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Copper(I)-catalysed transfer hydrogenations with ammonia borane†

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Highly Z-selective alkyne transfer semihydrogenations and conjugate transfer hydrogenations of enoates can be effected by employing a readily available and air-stable copper(I)/N-heterocyclic carbene (NHC) complex, [IPrCuOH]. As an easy to handle and potentially recyclable H₂ source, ammonia borane (H₃NBH₃) is used.

Ammonia borane (H₃NBH₃), an easy to prepare and manipulate solid, has attracted interest as a potential dihydrogen (H₂) storage material.^{1–6} Based on the fact that ammonia borane dehydrogenation products (e.g., borazines and polyiminoboranes) can potentially be recycled,⁷ and its exceptionally high H₂ content, many studies have been directed toward controlled catalytic release of H₂ (dehydrocoupling) from ammonia borane.^{1–6,8} Notably, to a much lesser extent H₃NBH₃ has been employed as a reagent for synthesis, for example in the reduction of functional groups or transfer hydrogenations.^{9–11}

Copper(I) hydride complexes have emerged as powerful catalysts for a wide variety of reductive transformations, generally relying on silanes^{12–15} or, to a much lesser extent, on tin hydrides¹⁶ or dihydrogen^{17,18} (Scheme 1). These terminal reducing agents differ conceptually: Si- and Sn-based compounds will deliver a hydride equivalent and the reaction products remain silylated or stannylated until hydrolysis or protonation. On the other hand, reactions with H₂ can directly deliver the hydride and proton equivalent, circumventing the need for an additional proton source and reducing the waste generated in these processes. As an example, copper(I)-catalysed alkyne semihydrogenations have recently demonstrated the potential replacement of silanes with H₂.^{18–20} However, most of these procedures require elevated H₂ pressure and/or have a limited substrate scope.

On the other hand, homogeneous transfer hydrogenations with copper catalysts are scarce,^{21,22} but circumvent the need

previous work:

- hydride transfer
- requires proton quench
- generates waste



- hydride and proton transfer
- high pressure equipment
- limited substrate scope



this work:

- transfer hydrogenation
- practical procedure



application in
catalysis

Scheme 1 Hydride and dihydrogen sources employed for copper(I) hydride catalysis.

for high pressure equipment. Given the synthetic utility of copper hydride complexes, we decided to investigate ammonia borane as a reagent for copper-catalysed transfer hydrogenations.

We decided to optimise the copper(I)-catalysed transfer hydrogenation with H₃NBH₃ employing alkyne transfer semihydrogenation of pentynol-derived alkyne **1** as a test reaction (Table 1).²³ Due to the reported reactivity with H₂,^{18–20} we chose well-defined copper(I)/NHC complexes bearing a copper–oxygen bond as the starting point for our studies. With a copper(I) *tert*-butanolate complex, generated *in situ* from [IMesCuCl] and NaOtBu, we observed no turnover to the corresponding alkene **Z-2** at room temperature, but full conversion at 50 °C (Table 1, entries 1 and 2). A similar trend was observed with a copper(I) hydroxide/NHC complex, [IPrCuOH],^{18b,24,25} which also underwent full conversion to **Z-2** at 50 °C (Table 1, entries 3–5). Increasing the temperature further to 60 °C led to diminished conversion of 60% (Table 1, entry 6). The use of [IPrCuOH] is particularly attractive, as it does not require the generally common *in situ* preactivation (to give activated copper(I) alkoxide complexes) and is air-stable for months. It should be noted that in all cases excellent *Z*-selectivity and negligible overreduction to the corresponding alkane **3** were maintained. This is synthetically

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Table 1 Optimisation of Cu-catalysed transfer semihydrogenation of alkynes

Entry	Conditions	Conv. ^a [%]	Z-2/E-2/3 ^a
1	[IMesCuCl], 7.5 mol% NaOtBu, 3 equiv. of H ₃ NBH ₃ , THF, rt, 16 h	—	—
2	[IMesCuCl], 7.5 mol% NaOtBu, 3 equiv. of H ₃ NBH ₃ , THF, 50 °C, 16 h	> 99	> 99:0:0
3	[IPrCuOH], 3 equiv. of H ₃ NBH ₃ , THF, rt, 16 h	9	> 99:0:0
4	As entry 3, 40 °C	68	> 99:0:0
5	As entry 3, 50 °C	> 99	> 99:0:0
6	As entry 3, 60 °C	60	> 99:0:0
7	As entry 5, in C ₆ H ₆	18	> 99:0:0
8	As entry 5, in CH ₂ Cl ₂	6	n.d.
9 ^b	[IPrCuOH], 3 equiv. of H ₃ NBH ₃ , THF, 50 °C, 4 h	> 99 ^c	> 99:0:0
10 ^b	[IPrCuOH], 2 equiv. of H ₃ NBH ₃ , THF, 50 °C, 4 h	88	> 99:0:0
11	[IPrCuOH], 20 mol% H ₃ NBH ₃ , 1 bar H ₂ , THF, 50 °C, 16 h	20	> 99:0:0

^a Determined by ¹H NMR and GC analysis. ^b Slow addition of H₃NBH₃ in THF solution over 3 h. ^c Isolated yield: 96%.

appealing, as alkyne semihydrogenations are commonly carried out with heterogeneous palladium catalysts which generally suffer from these particular limitations.²⁶ Furthermore, another report of alkyne transfer semihydrogenation with copper catalysts does not give turnover with internal alkynes,^{21c} which underscores the orthogonal reactivity of the present catalyst.²⁷

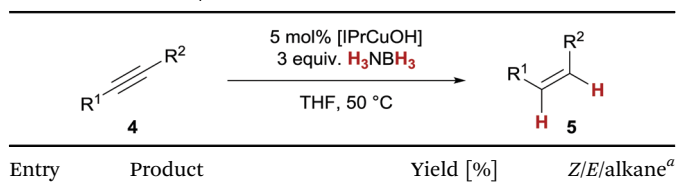
The use of THF as a solvent proved to be optimal, as the reactions in benzene and dichloromethane under otherwise similar reaction conditions gave lower conversion of **1** (Table 1, entries 7 and 8). At the onset of the described reactions in THF, gas evolution (most likely H₂) was observed, accompanied by a relatively slow reaction rate (>99% conversion was reached after ~12 h reaction time). This hinted at the loss of H₂ equivalents for the desired transfer semihydrogenation through concomitant dehydrocoupling. Consequently, when an H₃NBH₃ solution in THF was added slowly (3 h) to the reaction mixture, a generally faster conversion and higher yield of **Z-2** were observed (Table 1, entry 9, 96%). Nevertheless, a total of three equivalents of H₃NBH₃ were required for >99% conversion since slow addition of two equivalents of ammonia borane did not lead to complete consumption of **1** (Table 1, entry 10).²⁸ Based on these results, we decided to probe whether a direct transfer of a hydride equivalent to copper or H₂ activation after dehydrocoupling of ammonia borane was present. Therefore, we carried out the alkyne transfer semihydrogenation with a catalytic amount of H₃NBH₃ under a H₂ atmosphere. These reactions gave only 20% conversion of **1** (Table 1, entry 11); hence, we conclude that a direct hydride transfer mechanism is present.

With the optimised reaction conditions in hand, we explored the substrate scope of the copper-catalysed alkyne transfer

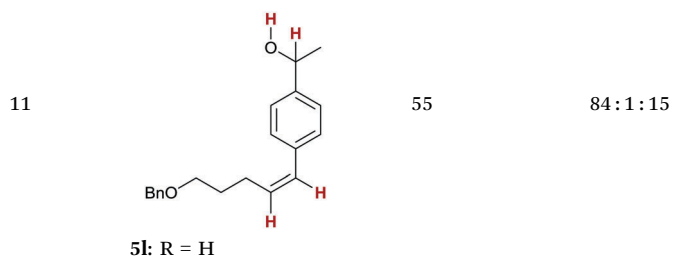
semihydrogenation. A variety of phenylacetylene derivatives with both electron-donating and electron-withdrawing functional groups **4a–4g** could successfully be transformed into alkenes **5a–5g** with good yields and consistently high *Z*-selectivity (Table 2, entries 1–7). Of note is the low conversion achieved with 3-anisole derivative **4d** (30%, Table 2, entry 4); we suspect substrate coordination detrimental to the catalyst, which is absent with the other two regioisomers **4c** and **4e** (Table 2, entries 3 and 5). With strongly electron-withdrawing substituents (as in **5h–5j**, Table 2, entries 8–10), we observed partial overreduction to the corresponding alkanes, although the *Z/E*-ratios remained high. The same is true for the acetophenone derivative **4k**, where overreduction to the benzyl alcohol **5k** was detected (Table 2, entry 11).²⁹ A variety of *Z*-stilbenes **5l–5o** can be synthesised from the corresponding tolane derivatives **4l–4o** with good yields (Table 2, entries 12–15). Here, a similarly detrimental influence of the electron-withdrawing trifluoromethyl substituent in **5o** was observed as 8% alkane was detected (Table 2, entry 15). The alkyne transfer semihydrogenation can also be applied to dialkylalkynes, as demonstrated by the formation of *Z*-alkenes **5p** and **5q** (Table 2, entries 16 and 17). For full consumption of the former, one more equivalent of H₃NBH₃ was added. Finally, protected allyl alcohols **5r** and **5s** can be obtained with high *Z*-selectivity from the corresponding propargylic silyl ethers **4r** and **4s** (Table 2, entries 18 and 19). Generally, the obtained yields are good; in some cases we observed undesired hydroboration/oxidation of the alkene products.²³ We attribute the somewhat diminished yields to this side reaction.

Finally, we investigated the transfer semihydrogenation of alkyl and aryl propiolates **6** (Scheme 2). In both cases the desired

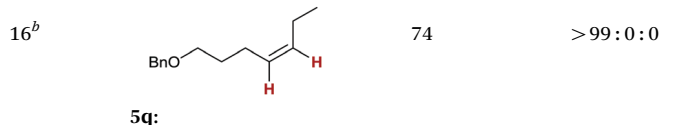


Table 2 Cu-catalysed transfer semihydrogenation of alkynes with ammonia borane: substrate scope

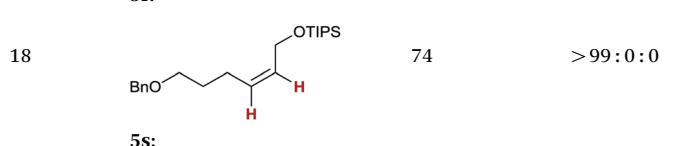
Entry	Product	Yield [%]	Z/E/alkane ^a
5a : R = 4-Me			
1		63	>99:0:0
2	5b : R = 4- <i>t</i> Bu	50	>99:0:0
3	5c : R = 4-OMe	69	95:5:0
4	5d : R = 3-OMe	30 conv.	>99:0:0
5	5e : R = 2-OMe	71	>99:0:0
6	5f : R = Br	80	>99:0:0
7	5g : R = Cl	63	>99:0:0
8	5h : R = CO ₂ Me	65	80:0:20
9	5i : R = CN	51	89:4:7
10	5j : R = CF ₃	71	81:0:19
5k :			



12		57	>99:0:0
13	5m : R = 4-OMe	66	>99:0:0
14	5n : R = 4-Cl	77	>99:0:0
15	5o : R = 4-CF ₃	50	92:0:8
5p :			

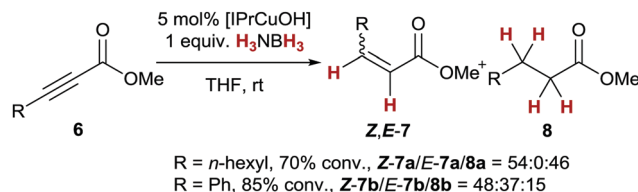


17		80 conv.	>99:0:0
5r :			



19		63	>99:0:0
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^a Determined by ¹H NMR and GC analysis. ^b 4 equiv. of H₃NBH₃ was used.

**Scheme 2** Transfer semihydrogenation of propiolates.

Z-acrylates **7a** and **7b** were found, however, in the presence of overreduced esters **8a** and **8b**. This indicated the viability of a conjugate transfer hydrogenation of the initially formed acrylates **7**. We therefore decided to investigate the conjugate transfer hydrogenation of α,β-unsaturated esters with ammonia borane next.

After initial optimisation we also found that conjugate transfer hydrogenation of enoates **9** with ammonia borane was feasible under similar reaction conditions (Table 3). Conjugate transfer hydrogenation of disubstituted ethyl esters **9a–9i** could be realised with generally good isolated yields when the reaction was carried out under previously optimised conditions (Table 3, entries 1–9). Isolated yields for esters **10** generally are higher than in the alkyne transfer semihydrogenation. This is possibly due to the absence of a reactive double bond in the products, which could lead to further reactions. Only thiophene derivative **10f** was obtained in a somewhat lower yield (Table 3, entry 6). Of note is the different reactivity of the catalyst toward double bond isomers: while the *E*-enoate **9d** underwent smooth conversion to the desired saturated product **10d** (Table 3, entry 4), the reaction with *Z*-enoate **9e** proceeded sluggishly (46% conversion to **10e** after 16 h) under otherwise similar conditions (Table 3, entry 5). Higher substituted acrylates **10j–10l** proved to be more difficult to obtain, most probably due to steric hindrance (Table 3, entries 10–12). However, by addition of 6 equivalents of H₃NBH₃, esters **10j–10l** could be isolated in good yields.³⁰

Table 3 Cu-catalysed conjugate transfer hydrogenation: substrate scope

Entry	Product	Yield [%]
1 ^a	10a : R ¹ = Ph, R ² = H	82
2 ^a	10b : R ¹ = 2-naphthyl, R ² = H	83
3 ^a	10c : R ¹ = 4-OMe-C ₆ H ₄ , R ² = H	83
4	10d : R ¹ = 4-OBn-C ₆ H ₄ , R ² = H	87
5 ^b	10e : R ¹ = H, R ² = 4-OBn-C ₆ H ₄	46 conv.
6	10f : R ¹ = 2-thiophenyl, R ² = H	66
7 ^a	10g : R ¹ = cyclohexyl, R ² = H	88
8	10h : R ¹ = 4-Br-C ₆ H ₄ , R ² = H	82
9	10i : R ¹ = 4-CF ₃ -C ₆ H ₄ , R ² = H	89
10 ^c	10j : R ¹ = 2-naphthyl, R ² = Me	93
11 ^c	10k : R ¹ = Ph, R ² = Me	76
12 ^c	10l : R ¹ = Ph, R ² = Ph	73 ^d

^a These reactions reached >99% conversion with 2 equiv. of H₃NBH₃.

^b The methyl ester of **9e** was employed. ^c 6 equiv. of H₃NBH₃ was used.

^d Contains 14% of the starting material.



In summary, we have developed copper(i) hydride-mediated transfer hydrogenation reactions employing ammonia borane (H_3NBH_3) as a H_2 equivalent. Alkynes can be converted into the corresponding alkenes with high Z-selectivities in transfer semihydrogenations. Also, the protocol enables a conjugate transfer hydrogenation of enoates. An air-stable and preactivated copper(i) hydroxide/NHC complex is used, circumventing the need for generally common *in situ* activation of the copper catalyst. We think that these results could be of importance with foresight to the use of readily available transition metal catalysts (such as copper) for transfer hydrogenations.

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- Reaction analysis by ^{11}B NMR shows the formation of polyamino-boranes (see the ESI† for details), which indicates the transfer of 1 equiv. of H_2 per molecule of ammonia borane. This could account for the high loadings of ammonia borane required.
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