

Cite this: *Dalton Trans.*, 2017, **46**,
3655

Stabilized lithium-ion battery anode performance by calcium-bridging of two dimensional siloxene layers†

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A Ca-bridged siloxene (Ca-siloxene) composed of two-dimensional siloxene planes with Ca bridging was synthesized *via* a solid state metathesis reaction using TaCl₅ to extract Ca from CaSi₂. Three different Ca-siloxenes synthesized at Cl₂/Ca molar ratios of 0.25, 1.25 and 2.5 (CS0.25, CS1.25 and CS2.5, respectively) were fabricated and investigated as anode active materials for lithium-ion batteries. Both secondary and primary Ca-siloxene particles, which serve to increase the contact interfaces with conductive materials and to generate accessible sites for lithium ions, respectively, were found to become smaller and to have increased pore volumes as the Cl₂/Ca molar ratio was increased. These Ca-siloxenes exhibited stable charge/discharge performance as anode materials, with 69–99% capacity retention after 50 charge/discharge cycles (compared with 36% retention for a conventional Kautsky-type siloxene). The charge capacity also increased with increases in the Cl₂/Ca molar ratio, such that the CS2.5 showed the highest capacity after 50 charge/discharge cycles. This may reflect the formation of Si₆Li₆ rather than SiLi_{4.4} and suggests the maintenance of layered Si planes for large capacity retention after charge/discharge cycling. The increase of contact interfaces between acetylene black (as a conductive material) and Ca-siloxenes was found to effectively increase the lithium-ion capacity of Ca-siloxene during high rate charge/discharge cycling.

Received 4th October 2016,
Accepted 18th February 2017

DOI: 10.1039/c6dt03837a

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1. Introduction

Certain anode materials have the potential to increase the capacity of lithium-ion batteries, so various graphite alternatives have been extensively studied for this purpose because of the limited theoretical capacity of graphite (372 mA h g⁻¹).^{1,2} Silicon (Si)-based compounds, such as SiLi_{4.4}, are among the possible alternatives and have shown high capacity values as anode materials (4200 mA h g⁻¹).³ One drawback associated with Si-based materials is structural decomposition in conjunction with volume expansion during charge/discharge cycling, leading to a loss of conductivity between the Si-based materials and conductive cell components. As a result, effects of binder have been reported for the prevention of loss of conductivity. For example, polyacrylate binders have been applied for improvement in the electrochemical performance of Si-based materials.⁴ Furthermore, Si-based materials having different structures and compositions have also been investi-

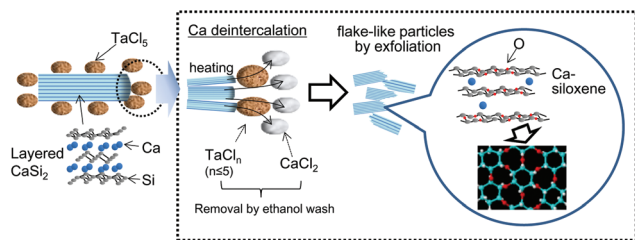
gated as a means of preventing this structural damage.^{5–14} For example, Si nanomaterials, including nanoparticles,^{5,6} nanowires⁷ and nanotubes,⁸ have been reported to exhibit effective capacity retention, especially in the presence of an active carbon matrix. Amorphous Si also shows promise in this regard,^{9,10} unlike crystalline Si, which is prone to high stresses and associated cracking. An interesting approach to fabricating porous Si materials has also been reported, with the aim of accommodating volume expansion during lithium insertion, using internal pores.^{11,12} Layered polysilane ((Si₆H₆)_x) has also been studied because its two-dimensional (2D) layered structure can serve as an analogue for graphite and contains sites that are accessible for lithium ion insertion/extraction.^{13,14}

Layered siloxene is another promising material since it incorporates 2D Si-based planes^{15–18} that may be applicable to the lithium storage reaction. Si planes with Kautsky-type siloxene structures are composed of Si₆O₃H₆ with Si₆ rings connected through Si–O–Si bridges.^{17,18} Unfortunately, in this structure, there are no bridges connecting the layered Si planes other than van der Waals forces, which may lead to cleavage between planes during charge/discharge cycles.¹⁹ In a previous study, we reported a convenient method for the synthesis of calcium-bridged siloxene (Ca-siloxene) *via* a solid state reaction²⁰ incorporating an ethanol wash¹⁹ (Scheme 1).

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† Electronic supplementary information (ESI) available: X-ray diffraction data and N₂ adsorption-desorption isotherm of Ca-siloxenes. See DOI: 10.1039/c6dt03837a



Scheme 1 The solid state reaction employed for Ca-siloxene synthesis.

This material shows potential for application as an anode active material in lithium-ion batteries.¹⁹ In this solid-state metathesis reaction, Ca is extracted from layered CaSi₂ with the aid of TaCl₅, acting as a Cl source to form CaCl₂, thus maintaining a partial Ca-bridging between the Si layers, the extent of which is controlled by varying the TaCl₅ concentration. This process forms Ca-siloxene owing to trace amounts of H₂O in the wash ethanol and HCl derived from tantalum chloride species, similar to the siloxene synthesis mechanism.^{17,18} Because of the Ca-bridging, this Ca-siloxene has demonstrated improved cyclability as an anode active material compared with the Kautsky-type siloxene.¹⁹

Herein, we report the effects of the morphology and composition of Ca-siloxene on its performance as an anode for lithium-ion batteries. In this work, primary Ca-siloxene particles were reduced in size to generate flake-like particles, varying with the extent of Ca extraction in the synthesis step, using a technique based on our previous work.¹⁹ This may not only facilitate internal pore formation to accommodate volume expansion, but could also improve cell performance by producing accessible sites for lithium ion insertion/extraction into the Si planes during charge/discharge cycles. In order to focus on the intrinsic anode performance, no binders were added for the comparison of capacity. The effect of the amount of conducting materials on capacity increase is also discussed herein in order to assess the usefulness of Ca-siloxene as a high capacity anode active material.

2. Experimental

2.1 Synthesis of Ca-siloxenes¹⁹

Samples were synthesized based on Cl₂/Ca molar ratios of 0.25, 0.5 and 2.5, and are denoted herein as CS0.25, CS1.25 and CS2.5, respectively. In these syntheses, CaSi₂ (3.2, 1.6 and 0.32 mmol for CS0.25, CS1.25 and CS2.5, Rare Metallic) and TaCl₅ (0.32 mmol, Wako Pure Chemical) were mixed together and packed in boron nitride crucibles. These mixtures were calcined at 215 °C for 5 h at a heating rate of 100 °C h⁻¹, which was sufficient to prevent overshooting the desired temperature. The mixtures, still in the boron nitride crucibles, were subsequently stored within sealed stainless steel cells, each with an inner volume of 10 cm³. The products were washed with anhydrous ethanol in order to remove the by-product CaCl₂ and excess TaCl₅ or their derivatives, followed by drying

under vacuum at 80 °C. All synthesis steps were performed under an Ar atmosphere. The Kautsky-type siloxene was synthesized for a reference sample as described in the literature.^{17,18}

2.2 Characterization of Ca-siloxenes

Scanning transmission electron microscopy (STEM) performed with an energy dispersive X-ray spectrometer (EDX) (JEOL 2100) was used to analyze the microstructures of the Ca-siloxenes. The secondary particle sizes were evaluated using a Microtrac MT3300EX particle size distribution analyzer, while the pore volumes were calculated by applying the N₂ adsorption method, employing a Quantachrome Autosorb-1. The crystalline structures of the samples were characterized by X-ray diffraction (XRD) using a Rigaku RINT-TTR (CuKα radiation, λ = 1.5418 Å, 50 kV, 300 mA). CaSi₂ was used as the reference sample, corresponding to Cl₂/Ca = nil.

2.3 Electrochemical properties

The lithium ion capacities of the Ca-siloxenes over 50 charge/discharge cycles were assessed using a Hokuto Denko HJ1010 galvanostat at 25 °C. The capacity was calculated based on a per Ca-siloxene mass basis with the subtraction of that for acetylene black under the same experimental conditions from the measured capacity, in order to confirm the capacity for Ca-siloxene. In these trials, each Ca-siloxene was mixed with conductive acetylene black (HS100, 39 m² g⁻¹, Denka Co., Ltd) at a Ca-siloxene/acetylene black ratio of 70/30 by weight, and then pressed onto a Ni foam to obtain electrodes in order to assess the essential Ca-siloxene performance. Li foil was used as the counter electrode and a 1 M solution of LiPF₆ in an ethylene carbonate/diethyl carbonate mixture (50/50 vol%) was employed as the electrolyte. A constant current of either 0.1 or 1 A g⁻¹ Ca-siloxene was applied with a voltage window of 0.02 to 1.50 V vs. Li/Li⁺. Scanning electron microscopy (SEM) images of electrodes composed of CS2.5/acetylene black with 5 wt% PTFE were acquired to assess volume expansion after 50 charge/discharge cycles, using a Hitachi SU-3500.

To determine the acetylene black proportion required to obtain optimal capacity and conductivity, CS2.5/acetylene black mixtures with mass proportions of 90/10, 80/20, 70/30, 50/50 and 30/70 were prepared, denoted as CS2.5-10, CS2.5-20, CS2.5-30, CS2.5-50 and CS2.5-70, respectively.

The electrical conductivity values of the CS2.5/acetylene black powder mixtures were measured using the two-probe method following uniaxial pressing at 10 MPa.

3. Results and discussion

3.1 Characterization of Ca-siloxene

The XRD spectra of the samples exhibit broad peaks derived from CaSi₂ in the region from 25 to 40°, and these peaks clearly increase in intensity according to the increase in the Cl₂/Ca ratio (Fig. S1†). This reflects the amorphous phase formation resulting from the fragmentation of layered CaSi₂ by



exfoliation, as has been reported in the literature.¹⁹ The presence of a Ca atom was confirmed in all Ca-siloxenes. The average Ca/Si molar ratios determined by EDX are 1.8/6, 0.8/6, and 0.3/6 for CS0.25, CS1.25, and CS2.5, respectively. The presence of Si–Ca bonding has been also confirmed by extended X-ray absorption fine structure spectra.¹⁹

Such a phenomenon is often accompanied by small particle formation and larger pore volume. Therefore, STEM images were used to examine the detailed morphology of the Ca-siloxene particles, as shown in Fig. 1(A)–(D). The secondary particles of CS0.25 consist of agglomerated primary particles with sizes on the order of several hundred nm. These are similar in shape to the CaSi₂ particles, although the secondary particles are evidently smaller than the CaSi₂ particles. The morphology is seen to gradually change with greater Cl₂/Ca molar ratios, due to the ongoing aggregation of smaller exfoliated primary particles, as seen in the case of the CS1.25 image in Fig. 1(B). Eventually, flake-like primary particles become dominant as in the CS2.5 (Fig. 1(C)), due to the exfoliation of the layered CaSi₂. The magnified image of the CS2.5 around flake-like primary particles is shown in Fig. 1(E). The thickness of flake-like primary particles is almost 15–20 nm. This flake-like structure has been also confirmed by an N₂ adsorption–desorption isotherm based on a typical H3-type hysteresis loop for a slit-like pore structure (Fig. S2†).¹⁹

The secondary particle sizes of the samples are shown in Fig. 2(A). The secondary particle size becomes smaller than that of the CaSi₂ (for which the Cl₂/Ca molar ratio is nil) as the Cl₂/Ca molar ratio is increased. This effect is controlled by the amount of TaCl₅ employed in the solid-state metathesis reaction, which in turn affects the exfoliation of the CaSi₂.¹⁹ Fig. 2(B) summarizes the pore volumes of the samples, which are

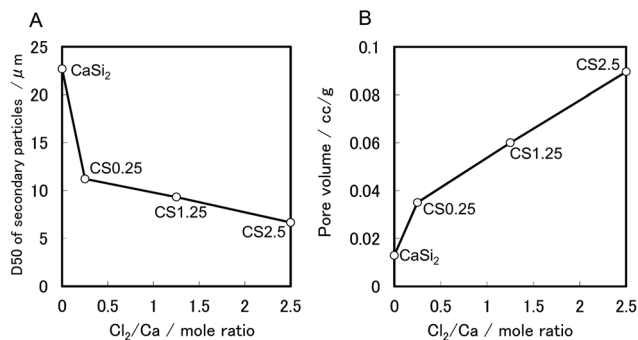


Fig. 2 (A) Secondary particle sizes, and (B) pore volumes of samples as functions of the Cl₂/Ca molar ratio.

seen to increase with increases in the Cl₂/Ca molar ratio. This transition also results from the morphological changes due to the formation of flake-like particles in the solid state exfoliation reaction of the CaSi₂. The peak of the pore size distribution is at approximately 4 nm (Fig. S2†), which may correspond to the size of the internal pores between the flake-like particles.¹⁹

These data demonstrate that it is possible to control the formation of nano-sized structures with increased Si-planes and pore volumes as a synergetic effect by varying the synthesis conditions, particularly the Cl₂/Ca molar ratio. This tailoring is capable of generating accessible sites for lithium ions within the Si planes by forming nano-sized flakes as primary particles. In addition, this technique produces increased contact interfaces between conductive particles through the formation of small secondary particles. Both these effects show promise in increasing the utility of this material as an anode in lithium-ion batteries, which will be discussed in the next section.

3.2 Charge/discharge cycling performance

The charge capacity values of the Ca-siloxene samples on a per Ca-siloxene mass basis over 50 charge/discharge cycles at a rate of 0.1 A g⁻¹, providing an indication of the degree of delithiation of the specimens, are shown in Fig. 3. All capacity data were derived from the Ca-siloxene capacity obtained after subtraction of the acetylene black capacity. All Ca-siloxenes exhibit stable charge/discharge performance as an anode material, and do not show rapid capacity fading after the first few cycles. In contrast, c-Si and Kautsky-type siloxenes are known to undergo capacity fading within the initial ten charge/discharge cycles.

The capacity of these Ca-siloxenes was evidently increased on increasing the degrees of Ca extraction owing to the exposed Si-planes and the larger pore volume that became available for lithiation/delithiation, such that the CS2.5 exhibited the highest capacity after 50 charge/discharge cycles. Interestingly, the capacity of the CS2.5 was higher than that of a Kautsky-type siloxene after 50 charge/discharge cycles, indicating the structural stability of the CS2.5 resulting from

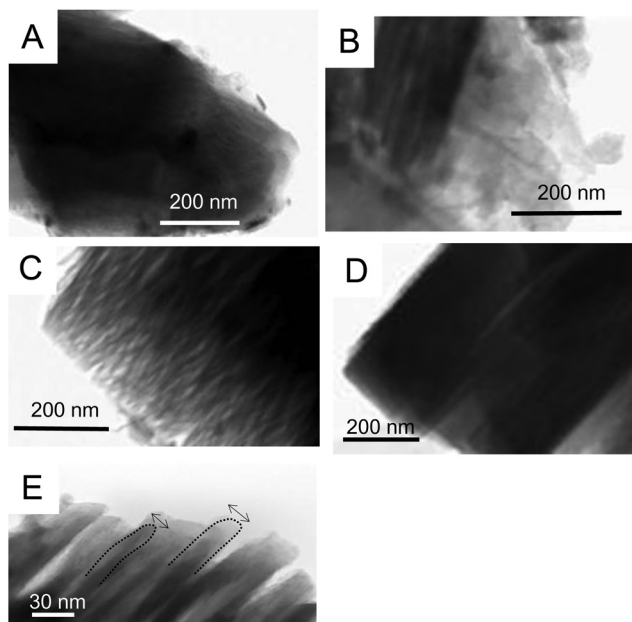


Fig. 1 STEM images of (A) CS0.25, (B) CS1.25, (C) CS2.5 (D) CaSi₂ and (E) CS2.5 (magnified image).



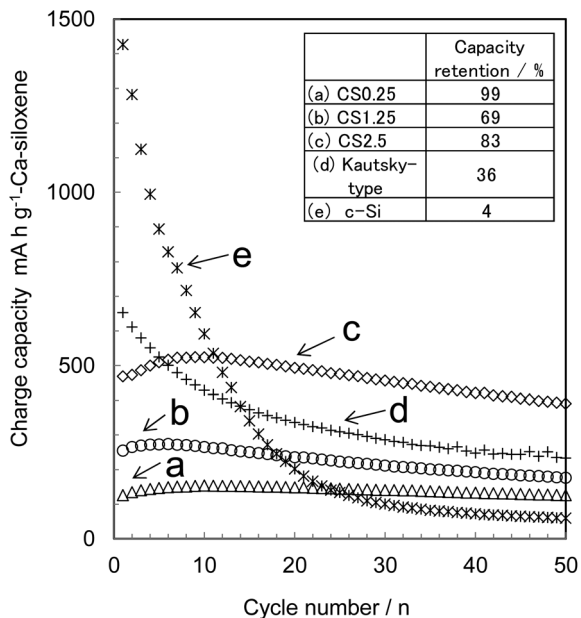


Fig. 3 Charge capacity per Ca-siloxene mass (*i.e.*, the capacity of acetylene black was subtracted) over 50 charge/discharge cycles at a rate of 0.1 A g^{-1} (inset: capacity retention after 50 cycles). Legend: (a) CS0.25, (b) CS1.25, (c) CS2.5, (d) Kautsky-type siloxene, and (e) c-Si.

Ca-bridging between Si planes,¹⁹ even though the Kautsky-type siloxene and the Ca-siloxene have similar Si plane structures. The CS0.25 maintained almost all of its original capacity throughout 50 charge/discharge cycles, exhibiting 99% capacity retention (Fig. 3, inset), although it did not show the same capacity as the CS2.5, primarily due to the morphological differences in the primary particles of these two materials. For example, the CS0.25 retained a high concentration of Ca atoms, which prevented the formation of Si–Li bonds, and was chiefly composed of large particles with lower pore volumes, as shown in Fig. 2. These conditions do not favor the approach of lithium ions towards Si atoms. The capacity retention of the CS1.25 was not increased to the same extent as that of the CS2.5, likely because of the balance between the formation of flake-like particles and the extent of Ca extraction required for effective lithium ion storage in the Si planes in this material. Therefore, it can be said that the morphology of the Ca-siloxene affects the capacity largely because the available Si-plane formation with larger pore volumes provides more lithium ion accessible sites in the flake-like primary particles, as discussed in section 3.1.

Fig. 4(A) shows the charge/discharge capacities and coulombic efficiency of CS2.5 over 50 cycles at a rate of 0.1 A g^{-1} . Some amount of irreversible capacity was observed in the first discharge step, probably due to solid electrolyte interface (SEI) formation on the CS2.5 surface, as often seen in the Si-based compounds.^{21,22} The voltage plateau was present at around 0.7 V, probably corresponding to the SEI formation (Fig. 4(B), charge/discharge curves). Therefore, such a formation also seems to contribute stable charge/discharge cycling in CS2.5.

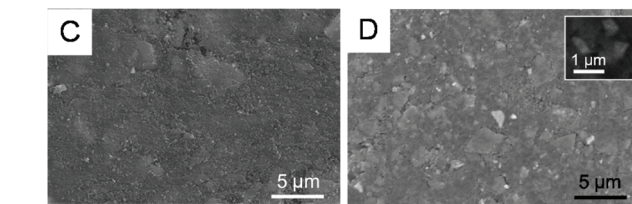
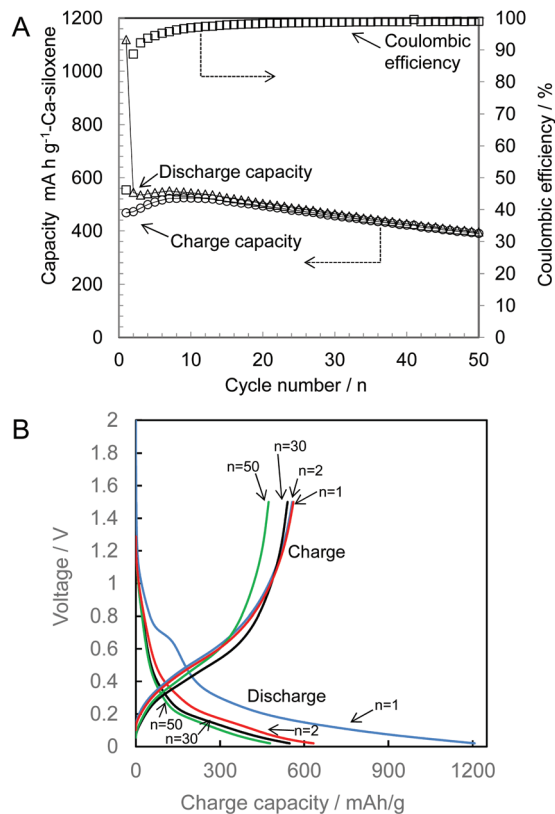


Fig. 4 (A) Charge/discharge capacity per Ca-siloxene mass and coulombic efficiency, (B) charge discharge curves of CS2.5 over 50 cycles, (C) SEM images of the CS2.5 anode before cycling, and (D) after 50 cycles (inset: magnified CS2.5 particles).

Although low coulombic efficiency at the first cycle was found in CS2.5 due to the SEI formation, a value more than 97% was maintained after the first several cycles.

Fig. 4(C) and (D) present the SEM images of the CS2.5 anode before and after 50 charge/discharge cycles, respectively. No obvious cracks are seen between the Ca-siloxene particles and the acetylene black after cycling. Although it is possible that volume expansion of the Ca-siloxene samples could occur during charge/discharge cycles, since this phenomenon is often seen with Si-based materials, this volume change may be accommodated owing to the internal pores between the flake-like primary particles. Concerning the morphology change of the CS2.5 particles after charge/discharge cycling, pulverization or cracking of particles was not observed (Fig. 4(D) inset, magnified image of CS2.5 particles), and a Ca atom was also present. In the case of the Si particles, clear particle pulverization due to $\text{SiLi}_{4.4}$ formation with huge volume expansion



(>400%) during the charge/discharge cycling was reported to be one of the reasons for the Si anode deterioration.²³ Therefore, the rapid decrease in the charge/discharge capacity is believed to have been prevented by successfully maintaining the contact interfaces between the Ca-siloxene and the conducting particles and also inside the particles, which may be one of the advantages over the Si anode.

Interestingly, the maximum capacity of the CS2.5 (524 mA h g⁻¹ Ca-siloxene, corresponding to 692 mA h g⁻¹ (Si mass basis)) was closer to the capacity of the Si₆Li₆ (approximately 957 mA h g⁻¹) based on the Si planes of a polysilane,¹⁴ rather than the 4200 mA h g⁻¹ value for SiLi_{4.4}. The smaller CS2.5 capacity relative to that of the Si₆Li₆ may be due to Si–Ca bonding, which is not available for Li storage. This result is in agreement with a Ca-siloxene structure including Si planes, and also suggests that these Si planes are sustained during the charge/discharge cycling compared with c-Si and Kautsky-type siloxenes. Similar results were theoretically predicted in the previous report: single and double layer silicone (Si monolayer sheet) showed little irreversible structural change during the charge and discharge cycles.²⁴

3.3 Effect of the conductive materials on capacity

The electrical conductivity of Ca-siloxene is known to be low: 6.4×10^{-4} S cm⁻¹ for CS0.25 and 4.8×10^{-8} S cm⁻¹ for CS2.5.¹⁹ Hence, the mixing of conductive materials with Ca-siloxene is an effective means of increasing the conductivity between the conductive material and the Ca-siloxene particles and also between the anode and the collector in a lithium-ion battery. This should lead to an increase in the charge/discharge capacity through optimization of the proportion of conductive material. For this reason, acetylene black, acting as a conductive material, was mixed with the CS2.5, which had demonstrated the best capacity among the Ca-siloxenes (see section 3.2). The charge capacity of CS2.5 with varying amounts of acetylene black is summarized in Fig. 5. The capacity was evidently increased with increases in the ratio of acetylene black to CS2.5, and the capacity was maintained in the CS2.5-30, -50 and -70 after 50 cycles. The CS2.5-70 also exhibited a constant charge capacity without any fading. In contrast, the CS2.5-20 and -30 showed a gradual decrease in capacity with prolonged cycling relative to the CS2.5-70, likely due to the slight ongoing decrease in the contact interface between the particles of CS2.5 and acetylene black during these charge/discharge cycles. The CS2.5-50 showed a slight decrease in capacity throughout 50 cycles, which is almost intermediate between CS2.5-30 and -70, just corresponding to the conductive material ratio. The CS2.5-10 had a rather small capacity, presumably because the contact interface between the CS2.5 and the acetylene black particles was insufficient to maintain electrical conductivity when employing the small amount of acetylene black in this material.

The conductivity of the CS2.5 pellets made with over 10 wt% acetylene black was almost 0.2 S cm⁻¹, a value that is almost identical to that of pure acetylene black under the same experimental conditions. Therefore, a conductive

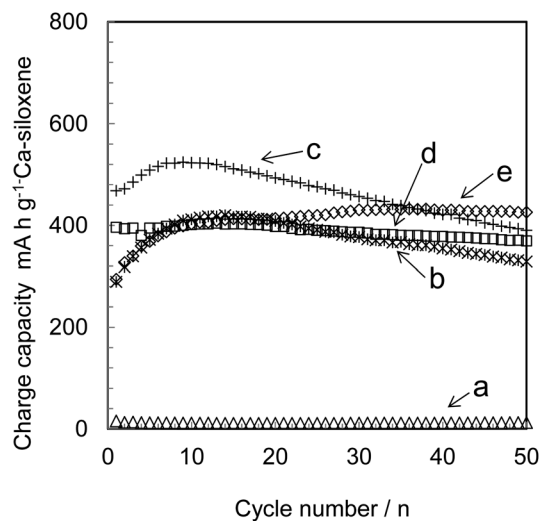


Fig. 5 Charge capacity per Ca-siloxene mass over 50 charge/discharge cycles at a rate of 0.1 A g^{-1} for samples with varying CS2.5/acetylene ratios. Legend: (a) CS2.5-10, (b) CS2.5-20, (c) CS2.5-30, (d) CS2.5-50 and (e) CS2.5-70.

pathway is evidently present in samples containing at least 10 wt% acetylene black, although intimate contact between the CS2.5 and the acetylene black particles may not occur until 20 wt% acetylene black is present, as seen in Fig. 5.

The formation of a contact interface between electrode active materials and conducting materials plays an important role, especially in a high charge/discharge rate, suggesting that the CS2.5-70 is expected to prevent capacity decrease under such a rate condition. Thus the effects of the contact interface between the Ca-siloxene and acetylene black particles on the current density of the CS2.5-30 and CS2.5-70 specimens are presented in Fig. 6. The capacity of the CS2.5-70 at a charge/

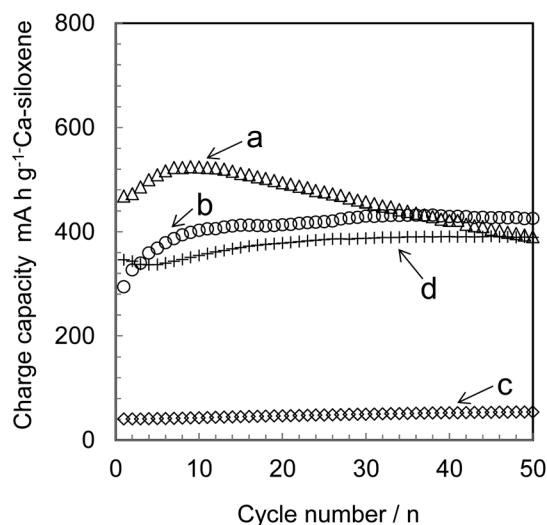


Fig. 6 Charge capacity per Ca-siloxene mass over 50 charge/discharge cycles at a rate of 0.1 A g^{-1} with sample/carbon mass ratios of (a) 70/30 and (b) 30/70, and at a rate of 1 A g^{-1} with ratios of (c) 70/30 and (d) 30/70.



discharge rate of 1 A g^{-1} remained at almost 90% of the value obtained at 0.1 A g^{-1} . In contrast, the capacity of the CS2.5-30 at 1 A g^{-1} dropped to only 15% of the capacity measured at 0.1 A g^{-1} . These data suggest that it is important to maintain the contact interface between the Ca-siloxene and the conducting material, especially during charge/discharge cycling at a high current density.

Ca-siloxene has the potential to generate a larger capacity than that of carbon, and exhibits a fast charge/discharge rate with structural stability during the charge and discharge process, as described above. Thus, an interface between the Ca-siloxene and the conducting material may be a key factor related to the future practical application of this material to lithium-ion batteries. Physical mixtures need a large amount of conducting material. Fabricating Ca-siloxene/carbon composite materials by adding organic precursors in the preparation step prior to calcination^{25,26} would be useful to reduce the amount of carbon. It is important to explore such Ca-siloxene coated with thin carbon layers in a further study.

4. Conclusions

Three different Ca-siloxenes, synthesized *via* a solid state metathesis reaction using TaCl_5 and CaSi_2 , were assessed as anode active materials for lithium-ion batteries. Decreases in the primary and secondary particle sizes of the Ca-siloxene and increases in the pore volume were observed at higher Cl_2/Ca molar ratios, leading to improved capacity after 50 charge/discharge cycles. In order to judge the maximum capacity and cyclability of Ca-siloxene, the mixing amount of acetylene black as a conductive material was changed. The formation of Si_6Li_6 and the structural stability of the Si planes during charge/discharge cycling were suggested. These findings reveal some of the fundamental properties of Ca-siloxenes with respect to their structure and composition, and should allow optimization of these compounds for future applications as anode materials in lithium-ion batteries.

Acknowledgements

The authors thank Ms Akiko Ueki of Toyota central R&D Labs. for the assistance with TEM analysis.

Notes and references

- H. Xia, W. Xiong, C. K. Lim, Q. Yao, Y. Wang and J. Xie, *Nano Res.*, 2014, **7**, 1797–1808.
- H. Xia, Y. Wan, G. Yuan, Y. Fu and X. Wang, *J. Power Sources*, 2013, **241**, 486–493.
- H. Wu and Y. Cui, *Nano Today*, 2012, **7**, 414–429.
- S. Komaba, T. Ozeki, N. Yabuuchi and K. Shimomura, *Electrochemistry*, 2011, **79**, 6–9.
- S. Zeng, D. Liu, Y. Chen, J. Qian, Y. Cao, H. Yan and X. Ai, *J. Mater. Chem. A*, 2015, **3**, 9938–9944.
- H. Kim, M. Seo, M.-H. Park and J. Cho, *Angew. Chem., Int. Ed.*, 2010, **49**, 2146–2149.
- M. R. Zamfir, H. T. Nguyen, E. Moyon, Y. H. Lee and D. Pribat, *J. Mater. Chem. A*, 2013, **1**, 9566–9586.
- H. Wu, G. Chan, J. W. Choi, I. Ryu, Y. Yao, M. T. McDowell, S. W. Lee, A. Jackson, Y. Yang, L. Hu and Y. Cui, *Nat. Nanotechnol.*, 2012, **7**, 310–315.
- U. Kasavajjula, C. S. Wang and A. J. Appleby, *J. Power Sources*, 2007, **163**, 1003–1039.
- M. T. McDowell, S. W. Lee, J. T. Harris, B. A. Korgel, C. Wang, W. D. Nix and Yi Cui, *Nano Lett.*, 2013, **13**, 758–764.
- A. Magasinski, P. Dixon, B. Hertzberg, A. Kvit, J. Ayala and G. Yushin, *Nat. Mater.*, 2010, **9**, 353–358.
- Z. Zhang, Y. Wang, W. Ren, Q. Tan, Y. Chen, H. Li, Z. Zhong and F. Su, *Angew. Chem., Int. Ed.*, 2014, **53**, 5165–5169.
- Y. Kumai, H. Kadoura, E. Sudo, M. Iwaki, H. Okamoto, Y. Sugiyama and H. Nakano, *J. Mater. Chem.*, 2011, **21**, 11941–11946.
- M. Ohashi, H. Nakano, T. Morishita, M. J. S. Spencer, Y. Ikemoto, C. Yogi and T. Ohta, *Chem. Commun.*, 2014, **50**, 9761–9764.
- A. Weiss, G. Beil and H. Meyer, *Z. Naturforsch., B: Anorg. Chem. Org. Chem.*, 1979, **34**, 25–30.
- H. Nakano, M. Ishii and H. Nakamura, *Chem. Commun.*, 2005, 2945–2947.
- S. Yamanaka, H. Matsu-ura and M. Ishikawa, *Mater. Res. Bull.*, 1996, **31**, 307–316.
- H. Kautsky, W. Vogell and F. Oeters, *Z. Naturforsch., B: Anorg. Chem. Org. Chem. Biochem. Biophys. Biol.*, 1955, **10**, 597–598.
- H. Imagawa, N. Takahashi, T. Nonaka, Y. Kato, K. Nishikawa and H. Itahara, *J. Mater. Chem. A*, 2015, **3**, 9411–9414.
- S. Y. Oh, H. Imagawa and H. Itahara, *J. Mater. Chem. A*, 2014, **2**, 12501–12506.
- S. Y. Oh, H. Imagawa and H. Itahara, *Chem. – Asian J.*, 2014, **9**, 3130–3135.
- X. Wu, Z. Wang, L. Chen and X. Huang, *Electrochem. Commun.*, 2003, **5**, 935–939.
- J. Wang, T. Xu, X. Huang, H. Li and T. Ma, *RSC Adv.*, 2016, **6**, 87778–87790.
- G. A. Tritsarlis, E. Kaxiras, S. Meng and E. Wang, *Nano Lett.*, 2013, **13**, 2258–2263.
- M. Thakur, S. L. Sinsabaugh, M. J. Isaacson, M. S. Wong and S. L. Biswal, *Sci. Rep.*, 2012, **2**, 795.
- G. Liu, S. Xun, N. Vukmirovic, X. Song, P. Olalde-Velasco, H. Zheng, V. S. Battaglia, L. Wang and W. Yang, *Adv. Mater.*, 2011, **23**, 4679–4683.

