



Cite this: *Dalton Trans.*, 2017, **46**, 10339

Synthesis and characterization of 1'-(diphenylphosphino)-1-isocyanoferrocene, an organometallic ligand combining two different soft donor moieties, and its Group 11 metal complexes†

Karel Škoch, ^a Ivana Císařová, ^a Jiří Schulz, ^a Ulrich Siemeling ^b and Petr Štěpnička ^{*a}

The development of a practical synthesis of 1'-(diphenylphosphino)-1-aminoferrocene (**2**) and its P-borane adduct (**2B**) allowed the facile preparation of 1'-(diphenylphosphino)-1-isocyanoferrocene (**1**). This compound combining two specific soft-donor moieties was studied as a ligand for univalent Group 11 metal ions. The reactions of **1** with AgCl at 1 : 1 and 2 : 1 molar ratios only led to the coordination polymer $[\text{Ag}_2(\mu\text{-Cl})_2(\mu(\text{P,C})\text{-1})_n]$ (**6**), while those with $\text{Ag}[\text{SbF}_6]$ provided the dimer $[\text{Ag}_2(\text{Me}_2\text{CO-}\kappa\text{O})_2(\mu(\text{P,C})\text{-1})_2][\text{SbF}_6]_2$ and the quadruply-bridged disilver complex $[\text{Ag}_2(\mu(\text{P,C})\text{-1})_4][\text{SbF}_6]_2$ (**8**), respectively. Addition of **1** to $[\text{AuCl}(\text{tht})]$ (tht = tetrahydrothiophene) afforded the mono- and the digold complex, $[\text{AuCl}(\text{1-}\kappa\text{P})]$ (**9**) and $[(\mu(\text{P,C})\text{-1})(\text{AuCl})_2]$ (**10**), depending on the reaction stoichiometry. Finally, the reaction of **1** with $[\text{Au}(\text{tht})_2][\text{SbF}_6]$ or halogenide removal from **9** with AgNTf_2 led to cationic dimers $[\text{Au}_2(\mu(\text{P,C})\text{-1})_2]\text{X}_2$ (**11**, X = SbF_6 (**a**) or NTf_2 (**b**)). Catalytic tests in the Au-mediated isomerization of (Z)-3-methylpent-2-en-4-yn-1-ol to 2,3-dimethylfuran revealed that **11a** and **11b** are substantially less catalytically active than their analogues containing 1'-(diphenylphosphino)-1-cyanoferrocene as the ligand, most likely due to a stronger coordination of the isonitrile moiety, which prevents dissociation of the dimeric complexes into catalytically active monomeric species.

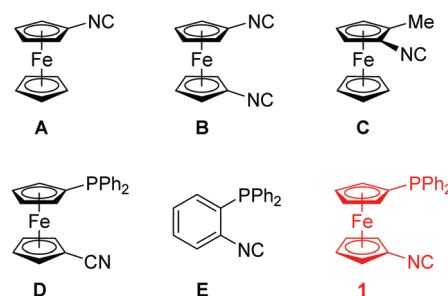
Received 28th June 2017,
Accepted 16th July 2017

DOI: 10.1039/c7dt02336g

rsc.li/dalton

Introduction

Isocyanides are useful synthetic intermediates¹ as well as attractive σ -donor/ π -acceptor ligands for coordination chemistry.² Even though the ferrocene motif has been frequently used in ligand design,³ ferrocene-based isocyanides were long overlooked. For instance, isocyanoferrocene (**A**, Scheme 1) was first reported in 1987^{4,5} (*i.e.*, relatively late for a simple ferrocene derivative⁶) and later studied as a redox-active metallo-ligand in transition metal complexes.^{4,7} More recent studies focused on its doubly functionalized congener, 1,1'-diisocyanoferrocene (**B**),⁸ disclosed unexpected coordination and reactivity of this compound in multinuclear complexes and for-



Scheme 1

^aDepartment of Inorganic Chemistry, Faculty of Science, Charles University, Hlavova 2030, 128 40 Prague 2, Czech Republic. E-mail: petr.stepnicka@natur.cuni.cz

^bInstitute of Chemistry, University of Kassel, Heinrich-Plett-Str. 40, 34132 Kassel, Germany

† Electronic supplementary information (ESI) available: Selected crystallographic data, additional structural drawings and copies of the NMR spectra. CCDC 1558580–1558590. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7dt02336g

mation of self-assembled monolayers on gold surfaces.⁹ Presumably the latest entry into the still narrow class of ferrocene-based isonitriles¹⁰ is the planar-chiral derivative, (*S_p*)-2-methyl-1-isocyanoferrocene (**C**).¹¹

Recently we reported the synthesis of 1'-(diphenylphosphino)-1-cyanoferrocene (**D**) and the coordination properties of this soft-soft hybrid donor toward Group 11 metals.¹² The often unusual structures of the isolated complexes as well as

excellent catalytic properties of the Au(I)-**D** complexes^{12b} led us to expand our studies toward the closest isomeric compound, 1'-(diphenylphosphino)-1-isocyanoferrocene (**1** in Scheme 1), which formally falls between diisocyanide **B** and the widely studied 1,1'-bis(diphenylphosphino)ferrocene (dppf).¹³

Although the transposition of the nitrogen and carbon atoms in the structures of **D** and **1** might appear marginal in terms of chemical constitution, it can be expected to dramatically alter the reactivity and coordination properties of these compounds. It is also noteworthy that similar ligands combining isocyanide and phosphine donor moieties remain limited to only a handful of compounds including Ph₂P(CH₂)_nNC (*n* = 2,¹⁴ 3¹⁵) and the recently reported compound **E** (Scheme 1).¹⁶

This contribution describes the preparation and structural characterization of the new phosphino-isonitrile ligand **1** and its complexes with Ag(I) and Au(I) ions. Also reported are the results of preliminary catalytic tests with Au(I)-**1** complexes aimed at a comparison with their Au(I)-**D** counterparts.

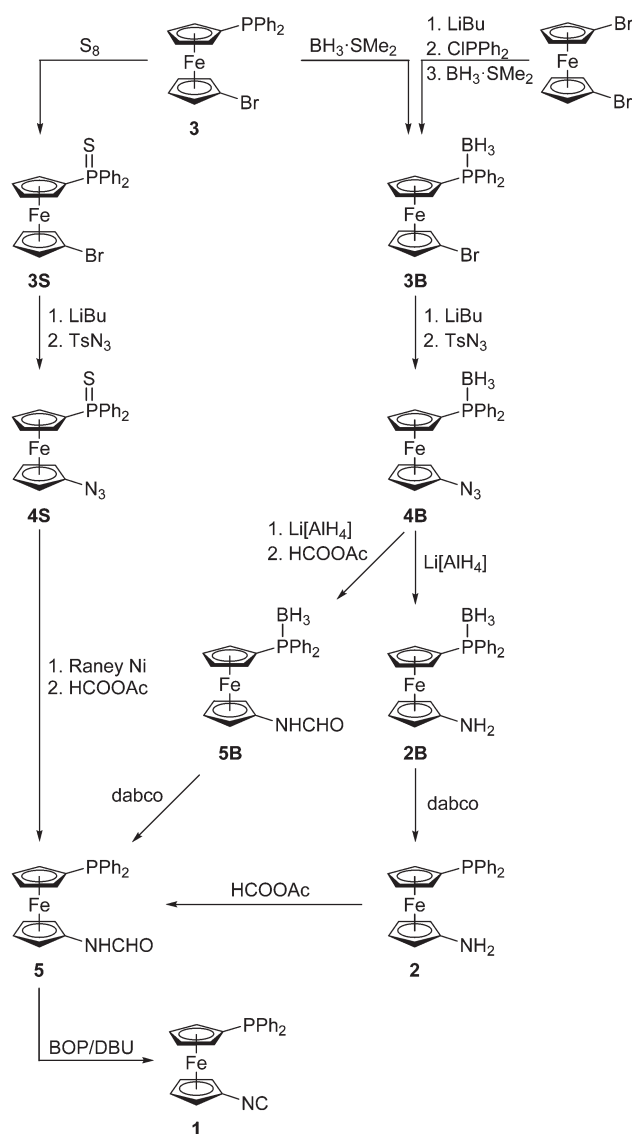
Results and discussion

Synthesis of phosphino-isonitrile **1**

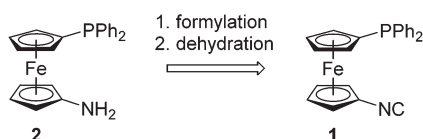
In analogy to the syntheses of organic isocyanides,¹⁷ we anticipated that our target compound **1** can be obtained from the known 1'-(diphenylphosphino)-1-aminoferrocene (**2**) by formylation and subsequent dehydration (Scheme 2). Although this assumption later proved correct, this straightforward plan had to be modified because the synthesis of amine **2** from 1'-(diphenylphosphino)-1-bromoferrocene (**3**)¹⁸ described in the literature¹⁹ was found to be irreproducible.^{20,21}

Considering the generally low stability of aminoferrocene derivatives, the often harsh reaction conditions required to obtain them,²² and also the sensitivity of phosphine groups to oxidation, the synthetic routes toward **1** were redesigned to include phosphine-protected intermediates that allow the introduction of the azide moiety (without an unwanted Staudinger reaction) and its subsequent reduction to amine group (Scheme 3).

Because some of us have successfully employed the protection of a phosphine group by thionation during the synthesis of Ph₂PfcNHCH₂*t*-Bu^{21a} (fc = ferrocene-1,1'-diyl), a similar approach was adopted also for the synthesis of **1** (Scheme 3, left branch) with the prospect that removal of the P-bound sulfur atom and reduction of the azide group (–N₃ → –NH₂) can be effected in a single step by reacting intermediate **4S** with Raney nickel. The starting phosphine bromide **3** was smoothly thionated with elemental sulfur and the stable



Scheme 3 Alternative syntheses of **1** exploiting different P-protected intermediates. Legend: TsN₃ = 4-toluenesulfonyl azide, BOP = (benzotriazol-1-yloxy)tris(dimethylamino)phosphonium hexafluorophosphate, dabco = 1,4-diazabicyclo[2.2.2]octane, DBU = 1,8-diazabicyclo[5.4.0]undec-7-ene.



Scheme 2 Intended preparation of **1** from amine **2**.

P-sulfide **3S** was lithiated and treated with 4-toluenesulfonyl azide (TsN₃) to afford azide **4S** in a good yield. A small amount of TsN₃ present in the reaction product did not interfere during the following step, after which it could be easily removed. The subsequent reaction of **4S** with Raney nickel and then with acetic formic anhydride provided the desired formamide **5** as a ca. 4 : 5 mixture of (*E*) and (*Z*) isomers (N.B. the acylation was performed directly to prevent a possible decomposition of intermediate amine **2**). However, the transformation of **4S** into **5** typically suffered from low yields (below 30%) and the product was contaminated by the dephosphorylated amide, FcNHCHO (Fc = ferrocenyl).²³

In the last step, which was the dehydration of formamide **5** to isonitrile **1**, the standard reagents (POCl₃/NEt₃, COCl₂/



$i\text{Pr}_2\text{NH}$)²⁴ failed, typically resulting in extensive darkening of the reaction mixture and oxidation of the phosphine moiety (confirmed by ³¹P NMR analysis). This led us to utilize the method previously used to convert oxime $\text{Ph}_2\text{PfcCH=NOH}$ into nitrile **D**.²⁵ Gratifyingly, addition of Castro's reagent (*i.e.*, (benzotriazol-1-yloxy)tris(dimethylamino)phosphonium hexafluorophosphate, BOP) and 1,8-diazabicyclo(5.4.0)undec-7-ene (DBU) to amide **5** ensued in a clean formation of isocyanide **1**, which was isolated in a 71% yield following chromatographic purification.

Although the synthesis of **1** was successful, we sought for a more efficient synthetic procedure, utilizing borane as the protecting group (Scheme 3, right branch). The necessary starting material **3B** was prepared by the reaction of phosphine-bromide **3** with $\text{BH}_3\cdot\text{SMe}_2$ or, alternatively, *via* a practical one-pot procedure (lithiation, phosphinylation, borylation) from 1,1'-dibromoferrocene.¹⁸ Lithiation of **3B** followed by quenching with TsN_3 afforded azide **4B** in a 89% yield. Compound **4B** was found to be stable when stored at low temperatures in the dark but gradually decomposed upon exposure to direct day light. The subsequent reduction with $\text{Li}[\text{AlH}_4]$ in THF smoothly converted azide **4B** into the protected amine **2B**. This compound was in turn formylated to give **5B** or isolated and converted the free phosphinoamine **2**. Deprotection of **5B** with 1,4-diazabicyclo[2.2.2]octane (dabco) or, alternatively, acylation of **2** with HCOOAc afforded formamide **5**, which was in turn used to prepare the target isocyanide **1** as outlined above (see Scheme 3).

Although longer by one step, the synthesis of **1** *via* borane-protected intermediates proved to be much more efficient in terms of the overall yield (compare the 70% yield of **1** from bromide **3** with *ca.* 25% yield for the $3 \rightarrow 1$ conversion *via* the $\text{P}=\text{S}$ intermediates) and selectivity. Besides, it provided a reliable access to amine **2** and its P-protected form **2B** that are both attractive for further syntheses. Compound **2B** appears to be particularly stable in both the solid state and solution. For instance, this adduct did not undergo any detectable decomposition or BH_3 migration between the Lewis basic amine and phosphine groups when its solutions in CDCl_3 , CD_3CN and C_6D_6 were stored at room temperature for 2 days. Accordingly, the reaction of **2** with 1 equiv. of $\text{BH}_3\cdot\text{SMe}_2$ (in CDCl_3) produced adduct **2B** as the sole product (N.B. adduct $2\cdot 2\text{BH}_3$ was detected only when larger amounts of the borane source were employed).

Isocyanide **1** and all intermediates occurring *en route* to this compound were characterized by multinuclear NMR and IR spectroscopy, mass spectrometry and elemental analysis or HR MS measurements. In addition, the solid state structures of **1**, **2**, **4S**, **4B**, and **5** were determined by single-crystal X-ray diffraction analysis. The ¹H and ¹³C NMR spectra confirmed the presence of an unsymmetrically 1,1'-disubstituted ferrocene moiety and the attached phosphorus groups in all cases. The manipulations at the phosphorus substituent were manifested in the ³¹P NMR spectra, showing broad signals at $\delta_{\text{P}} \approx 16$ for the BH_3 adducts, and sharp singlets at $\delta_{\text{P}} \approx -17$ and 42 for the free phosphines and phosphine sulfides, respectively. In addition to characteristic B–H stretching bands of the borane adducts

at 2340–2400 cm^{-1} ,²⁶ the IR spectra showed distinct features due to the polar substituents, namely intense azide $\nu(\text{N}_3)$ bands (**4S**: 2109 cm^{-1} , **4B**: 2108 cm^{-1}) and amide $\nu(\text{C}=\text{O})$ vibrations (**5** and **5B**: 1663 cm^{-1}). The band attributable to the $\nu(\text{N}=\text{C})$ vibration of **1** was seen at 2126 cm^{-1} (*cf.* 2125 cm^{-1} for PhNC).²⁷

The crystal structures of **1** and intermediates **4S**, **4B**, **2** and **5**

The molecular structures of azides **4S** and **4B** are presented in Fig. 1 and selected geometric parameters are given in Table 1. Compound **4S** crystallizes with two structurally independent molecules, which differ in the conformation of their ferrocene units (see the τ angles in Table 1) and in the orientation of the phenyl rings with respect to the ferrocene moiety (for an overlap of the independent molecules, see the ESI, Fig. S4†). Otherwise, the parameters describing the molecular geometry of the azides are similar to the respective data reported for FcN_3 ,²⁸ $\text{fc}(\text{N}_3)_3$,²⁹ $\text{FcP}(\text{S})\text{Ph}_2$,³⁰ dppfS_2 ,³¹ and $\text{dppf}\cdot 2\text{BH}_3$ ³² (Fc = ferrocenyl).

The structures of **1** and **2** are displayed in Fig. 2. Relevant structural data are given in Table 2. The ferrocene unit in the

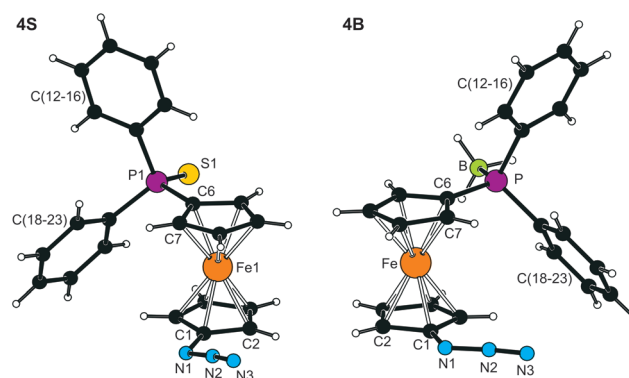


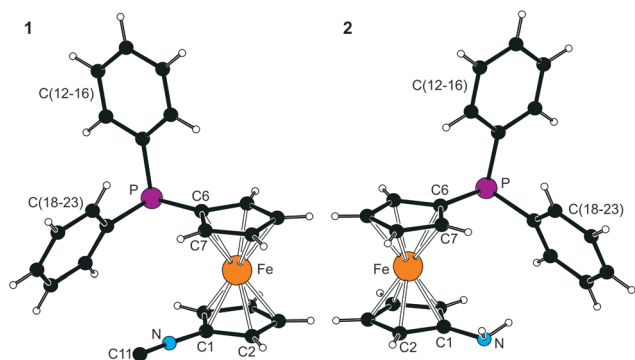
Fig. 1 Molecular structures of **4S** (molecule **1**) and **4B**.

Table 1 Selected distances and angles for **4S** and **4B** (in Å and °)^a

Parameter	4S (Y = S)		4B (Y = B)
	mol 1	mol 2	
Fe–Cg1	1.652(1)	1.649(2)	1.654(1)
Fe–Cg2	1.646(1)	1.632(1)	1.6464(9)
∠Cp1, Cp2	1.4(1)	1.3(2)	2.6(1)
τ	–122.8(2)	154.9(2)	108.1(2)
P–Cn6	1.790(2)	1.790(3)	1.790(2)
P–Y	1.9567(8)	1.950(1)	1.926(2)
Cn1–N ^α	1.415(4)	1.417(5)	1.422(3)
N ^α –N ^β	1.235(4)	1.251(4)	1.262(3)
N ^β –N ^γ	1.130(4)	1.143(5)	1.123(4)
Cn1–N ^α –N ^β	115.6(2)	114.9(3)	115.4(2)
N ^α –N ^β –N ^γ	173.3(3)	171.3(4)	172.1(3)

^a Definitions: Cp1 and Cp2 are the nitrogen- and phosphorus-substituted cyclopentadienyl rings, respectively. Cg1 and Cg2 denote their centroids. τ = torsion angle C1–Cg1–Cg2–C6. **4S**: n = void/5 for molecules 1/2, **4B**: n = void.



Fig. 2 Molecular structures of **1** and **2**.Table 2 Selected distances and angles for **1**, **2** and **5** (in Å and °)^a

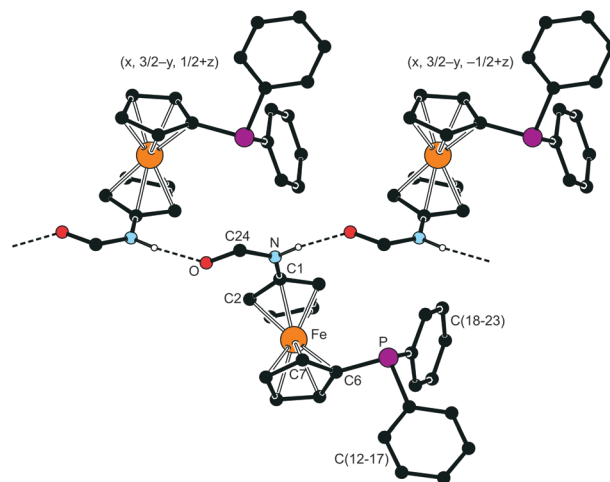
Parameter	1 ^b	2	5 ^c
Fe–Cg1	1.6438(8)	1.6585(7)	1.6482(7)
Fe–Cg2	1.6431(8)	1.6467(7)	1.6408(6)
∠Cp1,Cp2	1.4(1)	1.09(9)	2.79(9)
τ	−70.0(1)	69.0(1)	−86.3(1)
P–C6	1.814(2)	1.810(1)	1.813 (1)
C1–N	1.391(2)	1.410(2)	1.407(2)

^a Definitions: Cp1 and Cp2 are the cyclopentadienyl rings C(1–5) and C(6–10), respectively. Cg1 and Cg2 denote their centroids. τ = torsion angle C1–Cg1–Cg2–C6. ^b Further data: C11–N = 1.157(3), C1–N–C11 = 177.3(2). ^c Further data: N–C24 = 1.338(2), C1–N–C24 = 126.5(1), C24–O = 1.227(2), N–C24–O = 125.1(2).

structure of **1** has a synclinal eclipsed conformation (compare the τ angle in Table 2 with the ideal value of 72°)^{13a} and is negligibly tilted. Its phosphine substituent is oriented so that one phenyl group points above the ferrocene scaffold while the other as well as the lone pair of electrons are directed to the sides. The length of the N≡C bond of 1.157(3) is identical to that in FeNC (1.157(5) Å)³³ and the C1–N≡C11 moiety is linear (177.3(2)°).

The arrangement of the phosphinoferrrocene moiety in the structure of amine **2** is practically the same as in **1**, including the orientation of the cyclopentadienyl rings, and the C1–N bond in **2** has the same length as that in aminoferrrocene (1.405(5) Å).³⁴ Notably, the crystal structure of amine **2** lacks classical N–H...N hydrogen bonds. However, one of the amine hydrogen atoms (H1N) appears to be involved in an intra-molecular N–H...π interaction, being directed towards the phenyl ring C(18–23) with an H...Cg distance (Cg is the ring centroid) of 2.93 Å.

Recrystallization of formamide **5** from chloroform/hexane furnished crystals of the major (*Z*) isomer, which were used for the structure determination. In the crystal, compound **5** forms infinite chains *via* N–H1N...O hydrogen bridges (N...O = 2.808(2) Å) between molecules distributed around crystallographic glide planes (Fig. 3 and Table 2). The individual molecules show a minor tilting at their ferrocene units (*ca.* 3°) and an intermediate conformation between synclinal eclipsed and

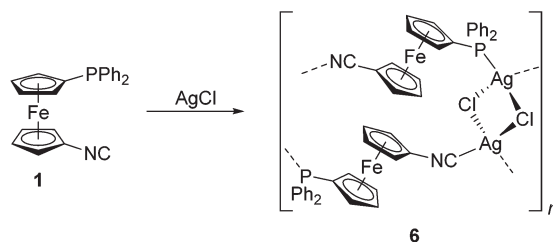
Fig. 3 Section of the hydrogen-bonded chain in the structure of **5** as viewed along the crystallographic *a* axis. For clarity, the CH_n hydrogen atoms are omitted.

anticlinal staggered. The amide moiety is coplanar with its parent cyclopentadienyl ring (dihedral angle: 0.8(2)°), which suggests conjugation of these molecular parts.

Complexes of **1** with Group 11 metals

With reference to our studies focused on the coordination behavior of phosphinonitrile donor **D** (Scheme 1),¹² we set out to investigate the coordination properties of the isomeric ligand **1** towards univalent Group 11 metal ions. Unfortunately, reactions of **1** with CuCl and [Cu(MeCN)₄]X (X = BF₄ and PF₆) as the common Cu(I) sources repeatedly failed to provide any defined products, typically giving rise to intractable precipitates (presumably of polymeric nature).

A defined, though also practically insoluble, product (N.B. the compound did not dissolve even in MeCN or DMSO) resulted upon reacting ligand **1** (1 or 2 equiv.) with freshly prepared AgCl (Scheme 4). X-ray diffraction analysis on crystalline **6** obtained by reaction of AgCl with 2 equiv. of **1** in chloroform and crystallization by addition of hexane revealed that the structure of **6** is built up from diamond-shaped {Ag₂(μ-Cl)₂} fragments that are interconnected by four P,C-bridging ligands **1** into an infinite one-dimensional chain (Fig. 4). Although central Ag₂Cl₂ motifs were found in the structures of, *e.g.*, [(Ph₃P)₂Ag(μ-Cl)]₂³⁵ and [(dppf-κ²P,P')Ag(μ-Cl)]₂,³⁶ the only

Scheme 4 Reaction of **1** with AgCl leading to coordination polymer **6**.

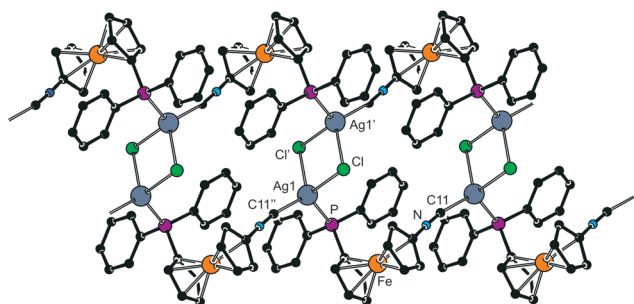


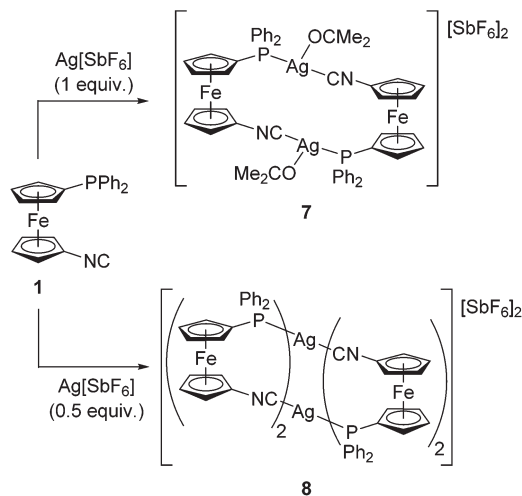
Fig. 4 A section of the infinite polymeric array in the structure of **6**. For clarity, all H atoms and the less populated position of the disordered silver(i) ion were omitted. Coordination geometry parameters (in Å and °): Ag1–P 2.4529(7), Ag1–C11 2.164(3), Ag1–Cl 2.6428(7), Ag1–Cl' 2.6358(8), P–Ag1–C11 117.71(7), P–Ag1–Cl 111.46(2), P–Ag1–Cl' 103.46(2), C11'–Ag1–Cl 110.35(7), C11'–Ag1–Cl' 120.30(8), Ag1–Cl–Ag1' 89.76(2), Cl–Ag1–Cl' 90.24(2).

related compound featuring two *different* donors whose structure has been reported to date is the dinuclear complex $[(\text{Ph}_3\text{P})(\text{MeNC})\text{Ag}(\mu\text{-Br})]_2$.³⁷ Indeed, the structures of **6** and the latter compound are similar except that in **6** the two donor moieties are tethered by the ferrocene unit, which in turn results in the formation of a polymeric assembly. It is also noteworthy that ligand **D** comprising the relatively harder cyano group reacted with AgCl to give heterocubane $[\text{Ag}(\mu_3\text{-Cl})(\text{D-}\kappa\text{P})]_4$ (Ag : D = 1 : 1) or the dimers $[\text{Ag}(\text{D-}\kappa\text{P})_2(\mu\text{-Cl})]_2$ and $[\text{AgCl}(\text{D-}\kappa\text{P})(\mu(\text{P},\text{N})\text{-D})]_2$ (Ag : D = 1 : 2).^{12c}

The structure of **6** is presented in Fig. 4 along with selected distances and angles. Although the silver atoms appear disordered over two positions differently positioned between the phosphine and isonitrile donors (refined ratios: $\approx 97 : 3$), the overall structure is rather symmetrical. The central Ag_2Cl_2 units are practically regular squares and lie on crystallographic inversion centers. Each silver(i) ion is further coordinated by two phosphino-isonitrile ligands **1**, once *via* the phosphorus atom and once through the isonitrile moiety. Both Ag–P and Ag–C bonds are shorter than the sum of the respective covalent radii ($\sum r_{\text{cov}}(\text{Ag}, \text{P}) = 2.52$ Å, $\sum r_{\text{cov}}(\text{Ag}, \text{C}_{\text{sp}}) = 2.14$ Å)³⁸ and the observed Ag–C distance (2.164(3) Å) is close to the average Ag–C distances in Ag–isonitrile complexes [2.14(8) Å (ref. 39)]. The coordination environment of the silver ions is distorted tetrahedral. Among the interligand angles, the Cl–Ag1–Cl' angle is the most acute (90°) and the C11'–Ag1–Cl (120°) and P–Ag1–C11' (118°) angles are the most opened.

The ferrocene moiety in the structure of **6** is roughly perpendicular to the Ag_2Cl_2 ring (*cf.* the dihedral angle of the C(1–5) ring and the Ag_2Cl_2 plane of 82.6(1)°) and has an intermediate conformation with τ of $-52.5(2)^\circ$. Notably, the length of the coordinated $\text{C}\equiv\text{N}$ bond (1.163(4) Å) is indistinguishable within experimental error from that in free **1**.

Addition of ligand **1** (1 equiv.) to $\text{Ag}[\text{SbF}_6]$ containing a weakly coordinating anion in acetone led to the dimeric complex $[\text{Ag}_2\{\mu(\text{P},\text{C})\text{-1}\}_2(\text{Me}_2\text{CO-}\kappa\text{O})][\text{SbF}_6]_2$ (**7** in Scheme 5). This compound is structurally similar to $[\text{Ag}_2\{\mu(\text{P},\text{N})\text{-D}\}_2(\text{AcOEt-}\kappa\text{O})][\text{SbF}_6]_2$ obtained from $\text{Ag}[\text{SbF}_6]$ and ligand **D**.^{12c} Unlike the



Scheme 5 Synthesis of Ag(i)-**1** complexes from $\text{Ag}[\text{SbF}_6]$.

latter compound, however, complex **7** appears to be stable in solution, which was manifested by a sharp ^{31}P NMR signal at δ_{P} 6.8 split into a pair of concentric doublets by interactions with ^{109}Ag and ^{107}Ag ($^1J_{\text{Ag},\text{P}} = 709$ and 615 Hz). Coordination of the isonitrile moiety was indicated by a shift of the $\nu(\text{N}\equiv\text{C})$ band in the IR spectrum to 2210 cm^{-1} (by 84 cm^{-1} to higher energies with respect to **1**). The IR spectrum further confirmed the presence of coordinated acetone through a sharp band of the $\nu(\text{C}=\text{O})$ vibration at 1695 cm^{-1} , which is 20 cm^{-1} less than for neat acetone due to coordination.⁴⁰

Upon increasing the amount of the phosphino-isonitrile ligand to 2 molar equiv., $\text{Ag}[\text{SbF}_6]$ was smoothly converted to the quadruply-bridged disilver(i) complex $[\text{Ag}_2\{\mu(\text{P},\text{C})\text{-1}\}_4][\text{SbF}_6]_2$ (**8**), which resembles an analogous Cu(i) complex obtained from ligand **D**, $[\text{Cu}_2\{\mu(\text{P},\text{N})\text{-D}\}_4][\text{SbF}_6]_2$.^{12a} In its ^{31}P NMR spectrum, complex **8** gave rise to a broad doublet at $\delta_{\text{P}} -1.2$ ($^1J_{\text{Ag},\text{P}} = 207$ Hz).⁴¹ The IR band due to coordinated isonitrile groups was observed at 2167 cm^{-1} , *i.e.* shifted by *ca.* 40 cm^{-1} to higher energies compared to uncoordinated ligand **1**.

Compounds **7** and **8** have been structurally authenticated by X-ray diffraction analysis. The complex cations in the crystal structure of **7** (Fig. 5) lie around crystallographic inversion centers, which makes only their halves structurally independent. Their tricoordinate silver ions exhibit an almost linear arrangement of the ligating P and C atoms, with a minor distortion due to the coordinated acetone ($\text{P-Ag-C} \approx 166^\circ$). The Ag–P and Ag–C bonds in **7** are shorter than those in **6** (by 0.07 and 0.05 Å), suggesting stronger dative bonds in the former cationic complex. In contrast, the Ag–O distance is substantially elongated (compare $\text{Ag-O1S} \approx 2.63$ Å with $\sum r_{\text{cov}}(\text{Ag}, \text{O}) = 2.11$ Å), presumably due to a relatively weaker coordination and steric reasons.

The cyclopentadienyl rings in **7** are tilted by 2.4(2)° and adopt a conformation near to synclinal eclipsed ($\tau = 83.8(1)^\circ$), which brings the phosphine moiety to the side of the fc-NC moiety (*cf.* the angle subtended by the P–C6 and C1–N bonds of 79.7(2)°). Consequently, the P–Ag bond points to the same

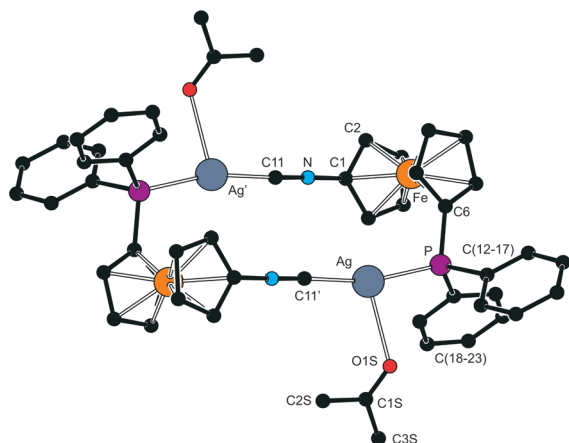


Fig. 5 View of the complex cation in the structure of **7**. The hydrogen atoms were omitted for clarity. Coordination geometry parameters (in Å and °): Ag–P 2.385(1), Ag–C11' 2.110(4), Ag–O1S 2.630(4), P–Ag–C11' 165.5(1), P–Ag–O1S 91.29(9), C11'–Ag–O1S 102.0(2).

direction as the $\text{N}\equiv\text{C}$ moiety, which in turn allows the formation of dimeric units $\text{Ag}_2(\mu(\text{P,C})\text{-1})_2$ consisting of two equivalent parts arranged in a side-by-side manner that laterally coordinate the acetone molecules.

The structure determination on $8\cdot 3\text{Me}_2\text{CO}$ (Fig. 6) revealed a compact and symmetrical complex molecule located around a crystallographic two-fold axis (space group $C2/c$; the silver atoms reside on this axis). Each silver(i) ion has a distorted tetrahedral P_2C_2 donor set being coordinated by four P,C-bridging ligands **1**. The lengths of the two structurally independent Ag–P bonds differ by *ca.* 0.036 Å. A more significant asymmetry observed for the Ag–CN bonds (0.059 Å) is not relayed further as the $\text{C}\equiv\text{N}$ bonds are similar in length and the Ag– $\text{C}\equiv\text{N}$ and Cp– $\text{N}\equiv\text{C}$ moieties remain linear. Despite the very different steric demands of the donor moieties, the interligand angles around the silver(i) ions do not depart much from the tetrahedral value, spanning a relatively narrow range of 103.35(9)–115.88(3)°.

The reactions of **1** with $[\text{AuCl}(\text{tht})]$ (tht = tetrahydrothiophene) paralleled the chemistry noted in the Au(i)–D system only partly (Scheme 6). Depending on the ligand-to-metal ratio, the reaction of **1** with $[\text{AuCl}(\text{tht})]$ afforded either the phosphine complex $[\text{AuCl}(\text{1-}\kappa\text{P})]$ (**9**) or the trinuclear Au_2Fe complex **10**. Both these products are poorly soluble (especially when crystallized), which makes their full characterization by solution techniques (*e.g.*, NMR) impossible. Nevertheless, the coordination of the phosphine moiety in **9** could be inferred from a shift of the ^{31}P NMR signal to a lower field (δ_{P} 29.0). The IR spectrum recorded with a solid sample (Nujol mull) displayed two bands attributable to the free (2126 cm^{-1}) and coordinated (2220 cm^{-1} , weak band; *cf.* $\nu_{\text{N}\equiv\text{C}}$ 2223 cm^{-1} for **10**) isonitrile moieties. Although unexpected, this observation is in line with the result of the structure determination revealing that a small fraction of the gold centers in the structure of **9** is displaced towards the isonitrile group of an adjacent molecule (*vide infra*). This phenomenon appears to be limited to

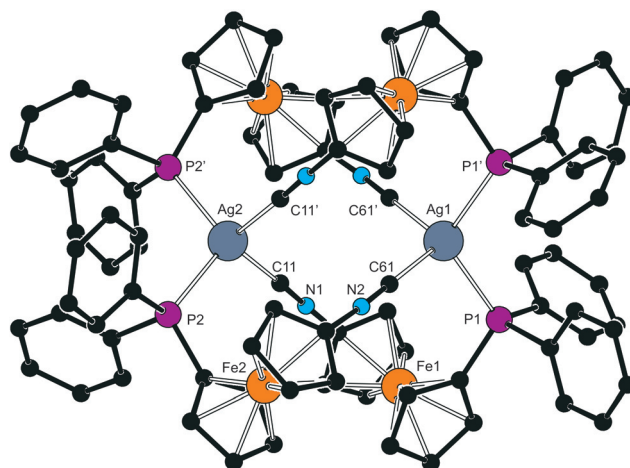
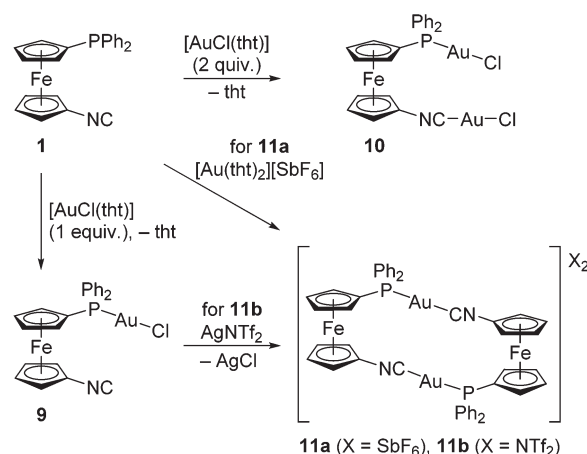


Fig. 6 View of the complex cation in the structure of $8\cdot 3\text{Me}_2\text{CO}$. The hydrogen atoms were omitted for clarity. Coordination geometry parameters (in Å and °): Ag1–P1 2.5376(9), Ag1–C61 2.264(3), Ag2–P2 2.5017(7), Ag2–C11 2.205(4), P1–Ag1–P1' 115.88(3), P1–Ag1–C61 103.35(9), P1–Ag1–C61' 112.44(9), C61–Ag1–C61' 109.5(1), P2–Ag2–P2' 109.65(2), P2–Ag2–C11 106.1(1), P2–Ag2–C11' 112.61(7), C11–Ag2–C11' 110.0(1); Ag2–C11–N1 177.9(3), N1–C11 1.150(5), C1–N1–C11 179.0(3), Ag1–C61–N2 178.5(3), N2–C61 1.148(3), C51–N2–C61 176.9(3). The prime-labeled atoms are generated by a crystallographic two-fold axis.



Scheme 6 Preparation of Au(i)–**1** complexes **9**–**11**.

the solid state as an NMR study ruled out the formation of isomeric monogold complexes (with P- and NC-bound ligand) and other species (*e.g.*, **1** and **10** formed by redistribution).

The structures of **9** and **10** are shown in Fig. 7 along with selected geometric data. As stated above, the gold(i) ions in the structure of **9** are disordered over two positions (refined occupancies: Au1 \approx 0.98, Au2 \approx 0.02; the ligand is considered invariant). In the dominant position, the gold(i) ion has the usual linear coordination whilst in the other it is approximately trigonal with the isonitrile moiety located at a bonding distance (Au2–C11 \approx 2.09 Å, $\sum r_{\text{cov}}(\text{Au}, \text{Csp}) = 2.05$ Å). The ferrocene unit in **9** has a synclinal eclipsed conformation ($\tau = 70.3(2)^\circ$) which, together with the spatial orientation of the phenyl



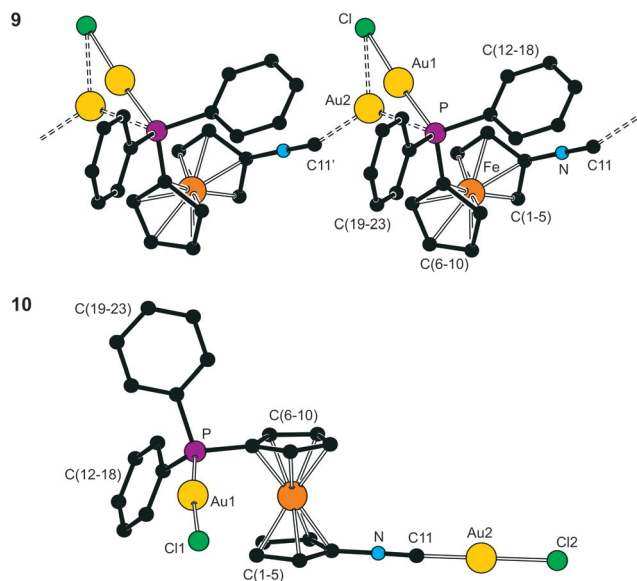


Fig. 7 Simplified views of the crystal structures of **9** (top) and **10** (bottom). Selected data for **9** (in Å and °): Au1–P 2.2298(7), Au1–Cl 2.2896(8), Au2–P 2.533(6), Au2–Cl 2.652(6), Au2–C11' 2.094(7), P–Au1–Cl 176.12(3), P–Au2–Cl 121.2(2), Au2'–N–C11 155.1(4), N–C11 1.153(4), C1–N–C11 178.9(3). Selected data for **10** (in Å and °): Au1–P 2.230(1), Au1–Cl1 2.288(1), P–Au1–Cl1 173.93(4), Au2–C11 1.930(5), Au2–Cl2 2.260(1), C11–Au2–Cl2 176.9(1), Au2–C11–N 178.2(4).

rings, allows for a compact arrangement of the linear assembly interlinked *via* the incidental P–Au2–NC interactions.

In the structure of **10**, both gold centers are linearly coordinated and averted from each other because the central ferrocene unit has an opened conformation with $\tau = -157.0(3)^\circ$ (N.B. the angle subtended by the Au1–Cl1 and Au2–Cl2 bonds is $139.26(6)^\circ$). Whereas the entire C1–N≡C11–Au2–Cl2 moiety is essentially linear, the Au1–Cl1 bond diverts from the plane of the parent cyclopentadienyl ring at an angle of $23.1(2)^\circ$. On the other hand, the coordination geometry parameters are not unexpected in view of the data reported for $[\text{Au}_2\text{Cl}_2(\mu(\text{P},\text{P}')\text{-dppf})]^{42}$ and $[\text{AuCl}(\text{D-}\kappa\text{P})]^{12c}$. In the crystal, the molecules of **10** assemble into head-to-tail dimers *via* weak Au1...Au2 aurophilic interactions (3.21 Å; see the ESI, Fig. S12†).⁴³

Finally, the reaction of $[\text{Au}(\text{tht})_2][\text{SbF}_6]$ with **1** (1 equiv.) or removal of the chloride ligand from **9** with AgNTf_2 resulted in the formation of the respective dimeric complexes **11** (Scheme 6), which are the structural isomers of the recently reported compounds $[\text{Au}_2(\mu(\text{P},\text{N})\text{-D})_2]\text{X}_2$ (X = SbF₆ and NTf₂).^{12b} The dimeric structure of **11a** and **11b** was corroborated by X-ray diffraction analysis (N.B. the structure of **11b** could not be adequately refined because of a severe disorder of the NTf₂[−] anion). The coordination of both donor moieties available in ligand **1** is indicated by a shift of the ³¹P NMR signals to low fields (δ_{P} 32–33) and through a shift of the $\nu_{\text{N}\equiv\text{C}}$ bands in the IR spectra to higher energies (2231–2233 cm^{−1}). Notably, both shifts are more pronounced than for the neutral complexes **9** and **10**, very likely owing to a stronger interaction of the donor moieties with the electron-poor cationic Au(I) centers.

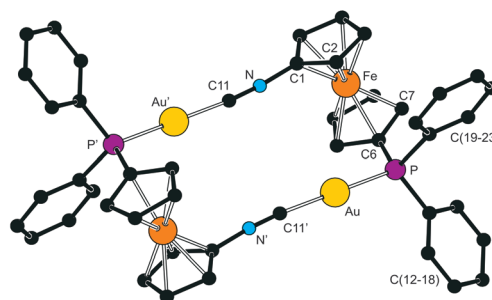


Fig. 8 View of the complex cation in the structure of **11a**·2Me₂CO. Hydrogen atoms, the counter anions and solvent molecules are omitted for clarity. Selected distances and angles (in Å and °): Au–P 2.2821(9), Au–C11' 2.008(4), P–Au–C11' 175.8(1), Au–C11'–N' 174.3(4), N–C11 1.135(5), C1–C11–N 173.4(4). The primed atoms are generated by crystallographic inversion.

The solid-state structure of **11a**·2Me₂CO shown in Fig. 8 resembles that of **7**, except that the coordination of the Au centers is practically undistorted linear (P–Au–C = $175.8(1)^\circ$). The ferrocene cyclopentadienyl rings adopt a conformation similar to that in **7** with τ of $-80.0(3)^\circ$, which brings the linear P–Au–NC subunits side-by-side (N.B. the angle subtended by the P–Au and C11–Au' bonds is only $4.17(1)^\circ$), though without any supporting intramolecular aurophilic interaction (Au...Au' = $5.4400(4)$ Å).

Electrochemistry

The electrochemical behavior of compound **1** and **2**, and of the Au(I) complexes **9–11** was studied by cyclic voltammetry on a glassy carbon disc electrode in dichloromethane containing Bu₄N[PF₆] as the supporting electrolyte.

Compound **1** (Fig. 9) was oxidized in a single irreversible step at 0.38 V vs. the ferrocene/ferrocenium reference couple (anodic peak potential, E_{pa} , is given at a scan rate of 100 mV s^{-1}),

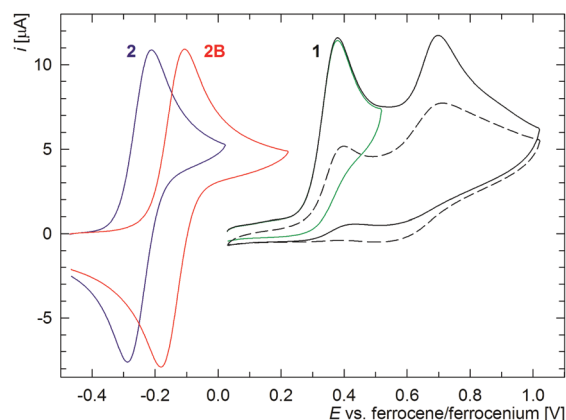


Fig. 9 Cyclic voltammograms of amine **2** (blue), its BH₃ adduct **2B** (red) and phosphino-isonitrile **1** (green – partial, and black – a wider scan) recorded with a glassy carbon disc electrode in CH₂Cl₂ (scan rate 0.1 V s^{-1}). The second sweep in the voltammogram of **1** is distinguished by a dashed line.



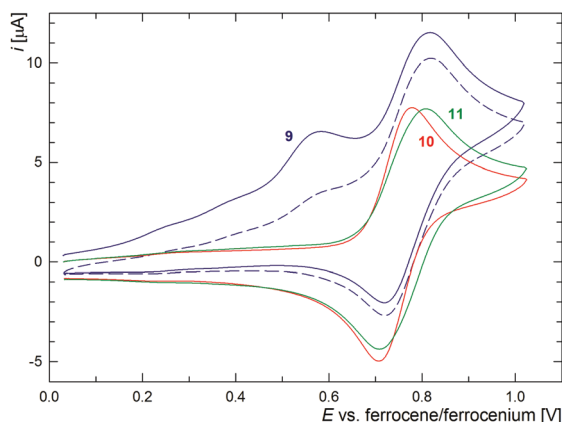
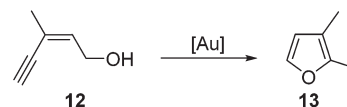


Fig. 10 Cyclic voltammograms of complexes **9** (blue, the second scan is shown as a dashed line), **10** (red) and **11a** (green) as recorded with a glassy carbon disc electrode in CH_2Cl_2 (scan rate 0.1 V s^{-1}).

which corresponds with an electron-withdrawing nature of the substituents at the ferrocene unit (*cf.* the Hammett σ_p constant: 0.46 for $-\text{NC}$, and 0.19 for $-\text{PPh}_2$).⁴⁴ This primary redox change attributable to the oxidation of the ferrocene moiety is followed by another irreversible redox event at a more positive potential ($E_{\text{pa}} \approx 0.75 \text{ V}$).

In contrast, the oxidation of amines **2** and **2B** was reversible (Fig. 9). However, the waves were shifted to less positive potentials ($E^{\circ'} = -0.24 \text{ V}$ and -0.13 V for **2** and **2B**, respectively) due to the electron-donating nature of the amine substituent ($\sigma_p = -0.66$). Even in this case, the first oxidations were followed by irreversible processes at higher potentials.

The cyclic voltammogram of the monogold complex **9** showed two redox events in the anodic region, namely an irreversible oxidation at $E_{\text{pa}} \approx 0.61 \text{ V}$ and a reversible redox process at $E^{\circ'} = 0.78 \text{ V}$ (Fig. 10). Since the digold complex **10** displayed a single reversible redox change at $E^{\circ'} = 0.76 \text{ V}$, it appears likely that the redox response of **9** can be affected by adsorption phenomena. A single reversible redox event was observed also for the dinuclear complexes **11a** and **11b**, whose behavior was expectedly identical ($E^{\circ'} = 0.79 \text{ V}$; see Fig. 10).



Scheme 7 Au-Catalyzed cycloisomerisation of enynol **12** to 2,3-dimethylfuran.

Table 3 Summary of the catalytic results^a

Complex	Au loading [mol.%]	Yield of 13 after 30 min [%]	Yield of 13 after 3 h [%]
9	1.0	0	0
10	1.0	82	89
10	0.1	64	75
11a	1.0	0	0
11b	1.0	0	0

^a Compound **12** (1 mmol) and 1,2-dichloroethane (internal standard; 1 mmol) were dissolved in CDCl_3 (2 mL). The respective catalyst was added and the reaction mixture was stirred at room temperature for a given time (the amount of the starting materials was doubled for reactions with 0.1 mol% Au). The yields were determined by integration of ^1H NMR spectra.

Catalytic tests with Au(I)-1 complexes

Considering the favorable catalytic properties of gold(i) complexes with ligand **D**,^{12b} we decided to probe the catalytic properties of the isomeric compounds featuring ligand **1**. For testing, we chose the gold-catalyzed cyclization of unsaturated alcohol **12** to give 2,3-dimethylfuran (**13**; Scheme 7) that has been used to evaluate the catalysts based on ligand **D**.

Unfortunately, the collected results (Table 3) indicate a generally very low (if any at all) catalytic activity of the prepared Au-1 complexes. For instance, compound **10** achieved a 82% yield of the cyclization product **13** after 30 min and a nearly 90% yield after 3 h at 1 mol% Au loading. These yields decreased substantially upon lowering the catalyst amount to 0.1 mol%. The monogold complex **9** and, surprisingly, also the dimers **11** proved to be catalytically inactive, which markedly contrasts with the very high catalytic activity of $[\text{Au}_2(\mu(\text{P,N})\text{-D})_2]\text{X}_2$ ($\text{X} = \text{SbF}_6, \text{NTf}_2$).^{12b}

Table 4 Comparison of the experimental (from X-ray diffraction analysis) and DFT computed structural parameters for cations $[\text{Au}_2(\mu(\text{P,N})\text{-L})_2]^{2+}$ ($\text{L} = \text{D}$ or **1**)^a

$[\text{Au}_2(\mu(\text{P,N})\text{-D})_2]^{2+}$			$[\text{Au}_2(\mu(\text{P,N})\text{-1})_2]^{2+}$		
Parameter	X-ray ^b	DFT	Parameter	X-Ray ^c	DFT
Au–N	2.035(4)	2.054	Au–C	2.008(4)	1.996
C≡N	1.139(8)	1.160	N≡C	1.135(5)	1.164
Au–P	2.225(2)	2.281	Au–P	2.282(1)	2.333
N–Au–P	175.1(2)	175.7	C–Au–P	175.8(1)	175.3
Au–N–C	168.2(5)	170.6	Au–C–N	174.3(4)	175.4

^a The calculations were performed at the PBE0/cc-pVDZ : sdd(Fe,Au) level of theory using the PCM model to treat possible solvent effects (CHCl_3). For details, see Experimental. ^b Data from ref. 12b, SbF_6^- salt. ^c This work; SbF_6^- salt.



Tentatively, we ascribed the observed low catalytic activity of the Au-1 complexes (particularly the dimers **11**) to a relatively stronger coordination of the isonitrile moiety, which prevents the formation of coordinatively unsaturated Au(I) species by dissociation of the Au–CN bonds. This assumption was supported by the results of DFT computations performed on the isolated cations $[\text{Au}_2(\mu(\text{P,N})\text{-L})_2]^{2+}$, where L = **D** or **1** (Table 4), which indeed revealed that the dissociation of the phosphino-nitrile-bridged dimers (L = **D**) to the monomeric species $[\text{Au}(\text{D})]^+$ is by approximately 12 kcal mol^{−1} less endergonic than for the corresponding dimeric cation containing the isomeric ligand **1** (27.2 vs. 39.4 kcal mol^{−1}).⁴⁵

Conclusions

The discovery of a practical and reliable synthesis of 1'-(diphenylphosphino)-1-aminoferrocene (**2**) and its P-protected analogues (the borane adduct **2B** in particular) reported in this contribution opens a way to further synthetic utilization of this versatile but long ignored phosphinoferrocene building block. Herein, the synthetic use of **2** is demonstrated by the synthesis of phosphino-isonitrile **1**, representing a new entry among the still uncommon mixed-donor ligands of this type. As revealed by our coordination study with Ag(I) and Au(I) ions with and without auxiliary chloride ligands (N.B. defined Cu(I) complexes were not isolated despite numerous attempts), the ferrocene moiety in the structure of **1** behaves as a molecular ball bearing connecting the two sterically distinct, soft donor moieties present in this ligand, allowing their orientation into positions appropriate for the formation of discrete, dimeric and even polymeric assemblies. Compared to the isomeric phosphino-nitrile donor **D**, ligand **1** becomes more easily (and also more strongly) coordinated through its isonitrile donor group. However, the stronger Au–CN bonding can be made responsible for a lower catalytic activity of the dimeric Au(I) complexes **11**, very likely due to a suppressed dissociation of these dimers into monogold(I) species. Further studies into the coordination behavior and possible synthetic applications of phosphino-isonitrile **1** focused mainly at its insertion chemistry are currently underway.

Experimental

Materials and methods

All syntheses were performed under an argon atmosphere using standard Schlenk techniques and with protection from the direct sun light. Compounds **3S**,⁴⁶ 4-toluenesulfonyl azide,⁴⁷ acetic formic anhydride,⁴⁸ $[\text{AuCl}(\text{tht})]^{49}$ and $[\text{Au}(\text{tht})_2][\text{SbF}_6]^{12b}$ were prepared according to the literature. Chloro-diphenylphosphine was distilled under vacuum. Anhydrous dichloromethane and tetrahydrofuran were obtained from a PureSolv MD5 solvent purification system (Innovative Technology, MA, USA). Toluene was dried over sodium metal and distilled. Other chemicals and solvents utilized for crystallizations and during

chromatography were of reagent grade (Sigma-Aldrich and Alfa-Aesar; solvents from Lach-Ner, Czech Republic) and were used without additional purification.

The NMR spectra were recorded at 25 °C on a Varian UNITY Inova 400 or a Bruker Avance III 400 spectrometer. Chemical shifts (δ in ppm) are given relative to an internal tetramethylsilane (¹H and ¹³C) and an external 85% H₃PO₄ standard (³¹P), all set to 0 ppm. In addition to the usual notation of signal multiplicity, vt and vq are used to distinguish virtual triplets and quartets due to ferrocene protons constituting the AA'BB' and AA'BB'X spin systems in the nitrogen- and PPh₂-substituted cyclopentadienyl rings, respectively (fc = ferrocene-1,1'-diyl). FTIR spectra were measured with a Nicolet 6700 spectrometer in a range of 400–4000 cm^{−1}. Electrospray ionization (ESI) mass spectra were obtained on an Esquire 3000 (Bruker) spectrometer using samples dissolved in HPLC-grade methanol. High resolution (HR) measurements were performed with an LTQ Orbitrap XL spectrometer (Thermo Fisher Scientific). The assignment of the observed ions was based on a comparison of the theoretical and experimental isotopic patterns. Elemental analyses were determined with a PerkinElmer PE 2400 CHN analyzer.

The cyclic voltammograms were recorded with a μ AUTOLAB III instrument (Eco Chemie, The Netherlands) at ambient temperature using samples dissolved in dry dichloromethane (sample concentration 0.5 mM or a saturated solution in the case of poorly soluble complexes) and Bu₄N[PF₆] as the supporting electrolyte (0.1 M). A glassy carbon disc (2 mm diameter) was employed as a working electrode, Ag/AgCl (3 M LiCl/EtOH) as a reference electrode and platinum sheet as a counter electrode. Decamethylferrocene was added as an internal reference for the final scans but the potentials were converted to the ferrocene/ferrocenium scale by subtracting by 0.548 V.⁵⁰

Synthesis of **4S**

An oven-dried two-necked flask was charged with **3S** (2.40 g, 5.0 mmol) and a magnetic stirring bar, flushed with argon and sealed a rubber septum. Dry THF (80 mL) was introduced under argon to dissolve the solid and the flask was cooled in an ice bath with continuous stirring. *n*-Butyllithium (2.0 mL 2.5 M, 5.0 mmol) was added and the mixture was allowed to react for 20 min with continuous cooling. Neat 4-toluenesulfonyl azide (TsN₃; 986 mg, 5.0 mmol) was introduced to the lithiated intermediate, causing the reaction mixture to turn red (Note: TsN₃ was stored over CaH₂ prior to the use. Reactions with the untreated commercial reagent gave considerably lower yields). The reaction mixture was stirred at 0 °C for 30 min and at room temperature for another 2 h. Then, it was diluted with saturated aqueous NaHCO₃ (40 mL) and diethyl ether (80 mL). The organic phase was separated and the aqueous one was extracted with diethyl ether (3 × 40 mL). The combined organic layers were washed with brine and dried over MgSO₄. The dried solution was evaporated with chromatographic silica gel and the crude pre-adsorbed product was transferred to a silica gel column. Elution with diethyl ether–hexane (1 : 3) provided an orange band containing the product. Evaporation afforded compound **4S** as an



orange oil (2.15 g, *ca.* 75%), which usually contained unreacted TsN_3 (less than 5 mol%) and FcP(S)Ph_2 (typically *ca.* 20 mol%; Fc = ferrocenyl) formed by protonation of the lithiated intermediate. Crystals used for single-crystal X-ray diffraction analysis were grown from diethyl ether. The product decomposes (darkens) upon prolonged storage (even when stored in a refrigerator and in the dark). If appropriate, it can be easily purified by chromatography as described above.

^1H NMR (CDCl_3): δ 4.07 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.23 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.54 (vq, $J' = 2.1$ Hz, 2 H, fc), 4.63 (vq, $J' = 2.1$ Hz, 2 H, fc), 7.40–7.51 (m, 6 H, P(S)Ph_2), 7.69–7.75 (m, 4 H, P(S)Ph_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 62.07 (CH of fc), 67.82 (CH of fc), 73.19 (d, $J = 10$ Hz, CH of fc), 73.84 (d, $J = 12$ Hz, CH of fc), 100.51 ($C^{\text{ipso}}\text{-N}_3$ of fc), 128.26 (d, $^2J_{\text{PC}} = 12$ Hz, CH^{ortho} of Ph), 131.31 (d, $^4J_{\text{PC}} = 2$ Hz, CH^{para} of Ph), 131.59 (d, $^3J_{\text{PC}} = 10$ Hz, CH^{meta} of Ph), 134.36 (d, $^1J_{\text{PC}} = 87$ Hz, C^{ipso} of Ph). The signals due to $C^{\text{ipso}}\text{-P}$ of fc was not found. $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ 41.7 (s). IR (Nujol): ν_{max} 3049 w, 2109 s (N_3), 1308 w, 1287 m, 1193 w, 1170 m, 1102 s, 1069 w, 1026 m, 998 w, 917 w, 837 m, 828 m, 754 m, 716 s, 695 s, 656 s, 629 w, 615 w, 541 m, 523 m, 496 m, 448 w cm^{-1} . ESI+ MS: m/z 438 ($[\text{M} - \text{N}_2 + \text{Na}]^+$), 466 ($[\text{M} + \text{Na}]^+$). HRMS calc. for $\text{C}_{22}\text{H}_{18}^{56}\text{FeN}_3\text{PS}$ (M^+): 443.0308, found 443.0309.

Synthesis of 5 from 4S

An aqueous suspension of active nickel (Sigma-Aldrich, RANEY®, type 2400) corresponding to 8 g of the solid material was transferred to a reaction flask and washed with reagent grade and then anhydrous THF (5-times each) under argon. The remaining solid was suspended in anhydrous THF (25 mL) and a solution of impure 4S (0.800 g; this amount corresponds to 1.8 mmol of pure 4S) in 10 mL of dry THF was added (an effervescence was observed during the addition). The reaction mixture was stirred at room temperature for 3 days and filtered through a Celite pad, which was then washed with THF (3×15 mL). Neat acetic formic anhydride (HCOOAc ; 0.8 mL, 10 mmol) was added to the filtrate and the resultant reaction mixture was stirred at ambient temperature for 3 h before quenching with saturated aqueous NaHCO_3 . After gas evolution had ceased, the mixture was extracted with diethyl ether (50 mL). The organic layer was washed with water and brine, dried over MgSO_4 and evaporated, leaving an orange residue, which was purified by column chromatography (silica gel, ethyl acetate–hexane 1 : 1). The first orange band containing FcPPh_2 and some tosylated impurities was followed by an orange band containing the desired product contaminated by traces of FcNHCHO .⁵ The second band was evaporated and the residue chromatographed again to afford pure 5 (orange solid; 0.253 g or 34% with respect to 2S available in the starting material) and FcNHCHO (orange solid; 55 mg, 13% with respect to 2S). The yield did not practically change when the reaction time was shortened to 1 day. Compound 5 was obtained as a mixture of (*Z*) and (*E*) isomers in 5 : 4 ratio. Crystals of (*Z*)-5 were obtained by recrystallization from chloroform/hexane.

Notably, when the reaction was performed in dry acetonitrile, the unwanted dephosphorylation leading to FcNHCHO

was suppressed but the yield of 5 was only 13%. Practically no desulfurization (only conversion of the azide to formamide) was observed in 1,2-dimethoxyethane under similar conditions (room temperature, 3 days), while in anhydrous ethanol the starting azide was converted to the *N*-alkylated products FcN(Et)CHO and $\text{Ph}_2\text{PfcN(Et)CHO}$ in *ca.* 1 : 1 ratio (ESI+ MS data for $\text{Ph}_2\text{PfcN(Et)CHO}$: m/z 442 ($[\text{M} + \text{H}]^+$), 464 ($[\text{M} + \text{Na}]^+$), 480 ($[\text{M} + \text{K}]^+$)).

Analytical data for 5

^1H NMR (CDCl_3): **major** – δ 3.96 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.08 (br s, 2 H, fc), 4.44 (vt, $J' = 1.8$ Hz, 2 H, fc), 4.52 (vt, $J' = 1.8$ Hz, 2 H, fc), 6.10 (br s, 1 H, NH), 7.30–7.42 (m, 10 H, PPh_2), 7.94 (d, $^3J_{\text{HH}} = 1.5$ Hz, 1 H, CHO); **minor** – δ 3.99 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.13 (vq, $J' = 1.7$ Hz, 2 H, fc), 4.16 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.42 (vt, $J' = 1.8$ Hz, 2 H, fc), 6.65 (br d, $^3J_{\text{HH}} = 11.7$ Hz, 1 H, NH), 7.30–7.42 (m, 10 H, PPh_2), 8.20 (d, $^3J_{\text{HH}} = 11.7$ Hz, 1 H, CHO). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): **major** – δ 63.22 (CH of fc), 65.97 (CH of fc), 72.23 (CH of fc), 74.26 (d, $J_{\text{PC}} = 8$ Hz, CH of fc), 75.99 (br s, $C^{\text{ipso}}\text{-P}$ of fc), 92.54 ($C^{\text{ipso}}\text{-N}$ of fc), 128.30 (d, $^3J_{\text{PC}} = 7$ Hz, CH^{meta} of Ph), 128.55 (CH^{para} of Ph), 133.47 (d, $^2J_{\text{PC}} = 20$ Hz, CH^{ortho} of Ph), 139.12 (d, $^1J_{\text{PC}} = 9$ Hz, C^{ipso} of Ph), 159.02 (CHO); **minor** – δ 62.88 (CH of fc), 66.76 (CH of fc), 72.21 (CH of fc), 74.16 (d, $J_{\text{PC}} = 8$ Hz, CH of fc), 93.74 ($C^{\text{ipso}}\text{-N}$ of fc), 128.30 (d, $^3J_{\text{PC}} = 7$ Hz, CH^{meta} of Ph), 128.82 (CH^{para} of Ph), 138.58 (d, $^2J_{\text{PC}} = 20$ Hz, CH^{ortho} of Ph), 138.58 (d, $^1J_{\text{PC}} = 9$ Hz, C^{ipso} of Ph), 163.06 (CHO). The signal due to $C^{\text{ipso}}\text{-P}$ of fc for the minor isomer appears to be obscured by the solvent resonance. $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ –17.5 (s, **major**), –16.8 (s, **minor**). IR (Nujol): ν_{max} 3260 m, 3095 m, 3057 w, 1663 s (CHO), 1561 m, 1263 m, 1211 w, 1193 w, 1164 m, 1094 w, 1069 w, 1033 m, 1023 m, 922 m, 853 m, 832 m, 805 m, 755 s, 741 s, 704 s, 697 s, 633 w, 529 w, 497 s, 486 s, 460 m, 435 w cm^{-1} . ESI+ MS: m/z 414 ($[\text{M} + \text{H}]^+$), 436 ($[\text{M} + \text{Na}]^+$), 452 ($[\text{M} + \text{K}]^+$). HRMS calc. for $\text{C}_{23}\text{H}_{20}^{56}\text{FeNOP}$ (M^+): 413.0632, found 413.0631. Anal. calc. for $\text{C}_{23}\text{H}_{20}\text{FeNOP}$ (413.2): C 66.85, H 4.88, N 3.39%. Found: C 66.83, H 4.91, N 3.37%.

Synthesis of 3B from 3

A $\text{BH}_3\text{-SMe}_2$ solution (6 mL of 2 M in THF, 12 mmol) was added to a solution of compound 3 (4.47 g, 10.0 mmol) in dry THF (30 mL) and the mixture was stirred at room temperature under argon for 90 min. Methanol (0.5 mL) was added to destroy an excess of borane and, after gas evolution had subsided (5 min), the reaction mixture was evaporated with chromatographic silica gel. The preadsorbed product was transferred to a short silica gel column. Elution with toluene–hexane (1 : 1) and evaporation of a single orange band provided the protected phosphine 3B as an orange solid (4.39 g, 95%). Note: Elution with toluene makes the workup faster but the product contains some foul smelling impurities.

^1H NMR (CDCl_3): δ 0.8–1.7 (very br m, 3 H, BH_3), 4.08 (vt, $J' = 1.9$ Hz, 2 H, fc), 4.33 (vt, $J' = 1.9$ Hz, 2 H, fc), 4.46 (vq, $J' = 1.8$ Hz, 2 H, fc), 4.55 (vdt, $J' = 1.8$ Hz, 1.2 Hz, 2 H, fc), 7.39–7.50 (m, 6 H, PPh_2), 7.55–7.61 (m, 4 H, PPh_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 69.44 (CH of fc), 70.49 (d, $^1J_{\text{PC}} = 68$ Hz, $C^{\text{ipso}}\text{-P}$ of fc),



71.67 (CH of fc), 74.87 (d, $J_{PC} = 9$ Hz, CH of fc), 75.47 (d, $J_{PC} = 7$ Hz, CH of fc), 77.72 (C^{ipso} -Br of fc), 128.50 (d, $^2J_{PC} = 10$ Hz, CH^{ortho} of Ph), 130.96 (d, $^1J_{PC} = 60$ Hz, C^{ipso} of Ph), 131.01 (d, $^4J_{PC} = 2$ Hz, CH^{para} of Ph), 132.61 (d, $^3J_{PC} = 9$ Hz, CH^{meta} of Ph). $^{31}P\{^1H\}$ NMR (CDCl₃): δ 16.2 (br d, $J \approx 75$ Hz). IR (Nujol): ν_{max} 3102 w, 3073 w, 3056 w, 2398 s (BH₃), 2373 s (BH₃), 2344 m (BH₃), 1350 m, 1309 m, 1195 m, 1181 m, 1172 s, 1157 w, 1132 w, 1108 s, 1063 s, 1026 s, 1000 m, 978 w, 930 w, 895 w, 872 s, 838 s, 828 w, 818 m, 766 w, 743 s, 703 s, 696 s, 643 s, 622 w, 610 m, 568 w, 527 m, 506 s, 496 s, 476 s, 461 s, 440 m (cm⁻¹). ESI+ MS: m/z 448 ([M - BH₃ + H]⁺), 485 ([M + Na]⁺). Anal. calc. for C₂₂H₂₁BBFeP (462.9): C 57.08, H 4.57. Found: C 57.32, H 4.55%.

Synthesis of 3B from 1,1'-dibromoferrocene

An oven-dried round bottom flask was charged with 1,1'-dibromoferrocene (17.1 g, 50 mmol), flushed with argon and sealed. Dry THF (200 mL) was introduced and the solution was cooled in a dry ice/ethanol bath to ca. -78 °C. *n*-Butyllithium (20 mL of 2.5 M in hexanes, 50 mmol) was introduced and the resultant mixture was stirred at -78 °C for 60 min during which time an orange precipitate formed. Chloro-diphenylphosphine (9.9 mL, 55 mmol) was added, whereupon the precipitate dissolved. The mixture was stirred at -78 °C for 30 min and then at room temperature overnight before it was treated with neat BH₃·SMe₂ (6.3 mL, 65 mmol). Stirring was continued for 1 h and the reaction mixture was diluted by saturated aqueous NaHCO₃ and ethyl acetate (50 mL each). The organic layer was separated, washed with water and brine, dried over MgSO₄ and, finally, evaporated with silica gel. The preadsorbed material was transferred to a silica gel column, which was eluted first with dichloromethane-hexane (1:5) to remove bromo- and 1,1'-dibromoferrocene (ca. 0.45 g). Then, the eluent was gradually changed to pure dichloromethane, which removed a major orange brown band containing 3B. Evaporation and crystallization from hot ethyl acetate-hexane (ca. 1:2 mixture) afforded analytically pure 3B as an orange crystalline solid. Yield: 19.5 g (84%), two crops.

Synthesis of 4B

Compound 3B (4.63 g, 10.0 mmol) was dissolved in dry THF (100 mL) in an oven-dried reaction flask, which was then cooled in dry ice/ethanol bath to ca. -78 °C. *n*-Butyllithium (4.0 mL of 2.5 M in hexanes, 10 mmol) was added and the reaction mixture, which gradually deposited an orange precipitate, was stirred at -78 °C for 30 min. Neat TsN₃ (1.97 g, 10 mmol; see comments below) was introduced and the resulting mixture was stirred at room temperature for 2 h (Note: The solid precipitate dissolved and the color of the reaction mixture changed to red). The reaction was terminated by addition of water (10 mL) and the mixture was transferred to a separatory funnel. The aqueous layer was separated and extracted with diethyl ether (3 × 30 mL). The organic layers were combined, washed with brine, dried over MgSO₄ and evaporated. The obtained brown oil was purified by column chromatography over silica gel using diethyl ether-hexane

(1:1) as the eluent (Note: The crude product should not be exhaustively evaporated to avoid formation of a gummy product, which is poorly soluble in the mobile phase). A single orange band was collected and evaporated under vacuum to afford 4B as an orange microcrystalline solid (3.78 g, 89%). Crystals of 4B used for structure determination were grown by diffusion of hexane into a diethyl ether solution at 4 °C.

Note: TsN₃ must be dried by stirring with CaH₂ freshly prior to the use (preferably overnight). Only if such carefully dried TsN₃ is used, the reaction product is contaminated by less than 5 mol% of FcPPh₂-BH₃. Otherwise, the amount of FcPPh₂-BH₃ and unreacted TsN₃ are higher. Azide 4B does not decompose when stored in the dark at 4 °C for weeks. If appropriate, however, it can be purified by column chromatography as specified above.

1H NMR (CDCl₃): δ 0.8–1.6 (very br m, 3 H, BH₃), 4.00 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.16 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.52 (vq, $J' = 1.8$ Hz, 2 H, fc), 4.65 (vdt, $J' = 1.8$ Hz, 1.1 Hz, 2 H, fc), 7.38–7.50 (m, 6 H, PPh₂), 7.56–7.62 (m, 4 H, PPh₂). $^{13}C\{^1H\}$ NMR (CDCl₃): δ 61.85 (CH of fc), 67.50 (CH of fc), 70.12 (d, $^1J_{PC} = 68$ Hz, C^{ipso} -P of fc), 73.33 (d, $J_{PC} = 8$ Hz, CH of fc), 73.70 (d, $J_{PC} = 10$ Hz, CH of fc), 100.41 (C^{ipso} -N₃ of fc), 128.49 (d, $^2J_{PC} = 10$ Hz, CH^{ortho} of Ph), 131.00 (d, $^4J_{PC} = 2$ Hz, CH^{para} of Ph), 131.01 (d, $^1J_{PC} = 60$ Hz, C^{ipso} of Ph), 132.58 (d, $^3J_{PC} = 9$ Hz, CH^{meta} of Ph). $^{31}P\{^1H\}$ NMR (CDCl₃): δ 16.2 (br d). IR (Nujol): ν_{max} 3077 w, 3054 w, 2382 s (BH₃), 2338 m (BH₃), 2108 s (N₃), 1314 w, 1285 s, 1223 w, 1176 s, 1161 m, 1134 m, 1107 s, 1061 s, 1036 m, 1028 s, 999 w, 920 w, 832 s, 820 m, 760 m, 741 s, 703 s, 692 m, 635 m, 625 m, 611 m, 595 w, 528 s, 489 s, 463 m, 446 w, 409 w cm⁻¹. ESI+ MS: m/z 420 ([M - N₂ + Na]⁺), 448 ([M + Na]⁺). HRMS calc. for C₂₂H₂₁B⁵⁶FeN₃P (M⁺): 425.0916, found 425.0921.

Synthesis of 5B for 4B

A solution of 4B (1.70 g, 4.0 mmol) in dry THF (40 mL) was slowly introduced to a reaction flask containing Li[AlH₄] (0.304 g, 8.0 mmol) while stirring and cooling on ice (gas evolution was noted). After the addition was complete, the cooling bath was removed and the mixture was stirred at room temperature for 3 h before an excess of the reducing agent was destroyed by a careful addition of 10% aqueous NaOH (0.5 mL) and stirring for 10 min. Anhydrous MgSO₄ (1 g) was added and the pasty mixture was transferred to a Celite pad on a glass frit. Washing with dry THF (60 mL in several portions) provided an orange filtrate, which was treated with formic acetic anhydride (1.0 mL, 12.8 mmol) in an argon-flushed vessel. The resultant mixture was stirred at room temperature for 3 h and then diluted with saturated aqueous NaHCO₃ and diethyl ether (50 mL each). The organic layer was separated, washed with water and brine, dried over MgSO₄ and evaporated. The crude product was purified by chromatography on a silica gel column with ethyl acetate-hexane (1:1) as the eluent. The first yellow band containing impurities was discarded and the second (major) one was collected and evaporated to afford 5B as an orange oil, which gradually solidified. Yield: 1.50 g (88%). The product is a mixture of (*E*) and (*Z*) isomers in ca. 2:1 ratio.



^1H NMR (CDCl_3): **major** – δ 0.8–1.7 (very br m, 3 H, BH_3), 3.90 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.29 (vq, $J' = 1.8$ Hz, 2 H, fc), 4.59 (vdt, $J' = 1.8$ Hz, 1.2 Hz, 2 H, fc), 4.69 (vt, $J' = 2.0$ Hz, 2 H, fc), 6.71 (br s, 1 H, NH), 7.39–7.52 (m, 6 H, PPh_2), 7.56–7.63 (m, 4 H, PPh_2), 8.06 (d, $^3J_{\text{HH}} = 1.1$ Hz, 1 H, CHO); **minor** – δ 0.8–1.7 (very br m, 3 H, BH_3), 4.01 (vt, $J' = 1.9$ Hz, 2 H, fc), 4.25 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.41 (vq, $J' = 1.8$ Hz, 2 H, fc), 4.58 (vdt, $J' = 1.9$ Hz, 1.2 Hz, 2 H, fc), 6.74 (br s, 1 H, NH), 7.39–7.52 (m, 6 H, PPh_2), 7.56–7.63 (m, 4 H, PPh_2), 8.14 (d, $^3J_{\text{HH}} = 11.6$ Hz, 1 H, CHO). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): **major** – δ 63.90 (CH of fc), 66.15 (CH of fc), 69.22 (d, $^1J_{\text{PC}} = 69$ Hz, $\text{C}^{\text{ipso}}\text{-P}$ of fc), 73.15 (d, $J_{\text{PC}} = 8$ Hz, CH of fc), 74.22 (d, $J_{\text{PC}} = 10$ Hz, CH of fc), 93.59 ($\text{C}^{\text{ipso}}\text{-N}$ of fc), 128.54 (d, $^2J_{\text{PC}} = 10$ Hz, CH^{ortho} of Ph), 130.95 (d, $^1J_{\text{PC}} = 60$ Hz, C^{ipso} of Ph), 131.04 (d, $^4J_{\text{PC}} = 2$ Hz, CH^{para} of Ph), 132.65 (d, $^3J_{\text{PC}} = 9$ Hz, CH^{meta} of Ph), 159.12 (CHO); **minor** – δ 62.56 (CH of fc), 67.17 (CH of fc), 70.18 (d, $^1J_{\text{PC}} = 68$ Hz, $\text{C}^{\text{ipso}}\text{-P}$ of fc), 73.28 (d, $J_{\text{PC}} = 8$ Hz, CH of fc), 74.21 (d, $J_{\text{PC}} = 10$ Hz, CH of fc), 95.04 ($\text{C}^{\text{ipso}}\text{-N}$ of fc), 128.59 (d, $^2J_{\text{PC}} = 10$ Hz, CH^{ortho} of Ph), 130.82 (d, $^1J_{\text{PC}} = 60$ Hz, C^{ipso} of Ph), 131.19 (d, $^4J_{\text{PC}} = 2$ Hz, CH^{para} of Ph), 132.57 (d, $^3J_{\text{PC}} = 10$ Hz, CH^{meta} of Ph), 162.39 (CHO). $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): 16.1 (br d, both isomers). IR (Nujol): ν_{max} 3275 br m, 3230 br m, 3082 m, 3054 m, 2374 s (BH_3), 2343 m (BH_3), 1687 m, 1663 s (CHO), 1582 m, 1564 s, 1353 w, 1310 w, 1267 w, 1258 w, 1173 s, 1133 w, 1107 m, 1057 s, 1029 s, 999 w, 927 w, 837 w, 830 w, 813 w, 759 w, 744 s, 701 s, 636 m, 623 w, 612 w, 528 m, 500 s, 475 s, 463 w, 447 w cm^{-1} . ESI+ MS: m/z 450 ($[\text{M} + \text{Na}]^+$), 466 ($[\text{M} + \text{K}]^+$). HRMS calc. for $\text{C}_{23}\text{H}_{23}\text{B}^{56}\text{FeNOP}$ (M^+): 427.0960, found 427.0965. Anal. calc. for $\text{C}_{23}\text{H}_{23}\text{BFeNOP}$ (427.0): C 64.68, H 5.43, N 3.28%. Found: C 64.50, H 5.08, N 3.23%.

Synthesis of 5 from 5B

A dry reaction flask was charged with **5B** (1.28 g, 3.0 mmol), 1,4-diazabicyclo[2.2.2]octane (0.449 g, 4.0 mmol) and a stirring bar, flushed with argon and sealed with a rubber septum. Dry toluene (60 mL) was added and the reaction flask was transferred to an oil bath maintained at 80 °C. After stirring for 20 h, the mixture was evaporated under reduced pressure. The residue was dissolved in a minimal amount of THF and purified by column chromatography (silica gel, ethyl acetate–hexane 1 : 1) to afford formamide **5** as an amber brown, slowly crystallizing oil. Yield: 1.17 g (94%).

Synthesis of 2B

A solution of azide **4B** (1.70 g, 4.0 mmol) in dry THF (40 mL) was added to an ice-cooled reaction flask containing $\text{Li}[\text{AlH}_4]$ (0.304 g, 8.0 mmol) and a stirring bar (an effervescence was observed). When the addition was complete, the cooling bath was removed and stirring was continued at room temperature for 3 h. The reaction was terminated by cautious addition of 10% aqueous NaOH (0.5 mL) and, after stirring for another 10 min, anhydrous MgSO_4 (1 g). The obtained semisolid mass was transferred to a Celite pad on a glass frit and washed with dry THF (60 mL in several portions). The orange filtrate was evaporated to afford an orange brown oil, which was purified by chromatography over a short silica gel using ethyl acetate–

hexane (1 : 3) as the eluent, which removed some non-polar impurities. A subsequent elution with ethyl acetate–hexane (1 : 1) provided a deep orange band containing the product. Evaporation of this band afforded amine **2B** as an orange oil, which gradually solidified. Yield: 1.49 g (93%). Note: solid **2B** does not show any signs of decomposition when stored at 4 °C in the dark but readily decomposes in chlorinated solvents.

^1H NMR (CDCl_3): δ 0.8–1.7 (very br m, 3 H, BH_3), 2.51 (br s, 2 H, NH_2), 3.71 (vt, $J' = 1.9$ Hz, 2 H, fc), 3.97 (vt, $J' = 1.9$ Hz, 2 H, fc), 4.25 (vq, 2 H, $J' = 1.8$ Hz, 2 H, fc), 4.48 (v dt, $J' = 1.8$ Hz, 1.1 Hz, 2 H, fc), 7.38–7.48 (m, 6 H, PPh_2), 7.58–7.64 (m, 4 H, PPh_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 59.76 (CH of fc), 64.76 (CH of fc), 68.42 (d, $^1J_{\text{PC}} = 70$ Hz, $\text{C}^{\text{ipso}}\text{-P}$ of fc), 72.48 (d, $^3J_{\text{PC}} = 8$ Hz, CH of fc), 73.78 (d, $^2J_{\text{PC}} = 10$ Hz, CH of fc), 107.37 ($\text{C}^{\text{ipso}}\text{-N}$ of fc), 128.39 (d, $^2J_{\text{PC}} = 10$ Hz, CH^{ortho} of Ph), 130.82 (d, $^4J_{\text{PC}} = 2$ Hz, CH^{para} of Ph), 131.54 (d, $^1J_{\text{PC}} = 59$ Hz, $\text{C}^{\text{ipso}}\text{-P}$ of Ph), 132.62 (d, $^3J_{\text{PC}} = 9$ Hz, CH^{meta} of Ph). $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ 16.4 (br d). IR (Nujol): ν_{max} 3416 m, 3385 m, 3332 m, 3076 m, 3058 w, 2380 s (BH_3), 2255 w (BH_3), 1503 s (NH_2), 1305 w, 1252 w, 1194 w, 1173 s, 1134 m, 1106 s, 1060 s, 1041 m, 1026 s, 998 w, 934 w, 850 m, 817 s, 796 m, 767 m, 744 s, 735 s, 700 s, 689 s, 637 s, 621 m, 613 m, 529 m, 502 s, 471 s, 460 m, 442 w, 417 w, 411 w cm^{-1} . ESI+ MS: m/z 398 ($[\text{M} - \text{H}]^+$). HRMS calc. for $\text{C}_{22}\text{H}_{23}\text{B}^{56}\text{FeNP}$ (M^+): 399.1011, found 399.1018. Anal. calc. for $\text{C}_{22}\text{H}_{23}\text{BFeNP}$ (399.0): C 66.21, H 5.81, N 3.51%. Found: C 66.10, H 5.93, N 3.50%.

Synthesis of amine 2

Compound **2B** (0.399 mg, 1.0 mmol) and 1,4-diazabicyclo[2.2.2]octane (0.224 g, 2.0 mmol) were dissolved in 20 mL of dry toluene in a dry reaction flask under argon. The reaction mixture was stirred at room temperature for 18 h and then evaporated under vacuum. A brown residue was dissolved in a minimal amount of THF and purified by column chromatography (silica gel, ethyl acetate–hexane 1 : 1). Evaporation of a single orange band furnished pure amine **2** as an orange solid (0.362 g, 94%). The compound is stable as a solid but gradually decomposes when dissolved in chlorinated solvents. It can be practically quantitatively recrystallized from hot hexane, which was also used to grow single crystals.

^1H NMR (CDCl_3): δ 2.25 (br s, 2 H, NH_2), 3.78 (vt, $J' = 1.8$ Hz, 2 H, fc), 3.90 (vt, $J' = 1.7$ Hz, 2 H, fc), 4.00 (vt, $J' = 1.8$ Hz, 2 H, fc), 4.33 (vt, $J' = 1.6$ Hz, 2 H, fc), 7.28–7.34 (m, 6 H, PPh_2), 7.36–7.43 (m, 4 H, PPh_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 59.47 (CH of fc), 64.42 (CH of fc), 71.57 (d, $^3J_{\text{PC}} = 3$ Hz, CH of fc), 73.78 (d, $^2J_{\text{PC}} = 15$ Hz, CH of fc), 75.41 (d, $^1J_{\text{PC}} = 4$ Hz, $\text{C}^{\text{ipso}}\text{-P}$ of fc), 106.05 ($\text{C}^{\text{ipso}}\text{-N}$ of fc), 128.13 (d, $^3J_{\text{PC}} = 6$ Hz, CH^{meta} of Ph), 128.46 (CH^{para} of Ph), 133.54 (d, $^2J_{\text{PC}} = 19$ Hz, CH^{ortho} of Ph), 139.47 (d, $^1J_{\text{PC}} = 9$ Hz, C^{ipso} of Ph). $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ –16.7 (s). The NMR data are in accordance with the literature.¹⁹

Synthesis of 5 from 2

Amine **2** (0.100 g, 0.26 mmol) was dissolved in dry THF (5 mL) under an argon atmosphere. Neat formic acetic anhydride (31 μL , 0.39 mmol) was added to the solution and the resulting mixture was stirred at room temperature for 3 h. The reaction



was terminated by addition of saturated aqueous NaHCO_3 and diethyl ether (10 mL each; an effervescence occurred). The aqueous layer was removed and the organic one was washed with water and with brine, and dried over MgSO_4 . Following evaporation under reduced pressure, the crude product was purified by flash column chromatography (silica gel, ethyl acetate–hexane 1 : 1) to afford pure formamide **5** as an orange glassy solid (0.099 g, 93%).

Preparation of 1'-(diphenylphosphino)-1-isocyanoferrrocene (**1**)

An oven-dried flask was charged with a stirring bar, amide **5** (413 mg, 1.0 mmol), (benzotriazol-1-yloxy)tris(dimethylamino) phosphonium hexafluorophosphate (BOP; 884 mg, 2.0 mmol) and dry THF (20 mL) under argon. Neat 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU; 0.38 mL, 2.5 mmol) was added and the mixture was stirred at room temperature for 3 h (the partly soluble BOP dissolved during this time and the mixture slightly darkened). The resulting solution was diluted with diethyl ether (10 mL), washed successively with saturated aqueous NaHCO_3 , water and brine, dried over MgSO_4 and concentrated under vacuum. The residue was taken up in a minimal amount of THF and transferred to a short silica gel column. Elution with diethyl ether–hexane (1 : 1) led to the development of an orange band, which was collected and evaporated under vacuum to afford **1** as an orange microcrystalline solid emitting an unpleasant odor typical of isonitriles. Yield: 281 mg (71%). Notes: A further elution with ethyl acetate recovered some unreacted (impure) amide **5**. The yield of **1** decreased when the reaction was carried out in incompletely dried THF and without rigorous exclusion of air and moisture.

^1H NMR (CDCl_3): δ 4.02 (vt, $J' = 1.9$ Hz, 2 H, fc), 4.22 (vq, $J' = 1.9$ Hz, 2 H, fc), 4.43 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.51 (vt, $J' = 1.9$ Hz, 2 H, fc), 7.31–7.38 (m, 10 H, PPh_2). $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3): δ 67.61 (CH of fc), 68.19 (CH of fc), 73.92 (d, $J_{\text{PC}} = 4$ Hz, CH of fc), 74.96 (d, $J_{\text{PC}} = 14$ Hz, CH of fc), 78.78 (d, $J_{\text{PC}} = 9$ Hz, $\text{C}^{\text{ipso}}\text{-P}$), 128.31 (d, $^3J_{\text{PC}} = 7$ Hz, CH^{meta} of Ph), 128.79 (CH^{para} of Ph), 133.42 (d, $^2J_{\text{PC}} = 20$ Hz, CH^{ortho} of Ph), 138.27 (d, $^1J_{\text{PC}} = 10$ Hz, C^{ipso} of Ph), 164.30 (NC). The resonance due to ferrocene $\text{C}^{\text{ipso}}\text{-NC}$ is probably obscured by the solvent signal. $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ –17.8 (s). ^1H NMR (acetone- d_6): δ 4.10 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.23 (vq, $J' = 1.8$ Hz, 2 H, fc), 4.57 (vt, $J' = 2.1$ Hz, 2 H, fc), 4.59 (vt, $J' = 1.8$ Hz, 2 H, fc), 7.37–7.41 (m, 10 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (acetone- d_6): δ –17.3 (s). IR (Nujol): ν_{max} 3114 w, 3054 w, 2126 s ($\text{N}\equiv\text{C}$), 1231 w, 1193 w, 1160 m, 1089 w, 1028 s, 999 w, 916 w, 836 m, 819 s, 747 s, 698 s, 639 w, 596 w, 531 w, 510 s, 502 m, 488 m, 458 w, 432 w, 414 w cm^{-1} . ESI+ MS: m/z 396 ($[\text{M} + \text{H}]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{FeNP}$ (395.2): C 69.90, H 4.59, N 3.54%. Found: C 69.79, H 4.63, N 3.35%.

Synthesis of complex **6**

A solution of **1** (20.0 mg, 50.6 μmol) in chloroform (1.5 mL) was added to a suspension of freshly precipitated AgCl (7.3 mg, 50.6 μmol) in the same solvent (0.5 mL) and the mixture was stirred in the dark for 24 h (a fine yellow solid sep-

arated). The precipitated product was filtered off, washed with chloroform and pentane, and dried under vacuum. Yield of **6**: 24.2 mg (89%), yellow virtually insoluble solid. Note: An identical product was obtained when the amount of **1** was increased to 2 equiv. with respect to AgCl .

IR (Nujol): ν_{max} 3114 w, 3081 w, 3052 w, 2160 s ($\text{N}\equiv\text{C}$), 2140 m ($\text{N}\equiv\text{C}$), 1311 w, 1233 w, 1194 m, 1168 m, 1100 s, 1070 w, 1036 m, 1025 m, 910 m, 832 m, 824 m, 751 m, 741 vs, 696 vs, 631 w, 613 w, 594 w, 535 m, 515 s, 503 s, 489 s, 470 s, 459 m, 442 w, 427 m cm^{-1} . ESI+ MS: m/z 502 ($[\text{Ag}(\text{1})]^+$), 704 ($[\text{Ag}_2(\text{1})\text{Cl}_2 + \text{Na}]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{AgClFeNP}\cdot 0.2\text{CHCl}_3$ (562.4): C 49.54, H 3.26, N 2.49%. Found: C 49.36, H 3.19, N 2.40%.

Synthesis of complex **7**

A solution of ligand **1** (20.0 mg, 50.6 μmol) in acetone (1.5 mL) was added to a suspension of $\text{Ag}[\text{SbF}_6]$ (17.4 mg, 50.6 μmol) in the same solvent (0.5 mL) and the mixture was stirred in the dark for 1 h (the solid silver(I) salt quickly dissolved). The resulting solution was filtered through a syringe filter (PTFE, 0.45 μm) and the filtrate was crystallized by layering with acetone (1 mL) and then with hexane (10 mL) as a top layer. The crystals, which formed during standing in the dark for approximately one week, were isolated by suction, rinsed with pentane and dried under vacuum. Yield: 31.5 mg (84%), orange needle-like crystals.

^1H NMR (*in situ*, acetone- d_6): δ 4.26 (br s, 2 H, fc), 4.60 (br s, 2 H, fc), 5.01 (vt, $J' = 1.7$ Hz, 2 H, fc), 5.14 (br s, 2 H, fc), 7.54–7.71 (m, 10 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (*in situ*, acetone- d_6): δ 6.8 (pair of d, $^1J_{\text{AgP}} = 709$ (^{109}Ag), 615 (^{107}Ag) Hz). IR (Nujol): ν_{max} IR (Nujol): ν_{max} 3119 w, 3050 w, 2210 m ($\text{N}\equiv\text{C}$), 1695 s (acetone, $\text{C}=\text{O}$), 1316 w, 1230 m, 1198 w, 1173 m, 1101 m, 1059 w, 1034 m, 1001 w, 920 m, 848 w, 821 w, 750 s, 699 s, 664 vs, 653 vs, 542 w, 513 s, 472 m, 445 w, 430 m cm^{-1} . ESI+ MS: m/z 395 (1^+), 502 ($[\text{Ag}(\text{1})]^+$), 646 ($[\text{Ag}_2(\text{1}) + \text{Cl}]^+$), 897 ($[\text{Ag}(\text{1})_2]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{AgF}_6\text{FeNPSb}\cdot (\text{CH}_3)_2\text{CO}$ (796.9): C 39.18, H 3.04, N 1.76%; found C 39.49, H 3.07, N 1.72%.

Preparation of complex **8**

A solution of compound **1** (20.0 mg, 50.6 μmol) in acetone (1.5 mL) was added to $\text{Ag}[\text{SbF}_6]$ in the same solvent (8.7 mg, 25.3 μmol in 0.5 mL). After stirring in the dark for 90 min, the reaction mixture was filtered through a PTFE syringe filter (0.45 μm pore size) and the filtrate was carefully layered with acetone (1 mL) and, finally, with hexane (10 mL) as the top layer. Crystallization by liquid-phase diffusion (in the dark) provided crystals which were isolated by suction, washed with pentane and dried under vacuum. Yield of **8**: 20.1 mg (70%), orange needles.

^1H NMR (*in situ*, acetone- d_6): δ 4.20 (br s, 2 H, fc), 4.71 (br s, 2 H, fc), 5.02 (br s, 2 H, fc), 5.10 (br s, 2 H, fc), 7.31–7.37 (m, 4 H, PPh_2), 7.42–7.50 (m, 6 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (*in situ*, acetone- d_6): δ –1.2 (br d, $^1J_{\text{AgP}} = 207$ Hz). IR (Nujol): ν_{max} 3054 w, 2167 m ($\text{N}\equiv\text{C}$), 1309 m, 1261 w, 1233 w, 1195 m, 1166 s, 1098 s, 1030 s, 999 m, 917 s, 890 w, 829 s, 743 s, 696 s, 658 vs, 584 w, 519 m, 508 s, 494 s, 476 m, 464 m, 426 w cm^{-1} . ESI+



MS: m/z 395 (1^+), 897 ($[\text{Ag}(1)_2]^+$), 1688 ($[\text{Ag}(1)_4]^+$). Anal. calc. for $\text{C}_{46}\text{H}_{36}\text{AgF}_6\text{Fe}_2\text{N}_2\text{P}_2\text{Sb}$ (1134.0): C 48.72, H 3.20, N 2.47%. Found: C 48.34, H 3.12, N 2.37%.

Preparation of complex 9

$[\text{AuCl}(\text{tht})]$ (16.2 mg, 50.6 μmol) and compound **1** (20.0 mg, 50.6 μmol) were reacted in chloroform (1 + 2 mL) as described above and the reaction mixture was filtered into pentane (30 mL). The precipitated solid was filtered off, washed with pentane and dried under vacuum. Yield of **9**: 27.4 mg (87%), yellow solid. Crystals used for structure determination were grown from diethyl ether, which slightly dissolves the product, layered with hexane.

^1H NMR (acetone- d_6): δ 4.31 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.63 (d vt, $^3J_{\text{PH}} = 3.1$ Hz, $J' = 2.0$ Hz, 2 H, fc), 4.73 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.88 (d vt, $^4J_{\text{PH}} = 1.1$ Hz, $J' = 2.0$ Hz, 2 H, fc), 7.57–7.74 (m, 10 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (acetone- d_6): δ 29.0 (s). IR (Nujol): ν_{max} 3087 w, 3052 w, 2220 w, 2125 s ($\text{N}\equiv\text{C}$), 1686 w, 1560 w, 1310 m, 1231 w, 1198 w, 1176 m, 1102 s, 1070 w, 1028 s, 998 m, 919 m, 893 w, 834 m, 823 m, 758 s, 743 s, 694 s, 637 w, 558 s, 532 s, 514 s, 494 m, 482 s, 453 w, 436 w cm^{-1} . ESI+ MS: m/z 592 ($[\text{M} - \text{Cl}]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{AuClFeNP}$ (627.6): C 44.01, H 2.89, N 2.23%. Found: C 43.76, H 2.96, N 1.99%.

Preparation of complex 10

A solution of $[\text{AuCl}(\text{tht})]$ (32.5 mg, 101 μmol) in chloroform (2 mL) was mixed with a solution of ligand **1** in the same solvent (20.0 mg, 50.6 μmol in 2 mL). After stirring for 1 h, the mixture was concentrated under reduced pressure (to approximately half the volume) and filtered through a syringe filter (0.45 μm PTFE). The filtrate was layered with chloroform (2 mL) and then with hexane (10 mL) as a top layer and set aside for crystallization. The separated material was filtered off, washed with pentane and dried under vacuum to afford **10** as an orange crystalline solid. Yield: 40.6 mg (94%). Single crystals used for structure determination were obtained when the filtered solution was layered with ethyl acetate and with hexane as a top layer. The crystalline product is practically insoluble.

IR (Nujol): ν_{max} 3128 w, 3082 w, 2223 s ($\text{N}\equiv\text{C}$), 1311 m, 1242 w, 1198 m, 1178 m, 1173 m, 1105 s, 1070 w, 1054 w, 1036 m, 1030 m, 997 w, 925 m, 841 m, 822 s, 756 s, 747 s, 705 s, 696 s, 635 w, 559 s, 534 s, 514 s, 495 s, 479 s, 470 m cm^{-1} . ESI+ MS: m/z 824 ($[\text{M} - \text{Cl}]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{Au}_2\text{Cl}_2\text{FeNP}$ (860.0): C 32.12, H 2.11, N 1.63%. Found: C 32.08, H 2.11, N 1.48%.

Synthesis of complex 11a

A chloroform solution of isonitrile **1** (20.0 mg, 50.6 μmol in 2 mL) was added to $[\text{Au}(\text{tht})_2][\text{SbF}_6]$ (30.8 mg, 50.6 μmol) dissolved in the same solvent (1 mL). The resulting mixture was stirred for 60 min and then evaporated. The residue was dissolved in acetone (2 mL) and the solution was crystallized by addition of hexane as the top layer (10 mL). The orange crystals, which separated during several days, were filtered off, washed with pentane and dried under vacuum. Yield of **11a**: 36.2 mg (79%).

^1H NMR (acetone- d_6): δ 4.40 (vt, $J' = 2.1$ Hz, 2 H, fc), 4.74 (d vt, $J = 3.1$, 1.9 Hz, 2 H, fc), 5.16 (d vt, $J = 1.2$, 1.9 Hz, 2 H, fc), 5.42 (vt, $J' = 2.1$ Hz, 2 H, fc), 7.65–7.71 (m, 4 H, PPh_2), 7.72–7.57 (m, 2 H, PPh_2), 7.80–7.87 (m, 4 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (acetone- d_6): δ 33.0 (s). IR (Nujol): ν_{max} 3118 w, 3056 w, 2231 s ($\text{N}\equiv\text{C}$), 1704 s, 1314 w, 1226 m, 1200 w, 1184 m, 1176 m, 1105 s, 1060 w, 1034 m, 999 w, 921 w, 848 m, 834 m, 823 m, 755 s, 749 s, 715 w, 696 m, 660 vs, 553 m, 531 m, 519 s, 493 m, 476 m, 435 w cm^{-1} . ESI+ MS: m/z 592 ($[\text{Au}(1)]^+$). Anal. calc. for $\text{C}_{23}\text{H}_{18}\text{AuF}_6\text{FeNPSb}$ (827.9): C 33.36, H 2.19, N 1.69%. Found: C 33.18, H 2.20, N 1.55%.

Preparation of complex 11b

A solution of ligand **1** (20.0 mg, 50.6 μmol) in dichloromethane (2 mL) was added to $[\text{AuCl}(\text{tht})]$ (16.2 mg, 50.6 μmol) dissolved in the same solvent (1 mL). After the mixture was stirred for 30 min, solid AgNTf_2 (19.6 mg, 50.6 μmol) was added to the *in situ* formed **9**, causing immediate separation of a white precipitate (AgCl). The mixture was stirred for another 30 min in the dark and then filtered through a syringe filter (0.45 μm PTFE). The filtrate was evaporated to give an orange residue, which was taken up with chloroform (1 mL). The solution was layered with chloroform (1 mL) and then hexane (10 mL) as a top layer, and allowed to crystallize by liquid-phase diffusion over several days. The orange crystals that formed were isolated by suction and dried under vacuum. Yield: 40.8 mg (92%).

^1H NMR (CDCl_3): δ 4.01 (vt, $J' = 2.0$ Hz, 2 H, fc), 4.76 (v dt, $J' = 3.2$ Hz, 2.0 Hz, 2 H, fc), 4.97 (v dt, $J' = 1.9$ Hz, 1.2 Hz, 2 H, fc), 5.12 (vt, $J' = 2.0$ Hz, 2 H, fc), 7.52–7.60 (m, 6 H, PPh_2), 7.63–7.70 (m, 4 H, PPh_2). $^{31}\text{P}\{^1\text{H}\}$ NMR (CDCl_3): δ 32.1 (s). ^{19}F NMR (CDCl_3): δ -78.6 (s). IR (Nujol): ν_{max} 3113 m, 3056 m, 2233 m ($\text{N}\equiv\text{C}$), 1351 s, 1227 s, 1193 s, 1134 s, 1104 m, 1056 s, 999 w, 922 w, 828 w, 788 m, 748 m, 694 m, 652 m, 615 m, 600 m, 570 m, 555 w, 530 m, 517 s, 476 m, 435 w cm^{-1} . ESI+ MS: m/z 592 ($[\text{Au}(1)]^+$). Anal. calc. for $\text{C}_{25}\text{H}_{18}\text{AuF}_6\text{FeN}_2\text{O}_4\text{PS}_2$ (871.9): C 34.42, H 2.08, N 3.21%. Found: C 34.33, H 2.29, N 3.18%.

Catalytic experiments

A reaction flask was charged with alcohol **12** (1.0 mmol; *ca.* 90 : 10 mixture of (*Z*)- and (*E*)-isomers, the (*E*)-isomer does not cyclize), 1,2-dichloroethane (1.0 mmol; internal standard) and CDCl_3 (2 or 4 mL). The catalyst was added and the resulting mixture was stirred in the dark for a given reaction time and then analyzed by ^1H NMR spectroscopy to determine the conversion.

X-ray crystallography

The diffraction data (completeness $\geq 99\%$, $\theta_{\text{max}} = 27.5^\circ$) were collected at 150(2) K using either Nonius Kappa CCD diffractometer with a Bruker Apex-II image plate detector or a Bruker D8 VENTURE Kappa Duo PHOTON100 instrument with a $\text{I}\mu\text{S}$ micro-focus X-ray tube, both equipped with a Cryostream Cooler (Oxford Cryosystems). Mo $\text{K}\alpha$ radiation was used throughout.



The structures were solved by direct methods (SHELXS-97) and refined by a full-matrix least squares procedure based on F^2 using SHELXL-97.⁵¹ The non-hydrogen atoms were refined with anisotropic thermal motion parameters. The NH hydrogen atoms in the structure of **2** and **5** were located on the difference Fourier maps, and their positions were freely refined with $U_{\text{iso}}(\text{H})$ set to 1.2 $U_{\text{eq}}(\text{N})$. All other hydrogen atoms were included in their calculated positions and refined using the “riding model”.

Selected crystallographic data, data collection and structure refinement parameters are presented in the ESI (Table S1†). All geometric calculations were performed with a recent version of the PLATON program,⁵² which was also used to prepare the structural diagrams.

DFT computations

DFT calculations were performed using the PBE0 density functional⁵³ combined with the SDD (Fe, Au) and cc-pVDZ (C, H, N, and P) basis sets. The solvent effects (CHCl_3) have been approximated by the polarized continuum model (PCM)⁵⁴ as implemented in the Gaussian 09 program package.⁵⁵ The geometry optimizations were started from the coordinates determined by X-ray diffraction analysis. The reported energies refer to the Gibbs free energies at 298 K. Computed energies and coordinates are presented in the ESI (Table S2 and mol 2 files†).

Acknowledgements

The research reported in this paper received support from the Czech Science Foundation (project no. 17-02495S).

References

- (a) *Isocyanide Chemistry: Applications in Synthesis and Material Science*, ed. V. G. Nenajdenko, Wiley-VCH, Weinheim, 2012; (b) A. Dömling, *Chem. Rev.*, 2006, **106**, 17; (c) G. Qiu, Q. Ding and J. Wu, *Chem. Soc. Rev.*, 2013, **42**, 5257; (d) V. P. Boyarskiy, N. A. Bokach, K. V. Luzyanin and V. Y. Kukushkin, *Chem. Rev.*, 2015, **115**, 2698.
- (a) Y. Yamamoto, *Coord. Chem. Rev.*, 1980, **32**, 193; (b) E. Singleton and H. E. Oosthuizen, *Adv. Organomet. Chem.*, 1983, **22**, 209.
- (a) *Ferrocenes: Ligands, Materials and Biomolecules*, ed. P. Štěpnička, Wiley, Chichester, 2008; (b) *Ferrocenes: Homogeneous Catalysis, Organic Synthesis, Materials Science*, ed. A. Togni and T. Hayashi, VCH, Weinheim, 1995.
- T. El-Shihi, F. Sigmüller, R. Herrmann, M. F. N. N. Carvalho and A. J. L. Pombeiro, *J. Organomet. Chem.*, 1987, **335**, 239.
- For an alternative synthesis of **A**, see: G. R. Knox, P. L. Pauson, D. Willison, E. Solčániová and Š. Toma, *Organometallics*, 1990, **9**, 301.
- Note that aminoferrocene, an obvious synthetic precursor to isocyanoferrocene, was prepared already in 1955 (ref. 6a), while cyanoferrocene was reported only two years later (ref. 6b–d): (a) F. S. Arimoto and A. C. Haven Jr., *J. Am. Chem. Soc.*, 1955, **77**, 6295; (b) G. D. Broadhead, J. M. Osgerby and P. L. Pauson, *Chem. Ind.*, 1957, 209, (*Chem. Abstr.*, 1957, 62321); (c) G. D. Broadhead, J. M. Osgerby and P. L. Pauson, *J. Chem. Soc.*, 1958, 650; (d) P. J. Graham, R. V. Lindsey, G. W. Parshall, M. L. Peterson and G. M. Whitman, *J. Am. Chem. Soc.*, 1957, **79**, 3416.
- (a) M. V. Barybin, T. C. Holovics, S. F. Deplazes, G. H. Lushington, D. R. Powell and M. Toriyama, *J. Am. Chem. Soc.*, 2002, **124**, 13668; (b) T. C. Holovics, S. F. Deplazes, M. Toriyama, D. R. Powell, G. H. Lushington and M. V. Barybin, *Organometallics*, 2004, **23**, 2927; (c) P. J. Shapiro, R. Zehnder, D. M. Foo, P. Perrotin, P. H. M. Budzelaar, S. Leitch and B. Twamley, *Organometallics*, 2006, **25**, 719; (d) V. N. Nemykin, A. A. Purchel, A. D. Spaeth and M. V. Barybin, *Inorg. Chem.*, 2013, **52**, 11004; (e) V. N. Nemykin, S. V. Dudkin, M. Fathi-Rasekh, A. D. Spaeth, H. M. Rhoda, R. V. Belosludov and M. V. Barybin, *Inorg. Chem.*, 2015, **54**, 10711.
- D. van Leusen and B. Hessen, *Organometallics*, 2001, **20**, 224.
- (a) U. Siemeling, D. Rother, C. Bruhn, H. Fink, T. Weidner, F. Träger, A. Rothenberger, D. Fenske, A. Priebe, J. Maurer and R. Winter, *J. Am. Chem. Soc.*, 2005, **127**, 1102; (b) T. Weidner, N. Ballav, M. Zharnikov, A. Priebe, N. J. Long, J. Maurer, R. Winter, A. Rothenberger, D. Fenske, D. Rother, C. Bruhn, H. Fink and U. Siemeling, *Chem. – Eur. J.*, 2008, **14**, 4346; (c) U. Siemeling, L. R. R. Klapp and C. Bruhn, *Z. Anorg. Allg. Chem.*, 2010, **636**, 539.
- A few more isocyanides bearing ferrocenyl substituents at the organic backbone were mentioned in ref. 4.
- D. M. McGinnis, S. F. Deplazes and M. V. Barybin, *J. Organomet. Chem.*, 2011, **696**, 3939.
- (a) K. Škoch, I. Císařová and P. Štěpnička, *Inorg. Chem.*, 2014, **53**, 568; (b) K. Škoch, I. Císařová and P. Štěpnička, *Chem. – Eur. J.*, 2015, **21**, 15998; (c) K. Škoch, F. Uhlík, I. Císařová and P. Štěpnička, *Dalton Trans.*, 2016, **45**, 10655.
- (a) K.-S. Gan and T. S. A. Hor, in *Ferrocenes: Homogeneous Catalysis, Organic Synthesis, Materials Science*, ed. A. Togni and T. Hayashi, Wiley-VCH, Weinheim, 1995, ch. 1, p. 3 ff; (b) S. W. Chien and T. S. A. Hor, in *Ferrocenes: Ligands, Materials and Biomolecules*, ed. P. Štěpnička, Wiley, Chichester, 2008, ch. 2, p. 33 ff; (c) G. Bandoli and A. Dolmella, *Coord. Chem. Rev.*, 2000, **209**, 161; (d) D. J. Young, S. W. Chien and T. S. A. Hor, *Dalton Trans.*, 2012, **41**, 12655.
- (a) R. B. King and A. Efraty, *J. Am. Chem. Soc.*, 1971, **93**, 564; (b) R. B. King and A. Efraty, *J. Chem. Soc., Chem. Commun.*, 1974, 1371.
- (a) C.-Y. Liu, D.-Y. Chen, M.-C. Cheng, S.-M. Peng and S.-T. Liu, *Organometallics*, 1995, **14**, 1983; (b) I. Yu,



- C. J. Wallis, B. O. Patrick, P. L. Diaconescu and P. Mehrkhodavandi, *Organometallics*, 2010, **29**, 6065.
- 16 L. Zhang, W. Yu, C. Liu, Y. Xu, Z. Duan and F. Mathey, *Organometallics*, 2015, **34**, 5697.
- 17 *Isonitrile Chemistry*, ed. I. Ugi, Academic Press, New York, 1971.
- 18 P. Štěpnička, in *Ferrocenes: Ligands, Materials and Biomolecules*, ed. P. Štěpnička, Wiley, Chichester, 2008, ch. 5, p. 177 ff.
- 19 I. R. Butler and S. C. Quayle, *J. Organomet. Chem.*, 1998, **552**, 63.
- 20 The reaction mixture resulting after addition of O-benzylhydroxylamine to *in situ* generated Ph₂PfcLi contained mostly (diphenylphosphino)ferrocene, unreacted amination reagent and only traces of amine 2.
- 21 Similar problems were noted before: (a) M. Ritte, C. Bruhn and U. Siemeling, *Z. Naturforsch.*, 2014, **69b**, 906. See also: (b) S. Dey, J. W. Quail and J. Müller, *Organometallics*, 2015, **34**, 3039.
- 22 A. Leonidova, T. Joshi, D. Nipkow, A. Frei, J.-E. Penner, S. Konatschnig, M. Patra and G. Gasser, *Organometallics*, 2013, **32**, 2037 and references cited therein.
- 23 We have noted that desulfuration of ferrocene P-sulfides by commercial Raney nickel can be accompanied by removal of the entire phosphine substituent already before: P. Štěpnička, B. Schneiderová, J. Schulz and I. Císařová, *Organometallics*, 2013, **32**, 5754.
- 24 I. Ugi, U. Fetzter, U. Eholzer, H. Knupfer and K. Offermann, *Angew. Chem., Int. Ed.*, 1965, **4**, 472.
- 25 The dehydration of oximes to nitriles with BOP/DBU was reported in: M. K. Singh and M. K. Lakshman, *J. Org. Chem.*, 2009, **74**, 3079.
- 26 M. K. Das and S. Roy, *Synth. React. Inorg. Met.-Org. Chem.*, 1986, **16**, 67.
- 27 K. BongSoo, J. M. Beebe, Y. Jun, X.-Y. Zhu and C. D. Frisbie, *J. Am. Chem. Soc.*, 2006, **128**, 4970.
- 28 P. Walla, V. B. Arion and U. H. Brinker, *J. Org. Chem.*, 2006, **71**, 3274.
- 29 A. Shafir, M. P. Power, G. D. Whitener and J. Arnold, *Organometallics*, 2000, **19**, 3978.
- 30 M. J. Verschoor-Kirss, O. Hendricks, C. M. Verschoor, R. Conry and R. U. Kirss, *Inorg. Chim. Acta*, 2016, **450**, 30.
- 31 Z.-G. Fang, T. S. A. Hor, Y.-S. Wen, L.-K. Liu and T. C. W. Mak, *Polyhedron*, 1995, **14**, 2403.
- 32 K. J. Donaghy, P. J. Carroll and L. G. Sneddon, *Inorg. Chem.*, 1997, **36**, 547.
- 33 B. Wrackmeyer, H. E. Maisel, O. L. Tok, W. Milius and M. Herberhold, *Z. Anorg. Allg. Chem.*, 2004, **630**, 2106.
- 34 K. Heinze and M. Schlenker, *Eur. J. Inorg. Chem.*, 2004, 2974.
- 35 A. Cassel, *Acta Crystallogr., Sect B: Struct. Crystallogr. Cryst. Chem.*, 1979, **35**, 174.
- 36 K. Yang, S. G. Bott and M. G. Richmond, *J. Chem. Crystallogr.*, 1995, **25**, 263.
- 37 M. M. Olmstead, M. Sheffrin and F. Jiang, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2004, **60**, m1142.
- 38 B. Cordero, V. Gómez, A. E. Platero-Prats, M. Revés, J. Echeverría, E. Cremades, F. Barragán and S. Alvarez, *Dalton Trans.*, 2008, 2832.
- 39 A search in Cambridge Crystallographic Database (version 5.38 of November 2016 with updates from November 2016 and February 2017) for Ag–C≡N–C fragments resulted in 63 hits and 129 Ag–C distances. Average value: 2.14(8) Å, median: 2.11 Å.
- 40 Spectral Database for Organic Compounds (SDBS): IR spectrum of neat acetone; SDBS No.: 319; <http://riodb01.ibase.aist.go.jp/sdbs/> (accessed April 21, 2017); $\nu_{\text{C=O}}$: 1715 cm⁻¹.
- 41 The broadening can be tentatively ascribed to a slow molecular dynamics of the sterically crowded complex cation in **8**. The ¹H NMR signals also display broad resonances.
- 42 O. Crespo, M. C. Gimeno, P. G. Jones and A. Laguna, *Acta Crystallogr., Sect. C: Struct. Chem.*, 2000, **56**, 1433.
- 43 (a) H. Schmidbaur, *Chem. Soc. Rev.*, 1995, 391; (b) H. Schmidbaur, *Gold Bull.*, 2000, **33**, 3; (c) H. Schmidbaur and A. Schier, *Chem. Soc. Rev.*, 2012, **41**, 370.
- 44 C. Hansch, A. Leo and R. W. Taft, *Chem. Rev.*, 1991, **91**, 165.
- 45 Our observations are in agreement with the theoretical computations on the series of cations $[(\eta^5\text{-C}_5\text{H}_5)_3\text{ML}]^+$, where L is MeC≡N or MeN≡C, and M = Ti, Zr or Hf, revealing that the isonitrile consistently forms stronger coordination bonds to the metal centre: H. Jacobsen, H. Berke, T. Brackemeyer, T. Eisenblätter, G. Erker, R. Fröhlich, O. Meyer and K. Bergander, *Helv. Chim. Acta*, 1998, **81**, 1692.
- 46 P. Štěpnička, J. Schulz and I. Císařová, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2009, **65**, m1210.
- 47 D. P. Day, T. Dann, R. J. Blagg and G. G. Wildgoose, *J. Organomet. Chem.*, 2014, **770**, 29.
- 48 L. I. Krimen, J. Savage and P. Yates, *Org. Synth. Coll.*, 1988, **6**, 8, DOI: 10.15227/orgsyn.050.0001 (accessed online).
- 49 R. Uson, A. Laguna and M. Laguna, *Inorg. Synth.*, 1989, **26**, 85.
- 50 F. Barrière and W. E. Geiger, *J. Am. Chem. Soc.*, 2006, **128**, 3980.
- 51 G. M. Sheldrick, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 2008, **64**, 112.
- 52 A. L. Spek, *Acta Crystallogr., Sect. D: Biol. Crystallogr.*, 2009, **65**, 148.
- 53 C. Adamo and V. Barone, *J. Chem. Phys.*, 1999, **110**, 6158.
- 54 (a) J. Tomasi, B. Mennucci and R. Cammi, *Chem. Rev.*, 2005, **105**, 2999; (b) G. Scalmani and M. J. Frisch, *J. Chem. Phys.*, 2010, **132**, 114110.
- 55 M. J. Frisch, *et al.*, *Gaussian 09, Revision A.02*, Gaussian, Inc., Wallingford, CT, 2009.

