25(7)	99.82(5)	100.47(6)
O1–V1–N1	73.13(5)	73.90(4)	74.21(5)
O1–V1–N3/4 ^a	87.05(5)	82.22(4)	86.46(5)
O3–V1–O5	103.10(7)	104.35(5)	102.93(6)
O3–V1–O6	92.46(6)	98.18(5)	93.36(6)
O3–V1–N1	84.68(5)	82.89(4)	84.40(5)
O3–V1–N3/4 ^a	73.89(5)	74.81(4)	74.85(5)
O5–V1–O6	104.18(8)	104.95(6)	105.65(7)
O5–V1–N1	160.75(6)	164.90(5)	161.66(6)
O5–V1–N3/4 ^a	89.19(6)	88.20(5)	89.34(6)
O6–V1–N1	92.89(7)	86.78(5)	90.47(6)
O6–V1–N3/4 ^a	162.91(6)	166.41(5)	162.85(6)
N1–V1–N3/4 ^a	75.93(5)	80.85(4)	76.23(5)

^a N3 for **3** and **4** and N4 for **5**.

pyrazine-2-carboxylate (prz, przNH₂) ligands coordinated with pyrazine nitrogen atoms *trans* to the oxido oxygen atoms [$\text{V}-\text{N} = 2.3021(12)$, $2.3075(12)$ Å for **4** and $2.2914(14)$, $2.2969(15)$ Å for **5**] and carboxylate oxygen atoms in a mutual *trans* orientation [$\text{V}-\text{O} = 2.0100(10)$, $1.9619(10)$ Å for **4** and $1.9831(12)$, $1.9624(12)$ Å for **5**]. The octahedral geometry displays a distortion with angles $\text{O5}-\text{V1}-\text{N1}$ of $164.90(5)^\circ$, $\text{O6}-\text{V1}-\text{N3}$ of $166.41(5)^\circ$ and $\text{O1}-\text{V1}-\text{O3}$ $149.58(4)^\circ$ in **4** and angles $\text{O5}-\text{V1}-\text{N1}$ of $161.66(6)^\circ$, $\text{O6}-\text{V1}-\text{N4}$ of $162.85(6)^\circ$ and $\text{O1}-\text{V1}-\text{O3}$ of $154.47(5)^\circ$ in **5**.

Hydrogen bonds between the ammonium cations and heteroatoms of the complex anions in **4** enable the formation of hydrogen-bonded double layers parallel to the *ab* plane (Fig. S2, ESI†). The hydrogen bond acceptors in **4** are carbonyl and oxido oxygen atoms and the pyrazine nitrogen N2. The crystal structure is further stabilized by weak $\text{C}-\text{H} \cdots \text{O}$ bonds and $\pi-\pi$ interactions with a $\text{Cg}4 \cdots \text{Cg}4$ centroid-to-centroid distance of $3.6544(11)$ Å connecting adjacent layers along the *c* axis (Tables S3 and S4, ESI†). Extended hydrogen bonding connects ammonium cations, the complex anion and lattice water molecules in the crystal structure of **5**· H_2O into a three dimensional network (Fig. S3, ESI†) enhanced also by the presence of an amino group on the przNH₂ ligand as a hydrogen bond donor. Crystal water acts as a hydrogen bond donor and acceptor, while carbonyl, carboxyl and oxido atoms and pyrazine N5 and amino N6 nitrogen atoms are involved as hydrogen bond acceptors. The crystal lattice is additionally stabilized by weak $\text{C4}-\text{H4} \cdots \text{O4}$ bonds and $\pi-\pi$ interactions with $\text{Cg}3 \cdots \text{Cg}3$ and $\text{Cg}4 \cdots \text{Cg}4$ centroid-to-centroid distances of $3.5669(10)$ and $3.9506(9)$ Å, respectively, connecting the complex anions along the *b* axis (Fig. S4, Tables S5 and S6, ESI†).

Crystal structures of $[\text{Zn}(\text{picCN})_2(\text{H}_2\text{O})_2]$ (6**), $[\text{Zn}(\text{picCN})_2(4\text{apy})_2] \cdot \text{C}_7\text{H}_8$ (**7**· C_7H_8), $[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**), $[\text{Zn}(\text{picCN})_2(\text{py})_2]$ (**9**), and $[\text{Zn}(\text{picCN})_2(\text{phen})] \cdot \text{C}_7\text{H}_8 \cdot 2\text{MeOH}$ (**10**· $\text{C}_7\text{H}_8 \cdot 2\text{MeOH}$).** Crystals of



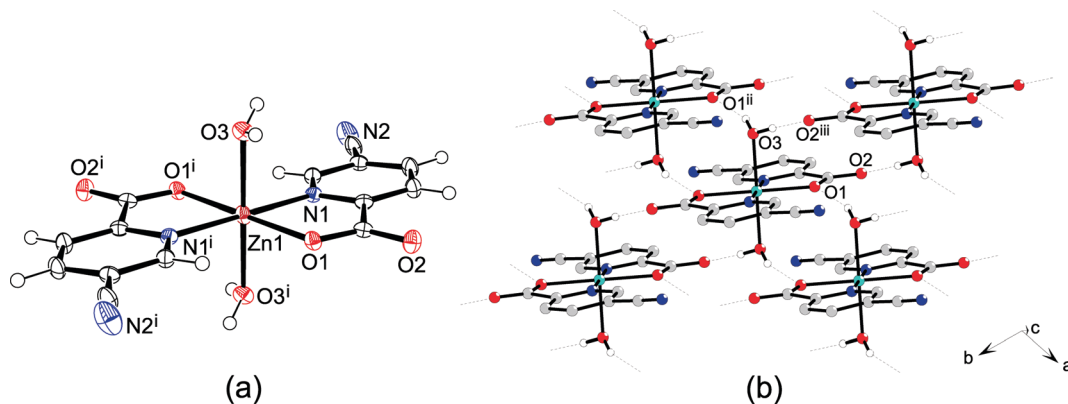


Fig. 4 (a) Structure of the compound $[\text{Zn}(\text{picCN})_2(\text{H}_2\text{O})_2]$ (**6**). Thermal ellipsoids are shown with a 50% probability level at 150 K. (b) The layer of hydrogen-bonded molecules in **6**. Dashed lines indicate $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds. Symmetry codes: (i) $-x+1, -y+1, -z+1$; (ii) $x-1, y, z$; (iii) $-x+1, -y, -z+1$.

compounds **6–10** were grown from the reaction mixtures after a few days of slow evaporation of the solvents. Diagrams of the structures of complexes **6–10** are shown in Fig. 4–6, and selected bond distances and angles are summarized in Tables 2–4.

The complex **6** crystallizes in the triclinic space group $P\bar{1}$. The asymmetric unit contains one half of the complex with the zinc atom sitting on the inversion center. The zinc atom is octahedrally coordinated with two bidentate picCN ligands in the equatorial position in a *trans* arrangement and with two water molecules in the axial positions. The distance between the zinc and the carboxyl oxygen ($\text{Zn1}-\text{O1} = 2.0895(15)$ Å) is shorter than the bond between the zinc and water ($\text{Zn1}-\text{O3} = 2.1118(18)$ Å). The $\text{Zn1}-\text{N1}$ distance is $2.1306(19)$ Å. The angles between the *trans* oriented atoms are 180° due to the center of symmetry. The chelate angle of the picCN ligand $\text{O1}-\text{Zn1}-\text{N1}$ is $78.65(7)^\circ$, the angle between the water and the carboxyl oxygen atom $\text{O1}-\text{Zn1}-\text{O3}$ is $90.23(7)^\circ$, and the angle between the water and the picolinato nitrogen atom $\text{O3}-\text{Zn1}-\text{N1}$ is $92.32(7)^\circ$.

The crystal structure of **6** is stabilized by hydrogen bonding between a coordinated water molecule acting as a hydrogen bond donor and a carbonyl and carboxyl oxygen atoms as hydrogen bond acceptors forming hydrogen-bonded layers parallel to the *ab* plane (Fig. 4 and Table S7, ESI†). Furthermore, the coordinated water molecule is also an acceptor of a weak $\text{C}-\text{H}\cdots\text{O}$ interaction formed with the picolinic C–H moiety.

By reacting contacting **6** and 4apy two different crystals were prepared depending on the solvent used. Both syntheses were performed in a 1:2 molar ratio, however, in toluene crystals with two coordinated 4apy molecules were obtained (**7**), while synthesis in methanol yielded **8** with only one 4apy molecule coordinated to the zinc atom.

Compound **7** crystallizes as a toluene solvate in the triclinic space group $P\bar{1}$ with the formula $[\text{Zn}(\text{picCN})_2(4\text{apy})_2]\cdot\text{C}_7\text{H}_8$. The asymmetric unit contains one molecule of the complex and one molecule of toluene. The complex has a distorted octahedral

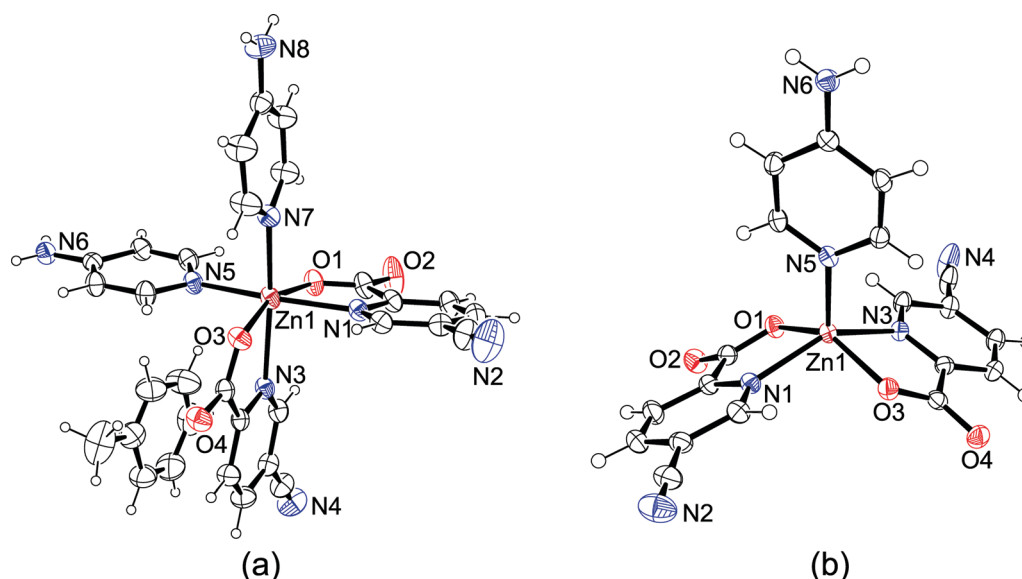


Fig. 5 Asymmetric unit of (a) $[\text{Zn}(\text{picCN})_2(4\text{apy})_2]\cdot\text{C}_7\text{H}_8$ (**7**· C_7H_8). Thermal ellipsoids are shown with a 30% probability level at 293 K. (b) $[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**). Thermal ellipsoids are shown with a 50% probability level at 150 K.



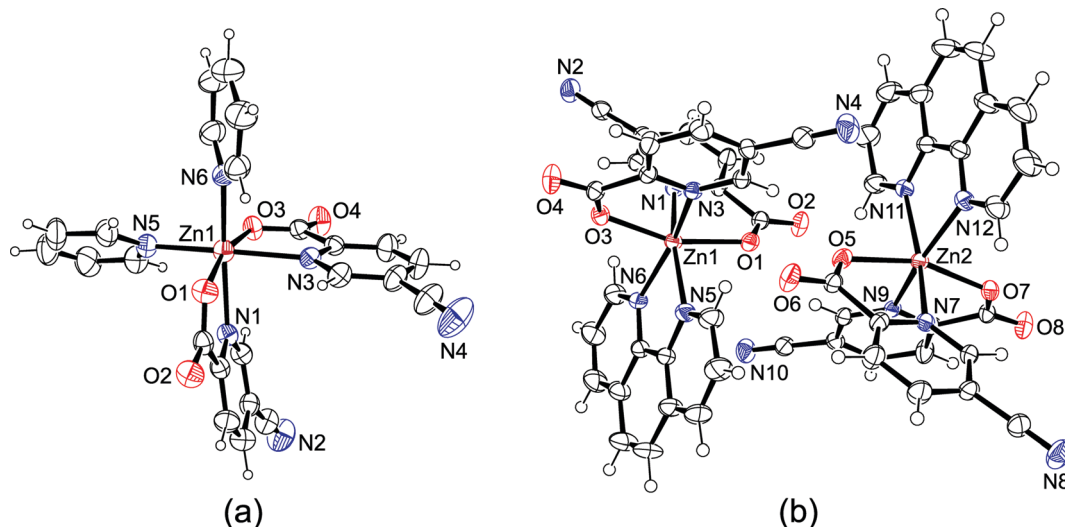


Fig. 6 (a) Asymmetric unit of $[\text{Zn}(\text{picCN})_2(\text{py})_2]$ (**9**). Thermal ellipsoids are shown with a 30% probability level at 293 K. (b) Crystallographically independent molecules of $[\text{Zn}(\text{picCN})_2(\text{phen})]$ (**10**). Solvate molecules have been removed for clarity. Thermal ellipsoids are shown with a 50% probability level at 150 K.

Table 2 Selected bond distances (Å) and angles (°) for $[\text{Zn}(\text{picCN})_2(\text{H}_2\text{O})_2]$ (**6**)

Distance (Å)	6		
Zn1–O1	2.0895(15)	Zn1–N1	2.1306(19)
Zn1–O3	2.1118(18)		
Angle (°)	6		
O1–Zn1–O1 ⁱ	180.00(8)	O3–Zn1–O3 ⁱ	180.0
O1–Zn1–O3	89.77(7)	O3–Zn1–N1	92.32(7)
O1–Zn1–O3 ⁱ	90.23(7)	O3–Zn1–N1 ⁱ	87.68(7)
O1–Zn1–N1	78.65(7)	N1–Zn1–N1 ⁱ	180.00(1)
O1–Zn1–N1 ⁱ	101.35(7)		

Symmetry code: (i) $-x + 1, -y + 1, -z + 1$.

geometry, with a *cis*-octahedral arrangement of 4apy ligands on the zinc central atom. Two 4apy molecules are bound to the zinc through the pyridine nitrogen atoms with Zn–N distances of 2.1256(18) and 2.1183(18) Å in *cis* positions with a N5–Zn1–N7 angle of 89.65(7)°, while the two bidentate picCN ligands are coordinated with the picolinato nitrogen atoms in a mutual *cis* position with Zn–N distances of 2.2246(17) and 2.2779(18) Å, being longer than those in compound **6**. Carboxyl oxygen atoms are bound to the zinc atom in mutual *trans* positions (O1–Zn1–O3 = 161.02(7)° with Zn–O distances of 2.0845(14) and 2.0992(14) Å).

The supramolecular structure of 7- C_7H_8 is stabilized by N–H...O hydrogen bonding between the amino groups of the 4apy ligands and the carbonyl oxygen atoms of the picCN ligands forming layers parallel to the *bc* plane (Fig. S5 and Table S8, ESI†). The crystal lattice is additionally stabilized by weak C–H... π and π ... π interactions with a Cg6...Cg7 centroid-to-centroid distance of 3.8672(1) Å between the toluene solvate molecule and the complex along the *b* axis (Table S9, ESI†).

$[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**) crystallizes in the monoclinic space group $P2_1/c$. The asymmetric unit contains one molecule of the complex. The central zinc atom is pentacoordinated with an

Table 3 Selected bond distances (Å) and angles (°) for $[\text{Zn}(\text{picCN})_2(4\text{apy})_2] \cdot \text{C}_7\text{H}_8$ (**7**· C_7H_8), $[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**) and $[\text{Zn}(\text{picCN})_2(\text{py})_2]$ (**9**)

Distance (Å)	7 · C_7H_8	8	9
Zn1–O1	2.0845(14)	2.0486(12)	2.056(2)
Zn1–O3	2.0992(14)	2.0518(12)	2.0483(18)
Zn1–N1	2.2246(17)	2.1309(15)	2.240(2)
Zn1–N3	2.2779(18)	2.0953(14)	2.222(2)
Zn1–N5	2.1256(18)	2.0088(14)	2.138(2)
Zn1–N7/6 ^a	2.1183(18)	—	2.133(2)
Angle (°)	7 · C_7H_8	8	9
O1–Zn1–O3	161.02(7)	144.58(5)	166.13(9)
O1–Zn1–N1	76.60(6)	78.59(5)	76.94(8)
O1–Zn1–N3	90.07(6)	88.50(5)	93.41(8)
O1–Zn1–N5	94.93(6)	113.92(6)	96.90(9)
O1–Zn1–N7/6 ^a	100.25(7)	—	92.45(9)
O3–Zn1–N1	91.40(6)	92.95(5)	91.87(8)
O3–Zn1–N3	75.05(6)	79.43(5)	77.30(8)
O3–Zn1–N5	96.94(7)	101.49(5)	91.25(8)
O3–Zn1–N7/6 ^a	94.59(6)	—	97.86(9)
N1–Zn1–N3	89.83(6)	145.82(6)	85.18(8)
N1–Zn1–N5	171.51(6)	104.35(6)	90.32(9)
N1–Zn1–N7/6 ^a	91.28(7)	—	168.29(9)
N3–Zn1–N5	90.78(7)	109.81(6)	167.53(9)
N3–Zn1–N7/6 ^a	169.60(6)	—	90.61(9)
N5–Zn1–N7/6 ^a	89.65(7)	—	95.92(9)

^a N7 for 7- C_7H_8 and N6 for **9**.

almost ideal square-pyramidal coordination sphere. Distortion of the square pyramid can be best described by the structural parameter $\tau = (\alpha - \beta)/60^\circ$, where α and β are the largest angles in the coordination sphere ($\tau = 0$ for a square-pyramid and $\tau = 1$ for a trigonal bipyramid),³² which in this case has a value of 0.02. The axial coordination site is occupied by 4apy with the shortest Zn–N distance among the reported compounds (2.0088(14) Å) and the bidentate picCN ligands are coordinated with a mutual *trans* arrangement in a basal plane with Zn–N distances of 2.1309(15) and 2.0953(14) Å and Zn–O distances of



Table 4 Selected bond distances (Å) and angles (°) for [Zn(picCN)₂(phen)]·C₇H₈·2MeOH (**10**·C₇H₈·2MeOH)

Distance (Å)			
Zn1–O1	2.073(2)	Zn2–O5	2.064(2)
Zn1–O3	2.061(2)	Zn2–O7	2.069(2)
Zn1–N1	2.186(2)	Zn2–N7	2.196(2)
Zn1–N3	2.193(2)	Zn2–N9	2.185(2)
Zn1–N5	2.162(2)	Zn2–N11	2.169(2)
Zn1–N6	2.159(3)	Zn2–N12	2.153(2)
Angle (°)			
O1–Zn1–O3	161.01(10)	O5–Zn2–O7	160.96(10)
O1–Zn1–N1	77.15(9)	O5–Zn2–N7	77.17(9)
O1–Zn1–N3	90.46(9)	O5–Zn2–N9	91.95(9)
O1–Zn1–N5	90.98(9)	O5–Zn2–N11	90.81(9)
O1–Zn1–N6	104.60(9)	O5–Zn2–N12	103.32(9)
O3–Zn1–N1	91.34(9)	O7–Zn2–N7	90.02(9)
O3–Zn1–N3	77.35(9)	O7–Zn2–N9	77.12(9)
O3–Zn1–N5	103.57(9)	O7–Zn2–N11	104.86(9)
O3–Zn1–N6	90.74(9)	O7–Zn2–N12	90.95(9)
N1–Zn1–N3	103.18(10)	N7–Zn2–N9	103.08(10)
N1–Zn1–N5	161.48(10)	N7–Zn2–N11	161.20(10)
N1–Zn1–N6	92.26(9)	N7–Zn2–N12	91.79(9)
N3–Zn1–N5	90.91(9)	N9–Zn2–N11	91.58(9)
N3–Zn1–N6	160.57(10)	N9–Zn2–N12	160.78(10)
N5–Zn1–N6	76.85(10)	N11–Zn2–N12	76.79(11)

2.0486(12) and 2.0518(12) Å. All the observed Zn–N/O distances in **8** are shorter than the corresponding distances in **7**. The angle between the picolinate nitrogen atoms (N1–Zn1–N3) is 145.82(6)°, and the angle between the carboxylate oxygen atoms (O1–Zn1–O3) is 144.58(5)°. Angles between the 4apy pyridine nitrogen and the atoms of the picCN ligands are in the range of 101.49(5)–113.92(6)°.

The supramolecular structure of **8** is stabilized with N–H···O hydrogen bonding between adjacent molecules through the amino group of the 4apy ligand and the carbonyl oxygen atoms of picCN ligands forming layers parallel to the *bc* plane (Fig. S6, ESI†). Additionally, weak C6–H6···N4 and C13–H13···O3 interactions are formed enhancing the layer structure, while adjacent layers are connected through weak C4–H4···O4 interactions (Table S8, ESI†).

[Zn(picCN)₂(py)₂] (**9**) crystallizes in the orthorhombic space group *Pna*2₁ and the asymmetric unit contains one molecule of the complex. The complex has a distorted octahedral geometry with the same *cis*-octahedral arrangement of the ligands as in **7**. Two py molecules are bound to the zinc through the pyridine nitrogen atoms with Zn–N distances of 2.138(2) and 2.133(2) Å in a *cis* position with a N5–Zn1–N6 angle of 95.92(9)°, while the two bidentate picCN ligands are coordinated with picolinate nitrogen atoms in a mutual *cis* position with Zn–N distances of 2.240(2) and 2.222(2) Å. Carboxyl oxygen atoms are bound to the zinc atom in mutual *trans* positions (O1–Zn1–O3 = 166.13(9)° with Zn–O distances of 2.056(2) and 2.0483(18) Å).

In the crystal structure of **9** no classical hydrogen bonding is present. Stabilization of the crystal lattice in all three dimensions is due to weak C–H···O interactions between the aromatic C–H and the carbonyl oxygen and the cyano nitrogen atoms (Fig. S7 and Table S8, ESI†).

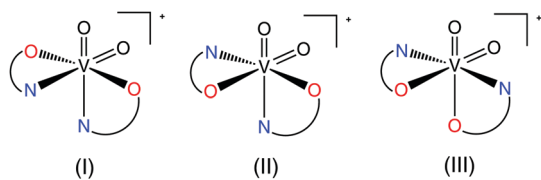
Compound **10** crystallizes in the monoclinic space group *P*2₁/*c* as a toluene methanol solvate (1:1:2) with the formula [Zn(picCN)₂(phen)]·C₇H₈·2MeOH. The asymmetric unit contains two crystallographically independent complex molecules presented in Fig. 6, two molecules of toluene and four molecules of methanol. The two complex molecules in the asymmetric unit are *A*- (Zn1) and *A*-isomers (Zn2) of **10**. The distances and angles of both isomers are almost the same, therefore only one is discussed. The zinc central atom is hexacoordinated with the same octahedral arrangement as in **7** and **9**. Phen is coordinated to the zinc with Zn–N distances of 2.162(2) and 2.159(3) Å and the picCN ligands with somewhat longer Zn–N distances of 2.186(2) and 2.193(2) Å and Zn–O distances of 2.073(2) and 2.061(2) Å. The two picolinate nitrogen atoms are *cis* to each other (N1–Zn1–N3 = 103.18(10)°), and the two carboxyl oxygen atoms are in a *trans* arrangement (O1–Zn1–O3 = 161.01(10)°).

Methanol molecules are hydrogen bonded to the carbonyl oxygen atoms of the complex. The supramolecular structure is further stabilized in all three dimensions by weak C–H···O and C–H···N interactions formed between the aromatic picCN or phen protons and the carbonyl or carboxyl oxygen or the cyano nitrogen atoms. Furthermore, toluene is embedded in the structure by C–H···π interactions between the phen ligands and toluene as well as by π···π stacking interactions with centroid-to-centroid distances of 3.8752(17) and 3.9270(19) Å between the picCN rings and toluene along the *b* axis (Fig. S8, Tables S9 and S10, ESI†).

Structural comparison

Structure has a significant influence on the stability and biotransformation of the complexes as has been demonstrated for example for VO²⁺ species and their interactions with serum proteins by the use of EPR spectroscopy.³³ We are focusing on the diverse structural characteristics and supramolecular topologies of potential antidiabetic vanadium compounds.³⁴ Recently, we have been studying VO²⁺ compounds with picolinic acid and pyrazine derivatives concerning their structures and interactions with proteins. Furthermore, we were able to establish the first crystallographic evidence for the formation of an isomer of a vanadyl bis(picolinate) complex with the two pyridine nitrogen atoms in a mutual *trans* position; the known structures possess a common arrangement with the two equatorial pyridine nitrogen atoms in a mutual *cis* position.²⁵ We have extended the research also on V(v) and zinc compounds in order to verify the flexibility and diversity of spatial arrangements available in these species as well. A search in the Cambridge Structural Database (CSD) for mononuclear bisoxido vanadate(v) compounds with two picolinate or pyrazine-2-carboxylate ligands or derivatives revealed only three crystal structures. A compound synthesized by Gätjens *et al.*^{19a} with 5-(methoxycarbonyl)picolinic acid and a compound prepared by Süss-Fink *et al.*³⁵ with pyrazine-2-carboxylic acid showing nitrogen atoms coordinated to the V metal center *trans* to the oxido oxygen atoms as observed also in **3–5** (Scheme 1(I)). Remarkably, different spatial orientation was found in NH₄[VO₂-(pic)₂]·H₂O synthesized by Melchior *et al.*³⁶ with one picolinate nitrogen atom coordinated to the V metal center *trans* to the

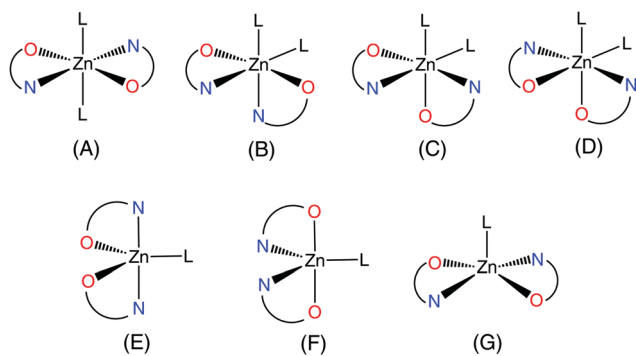




Scheme 1 Possible bis-chelated structures for VO_2^+ species formed by a bidentate ligand.

oxido oxygen and the other picolinate nitrogen atom coordinated in a *trans* position to the carboxyl oxygen atom (Scheme 1(II)). Structures with the spatial arrangement as in isomer (III) (Scheme 1) have not been reported yet.

A search in the CSD for mononuclear neutral zinc compounds with two picolinate ligands or derivatives revealed 27 hexa- and penta-coordinated crystal structures with seven different spatial arrangements. The most frequent, with 16 compounds, is a *trans*-octahedral arrangement of ligands (Scheme 2(A)) as found, for example, in *trans*-[Zn(pic)₂(H₂O)₂] (Hpic = picolinic acid),¹⁶ [Zn(5Mepic)₂(H₂O)₂] (H5Mepic = 5-methylpicolinic acid),³⁷ [Zn(hypic)₂(H₂O)₂] (Hhypic = 3-hydroxypicolinic acid)³⁸ and also in **6**. It is of interest to compare zinc compounds with the quinoline-2-carboxylato (q2c) ligand in the presence of 1-methylimidazole (Meim) or imidazole (im) with [Zn(q2c)₂(Meim)₂] possessing *trans* geometry³⁹ while [Zn(q2c)₂(im)₂] possesses *cis* geometry with the carboxylato groups in a mutual *trans* position⁴⁰ (Scheme 2(B)) as observed also in **7**, **9** and **10**. Another two possible *cis* spatial arrangements were found in *cis*-[Zn(pic)₂(H₂O)₂]⁴¹ (Scheme 2(C)) and in [Zn(bipyc)₂(H₂O)₂] (Hbipyc = 4,4'-bipyridyl-2-carboxylic acid) (Scheme 2(D)) with the nitrogen atoms in a mutual *trans* position.⁴² Additionally, two structures have a coordinated bidentate phen ligand enforcing a *cis* orientation⁴³ (Scheme 2(B)) as found also in **7**, **9** and **10**. Only 6 structures have been reported with a pentacoordinated arrangement. Out of these, five structures possess a trigonal-bipyramidal geometry: [Zn(6Rpic)₂(H₂O)] (H6Rpic = 6-methylpicolinic acid and 6-ethylpicolinic acid),^{17,44} [Zn(q2c)₂(H₂O)]⁴⁵ and [Zn(bp18c6)₂(H₂O)] (H₂bp18c6 = *N,N'*-bis[(6-carboxy-2-pyridyl)methyl]-4,13-diaza-18-crown-6)⁴⁶ with the nitrogen donor atoms in an axial position (Scheme 2(E)), and [Zn(Hpycc)₂(H₂O)] (H₂pycc = pyridine-5-carboxylato-2-carboxylic



Scheme 2 Observed bis-chelated structures for Zn^{2+} species formed with a bidentate picolinate ligand and analogues in the presence of a neutral ligand L.

acid)⁴⁷ with the carboxylic oxygen atoms in an axial position (Scheme 2(F)) while only [Zn(bupyc)₂(H₂O)] (Hbupyc = 5-butylpyridine-2-carboxylic acid)⁴⁸ possesses a square-pyramidal geometry (Scheme 2(G)) as observed in compound **8**. These findings show the rich structural diversity that has to be taken into consideration in drug research and design.

In vitro insulino-mimetic activity

The insulin-like activities of the vanadium(IV),²⁵ vanadium(V) and two zinc(II) compounds prepared were evaluated by an *in vitro* experiment. The inhibitory effects on free fatty acid (FFA) release from isolated rat adipocytes treated with epinephrine were measured. Fig. 7 and 8 show the effects of the vanadium and zinc complexes on FFA release compared with those of ionic VS (VOSO_4) and ZS (ZnSO_4), respectively. All metal complexes showed insulino-mimetic activity. The prz complex **2** was found to have a higher insulino-mimetic activity than the positive control (Fig. 7). The other vanadium compounds have activities similar to VOSO_4 . A comparison of **4** and **5** shows that the introduction of an amino group at position 3 of the pyrazine-2-carboxylato ligand has no significant influence on the activity of the complex. $\text{VO}(\text{IV})$ compounds exhibit higher activity than $\text{VO}_2(\text{V})$ species and this is in agreement with previous findings²⁴ and can be attributed to the different redox behavior of vanadium species,⁴⁹ the different mechanism of cellular uptake⁵⁰ or the different speciation in the cytosol.⁵¹ Additionally, it has been also proposed that $[\text{VO}(\text{IV})(\text{OH})_3]^-$ could be the active species.⁸

The zinc complexes also exhibited some insulino-mimetic activity, but the insulino-mimetic activity of these complexes was not more potent than that of ZnSO_4 (Fig. 8). Interestingly, the introduction of an N-donor 4apy to the zinc picCN complex

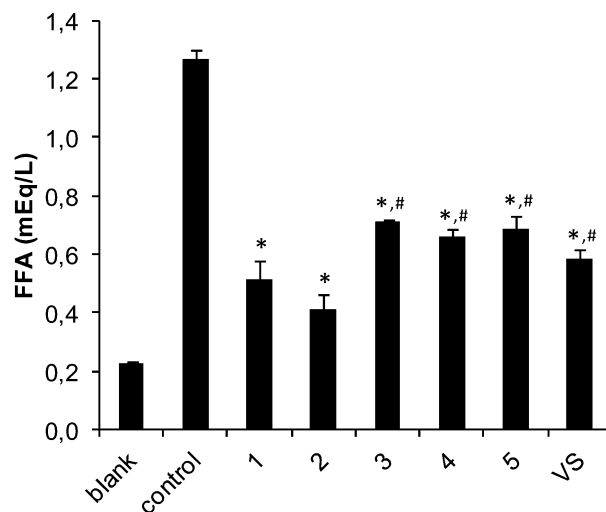


Fig. 7 Inhibitory effects of vanadium complexes (**1–5**) and VOSO_4 (VS) on free fatty acid release from isolated rat adipocytes treated with epinephrine. The blank represents cells only and the control represents cells plus 10 μM epinephrine solution. Data are expressed as the means \pm SDs for three experiments at a 0.5 mM concentration in water. Significant differences were analyzed by one-way analysis of the variance and Tukey test. *: $p < 0.01$ vs. control, and #: $p < 0.01$ vs. compound **2**.

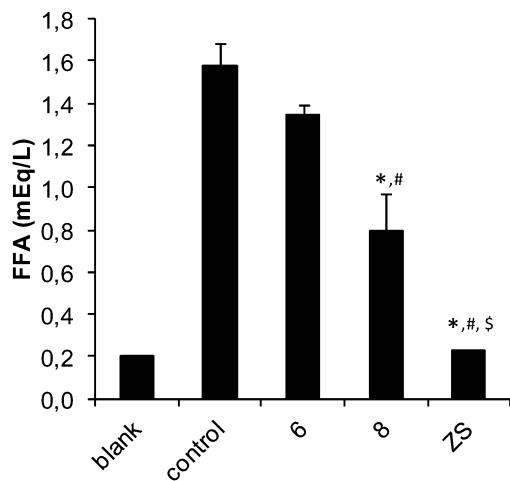


Fig. 8 Inhibitory effects of zinc complexes (**6** and **8**) and ZnSO_4 (ZS) on free fatty acid release from isolated rat adipocytes treated with epinephrine. The blank represents cells only and the control represents cells plus 10 μM epinephrine solution. Data are expressed as the means \pm SDs for three experiments at a 0.5 mM concentration in water. Significant differences were analyzed by one-way analysis of the variance and Tukey test. *: $p < 0.01$ vs. control, #: $p < 0.01$ vs. compound **6**, and \$: $p < 0.01$ vs. compound **8**.

8 significantly increased inhibition of the FFA release compared to **6**.

Conclusion

Three vanadium(v) compounds with 5-cyanopicolinato acid (HpicCN), pyrazine-2-carboxylic acid (Hprz) and 3-aminopyrazine-2-carboxylic acid (HprzNH_2) and five zinc compounds with HpicCN in the presence of water, 4-aminopyridine (4apy), pyridine (py) and 1,10-phenanthroline (phen) have been prepared and characterized. The crystal structures of $\text{NH}_4[\text{VO}_2(\text{picCN})_2]$ (**3**), $\text{NH}_4[\text{VO}_2(\text{prz})_2]$ (**4**), and $\text{NH}_4[\text{VO}_2(\text{przNH}_2)_2] \cdot \text{H}_2\text{O}$ (**5**· H_2O) possess similar spatial arrangements with the carboxylate oxygen atoms in a mutual *trans* orientation although the carboxylate oxygen atoms in the prototypical complex $\text{NH}_4[\text{VO}_2(\text{pic})_2]$ are positioned in a mutual *cis* orientation. In the zinc bis(5-cyanopicolinato) complexes $[\text{Zn}(\text{picCN})_2(\text{H}_2\text{O})_2]$ (**6**), $[\text{Zn}(\text{picCN})_2(4\text{apy})_2] \cdot \text{C}_7\text{H}_8$ (**7**· C_7H_8), $[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**), $[\text{Zn}(\text{picCN})_2(\text{py})_2]$ (**9**), and $[\text{Zn}(\text{picCN})_2(\text{phen})] \cdot \text{C}_7\text{H}_8 \cdot 2\text{MeOH}$ (**10**· $\text{C}_7\text{H}_8 \cdot 2\text{MeOH}$) three different arrangements were found with *trans* (**6**) and *cis* (**7**, **9** and **10**) octahedral and square-pyramidal (**8**) geometries. The “4apy” ligand forms two different compounds, $[\text{Zn}(\text{picCN})_2(4\text{apy})_2] \cdot \text{C}_7\text{H}_8$ (**7**· C_7H_8) with a *cis*-octahedral and $[\text{Zn}(\text{picCN})_2(4\text{apy})]$ (**8**) with a square-pyramidal geometry depending on the polarity of the solvent used (chloroform/toluene vs. methanol). Attempts were made to prepare analogues of **7** and **8** with a py ligand, however only *cis*-octahedral **9** was isolated with pyridine as the solvent and a diaqua complex **6** from a methanol solution. Interestingly, **8** transformed to **6** on standing in air. The diverse spatial arrangements observed indicate the structural flexibility, a property useful in the process of molecular recognition and formation of compound–protein complexes. The insulino-mimetic activity

of $\text{VO}(\text{IV})$ (**1**, **2**) $\text{VO}_2(\text{V})$ (**3**–**5**) and $\text{Zn}(\text{II})$ complexes (**6** and **8**) was studied by *in vitro* inhibition of free fatty acid (FFA) release from isolated rat adipocytes treated with epinephrine. All selected metal complexes showed insulino-mimetic activity. The vanadium(IV) complex **prz 2** was found to have a higher insulino-mimetic activity than the positive control. The other vanadium compounds have activities similar to VOSO_4 . The Zn complexes also exhibited some insulino-mimetic activity. The introduction of an N-donor 4apy to the zinc picCN complex **8** significantly increased the inhibition of FFA release compared to **6**.

Acknowledgements

The authors thank the Slovenian Research Agency for financial support (P1-0230-0175) as well as the EN-FIST Centre of Excellence, Slovenia for use of the Supernova diffractometer.

References

- 1 IDF Diabetes Atlas, International Diabetes Federation, 6th edn, 2013.
- 2 L. Guariguata, D. R. Whiting, I. Hambleton, J. Beagley, U. Linnenkamp and J. E. Shaw, *Diabetes Res. Clin. Pract.*, 2014, **103**, 137–149.
- 3 Y. Yoshikawa and H. Yasui, *Curr. Top. Med. Chem.*, 2012, **12**, 210–218.
- 4 C. T. Chasapis, A. C. Loutsidou, C. A. Spiliopoulou and M. E. Stefanidou, *Arch. Toxicol.*, 2011, **86**, 521–534.
- 5 K. Jurowski, B. Szewczyk, G. Nowak and W. Piekoszewski, *J. Biol. Inorg. Chem.*, 2014, **19**, 1–11.
- 6 G. Vardatsikos, N. R. Pandey and A. K. Srivastava, *J. Inorg. Biochem.*, 2013, **120**, 8–17.
- 7 E. Ciszak and G. D. Smith, *Biochemistry*, 1994, **33**, 1512–1517.
- 8 (a) D. C. Crans, J. J. Smee, E. Gaidamauskas and L. Yang, *Chem. Rev.*, 2004, **104**, 849–902; (b) D. Rehder, *Future Med. Chem.*, 2012, **4**, 1823–1837; (c) D. Rehder, *Dalton Trans.*, 2013, **42**, 11749–11761; (d) J. Costa Pessoa, E. Garribba, M. F. A. Santos and T. Santos-Silva, *Coord. Chem. Rev.*, 2015, **301–302**, 49–86; (e) D. Rehder, *Future Med. Chem.*, 2016, **8**, 325–338.
- 9 (a) E. Mrkalić, A. Zianna, G. Psomas, M. Gdaniec, A. Czapik, E. Coutouli-Argyropoulou and M. Lalia-Kantouri, *J. Inorg. Biochem.*, 2014, **134**, 66–75; (b) A. Zianna, G. Psomas, A. Hatzidimitriou, E. Coutouli-Argyropoulou and M. Lalia-Kantouri, *J. Inorg. Biochem.*, 2013, **127**, 116–126; (c) A. Tarushi, E. Polatoglou, J. Kljun, I. Turel, G. Psomas and D. P. Kessissoglou, *Dalton Trans.*, 2011, **40**, 9461–9473.
- 10 (a) A. Tarushi, X. Totta, A. Papadopoulos, J. Kljun, I. Turel, D. P. Kessissoglou and G. Psomas, *Eur. J. Med. Chem.*, 2014, **74**, 187–198; (b) A. Tarushi, Z. Karafllou, J. Kljun, I. Turel, G. Psomas, A. N. Papadopoulos and D. P. Kessissoglou, *J. Inorg. Biochem.*, 2013, **128**, 85–96.
- 11 A. Tarushi, K. Lafazanis, J. Kljun, I. Turel, A. A. Pantazaki, G. Psomas and D. P. Kessissoglou, *J. Inorg. Biochem.*, 2013, **121**, 53–65.



- 12 H. Kozłowski, M. Luczkowski, M. Remelli and D. Valensin, *Coord. Chem. Rev.*, 2012, **256**, 2129–2141.
- 13 D. Gambino, *Coord. Chem. Rev.*, 2011, **255**, 2193–2203.
- 14 (a) K. H. Thompson and C. Orvig, *J. Inorg. Biochem.*, 2006, **100**, 1925–1935; (b) K. H. Thompson, J. Lichter, C. LeBel, M. C. Scaife, J. H. McNeill and C. Orvig, *J. Inorg. Biochem.*, 2009, **103**, 554–558; (c) Y. Yoshikawa, E. Ueda, H. Miyake, H. Sakurai and Y. Kojima, *Biochem. Biophys. Res. Commun.*, 2001, **281**, 1190–1193; (d) Y. Adachi, J. Yoshida, Y. Kodera, A. Kato, Y. Yoshikawa, Y. Kojima and H. Sakurai, *J. Biol. Inorg. Chem.*, 2004, **9**, 885–893; (e) J. Costa Pessoa, S. Etcheverry and D. Gambino, *Coord. Chem. Rev.*, 2015, **301–302**, 24–48; (f) K. D. Mjos and C. Orvig, *Chem. Rev.*, 2014, **114**, 4540–4563.
- 15 H. Sakurai, K. Fujii, H. Watanabe and H. Tamura, *Biochem. Biophys. Res. Commun.*, 1995, **214**, 1095–1101.
- 16 Y. Yoshikawa, E. Ueda, K. Kawabe, H. Miyake, T. Takino, H. Sakurai and Y. Kojima, *J. Biol. Inorg. Chem.*, 2002, **7**, 68–73.
- 17 (a) S. Fujimoto, K. Fujii, H. Yasui, R. Matsushita, J. Takada and H. Sakurai, *J. Clin. Biochem. Nutr.*, 1997, **23**, 113–129; (b) Y. Fujisawa and H. Sakurai, *Chem. Pharm. Bull.*, 1999, **47**, 1668–1670.
- 18 (a) H. Sakurai, Y. Yoshikawa and H. Yasui, *Chem. Soc. Rev.*, 2008, **37**, 2383–2392; (b) H. Sakurai, Y. Kojima, Y. Yoshikawa, K. Kawabe and H. Yasui, *Coord. Chem. Rev.*, 2002, **226**, 187–198.
- 19 (a) J. Gätjens, B. Meier, T. Kiss, E. M. Nagy, P. Buglyó, H. Sakurai, K. Kawabe and D. Rehder, *Chem. – Eur. J.*, 2003, **9**, 4924–4935; (b) J. Gätjens, B. Meier, Y. Adachi, H. Sakurai and D. Rehder, *Eur. J. Inorg. Chem.*, 2006, 3575–3585; (c) H. Esbak, E. A. Enyedy, T. Kiss, Y. Yoshikawa, H. Sakurai, E. Garribba and D. Rehder, *J. Inorg. Biochem.*, 2009, **103**, 590–600.
- 20 Y. Wei, C. Zhang, P. Zhao, X. Yang and K. Wang, *J. Inorg. Biochem.*, 2011, **105**, 1081–1085.
- 21 (a) Y. Adachi, J. Yoshida, Y. Kodera, A. Katoh, J. Takada and H. Sakurai, *J. Med. Chem.*, 2006, **49**, 3251–3256; (b) M. Hiromura, Y. Adachi, M. Machida, M. Hattori and H. Sakurai, *Metallomics*, 2009, **1**, 92–100; (c) Y. Adachi, J. Yoshida, Y. Kodera and H. Sakurai, *Chem. Lett.*, 2005, **34**, 656–657; (d) Y. Yoshikawa, Y. Adachi and H. Sakurai, *Life Sci.*, 2007, **80**, 759–766.
- 22 (a) T. Moroki, Y. Yoshikawa, K. Yoshizawa, A. Tsubura and H. Yasui, *Metallomics*, 2014, **6**, 1632–1638; (b) M. E. López-Viseras, B. Fernández, S. Hilfiker, C. S. González, J. L. González, A. J. Calahorra, E. Colacio and A. Rodríguez-Diéguez, *J. Inorg. Biochem.*, 2014, **131**, 64–67; (c) N. Domingues, J. Pelletier, C.-G. Ostenson and M. M. C. A. Castro, *J. Inorg. Biochem.*, 2014, **131**, 115–122; (d) S. I. Pillai, S. P. Subramanian and M. Kandaswamy, *Eur. J. Med. Chem.*, 2013, **63**, 109–117; (e) Y. Yoshikawa, A. Murayama, Y. Adachi, H. Sakurai and H. Yasui, *Metallomics*, 2011, **3**, 686–692; (f) H. Kawarada, Y. Yoshikawa, H. Yasui, S. Kuwahara, Y. Habata and R. Saito, *Metallomics*, 2011, **3**, 675–679; (g) B. Fernández, A. Gómez-Vílchez, C. Sánchez-González, J. Bayón, E. San Sebastián, S. Gómez-Ruiz, C. López-Chaves, P. Aranda, J. Llopis and A. Rodríguez-Diéguez, *New J. Chem.*, 2016, **40**, 5387–5393; (h) Z.-Q. Han, S. Han and Y. Wang, *Acta Chim. Slov.*, 2016, **63**, 200–203; (i) J.-X. Lei, J. Wang, Y. Huo and Z. You, *Acta Chim. Slov.*, 2016, **63**, 670–677.
- 23 (a) M. Muneane, S. Motomura, S. Kamino, M. Ueda, H. Haba, Y. Yoshikawa, H. Yasui, M. Hiromura and S. Enomoto, *Biochem. Biophys. Rep.*, 2016, **5**, 211–215; (b) S. Fujimoto, H. Yasui and Y. Yoshikawa, *J. Inorg. Biochem.*, 2013, **121**, 10–15.
- 24 D. Rehder, J. C. Pessoa, C. F. Geraldes, M. M. Castro, T. Kabanos, T. Kiss, B. Meier, G. Micera, L. Pettersson, M. Rangel, A. Salifoglou, I. Turel and D. Wang, *J. Biol. Inorg. Chem.*, 2002, **7**, 384–396.
- 25 T. Koleša-Dobravc, E. Lodyga-Chruscinska, M. Symonowicz, D. Sanna, A. Meden, F. Perdih and E. Garribba, *Inorg. Chem.*, 2014, **53**, 7960–7976.
- 26 W. Trakarnpruk and P. Hoonsart, *Chin. J. Catal.*, 2007, **28**, 290–292.
- 27 (a) Z. Otwinowski and W. Minor, *Methods Enzymol.*, 1997, **276**, 307–326; (b) CrysAlis PRO, Agilent Technologies, Yarnton, England, 2013.
- 28 (a) G. M. Sheldrick, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2015, **71**, 3–8; (b) A. Altomare, M. C. Burla, M. Camalli, G. L. Cascarano, C. Giacovazzo, A. Guagliardi, A. G. G. Moliterni, G. Polidori and R. Spagna, *J. Appl. Crystallogr.*, 1999, **32**, 115–119.
- 29 M. Nakai, H. Watanabe, C. Fujiwara, H. Kakegawa, T. Satoh, J. Takada, R. Matsushita and H. Sakurai, *Biol. Pharm. Bull.*, 1995, **18**, 719–725.
- 30 (a) S. I. V. Judge and C. T. Bever Jr., *Pharmacol. Ther.*, 2006, **111**, 224–259; (b) M. H. Rabadi, K. Kreymborg and A. S. Vincent, *Drugs R&D*, 2013, **13**, 175–181.
- 31 K. Nakamoto, *Infrared and Raman Spectra of Inorganic and Coordination Compounds*, Wiley, New York, 4th edn, 1986.
- 32 A. W. Addison, T. N. Rao, J. Reedijk, J. van Rijn and G. C. Verschoor, *J. Chem. Soc., Dalton Trans.*, 1984, 1349–1356.
- 33 (a) T. Jakusch, J. Costa Pessoa and T. Kiss, *Coord. Chem. Rev.*, 2011, **255**, 2218–2226; (b) S. Mehtab, G. Gonçalves, S. Roy, A. I. Tomaz, T. Santos-Silva, M. F. A. Santos, M. J. Romão, T. Jakusch, T. Kiss and J. C. Pessoa, *J. Inorg. Biochem.*, 2013, **121**, 187–195; (c) G. Gonçalves, I. Tomaz, I. Correia, L. F. Veiros, M. M. C. A. Castro, F. Avecilla, L. Palacio, M. Maestro, T. Kiss, T. Jakusch, M. H. V. Garcia and J. C. Pessoa, *Dalton Trans.*, 2013, **42**, 11841–11861; (d) D. Sanna, M. Serra, G. Micera and E. Garribba, *Inorg. Chem.*, 2014, **53**, 1449–1464; (e) D. Sanna, G. Micera and E. Garribba, *Inorg. Chem.*, 2013, **52**, 11975–11985; (f) E. Lodyga-Chruscinska, G. Micera and E. Garribba, *Inorg. Chem.*, 2011, **50**, 883–899; (g) D. Sanna, G. Micera and E. Garribba, *Inorg. Chem.*, 2011, **50**, 3717–3728; (h) D. Sanna, G. Micera and E. Garribba, *Inorg. Chem.*, 2009, **48**, 5747–5757.
- 34 (a) T. Koleša-Dobravc, A. Meden and F. Perdih, *Monatsh. Chem.*, 2014, **145**, 1263–1275; (b) T. Koleša-Dobravc, A. Meden and F. Perdih, *Acta Chim. Slov.*, 2015, **62**, 261–271; (c) T. Koleša-Dobravc, A. Meden and F. Perdih, *New J. Chem.*, 2015, **39**, 4265–4277.
- 35 G. Süß-Fink, S. Stanislas, G. B. Shul'pin, G. V. Nizova, H. Stoeckli-Evans, A. Neels, C. Bobillier and S. Claude, *J. Chem. Soc., Dalton Trans.*, 1999, 3169–3175.



- 36 M. Melchior, K. H. Thompson, J. M. Jong, S. J. Rettig, E. Shuter, V. G. Yuen, Y. Zhou, J. H. McNeill and C. Orvig, *Inorg. Chem.*, 1999, **38**, 2288–2293.
- 37 L.-C. Du, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2009, **65**, m106.
- 38 C. Sun, X. Zheng and L. Jin, *J. Mol. Struct.*, 2003, **646**, 201–210.
- 39 T. A. Zevaco, H. Görls and E. Dinjus, *Inorg. Chim. Acta*, 1998, **269**, 283–286.
- 40 D. Dobrzyńska, T. Lis and L. B. Jerzykiewicz, *Inorg. Chem. Commun.*, 2005, **8**, 1090–1093.
- 41 L. Croitor, D. Chisca, E. B. Coropceanu and M. S. Fonari, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2013, **69**, m454.
- 42 C.-L. Chen, J. M. Ellsworth, A. M. Goforth, M. D. Smith, C.-Y. Su and H.-C. zur Loye, *Dalton Trans.*, 2006, 5278–5286.
- 43 (a) H.-L. Seng, S.-T. Von, K.-W. Tan, M. J. Maah, S.-W. Ng, R. N. Z. R. A. Rahman, I. Caracelli and C.-H. Ng, *BioMetals*, 2009, **23**, 99–118; (b) J.-C. Daran, P. Lemoine and B. Viossat, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 2002, **58**, m210–m212.
- 44 Y. Kojima, Y. Yoshikawa, E. Ueda, N. Kishimoto, M. Tadokoro and H. Sakurai, *Bull. Chem. Soc. Jpn.*, 2005, **78**, 451–455.
- 45 N. Okabe and Y. Muranishi, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2003, **59**, m244–m246.
- 46 R. Ferreira-Martínez, D. Esteban-Gómez, É. Tóth, A. de Blas, C. Platas-Iglesias and T. Rodríguez-Blas, *Inorg. Chem.*, 2011, **50**, 3772–3784.
- 47 Y.-C. Liang, M.-C. Hong, R. Cao and J.-B. Weng, *Chin. J. Struct. Chem.*, 2001, **20**, 455–458.
- 48 Y.-W. Tao and Y. Wang, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2011, **67**, m908.
- 49 Y. Yoshikawa, H. Sakurai, D. C. Crans, G. Micera and E. Garribba, *Dalton Trans.*, 2014, **43**, 6965–6972.
- 50 X. Yang, K. Wang, J. Lu and D. C. Crans, *Coord. Chem. Rev.*, 2003, **237**, 103–111.
- 51 A. Levina, A. I. McLeod, S. J. Gasparini, A. Nguyen, W. G. M. De Silva, J. B. Aitken, H. H. Harris, C. Glover, B. Johannessen and P. A. Lay, *Inorg. Chem.*, 2015, **54**, 7753–7766.

