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The effect of ring size on the selective carboxylation of cycloalkene oxides†

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Carbon dioxide utilisation technology can contribute to the reduction of atmospheric CO₂ levels both through its sequestration from flue gases and indirectly by relieving pressure on conventional feedstocks in chemical manufacturing. A promising approach is to employ CO₂ to produce valuable cyclic carbonates (CCs) in reaction with suitable epoxides. This also has the advantage that carbon dioxide replaces toxic and hazardous reactants such as phosgene. In earlier work we have investigated the synthesis of epoxides from cycloalkenes using supported gold and gold-palladium nanoparticles as catalysts and oxygen from air as the oxidant under solvent free conditions. A strong dependence of epoxide selectivity on ring size was observed with C5 < C6 < C7 ≪ C8. In this study we extend this work to the investigation of cycloaddition of CO₂ to different cycloalkene oxides with the ultimate aim of designing a process in which both epoxidation of an alkene and incorporation of CO₂ could be achieved in a single process. However, we have found the opposite trend for the selectivity to carbonates: smaller ring cycloalkene oxides giving the highest carbonate selectivities while large rings do not yield CCs at all. The product distributions suggest that an alternative ring opening of the epoxides to yield alcohols and ketones is preferred under all the experimental conditions explored for larger ring systems. Additionally, the mechanism of the CC synthesis using a quaternary ammonium salt and ZnBr₂ as the catalyst system was investigated using DFT methods. The results of the calculations support the experimental findings.

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Introduction

The main sources of anthropogenic CO₂ emissions are from the use of fossil fuels such as natural gas, oil and coal. These emissions lead to an increase in the concentration of carbon dioxide in the atmosphere. Carbon dioxide utilisation technology can contribute to reducing the CO₂ level by using carbon dioxide as a starting material and transforming it into valuable chemicals such as cyclic carbonates.¹ There is a wide range of applications of cyclic carbonates; they are extensively used as intermediates in the synthesis of pharmaceuticals, raw materials for engineering plastics, aprotic polar solvents and electrolytes for lithium-ion batteries.^{2–6} The use of cyclic carbonates (CCs) is one of the most effective routes for carbon dioxide fixation, and is usually prepared *via* the cycloaddition of carbon dioxide with epoxides to form a five-membered ring.^{3,7–10} Carbon dioxide is a C1 building block with 100%

atom economy and thus, a better alternative than toxic and hazardous reagents such as phosgene.¹¹ Many studies have reported quaternary ammonium salts to be active for the cycloaddition of CO₂ to epoxide.^{12–15} These are thought to act through nucleophilic attack by the anion of the salt to ring open the epoxide which then allows CO₂ to react with the resulting oxyanion.¹⁶ The nucleophile is eliminated on ring closure to form the cyclic carbonate. This requirement of the nucleophile to both facilitate ring opening and act as a leaving group on cyclic carbonate ring closure has been used to explain why Br[–] is a good choice of nucleophile.¹⁷ DFT calculations using ethene as a model reagent and NEt₄ Br have shown that the rate limiting step (calculated barrier 30.7 kcal mol^{–1}) is the initial nucleophilic attack for ring opening of the epoxide.¹⁸ Experimentally NBu₄ Br is commonly used as the longer alkane chains shield the cation positive centre more effectively and so increase the nucleophilicity of the anion. The use of co-catalysts such as Mg porphyrins¹⁹ and ionic liquids²⁰ have been shown to reduce the barrier to ring opening to 13.5 and 12.0 kcal mol^{–1} with respect to reagents, respectively. The calculations show that these act to stabilise the forming oxyanion in the epoxide ring opening transition state. A similar role also been found to be important for amino acid co-catalysts.²¹

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Alternative nucleophiles such as (4-dimethylamino)pyridine (DMAP) with (salen)Cr(III)²² and hydroxide anions with NBu_4 (ref. 23) have also been investigated and shown to follow similar mechanisms. For systems containing redox active metals such as Co in α -Keggin-type structures calculations have also shown that ring opening of the epoxide can occur without the intervention of a nucleophilic agent through electron transfer and oxidation of the metal centre.²⁴

To realise cyclic carbonate production on a large scale, epoxides themselves will have to be synthesised by the oxidation of alkenes. The direct synthesis of cyclic carbonates from olefins, avoiding additional work-up procedures, would be an interesting and economically feasible route. Cyclic carbonates can be produced directly from CO_2 and activated alkenes such as styrene using a supported Au catalyst to form the epoxide and zinc bromide with tetrabutylammonium as catalyst for the CO_2 insertion step, in one-pot.²⁵ However, to date the extension of this technology to the more challenging cycloalkenes has not been successful. In spite of the fact that the oxidative-carboxylation of olefins has been known since 1962,²⁶ little attention was paid to this method compared with the route that employs epoxide as a starting material. The combination of two reactions in a one-pot process requires compatibility between the reaction conditions such as temperature, pressure and suitable catalysts for each reaction. In previous studies, we have investigated the epoxidation of cycloalkenes, using supported gold and gold-palladium nanoparticles as catalysts under solvent free conditions with air as the oxidant and small amounts of a radical initiator.²⁷ A range of substrates was tested with varying ring size (C_5 to C_{12}). It was found that the selectivity to epoxide increases with increasing ring size of the cycloalkene. This has lead us to re-investigate the concept of coupling an epoxidation and carbonylation catalyst for these substrates. As a next step in that direction we now report the ring size dependence of the carbonylation stage of the reaction for a series of cyclic alkene oxides. We will show that cycloaddition of CO_2 to various sized cyclic alkene oxides can be completed using a quaternary ammonium salt and ZnBr_2 under green conditions, without solvent. We show that the reactivity of the cyclic alkene oxides is related to the ring size, with the smaller rings more reactive than larger ones at the same temperature (90 °C). Additionally, the selectivity to the cyclic carbonate also depends on the size of the cyclic alkene oxide ring; selectivity is reduced as the ring size of cycloalkene oxide is increased.

We combined our experimental data with DFT calculations for the example of cyclopentene oxide in order to obtain a complete understanding of the reaction route. This allows us to highlight the role of ZnBr_4 as a co-catalyst and to discuss the origin of the side products observed.

Experimental

Chemicals

Tetrabutylammonium bromide Bu_4NBr (98+%, Alfa Aesar), zinc bromide ZnBr_2 (99.9%, Alfa Aesar), cyclopentene oxide

$\text{C}_5\text{H}_8\text{O}$ (98%, Aldrich) cyclohexene oxide $\text{C}_6\text{H}_{10}\text{O}$ (98%, Aldrich), cyclooctene oxide $\text{C}_8\text{H}_{14}\text{O}$ (99%, Aldrich), cyclo-dodecene oxide $\text{C}_{12}\text{H}_{22}\text{O}$ (95%, Aldrich) were used as received from the suppliers without further purification.

Cycloalkene oxide reactions with CO_2

Catalytic experiments were performed in a stainless steel autoclave reactor (50 cm³, Parr Instrument), fitted with an inner Teflon liner. Typically, a specified amount of the Bu_4NBr and ZnBr_2 catalysts and the epoxide (5 ml) were added into the autoclave. The reactor was purged 3 times with CO_2 before being pressurized with CO_2 to the desired pressure. The autoclave was heated to required temperature, under continuous stirring. At the end of the reaction the reactor was cooled to 5 °C in an ice water bath. Products from the reaction were analysed using gas chromatograph (Varian star 3400 CX) fitted with a DB-5 column and a flame ionization detector (FID). For compounds identification GC-MS was used (Walters GCT Premier GC fitter with a HP 6890 N Mass spectrometer) which allowed further product quantification for some selected experiments. The experiments were repeated up to three times and the standard deviation was 1%.

DFT calculations

All calculations presented were carried out with the Gaussian09 software package.²⁸ We use the B3LYP hybrid functional^{29–32} with a 6-31G(d,p) basis^{33–38} throughout. Transition state geometries were identified through a combination of constrained co-ordinate scans and standard transition state search optimisation. A transition state was accepted only when a single negative mode corresponding to the expected reaction co-ordinate was found in a frequency calculation.

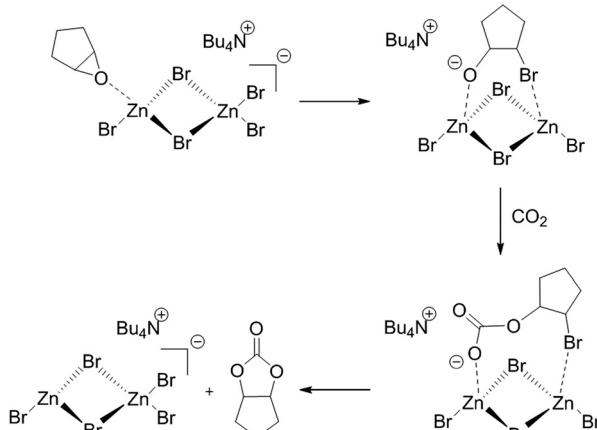
Results and discussion

Mechanism of cycloaddition of CO_2 to cyclopentene oxide

Darensbourg and Holcamp have suggested that the role of the ZnBr_2 based catalyst in this carbonylation reaction is to act as a Lewis acid stabilising the epoxide oxygen atom during the key epoxide ring opening step. CO_2 is then activated forming a linear carbonate intermediate before cyclization to a cyclic carbonate.³⁹ Scheme 1 shows an outline of the steps envisaged for the incorporation of CO_2 into cyclopentene oxide (as an example epoxide) catalysed by the quaternary ammonium salt and ZnBr_2 catalyst system.

This follows a similar mechanism to that proposed by North and Pasquale for the Al(Salen) catalysed incorporation of CO_2 into alkene oxides.⁴⁰ In the active form Br^- from the quaternary ammonium salt is added to one of the Zn centres of a dimer to form a $[\text{Zn}_2\text{Br}_5]^-$ anion. The epoxide first coordinates to the $[\text{Zn}_2\text{Br}_5]^-$ species via a dative bond with a Zn Lewis acid centre. Transfer of Br^- to one of the carbon atoms of the epoxide leads to epoxide ring opening with the resulting intermediate stabilised by interaction with both of





Scheme 1 Reaction sequence for the incorporation of CO_2 into cyclopentene oxide (as an example cyclic epoxide) to form a cyclic carbonate catalysed by $\text{Bu}_4\text{N}^+\text{Zn}_2\text{Br}_5^-$.

the Zn Lewis acid sites. CO_2 can now be incorporated to form a carbonate intermediate which ring closes to form the cyclic carbonate product. At the same time, Br^- is returned to the Zn dimer to reform the catalyst.

We initially used DFT calculations to probe the detail of this reaction scheme. The catalyst is modelled by a $[\text{Zn}_2\text{Br}_5^-]$ cluster with a charge balancing $[\text{NEt}_4]^+$ cation, we have used tetraethylammonium rather than tetrabutylammonium as the counter ion for computational efficiency. We also considered cyclopentene oxide as a model reagent. The calculated potential energy for each step of the reaction is shown in Fig. 1. Chemical structure diagrams and atomic co-ordinates for each structure are available from the ESI.[†]

The initial co-ordination of the epoxide to a Zn centre is energetically favourable, resulting in a lowering of the system energy by some 87 kJ mol⁻¹ (Int. 1). From this point, the ring opening of the epoxide group in isolation, following an $\text{S}_{\text{N}}1$ pathway, was found to lead to high energy transition states. The exact energy depends on the orientation of the resulting carbocation with respect to the $[\text{Zn}_2\text{Br}_5^-]$ complex with energies ranging between 38 and 46 kJ mol⁻¹ higher than the reference state of the isolated reagents ($\text{S}_{\text{N}}1$, TS-1t and $\text{S}_{\text{N}}1$, TS-1b) and 133 kJ mol⁻¹ above the co-ordinated epoxide (Int. 1). Co-ordination of a Br^- ion to the three co-ordinate C atom which is exposed on epoxide ring opening can then occur based on the addition of a terminal or a bridging Br^- ion. The addition of “t” and “b” labels to structural labels in Fig. 1 has been used to distinguish these cases throughout. Addition of a bridging Br^- along the $\text{S}_{\text{N}}1$ route requires a second barrier at 65 kJ mol⁻¹ ($\text{S}_{\text{N}}1$, TS-1b') to be surmounted, presumably due to breaking open of the Zn–Br–Zn bridge in this structure (Fig. 2b). The addition of a terminal Br^- is found to be barrierless leading to Int. 2t. Alternatively, the epoxide ring opening from Int. 1 can follow pathways leading to transition states for a concerted $\text{S}_{\text{N}}2$ type mechanism in which the new C–Br bond is formed as the epoxide ring opens and these were found to be considerably lower in energy. The bromination of the epoxide carbon atom can, again, takes place either using a bridging or terminal Br^- ion.

Reaction using a terminal Br^- ion was found to be preferred leading to a transition state in which the Br^- nucleophile can align with the receiving C atom at the same time that the epoxide oxygen atom is stabilised on the Lewis acid Zn centre, as shown in Fig. 2a. This type of transition state

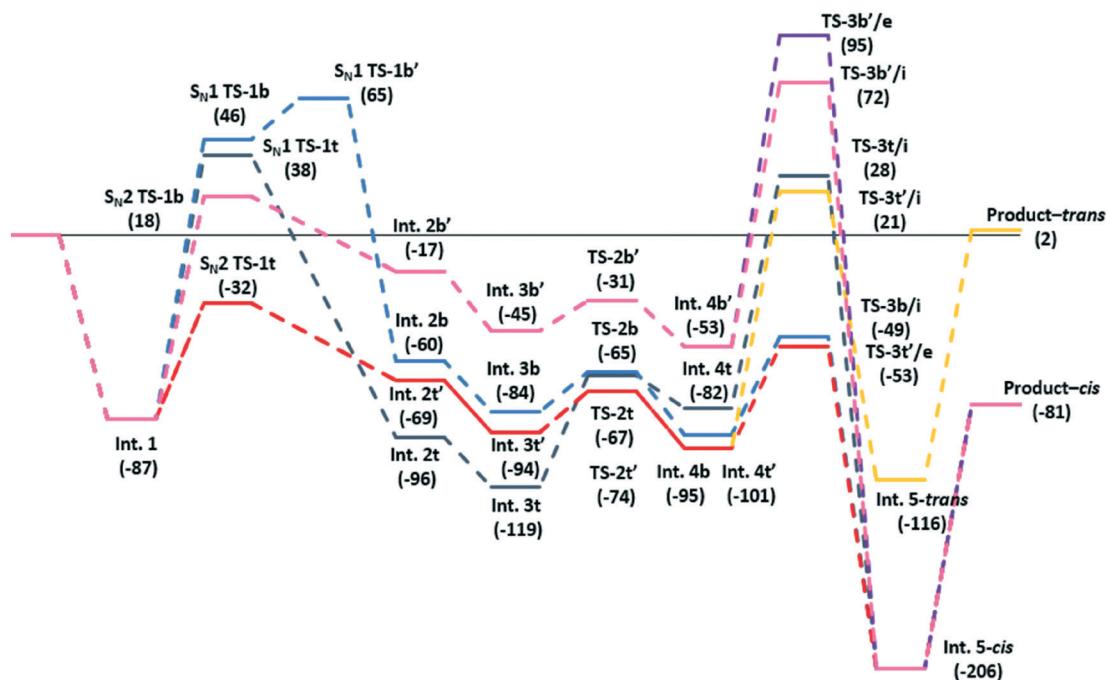


Fig. 1 Calculated pathways for the addition of CO_2 to pentene oxide to form a cyclic carbonate. Bracketed numbers are energies in kJ mol⁻¹ relative to the starting reagents shown to the left.

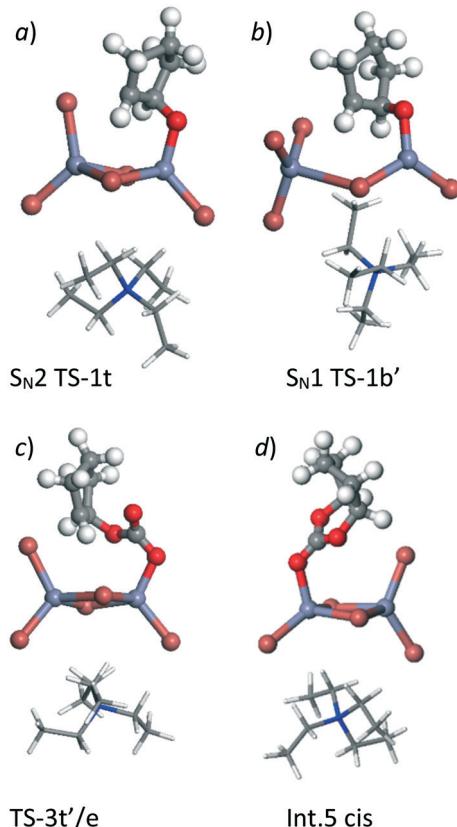


Fig. 2 Calculated structures for example points on the PES shown in Fig. 1. a) Lowest energy ring opening TS, S_{N2} TS-1t', b) second part of S_{N1} type epoxide ring opening mechanism S_{N1} TS-1b', c) lowest calculated barrier for carbonate ring closure, TS-3t'/e and d) product carbonate co-ordinated to Lewis acid centre, Int.5 cis. The NEt_4^+ cation is shown as wire all other atoms using ball and stick representations. Atom colours follow: C: grey, H: white, N: dark blue, O: red, Br: brown and Zn: light blue.

geometry has previously been shown to be important using Mg porphyrin based catalysts for the ring opening of the epoxide,¹⁹ in this case the structure has an energy of -32 kJ mol^{-1} (S_{N2} , TS-1t) relative to the initial reference state compared to 18 kJ mol^{-1} when a bridging Br^- ion is used in the S_{N2} step (S_{N2} , TS-1b). The intermediates formed from this point have both O and Br substituents on the cyclopentane ring stabilised by interaction with Zn centres. There are two possible chemical structures for the intermediate formed following the bromination of the epoxide, depending on the formation of the new C–Br bond from a bridging Br^- ion (Int. 2b) or from a terminal Br^- ion (Int. 2t). In Fig. 1 four distinct intermediates are shown as the detail of the configuration depends on the route taken to form the intermediate. Even so, both intermediates with a terminal Br^- added are lower in energy than those forming a bridging Br–C bond. We have followed the CO_2 insertion step for all four of the calculated structures. Initial co-ordination of CO_2 leads to a lowering in energy for all four cases to give the Int. 3 structures and relatively low barriers are found for formation of a C–O bond between the C atom of the carbon dioxide and the oxygen atom originating

from the epoxide (the TS-2 structure set). This addition of CO_2 takes place *via* insertion of the molecule into the O–Zn bond with the simultaneous co-ordination of a CO_2 oxygen atom to the Zn centre, so that the transition state is stabilised by interaction with the Lewis acid catalyst. The resulting structures contain a carbonate group which bridges between the cyclopentane and the Zn centre (Int. 4t and Int. 4b). Finally ring closure of the cyclic carbonate takes place and the Br^- anion returns to the $[ZrBr_5]^-$ cluster as a leaving group on the formation of the second C–O bond in the cyclic carbonate product. For this step, there is the additional choice of the oxygen atom that forms the second C–O bond in the cyclic carbonate product. From any of the Int. 4 structures we can envisage the O atom endocyclic in the metallocycle of the intermediate, *i.e.* co-ordinated to Zn, shifting to bond to the carbon atom. Alternatively a rotation of the C–O bond in the intermediate can bring the *exo*-cyclic oxygen atom into position for ring closure. This choice is recognised in Fig. 1 by the addition of /e or /i to the TS3 transition state structure labels for *exo* and *endo*-cyclic oxygen, respectively. The use of the *exo*-cyclic oxygen maintains co-ordination to the Zn Lewis acid centre in the transition state and, if this takes place from Int 4t the lowest energy transition state is found, at -53 kJ mol^{-1} (TS-3t'/e, Fig. 2c) relative to the starting point of the potential energy diagram in Fig. 1. The first endocyclic transition state is only marginally higher at -49 kJ mol^{-1} but this is arrived at following the S_{N1} route which has a high initial epoxide ring opening barrier (S_{N1} TS-1b'). Alternatives using the Zn co-ordinating oxygen for ring closure and/or the bridge site Br^- ion lead to significantly higher barriers which would imply that these pathways are not kinetically relevant.

For this cyclopentene oxide example the low energy route through the potential energy diagram consists of an S_{N2} ring opening step with a terminal Br^- ion and an exocyclic oxygen ring closure. This leads to a *cis*-arrangement of the cyclic carbonate and alkyl ring structures in the product. The *trans*-structure is only found for pathways containing high barriers to reaction.

Experimental results

Cyclic carbonate synthesis from epoxide and carbon dioxide has been an active field of research for a number of years.^{7,12–15} Here we investigate the effect of various reaction parameters such as reaction temperature and reaction time on the cycloaddition of carbon dioxide to cycloalkene oxides using a simple Zn based Lewis acid catalyst. These results will determine the most appropriate conditions for high yields of cyclic carbonate.

Table 1 summarises the experimental results obtained in this study using cyclohexene oxide as substrate. In the absence of Bu_4NBr and $ZnBr_2$ no reaction takes place at 80°C and only a 2% conversion with no observable cyclic carbonate product is seen at 125°C (Table 1, entries 1 and 2), confirming that the catalyst is required to observe carbonate formation. When the catalyst is present a conversion of 51% with 85% selectivity



Table 1 Cycloaddition of CO₂ to cyclohexene oxide

Entry	Catalyst	Temp./°C	Time/h	Conv./%	CC sel./%
1	Blank	80	4	0	0
2	Blank	125	4	2	0
3	Bu ₄ NBr + ZnBr ₂	90	4	51	85
4	Bu ₄ NBr + ZnBr ₂	125	4	80	83
5	Bu ₄ NBr + ZnBr ₂	125	8	89	81
6	Bu ₄ NBr + ZnBr ₂	125	16	98	80

Reaction conditions: 5 ml (49.4 mmol) cyclohexene oxide, 0.4 g Bu₄NBr, 0.16 g ZnBr₂, 20 bar CO₂.

is found at a temperature of 90 °C. Increasing the temperature to 125 °C leads to a conversion of 80%, while the selectivity decreases slightly to 83% (Table 1, entries 3 and 4). By allowing longer reaction times conversions up to 98% can be obtained after 16 h (Table 1, entries 5 and 6).

However, at such long times, the selectivity to cyclic carbonate dropped to 80% and the formation of small amounts of other by-products were identified by GC-MS. These products were for example 1,2-cyclohexanediol and 2-cyclohexene-1-ol. The spectroscopic data can be found in the ESI.† These products were observed as impurities in the starting materials, however at significantly lower quantities than those observed in a reaction run.

The overall goal is the direct synthesis of cyclic carbonate from cycloalkenes. We envisage a two stage process in which the alkene is first oxidised to the epoxide and then the cyclic carbonate is formed by reaction with CO₂. These initial reactions show how the cycloaddition step can be optimised for cyclohexene oxide. However, earlier work in our group has shown that epoxidation of cyclohexene is not very selective to cyclohexene oxide as the allylic ketone or alcohol is the preferred product using atmospheric oxygen with a radical initiator over a supported Au or Au-Pd catalyst.¹⁴ High epoxide selectivities were obtained when larger ring sizes, such as cyclooctene, were used as substrates in the oxidation step. Therefore, the cycloaddition of CO₂ to cyclooctene oxide was investigated next.

Reaction of cyclooctene oxide

Following a similar approach to that employed for cyclohexene oxide we attempted to study the effects of reaction temperature

and reaction time on the cycloaddition of CO₂ to cyclooctene oxide (Table 2). In this case no cyclooctene carbonate was observed under any of the conditions used and no consumption of CO₂ during the reactions was observed. The conversion of cyclooctene oxide requires higher temperatures when compared to cyclohexene oxide. After 4 h of reaction at 90 °C no conversion was observed (Table 2, entry 1). A conversion of 3% is seen at 130 °C after 4 h reaction time and at this temperature the reaction is practically completed after 48 h (Table 2, entries 2–4). Product identification by GC-MS indicated that cyclooctanone, cyclooctanol, and 1,2-cyclooctanediol are the major products (see ESI†). Furthermore, trace amounts of tributylamine, a decomposition product of the quaternary amine catalyst, were found, which may indicate reaction with the Bu₄N⁺ cation has taken place. The availability of a reference standard for cyclooctanone allowed us to quantify this product routinely and under long reaction times selectivity to the cyclooctanone is high. The reaction proceeds at a greater rate at 150 °C (Table 2, entries 5–8). At both 130 °C and 150 °C the selectivity to the cyclooctanone product at the end of the reaction maximises at 61%. Lewis acids are known to catalyse the conversion of epoxides to ketones following the Meinwald rearrangement mechanism. DFT calculations suggest that this can take place through an H shift in intermediates such as Int 2 (Fig. 1).⁴¹ Our results suggest that, under the reaction conditions used here, this H shift is considerably faster than CO₂ insertion for the cyclooctene oxide substrate.

Cycloaddition of CO₂ to different cycloalkene oxides

So far, cycloaddition of CO₂ to cyclohexene oxide and cyclooctene oxide have been studied and it was found that smaller

Table 2 Cycloaddition of CO₂ to cyclooctene oxide

Entry	Temp./°C	Time/h	Conv./%	Selectivity/%		
				Cyclooctene carbonate	Ketone	1,2-Diol
1	90	4	0	0	0	—
2	130	4	3	0	20	59
3		24	67	0	58	23
4		48	99	0	61	—
5	150	4	10	0	29	—
6		8	22	0	36	—
7		16	79	0	62	—
8		24	98	0	61	—

Reaction conditions: 5 g (39.6 mmol) cyclooctene oxide, 0.4 g Bu₄NBr, 0.16 g ZnBr₂, 20 bar CO₂; (—) no GC-MS measurement made. Ketone = cyclooctanone, 1,2-diol = 1,2-cyclooctanediol.



Table 3 Cycloaddition of CO₂ to different cycloalkene oxides

Entry	Substrate	Temp./°C	Time/h	Conv./%	CC sel./%
1	Cyclopentene oxide	90	4	58	91
2	Cyclohexene oxide	125	4	80	83
3 ^a	Cyclooctene oxide	150	8	22	0
4	Cyclododecene oxide	150	8	13	0
5	Cyclododecene oxide	150	16	48	0

Reaction conditions: 5 ml (26.0–57.3 mmol) cycloalkene oxide, 0.4 g Bu₄NBr, 0.16 g ZnBr₂, 20 bar CO₂. ^a 5 g (39.6 mmol).

ring size of the cycloalkene oxide convert easily to the corresponding cyclic carbonate. In order to confirm the observation of the effect of the ring size of cycloalkene oxide in this reaction, a range of different cycloalkene oxides have been studied (Table 3). As expected, smaller ring sizes such as cyclopentene oxide show high selectivity to cyclic carbonates (91%) at 90 °C after 4 h reaction time (Table 3, entry 1). Only small amounts of by-products were detected by GC-MS such as cyclopentanone, 1,2-cyclopentanediol and 2-bromocyclopentanol which may result from conversion of intermediates such as **Int 2** in Fig. 1 to alcohols before the Br[−] anion is returned to the catalyst. For the spectroscopic data please see the ESL.† In the case of the larger ring sizes, cyclododecene oxide, even with increasing the reaction time from 8 to 16 h, showed no cyclic carbonate product (Table 3, entries 4 and 5). Cyclododecanone was found to be the main product for this reaction (selectivity greater than 95%). This is in line with the cyclooctene oxide results. In general, it can be said that with increasing ring size of the cycloalkene oxide the selectivity towards the corresponding cyclic carbonate is decreasing and with ring sizes larger than C8 there is no formation of the cyclic carbonate at all.

Conclusions

We have studied the mechanism of carboxylation of cyclopentene using DFT methods which indicate that the Zn Lewis acid centres in the Bu₄NBr and ZnBr₂ catalyst system play an important role in stabilising the intermediates and transition states. The rate-determining step appears to be the ring opening of the epoxide in this case with a concerted ring opening and bromination from a terminal Br[−] anion in the [ZnBr₅][−] catalyst. Insertion of CO₂ into this intermediate is relatively straight forward with the second significant barrier for ring closure of the cyclic carbonate taking place *via* an *exo*-cyclic oxygen mechanism.

Experimentally, the carboxylation process for a range of different cycloalkene oxides has been studied and it has been found that more forcing conditions are required for the larger ring sizes with the consequence that ketone and diol side products become dominant for the ring sizes of cyclooctene oxide and above. Reference to the DFT derived mechanism suggests that the ketone product can be formed by an H shift in the first intermediate formed after epoxide ring opening. The trend with ring size for selectivity to carbonate is the opposite to that observed in our earlier work on the ox-

idation of cycloalkenes in which epoxide selectivity increased with ring size.¹⁴ This suggests that the direct conversion of cycloalkenes to carbonates is probably only possible for cycloalkenes with a ring size below 8 using this Lewis acid catalyst. Although currently the temperatures required for the two steps (epoxidation: 30–70 °C,¹⁴ carboxylation: 90–125 °C, Table 3) make these stages of the reaction incompatible for a one-pot process. Progress may be made by identifying alternative Lewis acid catalysts for the carboxylation step or identifying reaction conditions which favour carboxylation over re-arrangement of the intermediate to form ketones.

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Notes and references

- Q. Liu, L. Wu, R. Jackstell and M. Beller, *Nat. Commun.*, 2015, 6, 5933.
- S. Fukuoka, M. Kawamura, K. Komiya, M. Tojo, H. Hachiya, K. Hasegawa, M. Aminaka, H. Okamoto, I. Fukawa and S. Konno, *Green Chem.*, 2003, 5, 497–507.
- A.-A. G. Shaikh and S. Sivaram, *Chem. Rev.*, 1996, 96, 951–976.
- J. H. Clements, *Ind. Eng. Chem. Res.*, 2003, 42, 663–674.
- T. Sakakura and K. Kohno, *Chem. Commun.*, 2009, 1312–1330.
- A. Khan, L. Yang, J. Xu, L. Y. Jin and Y. J. Zhang, *Angew. Chem., Int. Ed.*, 2014, 53, 11257–11260.
- M. North, R. Pasquale and C. Young, *Green Chem.*, 2010, 12, 1514–1539.
- J. W. Comerford, I. D. V. Ingram, M. North and X. Wu, *Green Chem.*, 2015, 17, 1966–1987.
- C. Martín, G. Fiorani and A. W. Kleij, *ACS Catal.*, 2015, 5, 1353–1370.



10 V. B. Saptal and B. M. Bhanage, *Current Opinion in Green and Sustainable Chemistry*, 2017, **3**, 1–10.

11 S.-I. Fujita, B. M. Bhanage, Y. Ikushima, M. Shirai, K. Torii and M. Arai, *Catal. Lett.*, 2002, **79**, 95–98.

12 J. Sun, S.-I. Fujita, F. Zhao and M. Arai, *Appl. Catal., A*, 2005, **287**, 221–226.

13 J.-Q. Wang, D.-L. Kong, J.-Y. Chen, F. Cai and L.-N. He, *J. Mol. Catal. A: Chem.*, 2006, **249**, 143–148.

14 J. Sun, J. Ren, S. Zhang and W. Cheng, *Tetrahedron Lett.*, 2009, **50**, 423–426.

15 B. Song, L. Guo, R. Zhang, X. Zhao, H. Gan, C. Chen, J. Chen, W. Zhu and Z. Hou, *J. CO₂ Util.*, 2014, **6**, 62–68.

16 V. Caló, A. Nacci, A. Monopoli and A. Fanizzi, *Org. Lett.*, 2002, **4**, 2561–2563.

17 S. Foltran, R. Mereau and T. Tassaing, *Catal. Sci. Technol.*, 2014, **4**, 1585–1597.

18 J.-Q. Wang, K. Dong, W.-G. Cheng, J. Sun and S.-J. Zhang, *Catal. Sci. Technol.*, 2012, **2**, 1480–1484.

19 T. Ema, Y. Miyazaki, J. Shimonishi, C. Maeda and J.-Y. Hasegawa, *J. Am. Chem. Soc.*, 2014, **136**, 15270–15279.

20 A.-L. Girard, N. Simon, M. Zanatta, S. Marmitt, P. Goncalves and J. Dupont, *Green Chem.*, 2014, **16**, 2815–2825.

21 K. R. Roshan, A. C. Kathalikkattil, J. Tharun, D. W. Kim, Y. S. Won and D. W. Park, *Dalton Trans.*, 2014, **43**, 2023–2031.

22 D. Adhikari, S. T. Nguyen and M.-H. Baik, *Chem. Commun.*, 2014, **50**, 2676–2678.

23 T. Ema, K. Fukuwara, T. Sakai, M. Ohbo, F.-Q. Bai and J.-Y. Hasegawa, *Catal. Sci. Technol.*, 2015, **5**, 2314–2321.

24 F. Chen, X. Li, B. Wang, T. Xu, S.-L. Chen, P. Liu and C. Hu, *Chem. – Eur. J.*, 2012, **18**, 9870–9876.

25 J. Sun, S.-I. Fujita, F. Zhao, M. Hasegawa and M. Arai, *J. Catal.*, 2005, **230**, 398–405.

26 J. A. Verdol, *U.S. Pat.*, 3025305, 1962.

27 H. Alshammari, P. J. Miedziak, D. W. Knight, D. J. Willock and G. J. Hutchings, *Catal. Sci. Technol.*, 2013, **3**, 1531–1539.

28 G. W. T. M. J. Frisch, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision E.01*, Gaussian, Inc., Wallingford CT, 2009.

29 C. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Phys.*, 1988, **37**, 785–789.

30 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648–5652.

31 P. Stephens, F. Devlin, C. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, **98**, 11623–11627.

32 S. H. Vosko, L. Wilk and M. Nusair, *Can. J. Phys.*, 1980, **58**, 1200–1211.

33 W. J. Hehre, R. Ditchfield and J. A. Pople, *J. Chem. Phys.*, 1972, **56**, 2257–2261.

34 R. Ditchfield, W. J. Hehre and J. A. Pople, *J. Chem. Phys.*, 1971, **54**, 724–728.

35 P. C. Hariharan and J. A. Pople, *Theor. Chim. Acta*, 1973, **28**, 213–222.

36 P. C. Hariharan and J. A. Pople, *Mol. Phys.*, 1974, **27**, 209–214.

37 M. S. Gordon, *Chem. Phys. Lett.*, 1980, **76**, 163–168.

38 M. M. Franci, W. J. Pietro, W. J. Hehre, J. S. Binkley, M. S. Gordon, D. J. DeFrees and J. A. Pople, *J. Chem. Phys.*, 1982, **77**, 3654–3665.

39 D. J. Darensbourg and M. W. Holtcamp, *Coord. Chem. Rev.*, 1996, **153**, 155–174.

40 M. North and R. Pasquale, *Angew. Chem., Int. Ed.*, 2009, **48**, 2946–2948.

41 J. M. A. Fraile, J. A. Mayoral and L. Salvatella, *J. Org. Chem.*, 2014, **79**, 5993–5999.