The relationship between cis–trans isomerism and anticancer activity has been mainly addressed for square-planar metal complexes, in particular, for platinum(II), e.g., cis- and trans-[PCl₂(NH₃)₂], and a number of related compounds, of which, however, only cis-counterparts are in clinical use today. For octahedral metal complexes, this effect of geometrical isomerism on anticancer activity has not been investigated systematically, mainly because the relevant isomers are still unavailable. An example of such an octahedral complex is trans-[RuCl₄(Hind)₂]⁻, which is in clinical trials now as its indazolium (KP1019) or sodium salt (NKP1339), but the corresponding cis-isomers remain inaccessible. We report the synthesis of Na[cis-OsIVCl₄(kN2–1H–ind)]₂(Na[1]) suggesting a route to the cis-isomer of NKP1339. The procedure involves heating (H₂-ind)OCl₄Cl₄(kN1–2H–ind)] in a high boiling point organic solvent resulting in an Anderson rearrangement with the formation of cis-[OsIVCl₄(kN2–1H–ind)]₂ [1] in high yield. The transformation is accompanied by an indazole coordination mode switch from kN1 to kN2 and stabilization of the 1H-indazole tautomer. Fully reversible spectroelectrochemical reduction of [1] in acetonitrile at 0.46 V vs. NHE is accompanied by a change in electronic absorption bands indicating the formation of cis-[OsIVCl₄(kN2–1H–ind)]₂ [1] or [1]⁻. Chemical reduction of [1] in methanol with NaBH₄ followed by addition of nBu₄NCl afforded the osmium(iii) complex nBu₄N[cis-OsIVCl₄(kN1–2H–ind)]₂ (nBu₄N [1]). A metathesis reaction of nBu₄N[1] with an ion exchange resin led to the isolation of the water-soluble salt Na[1]. The X-ray diffraction crystal structure of [1]·Me₂CO was determined and compared with that of trans-[OsIVCl₄(kN2–1H–ind)]₂·Me₂SO (2·Me₂SO, also prepared in this work. EPR spectroscopy was performed on the Os(iii) complexes and the results were analyzed by ligand-field and quantum chemical theories. We furthermore assayed effects of [1] and Na[1] on cell viability and proliferation in comparison with trans-[OsIVCl₄(kN1–2H–ind)]₂ [3] and cisplatin and found a strong reduction of cell viability at concentrations between 30 and 300 μM in different cancer cell lines (HT29, H446, 4T1 and HEK293). HT-29 cells are less sensitive to cisplatin than 4T1 cells, but more sensitive to [1] and Na[1], as shown by decreased proliferation and viability as well as an increased late apoptotic/necrotic cell population.
Introduction

Indazole derivatives are characterized by a variety of applications as anti-inflammatory, antimicrobial, antihypertensive, antiprotozoal, antiobesity, and antifungal agents.\textsuperscript{1-3} They have been found to be active as inhibitors of nitric oxide synthase (NOS),\textsuperscript{4} bacterial gyrase,\textsuperscript{5} and various hydrolases,\textsuperscript{6} and various kinases.\textsuperscript{7-9} The ruthenium(II) complexes $\text{H}_2\text{ind}[\text{trans-RuCl}_4(\text{kN2-1H-indazole})_2]$ (KP1019) and Na[trans-RuCl$_4$(kN2-1H-indazole)$_2$] [NKP1339] are in clinical trials as potential anticancer agents.\textsuperscript{10} Indazole, usually referred to as 1H-indazole, exists in three tautomeric forms (Chart 1). 1H-Indazole is the dominant tautomer in the gas phase and aqueous solution,\textsuperscript{11,12} as well as in metal complexes, however, some substituted organic 2H-indazoles are known\textsuperscript{13} as well as rare, inorganic examples such as an organoruthenium(II) complex with 2-methyl-indazole,\textsuperscript{14} and the osmium(IV) complexes $\text{trans}[\text{OsCl}_4(\text{kN1-2H-ind})]$ and $\text{H}_2\text{ind}$. The synthesis of the latter complex along with $\text{trans}[\text{OsCl}_4(\text{kN1-2H-ind})]$ was achieved by coordination of an indazole-pendant azamacro cyclic scorpion to Cu\textsuperscript{II} or Zn\textsuperscript{II}.\textsuperscript{30}

Herein, we describe the synthesis of $\text{cis}[\text{OsCl}_4\text{(kN2-1H-ind})]$ and its one electron reduced species ($\text{nBu}_2\text{N}$)$[\text{I}]$ and Na[1] (Chart 2). These complexes are spectroscopically characterized, in particular by electron paramagnetic resonance (EPR), a technique which has been shown to be very useful in unraveling the solution structure of Ru\textsuperscript{III} anti-cancer drug candidates,\textsuperscript{31} as well as other Ru\textsuperscript{III} complexes.\textsuperscript{32} EPR has been less widely applied to Os\textsuperscript{III} complexes of any type because of the paucity of such compounds and frequently their inherently more challenging experimental EPR properties.\textsuperscript{33,34} In addition, the antiproliferative activity of the two octahedral cis-isomers [1] and Na[1], along with the X-ray diffraction structures of cis- and trans-$\text{OsCl}_4\text{(kN2-1H-ind})$ [1] and [2], respectively are reported. This discovery offers the pathway to cis-isomers of investigational anticancer drugs KP1019 and NKP1339 and a subsequent opportunity to investigate structure–activity relationships for octahedral cis/trans-$\text{RuCl}_4\text{(kN2-1H-ind})$ isomers.

Experimental section

Materials

OsO$_4$ (99.8%) was purchased from Johnson-Matthey, while $\text{nBu}_2\text{NBr}$, NaBH$_4$, 1H-indazole, KBr and Dowex Marathon C-resin Na$^+$ form were from Sigma-Aldrich. All chemicals were used without further purification.

Synthesis of osmium complexes

The starting compounds (H$_2$ind)$[\text{OsCl}_4]$ and (H$_2$ind)$[\text{OsCl}_3(k\text{N1-2H-ind})]$ were prepared according to the literature protocols.\textsuperscript{16,15}$\text{cis}[\text{OsCl}_4(k\text{N2-1H-ind})]_2\text{Me}_2\text{CO}$ ([1]$\text{Me}_2\text{CO}$). (H$_2$ind)$[\text{OsCl}_3(k\text{N1-2H-ind})]$ (160 mg, 0.26 mmol) was stirred in 1,1,2,2-tetrachloroethane (30 ml) at 128 °C for 10 h. Removal of the solvent at reduced pressure, washing the residue with diethyl ether and drying in vacuo afforded a black solid of [1]. Yield: 140 mg, 93%. Elemental analysis was performed on a sample crystallized from Me$_2$CO. Calcd for C$_{124}H_{142}Cl$_6$N$_{26}$Os$\cdot$Me$_2$CO ($M_r = 626.39$), %: C, 32.60; H, 2.90; N,
16.49 (s, 2H, H3) ppm. 13C{1H} NMR (Me2SO- d6, 125.82 MHz): δ = 77.76 (CH3), 162.76 (CH2), 125.82 (CH3). UV (CD3OD, 500.26 MHz): λmax, nm (ε1 cm−1): 370 (1665), 405 (1635), 543 (364), 618 (521), 724 (82). UV (ethanol), λmax, nm (ε1 cm−1): 228 (2663), 268 (2665), 364 (1533), 407 (1600), 562 (376), 729 (130). UV (acetonitrile), λmax, nm (ε1 cm−1): 243 (1228), 271 (1413), 286 sh, 365 (740), 409 (853), 559 (220), 725 (82). UV (acetone), λmax, nm (ε1 cm−1): 286 (2686), 369 (1727), 400 (1636), 507 (683), 618 (521), 724 (137). UV (ethanol), λmax, nm (ε1 cm−1): 370 (1665), 402 (1663), 562 (376), 608 sh, 725 (96). UV (THF), λmax, nm (ε1 cm−1): 296 (2423), 374 (1677), 400 (1797), 543 (364), 726 (83). 1H NMR (Me2SO-d6, 500.26 MHz): δ = −9.42 (s, 2H, H3), 1.35 (t, J = 7.96 Hz, 2H, H4), 4.66 (d, J = 8.98 Hz, 2H, H5), 8.04 (t, J = 8.06 Hz, 2H, H2), 10.87 (d, 2H, H3, J = 8.23 Hz), 16.49 (s, 2H, H3) ppm. 13C{1H} NMR (Me2SO-d6, 125.82 MHz): δ = 77.76 (CH3), 162.76 (CH2), 125.82 (CH3). Suitable crystals for the X-ray diffraction study were grown by slow evaporation of a solution of cis-[OsVCl3(H2Ind)] in Me2CO. For atom numbering and assignment of NMR resonances see Fig. S1 in ESI.‡

(nBu4N)[cis-OsVCl3H2Ind]2 (1). To a solution of cis-[OsVCl3H2Ind]2 (30 mg, 0.04 mmol) in methanol (1 ml) under nitrogen NaBH4 (3 mg, 0.079 mmol) was added. After stirring for 1 h nBu4NCl (16 mg) in water (1 ml) was added. The orange solution produced a yellow-brown precipitate, which was filtered, washed with cold water and diethyl ether and dried in vacuo to give a light-brown solid. Yield: 24 mg, 61%. Analytical data for nBu4N[1]: ESI-MS in MeOH (negative): m/z 567 [OsIVCl3(H2Ind)2]−, 332 [OsIVCl4]−. MIR, cm−1: 656, 752, 785, 833, 959, 1002, 1081, 1243, 1355, 1440, 1482, 1506, 1625, 2827, 2960 and 3285. 1H NMR (CD2OD, 500.26 MHz): δ = −22.16 (br. s, 2H), −7.06 (br. s, 2H), −0.63 (br. s, 2H), 0.32 (br. s, 2H), 1.12 (t, 2H, J = 7.5 Hz), 1.34 (br. s, 2H), 1.55 (sxt, 8H, J = 7.4 Hz), 1.96 (qui, 8H, J = 6.9 Hz), 2.33 (br. s, 2H), 3.55 (t, 8H, J = 7.8 Hz) ppm. Na2-cis-OsIIICl3H2Ind-H2O (Na[1]·H2O). Compound nBu4N[1] (50 mg, 0.06 mmol) was dissolved in a 1:1 mixture of methanol and water (50 ml) under a nitrogen atmosphere. Ion exchanger Dowex Marathon C Na+-form was added to the solution after soaking for 12 h in water. After stirring the suspension for 5 min the ion exchange resin was separated by filtration and the solution lyophilized to give a light-brown solid. Yield: 23 mg, 62%. Calcd for C14H12Cl4N4NaOs·H2O: %: C, 27.60; H, 2.32; N, 9.20. Found: %: C, 28.08; H, 2.15; N, 8.32. ESI-MS in MeOH (negative): m/z 567 [OsIIICl4(H2Ind)2]−. MIR, cm−1: 652, 749, 834, 964, 1002, 1080, 1241, 1356, 1438, 1471, 1510, 1627, 2817 and 3313, 3510. 1H NMR (MeOH-d4, 500.26 MHz): δ = −22.16 (br. s, 2H), −7.06 (br. s, 2H), −0.63 (br. s, 2H), 0.32 (br. s, 2H), 1.34 (br. s, 2H), 3.55 (t, 8H, J = 7.8 Hz) ppm.

Trans-OsCl3[(n-2H-ind)2]·2Me2SO (2)·2Me2SO. (H2Ind)OsVCl3H2Ind]·2Me2SO (2)·2Me2SO. (H2Ind)OsIIICl3H2Ind]·2Me2SO. (H2Ind)OsVCl3H2Ind]·2Me2SO.
Physical measurements

UV-vis spectra were recorded on a Jasco V-630 UV−vis spectrophotometer using samples dissolved in Me₂CO, MeCN, DMF, Me₂SO, EtOH, and THF. MIR spectra in the region 4000−600 cm⁻¹ were recorded with a Thermo Nicolet NEXUS 470 FT-IR spectrometer (Thermo Fisher Scientific, USA), in KBr pellets. Electrospray ionization mass spectrometry was carried out with a Thermo Scientific LCQ Fleet ion trap instrument (Thermo Scientific, Waltham, MA, USA) by using methanol as a solvent. High-resolution mass spectrometry (HRMS) was performed on a LTQ Orbitrap Velos instrument (Thermo Scientific, Waltham, MA, USA). Expected and measured isotope distributions were compared. The cyclic voltammetric studies were performed in a homemade microelectrochemical cell using a platinum wire as working and counter electrodes, and a silver wire as a pseudo-reference electrode. Ferrocene served as the internal potential standard. The redox potentials in Me₂SO solutions were recalculated vs. Standard Hydrogen Electrode (NHE) using the known redox potential of ferroocene (0.64 V) vs. NHE. A Heka PG310USB (Lambrecht, Germany) potentiostat with a PotMaster 2.73 software package was used in cyclic voltammetric and spectroelectrochemical studies. In situ spectroelectrochemical measurements were performed on a spectrometer Avantes, Model AvaSpec-2048x14-USBB2. Halogen and deuterium lamps were used as light sources (Avantes, Model AvaLight-DH-S-BAL) under an argon atmosphere in a spectroelectrochemical cell kit (AKSTCKIT3) with the Pt-microstructured honeycomb working electrode, purchased from Pine Research Instrumentation. The cell was positioned in the CUV-UV Cuvette Holder (Ocean Optics) connected to the diode-array UV−vis-NIR spectrophotometer by optical fibers. UV−vis-NIR spectra were processed using the AvaSoft 7.7 software package. The 1H and 13C NMR spectra were recorded at 500 and 125 MHz on a Bruker Avance III 500 MHz (Ultrashield Magnet) instrument by using MeOH-d₄ and Me₂SO-d₆ as solvents. The solvent residual peak for 1H and 13C was used as an internal reference. EPR spectra were recorded using an EMXplus X-band EPR spectrometer (Bruker, Germany) equipped with a high-sensitivity resonator ER 4119 HS and an ER 4141 VT variable temperature unit. The EPR spectra were analyzed and simulated with the Easyspin toolbox. Ligand-field parameters for LS d⁵ systems were calculated using a fitting program originally provided by B. R. McGarvey³³ and modified by J. Telser.³²

Crystallographic structure determination

X-ray diffraction measurements of [1]Me₂CO and [2]Me₂SO were performed on a Bruker X8 APEX-II CCD diffractometer at 199 and 100 K, respectively. The data were processed using SAINT software. The structures were solved by direct methods and refined by full-matrix least-squares techniques. Non-hydrogen atoms were refined with anisotropic displacement parameters. H atoms were placed at calculated positions and refined as riding atoms in the subsequent least squares model refinements. The isotropic thermal parameters were estimated to be 1.2 times the values of the equivalent isotropic thermal parameters of the non-hydrogen atoms to which hydrogen atoms are bonded. The following computer programs and equipment were used: solution SHELXS-2014 and refinement, SHELXL-2014; molecular diagrams, ORTEP; computer, Pentium IV. Crystal data, data collection parameters, and structure refinement details are given in Table 1. CCDC 1544885 and 1544886 contain the CIFs for this manuscript.

Computational details

The Gaussian09 package has been used for the geometry optimizations of the anion [1]− (total charge minus one and doublet spin state). The unrestricted B3LYP/d2-TZVP level of theory has been employed and the core electrons of osmium were treated via the SDD pseudo potential. Vibrational analysis showed that the optimized geometry of 2[1]− is free of imaginary (negative) frequencies.

The g-tensor values have been theoretically determined using two different approaches. The first attempt employed the ORCA software package at the DFT (with either BLYP or B3LYP functional) and complete active space self-consistent field (CASSCF) levels of theory. Here the scalar ZORA Hamiltonian and uncontracted double zeta (UDZ) basis sets have been used together with inclusion of solvent effects via the conductor-like screening model (COSMO) of Me₂SO. The spin-orbit coupling (SOC) effects have been included as a perturbation (with ORCA input settings: grid 7; scotype 3; socflags 1,3,3,1). The CASSCF calculation used five electrons in five orbitals and has been run as state specific (accounting formally for 10 roots). SOC treatment in CASSCF has been performed under default settings of the relativity module (dosoc

<table>
<thead>
<tr>
<th>Table 1 Crystal data and details of data collection for [1]Me₂CO and [2]Me₂SO</th>
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<tr>
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<tr>
<td>b [Å]</td>
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<tr>
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<tr>
<td>R₁</td>
</tr>
<tr>
<td>wR₂</td>
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<tr>
<td>GOFC</td>
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</tbody>
</table>

a R₁ = Σ|F₁| - |F₁||ΣF₁|, \ ŜwR₂ = Σ{[w(F₁^2) - F₁^2]}|Σ[w(F₁^2)]|^{1/2}, c GOFC = Σ{[w(F₁^2) - F₁^2]}(n - p) p^{1/2}, where n is the number of reflections and p is the total number of parameters refined.
true; g-tensor true). In the second approach, the ReSpect program package has been employed utilizing the four-component relativistic Dirac–Kohn–Sham (DKS) methodology. In this way, the SOC effects are included variationally and thus account for higher than linear effects, which can be especially crucial in calculations of g-tensor values of compounds containing heavy elements such as osmium. The non-collinear Kramers unrestricted formalism was used as the best method currently available for description of spin polarization at the relativistic DFT level of theory. The recent implementation of solvent models in the ReSpect program allowed us to take into account the effects of the environment (Me₂SO) in the framework of the COSMO model.

Cell lines and culture conditions
H446 small cell lung carcinoma (American Type Culture Collection, Manassas, VA, USA) and HEK293 embryonic kidney cells (a kind gift from the laboratory of Dr Ross Levine, MSKCC, New York, NY, USA) were maintained in minimal essential medium. 4T1 breast carcinoma (American Type Culture Collection, Manassas, VA, USA) cells were grown in RPMIM. HT-29 colorectal adenocarcinoma cells (a kind gift from the laboratory of Dr Vladimir Ponomarev, MSKCC, New York, NY, USA) were grown in McCoy’s medium. All media were supplemented with 10% fetal calf serum, penicillin (100 IU) and streptomycin (100 µg mL⁻¹). All the cell lines were grown in 150 cm² culture flasks (Corning Incorporated, New York, NY, USA) under standard sterile cell culture conditions at 37 °C under a humidified atmosphere with 5% CO₂ and 95% air.

Cytotoxicity in cancer cell lines
Cell viability was tested by the Alamar Blue assay (Invitrogen, Carlsbad, CA, USA). Stock solutions of the osmium compounds in Me₂SO were prepared at a concentration of 30 mM. Cisplatin was dissolved in 0.9% aqueous solution of NaCl. Na₃[OsIVCl₄(Hind)₂][Na[1]] stock solution was prepared under an argon atmosphere to prevent direct contact with oxygen. Serial dilutions of the osmium compounds and cisplatin were prepared in the corresponding nutrient medium, ranging from 0.003 µM to 300 µM. All dilutions contained a final Me₂SO concentration of 1%. For the Alamar Blue assay, cells were plated into 96-well plates (black, clear-bottom, Corning Incorporated) at a density of 4000 cells per well and allowed to adhere for 24 h. Then, the cells were incubated with osmium compounds or vehicle control (medium with Me₂SO) for 48 h. The Alamar Blue reagent was used as described by the manufacturer. The Alamar Blue reagent was diluted 1:10 in full culture medium to achieve the working solution. The medium was removed from 96-well plates and 100 µL Alamar Blue reagent was added. The plates were incubated at 37 °C for 3 h under protection from direct light. Fluorescence was read (bottom-read mode) using an excitation wavelength of 570 nm and an emission wavelength of 585 nm on a SpectraMax M5 microplate reader (Molecular Devices, Sunnyvale, CA, USA). The average reading values were analyzed by subtracting the background readings (wells with Alamar Blue but without cells) and compared against the untreated controls of each plate (defined as 100% viability) to determine the percentage of living cells. The experiment was carried out in six parallels of each compound.

Measurement of antiproliferative activity and apoptosis induction
We evaluated the potential of complexes [1] and Na[1] on cell viability, proliferation and apoptosis induction in comparison with trans-[OsVCl₄[kN1-2H-ind]₃] [3] cisplatin and untreated cells using an Annexin V Apoptosis Detection Kit (BioLegend). Thus, we plated 200 000 HT-29 or 4T1 cells per well in 6-well plates. After 24 h medium was removed and the cells were incubated with fresh medium or a 200 µM solution of [1], complex Na[1], [3] or cisplatin in medium for 48 h. After trypsinization, the supernatant and cells were collected in Falcon tubes. 1 ml of the suspension was analyzed in an automated cell counter (Beckman Coulter Counter) for cell count and viability. The remaining cells were stained with a FITC Annexin V (AV) and Propidium Iodide (PI) Kit as described by the manufacturer. The cells were washed twice with cell staining buffer (BioLegend), resuspended in Annexin binding buffer and transferred to a flow tube with a cell strainer cap (0.2–1 × 10⁶ cells in 100 µl). Then, 5 µl Annexin V were added per sample, followed by 10 µl PI. The tubes were incubated for 15 min in the dark and then analyzed on a LSR Fortessa III Flow Cytometer (BD Biosciences). Compensation was carried out using single stained controls. The cells that stained positive for Annexin V were considered early apoptotic, Annexin V and PI positive cells were considered late apoptotic or necrotic and the cells positive only for PI were considered necrotic.

Results and discussion
Synthesis
The complex cis-[OsIVCl₄[kN2-1H-ind]₃] [1] was prepared by Anderson rearrangement in 1,1,2,2-tetrachloroethane at 128 °C starting from (H₂ind)₃[OsIVCl₆]. This transformation is unprecedented in this family of compounds as it implies both a change of coordination mode from kN1 to kN2 of one indazole ligand and substitution of the chlorido ligand by an outer sphere indazole. The second indazole binds via N2 to osmium(v) with the stabilization of the 1H-indazole tautomeric form (see Scheme 1).

Starting from (H₂ind)₃[OsIVCl₆] and by exploring the Anderson-rearrangement reaction, different types of osmium(v) complexes, namely, two coordination isomers, (H₂ind)₃[OsIVCl₄(kN1-2H-ind)] and (H₂ind)[OsVCl₄(kN2-1H-ind)], were synthesized. These complexes were studied for their antiproliferative activity in vitro and in vivo (Scheme 2).

Solid state transformation of (H₂ind)₃[OsIVCl₆] at 150 °C yielded trans-[OsIVCl₄[kN2-1H-ind]₃] [2] as a side product and trans-[OsVCl₄[kN1-2H-ind]₃] [3] as the main species (Scheme 3). The minor product 2 could be identified by X-ray diffraction (vide infra).
While trans-configured osmium analogues of KP1019 in oxidation states +3 and +4, namely $\mathbf{3}^-$ and $\mathbf{3}$, have been reported in the literature, the corresponding cis-compounds were not known. Complex $\mathbf{3}$ was further used for the synthesis of trans-$\text{[Os}^\text{IVCl}_4(\text{Hazole})_2\text{]}$ by chemical or electrochemical reduction, building a link between the number of halido ligands, $n$, and Os formal oxidation state: trans-$\text{[Os}^\text{IVCl}_4(\text{Hazole})_2\text{]}$ with $n \geq 4$ and oxidation state +4, and species with $n < 4$ and oxidation state +3, as in mer-$\text{[Os}^\text{IIICl}_3(\text{Hazole})_3\text{]}$.\(^{35}\) trans-$\text{[Os}^\text{IIICl}_2(\text{Hazole})_2\text{]}$Cl, and cis-$\text{[Os}^\text{IIICl}_2(\text{Hazole})_2\text{]}$Cl.$^{58}$

This suggests an associative mechanism for the solid state reaction and a dissociative mechanism for the homogeneous reaction.$^{59}$ The reduction of osmium(n) to osmium(III) with NaBH$_4$ was performed as described previously for $\mathbf{3}^-$ (ref. 29) and $\mathbf{1}^-$ was isolated as its tetra-$n$-butylammonium salt ($n$Bu$_4$N$\mathbf{1}$). Cation metathesis employing an ion exchange resin$^{60}$ afforded the sodium salt Na$\mathbf{1}$. Compounds $\mathbf{1}$, $n$Bu$_4$N$\mathbf{1}$, and Na$\mathbf{1}$ were obtained in excellent to good yields of 93, 61, and 62%, respectively.

Crystal structures

The results of X-ray diffraction studies of complexes $\mathbf{1}$-Me$_2$CO and $\mathbf{2}$-2Me$_2$SO are shown in Fig. 1. Crystallographic data are summarized in Table 1, while selected bond lengths and angles are quoted in Table 2. To the best of our knowledge these are the first tetrachlorido-bis(1H-indazole)osmium(iv) complexes to be crystallographically characterized. $\mathbf{1}$-Me$_2$CO and $\mathbf{2}$-2Me$_2$SO crystallized in the monoclinic centrosymmetric space groups $\text{P}2_1/c$ and $\text{P}2_1/n$, respectively. The first contains one molecule of cis-$\text{[Os}^\text{IVCl}_4(1\text{H-indazole})_2\text{]}$ and one molecule of Me$_2$CO, while the second contains one molecule of trans-$\text{[Os}^\text{IVCl}_4(1\text{H-indazole})_2\text{]}$ and two molecules of Me$_2$SO in the asymmetric unit. The osmium atom in $\mathbf{1}$ displays a distorted octahedral coordination geometry with two chlorido ligands and two 1H-indazole ligands in equatorial positions, and two chlorido ligands bound axially. In $\mathbf{2}$ osmium(IV) adopts a compressed octahedral coordination geometry with four chlorido ligands occupying the equatorial plane and two 1H-indazole ligands in the axial positions. In a few cases reported in the literature, 1H-indazole was found to be deprotonated, acting as a bridging ligand in polynuclear metal complexes, or even more rarely it operates as a monodentate indazolato ligand, such as when coordinated via N2 to platinum and deprotonated at N1 in the square planar complex $\text{[PtCl(\text{N-indazolato})(PPh}_3)_2\text{]}$.\(^{62}$

A comparison of the bond lengths in the coordination sphere of osmium in $\mathbf{1}$, $\mathbf{2}$, and $\mathbf{3}$ is shown in Table 2. Inspection of geometric parameters in complexes $\mathbf{1}$ and $\mathbf{2}$ with the same coordination mode of indazole shows that the Os–N2 and Os–N11 bonds are significantly longer (0.02 Å; 5.5 × σ) in the cis- than in the trans-isomer (see Fig. 1 and Table 2). The same trend is clear when the Os–N bonds in cis-complex $\mathbf{1}$ are compared to those in trans-isomer $\mathbf{3}$ with a
dissimilar coordination mode of indazole (see Chart 2). This is presumably due to a stronger trans-influence of the chlorido ligand compared to nitrogen atom N2 or N1 (using the numbering for indazole in Chart 1) of indazole. Interestingly, an insignificant (<3 × σ) shortening of axial Os–N bonds in trans-complex [3] compared to that in trans-complex [2] (coordination modes are different) results in a significant (14 × σ) lengthening of equatorial Os–Cl bonds by ca. 0.03 Å (see Table 2) in [3] vs. those in [2].

Table 2 Selected bond lengths (Å) and angles (°) in the coordination polyhedron of osmium(IV) in [1], [2] and [3]

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<tr>
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<tr>
<th>Atom1–Atom2–Atom3</th>
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<th>[3]</th>
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</thead>
<tbody>
<tr>
<td>N1–Os–Cl1</td>
<td></td>
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<td>91.73(15)</td>
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<td>N1–Os–Cl2</td>
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<td>177.52(10)</td>
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<td>89.82(3)</td>
<td>90.22(6)</td>
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<td>C11–Os–N11</td>
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<tr>
<td>C11–Os–Cl4</td>
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<td>90.36(3)</td>
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<tr>
<td>N11–Os–Cl4</td>
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<tr>
<td>Cl2–Os–Cl4</td>
<td>91.91(2)</td>
<td>179.72(3)</td>
<td></td>
</tr>
</tbody>
</table>

Both trans-indazole ligands in [2] are almost coplanar in contrast to what was found in [(Ph3P)trans-RuCl4(1H-indazole)], wherein the 1H-indazole ligands were significantly twisted. A salient structural feature is the stabilization of indazole in its usual 1H-indazole form in [1] and [2] (Fig. 1) and coordination to osmium via nitrogen atom N2. This mode of coordination is typical of other metal ions, e.g. ruthenium, but rare for osmium(IV). The only previous example for osmium, namely H2ind[OsIVCl5(1H-ind)], was reported by some of us recently.67

Intermolecular hydrogen bonding interaction between complexes [1] and [2] and co-crystallized solvent molecules are shown in Fig. S2 and S3 in the ESI.

NMR spectra

All compounds were characterized by 1H NMR spectra showing signals due to the coordinated indazole. In the spectrum of nBu4N[1], signals of the Bu4N+ cation could also be assigned. The paramagnetic osmium(IV) in [1] caused some line broadening and spread of chemical shifts from −9.42 to +16.49 ppm (Fig. S4). It should, however, be noted that despite the paramagnetism of osmium(IV) in [1], the NMR lines remain sharp enough such that coupling constants could be easily determined. Qualitatively, the fact that the nuclear paramagnetic resonance spectrum of [1] is readily interpretable is unsurpris-
ing for OsIV (low-spin (LS) 5d⁴, S = 1) because this ion is expected to have very fast electron spin relaxation⁶⁴ due to the substantial orbital angular momentum of its ground state and the extremely large spin–orbit coupling (SOC) constant of OsIV (free-ion OsIV (Os(v) in spectroscopic notation) has ƞd = 4133.7 cm⁻¹,⁶⁵ ions with slow electronic relaxation, such as high-spin (HS) 3d⁵ S = 5/2 (e.g., FeIII), in contrast, typically give unusable NMR spectra.⁶⁴ The number of carbon resonances in ¹³C NMR spectra of [1] is in agreement with C₂ molecular symmetry of the complex. The broadening of resonances due to paramagnetism of osmium(III) (low-spin 5d⁵ S = 1/2) in nBu₄N [1] and Na[1] is more obvious compared to that for [1] and signal splitting for indazole protons is no more discernible. All signals are significantly upfield shifted being in the δ range from 3.55 to 22.16 ppm.

### UV–vis spectra and solvatochromism

The UV–vis spectra of [1] have been measured in a number of organic solvents (Fig. 2). Depending on the solvent used, a red shift of the absorption maximum from 543 (THF) to 618 (Me₂SO) nm was observed. This is in line with the relative permittivity values (εr) of solvents used [tetrahydrofuran (7.5),⁶⁶ acetone (21.30),⁶⁷ ethanol (25.02),⁶⁷ acetonitrile (38),⁶⁸ dimethyl formamide (37.31),⁶⁹ dimethyl sulfoxide (47.2)].⁷⁰ This can be explained by the better stabilization of (polar) excited states by likewise polar solvents.

### Electrochemical behavior

The cyclic voltammogram of [1] in a solution of 0.2 M nBu₄NPF₆ in acetonitrile starting from the anodic part using a platinum wire as the working electrode and a scan rate of 0.2 V s⁻¹ is shown in Fig. 3.

The first reduction step is reversible (E₁/₂ = 0.69 V vs. NHE or +0.01 V vs. Fe⁺/Fe⁺ at all scan rates used (from 10 mV s⁻¹ to 500 mV s⁻¹), and is followed by the less reversible one at a more negative potential (E₁/₂ = −0.78 V vs. NHE). Upon in situ reduction at a scan rate of 10 mV s⁻¹ in the region of the first reversible cathodic peak in acetonitrile, the UV–vis absorption bands at 408, 370 and 272 nm decreased, and, simultaneously, a new dominating band at 318 nm appeared (Fig. 4a). Fully reversible spectroelectrochemical behavior confirmed the high stability of cathodically generated cis-[OsIII(C₂Cl₄)(κN₂-1H-ind)₂](+ [1]) species. A very similar reversible cyclic voltammetric and spectroelectrochemical response was observed for [1] also in Me₂SO at the first reduction peak even at a very low scan rate of 5 mV s⁻¹ as illustrated in Fig. S5. Here the UV–vis absorption bands of [1] in Me₂SO at 415, 370 and 280 nm decreased, and simultaneously, a new dominating band at 325 nm was observed. The spectroelectrochemical data indicated that the isolation of the osmium(III) species in the solid state after reduction of the parent osmium(IV) complex is feasible. The chemically prepared and isolated osmium(III) complex nBu₄N [1] can be easily oxidized in Me₂SO back to [1]⁺ at anodic potential E = +1.0 V vs. NHE as shown in Fig. S6.⁶⁹

Fig. 4c shows the absorption spectra of the initial cis-isomer [1] (cyan trace) and one-electron reduced species, namely ([1]⁻) (dark yellow trace). The UV–vis spectrum of the osmium(III) species [1]⁻ generated electrochemically exhibits an absorption maximum at a slightly higher wavelength (318 nm in acetonitrile and 325 nm in Me₂SO) compared to its analogue trans-[OsIII(C₂Cl₄)(κN₁-2H-ind)₂]⁻ ([3] in Me₂SO) reported previously²⁹ with λmax at 301 nm. Of note is a more negative reduction potential with a shift of −0.20 mV reported recently for [3]²⁹ with E₁/₂ = +0.49 V vs. NHE or E₁/₂ = −0.19 V vs. Fe⁺/Fe⁺ for the first reduction step. A significant difference in the E₁/₂ values for [1] and [3] indicates a stronger effect of cis/trans isomerism on reduction potentials than that of the coordination mode of indazole in the compared species.

As mentioned previously, the OsIII/II couple is much less reversible than the OsIV/III couple and takes place at much more negative potentials. Upon repetitive redox cycling of sample 1 going up to the voltage of the second reduction peak,
irreversible changes were observed in cyclic voltammetric experiments in the thin layer cell at a scan rate of 10 mV s$^{-1}$ both for Me$_2$SO and acetonitrile solutions as illustrated in Fig. S7.$^{†}$ New redox active species are formed during the reduction of $[1]^{-}$ in the region of the second electron transfer indicating a lower stability of the generated osmium(II) complex. The changes in UV-vis spectra observed during cyclic voltammetry of $[1]^{-}$ going to the region of the second reduction in the Me$_2$SO solution (scan rate 5 mV s$^{-1}$) are shown in Fig. 5. A new optical band arises at ca. 440 nm in the region of the second electron transfer (see Fig. 5c and Fig. S8$^{†}$) which again disappears upon reoxidation. The irreversible changes in the initial optical spectra of $[1]$ were observed confirming the irreversible chemical changes of the corresponding osmium(ni) species $[1]^{2-}$ formed upon cathodic reduction of $[1]$ at the second reduction step. Interestingly, a newly formed product can be reversibly reduced at $E_{1/2} = +0.27$ V vs. NHE (see blue trace in Fig. S7$^{†}$). Cathodically induced metal dechlorination was previously mentioned for the trans-$\text{[Ru}^{III}\text{Cl}_4\text{(Hind)$_n$]^{-}}$ complex$^{21}$ upon reduction of ruthenium(ni) to ruthenium(ni) and for the trans-$\text{Os}^{III}$ analogues.$^{20}$ Similarly, our findings indicate that a chlorido ligand is replaced by a solvent molecule upon reduction of osmium(ni) species and a new redox active osmium(ni) complex containing a coordinated solvent molecule is formed at the second reduction potential of $[1]^{0}$ (respectively at the first reduction peak for $[1]^{+}$).

EPR of osmium(ni) complexes

As mentioned previously, the NMR resonances of the osmium(ni) complexes $\text{nBu_4N}[1]$ and Na[1] were considerably broadened due to the rapid nuclear relaxation induced by the paramagnetism of these compounds, which could be expected from the LS $d^5$ electron configuration of the central metal.$^{31b,c}$ EPR is a valuable technique to unravel the underlying mechanisms of action of paramagnetic anticancer drugs and, in particular, those of ruthenium(ni).$^{31}$ The X-band EPR spectra of frozen solutions of $\text{nBu}_4\text{N}[1]$ (in Me$_2$SO at 94 K) and Na[1] (in H$_2$O/ethylene glycol 5:1 v/v at 12 K) are shown in Fig. 6 and
confirm the paramagnetic nature of the complexes, while a Q-Band EPR spectrum of Na[1] and a powder EPR spectrum of nBu4N[1] are shown in Fig. S10 and S11, respectively.

Both spectra show rather broad lines with $\Delta B_{pp} > 150$ mT and not a very well resolved spectral pattern, characterized by a rhombic $g$-tensor. The $g$ values obtained by the simulation of the spectra (Table 3, $g_{\text{max}} \approx 2.50$, $g_{\text{mid}} \approx 1.45$ and $g_{\text{min}} \approx 1.35$) are very similar and can be considered virtually identical for both complexes, if we take into account the uncertainty of $g$-value determination induced by the large linewidth. Thus the $\text{cis-}[\text{Os}^{III}\text{Cl}_4(\text{enN-1H-ind})]^{-}$ complex anion geometry in frozen solution as well as the electronic configuration are not influenced by the different counterions that govern the solubility of nBu4N[1] and Na[1] in an organic or aqueous environment, respectively. An unambiguous conclusion about the spin state of a complex cannot be made based solely on a single frequency cw-EPR experiment. However, the interactions between the unpaired electrons in the high spin states usually result in a facilitated electron spin relaxation, and the consequent line broadening renders their EPR spectra often undetectable above liquid helium temperatures. The fact that we were able to record the spectrum of nBu4N[1] at 94 K and a comparable linewidth was observed in the spectrum of Na[1] at 12 K, indicate the absence of such electron–electron interactions. Moreover, the 35 GHz EPR spectrum of Na[1] (Fig. S10†) confirms the $g_{\text{max}}$ value from X-band. The osmium(II) ion is thus in the usual low spin $\text{d}^5$ ($S = \frac{5}{2}$) configuration, in both nBu4N[1] and Na[1], as reported for the vast majority of OsIII complexes in the literature.73

### Ligand field theory discussion of $[\text{I}^{-}]$ $g$-tensor

The $g$ values of LS nd5 complexes have been extensively analyzed over the years because of the importance of this electronic configuration in heme proteins and other bioinorganic systems.74 The model that we shall use for a simple ligand field theory (LFT) analysis of the $g$ tensor of $[\text{I}^{-}]$ is that developed by McGarvey.33,34,35,36 The details need not be discussed here, but the model treats this $t_2g_e$ configuration (strong-field notation) by ignoring any effects of excited states involving the $e_g$ orbitals and using the “hole” in the $t_{2g}$ orbitals as a basis set as follows:75

\[
|xy, \pm 1/2\rangle = dx_x^\pm dx_y^\pm dx_z^\pm dy_y^\pm dz_z^\pm; \\
|zx, \pm 1/2\rangle = idz_y^\pm dx_y^\pm dy_y^\pm dx_z^\pm; \\
|yz, \pm 1/2\rangle = idz_x^\pm dx_z^\pm dy_y^\pm dz_z^\pm.
\]

The (axial) splitting between the $d_{xy}$ and $d_{xz,yz}$ orbitals is given by $\Delta$ and the (rhombic) splitting between the $d_{xz}$ and $d_{yz}$ orbitals is given by $\gamma$. This analysis has been used previously for RuIII complexes of general type trans-[RuIII(\text{NH}_3)_3(L)(L')]$^{+}$,76 which could be considered as cationic analogues to the equatorially tetrachlorido coordinated anionic Ru/OsIII complexes. To the best of our knowledge, no EPR spectra of complexes of general formula cis-[OsIIIX_4(L)]$^{-}$ have been reported. Clarke and co-workers have reported EPR spectra for cis-[RuIII(\text{NH}_3)_3(\text{Him})_2]\text{Br}_2.77 It is possible to use analytical equations given by McGarvey to extract an internally consistent set of LFT parameters from the experimental $g$ values.78 This has been done here and the results are given in Table 4. In each case, there are two solutions, but usually, one can be discarded as it requires unreasonable parameters, e.g., the orbital reduction factor, $k$, is much too far from 1.0, or the splitting among the $t_{2g}$ orbitals is much too large. Note that in this model it is not possible to provide the absolute energy splitting among the $t_{2g}$ orbitals, only their energies normalized by the spin–orbit coupling (SOC) constant, $\zeta$. This parameter has recently been accurately determined for free-ion OsIII (g) (Os IV in spectroscopic notation) in its 5d5 electronic configuration: $\zeta_\text{d} = 3738.6$ cm$^{-1}$.79 The reduction of this parameter due to covalency in these complexes is unknown, but values of $\zeta = 2530$ and 2220 cm$^{-1}$ in trans-[OsCl_4(O)\text{py}]+ (py = pyridine) are found respectively for Cs$^+$ and tetraethylammonium counterions.34 The fact that the SOC constant is so sensitive even to counterions indicates the difficulty of quantifying this interaction, but an estimate of $\zeta \approx 2500$ cm$^{-1}$ seems reasonable in

### Table 3 The $g$-values of OsIII complexes nBu4N[1], Na[1] obtained from simulation of their EPR spectra

<table>
<thead>
<tr>
<th>$g$-Factor$^a$</th>
<th>nBu4N[1]$^b$</th>
<th>Na[1]$^c$</th>
</tr>
</thead>
<tbody>
<tr>
<td>$g_{\text{max}}$</td>
<td>2.44 ± 0.05</td>
<td>2.50 ± 0.05</td>
</tr>
<tr>
<td>$g_{\text{mid}}$</td>
<td>1.44 ± 0.05</td>
<td>1.45 ± 0.05</td>
</tr>
<tr>
<td>$g_{\text{min}}$</td>
<td>1.31 ± 0.05</td>
<td>1.37 ± 0.05</td>
</tr>
</tbody>
</table>

$^a$ While EPR is unable to provide the absolute sign of the $g$ factor the ligand field theory analysis indicates that the $g$ values are all negative for the investigated complexes (see Table 4). The assignment of $g$-tensor component directions is not made here; only ordering by magnitude (see Table 4 for such assignments).$^b$ In Me$_2$SO at 94 K.$^c$ In H$_2$O/ethylene glycol 5 : 1 v/v at 12 K.
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Table 4  The g-values of Ru\textsuperscript{III} and Os\textsuperscript{III} complexes relevant to the present work with results from LFT analysis

| Complex* | $g_x$, $g_y$, $g_z$ | $\Delta/\zeta$ | $V/\zeta$ | $V/\Delta$ | $k$
<table>
<thead>
<tr>
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</tr>
</thead>
<tbody>
<tr>
<td>eBu[N] \textsuperscript{1}</td>
<td>−1.31, −1.44, −2.44</td>
<td>−0.492</td>
<td>−0.043</td>
<td>0.088</td>
<td>0.8204</td>
</tr>
<tr>
<td>Na[1]</td>
<td>−1.37, −1.45, −2.50</td>
<td>−0.498</td>
<td>−0.026</td>
<td>0.053</td>
<td>0.8541</td>
</tr>
<tr>
<td>trans-[OsCl\textsubscript{4}(CO)[py]]\textsuperscript{−b}</td>
<td>−2.55, −2.55, −1.72</td>
<td>+0.228</td>
<td>0.0</td>
<td>0</td>
<td>1.220</td>
</tr>
<tr>
<td>trans-[OsBr\textsubscript{4}(CO)[py]]\textsuperscript{−b}</td>
<td>−2.50, −2.50, −1.80</td>
<td>+0.236</td>
<td>0.0</td>
<td>0</td>
<td>1.211</td>
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<tr>
<td>trans-[OsBr\textsubscript{4}(CO)]\textsuperscript{−}</td>
<td>−2.46, −2.46, −1.83</td>
<td>+0.221</td>
<td>0.0</td>
<td>0</td>
<td>1.192</td>
</tr>
<tr>
<td>trans-[RuCl\textsubscript{4}(Him)]\textsuperscript{−e}</td>
<td>−3.12, −2.44, −1.2</td>
<td>+0.536</td>
<td>+0.316</td>
<td>0.589</td>
<td>1.256</td>
</tr>
<tr>
<td>trans-</td>
<td>−2.35, −2.35, −1.87</td>
<td>+0.166</td>
<td>0.0</td>
<td>0</td>
<td>1.148</td>
</tr>
<tr>
<td>[RuCl\textsubscript{4}(Me\textsubscript{2}SO)]\textsuperscript{−} d</td>
<td>−3.04, −2.20, (−0.15); −3.04, −2.20, (+0.15); −2.20, −3.04; (+0.9),</td>
<td>+0.9635; +1.136; +0.6043; +0.7123; 0.627; 0.627; 1.002; 0.9667;</td>
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<td></td>
</tr>
<tr>
<td>(NH\textsubscript{3})\textsubscript{4}(Him)\textsubscript{2} \textsuperscript{−e}</td>
<td>−2.20, −3.04</td>
<td>−0.6392; −1.847; −0.4016; −1.160; 0.628; 0.628; 1.114; 0.9270;</td>
<td></td>
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<tr>
<td>cis-[Ru(NH\textsubscript{3})\textsubscript{4}(Him)]\textsuperscript{−e}</td>
<td>−2.88, −2.14, (−0.65); −2.88, −2.14, (+0.65); −2.88, −2.14, −2.14, (−0.15); −2.88, −2.14, (+0.15)</td>
<td>+0.6905; +1.463; +0.4265; +0.9017; 0.618; 0.616; 1.003; 0.8360;</td>
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<tr>
<td>Os(NH\textsubscript{3})\textsubscript{4}(H\textsubscript{2}O)\textsuperscript{−f}</td>
<td>−2.3, −2.3, −1.22; −2.2, −2.2, −1.08</td>
<td>+0.3855; +0.4097; 0.0; 0.0; 0.0; 0.0; 0.9795; 0.8954;</td>
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<tr>
<td>Ru(NH\textsubscript{3})\textsubscript{4}(H\textsubscript{2}O)\textsuperscript{−f}</td>
<td>−2.620, −2.620, −0.6; −2.620, −2.620, (0.0)</td>
<td>+0.7119; +0.9878; 0.0; 0.0; 0.0; 0.0; 1.052; 0.9678;</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>trans-[Ru(NH\textsubscript{3})\textsubscript{4}(4-pic)(H\textsubscript{2}O)]\textsuperscript{−f}</td>
<td>−1.20, −1.56, −2.81</td>
<td>−0.671</td>
<td>−0.112</td>
<td>0.167</td>
<td>0.9378</td>
</tr>
</tbody>
</table>

* Counterions are not given except for the two Os\textsuperscript{III} complexes reported here, but are noted for others below when relevant. \textsuperscript{1} Taken from Kremer; \textsuperscript{2} pyridine. The values presented are all composed with tetraethylammonium counterionation; for trans-Cs[OsCl\textsubscript{4}(CO)[py]]; $g_1$ = 2.46, $g_{\parallel}$ = 1.45. \textsuperscript{3} Taken from Diubghail et al.\textsuperscript{83}. \textsuperscript{4} Taken from de Paula et al.\textsuperscript{86}. The $g$ values reported are for the solid material; in aqueous or methanol solution, multiple species are reported due to replacement of Me\textsubscript{2}SO and/or chlorido ligands by solvent. \textsuperscript{5} Taken from Clarke et al.\textsuperscript{77}. The values for $|g_{\text{imin}}|$ were not observed (hence they are given in parentheses), but were proposed by their LFT analysis. In these systems, the proper choice of sign of the $g$ values is not apparent based on standard criteria, namely the magnitudes of $k$ or $\Delta$, so both choices are given. The results of additional calculations are given; for the trans complex, using a value for $|g_{\text{imin}}|$ that would likely be the largest possible, but difficult to observe experimentally: $|g_{\text{imin}}|$ = 0.9; for the cis complex using the a value for $|g_{\text{imin}}|$ that would likely be the smallest possible: $|g_{\text{imin}}|$ = 0.15 (a value of zero is possible, but does not provide enough information for the fitting process). We favor the first entry listed in each case, which is in agreement with the model originally favored by Clarke et al.\textsuperscript{77}. Taken from McGarvey et al.\textsuperscript{33} Two species are always observed in frozen aqueous solutions of [Os(NH\textsubscript{3})\textsubscript{4}(H\textsubscript{2}O)(C\textsubscript{2}H\textsubscript{4}O\textsubscript{4})]. \textsuperscript{4} Taken from McGarvey et al.\textsuperscript{33}. The experimentally reported range for $|g_{\text{imin}}|$ is 0.0 $\leq$ $|g_{\text{imin}}|$ $\leq$ 0.6; we thus present calculations for both 0.6 and 0.0 – the minimum magnitude. \textsuperscript{6} Taken from Souza et al.\textsuperscript{32} 4-pic = 4-picoline (4-methylpyridine). The specific $g$ values reported are those for frozen aqueous/ethylene glycol solution; those for a powder and in other solvents were similar. The signs of the $g$ values are not determined experimentally, but are determined here by the model of McGarvey,\textsuperscript{33b} which provides two choices of sign. Only the more physically plausible choice of signs is presented, except when both choices are acceptable.

Qualitatively for trans-[OsBr\textsubscript{4}(CO)]\textsuperscript{−}; the bromido ligands are expected to be $\pi$-donors (i.e., $g_{\pi}(X) > 0$) and the carbonyl ligands are expected to be strong $\pi$-acceptors (i.e., $g_{\pi}(Y) < 0$). This would give $\Delta > 0$, consistent with what is obtained from the $g$ value analysis. No cis-[MX\textsubscript{2}(Y, YZ)] complexes were known to McGarvey, but he proposed that $\Delta = -g_{\pi}(X) + g_{\pi}(Y)$ for this geometry.\textsuperscript{33b} The chlorido ligands (X) are expected to be $\pi$-donors and the indazole nitrogen atoms (Y) are expected to be weak $\pi$-acceptors ($g_{\pi}(X) \leq 0$), which would give $\Delta < 0$, again consistent with the $g$ values. A further discussion of the LFT parameters of these complexes is beyond the scope of this study, and would be in any case risky due to the overall relatively small number of Os\textsuperscript{III} coordination complexes characterized by EPR, and in particular, the lack of data on other Os\textsuperscript{III} cis/trans isomers. However, it is clear that the cis-Os\textsuperscript{III} complex, [1], as manifest in its “reversed” EPR spectrum, exhibits an electronic structure distinctively different from almost all of the related species listed in Table 4, in particular the trans-Os\textsuperscript{III} complexes. The wavefunction coefficients resulting from this analysis show that the “hole” is primarily in the $d_{xz}$ and $d_{yz}$ orbitals in the present system, while it is primarily in the $d_{xy}$ orbital in most other cases. The exception is [Ru(NH\textsubscript{3})\textsubscript{4}(4-pic)(H\textsubscript{2}O)]\textsuperscript{−} (4-pic = 4-picoline (4-methylpyridine)) wherein competing $\pi$-bonding involving the trans ligands was proposed to lead to the change in the
A quantum chemical theory (QCT) treatment presented in the next section provides further insight into $[^1]\text{I}^-$. Theoretical determination of $[^2]I^- g$-tensor

QCT calculations of EPR parameters of compounds containing heavy elements are challenging due to a dominating importance of spin–orbit relativistic effects. While in the case of compounds containing 3d transition metal complexes it is sufficient to treat the relativistic SOC effects in the evaluation of the $g$-tensor as a perturbation of the non-relativistic wavefunction, in the case of 4d- and 5d-transition metal compounds the explicit inclusion of both scalar and SOC relativistic effects appears necessary. Herein, we wish to compare the performance of the $g$-tensor evaluation of $[^2]I^-$ using a perturbative inclusion of SOC in the scalar relativistic ZORA wavefunction using either DFT or CASSCF levels of theory with the rigorous treatment of relativistic effects and $g$-tensor evaluation at the DKS level. Furthermore, the solvent polarization effects will be considered for further consistency via the COSMO approach.

The calculated $g$-tensor values are shifted from the experimental ones in the case of scalar ZORA DFT calculations with the perturbative inclusion of SOC. The solvent effects are large within the ZORA DFT treatment but do not lead to improvement with respect to experiments (Table 5). Considerably better agreement is achieved in the case of ZORA CASSCF calculations which yield a semi quantitative agreement with experiment [both in vacuo and via the COSMO(Me$_2$SO) solvent model]. Nevertheless, the CASSCF setup has been kept under conservative settings for the rotations out of the active space to enforce the d-electronic configuration of the open shell and/or

Table 5 Results of the theoretical determination by quantum chemical theory calculations of the $g$-tensor compared with experimental data

<table>
<thead>
<tr>
<th>Method</th>
<th>$g_x$</th>
<th>$g_y$</th>
<th>$g_z$</th>
<th>Details</th>
</tr>
</thead>
<tbody>
<tr>
<td>EXP. (94 K)</td>
<td>1.31</td>
<td>1.44</td>
<td>2.44</td>
<td>Me$_2$SO</td>
</tr>
<tr>
<td>EXP. (12 K)</td>
<td>1.37</td>
<td>1.45</td>
<td>2.50</td>
<td>H$_2$O/ethylene glycol 5 : 1 v/v</td>
</tr>
<tr>
<td>ZORA/BLYP</td>
<td>2.056</td>
<td>3.649</td>
<td>5.087</td>
<td>In vacuo</td>
</tr>
<tr>
<td></td>
<td>2.030</td>
<td>4.731</td>
<td>5.507</td>
<td>COSMO(Me$_2$SO)</td>
</tr>
<tr>
<td>ZORA/B3LYP</td>
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<td>4.428</td>
<td>5.094</td>
<td>In vacuo</td>
</tr>
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<td>2.021</td>
<td>4.908</td>
<td>7.258</td>
<td>COSMO(Me$_2$SO)</td>
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<tr>
<td>ZORA/CASSCF</td>
<td>1.665</td>
<td>1.780</td>
<td>3.090</td>
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<tr>
<td></td>
<td>1.594</td>
<td>1.702</td>
<td>3.188</td>
<td>COSMO(Me$_2$SO)</td>
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<tr>
<td>DKS/B3LYP</td>
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<td>2.536</td>
<td>In vacuo</td>
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<td></td>
<td>1.262</td>
<td>1.797</td>
<td>2.471</td>
<td>COSMO(Me$_2$SO)</td>
</tr>
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Fig. 7 Cell viability of human cancer cell lines after a 48 h treatment with cisplatin, $[^1]\text{I}$, Na$[^1]$ and $[^3]$, in relation to the untreated control (1% Me$_2$SO) using the Alamar Blue assay kit. Displayed are means ± standard deviation of six parallels.
in the active space. Although the rigorous DKS method does not yield the best agreement in vacuo, introducing a solvent model leads to a large improvement of the calculated results providing almost quantitative agreement with the experiment. This once again shows the need for a suitable treatment of both relativistic and spin polarization effects in the calculation of EPR parameters for 5d ions. As such, the calculated solvent effects on the electronic g-tensor parameters are so far the largest ones observed at the DKS level of theory and to the best of our knowledge, they are the largest ones calculated for systems containing heavy elements. Furthermore, the excellent agreement of DKS/COSMO theory with experimentally determined values is promising for a more systematic theoretical exploration of experimentally determined g-tensors for osmium(III) compounds, and potentially other paramagnetic 5d ion complexes, such as of rhenium(II) or iridium(IV).

Effective concentrations to affect cell viability

Having now the availability of well-characterized cis- as well as trans-osmium(III) complexes it is possible to make a direct comparison of their anti-cancer efficacy. Specifically, first to determine in which concentration range these complexes mediate cytotoxic effects, we conducted a cytotoxicity screen in four different cancer cell lines (HT29, 4T1, HEK293 and H446) and compared the effects to an untreated control and cisplatin. The results for select drug concentrations are displayed in Fig. 7. Strong effects on cell viability were only observed for the highest concentration of 300 µM, where all cell lines showed a viability of below 30% compared to the control, with the exception of 4T1 cells to treatment with trans-[OsIVCl4(kN1-2H-ind)2] ([3]). At 30 µM cisplatin reduced cell viability to 30–60%, while the osmium compounds did not reduce viability. HT29 cells were the only cells where the osmium compounds led to a stronger reduction of cell viability than cisplatin. The corresponding IC50 values are quoted in Table S1.

Antiproliferative activity in cancer cell lines

We further determined the potential of complexes [1] and Na[1] to inhibit proliferation and induce apoptosis in comparison with trans-[OsIVCl4(kN1-2H-ind)2] ([3]) and cisplatin in HT29 and 4T1 cells at a concentration of 200 µM. Here, we used an apoptosis detection kit to quantify the percentage of a cell population that has entered an early apoptotic (AV+/PI−) or late apoptotic/necrotic state (AV+/PI+) using flow cytometry (Fig. 8a). We found that the 4T1 (breast cancer) cells were very sensitive to cisplatin (more than 60% increase in AV+ and AV+/PI+ cells compared to untreated cells), while HT29 (colorectal) cells showed no significant difference between untreated and cisplatin treated cells (Fig. 8b). Interestingly, the cisplatin resistant HT29 cells were more sensitive to treatment...
with the osmium complexes than 4T1 cells. In particular, the 
AV+/Pi+ population increased significantly to 40.3% ([1]),
36.0% (Na[1]) and 57.2% ([3]) upon treatment (vs. 15.1% 
untreated, p < 0.05), indicating that the trans-isomer induced 
apoptosis most effectively in HT29 cells (Fig. 8b).

In the cisplatin sensitive 4T1 cells, we also observed a 
significant increase in the AV+/Pi+ population for all osmium 
compounds, but it was less pronounced ([1]: 27.7%, Na[1]: 
33.9%, trans-[OsIVCl4(kN-1-2H-ind)])2: 27.6%, untreated: 16.8%) 
and the cis- and trans-isomers [1] and [3] showed no difference 
(Fig. 8b).

In addition to flow cytometry data, we also measured cell 
viability and cell number of each sample in an automated cell 
counter. We found more pronounced effects on cell count 
than on cell viability, indicating that proliferation was affected, 
but the cells remained viable (Fig. 8c). In HT29 cells, cell viabi-
licity was significantly affected by cisplatin, Na[1] and [3]. The 
cell number was strongly reduced by over 75% by all osmium 
compounds and to 41% by cisplatin (Fig. 8c). In 4T1 cells, the 
strongest effect on cell count was observed for the cis-com-
 pound Na[1] (20.9% of untreated), followed by cisplatin at 
39.5% (Fig. 8c).

Conclusion

In this work we reported on the synthesis and crystal structure of 
cis-[OsIVCl4(N2-1H-ind)] ( [1]), the first cis-configured 
isomer in the series of compounds of the general formula 
[OsIVCl4(Hazole)2], where Hazole = indazole, pyrazole, 
benzimidazole, and imidazole. In addition, we showed that 
trans-[OsIVCl4(N2-1H-ind)] ( [2]) is also available, although 
more effective synthetic routes still need to be developed. Both 
compounds are complementary to their elsewhere reported 
trans-isomer [OsIVCl4(N2-1H-ind)]2 ( [3]). Electrochemical and 
chemical one-electron reduction of [1] afforded cis-configured 
osmium[II] species isolated as tetrabutylammonium and sodium 
salts, nBu4N[1] and Na[1], respectively. X-band EPR 
spectra of nBu4N[1] in Me2SO at 94 K and Na[1] in H2O/ethyl-
ene glycol 5:1 v/v at 12 K are characterized by ill-resolved 
broad lines with a rhombic g-tensor indicating the low-spin 
5d5 (S = ½) electronic configuration of osmium[II] in these 
complexes. A simple ligand-field theory analysis of the g values 
indicates that in these cis-isomers, the electronic configuration 
differs from that of trans-analogs in that the electron “hole” 
is primarily in a d_{xz,yz} orbital as opposed to a d_{xy} orbital in the 
much more commonly found species with trans-geometry. The 
g-tensor of the OsIII complexes of interest was further probed 
by quantum chemical theory. Specifically, the g-tensor values 
calculated by DKS/COSMO theory (with involvement of the 
solvent model) are in quite good agreement with those deter-
mined experimentally. Access to both cis- and trans-isomers of 
osmium[IV] and osmium[II] that are water soluble prompted us 
to investigate their in vitro anti-cancer activities. The complex-
es [1], Na[1] and [3] showed antiproliferative effects on 
cancer cells which were not coupled to their sensitivity to cis-
platin. Cytostatic effects were stronger than actual cell killing 
at 48 h incubation time. HT29 cells showed a higher sensitivity 
to treatment with osmium complexes than cisplatin. 4T1 cells 
showed lower rates of apoptosis induction by the osmium 
complexes than cisplatin, but a very strong cytostatic effect was 
observed for the cis-isomer Na[1]. The unavailability of a larger 
 scale procedure for the synthesis of the trans-isomer [2] makes 
the direct disclosure of the effect of geometrical isomerism on 
cytotoxicity in this case not possible. From the other side, a 
direct extrapolation of the results of the antiproliferative 
activity and cytostatic effects of the cis-isomer Na[1] on the 
ruthenium analogues is not possible as well, if one takes into 
account the body of accumulated evidence and cytotoxicity 
data for related ruthenium and osmium compounds 
and complexity of the problem. Even though the analogous cis-isomers of KP1019/NK1339 
remain unknown, their synthesis appears to be imminent, 
giving an opportunity that deserves to be exploited in the 
search for more effective metal-based anticancer drugs than 
those currently in clinical trials.

Conflicts of interest

The authors declare no competing financial interest.

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