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CORRECTION

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Correction: Mono- and tri-ester hydrogenolysis using tandem catalysis. Scope and mechanism

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Correction for 'Mono- and tri-ester hydrogenolysis using tandem catalysis. Scope and mechanism' by Tracy L. Lohr *et al.*, *Energy Environ. Sci.*, 2016, **9**, 550–564.

The authors wish to correct Fig. 4 and associated discussion on p. 555 of the manuscript. The authors discovered a small Excel error in the calculation of Nt. We have therefore replaced Fig. 4 with a new graph that plots % conversion after 1 h. This correction does not affect any other data or change any of the scientific conclusions in the article.

The paragraph on p. 555 should read as follows:

"Substituent effects on the acyl and alkoxy group were also investigated (Fig. 4 and 5). Sequentially replacing the methyl hydrogen atoms of cyclohexyl acetate with methyl groups (R = Me < Et < ^{i}Pr < ^{t}Bu; Fig. 4, blue columns) monotonically depresses the conversion after 1 h, under identical reaction conditions, from 89% (Me) to 33% (^{t}Bu). Note however that replacing an acyl H with a single electron-withdrawing chloro group increases the conversion to 93% (CH_2Cl), while including a second chloro group ($CHCl_2$) depresses the conversion to 61%. Fully chlorinating the methyl group (CCl_3) further depresses the conversion to 44%. These results indicate that the most reactive acyl moieties are those with relatively unencumbered electron-withdrawing substituents. Furthermore, appending even more strongly electron-withdrawing substituents such as trifluoro (CF_3) substantially increases the conversion relative to R = CH_3. In sum, these results indicate that electron-withdrawing groups which stabilize the $RCOO^-$ negative charge most enhance the ester hydrogenolysis turnover frequency."

Fig. 4 should appear as follows:

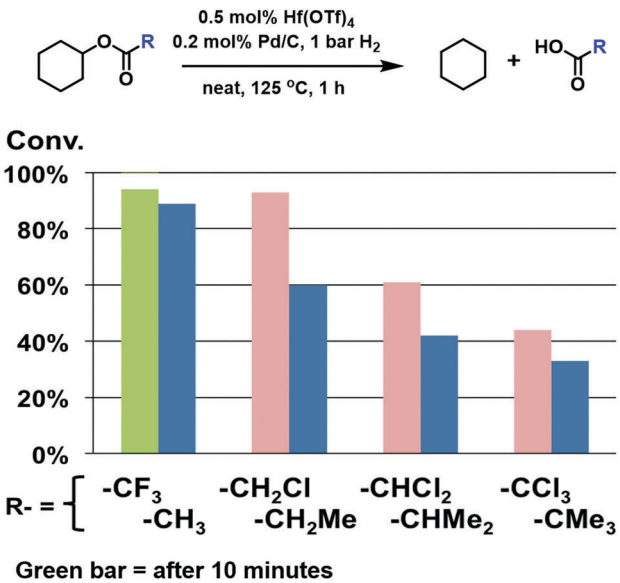

Fig. 4 Influence of acyl substituent on ester hydrogenolysis activity.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

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