



Cite this: *Chem. Commun.*, 2017, 53, 11877

Received 19th September 2017,  
Accepted 11th October 2017

DOI: 10.1039/c7cc07310k

rsc.li/chemcomm

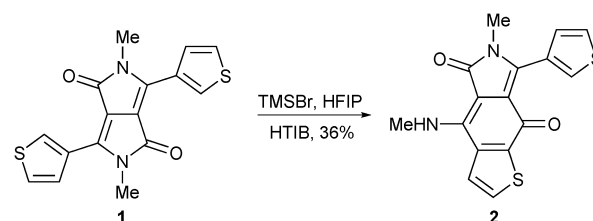
## Unprecedented rearrangement of diketopyrrolopyrroles leads to structurally unique chromophores†

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Diketopyrrolopyrroles possessing thienyl, furyl and benzofuryl substituents undergo unprecedented skeletal rearrangement in the presence of trimethylsilyl bromide resulting in the formation of thieno[2,3-*f*]isoindole-5,8-diones and furo[2,3-*f*]isoindole-5,8-diones. These relatively small dyes possess favorable photophysical properties with the emission maxima within the range of 573–624 nm, large fluorescence quantum yields, moderate sensitivity of emission to solvent polarity and a HOMO–LUMO gap of ca. 1.8 eV.

Organic electronics and fluorescence imaging require heterocycles with  $\pi$ -extended conjugation.<sup>1</sup> Rational design, aided by combinatorial and diversity-orientated approaches, provides invaluable routes toward novel functional materials.<sup>2</sup> Nonetheless, certain sets of chromophores evade even the best designed retrosynthetic analysis and the serendipity factor tends to play a role in the discovery of new dyes.<sup>3</sup> Depending on the targeted application, fluorophores possessing different combinations of properties are desirable. Small red-emitters are pivotal for applications in live fluorescence imaging.<sup>4</sup>

As a part of a larger project aimed at the synthesis of novel fluorophores, we investigated oxidative aromatic coupling of diketopyrrolopyrroles (DPPs, Fig. S1, ESI†).<sup>5</sup> We focused on iodine(III) reagents which efficiently promote such transformations, making them often the reagents of choice.<sup>6</sup> During this study, we observed that subjecting 2,5-dimethyl-1,4-diketo-3,6-di(thien-3-yl)pyrrolo[3,4-*c*]pyrrole (**1**) to hydroxy(tosyloxy)iodobenzene (Koser's reagent, HTIB) and TMSBr led to the formation of a purple compound as the sole isolable product (Scheme 1 and Table 1,



Scheme 1 The rearrangement of DPP **1** into thieno[2,3-*f*]isoindole-5,8-dione **2**.

entry 1).<sup>7</sup> Surprisingly, while the exact mol. mass of a new compound turned out to be equal to the mol. mass of the substrate, the <sup>1</sup>H NMR data revealed the loss of symmetry in the product (Fig. S2, ESI†). Such findings are consistent only with the rearrangement of the DPP **1**.

NMR analysis provides structural information about the synthesized compound. The <sup>13</sup>C NMR spectrum shows two signals in the carbonyl region (171.1 ppm and 165.6 ppm) along with two signals in the aliphatic region (32.8 ppm and 28.7 ppm). Moreover, the <sup>1</sup>H NMR spectrum shows a singlet at 3.38 ppm and a doublet at 3.42 ppm, which we ascribe as methyl groups. The difference in the chemical shifts as well as the multiplicity of the signals reconfirms the formation of the asymmetric product. The character of the proton signal at 8.82 ppm (<sup>1</sup>H NMR, broad quartet) reveals the opening of one  $\gamma$ -lactam ring and the presence of an NH–CH<sub>3</sub> moiety in the product. Moreover, COSY experiments support a 3 + 2 spin system with one uncorrelated downfield signal. The reaction on one of the carbons adjacent to the sulfur atom can rationalize this finding. Finally, thorough analysis of <sup>1</sup>H<sup>13</sup>C HSQC and <sup>1</sup>H<sup>13</sup>C HMBC spectra revealed that the structure of **2** possesses the thieno[2,3-*f*]isoindole-5,8-dione core, which has so far never been described in the literature (Scheme 1). The final proof was obtained from single crystal X-ray analysis (Fig. S2 and S3, ESI†).

Encouraged by the unique structure of the prepared dye, we started a systematic study of the discovered transformation. The mixture of TMSBr and HTIB provides only moderate yields

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† Electronic supplementary information (ESI) available. CCDC 1574216. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7cc07310k

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**Table 1** Influence of the reaction conditions on the rearrangement of DPP **1** into dye **2**<sup>a</sup>

Entry	Lewis acid	Additive, eq.	Solvent	Yield, <sup>b</sup> %
1	TMSBr	HTIB (2)	HFIP	36
2	TMSBr	HTIB (0.05)	HFIP	75
3	TMSBr	—	HFIP	73
4	—	HTIB (2)	HFIP	0
5	TMSBr	—	TFE <sup>c</sup>	75
6	TMSBr	—	MeOH	0
7	TMSBr	—	DCM	0
8	BF <sub>3</sub> ·Et <sub>2</sub> O	—	HFIP	53
9	BF <sub>3</sub> ·Et <sub>2</sub> O <sup>d</sup>	—	TFE <sup>c</sup>	72

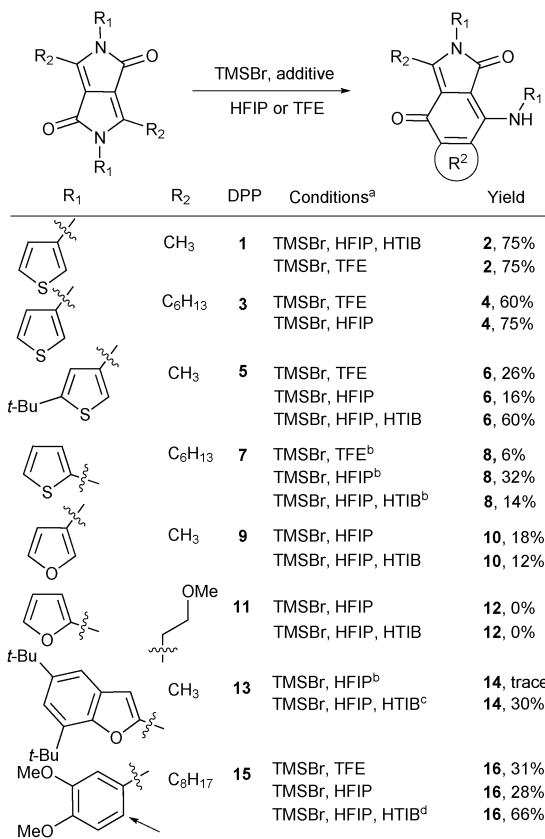
<sup>a</sup> Unless otherwise noted, reactions were performed with **1** (0.15 mmol), Lewis acid (0.15 mmol), solvent (1 ml), 50 °C, 17 h. <sup>b</sup> Isolated yields are reported. <sup>c</sup> Reaction was performed at 70 °C. <sup>d</sup> 2 eq. of the Lewis acid was utilized.

of the expected product, strongly emphasizing the need for optimization of the reaction conditions (Table 1, entry 1). In fact, the iodine(III) reagent does not support the formation of dye **2** product which suggests the non-oxidative character of this transformation. Rather, HTIB causes the decomposition of the desired product when added in excess (Table 1, entries 1–3). Conversely, the reaction does not proceed in the absence of TMSBr (Table 1, entry 4). We also elaborated the study of the discovered transformation to exhibit strong solvent dependence, which points to the importance of an acidic medium (Table 1, entries 3 and 5–7). In addition, improved solubility of the parent DPP aids a favorable outcome.

Aiming to increase the yield of the desired product, we tested a stronger Lewis acid, BF<sub>3</sub>·Et<sub>2</sub>O. Experimental investigations have proven the formation of dye **2** in up to 72% yield (Table 1, entries 8 and 9). It is noteworthy that the use of BF<sub>3</sub>·Et<sub>2</sub>O requires a higher loading and elevated temperature to achieve a reasonable yield of the rearranged product. Other Lewis acids such as POCl<sub>3</sub>, TFOH, TFA or BCl<sub>3</sub> do not mediate this rearrangement (Table S1, ESI<sup>†</sup>).

By adopting conditions 2 and 5 (Table 1), we undertook the synthesis of analogues of heterocycle **2**. Experimental findings showed a negligible effect of the alkyl substituents on the discovered transformation (Scheme 2). Notably, in this case the reaction medium plays an important role. We observed a significant improvement in the yield of dye **4** (60% *versus* 75%) when HFIP was used in place of TFE. Variations in the solubility of the substrate in the tested solvents can probably rationalize this finding. Surprisingly, the transformation of DPP **5**, bearing additional *tert*-butyl groups, does not proceed as expected in either TFE or HFIP. Unexpectedly, the presence of HTIB in a catalytic amount leads to significant improvement in the reaction outcome (Scheme 2).

2,5-Di(*n*-hexyl)-1,4-diketo-3,6-di(thien-2-yl)pyrrolo[3,4-*c*]pyrrole (**7**) affords the desired product **8** with a minuscule 6% yield even if the reaction is conducted for 3 days. The addition of Koser's reagent improves the reactivity of DPP **7**, however, probably due to the oxidative character, this reagent induces the decomposition of the product or intermediates. Luckily, the replacement of TFE with HFIP and lowering the reaction temperature to 50 °C led to the desired product in 32% yield.



**Scheme 2** Scope of the diketopyrrolopyrroles' rearrangement. <sup>a</sup> Unless otherwise noted, reactions were performed with TMSBr (0.15 mmol), solvent (1 ml), for 17 h at 50 °C (for HFIP) or 70 °C (for TFE). Isolated yields are reported; <sup>b</sup> reaction was conducted for 2 days; <sup>c</sup> reaction was conducted for 3 days; <sup>d</sup> reaction was performed with 2 eq. of Lewis acid.

Diketopyrrolopyrroles bearing 2-furyl and 3-furyl groups (**9** and **11**) display a lower reactivity compared to **1** and **7**, respectively. The desired product was obtained in low yield only for DPP **9**. The attempts to improve the reaction outcome (HTIB, prolongation of the reaction time) did not bring about positive results. DPP **13**, possessing a peripheral benzofuran moiety, displays poor reactivity as well. Fortunately, significant stability of the rearranged product allowed for the prolongation of the reaction time as well as utilization of the iodine(III) catalyst and, as a consequence, we could achieve a higher yield of **14**. Under similar conditions, the rearrangement occurs for DPP **15** (Scheme 2), bearing 3,4-dimethoxyphenyl groups. *N*-Alkylated DPPs bearing less electron-rich aryl substituents do not undergo this rearrangement.

With the library of new dyes in hand, we conducted a series of photophysical studies, which reveal that the absorption bands corresponding to the lowest energy absorption are located in the 530–606 nm region (Table 2 and Fig. 1), displaying bathochromic shifts in comparison to the parent DPPs. Emission maxima are red-shifted to the orange-red part of the visible spectra. Compounds **2**, **4** and **6** exhibit only small differences in absorption and emission maxima. Nevertheless, the type and position of the heteroatom influence  $\lambda_{\text{abs}}$  and  $\lambda_{\text{em}}$  significantly. Different charge distributions in the molecules



Table 2 Photophysical properties of novel dyes measured in dichloromethane

Dye	$\lambda_{\text{abs}}$ [nm]	$\lambda_{\text{em}}$ [nm]	$\epsilon_{\text{max}}$ [M <sup>-1</sup> cm <sup>-1</sup> ]	$\Phi$ [%]	$\Delta S$ [cm <sup>-1</sup> ]	$\tau$ [ns]	$k_{\text{r}}$ [10 <sup>7</sup> s <sup>-1</sup> ]
2	570	601	25 500	48 <sup>a</sup>	900	8.21	5.9
4	569	605	17 200	49 <sup>a</sup>	1100	7.78	6.3
6	571	602	17 300	47 <sup>a</sup>	900	6.45	7.3
8	582	619	16 000	15 <sup>a</sup>	1000	2.25	6.7
10	549	573	17 500	13 <sup>b</sup>	800	2.33	5.6
14	606	624	59 400	20 <sup>c</sup>	500	2.24	8.9
16	541	582	18 200	57 <sup>b</sup>	1300	7.25	7.9

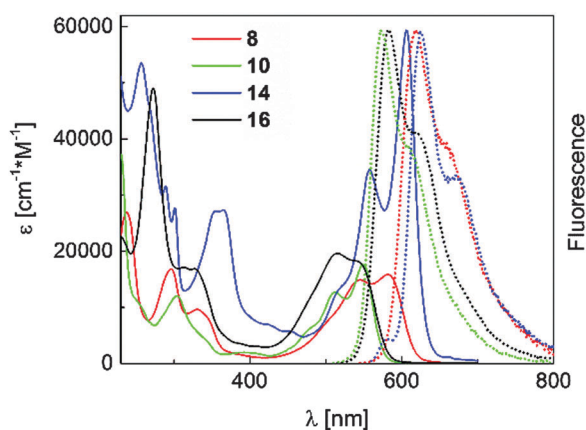
<sup>a</sup> Calculated with respect to Rh101 in MeOH. <sup>b</sup> Calculated with respect to Rh6G in EtOH. <sup>c</sup> Calculated with respect to SRh101 in EtOH.

can rationalize both the red-shift of the absorption and emission maxima when regioisomeric products **2** and **8** are superimposed and the blue-shift of  $\lambda_{\text{abs}}$  and  $\lambda_{\text{em}}$  for dyes possessing a furan unit (**2** versus **10**). Dye **14** possessing a  $\pi$ -expanded core is characterized by the most bathochromically shifted emission and absorption in the analyzed series.

Fluorescence responses for the majority of new dyes are rather high (Table 2). Fluorescence quantum yields ( $\Phi$ ) and  $\lambda_{\text{em}}$  are modestly influenced by the solvent polarity (Table S2, ESI<sup>†</sup>). In the exemplary case of **2**,  $\Phi$  changes from 56% in toluene to 28% in methanol (Table 3). This outcome is due to their moderately polarized electronic structures.

To rationalize these spectra, we performed additional quantum chemistry calculations for dye **2**, which concurred with experimental data (Fig. 2). According to our findings, the second excited  $S_2$  state locates around 7000 cm<sup>-1</sup> above the  $S_1$ , which means that the lowest excited  $S_1$  state is well separated from the higher located, electronically excited states (Fig. 2 and Fig. S5 and Tables S3, S4, ESI<sup>†</sup>). Moreover, the electronic transition  $S_0 \rightarrow S_1$  is of  $\pi \rightarrow \pi^*$  character and corresponds to excitation from the HOMO to the LUMO level (Fig. 3).

In order to get more information about the reasons for smoothing the fluorescence spectra (lack any well resolved vibrational structure even at 5 K) and the possible depopulation channels we simulated the vibrational structure of the fluorescence spectrum of **2** (Fig. S5, ESI<sup>†</sup>). Calculation of the Franck-Condon (FC) factors (Fig. S6, ESI<sup>†</sup>) shows that the vibrational

Fig. 1 Absorption and fluorescence spectra for **8**, **10**, **14** and **16** in DCM.Table 3 Optical properties of **2** in various solvents

Solvent	$\lambda_{\text{abs}}$ [nm]	$\lambda_{\text{em}}$ [nm]	$\Delta S$ [cm <sup>-1</sup> ]	$\Phi^a$ [%]
Hexane	566	581	500	—
Toluene	573	599	800	56
DCM	570	601	900	48
ACN	566	603	1100	42
MeOH	565	617	1500	28

<sup>a</sup> Calculated with respect to Rh101 in MeOH.

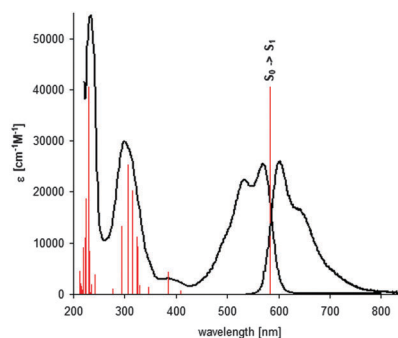


Fig. 2 Experimental absorption and fluorescence spectra of **2** in DCM. Calculated transition energies from the ground  $S_0$  to the electronically excited states are given by vertical solid lines with the heights proportional to the oscillator strength of the transition (see Table S3, ESI<sup>†</sup> for details). Vibronic structure is smoothed by low frequency vibrations (for interpretation see Fig. S6, ESI<sup>†</sup>).

structure of the absorption and fluorescence spectra are obscured by the three low frequency vibrations: 37, 59 and 100 cm<sup>-1</sup>. These vibrations are characterized by high shift parameters<sup>8</sup> and contribute to the spectrum in their common combinations as well as in combination with other vibrations.

To elucidate the influence of the solvent polarity on the absorption and fluorescence spectra, we calculated dipole moments for **2** in the isolated form as well as in solution. The comparison of the HOMO and LUMO orbitals clearly depicts the charge shift in the excited  $S_1$  state (Fig. 3). This shift causes an increase of the dipole moment (for ca. 1.5–2.2 D) and the bathochromic shift of both absorption and emission maxima in the polar solvents (Table S5, ESI<sup>†</sup>).

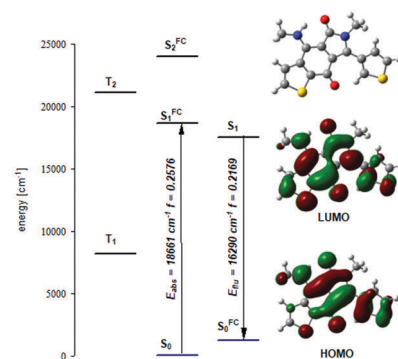


Fig. 3 Energy diagram with transition energies for absorption from the ground  $S_0$  state to the  $S_1^{\text{FC}}$  and fluorescence from the relaxed  $S_1$  state to  $S_0^{\text{FC}}$  for **2**. On the right – the orbitals HOMO and LUMO of **2**. Electronic configuration HOMO  $\rightarrow$  LUMO describes the transition  $S_0 \rightarrow S_1$ .



**Table 4** Redox potentials of the dyes **2**, **4**, **6**, **8**, **10**, **14** and **16** measured in dichloromethane<sup>a</sup>

Dye	$E_{\text{ox}}^{\text{pa}}$ , V	IP, eV	$E_{\text{red}}^{1/2}$ , V	$E_{\text{red}}^{\text{onset}}$ , V	EA, eV
<b>2</b>	0.90	-5.1	-1.10	-1.02	-3.3
<b>4</b>	0.96	-5.1	-1.15	-1.06	-3.3
<b>6</b>	0.86	-5.1	-1.17	-1.07	-3.3
<b>8</b>	0.92	-5.1	-1.15	-1.03	-3.3
<b>10</b>	0.94	-5.2	-1.20 <sup>b</sup>	-1.05	-3.3
<b>14</b>	0.89	-5.1	-0.93	-0.81	-3.5
<b>16</b>	0.80	-5.0	-1.28	-1.19	-3.2

<sup>a</sup> Measurement conditions: electrolyte (NBu<sub>4</sub>ClO<sub>4</sub>,  $c = 0.1$  M); DCM<sub>dry</sub>, potential sweep rate: 100 mV s<sup>-1</sup>, working electrode: glassy carbon (GC); auxiliary electrode: Pt wire; reference electrode: Ag/AgCl; all measurements were conducted at room temperature. <sup>b</sup>  $E_{\text{red}}^{\text{pc}}$  in V for **10**.

The performed calculations also provide information on the origin of the remaining non-radiative depopulation channel. We found that the excited S<sub>1</sub> state is located about 9500 cm<sup>-1</sup> above the triplet T<sub>1</sub> and 2500 cm<sup>-1</sup> below the T<sub>2</sub> states. Such an energetic situation precludes an efficient intersystem crossing pathway for the S<sub>1</sub> → T<sub>1</sub> depopulation. Thus, the main non-radiative depopulation channel of the S<sub>1</sub> state should proceed *via* internal conversion. This result concurs with our unsuccessful effort to detect phosphorescence emission at 5 K.

Our calculations also indicate that **2** possesses a planar core with the peripheral thiophene ring rotated around a single bond by around 25° with respect to the molecular plane (Table S4, ESI<sup>†</sup>). Thus, new heterocycles are planar, structurally rigid fluorophores, the structure of which does not change substantially upon electronic excitation. As a consequence of the small conformational changes following the transition between S<sub>0</sub> and S<sub>1</sub> states in these compounds (additional computational data in Table S6, ESI<sup>†</sup>) the Stokes' shifts are small.

To verify the potential use of the discovered dyes for optoelectronic applications we performed a series of cyclic voltammetry (CV) measurements, which depict the influence of structural changes on their susceptibility towards oxidation and reduction. In the majority of cases, the investigated molecules show reversible reduction and quasi-reversible oxidation waves (Table 4 and Fig. S5, ESI<sup>†</sup>). The exception to the rule is observed only for **10** (irreversible oxidation). It is noteworthy that no changes were perceived with increasing number of redox cycles and the direction of the CV-measurement.

We calculated both ionization potential (IP) and electron affinity (EA) values from the corresponding onset potentials. The results clearly indicate that peripheral substituents, as well as the electronic nature of the aromatic rings participating in the rearrangement, have no significant influence on the ionic potential (-5.1 ± 0.1 eV) or on the electron affinity (-3.3 ± 0.2 eV).

Herein, we have introduced structurally unique dyes possessing an isoindole-dione core, linearly fused with a thiophene, furan, or benzene ring. We have demonstrated that diketopyrrolopyrroles possessing moderately electron-rich aryl substituents undergo heretofore unknown rearrangements initiated by mild Lewis acids.

Strong emission of red light combined with straightforward synthesis, relatively small molecular mass, and potential chemical reactivity may open doors for these dyes for applications in fluorescence microscopy (bioimaging).

The financial support of the Polish National Science Centre (MAESTRO-2012/06/A/ST5/00216), Foundation for Polish Science (TEAM/2016-3/22) and the Global Research Laboratory Program (2014K1A1A2064569) through the National Research Foundation (NRF) funded by the Ministry of Science, ICT & Future Planning (Korea) is cordially acknowledged. Theoretical calculations were performed at the Interdisciplinary Center of Mathematical and Computer Modeling (ICM) of the Warsaw University (Poland) under the computational grant No. G-32-10. S. O. acknowledges JSPS fellowships.

## Conflicts of interest

There are no conflicts to declare.

## Notes and references

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