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# Exploring the effect of fluorinated anions on the CO<sub>2</sub>/N<sub>2</sub> separation of supported ionic liquid membranes†

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The CO<sub>2</sub> and N<sub>2</sub> permeation properties of ionic liquids (ILs) based on the 1-ethyl-3-methylimidazolium cation ([C<sub>2</sub>mim]<sup>+</sup>) and different fluorinated anions, namely 2,2,2-trifluoromethylsulfonyl-*N*-cyanoamide ([TFSAM]<sup>−</sup>), bis(fluorosulfonyl) imide ([FSI]<sup>−</sup>), nonafluorobutanesulfonate ([C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]<sup>−</sup>), tris(pentafluoroethyl)trifluorophosphate ([FAP]<sup>−</sup>), and bis(pentafluoroethylsulfonyl)imide ([BETI]<sup>−</sup>) anions, were measured using supported ionic liquid membranes (SILMs). The results show that pure ILs containing [TFSAM]<sup>−</sup> and [FSI]<sup>−</sup> anions present the highest CO<sub>2</sub> permeabilities, 753 and 843 Barrer, as well as the greatest CO<sub>2</sub>/N<sub>2</sub> permselectivities of 43.9 and 46.1, respectively, with CO<sub>2</sub>/N<sub>2</sub> separation performances on top of or above the Robeson 2008 upper bound. The re-design of the [TFSAM]<sup>−</sup> anion by structural unfolding was investigated through the use of IL mixtures. The gas transport and CO<sub>2</sub>/N<sub>2</sub> separation properties through a pure [C<sub>2</sub>mim][TFSAM] SILM are compared to those of two different binary IL mixtures containing fluorinated and cyano-functionalized groups in the anions. Although the use of IL mixtures is a promising strategy to tailor gas permeation through SILMs, the pure [C<sub>2</sub>mim][TFSAM] SILM displays higher CO<sub>2</sub> permeability, diffusivity and solubility than the selected IL mixtures. Nevertheless, both the prepared mixtures present CO<sub>2</sub> separation performances that are on top of or above the Robeson plot.

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## Introduction

The development of supported ionic liquid membranes (SILMs) for CO<sub>2</sub> separation has been widely investigated in recent years mainly due to their easy preparation and versatility.<sup>1–3</sup> In contrast to traditional liquid membranes, which are produced by impregnating a porous membrane support with common organic solvents,<sup>4</sup> SILMs use ionic liquids (ILs) and thus benefit from negligible displacement of the liquid phase from the membrane pores through evaporation,<sup>5,6</sup> due to the low volatility of ILs.<sup>7</sup> It should also be emphasized that within the CO<sub>2</sub> separation context, the most important features of ILs are their high CO<sub>2</sub> affinity over light gases<sup>8–10</sup> and their inherent designer nature that enables the tailoring of IL properties by proper selection

of cations and/or anions or *via* the addition of specific functional groups.

Numerous works have investigated the effect of IL chemical structure on the gas permeation properties of SILMs. A broad diversity of cations, such as imidazolium,<sup>11</sup> triazolium,<sup>12</sup> thiazolium,<sup>13</sup> pyridinium,<sup>14</sup> cholinium,<sup>15</sup> ammonium,<sup>16</sup> and phosphonium,<sup>17</sup> combined with halogens, sulfonates, carboxylates, fluorinated or cyano-functionalized anions, have been studied. Other works, mostly focusing on imidazolium-based ILs, have also explored the effect of alkyl,<sup>18</sup> fluoroalkyl,<sup>19</sup> etoxyalkyl,<sup>20</sup> and aminoalkyl<sup>21</sup>-functionalized cations. Since IL anions have a stronger influence on the CO<sub>2</sub> separation performance of SILMs than IL cations,<sup>1</sup> they deserved from the start a closer look. The first studies on SILMs made use of fluorinated anions such as bis(trifluoromethylsulfonyl)imide [NTf<sub>2</sub>]<sup>−</sup>, tetrafluoroborate [BF<sub>4</sub>]<sup>−</sup>, and hexafluorophosphate [PF<sub>6</sub>]<sup>−</sup> and enabled drawing conclusions about the CO<sub>2</sub>-phyllic behaviour and high CO<sub>2</sub> permeabilities of these anions.<sup>22</sup> More recently, low viscous ILs with cyano-functionalized anions, such as tricyanomethanide [C(CN)<sub>3</sub>]<sup>−</sup> and tetracyanoborate [B(CN)<sub>4</sub>]<sup>−</sup>,<sup>23–25</sup> have been recognized as better candidates for the development of improved SILMs, because of their superior CO<sub>2</sub> permeabilities and permselectivities when compared to the most used [NTf<sub>2</sub>]<sup>−</sup> anion. Task-specific ILs bearing amine groups, such as those containing amino acid

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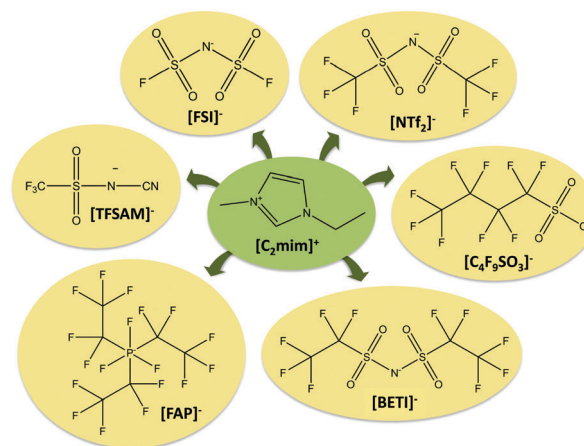
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Additionally, this work investigates the impact on gas transport through SILMs of using a pure IL *versus* a structurally similar IL mixture as the liquid phase. Inspired by the fact that the [TFSAM]<sup>−</sup> anion has an unusual asymmetric chemical structure, which combines both fluorinated and cyano functionalities, the re-design of the chemical structure of a pure [C<sub>2</sub>mim][TFSAM] IL through the use of IL mixtures is explored. For that purpose, different pairs of ILs, based on the [C<sub>2</sub>mim]<sup>+</sup> cation and anions containing fluorinated or cyano functionalities, were selected and

The experimental gas permeability ( $P$ ) values obtained through the prepared SILMs having ILs with fluorinated anions, measured at 293 K with a *trans*-membrane pressure differential of 100 kPa, are shown in Table 2. To the best of our knowledge, the CO<sub>2</sub>/N<sub>2</sub> separation properties of [C<sub>2</sub>mim][TFSAM], [C<sub>2</sub>mim][FSI], [C<sub>2</sub>mim][FAP], [C<sub>2</sub>mim][BETI] and [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] SILMs are reported here for the first time, while those of the [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM were previously determined using the same experimental conditions.<sup>29</sup> It is important to mention that in order to attain stable SILMs, both hydrophilic and hydrophobic supports were used according to the hydrophobicity of ILs, and the results are compared in this section, irrespective of the support used.



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**Table 1** Water contents and thermophysical properties of the pure ILs

IL sample	H <sub>2</sub> O (wt%)	<i>M</i> (g mol <sup>-1</sup> )	$\eta^a$ (mPa s)	$\rho^b$ (g cm <sup>-3</sup> )	<i>V<sub>m</sub></i> (cm <sup>3</sup> mol <sup>-1</sup> )
[C <sub>2</sub> mim][TFSAM]	0.02	284.26	23.7	1.3518	210.3
[C <sub>2</sub> mim][FSI]	0.09	291.30	22.4	1.4480	201.2
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>c</sup>	0.02	391.31	39.1	1.5240	256.8
[C <sub>2</sub> mim][FAP] <sup>d</sup>	—	556.17	76.4	1.7081	325.6
[C <sub>2</sub> mim][BETI]	0.02	491.33	85.5	1.5989	307.3
[C <sub>2</sub> mim][C <sub>4</sub> F <sub>9</sub> SO <sub>3</sub> ]	0.08	410.26	109.7	1.5420	267.1

<sup>a</sup> Viscosity ( $\eta$ ) and density ( $\rho$ ) measured at 293 K, except for [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] IL (303 K). <sup>b</sup> Molar volume (*V<sub>m</sub>*) obtained for 293 K, except for [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] IL (303 K). <sup>c</sup> The values of [C<sub>2</sub>mim][NTf<sub>2</sub>] were taken from Tomé *et al.*<sup>29</sup> <sup>d</sup> The density and viscosity values of [C<sub>2</sub>mim][FAP] were taken from Neves *et al.*<sup>52</sup>

**Table 2** Gas permeabilities (*P*)<sup>a</sup> through the prepared SILMs of pure ILs<sup>b</sup>

SILM sample	<i>P</i> (CO <sub>2</sub> )	<i>P</i> (N <sub>2</sub> )	$\alpha$ (CO <sub>2</sub> /N <sub>2</sub> )
[C <sub>2</sub> mim][TFSAM]	753 ± 0.2	17 ± 0.1	43.9 ± 0.1
[C <sub>2</sub> mim][FSI]	843 ± 0.5	18 ± 0.2	46.1 ± 0.5
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>c</sup>	589 ± 1.0	17 ± 0.1	35.5 ± 0.3
[C <sub>2</sub> mim][FAP]	624 ± 0.4	24 ± 0.1	26.0 ± 0.1
[C <sub>2</sub> mim][BETI]	437 ± 1.9	18 ± 0.1	24.8 ± 0.2
[C <sub>2</sub> mim][C <sub>4</sub> F <sub>9</sub> SO <sub>3</sub> ]	32 ± 0.2	6 ± 0.1	5.5 ± 0.1

<sup>a</sup> Barrer (1 Barrer = 10<sup>-10</sup> cm<sub>(STP)</sub><sup>3</sup> cm cm<sup>-2</sup> s<sup>-1</sup> cmHg<sup>-1</sup>). <sup>b</sup> The listed uncertainties represent the standard deviations, based on three experiments. <sup>c</sup> Values taken from Tomé *et al.*<sup>29</sup>

From Table 2, the same trend in gas permeability is valid for all the studied SILMs: *PCO*<sub>2</sub> >> *PN*<sub>2</sub>, as expected. Regarding the influence of the fluorinated-based anions, SILMs having the [FSI]<sup>-</sup>, [TFSAM]<sup>-</sup> and [FAP]<sup>-</sup> anions present higher CO<sub>2</sub> permeabilities of 843, 753 and 624 Barrer, respectively, than the SILM containing the [NTf<sub>2</sub>]<sup>-</sup> anion, which is well-known for its high CO<sub>2</sub> permeability (589 Barrer).<sup>29</sup> It should be noted that in spite of the similar structures of [NTf<sub>2</sub>]<sup>-</sup> and [FSI]<sup>-</sup> anions, in which the difference consists in two extra <sup>-</sup>CF<sub>3</sub> groups in the [NTf<sub>2</sub>]<sup>-</sup> anion structure (Fig. 1), the CO<sub>2</sub> permeability through the [C<sub>2</sub>mim][FSI] SILM is ~1.5 times higher than that through the [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM. Generally, CO<sub>2</sub> permeabilities through the studied SILMs are found to decrease in the following IL anion order: [FSI]<sup>-</sup> > [TFSAM]<sup>-</sup> > [FAP]<sup>-</sup> > [NTf<sub>2</sub>]<sup>-</sup> > [BETI]<sup>-</sup> > [C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]<sup>-</sup> (Table 2). Considering the IL anion viscosity trend, obtained at 293 K (Table 1), [FSI]<sup>-</sup> < [TFSAM]<sup>-</sup> < [NTf<sub>2</sub>]<sup>-</sup> < [FAP]<sup>-</sup> < [BETI]<sup>-</sup> < [C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]<sup>-</sup>, it can be concluded that these experimental data are in agreement with the general trend usually observed in the literature, where ILs with high viscosities yield SILMs with low gas permeabilities.<sup>3,14,15,23,40</sup> However, [C<sub>2</sub>mim][FAP] is the only exception since it presents a different behaviour (Table 2): despite its high viscosity (76.4 mPa s), it also exhibits high CO<sub>2</sub> permeability (624 Barrer), higher than those of the [C<sub>2</sub>mim][NTf<sub>2</sub>] IL (39.1 mPa s and 589 Barrer). Notice that the [FAP]<sup>-</sup> anion has the most different chemical structure among all the IL anions studied in this work, consisting of a phosphorus atom surrounded by fluorine atoms, without sulfonyl functional groups (Fig. 1). Moreover, taking a closer look at the gas permeabilities obtained through SILMs immobilized with the remaining ILs, it can be seen that

higher CO<sub>2</sub> permeabilities are achieved for ILs with anions bearing a smaller number of fluorine elements, such as [TFSAM]<sup>-</sup> and [FSI]<sup>-</sup> anions (Table 2).

Gas diffusivity (*D*) is a mass transfer property that directly accounts for gas permeability (eqn (1)). Typically, the higher the gas diffusivity, the faster is the gas flux through the SILM. The experimental CO<sub>2</sub> and N<sub>2</sub> diffusivity values obtained through the prepared SILMs are presented in Table 3. The CO<sub>2</sub> diffusivities of the SILMs with fluorinated anions can be ordered as follows: [FSI]<sup>-</sup> > [TFSAM]<sup>-</sup> > [FAP]<sup>-</sup> > [NTf<sub>2</sub>]<sup>-</sup> > [BETI]<sup>-</sup> > [C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]<sup>-</sup>, which fully corresponds to the IL anion order observed for CO<sub>2</sub> permeabilities (Table 2). As for N<sub>2</sub> diffusivities the subsequent order is attained: ([FAP]<sup>-</sup> > [FSI]<sup>-</sup> > [TFSAM]<sup>-</sup> > [BETI]<sup>-</sup> > [NTf<sub>2</sub>]<sup>-</sup> > [C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]<sup>-</sup>), which is nearly the same IL anion order observed for N<sub>2</sub> permeabilities, but different from that obtained for CO<sub>2</sub> diffusivities and permeabilities.

The relationship between IL viscosity and gas diffusion is in fact the basis of the dependence of permeability on viscosity that we have shown above. A number of works have reported the inversely proportional relationship between gas diffusivity and IL viscosity.<sup>3,16,17,41</sup> Along this line of thought, the relationship between experimental CO<sub>2</sub> diffusivities and IL viscosity for the studied SILMs is depicted in Fig. 2.

In agreement to what was previously observed in the literature for other SILMs,<sup>4,21–23</sup> the CO<sub>2</sub> diffusivity through SILMs having fluorinated anions decreases as the IL viscosity increases. The SILMs with the lowest CO<sub>2</sub> diffusivities are [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] (55 × 10<sup>-12</sup> m<sup>2</sup> s<sup>-1</sup>) and [C<sub>2</sub>mim][BETI] (167 × 10<sup>-12</sup> m<sup>2</sup> s<sup>-1</sup>), which also have the lowest CO<sub>2</sub> permeabilities (437 and 32 Barrer), depicting the highest viscosities (109.7 and 85.5 mPa). Similarly to what is mentioned above for CO<sub>2</sub> permeabilities, again a deviant behaviour can be observed for the [C<sub>2</sub>mim][FAP] SILM, since its CO<sub>2</sub> diffusivity is in between those of [C<sub>2</sub>mim][NTf<sub>2</sub>] and [C<sub>2</sub>mim][TFSAM] (Table 3), but its viscosity values (76.4 mPa s) are higher than those of [C<sub>2</sub>mim][NTf<sub>2</sub>] (39.1 mPa s) and [C<sub>2</sub>mim][TFSAM] (23.7 mPa s) (Fig. 2).

The gas solubility (*S*) values calculated using Eqn 1 are listed in Table 4. It can be seen that the [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] SILM presents the lowest CO<sub>2</sub> solubility (4 × 10<sup>-6</sup> m<sub>(STP)</sub><sup>3</sup> m<sup>-3</sup> Pa<sup>-1</sup>), while the previously reported [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM has the highest CO<sub>2</sub> solubility (26 × 10<sup>-6</sup> m<sub>(STP)</sub><sup>3</sup> m<sup>-3</sup> Pa<sup>-1</sup>). The CO<sub>2</sub> solubility values of the remaining studied SILMs having different fluorinated anions are very similar (ranging from 20 × 10<sup>-6</sup>

**Table 3** Gas diffusivity (*D*) values through the prepared SILMs of pure ILs

SILM sample	Gas diffusivity (×10 <sup>12</sup> ) (m <sup>2</sup> s <sup>-1</sup> )	
	<i>D</i> (CO <sub>2</sub> )	<i>D</i> (N <sub>2</sub> )
[C <sub>2</sub> mim][TFSAM]	258 ± 1.1	291 ± 4.5
[C <sub>2</sub> mim][FSI]	315 ± 4.4	353 ± 7.5
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>a</sup>	172 ± 1.7	203 ± 2.5
[C <sub>2</sub> mim][FAP]	214 ± 1.9	326 ± 0.2
[C <sub>2</sub> mim][BETI]	167 ± 2.1	216 ± 2.4
[C <sub>2</sub> mim][C <sub>4</sub> F <sub>9</sub> SO <sub>3</sub> ]	55 ± 1.2	88 ± 2.0

<sup>a</sup> The gas diffusivity values through the [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM were taken from Tomé *et al.*<sup>29</sup>



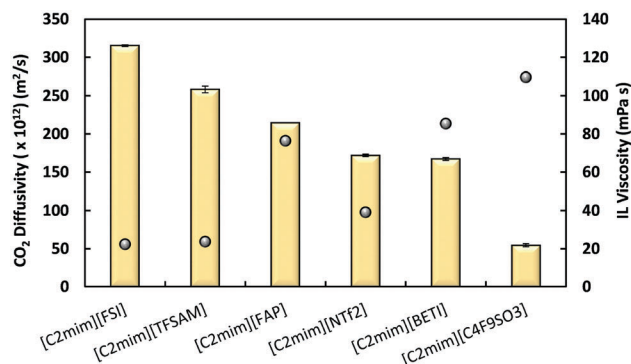


Fig. 2 Relationship between the CO<sub>2</sub> diffusivities determined through the SILMs prepared with pure ILs and the respective IL viscosities measured at 293 K. Error bars represent standard deviations based on three experimental replicas.

to  $22 \times 10^{-6} \text{ m}_{(\text{STP})}^3 \text{ m}^{-3} \text{ Pa}^{-1}$ ) and can be ordered as follows:  $[\text{TFSAM}]^- \approx [\text{FAP}]^- > [\text{FSI}]^- \approx [\text{BETI}]^-$ . Regarding the N<sub>2</sub> solubility, its values are always significantly lower (between  $0.39 \times 10^{-6}$  and  $0.61 \times 10^{-6} \text{ m}_{(\text{STP})}^3 \text{ m}^{-3} \text{ Pa}^{-1}$ ) than those of CO<sub>2</sub> for all the SILMs studied.

Over the past few years, a number of correlations have been proposed with the intention of understanding the relationships between CO<sub>2</sub> solubility and the intrinsic properties of ILs.<sup>42–44</sup> The proposed models showed that CO<sub>2</sub> solubility increases with increasing IL molecular weight, molar volume and free volume.<sup>10</sup> Taking into consideration the gas solubility results obtained in this work (Table 4), as well as the range of molar volumes (from 201.2 up to 325.6 cm<sup>3</sup> mol<sup>−1</sup>) and molecular weights (from 284.3 up to 556.2 g mol<sup>−1</sup>) of the ILs used, divergences from the abovementioned trends can be found for the studied SILMs having fluorinated anions. For example, the [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM shows the highest CO<sub>2</sub> solubility ( $26 \times 10^{-6} \text{ m}_{(\text{STP})}^3 \text{ m}^{-3} \text{ Pa}^{-1}$ ), but it does not have the highest IL molar volume and molecular weight (Table 1). Likewise, the lowest CO<sub>2</sub> solubility ( $4 \times 10^{-6} \text{ m}_{(\text{STP})}^3 \text{ m}^{-3} \text{ Pa}^{-1}$ ) belongs to the [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] SILM, although it does not present the lowest IL molar volume and molecular weight (Table 1). The effect of fluorination, either in the IL cation or anion, on CO<sub>2</sub> solubility has been studied by different researchers.<sup>19,45–47</sup> Tagiuri *et al.*<sup>48</sup> explored the effect of cation on the CO<sub>2</sub> solubility of three different ILs combining the [FSI]<sup>−</sup> anion. Moreover, Kroon *et al.*<sup>49</sup> determined the CO<sub>2</sub> solubility in the [C<sub>2</sub>mim][FAP] IL by measuring the bubble point

pressures of the binary mixture of [C<sub>2</sub>mim][FAP] + CO<sub>2</sub>. The results showed that the CO<sub>2</sub> solubility in [C<sub>2</sub>mim][FAP] is higher when compared to that of ILs having the same cation combined with other fluorinated anions such as [NTf<sub>2</sub>]<sup>−</sup>, [BF<sub>4</sub>]<sup>−</sup> and [PF<sub>6</sub>]<sup>−</sup>, due to the fact that [FAP]<sup>−</sup> has a large size and it is highly fluorinated. Although it has been recognized that introducing fluorination into the cation and/or anion can effectively improve CO<sub>2</sub> solubility,<sup>46</sup> it was recently reported after critical analysis that no special effect of the fluorination upon the CO<sub>2</sub> solubility has been observed for both perfluorocarbon and heavily fluorinated ILs.<sup>50</sup> In fact, the introduction of fluorination into the anions of the ILs studied in this work does not significantly affect the obtained gas solubility values (Table 4), except for the case of the [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] IL that displays a very low CO<sub>2</sub> solubility.

### Re-designing the [TFSAM]<sup>−</sup> anion by structural unfolding: effect on gas permeation

Taking into account that the [TFSAM]<sup>−</sup> anion has an unconventional and asymmetric chemical structure, combining both fluorinated and cyano functionalities, which have both been recognized to be responsible for high CO<sub>2</sub> separation performance, we explore here the effect of structural unfolding of the pure [C<sub>2</sub>mim][TFSAM] IL on gas permeation properties of SILMs using IL mixtures. Thus, two equimolar IL mixtures were used for this purpose: [C<sub>2</sub>mim][SCN][OTf] (Fig. 3), which is studied here for the first time, and [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>], whose gas permeation and thermophysical properties were previously determined by us.<sup>29</sup> Both these mixtures have IL anions that show structural similarities to the [TFSAM]<sup>−</sup> anion (Fig. 3). The composition description, water content (wt%), molar mass (*M*), viscosity (*η*), density (*ρ*) and molar volume (*V<sub>m</sub>*) values of the pure ILs, [C<sub>2</sub>mim][TFSAM], [C<sub>2</sub>mim][SCN], [C<sub>2</sub>mim][OTf], [C<sub>2</sub>mim][N(CN)<sub>2</sub>], [C<sub>2</sub>mim][NTf<sub>2</sub>], and the selected IL mixtures are listed in Table 5, while their gas permeability, diffusivity and solubility values are depicted in Fig. 4(a)–(c), respectively.

Fig. 4(a) and (b) show that the pure [C<sub>2</sub>mim][TFSAM] SILM exhibits higher gas permeabilities and diffusivities than SILMs

Table 4 Gas solubility (*S*) values through the prepared SILMs of pure ILs

SILM sample	Gas solubility ( $\times 10^6$ ) ( $\text{m}_{(\text{STP})}^3 \text{ m}^{-3} \text{ Pa}^{-1}$ )	
	<i>S</i> (CO <sub>2</sub> )	<i>S</i> (N <sub>2</sub> )
[C <sub>2</sub> mim][TFSAM]	22 ± 0.09	0.44 ± 0.01
[C <sub>2</sub> mim][FSI]	20 ± 0.29	0.39 ± 0.01
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>a</sup>	26 ± 0.28	0.61 ± 0.01
[C <sub>2</sub> mim][FAP]	22 ± 0.21	0.55 ± 0.001
[C <sub>2</sub> mim][BETI]	20 ± 0.17	0.61 ± 0.01
[C <sub>2</sub> mim][C <sub>4</sub> F <sub>9</sub> SO <sub>3</sub> ]	4 ± 0.07	0.49 ± 0.002

<sup>a</sup> The gas solubility values through the [C<sub>2</sub>mim][NTf<sub>2</sub>] SILM were taken from Tomé *et al.*<sup>29</sup>

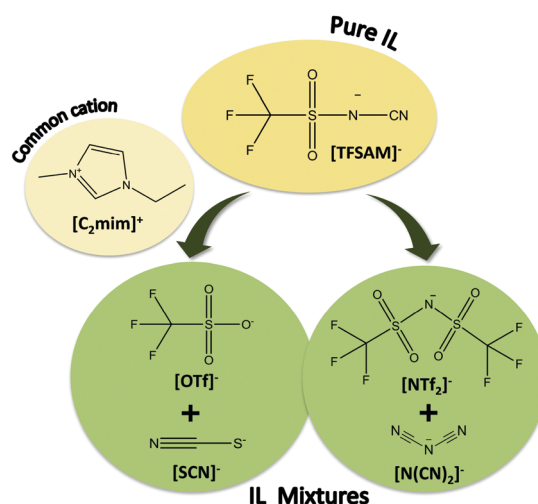


Fig. 3 Chemical structures of the pure IL and IL mixtures studied in this work.





**Table 5** Composition descriptions, water contents and thermophysical properties of the pure ILs and IL mixtures

IL sample	Composition description (mole fraction)	H <sub>2</sub> O (wt%)	<i>M</i> (g mol <sup>-1</sup> )	$\eta^a$ (mPa s)	$\rho^b$ (g cm <sup>-3</sup> )	<i>V<sub>m</sub></i> (cm <sup>3</sup> mol <sup>-1</sup> )
[C <sub>2</sub> mim][TFSAM]	Pure	0.02	284.26	23.7	1.3518	210.3
[C <sub>2</sub> mim][SCN] <sup>c</sup>	Pure	0.09	169.25	27.9	1.1190	151.2
[C <sub>2</sub> mim][SCN][OTf]	0.5 [C <sub>2</sub> mim][SCN] + 0.5 [C <sub>2</sub> mim][OTf]	0.16	214.74	36.4	1.2638	169.9
[C <sub>2</sub> mim][OTf]	Pure	0.10	260.23	49.5	1.3850	187.9
[C <sub>2</sub> mim][N(CN) <sub>2</sub> ] <sup>c</sup>	Pure	0.09	177.21	18.0	1.1060	160.2
[C <sub>2</sub> mim][N(CN) <sub>2</sub> ][NTf <sub>2</sub> ] <sup>c</sup>	0.5[C <sub>2</sub> mim][N(CN) <sub>2</sub> ] + 0.5 [C <sub>2</sub> mim][NTf <sub>2</sub> ]	0.12	284.26	29.2	1.3620	208.7
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>c</sup>	Pure	0.02	391.31	39.1	1.5240	256.8

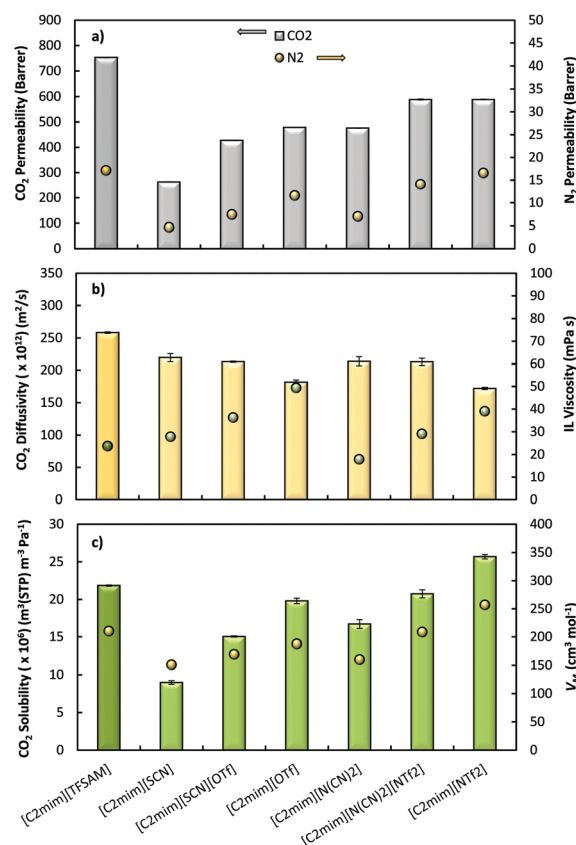
<sup>a</sup> Viscosity ( $\eta$ ) and density ( $\rho$ ) measured at 293 K. <sup>b</sup> Molar volume (*V<sub>m</sub>*) obtained at 293 K. <sup>c</sup> The values of [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] were taken from Tomé *et al.*<sup>29</sup>

composed of both [C<sub>2</sub>mim][SCN][OTf] and [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] IL mixtures. Moreover, the addition of 0.5 mole fraction of [C<sub>2</sub>mim][SCN] did not significantly affect CO<sub>2</sub> permeability but decreased to almost half the N<sub>2</sub> permeability compared to those of the pure [C<sub>2</sub>mim][OTf] SILM. From Fig. 4(b), where the relationship between CO<sub>2</sub> diffusivity and IL viscosity is illustrated, it can also be seen that both [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] and [C<sub>2</sub>mim][SCN][OTf] IL mixtures present slightly higher viscosities (Table 5) and lower CO<sub>2</sub> diffusivities than those of the pure [C<sub>2</sub>mim][TFSAM] SILM. Also, the presence of 0.5 mole fraction of [C<sub>2</sub>mim][SCN] leads to a decrease in IL viscosity and an increase in CO<sub>2</sub> diffusivity. The same behaviour was found for [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] with the addition of 0.5 mole fraction of [C<sub>2</sub>mim][N(CN)<sub>2</sub>]. These results are in accordance with the general trend observed in the literature: the CO<sub>2</sub> diffusivity decreases with the increase in IL viscosity.<sup>4,21–23</sup>

Concerning gas solubility, from Fig. 4(c), it can be observed that CO<sub>2</sub> solubilities of both IL mixtures are in between those of the individual IL components. Moreover, the presence of 0.5 mole fraction of [C<sub>2</sub>mim][OTf] or [C<sub>2</sub>mim][NTf<sub>2</sub>] in the corresponding mixtures leads to an increase in the CO<sub>2</sub> solubilities, probably due to the fact that they present higher molar volume compared to the pure [C<sub>2</sub>mim][SCN] or [C<sub>2</sub>mim][N(CN)<sub>2</sub>], respectively. Furthermore, the [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] mixture displays a CO<sub>2</sub> solubility ( $20 \times 10^{-6}$  m<sub>(STP)</sub><sup>3</sup> m<sup>-3</sup> Pa<sup>-1</sup>) closer to that of the pure [C<sub>2</sub>mim][TFSAM] SILM ( $22 \times 10^{-6}$  m<sub>(STP)</sub><sup>3</sup> m<sup>-3</sup> Pa<sup>-1</sup>), probably due to the fact that these IL phases have very close molar volumes (208.7 and 210.3 cm<sup>3</sup> mol<sup>-1</sup>, respectively) and the same molecular weight (284.3 g mol<sup>-1</sup>). Conversely, the [C<sub>2</sub>mim][SCN][OTf] mixture has lower CO<sub>2</sub> solubility ( $15 \times 10^{-6}$  m<sub>(STP)</sub><sup>3</sup> m<sup>-3</sup> Pa<sup>-1</sup>) than that observed for the pure [C<sub>2</sub>mim][TFSAM] SILM, which is in agreement with the lower molar volume (169.9 cm<sup>3</sup> mol<sup>-1</sup>) and molecular weight (214.7 cm<sup>3</sup> mol<sup>-1</sup>) presented by this IL mixture containing the [SCN]<sup>-</sup> and [OTf]<sup>-</sup> anions. Despite the fact that it is possible to obtain similar viscosities and molar volumes of the pure [C<sub>2</sub>mim][TFSAM] IL by mixing the [C<sub>2</sub>mim][N(CN)<sub>2</sub>] and [C<sub>2</sub>mim][NTf<sub>2</sub>] ILs, improved gas permeabilities, diffusivities and solubilities were obtained through the pure [C<sub>2</sub>mim][TFSAM] SILM.

### Comparison of CO<sub>2</sub>/N<sub>2</sub> separation performance

The gas permeabilities and CO<sub>2</sub>/N<sub>2</sub> permselectivities of all the studied SILMs are listed in Table 6. Amongst the pure SILMs with fluorinated anions, the [C<sub>2</sub>mim][FSI] and [C<sub>2</sub>mim][TFSAM] SILMs not only show the highest gas permeabilities (843 and



**Fig. 4** (a) Gas permeability, (b) CO<sub>2</sub> diffusivity as a function of IL viscosity and (c) CO<sub>2</sub> solubility as a function of IL molar volume for the studied SILMs. Error bars represent standard deviations based on three experimental replicas. The gas permeability, diffusivity and solubility values of [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] and [C<sub>2</sub>mim][NTf<sub>2</sub>] were taken from Tomé *et al.*<sup>29</sup>

753 Barrer, respectively), but also have the largest CO<sub>2</sub>/N<sub>2</sub> permselectivities (43.9 and 46.1, respectively). In contrast, the lowest gas permeabilities and CO<sub>2</sub>/N<sub>2</sub> permselectivity belong to the [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] SILM. The CO<sub>2</sub>/N<sub>2</sub> permselectivities of the pure SILMs decrease as the fluorinated chain increases in the IL anion: [FSI]<sup>-</sup> > [TFSAM]<sup>-</sup> > [NTf<sub>2</sub>]<sup>-</sup> > [BETI]<sup>-</sup>, the [C<sub>2</sub>mim][FAP] and [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] SILMs being the only exceptions.

Concerning the effect of the structural unfolding of the [TFSAM]<sup>-</sup> anion, and as previously discussed, the pure [C<sub>2</sub>mim][TFSAM] SILM presents higher gas permeabilities compared to its structurally similar IL mixtures (Table 6). Nevertheless, the greatest CO<sub>2</sub>/N<sub>2</sub>



**Table 6** Single gas permeability ( $P$ )<sup>a</sup> and ideal permselectivities ( $\alpha$ ) of all the studied SILMs<sup>b</sup>

SILM sample	$P$ (CO <sub>2</sub> )	$P$ (N <sub>2</sub> )	$\alpha$ (CO <sub>2</sub> /N <sub>2</sub> )
[C <sub>2</sub> mim][TFSAM]	753 ± 0.2	17 ± 0.1	43.9 ± 0.1
[C <sub>2</sub> mim][FSI]	843 ± 0.5	18 ± 0.2	46.1 ± 0.5
[C <sub>2</sub> mim][FAP]	624 ± 0.4	24 ± 0.1	26.0 ± 0.1
[C <sub>2</sub> mim][BETI]	437 ± 1.9	18 ± 0.1	24.8 ± 0.2
[C <sub>2</sub> mim][C <sub>4</sub> F <sub>9</sub> SO <sub>3</sub> ]	32 ± 0.2	6 ± 0.1	5.5 ± 0.1
[C <sub>2</sub> mim][SCN] <sup>c</sup>	263 ± 0.6	5 ± 0.2	56.6 ± 1.9
[C <sub>2</sub> mim][SCN][OTf]	428 ± 0.5	7 ± 0.1	57.6 ± 0.4
[C <sub>2</sub> mim][OTf]	479 ± 0.6	12 ± 0.1	41.4 ± 0.3
[C <sub>2</sub> mim][N(CN) <sub>2</sub> ] <sup>c</sup>	476 ± 0.8	7 ± 0.1	67.8 ± 0.6
[C <sub>2</sub> mim][N(CN) <sub>2</sub> ][NTf <sub>2</sub> ] <sup>c</sup>	589 ± 1.9	14 ± 0.2	41.8 ± 0.7
[C <sub>2</sub> mim][NTf <sub>2</sub> ] <sup>c</sup>	589 ± 1.0	17 ± 0.1	35.5 ± 0.3

<sup>a</sup> Barrer (1 Barrer = 10<sup>-10</sup> cm<sub>(STP)</sub><sup>3</sup> cm cm<sup>-2</sup> s<sup>-1</sup> cmHg<sup>-1</sup>). <sup>b</sup> The listed uncertainties represent the standard deviations, based on three experiments. <sup>c</sup> Values taken from Tomé *et al.*<sup>29</sup>

permselectivity (57.6) was achieved for the SILM containing the [C<sub>2</sub>mim][SCN][OTf] IL mixture.

With the purpose of comparing the performance results obtained in this work to those reported in the literature for other SILMs, Fig. 5 displays the Robeson plot for CO<sub>2</sub>/N<sub>2</sub> separation, where the CO<sub>2</sub>/N<sub>2</sub> permselectivity is plotted against CO<sub>2</sub> permeability and the solid black line represents the empirical 2008 upper bound for this gas pair.<sup>51</sup> It can be seen that among the SILMs immobilized with the pure IL having fluorinated anions, both the [C<sub>2</sub>mim][TFSAM] and [C<sub>2</sub>mim][FSI] SILMs fall or exceed the Robeson 2008 upper bound, meaning that these two ILs are the most promising candidates for CO<sub>2</sub>/N<sub>2</sub> separation processes. Comparing the results of the pure [C<sub>2</sub>mim][TFSAM] SILM with those of selected IL mixtures, Fig. 5 clearly shows that the CO<sub>2</sub>/N<sub>2</sub> separation performance of the SILM immobilized with the [C<sub>2</sub>mim][SCN][OTf] IL mixture surpasses the upper bound, primarily due to its higher CO<sub>2</sub>/N<sub>2</sub> permselectivity (57.6). On the other hand, the CO<sub>2</sub>/N<sub>2</sub> separation efficiency of the [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] SILM is on top of the

upper bound, since it presents lower permselectivity (41.8), despite its high CO<sub>2</sub> permeability (589 Barrer) in comparison to that of the [C<sub>2</sub>mim][SCN][OTf] SILM (428 Barrer). Actually, it is the [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] SILM that discloses the most similar CO<sub>2</sub>/N<sub>2</sub> separation performance results to the pure synthesized [C<sub>2</sub>mim][TFSAM] IL (Fig. 5).

## Experimental

### Materials

Lithium bis(pentafluoroethylsulfonyle)imide (Li(CF<sub>3</sub>CF<sub>2</sub>SO<sub>2</sub>)<sub>2</sub>N, LiBETI, 98%, Chameleon Reagent) and lithium nonafluoro-1-butananesulfonate (LiC<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>, > 95%, TCI Chemicals) were used without purification. Reagent-grade dichloromethane, acetonitrile, hexane and ethyl acetate were obtained from Aldrich or Merck and were dried by vacuum distillation over P<sub>2</sub>O<sub>5</sub>. *N*-Methylimidazole (98%, Aldrich) and bromoethane (98%, Acros) were distilled under an inert atmosphere over CaH<sub>2</sub>.

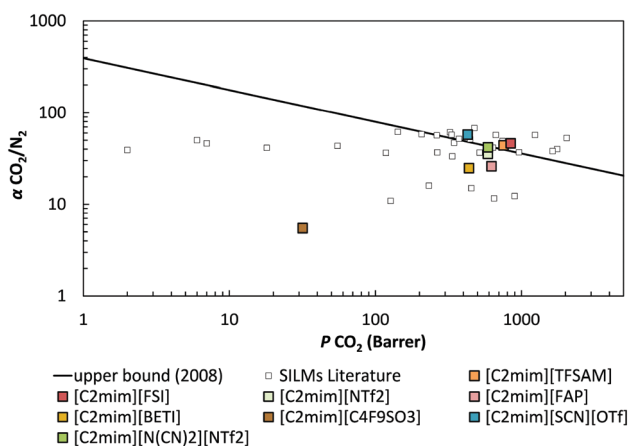
1-Ethyl-3-methylimidazolium bis(fluorosulfonyle)imide ([C<sub>2</sub>mim][FSI], 99.5 wt%, Solvionic), 1-ethyl-3-methylimidazolium tris(pentafluoro-ethyl)trifluorophosphate ([C<sub>2</sub>mim][FAP], 98 wt%, Merck), 1-ethyl-3-methylimidazolium thiocyanate ([C<sub>2</sub>mim][SCN], > 98 wt%, IoLiTec) and 1-ethyl-3-methylimidazolium trifluoromethanesulfonate ([C<sub>2</sub>mim][OTf], ≥ 98 wt%, Aldrich) were obtained from the specified suppliers. To reduce the content of water and other volatile substances, the pure ILs were dried at approximately 1 Pa and 318 K for at least 4 days.

### IL synthesis and characterization

**1-Ethyl-3-methylimidazolium bromide ([C<sub>2</sub>mim][Br]).** [C<sub>2</sub>mim][Br] was synthesized by the reaction between *N*-methylimidazole and an excess of bromoethane following the method used in a previous work.<sup>34</sup> Spectroscopic data of the target compound were in accordance with those reported in the literature.

**1-Ethyl-3-methylimidazolium 2,2,2-trifluoromethylsulfonyle-*N*-cyanoamide ([C<sub>2</sub>mim][TFSAM]).** [C<sub>2</sub>mim][TFSAM] was prepared by ion exchange between [C<sub>2</sub>mim][Br] and KTFSAM, in an aqueous medium in accordance with a described procedure.<sup>35</sup> Yield: 85%; anal. calcd for C<sub>8</sub>H<sub>11</sub>N<sub>4</sub>F<sub>3</sub>SO<sub>2</sub> (284.26), %: N, 19.71%; C, 33.80%; H, 3.90%; Found, %: N, 19.68%; C, 33.78%; H, 3.99%; <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>): 9.08 (s, 1H, 2H (Im)), 7.72 (s, 1H, H4 (Im)), 7.64 (s, 1H, H5 (Im)), 4.22–4.15 (m, 2H, CH<sub>2</sub>CH<sub>3</sub>), 3.84 (s, 3H, CH<sub>3</sub>), 1.42 (t, 3H, CH<sub>2</sub>CH<sub>3</sub>, *J*<sub>HH</sub> = 7.5 Hz); <sup>13</sup>C NMR (100.6 MHz, DMSO-*d*<sub>6</sub>): 136.1, 125.0–115.4 (q, <sup>1</sup>*J*<sub>CF</sub> = 325 Hz), 123.4, 121.8, 44.1, 35.5, 14.8; <sup>19</sup>F NMR (282.4 MHz, DMSO-*d*<sub>6</sub>): −77.8; IR (KBr pellet): 3158 (m,  $\nu_{C-H}$ ), 3117 (m,  $\nu_{C-H}$ ), 2192 (vs,  $\nu_{C-N}$ ), 1573 (m), 1468 (w), 1333 (vs,  $\nu_{asSO_2}$ ), 1236 (s), 1217 (vs,  $\nu_{CF}$ ), 1170 (vs,  $\nu_{sSO_2}$ ), 1119 (s,  $\nu_{CF}$ ), 832 (s), 752 (w), 639 (m), 595 (s), 479 (m) cm<sup>-1</sup>.

**1-Ethyl-3-methylimidazolium nonafluorobutanesulfonate ([C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>]).** Lithium nonafluoro-1-butananesulfonate (8.00 g, 0.026 mol) was dissolved in 20 mL of distilled water and added dropwise to a solution of [C<sub>2</sub>mim][Br] (3.84 g, 0.020 mol) in 15 mL of H<sub>2</sub>O at ambient temperature. The solution was stirred for 2 h at room temperature and then [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] was extracted



**Fig. 5** CO<sub>2</sub>/N<sub>2</sub> separation performance of all the studied SILMs. The experimental error is within the data points. Data are plotted on a log–log scale and the upper bound is adapted from Robeson.<sup>51</sup> The literature data previously reported for other SILMs (□)<sup>15,17,22,24,29,30,32,53</sup> are also illustrated for comparison. The values of [C<sub>2</sub>mim][N(CN)<sub>2</sub>][NTf<sub>2</sub>] and [C<sub>2</sub>mim][NTf<sub>2</sub>] SILMs were taken from Tomé *et al.*<sup>29</sup>



with dichloromethane (4 × 40 mL). The combined CH<sub>2</sub>Cl<sub>2</sub> solution was washed with a small amount of water and dried over anhydrous MgSO<sub>4</sub>. The magnesium sulfate was filtered off and dichloromethane was stripped off under reduced pressure. The product was obtained as slightly yellow transparent fluid oil, which was finally dried at 323 K and 100 Pa for 12 h using a special flask filled with P<sub>2</sub>O<sub>5</sub> and introduced into the vacuum line. Yield: 5.61 g (68%); anal. calcd for C<sub>10</sub>H<sub>11</sub>N<sub>2</sub>F<sub>9</sub>SO<sub>3</sub> (410.26), %: C, 29.28%; H, 2.70%; F, 41.68%; Found, %: C, 28.99%; H, 2.83%; F, 41.39%; <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>): 9.09 (s, 1H, H<sub>2</sub> (Im)), 7.77 (s, 1H, H<sub>4</sub> (Im)), 7.68 (s, 1H, H<sub>5</sub> (Im)), 4.22–4.17 (m, 2H, CH<sub>2</sub>CH<sub>3</sub>), 3.85 (s, 3H, CH<sub>3</sub>), 1.43–1.39 (m, 3H, CH<sub>2</sub>CH<sub>3</sub>); <sup>13</sup>C NMR (100.6 MHz, DMSO-*d*<sub>6</sub>): 136.2, 123.5, 121.9, 117.1 (qt, –CF<sub>2</sub>–CF<sub>3</sub>, <sup>1</sup>J = 288 Hz, <sup>2</sup>J = 34 Hz), 113.4 (tt, (–)O<sub>3</sub>S–CF<sub>2</sub>–, <sup>1</sup>J = 288 Hz, <sup>2</sup>J = 34 Hz), 110.4 (tp, (–)O<sub>3</sub>S–CF<sub>2</sub>–CF<sub>2</sub>–, <sup>1</sup>J = 266, <sup>2</sup>J = 33 Hz), 109.8 (qq, –CF<sub>2</sub>–CF<sub>3</sub>, <sup>1</sup>J = 268 Hz, <sup>2</sup>J = 38 Hz), 44.1, 35.6, 14.9; <sup>19</sup>F NMR (282.4 MHz, DMSO-*d*<sub>6</sub>): –80.8, –114.9, –121.6, –125.9; IR (KBr pellet): 3156 (s, ν<sub>C–H</sub>), 3117 (s, ν<sub>C–H</sub>), 2993 (m, ν<sub>C–H</sub>), 1574 (s), 1462 (m), 1432 (w), 1393 (w), 1353 (s), 1261 (vs, ν<sub>as</sub>SO<sub>2</sub>), 1236 (vs), 1214 (vs, ν<sub>CF</sub>), 1170 (vs, ν<sub>ss</sub>SO<sub>2</sub>), 1134 (s, ν<sub>CF</sub>), 1057 (vs), 1119 (m), 1006 (m), 988 (w), 870 (m), 845 (m), 802 (m), 736 (m), 699 (m), 679 (w), 656 (s), 620 (s), 596 (m), 564 (m), 532 (s) cm<sup>–1</sup>.

**1-Ethyl-3-methylimidazolium bis(pentafluoroethylsulfonyl)imide ([C<sub>2</sub>mim][BETI]).** The procedure previously described for [C<sub>2</sub>mim][C<sub>4</sub>F<sub>9</sub>SO<sub>3</sub>] was also used for the synthesis of [C<sub>2</sub>mim][BETI]. After purification and drying, [C<sub>2</sub>mim][BETI] was obtained as colorless transparent fluid liquid. Yield: 82%; anal. calcd for C<sub>10</sub>H<sub>11</sub>N<sub>3</sub>F<sub>10</sub>S<sub>2</sub>O<sub>4</sub> (491.32), %: C, 24.45%; H, 2.26%; N, 8.55%; Found, %: C, 24.50%; H, 2.03%; N, 8.49%; <sup>1</sup>H NMR (300 MHz, DMSO-*d*<sub>6</sub>): 9.11 (s, 1H, H<sub>2</sub> (Im)), 7.76 (s, 1H, H<sub>4</sub> (Im)), 7.68 (s, 1H, H<sub>5</sub> (Im)), 4.22–4.16 (m, 2H, CH<sub>2</sub>CH<sub>3</sub>), 3.84 (s, 3H, CH<sub>3</sub>), 1.43–1.40 (m, 3H, CH<sub>2</sub>CH<sub>3</sub>); <sup>13</sup>C NMR (100.6 MHz, DMSO-*d*<sub>6</sub>): 136.7, 123.9, 122.4, 117.7 (qt, –CF<sub>2</sub>–CF<sub>3</sub>, <sup>1</sup>J = 287 Hz, <sup>2</sup>J = 34 Hz), 110.1 (tq, –CF<sub>2</sub>–CF<sub>3</sub>, <sup>1</sup>J = 293 Hz, <sup>2</sup>J = 37 Hz), 44.6, 36.0, 15.4; <sup>19</sup>F NMR (282.4 MHz, DMSO-*d*<sub>6</sub>): –78.8, –117.6; IR (KBr pellet): 3160 (m, ν<sub>C–H</sub>), 3124 (m, ν<sub>C–H</sub>), 2993 (w, ν<sub>C–H</sub>), 1574 (m), 1472 (w), 1432 (w), 1355 (vs, ν<sub>as</sub>SO<sub>2</sub>), 1331 (vs, ν<sub>CF</sub>), 1223 (vs, ν<sub>CF</sub>), 1172 (vs, ν<sub>ss</sub>SO<sub>2</sub>), 1087 (s, ν<sub>CF</sub>), 978 (s), 824 (w), 775 (w), 755 (m), 742 (m), 701 (w), 644 (m), 616 (s), 536 (m), 525 (m) cm<sup>–1</sup>.

NMR spectra were recorded on an AMX-400 spectrometer (Bruker) at 298 K in the indicated deuterated solvent and are listed in ppm. The signal corresponding to the residual protons of the deuterated solvent was used as an internal standard for <sup>1</sup>H and <sup>13</sup>C NMR, while for <sup>19</sup>F NMR CHCl<sub>2</sub>F was used as an external standard. IR spectra were acquired on a Nicolet Magna-750 Fourier IR-spectrometer using KBr pellets (128 scans, resolution was 2 cm<sup>–1</sup>).

**IL mixture preparation.** The IL mixture, [C<sub>2</sub>mim][SCN][OTf], containing 0.5 mole fraction of both [C<sub>2</sub>mim][OTf] and [C<sub>2</sub>mim][SCN], was prepared using an analytical high precision balance with an uncertainty of ±10<sup>–5</sup> g by syringing known masses of the IL components into a glass vial. Good mixing was ensured by magnetic stirring for 30 min at 298 K. Then, the IL mixture was dried at roughly 1 Pa and 318 K for at least 4 days immediately prior to use. The water contents of all IL samples were determined by Karl Fischer titration using a 831 KF Coulometer (Metrohm).

**Density and viscosity determination.** The density and viscosity measurements of the pure ILs and the [C<sub>2</sub>mim][SCN][OTf] IL mixture were performed at 293 K and atmospheric pressure using an SVM 3000 Anton Paar rotational Stabinger viscometer-densimeter, where the standard uncertainty for the temperature was 0.02 K. The repeatability of density and dynamic viscosity of this equipment was 0.0005 g cm<sup>–3</sup> and 0.35%, respectively. Measurements of each sample were performed in triplicate to ensure accuracy and the reported results are average values. The highest relative standard uncertainty registered for the density and dynamic viscosity measurements was 1 × 10<sup>–4</sup> and 0.03, respectively.

**Gas permeation measurements.** Porous hydrophobic poly(vinylidene fluoride) (PVDF) membranes supplied by Millipore Corporation (USA), with a pore size of 0.22 μm and an average thickness of 125 μm, were used to support [C<sub>2</sub>mim][FSI], [C<sub>2</sub>mim][FAP] and [C<sub>2</sub>mim][BETI]. Since the impregnation of the remaining IL samples into hydrophobic PVDF resulted in unstable SILMs, the other IL samples were supported on porous hydrophilic poly(tetrafluoroethylene) (PTFE) membranes acquired from Merck Millipore, with a pore size of 0.2 μm and an average thickness of 65 μm. All the SILM configurations were prepared by the vacuum method.<sup>29</sup>

Ideal gas permeabilities and diffusivities through the prepared SILMs were measured using a time-lag apparatus.<sup>36</sup> First, each SILM was degassed under vacuum inside the permeation cell for 12 h. Then, CO<sub>2</sub> and N<sub>2</sub> permeation experiments were carried out at 293 K with a *trans*-membrane pressure differential of 100 kPa. All the permeation data were measured at least in triplicate on a single SILM sample. The highest relative standard uncertainty registered for gas permeability measurements was 0.03. The permeation cell and lines were evacuated until the pressure was below 0.1 kPa before each run. No residual IL was found inside the permeation cell at the end of the experiments. The thickness of the SILMs was assumed to be equivalent to the membrane filter thickness.

Gas transport through the prepared SILMs was assumed to follow a solution–diffusion mass transfer mechanism,<sup>37</sup> where the permeability (*P*) is related to diffusivity (*D*) and solubility (*S*) as follows:

$$P = D \times S \quad (1)$$

The permeate flux of each gas (*J<sub>i</sub>*) was determined experimentally using eqn (2),<sup>38</sup> where *V<sub>p</sub>* is the permeate volume, Δ*p<sub>d</sub>* is the variation of downstream pressure, *A* is the effective membrane surface area, *t* is the experimental time, *R* is the gas constant and *T* is the temperature.

$$J_i = \frac{V_p \Delta p_d}{A t R T} \quad (2)$$

Ideal gas permeability (*P<sub>i</sub>*) was then determined from the steady-state gas flux (*J<sub>i</sub>*), the membrane thickness (*ℓ*) and the trans-membrane pressure difference (Δ*p<sub>i</sub>*), as shown in eqn (3).<sup>38</sup>

$$P_i = \frac{J_i}{\Delta p_i / \ell} \quad (3)$$



Gas diffusivity ( $D_i$ ) was determined according to eqn (4). The time-lag parameter ( $\theta$ ) was calculated by extrapolating the slope of the linear portion of the  $p_d$  vs.  $t$  curve back to the time axis, where the intercept was equal to  $\theta$ .<sup>39</sup>

$$D_i = \frac{\ell^2}{6\theta} \quad (4)$$

After  $P_i$  and  $D_i$  were known, the gas solubility ( $S_i$ ) was calculated using the relationship shown in eqn (1). The ideal permeability selectivity (or permselectivity),  $\alpha_{ij}$ , was obtained by dividing the permeability of the more permeable species  $i$  to the permeability of the less permeable species  $j$ . The permselectivity can also be expressed as the product of the diffusivity selectivity and the solubility selectivity:

$$\alpha_{ij} = \frac{P_i}{P_j} = \left( \frac{D_i}{D_j} \right) \times \left( \frac{S_i}{S_j} \right) \quad (5)$$

## Conclusions

In this work, ILs containing a common cation ( $[\text{C}_2\text{mim}]^+$ ) and different fluorinated anions ( $[\text{TFSAM}]^-$ ,  $[\text{FSI}]^-$ ,  $[\text{C}_4\text{F}_9\text{SO}_3]^-$ ,  $[\text{BETI}]^-$ ,  $[\text{FAP}]^-$ ) were synthesized and used as liquid phases to prepare SILMs for flue gas separation ( $\text{CO}_2/\text{N}_2$ ). The single  $\text{CO}_2$  and  $\text{N}_2$  permeation properties through the prepared SILMs were determined. The viscosity and density of the IL phases were also evaluated. The results showed that  $\text{CO}_2$  permeabilities and diffusivities through the studied SILMs follow the same fluorinated anion order:  $[\text{FSI}]^- > [\text{TFSAM}]^- > [\text{FAP}]^- > [\text{NTf}_2]^- > [\text{BETI}]^- > [\text{C}_4\text{F}_9\text{SO}_3]^-$ , which is inversely related to IL viscosity, with the only outlier being  $[\text{C}_2\text{mim}][\text{FAP}]$ . Conversely, the introduction of fluorination in the IL anions did not significantly affect gas solubility, except for the case of the  $[\text{C}_2\text{mim}][\text{C}_4\text{F}_9\text{SO}_3]$  SILM displaying a very low  $\text{CO}_2$  solubility. Among the pure SILMs, it is worth noting that the best separation performances were achieved for  $[\text{C}_2\text{mim}][\text{TFSAM}]$  and  $[\text{C}_2\text{mim}][\text{FSI}]$  SILMs that fall on top of or surpassed the Robeson 2008 upper bound, with  $\text{CO}_2$  permeabilities of 753 and 843 Barrer and  $\text{CO}_2/\text{N}_2$  permselectivities of 43.9 and 46.1, respectively.

Furthermore, the effect of structural unfolding of the  $[\text{TFSAM}]^-$  anion on gas permeation properties of SILMs was investigated using IL mixtures comprising both fluorinated and cyano functionalities in the anions. The pure  $[\text{C}_2\text{mim}][\text{TFSAM}]$  IL provided a membrane with improved  $\text{CO}_2$  permeabilities, diffusivities and solubilities compared to those of the SILMs based on the selected  $[\text{C}_2\text{mim}][\text{SCN}][\text{OTf}]$  and  $[\text{C}_2\text{mim}][\text{N}(\text{CN})_2][\text{NTf}_2]$  IL mixtures. Overall, and despite the fact that the  $[\text{C}_2\text{mim}][\text{SCN}][\text{OTf}]$  SILM revealed better  $\text{CO}_2/\text{N}_2$  separation performance essentially due to its higher  $\text{CO}_2/\text{N}_2$  permselectivity (57.6), the  $[\text{C}_2\text{mim}][\text{N}(\text{CN})_2][\text{NTf}_2]$  IL mixture disclosed the most similar results to the pure synthesized  $[\text{C}_2\text{mim}][\text{TFSAM}]$  IL, not only in terms of thermophysical properties, but also regarding gas transport and  $\text{CO}_2/\text{N}_2$  separation performance.

## Conflicts of interest

There are no conflicts to declare.

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