



Cite this: *Phys. Chem. Chem. Phys.*,
2017, **19**, 6095

Benzimidazolylquinoxalines: novel fluorophores with tuneable sensitivity to solvent effects†

Timur I. Burganov,^a Nataliya A. Zhukova,^a Vakhid A. Mamedov,^a
Christoph Bannwarth,^b Stefan Grimme^b and Sergey A. Katsyuba*^a

We report on the photophysical properties, conjugation, conformational behavior, intra- and intermolecular hydrogen bonds (HBs) of a series of novel fluorophores, consisting of 3-arylquinoxaline and benzimidazole moieties linked by a single CC bond. Computations employing density functional theory (DFT) reveal that conjugation between these moieties stabilizes *syn*-conformers with two HB centers located on the same side of the molecule. *Anti*-conformers form stronger intermolecular HBs with DMSO and DMF than *syn*-conformers, and this influences the energy gap between *syn*- and *anti*-forms, especially upon excitation of the molecules to the S_1 state. Substituents introduced in various positions of the molecules modify their conformational behavior, and mutual disposition of excited singlet states relative to the ground states. Various substitution patterns produce very different effects on relative quantum yield of luminescence: from a moderate increase in polar DMSO and DMF relative to 1,2-dichloroethane solutions to complete quenching of emission which is observable in polar media. The observed behavior is understood with the aid of DFT and time-dependent DFT calculations. The tuneability of the spectroscopic range of the luminescence and especially of its sensitivity to environmental effects allows rational design of the novel fluorophores of this family for various applications.

Received 28th September 2016,
Accepted 27th January 2017

DOI: 10.1039/c6cp06658e

rsc.li/pccp

Introduction

Bifunctional azaaromatic compounds, containing a hydrogen bond (HB) donor and acceptor located in separate moieties linked by a single bond, attract considerable attention as fluorophores often demonstrating a remarkable sensitivity to environmental effects, which makes them suitable for fluorescence sensing in a wide range of chemical, biochemical and biophysical applications (see, for example, ref. 1 and 2 and references cited herein). Such a topology often leads to a solvent-dependent appearance of two conformers, *syn* and *anti*, with the two HB centers located either on the same or on the opposite sides of the molecule, respectively (Fig. 1). It is believed that an intramolecular HB stabilizes the *syn* conformer, which usually dominates in nonpolar solvents. In polar solvents, where formation of intermolecular HBs with HB donor and/or acceptor sites is possible, the fraction of the more polar *anti* conformer may become significant. In spite of the vast number of studies of these conformationally flexible systems, they were limited to systems, comprising either a pyridyl or a

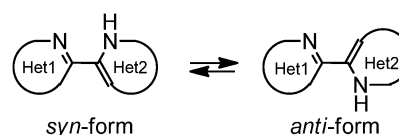


Fig. 1 *Syn*- and *anti*-conformers of bifunctional azaaromatic compounds.

naphthyridine moiety as a HB acceptor and a pyrrole, pyrazole or indole unit as an intramolecular HB donor (see, for example, pyridylpyrroles,³ -pyrazoles^{4,5} and -indoles,⁵ or indole-naphthyridines⁷). These latter units typically contain the only heteroatom in the immediate vicinity of the C–C bond linking the HB donor and acceptor groups (Fig. 1).

Recently we have synthesized a series of 2-(benzimidazol-2-yl)-3-arylquinoxalines⁸ (Fig. 2), where the benzimidazole moiety, being the HB donor owing to its NH group, also includes another nitrogen atom immediately adjacent to the C–C link to the HB acceptor, *i.e.*, the 3-arylquinoxaline moiety.

This structural peculiarity may result in inversion of relative polarity of *syn* and *anti* forms as compared with other bifunctional azaaromatic compounds mentioned above. In the present work we study internal rotation, conjugation, HB and photophysical characteristics of these novel compounds in polar and nonpolar solvents. The results demonstrate that absorption and emission of the benzimidazolylquinoxalines (BQ) and environmental effects

^a A. E. Arbusov Institute of Organic and Physical Chemistry, Kazan Scientific Centre of the Russian Academy of Sciences, Arbuzov str. 8, 420088 Kazan, Russia.
E-mail: katsyuba@iopc.ru, skatsyuba@yahoo.com

^b Mulliken Center for Theoretical Chemistry, Institut für Physikalische und Theoretische Chemie der Universität Bonn, Beringstr. 4, 53115 Bonn, Germany

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c6cp06658e



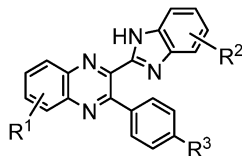


Fig. 2 General formula of novel 2-(benzimidazol-2-yl)-3-arylquinoxalines.

influencing their luminescence are strongly dependent on substituents introduced in various positions of the aromatic moieties of these compounds.

Experimental and computational details

All compounds considered were synthesized according to the protocols described earlier.^{8,9} Electronic absorption (UV-Vis) and steady state fluorescence spectra were recorded at room temperature on a Perkin-Elmer Lambda 35 spectrometer and a Cary Eclipse Fluorescence Spectrophotometer (Varian), respectively, using 10 mm quartz cells. Absorption spectra were registered with a scan speed of 480 nm min⁻¹, using a spectral width of 1 nm. All samples were prepared as solutions in three solvents: 1,2-dichloroethane (DCE), dimethylsulfoxide (DMSO), and dimethylformamide (DMF), with the concentrations ranging from ~10⁻⁵ to ~10⁻⁴ mol L⁻¹. The absorbance at excitation wavelength was less than 0.1 to avoid the “inner filter effect”. Fluorescence quantum yields were measured using aqueous solution of quinine sulfate as the standard ($\lambda_{\text{ex}} = 355$ nm, $\phi = 0.57$ in 0.1 M H₂SO₄). Appropriate corrections were made for the optical density of the solutions and the refractive index of the medium.¹⁰

All density functional theory (DFT) calculations were performed with the Turbomole 6.4 program package.¹¹ In the ground state self-consistent field (SCF) calculations, the resolution-of-the-identity approach for Coulomb and exchange integrals (RI-JK)¹²⁻¹⁵ was employed. In general the spherical def-TZVP atomic orbital (AO) basis set was used.¹⁶ All structures were optimized with the hybrid PBE0 functional.¹⁷ In all geometry optimizations the D3 approach¹⁸ to describe the London dispersion interactions together with the Becke–Johnson (BJ) damping function¹⁹ was employed as implemented in the TURBOMOLE program. Stationary points were characterized as minima by frequency analyses. The same frequencies were used in the computations of the free energy within the framework of a modified rigid rotor, harmonic oscillator statistical treatment (ΔG_{RRHO}).²⁰ The optimized geometries were further used for the single-point calculations of the electronic energies by applying the PW6B95 meta-hybrid density functional²¹ in combination with the quadruple-zeta def2-QZVP Ahlrich’s basis set.^{22,23} The computed free energies are then obtained from $\Delta G_{\text{gas}} = \Delta E(\text{PW6B95/def2-QZVP}) + \Delta G_{\text{RRHO}}(\text{PBE0/def-TZVP})$. Solvent effects on the thermochemical properties have been obtained by the COSMO-RS method²⁴ (COSMOtherm software package²⁵) based on BP86/TZVP²⁶ calculations (parameterization from 2012). Solvation contributions to free energies at 298.15 K are computed from the gas phase structures obtained at the abovementioned

levels of theory. The computed free energies are then obtained by $\Delta G_{\text{solution}} = \Delta G_{\text{gas}} + \Delta \delta G_{\text{COSMO-RS}}$.

Time-dependent density functional response theory (TD-DFT)²⁷⁻²⁹ has been employed to compute the vertical excitation energy (*i.e.*, absorption and emission wavelengths) and oscillator strength on the ground state and first excited state geometries, each optimized in the gas phase as well as in the 1 : 1 complexes with DMSO molecules. The equilibrium geometries of the lowest excited singlet states have been determined at the PBE0-D3(BJ)/def-TZVP level, making use of Tamm–Dancoff approximated TD-DFT³⁰ for calculation of the vertical excitation energies. In most cases, the spectra were broadened by Gaussian functions with a full-width at 1/*e* height of 0.4 eV. No energy shift has been applied. The dipole length representation is used to calculate oscillator strengths discussed in the present paper. The hybrid B3LYP^{31,32} functional in combination with the 6-31+G* basis set³³⁻³⁶ was used for computation of the Raman spectra according to our previous experience in studies of conjugational effects.³⁷ Molecular orbitals based on PBE0/def-TZVP were visualized in the ChemCraft 1.6 program³⁸ with a 0.03 contour value.

Results and discussion

Conformations in the ground (S₀) and excited (S₁) states

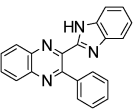
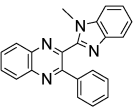
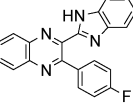
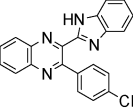
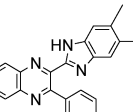
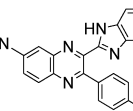
The quinoxaline moiety (Q) of the studied molecules is linked by C_Q–C_B and C_Q–C_{Ph} single bonds with benzimidazole (B) and phenyl (Ph) moieties, respectively. According to quantum chemical computations, internal rotation about the C_Q–C_{Ph} bond results in the only stable conformation with a dihedral angle of *ca.* 40–50° between the Q and Ph planes, depending on the specific molecule and the conformation about the C_Q–C_B bond. In spite of non-coplanarity of the Q and Ph rings, both aromatic moieties are conjugated, which is clearly indicated by *ca.* a ten- to fifteen-fold increase of the Raman intensity of the band of the phenyl ν_{sa} mode at *ca.* 1600 cm⁻¹ (*I*_{sa}) relative to the corresponding band of toluene (for 1, 2 and 5, see Table 1 for numbering of compounds) or *para*-fluoro- (for 3) and *para*-chloro-toluenes (for 4 and 6).^{37,39} These toluenes are regarded as “external standard” molecules, where the aromatic ring does not conjugate with the methyl group.

Internal rotation about the C_Q–C_B bond results in two stable conformations: *syn* and *anti* (*vide supra*). Q and B moieties lie practically in the same plane in the *syn*-conformation (Fig. 1S, ESI†), but are not coplanar in the *anti*-conformation for steric reasons. Conjugation between the almost coplanar Q and B moieties in the *syn*-conformers is more pronounced than in the non-planar *anti*-conformers, which is suggested by the highest occupied molecular orbital (HOMO) pictures (Fig. 3) and red shift of the electronic absorption maxima in the TDDFT calculated spectra of the *syn*-conformers relative to the *anti*-forms (Fig. 4).

To estimate the possible role of intramolecular HB NH⋯N in stabilization of *syn*-conformers relative to the *anti*-conformers we have compared structural, spectroscopic and electron density characteristics of the conformers. The intramolecular HB can be quantitatively characterized by elongation of the N–H bond



Table 1 Electronic energies ΔE (kcal mol⁻¹) and Gibbs free energies ΔG (kcal mol⁻¹) of *anti*-conformers relative to *syn*-conformers calculated for isolated molecules under study (gas phase) or for their H-complexes with dimethylsulfoxide (DMSO). ΔG values estimated for 1,2-dichloroethane (DCE) and DMSO solutions with the use of COSMO-RS method²⁴ are given in parentheses (in italics and bold, respectively)

Compound	S ₀ (gas phase)		S ₀ (DMSO)	S ₁ (gas phase)	S ₁ (DMSO)
	ΔE	ΔG	ΔE	ΔE	ΔE
 1	3.2	3.2 (1.4 ^a /0.8 ^b)	1.0	3.6	-1.0
 2	-0.3	0.2	—	-0.4	—
 3	3.7	3.7	—	4.4	—
 4	3.8	3.8 (1.3 ^a /0.9 ^b)	0.9	4.2	-1.0
 5	3.3	3.4	—	4.0	—
 6	4.2	4.1 (2.1 ^a /0.9 ^b)	2.8	5.8	0.1

^a DCE solution. ^b DMSO solution.

of the *syn*-conformer relative to the *anti*-conformer, $\Delta\text{NH}_{\text{syn-anti}}$ and a red shift and an increase of the intensity of the infrared (IR) band of NH stretching vibrations $\Delta\nu\text{NH}_{\text{syn-anti}}$ and $\Delta I(\nu\text{NH})_{\text{syn-anti}}$ respectively.⁴⁰ Almost negligibly small values found for all three characteristics of HB strength (Table 1S, ESI[†]) suggest that NH...N intramolecular H-bonding is either very weak or absent. *E.g.*, $\Delta\text{NH}_{\text{syn-anti}} = 0.00123 \text{ \AA}$ in the case of **1**. Following the Bader analysis of the electronic density (see, *e.g.*, ref. 41 and references cited herein) a stabilizing HB could be indicated by a NH...N bond critical point. Nevertheless, the critical points at the NH...N intramolecular contacts were localized in none of the *syn*-conformers of molecules **1**, **3–6**. This fact is additional evidence of the absence or the insignificant role of intramolecular HBs in conformational behavior of the studied compounds. Hence, the energetic preferability of *syn*-conformers (Table 1) results from

the other factors, *e.g.*, the abovementioned difference in conjugation of *syn*- and *anti*-conformers, as well as steric congestion in the *anti*-conformation.

The minimal difference in Gibbs free energy (ΔG) of the *anti*-conformer relative to the corresponding *syn*-conformer, calculated for the gas phase, is found to be 3.2 kcal mol⁻¹ for compound **1**. The barrier ΔG^\ddagger for *syn*→*anti* transformation during internal rotation about the C_Q–C_B bond amounts to 5.9 kcal mol⁻¹ for **1**, while the *anti*→*syn* barrier is only 2.8 kcal mol⁻¹. Thus, population of *syn*-conformations is favorable thermodynamically. A minimal ΔG value is found for molecule **2**, where replacement of the N–H moiety of molecule **1** by an N–Me moiety results in steric strain induced by the methyl group, which, in turn, essentially increases non-planarity of the *syn*-conformer of **2** (dihedral angle between Q and B moieties = 44°) and strongly destabilizes this conformer relative to the *anti*-form (Table 1).

The above conformational behavior of the isolated molecules suggests that in non-polar media benzimidazolylquinoxalines **1**, **3–6** should exist mainly in the *syn*-form. Solvent effects of the moderately polar aprotic solvent, 1,2-dichloroethane (DCE), on the thermochemical properties obtained by the COSMO-RS²⁴ method do not change the above picture essentially: ΔG (*syn*→*anti*) estimated for DCE solutions varies between 1.3 and 2.1 kcal mol⁻¹ (Table 1). To estimate the possible influence of polar solvents, dimethylsulfoxide (DMSO) and dimethylformamide (DMF), on conformational equilibria of the compounds under study, we optimized H-bonded 1 : 1 complexes of *syn*- and *anti*-conformers with the respective solvent molecules (*e.g.*, Fig. 5).

According to this explicit solvation approach, HB-complexes formed by *anti*-conformers with the solvent molecules are stronger than the HBs formed by *syn*-conformers (Table 1S, ESI[†]), and this difference results in stabilization of the *anti*-forms relative to the *syn*-conformations. It should be noted that the influence of DMSO, taken into account implicitly within the framework of the COSMO-RS²⁴ model, produces very similar thermodynamic effects: even more polar *syn*-conformers are destabilized in this polar solvent relative to less polar *anti*-conformers (Table 1). In spite of this decrease of ΔG (*syn*→*anti*) values to *ca.* 1–3 kcal mol⁻¹, the *syn*-conformers are still expected to dominate even in the polar solutions, though the presence of the *anti*-conformers cannot be excluded in these cases. The domination of *syn*-conformers in polar media is confirmed by comparison of the TDDFT simulated absorption spectra of *syn*- and *anti*-conformers with the experimental spectra of **1–6** registered for DMF or DMSO solutions (see, *e.g.*, Fig. 4). In contrast, similar comparison for **2** (Fig. 2S, ESI[†]) suggests domination of *anti*-conformers both in DCE and in polar solutions.

According to our quantum chemical computations, the conformational behavior of the isolated molecules in the first excited state, S₁, is qualitatively similar to that described above for the ground state, S₀ (Table 1). Energy differences, ΔE , between *syn*- and *anti*-conformers in the S₁ state moderately increase relative to the S₀ state because *syn*-conformers are additionally stabilized by the NH...N intramolecular HB, which is quite pronounced in the excited state.⁴² The strengthening of the intramolecular HBs can be ascribed to the electron density redistribution upon



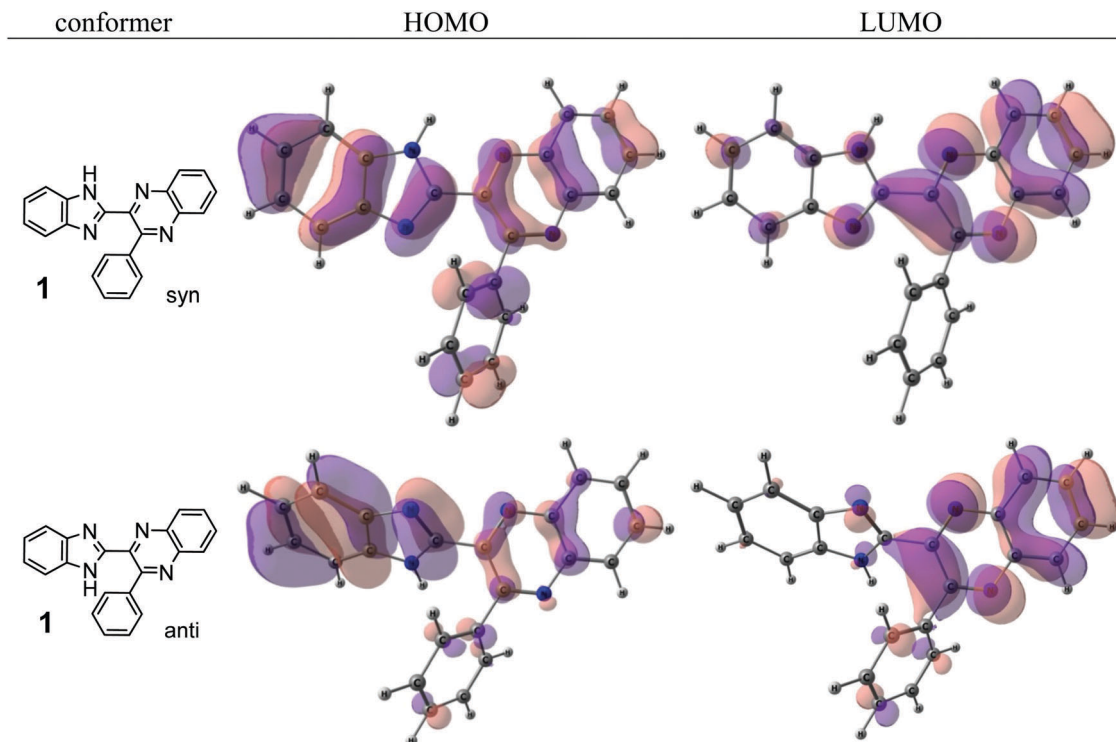


Fig. 3 The highest occupied (HOMO) and lowest unoccupied (LUMO) molecular orbitals of the *syn*- and *anti*-conformers of **1**.

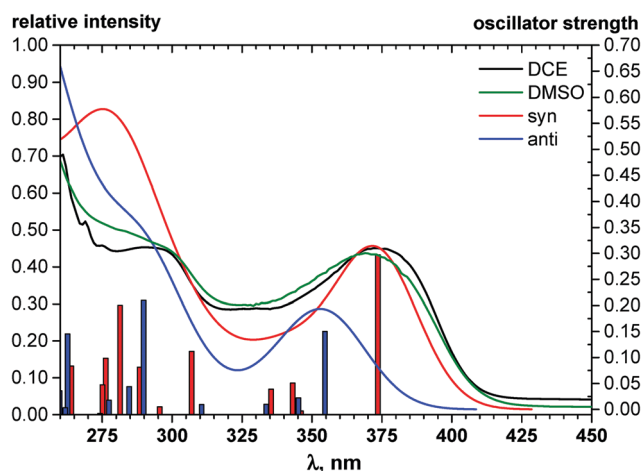


Fig. 4 TDDFT simulated absorption spectra of *syn*- (red lines) and *anti*-conformers (blue lines) of **3** in comparison with the experimental spectrum registered for **3** solutions in DCE (black line) and DMSO (green line). The heights of the vertical straight lines correspond to the calculated oscillator strengths of the corresponding vertical transitions. Red and blue envelope curves are obtained by broadening of the vertical straight lines by Gaussian functions with a full-width at 1/e height of 0.4 eV.

excitation, resulting in an increase in the excited state basicity of the proton acceptor and/or an increase in the excited state acidity of the proton donor. Similar strengthening of the $\text{NH} \cdots \text{O}$ intermolecular HBs with the DMSO or DMF solvent molecules in the excited state results in the inversion of relative energetic stability of *syn*- and *anti*-conformers of BQs in the S_1 state (Table 1).⁴³

As *anti*-conformers form stronger HBs with DMSO or DMF than *syn*-conformers (Table 1S, ESI[†]), the former conformers are more effectively stabilized in the S_1 state than the latter. These results suggest that the emission spectra of **1**, **3–6** should be assigned mainly to *syn*-conformers in the case of DCE solutions, while the possible presence of *anti*-conformers should be taken into account in the case of DMSO and DMF solutions. In contrast, both *anti*- and *syn*-conformers of **2** are expected to determine emission of all solutions of this compound. It should be noted though that according to TD-DFT computations, the energies of vertical $S_1 \rightarrow S_0$ transitions of *syn*- and *anti*-conformers practically coincide (Table 2S, ESI[†]).

Electronic absorption and fluorescence spectra

The electronic absorption and fluorescence spectra were recorded in three aprotic solvents: intermediately-polar DCE ($\epsilon = 10.36$), and polar DMF ($\epsilon = 36.7$) and DMSO ($\epsilon = 46.7$). The spectra are compared in Fig. 6 and in Table 2. On the basis of TDDFT computations the lowest-energy absorptions in the spectra (see, e.g., Fig. 4) are assigned mainly to transitions between the highest occupied and the lowest unoccupied molecular orbitals (HOMO and LUMO, respectively) of *syn*-conformers of **1**, **3–6**, and of *anti*-conformers of **2**. Visualization of these orbitals (Fig. 3) demonstrates that the HOMO of **1** is spread over π -systems of all three aromatic rings: B, Q, and Ph. The LUMO of **1** is represented mainly by the π^* -orbital of the Q moiety. This visual analysis of frontier MOs suggests that the first vertical transition results in partial charge transfer from the donor B moiety to the acceptor Q moiety. The charge-transfer character of



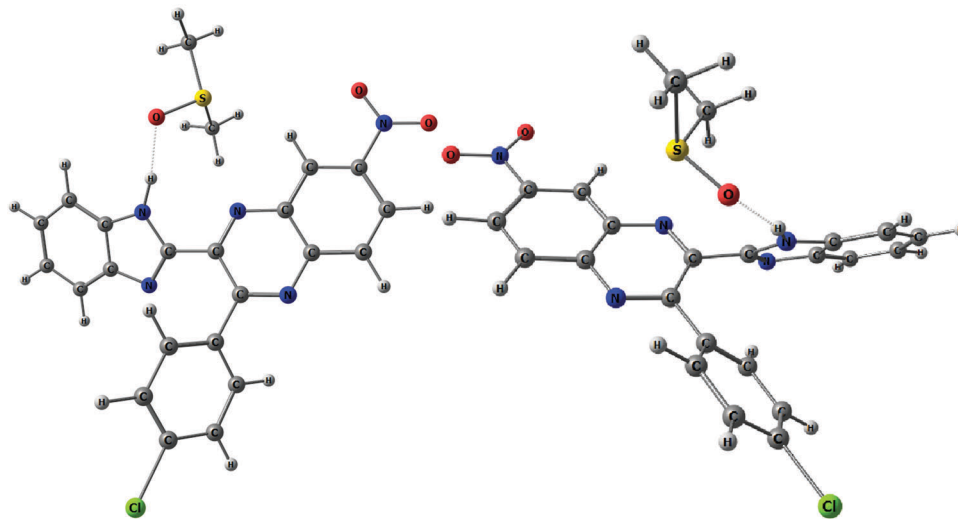


Fig. 5 H-bonded 1:1 complexes of *syn*- (left) and *anti*-conformers (right) of **6** with DMSO molecules. H-bonds are shown with dotted lines.

this transition is confirmed also by a *ca.* 3- to *ca.* 5-fold increase of dipole moments of the molecules in the S_1 state relative to the ground S_0 state (Table 3S, ESI†). In spite of this, all main features of the experimental absorption and emission spectra of the compounds under study were nicely reproduced (Table 2) by TDDFT computations at the PBE0/def-TZVP level of approximation.

As can be seen from Table 2, the wavelengths of lowest-energy absorptions in the spectra of various solutions of the BQs vary in the range of 352 to 396 nm (3.55 to 3.13 eV) and practically do not depend on the solvents used. Both visual analysis of the frontier MOs of the molecules (*vide supra*) and rather large values of the oscillator strengths computed for the vertical $S_0 \rightarrow S_1$ transitions (Table 2) suggest that these absorptions are dominantly of $\pi\text{-}\pi^*$ character.

The fluorescence excitation spectra of the compounds match well the absorption. The wavelengths of the emission maxima vary in the rather broad range of 440 to 519 nm (2.82 to 2.39 eV) and depend on the solvent used: a replacement of DCE by the much more polar DMF and DMSO causes red shifts of *ca.* 20–25 nm (Table 2). Similar dependencies on solvent polarity were reported also for other bifunctional HB donor/acceptor azaaromatic compounds, see *e.g.* ref. 3, 5–7, 44 and references cited therein. The above mentioned shifts in the present case can be, at least partly, ascribed to stronger stabilization of the more polar S_1 states in polar media relative to the less polar S_0 species. Intermolecular HBs $\text{N-H}\cdots\text{O}=\text{S}$ or $\text{N-H}\cdots\text{O}=\text{C}$ with DMSO or DMF molecules, respectively, also decrease the $S_0\text{-}S_1$ energy gap, as they are stronger in the S_1 state than in the ground S_0 state (Table 1S, ESI†).⁴³

Quantum yields of emission collected in Table 2 vary between 0.03 and 0.29, being quite comparable to the yields reported for the other related azaaromatic chromophores.^{44–63} A distinguishing feature of the studied compounds is that the intensity of their luminescence varies greatly with the used solvent. In the case of **1**, **3**–**5** (Table 2) it is maximal in the polar DMSO and DMF solvents, while in the less polar DCE the quantum yield of emission drops to *ca.* 30–70% of its maximal value. In contrast, fluorescence of **6** is easily observable in DCE

solutions and is completely quenched in DMF and DMSO. Such a behavior drastically differs from behavior of **1**, **3**–**5** and the other related azaaromatic chromophores.^{44–63}

As no luminescence quenching in the polar solvents is found for other BQs, in particular, for the closely related **4**, obviously, the only structural difference between **4** and **6**, *viz.* the NO_2 moiety of **6**, plays a key role in the above effect. It is known that the presence of nitro groups typically quenches the fluorescence of organic compounds to the level below the detection limit.⁶⁴ The non-radiative decay pathways may involve internal coordinates of the nitro group.⁶⁵ A complete mechanistic picture is rather complicated and very high-level single- and multi-reference methods are needed to assure sufficiently accurate theoretical description.^{65–67} This extremely computationally demanding task is beyond the scope of the present study. Nevertheless, it must be emphasized that the NO_2 moiety in molecule **6**, indeed, demonstrates remarkable structural changes upon excitation (Fig. 1S, ESI†).

There were a few reports showing the dependence of quantum yield of some fluorophores, containing the nitrophenyl group, on solvent polarity, see, *e.g.* ref. 68 and 69. In contrast, in our case even the addition of small amounts of DMF or DMSO in solution of **6** in DCE, which should not strongly affect the dielectric properties of the solvent, results in a dramatic quenching of the fluorescence (Fig. 7).⁷⁰

This suggests that the polarities of media play only a minor role in the above effect, and instead, HBs of **6** with DMF or DMSO are a key factor of the quenching. The primary effect of H-bonding with DMSO or DMF is an energetic shift of the singly excited states. As $\pi\text{-}\pi^*$ and $n\text{-}\pi^*$ states have different sensitivities to H-bonding, the shift results in a change of the energy gap between these states and a decreasing gap can enhance the internal conversion to the ground state.⁷¹

Influence of the structure of BQs on their conformational and photophysical properties

The structure of the parent BQ molecule **1** was modified in two ways: first, by the introduction of $-\text{Cl}$ (compound **4**), $-\text{F}$ (**3**)



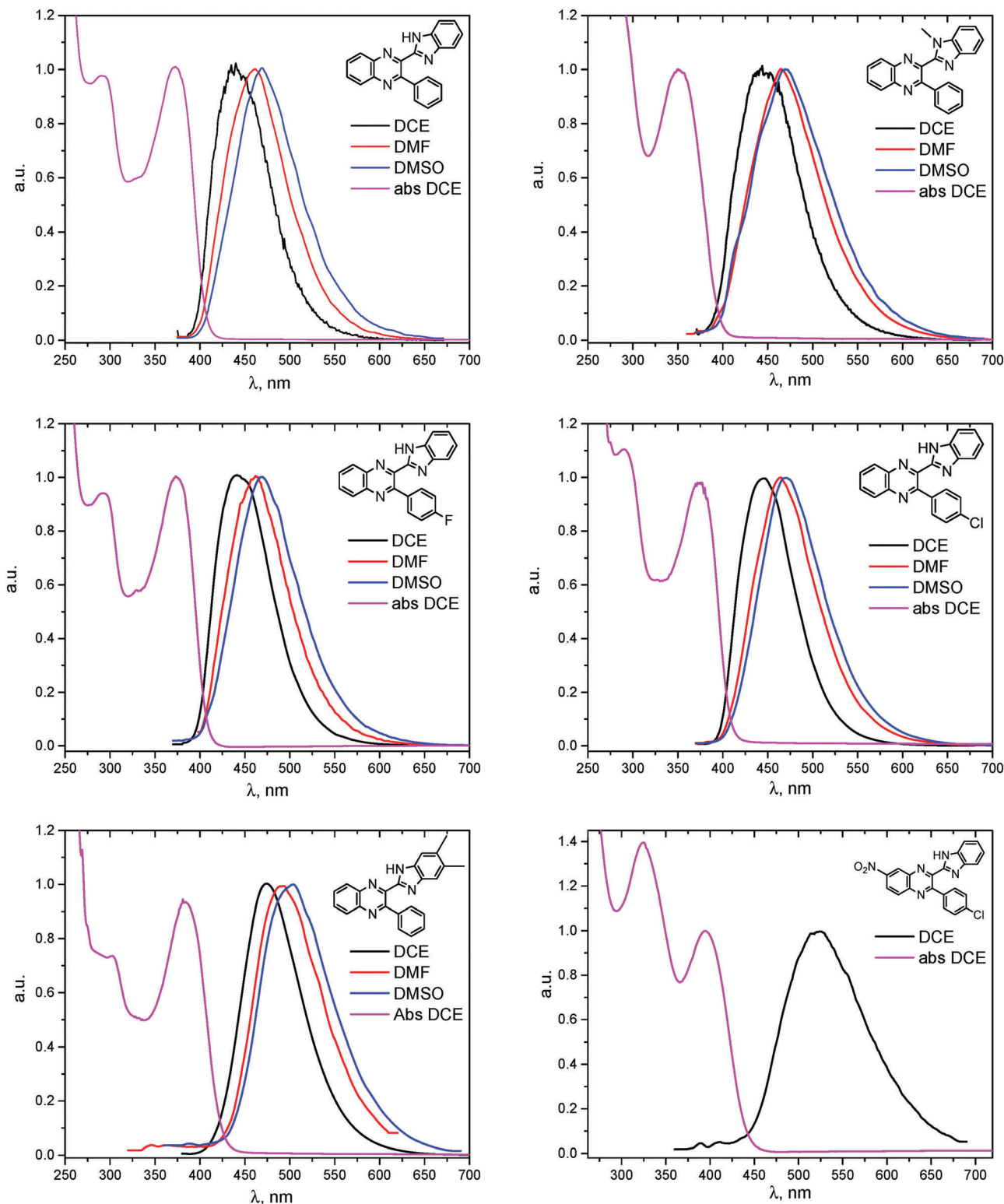


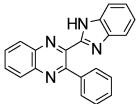
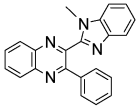
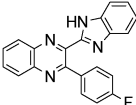
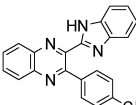
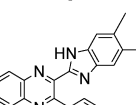
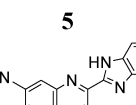
Fig. 6 Absorption of BQs **1–6** in DCE (magenta) and fluorescence in DCE (black), DMF (red) and DMSO (blue).

and $-\text{NO}_2$ (**6**) substituents to the acceptor 3-phenylquinoxaline moiety; second, by the introduction of two methyl groups (compound **5**) to the donor benzimidazole moiety. In spite of these modifications, the conformational behavior of all the

studied BQs remained qualitatively the same: in the ground state *syn*-conformers strongly dominate in the gas phase and in intermediately-polar DCE solutions. The energetic preferability of the *syn*-conformers modestly increases along with the



Table 2 TDDFT calculated characteristics of vertical S_0 – S_1 transitions and experimental characteristics of the absorption and emission bands obtained for **1–6** in various solvents at 293 K

Compound	Calculations ^a		Experiment ^b			λ_{emi} , nm (ϕ)		
	λ_{abs} , nm (f)	λ_{emi} , nm (f)	λ_{abs} , nm ($\log \epsilon$)			DCE ^c	DMF ^d	DMSO ^e
 1	368 (0.351)	438 (0.111)	373 (4.24)	367 (4.15)	367 (4.10)	440 (0.06)	460 (0.13)	470 (0.11)
 2	367 (0.220)	467 (0.059)	352 (4.03)	352 (4.01)	352 (4.02)	446 (0.03)	465 (0.03)	470 (0.04)
 3	373 (0.321)	454 (0.133)	375 (4.20)	370 (4.13)	370 (4.18)	444 (0.03)	462 (0.16)	468 (0.12)
 4	373 (0.320)	442 (0.103)	375 (4.24)	371 (4.15)	371 (4.18)	446 (0.05)	468 (0.20)	470 (0.18)
 5	382 (0.429)	466 (0.129)	383 (4.25)	383 (4.20)	380 (4.20)	473 (0.06)	492 (0.29)	501 (0.25)
 6	412 (0.262)	513 (0.133)	396 (4.16)	391 (4.12)	391 (4.09)	519 (0.03)	—	—

^a Gas-phase calculations for isolated *syn*-conformers; f – calculated oscillator strengths (length representation). ^b λ_{abs} – the longest wavelength absorption maximum; ϵ – extinction at λ_{abs} ; λ_{emi} – the wavelength of emission maximum; ϕ – the integral quantum yield of emission. ^c For solutions in DCE. ^d For solutions in DMF. ^e For solutions in DMSO.

strengthening of the acceptor capacity of the substituted 3-arylquinoxaline moiety (**1** < **3** \approx **4** < **6**, according to computations, Table 1), while the modification of the benzimidazole moiety seems to play only a minor role (*cf.* **1** and **5**). Replacement of the NH moiety of **1** by the NMe group in **2** increases steric strain in the *syn*-conformer, and this results in a qualitatively different conformational behavior: the (free) energies of the *syn*- and *anti*-conformers of the latter molecule are almost the same.

The position of the lowest-energy absorption band of the π – π^* transition in the spectra of **1–6** is influenced by both types of structural modifications of the BQs: the stronger the acceptor capacity of the 3-arylquinoxaline moiety or the donor capacity of the benzimidazole moiety, the larger is λ_{abs} (Table 2).

The position of the emission band is red-shifted by introduction of donor methyl groups in the case of **5**, but the strongest

effect is produced by introduction of NO_2 groups, which results in a maximal λ_{emi} value for **6** (Table 2). In the latter case the nitro group is located in the electron accepting moiety, and demonstrates much more flexibility than the remaining heterocyclic system (Fig. 1S, ESI[†]). *I.e.*, the NO_2 group presence induces pronounced geometry differences between S_0 and S_1 states and, hence, results in the largest Stokes shift among the whole series. Moreover, the presence of NO_2 groups seems to be a key factor responsible for exceptional sensitivity of the quantum yield of luminescence to HBs with solvent molecules (*vide supra*).

Thus, introduction of various substituents in donor or acceptor parts of the BQ molecules influences their conformational equilibria as well as their photophysical properties. Moreover, the good general agreement of the presented computations with the available experimental data regarding the above mentioned



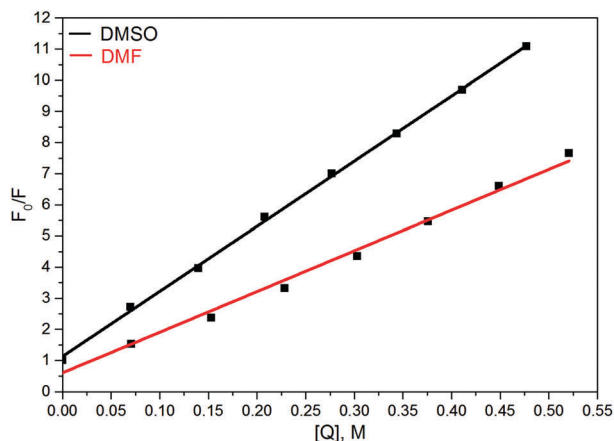


Fig. 7 Stern–Volmer plot of fluorescence quenching by DMF and DMSO for solution of compound **6** in DCE. F_0 and F are the fluorescence intensities in the absence and presence of quencher $[Q]$, respectively.

properties suggests that these quantum chemical methods can be of great help in rational tuning of absorption and emission characteristics of these compounds.

Conclusions

We have introduced a series of novel fluorophores, consisting of a 3-arylquinoxaline moiety and a benzimidazole moiety linked by a single bond. Conjugation between these moieties stabilizes the *syn*-conformers of the BQs relative to the *anti*-conformers, the former species being strongly thermodynamically preferable in the gas phase and in DCE solutions. *Anti*-conformers form stronger intermolecular HBs with DMSO and DMF than *syn*-conformers, and this difference decreases the stability of *syn*- relative to *anti*-conformers in these polar solvents. The above effect strengthens upon electronic excitation to the S_1 state and may cause an inversion of relative energetic stability of the two conformers in DMF and DMSO solutions. Thus, the absorption spectra of the BQs in all solvents and emission in DCE solutions are dominated by *syn*-conformers, while the possible presence of *anti*-forms may be taken into account for the interpretation of the emission spectra of DMF and DMSO solutions. Substituents introduced in various positions of the 3-arylquinoxaline and benzimidazole moieties influence conformational behavior and mutual energetic displacement of excited singlet states relative to their ground state. This in turn strongly affects their luminescence. Various substitution patterns produce very different effects on quantum yield of luminescence: from a moderate increase in polar DMSO and DMF relative to the DCE solutions to complete quenching of emission in these polar media. This tuneability of the BQ's luminescence and especially its sensitivity to environmental effects allows rational design of novel fluorophores of this family for various applications.

Acknowledgements

The financial support from scholarship programs of the German Academic Exchange Service (DAAD, personal ref. no. 91577762)

and of the government of the Republic of Tatarstan “Algarish” is gratefully acknowledged (SAK).

References

- 1 A. Kyrychenko, J. Herbich and J. Waluk, in *Tautomerism: Methods and Theories*, ed. L. Antonov, Wiley, Weinheim, 2013, pp. 49–78.
- 2 J. E. Kwon and S. Y. Park, *Adv. Mater.*, 2011, **23**, 3615–3642.
- 3 J. Herbich, M. Kijak, A. Zielinska, R. P. Thummel and J. Waluk, *J. Phys. Chem. A*, 2002, **106**, 2158–2163.
- 4 W. S. Yu, C. C. Cheng, Y. M. Cheng, P. C. Wu, Y. H. Song, Y. Chi and P. T. Chou, *J. Am. Chem. Soc.*, 2003, **125**, 10800–10801.
- 5 V. Vetokhina, K. Dobek, M. Kijak, I. I. Kaminska, K. Muller, W. R. Thiel, J. Waluk and J. Herbich, *ChemPhysChem*, 2012, **13**, 3661–3671.
- 6 J. Herbich, C.-Y. Hung, R. P. Thummel and J. Waluk, *J. Am. Chem. Soc.*, 1996, **118**, 3508–3518.
- 7 B. Golec, M. Kijak, V. Vetokhina, A. Gorski, R. P. Thummel, J. Herbich and J. Waluk, *J. Phys. Chem. B*, 2015, **119**, 7283–7293.
- 8 V. A. Mamedov, N. A. Zhukova, T. N. Beschastnova, A. T. Gubaidullin, A. A. Balandina and S. K. Latypov, *Tetrahedron*, 2010, **66**, 9745–9753.
- 9 V. A. Mamedov, A. A. Kalinin, A. T. Gubaidullin, A. V. Chernova, I. A. Litvinov, Y. A. Levin and R. R. Shagidullin, *Russ. Chem. Bull., Int. Ed.*, 2004, **53**, 164–175.
- 10 J. R. Lakowicz, *Principles of Fluorescence Spectroscopy*, Kluwer Academic, New York, 2nd edn, 1999, pp. 52–54.
- 11 TURBOMOLE V6.4 2012, a program package developed at the University of Karlsruhe and at the Forschungszentrum Karlsruhe GmbH from 1989–2007, and at TURBOMOLE GmbH since 2007; available from <http://www.turbomole.com>.
- 12 R. Bauernschmitt, M. Haeser, O. Treutler and R. Ahlrichs, *Chem. Phys. Lett.*, 1997, **264**, 573–578.
- 13 R. Bauernschmitt and R. Ahlrichs, *Chem. Phys. Lett.*, 1996, **256**, 454–464.
- 14 R. Bauernschmitt and R. Ahlrichs, *J. Chem. Phys.*, 1996, **104**, 9047–9052.
- 15 J. Wang, R. M. Wolf, J. W. Caldwell, P. A. Kollman and D. A. Case, *J. Comput. Chem.*, 2004, **25**, 1157–1174.
- 16 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 17 C. Adamo and V. Barone, *J. Chem. Phys.*, 1999, **110**, 6158–6170.
- 18 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 19 (a) A. D. Becke and E. R. Johnson, *J. Chem. Phys.*, 2005, **122**, 154104; (b) E. R. Johnson and A. D. Becke, *J. Chem. Phys.*, 2005, **123**, 024101; (c) S. Grimme, S. Ehrlich and L. Goerigk, *J. Comput. Chem.*, 2011, **32**, 1456.
- 20 S. Grimme, *Chem. – Eur. J.*, 2012, **18**, 9955–9964.
- 21 Y. Zhao and D. G. Truhlar, *J. Phys. Chem. A*, 2005, **109**, 5656–5667.
- 22 F. Weigend, F. Furche and R. Ahlrichs, *J. Chem. Phys.*, 2003, **119**, 12753–12762.



- 23 F. Weigend and R. Ahlrichs, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3297–3305.
- 24 F. Eckert and A. Klamt, *AIChE J.*, 2002, **48**, 369–385.
- 25 F. Eckert and A. Klamt, *COSMOtherm, Version C3.0, Release 12.01*, COSMOlogic GmbH & Co. KG, Leverkusen, Germany, 2012.
- 26 A. Schäfer, C. Huber and R. Ahlrichs, *J. Chem. Phys.*, 1994, **100**, 5829–5835; A. D. Becke, *Phys. Rev. A: At., Mol., Opt. Phys.*, 1988, **38**, 3098–3100; J. P. Perdew, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **33**, 8822–8824.
- 27 E. U. K. Gross, J. F. Dobson and M. Petersilka, in *Density Functional Theory II, Springer Series in Topics in Current Chemistry*, ed. R. F. Nalewajski, Springer, Heidelberg, Germany, 1996.
- 28 M. E. Casida, in *Recent Advances in Density Functional Methods*, ed. D. P. Chong, World Scientific, Singapore, 1995, vol. 1(27), pp. 155–192.
- 29 F. Furche, *J. Chem. Phys.*, 2001, **114**, 5982–5992.
- 30 S. Hirata and M. Head-Gordon, *Chem. Phys. Lett.*, 1999, **314**, 291–299.
- 31 A. D. Becke, *J. Chem. Phys.*, 1993, **98**, 5648–5652.
- 32 C. Lee, W. Yang and R. G. Parr, *Phys. Rev.*, 1988, **B37**, 785–789.
- 33 W. J. Hehre, R. Ditchfield and J. A. Pople, *J. Chem. Phys.*, 1972, **56**, 2257–2261.
- 34 P. C. Hariharan and J. A. Pople, *Theor. Chim. Acta*, 1973, **28**, 213–222.
- 35 T. Clark, J. Chandrasekhar, G. W. Spitznagel and P. v. R. Schleyer, *J. Comput. Chem.*, 1983, **4**, 294–301.
- 36 M. J. Frisch, J. A. Pople and J. S. Binkley, *J. Chem. Phys.*, 1984, **80**, 3265–3269.
- 37 S. A. Katsyuba, T. I. Burganov, E. E. Zvereva, A. A. Zagidullin, V. A. Miluykov, P. Lönnecke, E. Hey-Hawkins and O. G. Sinyashin, *J. Phys. Chem. A*, 2014, **118**, 12168–12177.
- 38 <http://www.chemcraftprog.com>.
- 39 It is well established that conjugation leads to a many-fold increase in the intensity of polarized lines, corresponding to the symmetric stretching vibrations of the bonds participating in conjugation. The effect is discussed in depth in, e.g.: P. P. Shorygin, *Russ. Chem. Rev.*, 1971, **40**, 367–392; E. D. Schmid and R. D. Topsom, *J. Am. Chem. Soc.*, 1981, **103**, 1628–1632.
- 40 A. V. Iogansen, *Spectrochim. Acta, Part A*, 1999, **55**, 1585–1612.
- 41 S. A. Katsyuba, M. V. Vener, E. E. Zvereva, Z. Fei, R. Scopelliti, J. G. Brandenburg, S. Siankevich and P. J. Dyson, *J. Phys. Chem. Lett.*, 2015, **6**, 4431–4436.
- 42 In contrast to S_0 state, critical point is easily localized on N–H...N intramolecular HB for *syn*-conformers in S_1 state, except for molecule **6**. Strengthening of the N–H...N intramolecular HB in S_1 state relative to S_0 state is also clearly seen from comparison of $\Delta N H_{syn-anti}$ values for S_1 and S_0 states in Table 1S (ESI[†]). Moderate difference of $\Delta N H_{syn-anti}$ values for S_1 and S_0 states (Table 1S, ESI[†]), pronounced non-coplanarity of B and Q moieties of *syn*-conformer of **6** in S_1 state (Fig. 1S, ESI[†]), as well as the absence of critical point on N–H...N bond, suggest the absence of intramolecular HB even in S_1 state of this molecule.
- 43 Present TDDFT computations seem to overestimate this effect, as S_1 states of H-complexes formed by DMSO with the BQs are found more energetically stable than in the experiment: e.g., $S_1 \rightarrow S_0$ vertical transition energy computed for the H-complex of DMSO with **1** amounts to 2.10 or 2.05 eV for *syn*- or *anti*-conformers, respectively, while 2.64 eV is found experimentally for solution of **1** in DMSO (Table 2).
- 44 M. Kijak, I. Petkova, M. Toczek, G. Viosna-Salyga, A. Zielinska, J. Herbich, R. P. Thummel and J. Waluk, *Acta Phys. Pol., A*, 2007, **112**, 105–119.
- 45 J. Hidalgo, A. Sánchez-Coronilla, M. A. Muñoz, C. Carmona and M. Balon, *J. Lumin.*, 2007, **127**, 671–677.
- 46 G. Wiosna, I. Petkova, M. S. Mudadu, R. P. Thummel and J. Waluk, *Chem. Phys. Lett.*, 2004, **400**, 379–383.
- 47 P. M. Froehlich and K. Nelson, *J. Phys. Chem.*, 1978, **82**, 2401–2403.
- 48 R. W. Ricci and J. M. Nesta, *J. Phys. Chem.*, 1976, **80**, 974–980.
- 49 P. Chowdhury, S. Panja, A. Chatterjee, P. Bhattacharya and S. Chakravorti, *J. Photochem. Photobiol., A*, 2005, **170**, 131–141.
- 50 A. Kyrchenko, I. Y. Sevriukov, Z. A. Syzova, A. S. Ladokhin and A. O. Doroshenko, *Biophys. Chem.*, 2011, **154**, 8–17.
- 51 J. Waluk and S. J. Komorowski, *J. Mol. Struct.*, 1986, **142**, 159–162.
- 52 J. Waluk and S. J. Komorowski, *Chem. Phys. Lett.*, 1987, **4**, 368–372.
- 53 I. Petkova, M. S. Mudadu, A. Singh, R. P. Thummel, I. H. M. van Stokkum, W. J. Buma and J. Waluk, *J. Phys. Chem. A*, 2007, **111**, 11400–11409.
- 54 S. K. Behera, G. Sadhuragiri, P. Elumalai, M. Sathiyendiran and G. Krishnamoorthya, *RSC Adv.*, 2016, **6**, 59708–59717.
- 55 X. Q. Cao, X. H. Lin, Y. Zhu, Y. Q. Ge and J. W. Wang, *Spectrochim. Acta, Part A*, 2012, **98**, 76–80.
- 56 A. Kanungo, D. Patra, S. Mukherjee, T. Mahata, P. R. Maulik and S. Dutta, *RSC Adv.*, 2015, **5**, 70958–70967.
- 57 A. Mishra, S. Chatterjee and G. Krishnamoorthy, *J. Photochem. Photobiol., A*, 2013, **260**, 50–58.
- 58 Y. Wu, P. V. Lawson, M. M. Henary, K. Schmidt, J.-L. Brédas and C. J. Fahrni, *J. Phys. Chem. A*, 2007, **111**, 4584–4595.
- 59 P. E. Kruger, P. R. Mackie and M. Nieuwenhuyzen, *J. Chem. Soc., Perkin Trans. 2*, 2001, 1079–1083.
- 60 S. Achelle, C. Baudequin and N. Plé, *Dyes Pigm.*, 2013, **98**, 575–600.
- 61 H.-J. Son, W.-S. Han, K.-R. Wee, D.-H. Yoo, J.-H. Lee, S.-N. Kwon, J. Ko and S. O. Kang, *Org. Lett.*, 2008, **10**, 5401–5404.
- 62 *Ca.* ten-fold decrease of quantum yield in DMSO solutions relative to CH_2Cl_2 solutions was reported for two 2-aryl,6-(aryleneethynylene)-1H-indoles: M. Flores-Jarillo, A. Alvarez-Hernandez, R. A. Vazquez-García, E. Arias, I. Moggio and J. R. Torres, *Dyes Pigm.*, 2016, **133**, 41–50; only much less pronounced effect was found for *N*-alkylated derivatives of the 2-aryl,6-(aryleneethynylene)-1H-indoles, which suggests an important role of intermolecular HB NH...O=S in the luminescence quenching.
- 63 *Ca.* 5-fold decrease of quantum yield in DMF solutions relative to CH_2Cl_2 solutions was reported for some



- benzimidazole-borondipyrromethenes: Z. Li, L.-J. Li, T. Sun, L. Liu and Zh. Xie, *Dyes Pigm.*, 2016, **128**, 165–169.
- 64 B. Valeur, *Molecular Fluorescence Principles and Applications*, Wiley-VCH Verlag GmbH, 2001, p. 381.
- 65 J.-M. Mewes, V. Jovanovic, C. M. Marian and A. Dreuw, *Phys. Chem. Chem. Phys.*, 2014, **16**, 12393–12406.
- 66 S. Gozem, F. Melaccio, A. Valentini, M. Filatov, M. Huix-Rottlant, N. Ferre, L. M. Frutos, C. Angeli, A. I. Krylov, A. A. Granovsky, R. Lindh and M. Olivucci, *J. Chem. Theory Comput.*, 2014, **10**, 3074–3084.
- 67 M. Huix-Rottlant, A. Nikiforov, W. Thiel and M. Filatov, *Top. Curr. Chem.*, 2016, **368**, 445–476.
- 68 B. Bursa, D. Wróbel, B. Barszcz, M. Kotkowiak, O. Valulyuk, D. Gryko, L. Kolanowski, M. Baraniak and G. Lota, *Phys. Chem. Chem. Phys.*, 2016, **18**, 7216–7228.
- 69 S. Hachiya, K. Asai and G. Konishi, *Tetrahedron Lett.*, 2013, **54**, 1839–1841.
- 70 We were not able to find reports of comparable rates of fluorescence quenching by DMSO. *Ca.* 6-fold decrease of fluorescence intensity of corrole dye substituted with 4-NO₂C₆H₄ group was found at addition of 20% of DMF to toluene solution of the dye (Fig. 5 in ref. 68). In case of compound **6** (Fig. 7) similar effect was achieved at addition of *ca.* 2% of DMF.
- 71 E. C. Lim, *J. Phys. Chem.*, 1986, **90**, 6770–6777.

