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Non-classical growth of water-redispersible spheroidal gold nanoparticles assisted by leonardite humate†

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The growth of gold nanoparticles (AuNPs) assisted by humate – a natural hyperbranched polyelectrolyte – was studied using *in situ* and *ex situ* techniques. The conditions for formation of almost monodisperse gold nanospheres of 13 ± 3 nm diameter with pronounced plasmonic properties were defined. A striking similarity was found between humate- and citrate-mediated growth of AuNPs: the formation of gold nanospheres involved rapid nucleation of gold seeds, slow growth within the intermediate agglomerates (visible as nanoworms after drying on a TEM grid), and rapid peptization into the final gold nanospheres. Both humate and citrate syntheses produced ultrastable gold sols with pronounced plasmonic properties. The substantial difference was slower kinetics of humate synthesis, 240 min *versus* 15 min for citrate, and the water redispersible properties of the humate-capped gold nanospheres after freeze drying, which was not seen with the citrate AuNPs. Theoretical calculations revealed a leading role of steric factors in the formation of intermediate aggregates of capped AuNPs at the stage of their slow growth in the case of both citrate and humate. We suggested that it was the polyelectrolyte nature of humate which enabled the water-redispersibility of humate- *versus* citrate-capped gold nanoparticles.

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Introduction

Gold nanoparticles (AuNPs) have attracted substantial attention due to their size- and shape-tunable optoelectronic properties and catalytic and biological stability, which offer a wide range of applications in nano- and biomedical technologies.^{1,2} AuNPs are frequently produced using toxic reducing agents, solvents, or stabilizers for tailoring unique

properties.^{3–9} As a result, a whole suite of greener synthetic routes has been developed, which implies the use of ecofriendly natural compounds and biopolymers.^{10–16} Among these are natural polyelectrolytes – humic substances (HS) – which have a very good safety profile and abundant resources (*e.g.*, lignites).¹⁷ The presence of multiple carboxyl, hydroxyl, and carbonyl groups results in remarkable macroligand, surface-active and redox properties of HS.^{18–22} They are used as capping and stabilizing agents for engineered nanoparticles including AuNPs.^{23–29} The use of different fractions of soil HS enabled fabrication of anisotropic gold nanoplates with enhanced optical properties.³⁰ To improve control over the size and shape of optically active AuNPs, synthetic HS were employed as reducing and capping agents.³¹ However, the use of humates for the preparation of monodisperse spheroidal gold nanoparticles similar in properties to citrate-AuNPs remains an unresolved problem.

In our previous studies we demonstrated how humic poly-anions derived from leonardite could be efficiently used to control the growth (sizes) and morphology of iron oxo-hydroxide nanoparticles.^{32,33} We have also shown the excellent colloidal stability and biocompatibility of humate-capped NPs.^{34,35}

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† Electronic supplementary information (ESI) available: Additional information about real-time UV-vis spectroscopy, fluorescence spectroscopy, Fourier transform ion cyclotron resonance mass spectrometry and theoretical calculations; HRTEM and XRD analysis of the obtained AuNPs; figures demonstrating the colloidal stability of the prepared AuNP sols; considerations on the non-redispersible behavior of citrate-capped AuNPs. See DOI: 10.1039/c6ce02149b



In this work, we used leonardite-derived humate as a capping agent to achieve control over the morphology, size distribution, plasmonic properties, and water redispersibility of AuNPs. Firstly, we defined the conditions for humate-mediated synthesis of spheroidal AuNPs with a narrow size distribution and enhanced optical properties; secondly, we studied the growth of spheroidal humate AuNPs as compared to citrate AuNPs; thirdly, we surmised the mechanisms which underlie the ultrastability and water-redispersibility of AuNPs capped with polyelectrolytic humate as opposed to low molecular weight citrate.

Experimental

Gold nanoparticle synthesis

Tetrachloroauric(III) acid was from Sigma-Aldrich (G4022 product, ACS reagent, $\geq 49.0\%$ Au basis). All other reagents were of analytical grade. All solutions were prepared using distilled water (Milli-Q RG, Millipore). Glassware and magnetic stirring bars were pre-washed with *aqua regia* (a mixture of hydrochloric (HCl) and nitric (HNO₃) acids at a molar ratio of 3 : 1).

Commercially available potassium humate from leonardite (a highly oxidized variety of lignites) (CHP) was used in our studies. It was characterized using elemental and functional group analyses, ¹³C NMR spectroscopy, and size exclusion chromatography (SEC). The H/C and O/C atomic ratios were 0.84 and 0.38, respectively, showing the highly aromatic character of the humic material used. This was also confirmed by the ¹³C NMR data on carbon distribution among the structural fragments (in % of the total C): 15% aliphatic C, 7% aliphatic *O*-substituted C, 47% aromatic C, 12% aromatic *O*-substituted C, 13% carboxylic and ester C, 5% ketone C. The functional group analysis yielded a total acidity of 5.3 ± 0.9 mmol g⁻¹ and carboxylic acidity of 2.9 ± 0.5 mmol g⁻¹. The SEC-determined weight-average molecular weight (M_w) was 10.0 ± 0.2 kDa, and the number-average molecular weight (M_n) was 5.4 ± 0.1 kDa.

An amount of the CHP sample was dissolved in MilliQ water and centrifuged to prepare a HS stock solution with a concentration of 212 mg L⁻¹.

The synthesis of AuNPs was carried out under pH control and the pH value was monitored and logged throughout the reaction using Expert-001 pH monitoring system and Expert-00x software (Econics-Expert Ltd., Russia). For this purpose, a working solution of CHP with a concentration of 85 mg L⁻¹ was placed in a foil-wrapped glass vessel and heated up to 70 °C. HAuCl₄ solution was added under vigorous stirring to give 5×10^{-4} M HAuCl₄. Immediately after HAuCl₄ addition, the pH was adjusted to 7.0 with 0.1 M KOH (no further pH adjustment was made). The reaction mixture was stirred continuously for 4 hours at 70 °C. The same synthesis procedure was carried out at initial pH values of 3, 5, and 9. The samples were designated as Au-HS-*x*, where *x* is the initial pH value of the synthesis. A working humate solution (85 mg L⁻¹) without HAuCl₄ addition was used in a blank experi-

ment. It was brought to 70 °C and stirred for 4 hours at pH 7.0.

An additional synthesis of humate-AuNPs at pH 7.0 was conducted under nitrogen atmosphere. The working CHP solution was purged by nitrogen gas for an hour until HAuCl₄ was added. Then, the pH value was adjusted to 7.0, and the reaction mixture was vigorously stirred for 4 hours at 70 °C. The sample was designated as Au-HS-7-N₂.

The synthesis of citrate-AuNPs is described in the ESI.†

Dry powders of humate- and citrate-stabilized AuNPs were prepared by freeze-drying in a Labconco FreeZone freeze dry system (-50 °C, 0.03 mbar pressure). For redispersion, a dried AuNP sample was dissolved in 0.1 M KOH under sonication (10 min); then the pH was adjusted to 7 using 0.1 M HCl.

Analytical methods

UV-visible absorption spectra were recorded using a Perkin-Elmer Lambda 950 spectrophotometer (transmission geometry, wavelength range of 200–1000 nm, 1 nm step). The plasmonic spectra of AuNPs were collected immediately after preparation and after storing in the dark for two months and two years. A custom-built set-up was used for real-time recording of UV-vis absorption spectra, as described in the ESI.†

The NP morphology was analyzed using a Carl Zeiss Libra 200 MC transmission electron microscope operating at 200 kV. The reaction mixture was dripped on lacey-carbon copper grids (SPI, USA) at different time-delays after HAuCl₄ addition, blotted and dried. Images were processed using open-source Gwyddion software,³⁶ and particle size distributions were calculated using 300+ particles of each sample.

Fluorescence spectra were recorded using a Fluoromax-4 spectrofluorometer (Horiba Jobin Yvon) at 266 nm excitation; excitation and emission slit widths were set to 5 nm. For fluorescence measurements, 25 ml of reaction mixtures were taken and dissolved in 3 ml of distilled water.

Fourier transform ion cyclotron resonance mass spectrometry (FTICR MS) analysis was conducted using facilities of the Emanuel Institute of Biochemical Physics of RAS (Moscow). The analytical conditions and data interpretation are described in the ESI.†

Results and discussion

Impact of pH on the size and shape of the humate-AuNPs

We have carried out humate-mediated syntheses of AuNPs under pH control (initial pH values were adjusted to 3, 5, 7, and 9) to define the conditions which favor the size and shape homogeneity of AuNPs. The change in pH altered the macroligand and redox properties of humic polyanions. At pH 3, the major portion of acidic groups is protonated, which minimizes the charge of humic polyanions and maximizes their coiling. An increase in pH brings about a growth in charge density due to deprotonation of carboxyl groups, which reaches 100% at pH 7. This pH also favors the maximum oxidizing ability of phenolic groups.³⁷ It should be



noted that speciation of Au^{III} is also governed by pH in aqueous chloride-containing solutions: an increase in pH yields more hydrolyzed auric species with lesser chemical reactivities.^{38,39}

Fig. 1 demonstrates the UV-vis absorption spectra, sizes and morphologies of the humate-AuNPs synthesized at different pH values. For comparison, the corresponding data for citrate-AuNPs are shown.⁴⁰ The closest similarity with citrate-AuNPs with respect to size, shape and optical properties (Fig. 1b) was seen for AuNPs synthesized at pH 7 (Fig. 1c and d): they had the smallest size (13 ± 3 nm), a spheroidal shape, and the narrowest SPR peak located at 524 nm. All other pH values (3, 5 and 9) brought about the formation of irregularly shaped AuNPs (Fig. 1e–g). The HRTEM image of a single AuNP reveals a crystalline structure whose interplanar distances (2.35 \AA) correspond to that of fcc metallic gold (Fig. 1d). HRTEM allowed visualization of numerous dislocations and twin boundaries in the synthesized AuNPs (Fig. 1d and S1†). The particle phase compositions were determined using XRD (Fig. S2†).

The appearance of ellipsoidal or rod-like AuNPs at pH 5 and 9 led to a red shift of the SPR peak and caused its broadening (Fig. 1a, e and f). The shoulder at 625 nm was observed

for the Au-HS-3 sample, which could be explained by the formation of triangular/hexagonal gold nanoprisms (Fig. 1g).

All Au-HS sols obtained in this study showed high colloidal stability: the position and intensity of the SPR peak did not change over two years (Fig. S4†). This corroborates with the reported findings on HS as an effective reducing and stabilizing agent for the production of gold sols.^{29–31} However, only AuNPs synthesized at pH 7 had a spheroidal shape and narrow size distribution closely resembling those of citrate AuNPs.^{40,41} To understand whether the resemblance of AuHS-7 and citrate-AuNPs was a “happy coincidence”, or both processes are governed by similar mechanisms, we have conducted detailed kinetic studies on humate-AuNP formation at pH 7 and compared them with the citrate synthesis.

Kinetics and chemistry of AuNP formation assisted by leonardite humate at pH 7

The TEM images of humate-AuNPs in the course of their formation at pH 7 were obtained *ex situ* and shown in Fig. 2a–c. The UV-vis spectra were recorded *in situ* and processed to calculate the maximum wavelength, FWHM, and area of the SPR peak of the humate-AuNPs. The corresponding data are

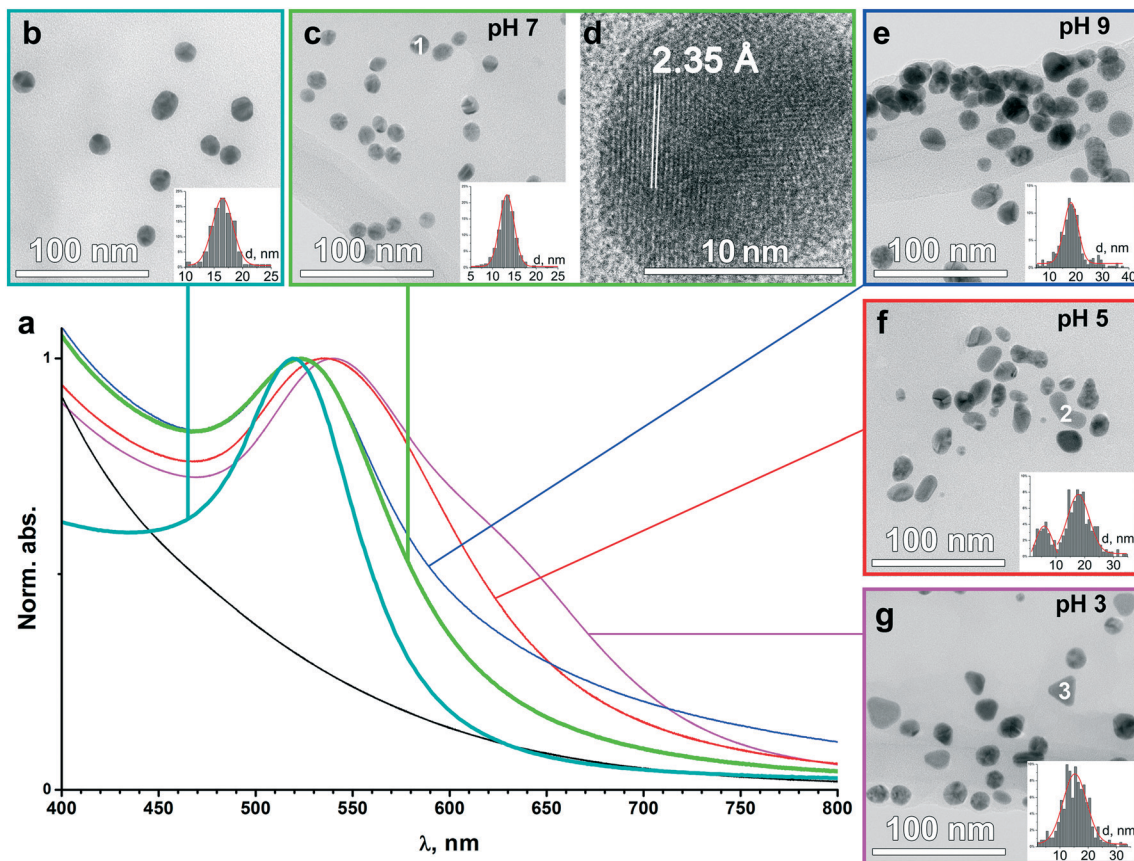


Fig. 1 Humate-AuNPs obtained at different initial pH values. (a) UV-vis absorption spectra: black – initial HS solution, dark cyan – citrate-AuNPs, green – Au-HS-7, blue – Au-HS-9, red – Au-HS-5, magenta – Au-HS-3. TEM micrographs: (b) citrate-AuNPs (0.25 mM HAuCl₄ + 1.75 mM Na₃C₆H₅O₇, 100 °C); (c) Au-HS-7; (d) magnified view of a single Au-HS-7 particle; the atomic interplanar distance of metallic gold is highlighted; (e) Au-HS-9; (f) Au-HS-5; (g) Au-HS-3. Three major types of humate-AuNP morphologies are shown with numbers: 1 – nanospheres (c); 2 – irregularly shaped NPs (f); 3 – prismatic nanoplatelets (g). Full-size graphs of particle size distributions for all Au-HS samples are given in the ESI† (Fig. S3).



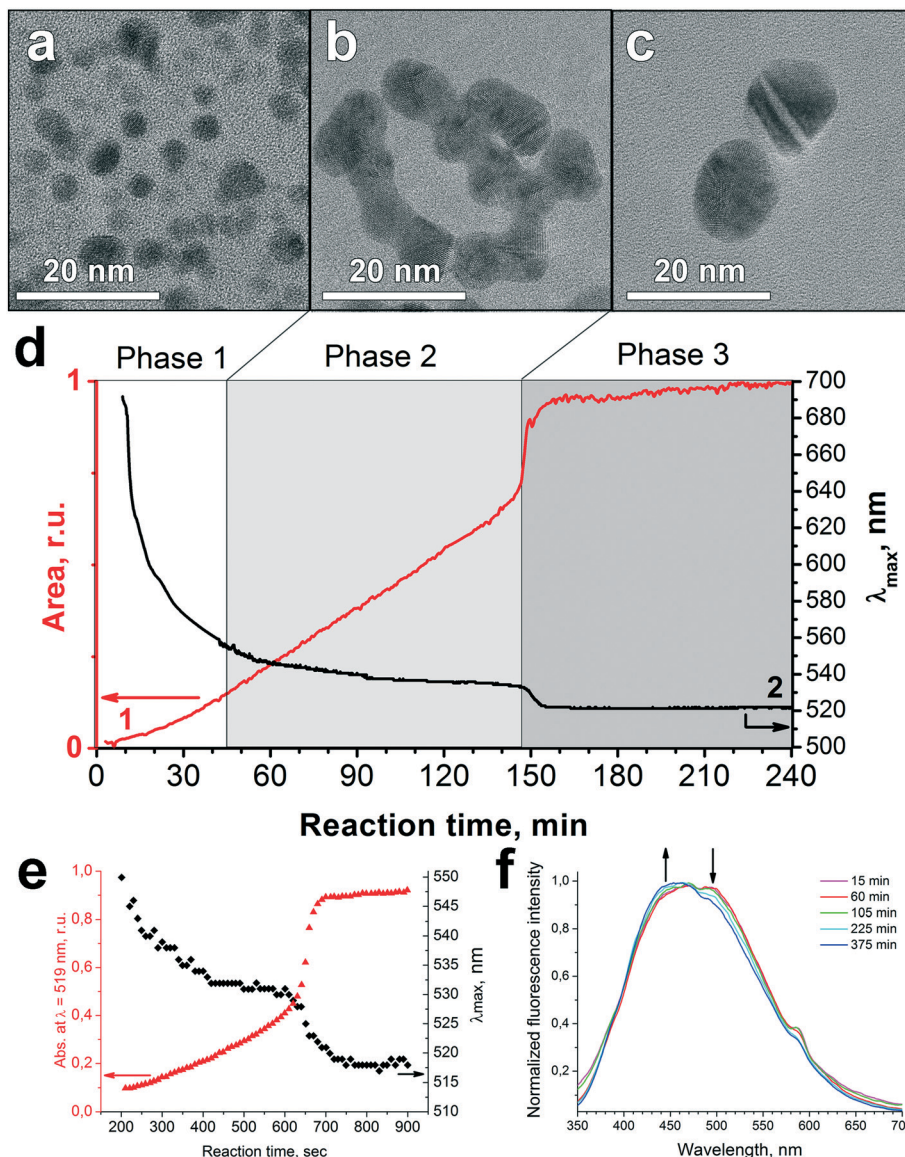


Fig. 2 Evolution of humate-AuNP formation at pH 7: TEM images of (a) AuNP seeds at the initial step, (b) worm-like intermediate aggregates, and (c) final spheroidal NPs. (d) Time evolution of SPR peak parameters (1 – area, 2 – λ_{\max}). (e) Time evolution of the SPR band during the reaction of 2.53×10^{-4} M HAuCl₄ with 1.70×10^{-3} M citrate at 100 °C (numerical data adopted from ref. 42). (f) Change in fluorescence spectra of HS over reaction time.

plotted in Fig. 2d and S5† (the latter includes FWHM and pH curves recorded during the synthesis). Similar SPR data reported for the citrate-AuNPs are given in Fig. 2e for comparison. The fluorescence spectra of the Au-HS reaction mixture were recorded over the entire reaction time (Fig. 2f) to monitor changes in humate optical properties.

Evolution of the SPR band during the reaction of HAuCl₄ with potassium humate revealed three stages in humate-AuNP formation as indicated in Fig. 2d. The first phase, nucleation, started immediately after HAuCl₄ addition to humate solution with the formation of numerous gold seeds of 3.0 ± 0.5 nm diameter (Fig. 3a). At this time some large dendritic aggregates (up to 150 nm) were formed as well (Fig. S6†), which could originate from local supersaturations upon the addition of concentrated HAuCl₄. These aggregates might

be responsible for the low-intensity SPR signal at 600–700 nm at the first minutes of the reaction. These large structures were not observed at later reaction times, probably due to dissolution or sedimentation.

The second phase was characterized by the appearance of a network of worm-like AuNP-structures (Fig. 2b). The latter seem to be artifacts of drying on the TEM grid, while in the reaction mixture the 5 ± 2 nm AuNPs did not merge within these agglomerates. However, their close proximity gave rise to coupling of plasmonic oscillations, which explains the position of the SPR peak at *ca.* 550 nm (Fig. 2d), while the expected SPR wavelength for 5 nm gold nanospheres in water is *ca.* 517 nm.⁴³ At this phase the SPR peak shifted slowly from 550 to 535 nm and its area increased steadily indicating the growth of humate-AuNPs over 145 min (Fig. 2d).



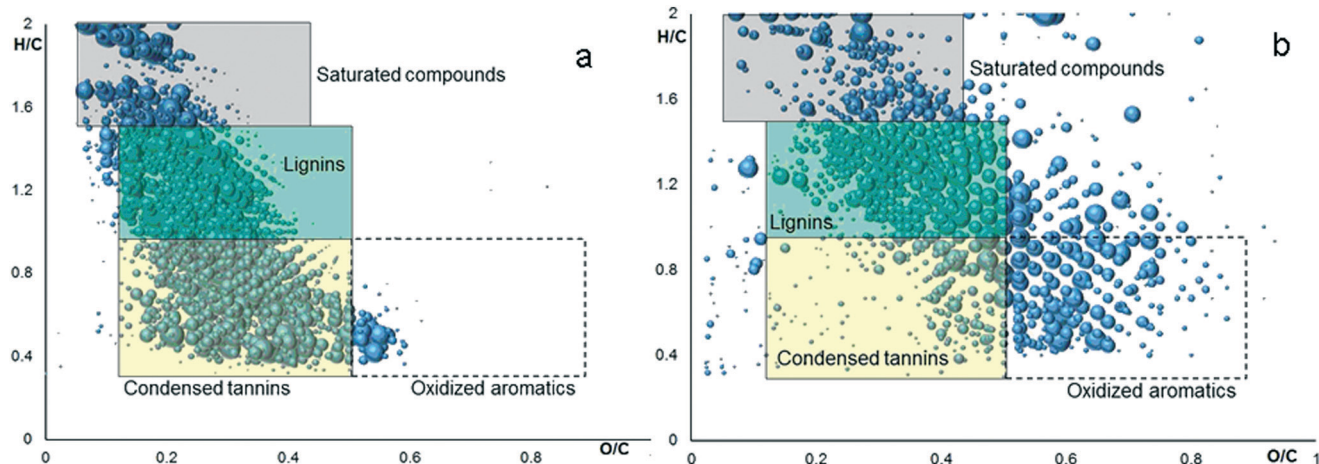


Fig. 3 Van Krevelen diagrams plotted from FTICR-MS identification of potassium humate which was obtained before (a) and after (b) its reaction with HAuCl_4 at pH 7 (AuHS-7).

After reaching 145 min, the SPR peak moved suddenly to 524 nm, which was followed by its significant narrowing and 30% increase in intensity (Phase 3). TEM imaging suggests that such abrupt and rapid changes were caused by peptization of the worm-like agglomerates into nearly monodisperse spheroidal AuNPs with a mean diameter of 13 ± 3 nm (Fig. 2c).

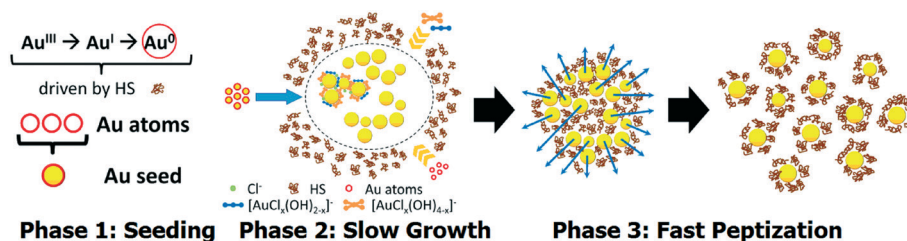
The evolution of plasmonic spectra of humate-AuNPs resembles closely the citrate-mediated synthesis shown in Fig. 2e.⁴² Both syntheses are characterized by a steady blue shift of λ_{max} over the reaction time until an abrupt change takes place in both the SPR wavelength and its intensity. This suggests that both syntheses may have similar particle growth mechanisms. Similarly to citrate, humate plays a role of acid-base buffer at pH 7, producing multidentate poly-anions due to deprotonation of carboxyl groups. These carboxylates are capable of ligand exchange with $[\text{AuCl}_{4-n}\text{OH}_n]^-$ species stabilizing them in solution. Simultaneously, the phenolic groups of humic macroligands act as reductants for $[\text{AuCl}_4]^-$ ions. In contrast to citrate, the oxidation of phenolic groups is a two stage reversible process yielding the phenoxyl radical. The rate of oxidation is much slower as compared to that of citrate due to the formation of stable semiquinone radicals. Ultimately, further oxidation of phenoxyl radicals would lead to irreversible cleavage of aromatic rings followed by formation of dibasic acids.

We used fluorescence spectroscopy for monitoring humate transformation during the growth of AuNPs. The normalized spectra are shown in Fig. 2f. They did not reveal any change in the humate fluorescence spectral band shape ($\lambda_{\text{exc}} = 266$ nm) during the first 60 minutes of reaction time. After that a

gradual decrease was observed in the relative intensity of the red edge fluorescence. These changes could not be attributed to a pH increase during the synthesis, which we checked by conducting fluorescence measurements at different pH values (Fig. S7[†]). Instead, they could be related to partial degradation of the conjugated system of fluorophores contributing to the red edge of HS fluorescence.⁴⁴ Direct evidence of partial HS oxidation was obtained using FTICR MS analysis of the humate before and after its reaction with HAuCl_4 (Fig. 3).

Fig. 3 shows FTICR MS identifications plotted in Van Krevelen diagrams, which establish the relationships between H/C and O/C atomic ratios for molecular constituents of the analyzed mixture. Color mapping relates the diagram areas to predominant chemical classes observed in these regions. This enables comparison of the molecular composition of the CHP before and after reaction with HAuCl_4 (Fig. 3a and b, respectively). Thus, the contribution of condensed phenolic constituents, which populate the yellow field of the diagram, decreased from 46% to 25% after reaction, whereas the content of oxygen rich aromatic compounds (white field) increased from 2% to 18%. This is indicative of oxidative transformation of humic phenolic constituents. The content of other humic components underwent very little change. This provides for stability of the humic matrix over the entire reaction time. The high-resolution views, references, and details of FTICR MS experiments are given in the ESI.[†]

The kinetic data allowed us to propose the following scheme of the three-phase growth of AuNPs mediated with humate which leads to the formation of stable sols of spherical AuNPs:



The given scheme was further backed up by considerations of the underlying chemistry including the redox chemistry of HS, and the reported data on the chemistry of HAuCl₄ at pH 7.^{38,39}

At stage 1, highly reactive species of [AuCl₄]⁻ are reduced by oxidizing phenolic units of humate to Au⁰ or Au^I which is accompanied by proton release causing a drop in the pH curve down to 5.8 within the first 15 min (Fig. S5†). At the same time the hydrolysis of [AuCl₄]⁻ (also lowering the pH value) gives much less reactive species of [AuCl_{4-n}(OH)_n]⁻ while the formation of complexes with humate ions slows down new Au⁰ generation, turning the reaction to a surface growth mode.

At stage 2, the metallic gold surface of the formed seeds absorbs both Au^{III} and Au^I species and autocatalyses their reduction and Au^I disproportionation.⁴⁵ Complexation with oxidized HS can also facilitate Au^I disproportionation similarly to dicarboxyl acetone in citrate-mediated synthesis.^{40,46} Over this time random flocculation of Au⁰ seeds and particles occurs, yielding intermediate agglomerates visualized by *ex situ* TEM as worm-like structures (Fig. 2b).

At stage 3, peptization of AuNPs observed in the humate-mediated system after 145 minutes is similar to that described by Biggs *et al.*⁴⁷ and Wall *et al.*⁴⁸ for citrate-gold surface interactions: they detected a change from attractive to repulsive interactions at the gold surface resulting from a gradual decrease of auric ion concentration and accumulation of citrate in the electrical double layer of the particles. In our system, formation of a large area of the gold surface with a positive charge was reflected in the pH curve as a slight decrease which could result from the sorption of OH⁻ anions (Fig. S5†).

Of interest is that the widely used citrate-mediated growth of gold nanospheres is also non-classical and involves agglomeration of the primary gold seeds and particle growth within the intermediate agglomerates.⁴⁹ The appearance of intermediate worm-like structures on *ex situ* TEM images was reported elsewhere.^{39,50,51} At the same time, some SAXS studies did not detect formation of the aggregates during the citrate-mediated synthesis.^{52,53}

It should be stressed that the non-classical growth of AuNPs seems to be feasible only within a certain “pH window” in the case of both citrate and humate. Ji *et al.*³⁹ observed intermediate aggregation only in the pH range from 3.7 to 6.5. Our observations show higher pH values: from 6.5 to 8.0. We surmised that in the case of humate, the pH drop caused by Au(III) hydrolysis was compensated for by humate decarboxylation (Fig. S5†), which sustained the higher pH values. The humate decarboxylation might be triggered by oxidation of its alcoholic groups to β-keto acids by dissolved oxygen.⁵⁴ We checked this hypothesis by conducting Au-HS-7 synthesis under nitrogen atmosphere. Indeed, in the absence of oxygen we observed a pH drop to 4.85 by the time of reaction completion (Fig. S9†), which yielded randomly shaped AuNPs (Fig. S10†). The obtained results corroborate well with the findings on the important role of dissolved oxygen in re-

dox reactions and particle growth processes.⁵⁵ The remarkable impact of a N₂ atmosphere was also reported for radiolytic synthesis of citrate AuNPs.⁵⁶

Further, we undertook theoretical studies on gold colloidal stability during the humate synthesis to get deeper insight into the peptization mechanism of intermediate AuNP agglomerates.

Theoretical calculations on the colloidal stability of AuNPs stabilized by humate and citrate

Chow and Zukoski⁴⁹ developed a model for gold sol formation in citrate-mediated synthesis incorporating colloidal stability. They suggested the formation of intermediate fluffy flocs where direct contact and irreversible aggregation of gold subunits are limited by a steric barrier produced by adsorbed ions. This is in accordance with *in situ* SAXS studies,^{51–53} which could not detect merged AuNP aggregates in a reaction mixture, leading the authors to conclude that the merging of NPs within dendrites, worm-like intermediates, *etc.*, observed by *ex situ* TEM^{39,50,51} is a result of a drying process. To define the factors providing for reversible AuNP flocculation and stabilization of intermediate aggregates with remaining interparticle separation, the pairwise particle interaction potential of AuNPs was calculated theoretically.

For this purpose, we considered the interaction of two identical charged spherical AuNPs with a radius a . Let r be the distance between their centers; thus $h = r - 2a$ is the shortest distance between their surfaces. The total energy of interaction of colloidal NPs (V_{tot}) expressed in $k_{\text{B}}T$ units (k_{B} and T are the Boltzmann constant and absolute temperature, respectively) is composed of three terms:

$$V_{\text{tot}} = V_{\text{vdw}} + V_{\text{el-st}} + V_{\text{solv}}$$

In accordance with the classical DLVO theory,⁵⁷ the first two terms correspond to van der Waals attraction and electrostatic repulsion, respectively, while the third one accounts for non-DLVO solvation forces. The attraction potential (V_{vdw}) is given by the Hamaker expression for spherical particles⁵⁸ where we used the reported value of the gold/water/gold Hamaker constant $A = 2.5 \times 10^{-19}$ J measured by AFM.⁵⁹ We used the Ohshima expression^{60–62} for calculation of electrostatic repulsion potential ($V_{\text{el-st}}$), since the condition $a/r_{\text{D}} < 5$ was fulfilled. Here a is the particle radius assumed to be 8.5 nm (estimated from our TEM results, Fig. 2), and r_{D} is the Debye radius equal to 1.8 nm for the ionic strength of the citrate-based reaction mixture of 2.5×10^{-2} M. The Bjerrum length l_{b} was set to 0.6 nm because the syntheses were performed at $T = 350$ K. At particle separations lower than a few nanometers, solvation forces may substantially influence interparticle interaction and play a key role in dispersion stabilization. We approximated the steric repulsion of soft adsorbed layers by a smooth repulsive potential V_{solv} instead of the infinite vertical potential wall, which was used by Chou and Zukoski.⁴⁹ This allowed us to get both primary and



secondary minima at the interaction potential curve. Thus, the last term calculated in the framework of Derjaguin approximation ($h \ll a$)⁵⁷ reads

$$V_{\text{solv}} = 2\pi \frac{B_s a l}{k_B T} \exp\left(\frac{h}{l}\right)$$

with the decay length $l = 0.5$ nm and parameter $B_s = 22 \times 10^{-3}$ J m⁻². Full mathematical expressions and corresponding assumptions are provided in the ESI†

Given that humate (or citrate) ions replace gold ions at the gold–water interface and cause an increase in the surface potential ψ_0 , a model was developed for the qualitative explanation of how particle surface potential affects the dispersion stability. Fig. 4 shows a set of pairwise interaction potentials which correspond to different ψ_0 values.

At medium ψ_0 values (e.g. 35 mV, highlighted in cyan in Fig. 4), the secondary minimum arises at the potential curve, and particles are held closely together with a characteristic distance of 2–3 nm corresponding to the minimum position. A rather high activation energy for transition between secondary and primary minima prevents irreversible particle aggregation. This might correspond to TEM images related to Phase 2 in Fig. 2. According to the theoretical analysis, the secondary minimum at potential curves exists even in the case of non-charged particles (near-zero ψ_0). However, the activation energy for transition between secondary and primary minima became low ($\sim 2k_B T$). This should lead to irreversible flocculation of the particles. However, this is not the case for a real reaction mixture since the AuNP surface is always covered by either gold chloride complex anions or by citrate/humate anions.

The further growth of ψ_0 strengthens the repulsion of particle double layers; the secondary minimum shifts to higher h values, becomes shallower, and gradually disappears (compare the curves for 40, 50, and 70 mV highlighted in Fig. 4 in green, yellow, and red, respectively). The thermal motion of particles allows them to abandon this shallow potential well, and they rapidly become redispersed again (compare with Phase 3 in Fig. 2). Conceptual results of calculations are virtually independent on l and B_s values (see the ESI†). The

growing surface potential ψ_0 has an impact on the particle stability (and pairwise potentials) similar to that in the case of decreasing salt concentration.⁶³

Direct calculation of the potential interaction curves in the case of humate-mediated synthesis is more complicated. Firstly, humates are polyelectrolytes of high polydispersity. The exact charge distribution of humate molecules is unknown, which hinders proper calculation of solution ionic strength. Secondly, both electrostatic and solvation terms are valid for the solutions of low-molecular-weight electrolytes, while the behavior of polyelectrolytes in solution is much more complex. Nevertheless, the estimated ionic strength of the humate-based reaction mixture was quite similar to that for the citrate-mediated synthesis. Thus, the qualitative estimates of formation and stabilization of intermediate AuNP aggregates (Fig. 4) ought to be applicable for both citrate and humate-mediated syntheses. Humate sorption on the AuNP surface is enhanced due to multi-site cooperative interactions which contribute to the aggregates' disassembly. These steric effects result in ultra-stable AuNP sols (Fig. S4†).

The above analysis refers to the pairwise interaction of colloidal NPs. In the real system, particle agglomerates with several neighbors are observed. Simple evaluation (similar to the case of liquid–vapor equilibrium) can be done in order to estimate the depth of the secondary minimum needed to make the aggregated state of AuNPs more favorable as compared to the separated one. The equality of chemical potentials of the separated and aggregated particles results in energy gain upon flocculation of about $k_B T \ln(n_{\text{agg}}/n_{\text{free}})$ per particle, where n_{agg} is the NP concentration in agglomerates, and n_{free} is the concentration of the separated NPs in solution. Since the $n_{\text{agg}}/n_{\text{free}}$ ratio is $\sim 10^6$ (roughly estimated from the aggregates' morphology and the presence of some non-aggregated AuNPs even at the early stages of synthesis as observed by TEM (Fig. 2)), we obtain $V_{\text{min}} \sim 14k_B T$. If the aggregates are large enough and each particle has several neighbors ($n \sim 10$), this value corresponds to a rather mild energy gain of about a few $k_B T$.

The obtained conceptual results are in reasonable agreement with the experimental data. Of particular importance is the appearance of the secondary minimum at the interparticle interaction potential curve at separations of 2–3 nm, providing rationalization for reversible flocculation of AuNPs and intermediate aggregate stabilization with no particle fusion in both humate- and citrate-mediated syntheses. The van der Waals attraction of particles trapped in this minimum is equilibrated both by a repulsion of overlapping electric double layers and a steric repulsion of adsorbed moieties and hydration shells, preventing particles from irreversible flocculation. Simple DLVO calculations without taking into account this steric component would not reveal the existence of the secondary minimum at distances of a few nanometers. At the same time it is the secondary minimum that facilitates stabilization of non-fused particle aggregates producing dendritic and worm-like structures upon drying as observed by TEM.

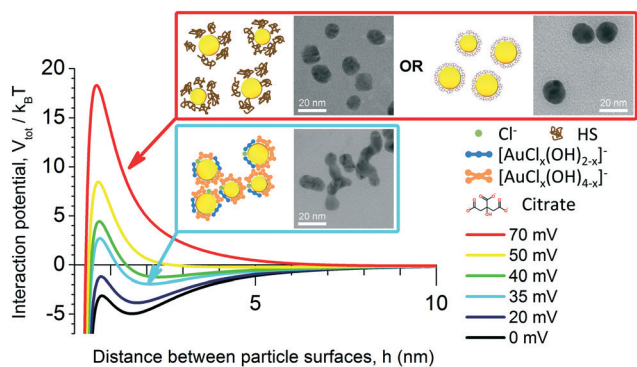


Fig. 4 Pairwise AuNP interaction potentials at different values of surface potential ψ_0 .



The obtained results on the leading role of steric factors in the stabilization of aqueous sols of capped AuNPs provide a reasonable explanation for the water-redispersibility of humate-AuNPs which was examined in the experiments described below.

Water-redispersible properties of humate-stabilized gold nanoparticles

We have freeze dried stable sols of humate-capped AuHS-7 and citrate-AuNPs. Both sols produced powders, which were redispersed in alkaline aqueous solution upon sonication. The resultant solutions are shown in Fig. 5a and b.

The humate-AuNPs produced a colloidal suspension with typical red color (Fig. 5a), whereas the citrate-AuNPs did not redisperse and gave a gray color to the solution (Fig. 5b). This is indicative of irreversible flocculation of citrate-AuNPs during freeze-drying, which was not the case for the humate-AuNPs. We have measured the UV-vis absorbance spectrum of redispersed humate-AuNPs, obtained their TEM image and compared them to those for the initial sols as shown in Fig. 5c and d. It can be seen that the redispersed humate-AuNPs had the same size distribution and SPR-peak location and intensity as the initial aqueous sol of these AuNPs.

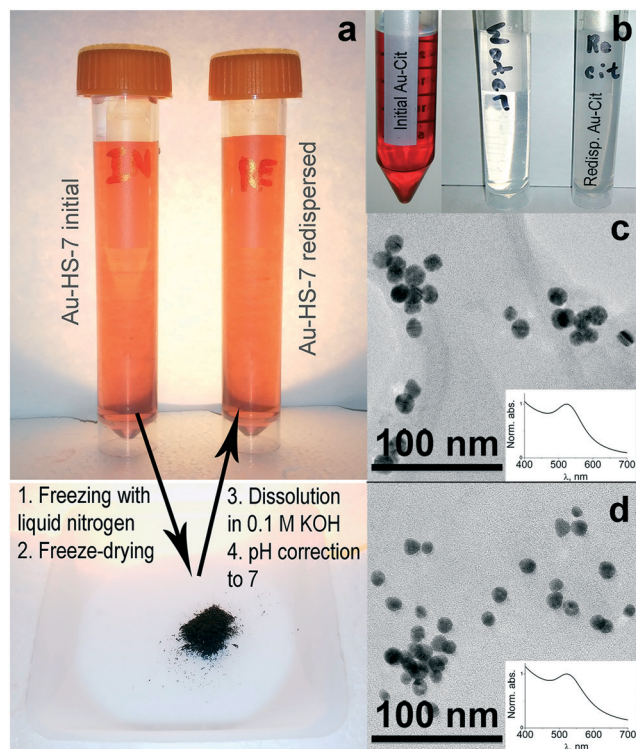


Fig. 5 Redispersion of the humate- and citrate-capped AuNPs in water: (a) humate-capped AuNP sols before and after redispersion and (b) citrate-capped AuNPs before and after redispersion. (c and d) TEM images and SPR peaks of the initial and redispersed humate-capped AuNPs (Au-HS-7 sol), respectively.

The observed difference between humate- and citrate-AuNPs could be connected to the polyelectrolytic nature of humate. Macromolecular humate will remain at the metallic surface during freeze-drying, whereas low molecular weight citrate will sublime. The high hydration affinity of humic macroligands could bring about water-redispersibility of humate-AuNPs.

Conclusions

The non-classical growth of spheroidal gold nanoparticles (AuNPs) was induced by humates, which are natural hyperbranched polyelectrolytes with carboxylic and phenolic functionalities. Striking similarities between the humate- and the citrate-mediated growth of spheroidal AuNPs have been revealed. Theoretical calculations of pairwise particle interaction potentials revealed the appearance of the secondary minimum at the particle separation of 2–3 nm which explained the formation of stable intermediate aggregates in the case of both citrate and humate and their follow-up disintegration upon an increase of the charge surface caused by sorption of citrate or humate. These results do not support the recently published opinion of Wuthschick *et al.*,⁵³ who referred to citrate-mediated formation of spheroidal monodisperse AuNPs as a “fortunate coincidence created by a favorable interplay of several chemical and physicochemical processes”. We believe that in-depth USAXS comparative studies of the citrate and humate mediated growth of spheroidal plasmonic AuNPs, which we plan in the nearest future, might shed new light on the mechanism of AuNP growth mediated by low- and high-molecular weight carboxylic ligands. They will provide more precise tools for directed design of HS-based gold nanomaterials with tailored properties. A unique feature of the proposed synthesis is the formation of water-redispersible AuNPs, which has never been reported for this system before. The polyelectrolyte nature of humate was suggested to be a key factor contributing to strong sorption of humate at the water-metallic gold interface, which enabled water-redispersion of the humate-capped AuNPs. This opens a way to a much wider application of humate-AuNPs in biomedical diagnostics and environmental sensing technologies. To demonstrate the scope of these applications is our immediate goal.

Abbreviations

AuNP	Gold nanoparticle
CHP	Coal humate isolated from Powhumus™
DLVO	Derjaguin–Landau–Verwey–Overbeek (theory)
FT ICRMS	Fourier transform ion cyclotron resonance mass spectrometry
FWHM	Full width at half maximum
HRTEM	High-resolution transmission electron microscopy
HS	Humic substances
MW	Molecular weight
NMR	Nuclear magnetic resonance



NP	Nanoparticle
pK	Dissociation constant
PSD	Particle size distribution
SAXS	Small-angle X-ray scattering
SEC	Size exclusion chromatography
SERS	Surface enhanced Raman scattering
SPI	Structure Probe, Inc.
SPR	Surface plasmon resonance
TEM	Transmission electron microscopy (microscope)
UV-vis	Ultraviolet-visible (range of electromagnetic spectrum)
XRD	X-ray diffraction

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