Synthesis of a Fe3O4@P4VP@metal–organic framework core–shell structure and studies of its aerobic oxidation reactivity†

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A novel core–shell magnetic Fe3O4@P4VP(poly(4-vinylpyridine))@MIL-100(Fe) composite material has been successfully synthesized. The Fe3O4@P4VP core was initially prepared via polymerization of P4VP on the surface of pre-made Fe3O4 microspheres. The MIL-100(Fe) (MIL, Matériel Institut Lavoisier) metal–organic framework shell was then introduced through a layer by layer strategy. The obtained Fe3O4@P4VP@MIL-100(Fe) catalyst was applied in the selective oxidation of alcohols using molecular oxygen as the oxidant, and a variety of substrates were tolerated. The initial activity of the Fe3O4@P4VP@MIL-100(Fe) core–shell catalyst was maintained after ten reuses.

1. Introduction

Catalytic oxidations, such as conversions of alcohols to their corresponding carbonyl compounds, are of considerable importance in fine and industrial organic chemistry.1 Recently, aerobic transition metal-promoted selective oxidations have received increasing attention because of the inexpensiveness and environmental friendliness of molecular oxygen.2 Studies about homogenous catalysts are abundant in the literature.3 However, homogeneous catalysts are inherently difficult to recover from the reaction mixture, causing industrial wastes and additional costs.4 For this reason, heterogeneous catalysts are preferable for organic transformations, as they display several advantages during the purification and recovery steps of products and catalysts.5

As a result, numerous heterogeneous catalysts, such as molecular sieve,6 polymer supported7 and inorganic microspheres supported ones, have been developed in recent years.8 Among these solid catalysts, core–shell structured ones have been considered the most ideal ones, thanks to their unique physicochemical and multi-functional properties.9 Core–shell structures bearing magnetic functionality are of particular interest as an effective separation method because of their easy and quick response in an external magnetic field.10 A variety of magnetic core–shell composites have been developed utilizing different coatings, such as silica, carbon, polymer and porous materials on a Fe3O4 core.11 Several hybrid magnetic core–shell nanocatalysts have been utilized in a variety of catalytic reactions including reductions, oxidations, epoxidations, coupling reactions and photocatalysis.12 Many studies have reported immobilization of transition metal salts into core–shell structures.13 On the other hand, significant metal leaching and poor catalyst recyclability have been recurring issues in these metal immobilized systems. It is difficult to address such limitations because of the low molecular interactions present in the coordination bonds between the metal and the core–shell structure (Scheme 1).14

Metal–organic frameworks (MOFs)15 have gained increasing focus in catalysis, as a kind of porous materials with extremely large surface area, a high number of active sites and good stability.16 Therefore, fabrication of MOFs coated magnetic core–shell catalysts could address the problem of catalyst leaching and the need for recyclability.17 In this work, we wish to report a core–shell–shell structured Fe3O4@P4VP@MIL-100(Fe) composite core–shell–shell composite (Scheme 1).


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Scheme 1 Schematic illustration of the synthesis of Fe3O4@P4VP@MIL-100(Fe).
catalyst, which was assembled by coating MIL-100(Fe) on the surface of Fe3O4@P4VP microspheres. The inner shell of the P4VP polymer layer played an important role in the absorption of Fe2+ ions for the formation of the MIL-100(Fe) shell, as well as in protecting the magnetic core from destruction and aggregation during the oxidation process. The synthesized Fe3O4@P4VP@MIL-100(Fe) was characterized by TEM, field-emission SEM (FESEM), powder X-ray diffraction (PXRD), nitrogen adsorption/desorption analysis, FTIR spectroscopy, thermogravimetric analysis (TGA) and vibrating sample magnetometry (VSM). The aerobic oxidation of alcohols was investigated using our synthesized Fe3O4@P4VP@MIL-100(Fe) catalyst and great yields were achieved. A variety of alcohols was compatible with our oxidation system and the core–shell catalyst could be rapidly separated by an externally applied magnetic field.

2. Experimental section

2.1 Materials

Ferric chloride hexahydrate (FeCl3·6H2O), polyvinyl pyrrolidone (PVP; Mw = 58,000), 4-vinyl pyridine (4-VP, 96%), divinylbenzene (DVB, 80%), 1,3,5-benzenetricarboxylic acid (H3BTC), potassium persulfate (KPS, 99%), 2,2,6,6-tetramethyl-1-piperidinylxoy (TEMPO), trimethylacetaldehyde, benzyl alcohol, substituted benzyl alcohols, 2-pyrldinemethanol, cinnamic alcohol, 3-methyl-2-buten-1-ol, 3-methyl-2-buten-1-ol, 1-ocetyl alcohol, 1-phenyl ethanol, 2-cyclohexen-1-ol, cyclopentanol, cyclooctene, cyclohexene, cyclododecene, norborne and 2-pinene were purchased from Alfa Aesar. Poly(acrylic acid) (PAA; Mw = 1800) was obtained from Sigma-Aldrich. Anhydrous ethanol, ethylene glycol (EG), sodium acetate and potassium nitrite were purchased from Beijing Chemical Reagents Company (China). 4-Vinyl pyridine was purified by distillation under reduced pressure to remove inhibitor hydroquinone before use.

2.2 Synthesis of Fe3O4 nanoparticles with PAA

The Fe3O4 nanoparticles modified on the surface with poly (acrylic acid) (PAA) were synthesized by solvothermal method according to reported literature. First, FeCl3·6H2O (1.08 g) and PAA (0.108 g) were dissolved in ethylene glycol (40 mL) under magnetic stirring at 60 °C, followed by adding sodium acetate (9 g) under vigorous stirring to form a uniform solution. Then, the solution was decanted to a stainless-steel autoclave with Teflon-lining (50 mL) and maintained at 200 °C for 12 h. After the autoclave was cooled to room temperature, the black product was collected using a magnet. The Fe3O4 nanoparticles modified with PAA were washed with deionized water and ethanol several times, collected and dried under vacuum.

2.3 Preparation of core–shell Fe3O4@P4VP microspheres

The core–shell Fe3O4@P4VP microspheres were prepared by coating P4VP shell on the surface of Fe3O4 nanoparticles through radical initiated polymerization. Fe3O4 nanoparticles modified with PAA (0.1 g) were dispersed in 100 mL of a 0.15 wt% PVP aqueous solution under ultrasonication. The uniform solution was mixed with emulsion of 4-VP (0.125 g) and DVB (0.125 g) in aqueous solution (20 mL) that contains PVP (0.05 g). Then it was emulsionized by ultrasonic after placed in a 250 mL four-necked flask. The liquid was mechanical stirred for 4 hours under a flow of nitrogen and then polymerization of 4-VP was initiated by KPS (10 mg); at this point, the solution was heated to 70 °C. Core–shell Fe3O4@P4VP microspheres were obtained after 4 h, washed with deionized water and ethanol three times and dried under vacuum.

2.4 Synthesis of core–shell–shell Fe3O4@P4VP@MIL-100(Fe) microspheres

The core–shell–shell Fe3O4@P4VP@MIL-100(Fe) microspheres were synthesized by a step-by-step strategy. Typically, 0.05 g of Fe3O4@P4VP microspheres were dispersed in 10 mL of a 4.0 mM FeCl3·6H2O ethanol solution for 15 minutes and collected with a magnet. Then the microspheres were dispersed in a H2BTC ethanol solution (4 mL, 10 mM) and stirred at 70 °C for 30 minutes. The product was washed with ethanol and vacuum dried after twenty cycles.

2.5 Typical catalytic procedure for aerobic oxidation of alcohol

Alcohol (1.0 mmol), Fe3O4@P4VP@MIL-100(Fe) catalyst (5 mol%) were mixed in acetonitrile (5 mL); TEMPO (0.5 equiv.) and KNO3 (0.2 equiv.) were added in a 25 mL flask equipped with a condenser. The catalytic reaction was stirred at 75 °C under 1 atm of O2. After the reaction, the Fe3O4@P4VP@MIL-100(Fe) catalyst was recovered by applying an external magnetic field and the yield of product was determined by GC-MS.

2.6 Leaching test and catalyst recycling

The hot filtration test for the aerobic oxidation of benzyl alcohol was performed on a sample taken from the mixture after letting the reaction run for 4 hours. The catalyst was separated by a magnet and the supernatant solution was transferred into a new flask letting the reaction continue in absence of the catalyst. After certain reaction time, the reaction solution was analyzed by GC-MS. The recyclability of the heterogeneous catalyst was tested for the oxidation of alcohols and epoxidation of olefins. The reaction run with a batch of recycled catalyst was performed under the conditions described above.

3. Results and discussion

The Fe3O4 microspheres modified with PAA were synthesized by a solvothermal method reported in literature. These microspheres were composed of tiny Fe3O4 nanocrystals within 10 nm (Fig. 1a). The P4VP shell was successfully grafted on Fe3O4 to form core–shell composite microspheres held together by hydrogen bonds between carboxyl groups of PAA chains and pyridine. The diameters of the magnetic core and polymer shell were about 200 and 38 nm, respectively (Fig. 1b). Fe3O4@P4VP@MIL-100(Fe) core–shell–shell catalysts were prepared by a layer-by-layer method with 5 to 20 assembling cycles (Fig. 1c–f). The Fe3+ ions were adsorbed on the surface of P4VP through
pyridine coordination, followed by reacting with 1,3,5-benzene
tricarboxylic acid to complete each assembling cycle. The
thickness of the MIL-100(Fe) shell was controlled by adjusting
the assembling cycle. The MOF shells could be increased
significantly to achieve a thickness ranging between 15 and
90 nm.

To further confirm the crystalline structures of the Fe3O4,
Fe3O4@P4VP and Fe3O4@P4VP@MIL-100(Fe) microspheres,
powder XRD were performed and the patterns are shown in
Fig. 2. The diffraction peaks of Fe3O4 (Fig. 2a) agree with stan-
dard JCPDS 75-1609.39 indicating a face-centered cubic lattice.
After coating with P4VP, the peaks of the core–shell Fe3O4@P4VP
microspheres are almost the same as those for Fe3O4, be-
cause of the low-crystalline nature of the polymeric shell
(Fig. 2b). However, after the introduction of the MIL-100(Fe)
layer, several new peaks appeared on the XRD pattern
(Fig. 2c). These new peak match the peaks present in the
simulated MIL-100(Fe) pattern (Fig. S1†), which indicates that
the formation of MIL-100(Fe) was successful.

The EDX elemental maps further support the successf
of the synthesis of Fe3O4@P4VP@MIL-100(Fe), and that the iron
MOF was well grafted on the surface of Fe3O4@P4VP (Fig. 3).
The iron elemental maps of the Fe3O4@P4VP@MIL-100(Fe) also
showed an excellent distribution of iron content on the surface.
The existence of iron in the Fe3O4 core structure was also
observed, as shown in Fig. 3.

The FTIR spectrum of the product is shown in Fig. 4. The
relatively high intensity of the band at 592 cm⁻¹ is characteristic
of Fe–O vibrations. The characteristic absorption bands at 1707,
1250, and 1165 cm⁻¹ correspond to the C=O stretching of the
coxylic group, to the in-plane bending of C–O–H, and to the
–(C–O)H stretching, proving the existence of rich –COOH
functional groups on the surface of the Fe3O4 particles. As for
the Fe3O4@P(4-VP-DVB) microspheres, the characteristic
absorptions at 1603, 1562, and 1417 cm⁻¹ are attributed to the
vibrational modes of the pyridine ring. The band at 1603 cm⁻¹
 corresponds to the stretching vibration absorption of the C–N
bond, and the bands at around 1562 and 1417 cm⁻¹ are
attributed to the stretching vibration absorption of the C=C
bond; this further proves the encapsulation of P4VP on the
surface of the Fe3O4 particles. The peaks at 1435 and 1575 cm⁻¹
are due to the stretching vibrations of the C-C bonds of the
benzene ring, while the peaks at 1376 and 1618 cm⁻¹ are

![Graph 1](Fig. 1 HRTEM images of (a) Fe3O4, (b) Fe3O4@P4VP and Fe3O4@P4VP@MIL-100(Fe) core–shell after (c) 5, (d) 10, (e) 15, (f) 20 assem-
bling cycles.)

![Graph 2](Fig. 2 XRD patterns of (a) Fe3O4, (b) Fe3O4@P4VP, (c) Fe3O4@P4VP@MIL-100(Fe) and (d) recycled Fe3O4@P4VP@MIL-100(Fe).)

![Graph 3](Fig. 3 The TEM images of the Fe3O4@P4VP@MIL-100(Fe) and EDX elemental maps of Fe, O and C, respectively.)

![Graph 4](Fig. 4 FTIR spectra of (a) Fe3O4, (b) Fe3O4@P4VP and (c) Fe3O4@P4VP@MIL-100(Fe).)
attributed to the C–O stretching vibrations of the carboxylic moiety, suggesting that MIL-100 (Fe) shell has been successfully grafted on the surface of Fe$_3$O$_4$@P4VP.

Surface area and pore size distribution were characterized by nitrogen adsorption/desorption experiment. The N$_2$ adsorption/desorption isotherms are shown in Fig. 5 and the curves are formed by the type I isotherms (Fig. 5). In addition, the isotherms are of type II at the range of high $P/P_0$. The surface area and pore volume of Fe$_3$O$_4$@P4VP@MIL-100(Fe) were calculated by the BET model and the values were determined to be 346.2 and 2.31 m$^3$ g$^{-1}$, respectively. The surface area is significant improved, when compared to its Fe$_3$O$_4$@P4VP precursor, which has a surface area of 31.1 m$^2$ g$^{-1}$.

TGA curves of Fe$_3$O$_4$, Fe$_3$O$_4$@P4VP and Fe$_3$O$_4$@P4VP@MIL-100(Fe) are shown in Fig. 6. For PAA modified Fe$_3$O$_4$ microspheres, the weight loss was about 17 wt% over two steps, attributed to the desorption of adsorbed water and PAA chains in Fe$_3$O$_4$ microspheres (Fig. 6a). The TGA curves of the Fe$_3$O$_4$@P4VP microspheres exhibited three steps of weight loss with 52 wt% of the initial weight remaining. Elimination of water, decomposition of the oligomeric structure and the desorption of P4VP shell correspond to the three weight losses (Fig. 6b). As expected, the mass loss of Fe$_3$O$_4$@P4VP@MIL-100(Fe) was lower than that of Fe$_3$O$_4$@P4VP, likely due to the formation of iron oxide through calcination (Fig. 6c).

The magnetic properties of Fe$_3$O$_4$ (PAA), Fe$_3$O$_4$@P4VP and Fe$_3$O$_4$@P4VP@MIL-100(Fe) were measured by vibrating sample magnetometry (VSM). All three samples exhibited super-paramagnetism, presenting no remanence or coercitive forces (Fig. 7). The resultant magnetization saturation values of Fe$_3$O$_4$ (PAA), Fe$_3$O$_4$@P4VP and Fe$_3$O$_4$@P4VP@MIL-100(Fe) were 73.90, 47.35 and 28.21 emu g$^{-1}$, respectively, indicating strong magnetic response by the materials. The saturation magnetic moment of Fe$_3$O$_4$@P4VP and Fe$_3$O$_4$@P4VP@MIL-100(Fe) are lower than that of Fe$_3$O$_4$ because the P4VP and MOF shells are non-magnetic materials. Moreover, the saturation magnetic moment shows linear dependence on the content of magnetite. Therefore, the amount of P4VP and MIL-100(Fe) calculated according to the VSM results were 21.41 wt% and 40.42 wt%, respectively.

The catalytic activity of the synthesized Fe$_3$O$_4$@P4VP@MIL-100(Fe) toward the aerobic oxidation reaction was evaluated employing alcohol substrates, and using TEMPO as the radical initiator. No benzaldehyde was detected in the absence of the catalyst under our reaction conditions (Table 1, entry 1). Also, the employment of Fe$_3$O$_4$@P4VP as the catalyst failed to oxidize any benzyl alcohol to its corresponding aldehyde (Table 1, entry 2). Homogeneous iron catalysts were evaluated and almost quantitative conversion was observed. However, slightly lowered selectivity was observed due to the formation of benzoic acid (Table 1, entries 3 and 4). 82% conversion and 99% selectivity were achieved using MIL-100(Fe) MOF catalyst at 5 mol% catalyst loading (Table 1, entry 5). The amounts of MIL-100(Fe) and Fe$_3$O$_4$@P4VP@MIL-100(Fe) added in this reaction were 5 mol%, based on the same iron loading (Table 1, entries 5 and 6). The Fe$_3$O$_4$@P4VP@MIL-100(Fe) catalyst gave an excellent conversion and yield due to the efficient utilization of the
porous MOF structure, which was coated as a nanolayer on the magnetic core. Hydrogen peroxide has been utilized as the oxidant, only low yield was obtained (Table 1, entry 7). Our reaction condition is not compatible with aqueous hydrogen peroxide oxidant. Furthermore, air was also tested as the low concentration form of oxygen and moderate yield was obtained (Table 1, entry 8). Further solvent screenings showed that CH₃CN is the most suitable solvent for the benzyl alcohol oxidation (Table 1, entries 6, 9–11). Aromatic solvents such as toluene gave slightly compromised yields, presumable due to their lower polarity (Table 1, entry 9). Moreover, oxygen containing solvents, such as THF and ethanol, were not suitable oxidation reaction solvent due to possible metal coordination (Table 1, entries 10 and 11).

Conversion was monitored over time employing various catalyst loadings of Fe₃O₄@P4VP@MIL-100(Fe); the results are summarized in Fig. 8. Relatively low conversion was observed at 1–3 mol% of catalyst loading, indicating the crucial role of the Fe-derived catalyst. However, conversion and yield were not satisfying until the catalyst loading was increased to 5 mol%. As a result, 5 mol% loading of Fe₃O₄@P4VP@MIL-100(Fe) was chosen to be the optimal reaction condition for future studies.

With the optimal reaction conditions in hand, several alcohol substrates were chosen to study the aerobic oxidation catalytic performance of Fe₃O₄@P4VP@MIL-100(Fe). Benzyl alcohol was transformed to the corresponding benzaldehyde in 99% yield after 12 h (Table 2, entry 1). For the aerobic oxidation of benzyl alcohol in presence of Fe₃O₄@P4VP@MIL-100(Fe) as the catalyst, the turnover number (TON) was calculated to be 20. Electron-rich benzyl alcohols, such as p-methyl benzyl alcohol and p-methoxy benzyl alcohol were reactive under the optimal reaction conditions, affording 99% and 92% yield, respectively (Table 2, entries 2 and 3). Alcohol substrates bearing electron-withdrawing functional groups were also tolerated; although a slightly lowered yield was observed for p-fluoro benzyl alcohols (Table 2, entry 4). The heterocyclic alcohol pyridin-2-ylmethanol was converted smoothly to its corresponding aldehyde in the presence of the Fe derived core–shell catalyst (Table 2, entry 5). Cinnamyl alcohol was also evaluated as an example of allylic alcohol; cinnamaldehyde was formed as the only product in 92% yield (Table 2, entry 6). In order to test the transformation of secondary alcohols to ketones, 1-phenylethanol and cyclohexanol were both tested; unfortunately, only low conversions were observed (Table 2, entries 7 and 8).

Table 1  Optimization of the aerobic oxidation reaction conditions

<table>
<thead>
<tr>
<th>Entry</th>
<th>Catalyst</th>
<th>Solvent</th>
<th>Conv.</th>
<th>Sel.</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>—</td>
<td>CH₃CN</td>
<td>&lt;5%</td>
<td>—</td>
</tr>
<tr>
<td>2</td>
<td>Fe₃O₄@P4VP</td>
<td>CH₃CN</td>
<td>&lt;5%</td>
<td>—</td>
</tr>
<tr>
<td>3</td>
<td>FeCl₃</td>
<td>CH₃CN</td>
<td>99%</td>
<td>92%</td>
</tr>
<tr>
<td>4</td>
<td>Fe(NO₃)₃</td>
<td>CH₃CN</td>
<td>99%</td>
<td>93%</td>
</tr>
<tr>
<td>5</td>
<td>MIL-100(Fe)</td>
<td>CH₃CN</td>
<td>86%</td>
<td>99%</td>
</tr>
<tr>
<td>6</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>CH₃CN</td>
<td>99%</td>
<td>99%</td>
</tr>
<tr>
<td>7*</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>CH₃CN</td>
<td>36%</td>
<td>99%</td>
</tr>
<tr>
<td>8**</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>CH₃CN</td>
<td>62%</td>
<td>99%</td>
</tr>
<tr>
<td>9</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>PhCH₃</td>
<td>82%</td>
<td>99%</td>
</tr>
<tr>
<td>10</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>THF</td>
<td>67%</td>
<td>99%</td>
</tr>
<tr>
<td>11</td>
<td>Fe₃O₄@P4VP@MIL-100(Fe)</td>
<td>EtOH</td>
<td>25%</td>
<td>99%</td>
</tr>
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</table>

* Reaction conditions: 1.0 mmol of benzyl alcohol, 5 mol% of Fe₃O₄@P4VP@MIL-100(Fe) catalyst, 5.0 mL of solvent, 0.2 mmol of TEMPO, and 0.2 mmol of KNO₂; 60 °C, 12 h under 1 atm O₂.
** Conversions and selectivities were calculated by GC-MS using nitrobenzene as the internal standard.  

Table 2  Aerobic oxidation of a variety of alcohols

<table>
<thead>
<tr>
<th>Alcohol</th>
<th>Conv.</th>
<th>Sel.</th>
</tr>
</thead>
<tbody>
<tr>
<td>Benzyl alcohol</td>
<td>99%</td>
<td>99%</td>
</tr>
<tr>
<td>p-Methyl benzyl alcohol</td>
<td>99%</td>
<td>99%</td>
</tr>
<tr>
<td>p-Methoxy benzyl alcohol</td>
<td>99%</td>
<td>99%</td>
</tr>
<tr>
<td>p-Fluoro benzyl alcohol</td>
<td>93%</td>
<td>99%</td>
</tr>
<tr>
<td>Benzaldehyde</td>
<td>94%</td>
<td>99%</td>
</tr>
</tbody>
</table>

* Reaction conditions: 1.0 mmol of alcohol, 5 mol% of Fe₃O₄@P4VP@MIL-100(Fe) catalyst, 5.0 mL of solvent, 0.5 mmol of TEMPO, and 0.2 mmol of KNO₂; stirred at 60 °C for 12 h under 1 atm O₂.
characterized. The composite microspheres were composed by a magnetic Fe$_3$O$_4$(PAA) core, a P4VP middle layer, and a MIL-100(Fe) MOF shell. Taking advantage of their porous Fe-derived layer, the magnetic composite microspheres were utilized as an efficient catalyst for the selective aerobic oxidation of alcohols. A variety of alcohol substrates were tolerated under our optimized conditions. The initial catalytic activity of the Fe$_3$O$_4$P4VP@MIL-100(Fe) catalyst was retained after at least ten consecutive reaction cycles. A hot filtration test suggested extremely low leaching of iron into the solution. Further applications of the prepared core–shell MOF catalyst are currently under investigation.

Acknowledgements

We thank the Science Research Foundation of Xijing University (Grant No. XJ16T02), the National Natural Science Foundation of China (No. 51673157) and the BUCT Fund for Disciplines Construction and Development (No. XK1529) for financial support.

Notes and references


