



# Cooperative, ion-sensitive co-assembly of tripeptide hydrogels†

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**Peptide co-assembly is of interest for the development of functional supramolecular biomaterials. Herein, computational simulations were combined with experimental validation to aid the design and understanding of cooperative co-assembly of a structure-forming tripeptide (FFD) and a functional copper-binding tripeptide (GHK) leading to hydrogel formation in response to complexation with copper ions.**

Supramolecular self-assembly<sup>1,2</sup> provides simple functional mimics of biological materials.<sup>3–5</sup> Co-assembly of two or more components provides options for control of morphological, physical and mechanical properties where the co-assembled material may have properties that are not found in the individual components. This cooperativity may be achieved by combining low molecular weight gelators with polymers<sup>6</sup> or surfactants;<sup>7</sup> or (aromatic) peptide amphiphiles with surfactants,<sup>8–10</sup> polysaccharides,<sup>11,12</sup> proteins<sup>13,14</sup> or light responsive gelators.<sup>15</sup> Co-assembly in these systems is typically governed by non-covalent interactions, in particular electrostatics.

While most existing supramolecular peptide biomaterials are based on long peptides,<sup>16</sup> typically containing eight or more amino acids, it is increasingly recognized that di-<sup>17,18</sup> or tripeptides contain sufficient molecular information to enable formation of self-assembling materials, including hydrogels,<sup>19–22</sup> tunable emulsions,<sup>23</sup> stable surface coatings<sup>24</sup> and most recently, supramolecular substrates for formation of polymeric pigments with tunable properties.<sup>25</sup> Thus, supramolecular structures based on short peptides hold promise as functional biomaterials.<sup>26–29</sup>

A number of short peptides are known to possess specific biological functions. These short peptides have been effectively introduced into supramolecular biomaterials through covalent linking to the structure-forming component of the system. Examples include materials bearing the fibronectin-derived RGDS,<sup>30,31</sup> or RGD<sup>32,33</sup> the laminin-derived IKVAV<sup>34</sup> the tyrosine kinase nerve growth factor receptor-derived KFG,<sup>27</sup> and the catalytically relevant<sup>35</sup> histidine containing tripeptides (such as HFF).<sup>36</sup> Another example is GHK, which is a copper-binding extracellular matrix derived peptide sequence that plays a role in wound healing.<sup>37,38</sup>

In this communication, we demonstrate how non-covalent incorporation of short functional peptides (exemplified by glycine-histidine-lysine, GHK) into supramolecular structures can be used to functionalise short peptide hydrogels. In particular, this work focuses on the design and co-assembly of a structure-forming tripeptide (diphenylalanine-aspartic acid, FFD)<sup>21,23</sup> and GHK leading to hydrogel formation in response to complexation with copper ions.

GHK is itself not a gelator, and FFD and DFF were selected as candidate structural tripeptides due to their previously described self-assembly and charges that complement that of GHK. FFD and DFF were previously shown to form bilayer-like aggregates, but do not form gels in isolation at neutral pH (FFD has been shown to form a hydrogel at pH 5).<sup>23</sup>

30 mM of GHK and 30 mM of FFD were dissolved separately in PBS buffer (pH 7.4) at room temperature. Samples were sonicated and vortexed in order to improve solubility and homogeneity. Both samples formed clear solutions. To investigate whether electrostatic interactions of the positively charged lysine (K) and the negatively charged aspartic acid (D) could induce co-assembly, a mixture of GHK/FFD was prepared at the same conditions (*i.e.* 30 mM of each, 60 mM total) which formed a clear solution (no hydrogel was formed). Interestingly, spontaneous hydrogel formation was observed in the presence of 30 mM CuCl<sub>2</sub>, suggesting that GHK complexation with copper ions leads to the co-assembly and gelation observed.

Transmission electron microscopy (TEM) was used to observe the structures formed. As shown in Fig. 1, GHK on its own

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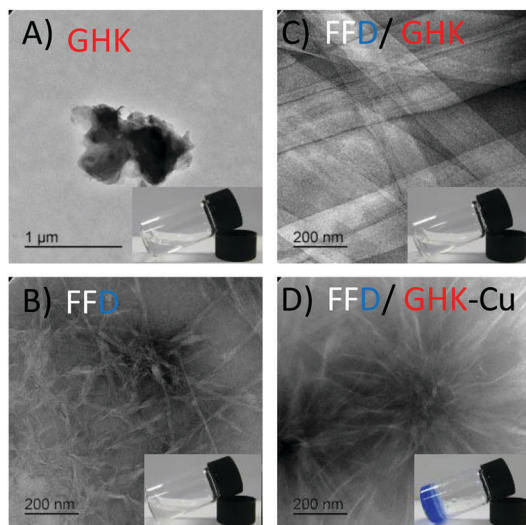
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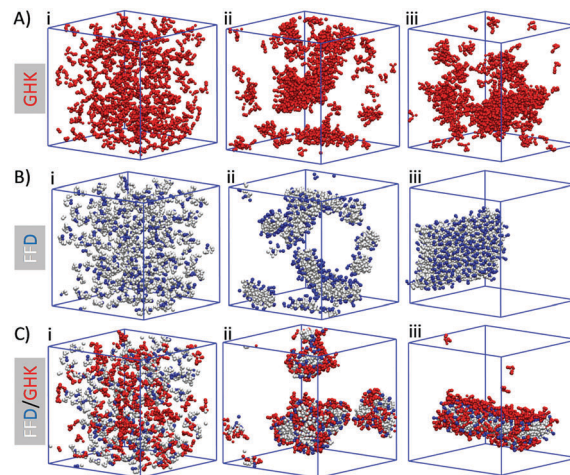




**Fig. 1** TEM images of all tripeptides showing random aggregates of GHK (A), self-assembly of FFD nanofibers (B), nanotape formation due to the co-assembly of FFD/GHK (C) and nanofiber formation due to the co-assembly of FFD/GHK after complexation with copper ions (D). Insets show the optical images of the corresponding peptides in inverted vials.

formed random aggregates, Fig. 1A. FFD (which was previously observed to form bi-layered nanotubes in acidic conditions)<sup>23</sup> formed nanofibers at pH 7.4, Fig. 1B. Upon mixing of both FFD and GHK, the peptides co-assemble to form tape-like structures, Fig. 1C. However, this co-assembly did not lead to gelation as is typically observed for 2D self-assembled structures, such as tapes and sheets<sup>39</sup> that are usually not able to entrap the solvent (in this case water) within their networks. When  $\text{CuCl}_2$  was mixed with the FFD/GHK mixture, networks of nanofibers were formed leading to hydrogel formation, Fig. 1D. The TEM images suggest a rapid nucleation and growth mechanism giving rise to spherulite-like structures.<sup>40</sup> Additional TEM images with different magnifications are presented in Fig. S1 (ESI<sup>†</sup>).

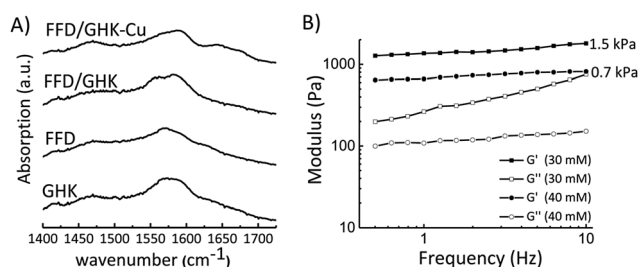
Having established the co-assembly conditions for the successful formation of hydrogels, we used computational simulations to provide further understanding of the molecular co-assembly. Computational results were obtained using GROMACS 4.5.3 utilizing the MARTINI coarse grain (CG) force field. In each system, 300 peptide molecules are added (150 of each for mixed systems) to a  $12.5 \text{ nm}^3$  box solvated with MARTINI coarse grained water. This gives an approximate peptide concentration of 200 mM, which is a factor of 10 greater than normal experimental conditions, consistent with previous work in this field where the increased concentration is required to accelerate the simulated aggregation process.<sup>21,41,42</sup> Each simulation is equilibrated for a total of 9.6  $\mu\text{s}$ , with snapshots taken at initial, mid-point and final stages, Fig. 2. As can be seen, for GHK (Fig. 2A) the peptides show some clustering but overall maintain a random arrangement. For FFD, (Fig. 2B), the peptides arrange themselves with the phenylalanine rings interacting with aromatic side chains of other peptides. This induces the formation of a bilayer structure, with a hydrophobic core (FF) and negatively charged hydrophilic (D) surface. This formation is already



**Fig. 2** Computational time course for the self-assembly of the tripeptides in isolation (A and B) and co-assembled nanostructures (C). (i) Initial frame (0  $\mu\text{s}$ ), (ii) mid-point ( $\sim 3 \mu\text{s}$ ), (iii) final frame ( $\sim 9.6 \mu\text{s}$ ).

observed at the start of the simulation, indicating a rapid hydrophobic collapse. When introducing GHK into a FFD system (Fig. 2C) a co-assembly is observed where GHK peptides are organized on the surface of the FFD structure. The computational results suggest that this occurs due to the electrostatic interaction between the lysine and the anionic surface of the FFD.

To assess the effect of hydrogen bonding interactions experimentally, and to assess the formation of salt bridges between D and K residues, FTIR measurements were performed in  $\text{D}_2\text{O}$  for individual components (GHK and FFD) and for mixtures of both before and after complexation with copper ions, Fig. 3A. For both GHK and FFD broad FTIR peaks at around  $1560 \text{ cm}^{-1}$  is assigned for the deprotonated  $\text{COO}^-$ . The shift of this peak towards a lower wavenumber for FFD/GHK is indicative of the formation of more ordered structures with possible salt-bridge formation in co-assembled mixtures.<sup>21,23</sup> It is worth noting that, although complexation with  $\text{Cu}^{2+}$  leads to fiber/hydrogel formation, the FTIR spectrum does not show a clear shift of peak at  $1560 \text{ cm}^{-1}$ . This observation might be explained as that salt-bridge formation becomes less significant after complexation with copper ions, thus reducing the effective salt-bridge. The appearance of the peaks in the amide I region indicated the presence of hydrogen bonding between the amide



**Fig. 3** (A) FTIR spectra of tripeptides in isolation and after co-assembly, samples (30 mM of each component) were prepared in deuterated phosphate buffer pH 7.4 at room temperature. (B) Rheology data for the FFD/GHK-Cu hydrogel with associated storage and loss moduli.



carbonyls. However, the lower intensity of these peaks suggest that these are weak hydrogen bonds. As these peaks are present in only samples containing the copper, we can conclude that the presence of the copper helps the assembling propensity of the tripeptides resulting in hydrogelation, whilst forming highly ordered nanostructures. There is also a shift in the signal from FFD in the presence of both GHK and copper which indicates the formation of more stable structures.

In order to assess the role of copper ions in specific binding and hydrogelation, we studied the gelation ability of GHK/FFD at different conditions. For instance,  $\text{Cu}^{2+}$  concentrations were varied (*i.e.* 20, 30, 40, 50 mM) and different metal ions ( $\text{Zn}^{2+}$ ,  $\text{Co}^{2+}$ ) were used as summarized in Table S1 (ESI<sup>†</sup>) (all added as chloride salts). We observed that gelation requires stoichiometric  $\text{Cu}^{2+}$  ions concentration (30 mM) and the gel was not formed when a lower  $\text{Cu}^{2+}$  ion concentration (20 mM) or much higher concentration (50 mM), most probably due to super saturation of the peptide-copper complexes in solution, thus preventing assembly. The lack of gelation for  $\text{Zn}^{2+}$  and  $\text{Co}^{2+}$  demonstrates that GHK binds specifically to  $\text{Cu}^{2+}$ <sup>43,44</sup> and this is required to enhance the peptide co-assembly leading to hydrogel formation. As previously reported, single crystal X-ray diffraction studies revealed that copper ions binding to GHK involves the N-terminal group of glycine, the nitrogen atom of the first amide bond, the unprotonated imidazole nitrogen from histidine, and two oxygen atoms in water (see also Scheme S1, ESI<sup>†</sup>).<sup>45</sup>

Having established the co-assembly conditions and ion-selectivity of gelation, we studied the mechanical properties of the formed hydrogels by rheology as shown in Fig. 3B. The rheology data shows that the hydrogel formed at 30 mM  $\text{Cu}^{2+}$  has an average loss modulus value of 1.5 kPa, while the hydrogel formed at 40 mM  $\text{Cu}^{2+}$  has a weaker gel with a storage modulus of about 0.7 kPa. The slope of the modulus might indicate thickening instability of the sample when 30 mM of  $\text{CuCl}_2$  was used as previously reported for similar systems.<sup>46</sup>

To assess the importance of peptide sequence, we also studied the co-assembly process in the presence of DFF instead of FFD, Fig. 4. Consistent with previously reported structure forming tripeptides (*i.e.*, anionic amino acids are preferred at the C-terminus),<sup>21</sup> no gel formation was observed for individual components as well as for mixtures before and after the addition of copper ions, Fig. 4A–C. TEM images (Fig. 4D–F and Fig. S2, ESI<sup>†</sup>) also suggest that both individual peptide solutions and mixtures are unable to self-assemble to form highly organized structures both in the presence and absence of copper ions. These results are in good agreement with our previous<sup>21</sup> and current computational predictions (Fig. S3, ESI<sup>†</sup>).

Molecular dynamics simulations of DFF and GHK mixtures show clear differences when compared with FFD/GHK, Fig. S3 (ESI<sup>†</sup>). The DFF structure compared with the FFD shows similar bilayer formation. The introduction of GHK appears to cause the collapse of the DFF structure resulting in random aggregates of DFF/GHK, as a result, no hydrogelation is observed in the presence of  $\text{Cu}^{2+}$ . Similarly, MD simulations of FFD self-assembly have resulted in a structure formation similar to that of DFF and FFD (Fig. S4, ESI<sup>†</sup>). However, the presence of aspartic acid in the middle position in the co-assembled system disrupts the packing

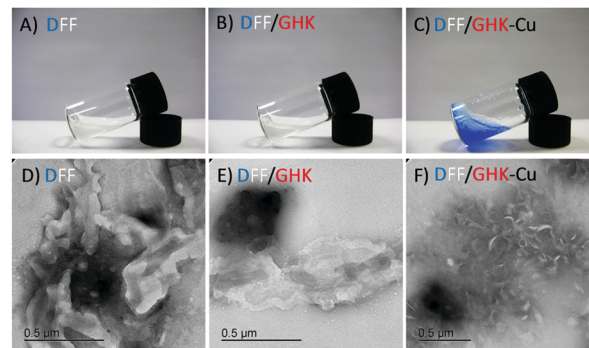


Fig. 4 Optical images (upper panel) and TEM images (lower panel) of the control (DFF) aggregation (A and D), mixed with GHK (B and E) and co-assembled sample after complexation with copper ions (C and F).

of the FF dyad within the tripeptide. The loss of the intramolecular  $\pi$ -interactions between the phenyl groups indicates that the FFD structure is not stable enough to form an elongated fibrillar structure when GHK is present.

It is worth noting that GHK has generated significant interest from the cosmetics industry, in particular as a complex with  $\text{Cu}^{2+}$  which facilitates copper intake by cells to enhance enzymatic processes, wound healing and promotes angiogenesis.<sup>47,48</sup> It also improves skin firmness and texture, fine lines and hyperpigmentation.<sup>37,38</sup> Thus, there is clear evidence of the biological activity of GHK and GHK–Cu complex and it is therefore considered as an additive for cosmetics and skin treatments. Formulating GHK into a biocompatible tripeptide hydrogel may be of direct interest for cosmetics industry.

In summary, the cooperative co-assembly of functional and structural tripeptides has been studied by combining both experimental validation and computational simulations. We have found that the co-assembled peptides can selectively complex with copper ions, leading to structural reconfiguration from clear solution of nanotapes to hydrogel formation of nanofibers. GHK is a known bioactive peptide and the hydrogels developed here may have applications in wound care and cell culture, which will require further investigation. The combined computation/experimental approach followed here provides a generally useful approach to identification of co-assembled structures of structural and functional components based on short peptides.

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## Conflicts of interest

There are no conflicts to declare.

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