The reaction of the bis(pentalene)dititanium complex Ti₂(μ-η⁵,η⁵-Pn)₂ (Pn = C₈H₈(1,4-SiPr₃)₂) with the N-heterocyclic carbene 1,3,4,5-tetramethylimidazol-2-ylidene results in intramolecular C–H activation of one of the Pr methyl groups of a Pn ligand and formation of a "tucked-in" bridging hydride. The "tuck-in" process is reversed by the addition of hydrogen, which yields a dihydride featuring terminal and bridging hydrides.

Group IV metalloocene chemistry has been instrumental in the development of organometallic chemistry, elucidating fundamental aspects of bonding and reactivity, especially via the synthesis, isolation and study of low valent metalloocene complexes. In this context, the synthesis and isolation of such complexes bearing hydride ligands has been important; for example, such low or mixed valence hydride complexes have been shown to promote or be involved in the fixation of N₂ to NH₃. In this communication, we present the first examples of the synthesis of bridging titanium hydrides under non-reducing conditions, via the reaction of bis[pentalene]dititanium complex Ti₂(μ-η⁵,η⁵-Pn)₂ (Pn = C₈H₈(1,4-SiPr₃)₂) with an N-heterocyclic carbene and subsequent hydrogenolysis.

We have previously reported on the reactivity of Ti₂(μ-η⁵,η⁵-Pn)₂ (Pn = C₈H₈(1,4-SiPr₃)₂) (1) towards a variety of small molecules and π-acceptor ligands. In order to gain a better insight into the reactivity of (1), we decided to study its interaction with strong σ-donor ligands. When (1) was treated with an excess of PMe₃ no reaction was observed. However, addition of 1,3,4,5-tetramethylimidazol-2-ylidine (2) to (1) in toluene at 0 °C resulted in an immediate colour change from crimson red to a dark pine green (Scheme 1). The formulation of the new complex (3) as a "tucked-in" hydride resulting from C–H activation of one of the Pr methyl groups of a Pn ligand was initially confirmed by NMR. In particular, the ¹H-NMR spectrum showed two inequivalent pentalene ring environments, a sharp singlet at −7.91 ppm for the bridging hydride (T₁ 479 ms), with one of the diastereotopic protons of the "tucked-in" CH₂ group appearing as an overlapping dd at −2.91 ppm whilst the other was largely obscured by the complex aliphatic region of the spectrum at ca 1.37 ppm; the coordination of the NHC was confirmed by the observation of a peak at 197.78 ppm in the ¹³C{¹H}-NMR spectrum.

The structure of (3) was confirmed by X-ray diffraction and is shown in Fig. 1. The NH complex coordinates to one of the Ti centres (Ti2 in Fig. 1), while one of the methyl groups on a TIPS substituent has been cyclometallated on the other Ti centre (Ti1 in Fig. 1) with concurrent formation of a bridging hydride. The Ti–Ti bond has been retained but lengthened to 2.5610(8) Å (from 2.399(2) Å in (1)) and is typical of a single bond. The Ti–C(carbene) bond (2.300(2) Å; Ti2–C1 in Fig. 1) is within the range of 2.2–2.35 Å reported for other Ti–NHC complexes. The Ti–H bond lengths (i.e. Ti2–H1: 1.72(3) Å, Ti1–H1: 1.79(3) Å) are identical within esd's and are similar to previously reported monomeric and dimeric titanium hydrides as well as Ti(µ) alumino- and borohydrides, although it has to be noted that, probably due to the topology of the hydride ligand in (3), these Ti–H bond distances fall at the shorter end of the known range. Due to this unique topology, the Ti–H–Ti bond angle (93.4(13)°) approaches a right angle and is the most acute ever observed in dimeric titanium hydrides.

**Scheme 1** Synthesis of a new syn-bimetallic hydride titanium cluster.
Hydrogenolysis of (3) to afford (4)

Fig. 1 ORTEP diagram of the molecular structure of (3) displaying 50% probability ellipsoids. iPr groups omitted for clarity.

It is also worth noting that the addition of (2) to (1) results in the formal oxidation of the two Ti centres (i.e. from +2 to +3), and employing the CBC model each Ti has a count of 18 e (16 e in (1)).

With a view to synthesising a new hydride derivative via σ-bond metathesis of the Ti–CH₂ bond in (3) with dihydrogen, 2 bar of H₂ was added to a C₇D₈ solution of (3) in an NMR tube. This indeed resulted in clean conversion of (3) (100% spectroscopic yield) to a new complex (4) (Scheme 2).

Scheme 2 Hydrogenolysis of (3) to afford (4)

Compared to the ¹H NMR spectrum of (3), (4) displays a new, broader hydride peak (Δν₁/₂ = 29 Hz) at −8.82 ppm at room temperature, whilst the signal for the “tucked-in” CH₂ group has disappeared completely; the NHC is still coordinated (¹³C[⁴H] δ 198.27 ppm). Removal of the H₂ overpressure by freeze–thaw-degassing showed that (4) is persistent in solution, although some regeneration of (3) was observed (Scheme 2). Addition of H₂ to a solution of (3), via a Toeppler pump, showed that for the conversion of (3) to (4) to occur quickly (minutes) 5 eq of H₂ are required (when 1–2 equivalents of H₂ were added, complete conversion to (4) occurred after ca. 1 week). The rate of reaction was also found to be pressure dependent: when (3) was exposed to an atmosphere of 10% H₂ in N₂ at 1.5 bar but in an amount corresponding to only 1 equivalent of H₂ the reaction was again complete in minutes.

Variable temperature ¹H NMR studies showed that the broad hydride peak at −8.82 ppm in (4) becomes fully resolved into a doublet at 0 °C (with no further change below that temperature and down to −70 °C) with a T₁ of 310 ms, with the concomitent appearance of a second doublet centred at 2.17 ppm (T₁ 336 ms), which is too broad to be observable at room temperature (Fig. 2); these two signals are related by a coupling constant of J₁HH = 11 Hz. EXSY spectroscopy (in both the presence and absence of an H₂ overpressure) confirmed that these two protons exchange at 30 °C while at 0 °C the process is quenched. Thus the peak at −8.82 ppm is assigned to the bridging hydride in (4) and that at 2.17 ppm to the terminal one (Scheme 2).

Initial attempts to crystallise (4) by standard methods (i.e. removal of volatiles and recrystallisation) were frustrated by the preferential isolation of crystalline (3) (as it is less soluble than (4)) with the mother liquor consisting of a mixture of (3) and (4) (ca. 20:80 by NMR), due to the partial reversibility of the reaction. However, the solid state molecular structure of (4) (Fig. 3) was eventually determined from single crystals grown by cooling slowly a freshly prepared solution of (4) at −78 °C under an overpressure (1.5 atm) of H₂ and confirms the spectroscopic assignment.

The Ti–C(carbene) bond length in (4) is 2.291(4) Å and is identical to that found in (3). On the other hand, the Ti–Ti bond is slightly shortened in (4) from 2.5610(8) Å in (3) to 2.5413(8) Å possibly due to the negligible steric requirements of the terminal hydride ligand. The Ti–H bridging bond distances (Ti1–H1 = 1.84(5) Å; Ti2–H1 = 1.79(5) Å) in (4) are similar within esd’s and compare with the ones found in (3); the same applies to the Ti–H(terminal) (i.e. Ti2–H2 = 1.74(4) Å in Fig. 2) bond length. The Ti1–H1–Ti2 bond angle in (4) again approaches 90° (89(2)°) and is very similar to that found in (3).

When (3) was treated with an excess of D₂ (5 eq.), the formation of (4-D) was observed, but deuterium was found to be only incorporated in the hydridic positions, and not in the new Me group derived from the previously “tucked-in” CH₂ group (confirmed by ²H-NMR, DEPT-135 and gHSQC). Hence the reaction of (3) with H₂ to form (4)
Table 1  Electronic binding energies (\(\Delta E\)) and Gibbs energy changes (\(\Delta G\)) for the reaction Ti\(_2\)Pn\(_2\) + NHC = Ti\(_2\)Pn\(_2\)NHC

<table>
<thead>
<tr>
<th>Ti pentatene dimer</th>
<th>NHC</th>
<th>(\Delta E) (eV)</th>
<th>(\Delta G^a) (kJ mol(^{-1}))</th>
<th>Ti–Ti–C (Å)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ti(_2)(C(_8)H(_6))(_2)</td>
<td>C(_3)H(_4)N(_2)</td>
<td>-1.31</td>
<td>-66</td>
<td>100</td>
</tr>
<tr>
<td>Ti(_2)(C(_8)H(_6))(_2)</td>
<td>C(_3)Me(_2)N(_2)</td>
<td>-1.26</td>
<td>-65</td>
<td>114</td>
</tr>
<tr>
<td>Ti(_2)Pn(_2)</td>
<td>C(_3)Me(_2)N(_2)</td>
<td>-0.48</td>
<td>+21</td>
<td>127</td>
</tr>
</tbody>
</table>

\(a\) Calculations are on gas phase species which leads to an overestimation of the entropy change when there is a change in the number of species.

does not go via \(\sigma\)-bond metathesis (which would lead to D incorporation in the Me group in the reaction with D\(_2\)). Hence the formation of (4) (and also (3)) was probed computationally (ADF:BP/TZP: details are given in the ESI\(^*\)). Preliminary studies suggested that steric was important in determining the reaction energies. For example energy of binding an NHC to a Ti\(_2\)Pn\(_2\) dimer depended critically on the substituents. Introduction of the methyl substituents on the NHC made very little difference to its binding energy. Calculations are on gas phase species which leads to an overestimation of the entropy change when there is a change in the number of species. The degree to which the tuck-in reaction was favoured in the absence of an NHC was also investigated. The formation of Ti\(_2\)Pn\(_2\)(Pn\(_3\)-H)(μ-H) from Ti\(_2\)Pn\(_2\) was calculated to have \(\Delta E = -0.03\) eV and \(\Delta G = 11\) kJ mol\(^{-1}\). However, the energies of the observed tuck-in reaction with the methylated NHC present were calculated to be \(\Delta E = -0.69\) eV and \(\Delta G = 9\) kJ mol\(^{-1}\). Thus the presence of the base improves the energetics of the tuck-in reaction. The significant entropy disadvantage in the gas phase would be lessened in solution.

The HOMO of 3 (Fig. 4) shows a Ti–Ti \(\sigma\)-bond. The calculated Ti–Ti distance is 2.56 Å in excellent agreement with experiment. The Ti–H distances are 1.83 Å and the angle at the bridging hydrogen 89°. Such discrepancies from the experimental values are not unusual when comparing distances to bound hydrogen between theory and X-ray diffraction experiments.

Addition of H\(_2\) to 3 to form 4 is calculated to have reaction energies \(\Delta E = -0.77\) eV and \(\Delta G = -31\) kJ mol\(^{-1}\). The calculated Ti–Ti distance for 4 is 2.54 Å reproducing the shortening from 3 found experimentally. The Ti–H(terminal) distance is 1.74 Å, the Ti–H(bridging) distances 1.81 and 1.82 Å and the Ti–H–Ti angle unchanged at 89°.

A transition state for this reaction was modelled using just one SiPn\(_3\) substituent on one of the pentatene ligands and C\(_3\)H\(_4\)N as the NHC for computational efficiency. The free energy of activation was estimated as 84 kJ mol\(^{-1}\) for such a system. The transition state structure is shown in Fig. 5.

The Ti distance to the previously bridging H is 3.53 Å and the Ti distances to the reacting H\(_2\) are 2.74 and 3.24 Å, the H–H distance being 0.76 Å. Such a geometry indicates that the tuck-in process is reversed before complete H\(_2\) addition, consistent with the lack of deuterium incorporation into iPr groups and the conclusion that \(\sigma\) bond metathesis is not in play. It may be that the steric compression induced by the mere approach of the H\(_2\) molecule is sufficient to reverse the tuck-in process, an idea given some credence by the pressure dependence of the reaction of (3) with H\(_2\) (vide infra).

In conclusion, we have described the facile preparation of the first example of a syn-bimetallic Ti complex (3) featuring a bridging hydride, originating from the C–H activation of an iPr substituent induced by addition of the strong Lewis base 1,3,4,5-tetramethylimidazol-2-ylidene. Preliminary studies show that this transformation is also effected by other, effectively “planar” Lewis bases, e.g. dimethylaminopyridine (DMAP). The resultant C–H activation product (3) readily reacts with an excess of H\(_2\) to produce very cleanly a unique syn-bimetallic di-hydride complex (4) featuring bridging and terminal hydride ligands. Labelling experiments and computational studies strongly suggest that the latter reaction does not proceed via a \(\sigma\)-bond metathesis mechanism.

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Conflicts of interest

There are no conflicts to declare.
Notes and references

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8 The hydride was found in the Fourier difference map and refined freely. We recognize the difficulties associated with the location of hydrogen atoms next to heavy atoms as Fourier ripples can be erroneously misinterpreted for hydrogen atoms due to the sharp cut-off at high angles. Nevertheless, based on the spectroscopic evidence the hydrogen atoms have been included in the supplied models.


15 Data were collected up to 0.82 Å resolution using Cu/Kα as well up to 0.78 Å using Mo/Kα with identical metric parameters.