Metal-mediated generation of triazapentadienate-terminated di- and trinuclear \( \mu_2 \)-pyrazolate \( \text{Ni}^{\text{II}} \) species and control of their nuclearity†

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1,3,5-Triazapentadienate-terminated di- and trinuclear nickel(II) complexes featuring bridging azolate ligands, \([\text{Ni}_2(\mu_2\text{-aze})_2(\text{TAP})_2]\) (TAP = \( \text{H} = \text{C(O)Me} \text{NC(O)Me} = \text{NH} \), azole = 3,5-Me$_2$pyrazole 2, 3,5-Ph$_2$pyrazole 3) and \([\text{Ni}_3(\mu_2\text{-aze})_4(\text{TAP})_2]\) (azole = 3,5-Me$_2$pyrazole 4, indazole 5), were obtained from systems Ni$^{II}$/NCNR$_2$/azole systems in MeOH. The terminal TAP ligands in the \([\text{Ni}_2(\mu_2\text{-aze})_2(\text{TAP})_2]\) and \([\text{Ni}_3(\mu_2\text{-aze})_4(\text{TAP})_2]\) species originate from the previously unreported cascade Ni$^{II}$-mediated and chelation-driven reaction between cyanamides and methanol. The oligomeric species and also \([\text{Ni(TAP)}_2]\) (1) are subject to interconversions that depend on the reactants involved and the reaction conditions. The control of the nuclearity of these complexes can be achieved by changing the amount of azoles or by their protonation, alteration of the steric hindrance of the substituents in the heterocycles, and by changing the reaction temperature. Complexes 1–4 were characterized using elemental (C, H, N) analyses, $^1$H, $^{13}$C($^1$H) NMR, FTIR, HRESI-MS, TG-DTA, X-ray crystallography, and 5 was characterized using HRESI-MS and X-ray crystallography. Unconventional metallophilic contacts Ni$^{II}$...Ni$^{II}$ were observed in dimer 3 in the solid state (the distance for Ni...Ni is 2.99 Å, whereas the double Bondi's vdW radius for Ni is 3.26 Å) and the reality of these interactions was confirmed theoretically by the topological analysis of the electron density distribution (AIM method). The estimated energy for these non-covalent Ni...Ni interactions (ca. 4 kcal mol$^{-1}$) fills the gap in the reported energies of the metal...metal interactions in a series comprising of Ni$^{II}$...Ni$^{II}$ (this work), Pd$^{II}$...Pd$^{II}$ (4.3–6.0 kcal mol$^{-1}$), and Pt$^{II}$...Pt$^{II}$ (3.9–11.7 kcal mol$^{-1}$).

Introduction

The abundance of pyrazole(ate)-based ligands is due to the ability of these species to bind metals in a variety of coordination modes, in particular, as bridging ligands. In addition, pyrazole(ate) fragments can be integrated into metal species as they are or as a functionality of more complex polydentate ligands, e.g. polypyrrozolylborates.\(^a\) Metal complexes bearing pyrazole(ates) of different types found applications as catalysts,\(^b\) magnetic,\(^c\)–\(^d\) luminescent,\(^e\)–\(^f\) gas storage or separation\(^g\)–\(^h\) and also as precursors for CVD.\(^i\) Several excellent reviews provide a general picture of the syntheses and application of pyrazole(ate) metal species.\(^1\),\(^2\),\(^3\),\(^4\),\(^5\)

When pyrazolates act as bridging ligands, they are capable of forming metal complexes exhibiting different nuclearity.\(^4\) In particular, nickel(II) centers react with unsubstituted pyrazole (HPz) giving the 1D-chain polymers, \([\text{Ni(HPz)}_2]_n\) (Fig. 1, type A), as two thermoehromic\(^j\) crystalline phases.\(^7\) Nickel(II) complexes featuring bis-pyrazolate ligands form porous MOFs possessing high thermal robustness and exhibiting adsorptive properties and, in particular, MOFs based upon Ni(bpb) (H$_2$bpb = 1,4-(4-bispyrazolyl)benzene) act as a moisture resistant adsorbent toward thiophene.\(^9\) Nickel(II) and 3(5)-arylpyrazoles (H$_4$Pz) form binary cyclic trinuclear complexes \([\text{Ni}_3(\text{H}_4\text{Pz})_3]_n\) (type B).\(^10\)

Examples of oligomeric species incorporating the \([\text{Ni}(\mu_2\text{-aze})_2\text{Ni}]\) fragment supported by other ligands include the dinuclear \([\text{Ni}_3\text{Cp}_2(\text{H}_{12}Pz)_2]_n\) (L = 1/2Cp, type C) and the trinuclear \([\text{Ni}_3\text{Cp}_2(\text{H}_{12}Pz)_2]_n\) (L = 1/2Cp, type D) species that were obtained from the 3,4,5-substituted pyrazoles H$_{12}$Pz and Cp$_2$Ni.\(^11\) The nuclearity of these complexes depends on the bulkiness of the 3- and 5-substituents in the pyrazolate moieties\(^11\) and also on the reaction temperature.\(^11\) It is reported that the sterically hindered substituents and low synthesis temperature favor dimetallic rather than trimetallic species. It is noteworthy that...
Results and discussion

To study the NiII-involving reactions between cyanamides and azole-type heterocycles we chose nickel salts NiX2-nH2O (X = Cl, OTf, n = 0; Br, n = 3; I, n = 6), N,N-disubstituted cyanamides NCNR4 (R2 = Me2, Et2, C6H10, C2H4OC2H4), and also pyrazoles and indazole as depicted in Fig. 2.

We studied the reaction of dialkylcyanamides with the azoles in the presence of nickel(II) salts in methanol, where the nickel salts are well soluble. In these reactions, di- and trinuclear nickel(II) complexes bearing bridging μ3-azolate and terminal 1,3,5-triazapentadione (abbreviated as TAP) ligands were obtained as major products. We observed (see later) that the TAP ligand originates from an unusual cascade of NiII-mediated and chelation-driven cyanamide–methanol integration, and that TAP generation is not affected by the presence of any one of the azoles. Methanol in the reaction serves as both a reagent and a solvent and in view of that we decided to study the NiII-mediated reaction of dialkylcyanamides with different alcohols, whereupon we included the azoles in the tested systems. Our results are sequentially disclosed in the sections that follow.

Nickel(II)-mediated and chelation-driven cascade reaction of NCNR4 and MeOH

The NiII-mediated integration between cyanamides and alcohols was conducted in alcohols R′OH (R′ = Me, Et, Pr). The reactions of NiX2-nH2O and NCNR4 (Scheme 1) were attempted at molar ratios in the range between 1:2 and 1:10 and all tests were conducted either at RT or at 50 °C (for a detailed discussion of the reaction conditions see the ESI†).

We observed that the cascade generation of (TAP)NiII from NCNR4 proceeds in MeOH after maintaining the reaction mixture for two weeks at RT or for 3 d at 50 °C and this reaction gives the complex [Ni(TAP)2] (1; TAP = HN−C(OMe)NC(OMe)−NH) via column chromatography (yields up to 40%; for detailed characterization of 1 see the ESI†).

Bis-chelated complex 1 can be formed in a NiII-mediated and chelation-driven cascade reaction of dimethylcyanamide with...
MeOH. Moreover, MeOH plays the role of an alkylation reagent supporting NR₂/OMe replacement in the formed species, and other products derived from the methylation of the NR₂ group, viz. [Me₂NR₂]X were identified using HRESI-MS (for a detailed discussion of the postulated reaction mechanism see the ESI†).

Complexes of the type M{HN=C(OR)NC(OR)=NH} have been reported for late transition metals, e.g. Co III,22 Ni II,22 Cu II,22,23 Zn II 22b,23ab,24 and Pd II.25 These chelates were obtained either through a chelation-driven reaction of the dicyanamides, (NC)₂NNa or (NC)₂NH, with alcohols in the presence of a transition metal salt (Cu II,14, Zn II,10,12 or Ni II) reacted with HMe₂Pz and NCNR₂ at either RT, or at 50 °C (yield up to 25%). The HPh₂Pz forms dimer 3 in all tested combinations of NCNR₂ and NiX₂ and the highest yield of 3 (34%) was obtained for the system NiCl₂/NCNMe₂/HPh₂Pz (molar ratio 1:4:2). In other cases, the obtained optimal molar ratios between NCNR₂ and the azoles were 4:2 (2) and 1:4 (4) at a 0.1 M concentration of NiX₂ (1 equiv.).

The reaction that involves the unsymmetrical indazole and NCNR₂ should potentially give a mixture of regiosomeric species. Indeed, we observed using both NMR and HRESI-MS the generation of a broad spectrum of products that we failed to separate using column chromatography due to similar retention or to isolate the products in the pure form via fractional crystallization. A few single-crystals of [Ni₂(μ₂-Ind)₄(TAP)]₃ (5) were mechanically picked out from the NiCl₂/NCNMe₂/HPh₂Pz (1:1:4:5, d, 50 °C) reaction mixture and were characterized by XRD (see later). The solid residue obtained after evaporation of the solvent contains a mixture of 5 with other yet unidentified products. The HRESI-MS of this residue displays several groups of peaks corresponding to [HN=C(OMe or NMe₂)NH(NMe₂)₂=NH₃]⁻,

[Ni(C₂H₂N₂)(OMe)₃(NMe₂)₄- (NMe₂)₄-x + H⁺]⁻ (x = 1–4), [Ni₂(μ₂-Ind)₂- (C₂H₂N₂)(OMe)₃(NMe₂)₄-x + H⁺]⁻ (x = 1–3), and [Ni₃(μ₂-Ind)₄- (C₂H₂N₂)(OMe)₃(NMe₂)₄-x + H⁺]⁻ (x = 1–3) (ESI† Fig. S8).

X-ray crystallographic studies for 2–5

Single-crystal XRD studies were performed for 2–5 (Fig. 3–6). The structure of 3 is built up of two crystallographically independent molecular units, whereas all other structures consist of one molecular unit. In the structures of the dimers and trimers, the nickel atoms exhibit square-planar environments formed with four N atoms from the TAP ligand or a substituted azolate (see more information regarding the X-ray crystallographic studies in the ESI†).

Consideration of the crystallographic data suggests the presence of Ni···Ni metalophilic interactions in the di- and trinuclear complexes.

### Nickel(II)-mediated reaction of NCNR₂ with the azoles in alcohols

Mixing NiX₂·nH₂O (X = Cl, OTf, n = 0; Br, n = 3; I, n = 6) with NCNR₂ (R = Me₂, Et₂, C₂H₄, C₆H₄O₂C₂H₄) and the azoles in methanol at RT or 50 °C (see the more detailed discussion later) results in the generation of the di- and trinuclear complexes [Ni₂(μ₂-R'Pz)₂(TAP)₂] (R' = Me 2, Ph 3) and [Ni₃(μ₂-MePz)₄(TAP)] (4) (Scheme 1, b and c). The reactions were performed in all possible combinations of the reactants and molar ratios between NiX₂·nH₂O and NCNR₂ that range from 1:2 to 1:6 and between NCNR₂ and the azoles that are in the interval from 0.5 to 1:4 and two temperatures, i.e. RT and 50 °C.

Although all four cyanamides underwent the reaction, two of them (R₂ = Me₂, C₂H₄O₂C₂H₄) demonstrated greater reactivity and gave higher yields of 2–4; these two cyanamides were also the most efficient in the generation of the TAP ligand (see previous section). The di- and trinuclear complexes were obtained for the substituted pyrazoles, i.e. HMe₂Pz and HPh₂Pz. Among the tested nickel salts, the highest yields of 2–4 were obtained with NiX₂ (X = Cl, OTf). Concurrently, the reaction with the unsubstituted pyrazole (R' = H) after 10 min (MeOH, either RT or 50 °C) gave the previously reported crystalline polymeric α-[Ni(Pz)₂]n (for its identification by powder XRD see Fig. S25, ESI†), which is poorly soluble in all commonly used solvents.

For the systems NiX₂/NCNR₂/azole (X = Cl, OTf, R₂ = Me₂), the composition of the product depends on the temperature, anion X, and substituent R’ in the heterocycle. For HMe₂Pz and NiCl₂ at 50 °C, only trimer 4 was isolated (43%), while at RT both 2 and 4 species were obtained. Pure dimer 2 was obtained when excess of NCNR₂ was applied in the reaction with NiCl₂ (NiCl₂:NCNR₂ ratio was 1:4 or 1:6) at RT and when Ni(OTf)₂ reacted with HMe₂Pz and NCNR₂ at either RT, or at 50 °C (yield up to 25%). The HPh₂Pz forms dimer 3 in all tested combinations of NCNR₂ and NiX₂ and the highest yield of 3 (34%) was obtained for the system NiCl₂/NCNMe₂/HPh₂Pz (molar ratio 1:4:2). In other cases, the obtained optimal molar ratios between NCNR₂ and the azoles were 4:2 (2) and 1:4 (4) at a 0.1 M concentration of NiX₂ (1 equiv.).
trinuclear structures. Indeed, inspection of the XRD data reveals that the interatomic distances between the Ni atoms in 2 (3.10 Å), 3 (2.99 Å), and 4 (3.25 Å) are less than the sum of their Bondi’s (the shortest) van der Waals radii (3.26 Å for Ni + Ni atoms). In order to confirm or deny the assumption on the presence of Ni···Ni metallophilic interactions in 2–4 and to quantitatively estimate their energies we carried out theoretical DFT calculations and performed topological analysis of the electron density distribution within the formalism of Bader’s theory (AIM method; for more detailed information see the ESI†). This approach has already been successfully used by us to study the non-covalent interactions and properties of the coordination bonds in transition metal complexes.28

Theoretical consideration of Ni···Ni metallophilic interactions in 3

Metallophilic interactions impart important properties to functional materials such as luminescence,29 polychromism,30 magnetism,31 or one-dimensional electrical conductivity.32 This phenomenon is generally considered as weak attractive interactions between adjacent metal centers due to overlapping of their d^z^2 and p_z orbitals33 and it has been intensively studied in recent years. Most reports in this area are related to the solid state, where structural parameters obtained by XRD provide precise information on metal···metal separations.

The strength of such M···M interactions is comparable to weak or medium hydrogen bonding, e.g., for Au···Au aurophilic interactions (the strongest one) the appropriate bond energy is ca. 7–12 kcal mol^−1,34 whereas other metallophilic interactions are usually weaker because the influence of relativistic effects in such cases is lower, and the ability of atoms to be involved in this type of bonding increases down through the group. The M^I···M^I (M = Cu, Ag, Au) metallophilic interactions are certainly most studied.35 For group 10 elements, metallophilic interactions of heavy elements (i.e. Pt^II···Pt^II (ref. 36) and Pd^II···Pd^II (ref. 37)) are quite common, whereas Ni^II···Ni^II short contacts were detected for a significantly lower number of systems.12,18 Their metallophilic nature was never confirmed using theoretical studies and, consequently, their energies were not estimated.

The AIM analysis (ESI† Fig. S7) did not reveal Ni···Ni interactions in 2 and 4. However, in 3, exhibiting the shortest Ni···Ni distance, we succeeded to find an appropriate bond critical point (3, −1) (BCP) (Fig. 7). The low magnitude of the
reacting with the appropriate azole and NiCl₂. Thus, complexes 1–4 were converted to monomeric species. Interconversions of 1–4 in 3 is, at least partially, determined by crystal packing effects. Although the analysis of the XRD data for 3 gives dimeric "Ng(NO)₂" complexes and allows the synthesis of well-defined (azolate)₂Ni²⁺ species. Partially our observations are in agreement with those reported previously for azolate systems, whereas the addition of an acid promotes conversion of trimers to dimers and then to [µ²-azolate]²Ni²⁺ and NiCl₂. As it is more sterically demanding, due to the restricted rotational freedom in the complexes, H(3,5-Ph₂)Pz forms only dimer 3, whereas H(3,5-Me₂)Pz gives dimeric 2 and trimeric 4 species. Higher temperature favors the generation of trimeric 4 rather than dimeric 2.

All these observations open up a route to control both the nuclearity of nickel(II)-azole complexes and allows the synthesis of well-defined (azolate)₂Ni²⁺ species. Partially our observations are in agreement with those reported previously for azolate species bearing the terminal CpNi fragment. Although in these two studies the dependence of nuclearity on the amount ofazole/acid was not established, the authors observed the dependence of the nuclearity of azolateNi²⁺ species on both the steric hindrance of azoles and temperature. We expect that our demonstrations the presence of unreacted 1 and some intermediate products, e.g. the cationic [Ni(Ph₂Pz)₂(TAP)]³⁺ species. The HRESI-MS spectra of the reaction mixture comprising of 1, NiCl₂, and IndH (a molar ratio of 1 : 1 : 8; 7 d, 50 ºC, MeOH) exhibit several groups of peaks corresponding to the [TAP + 2H⁺], [Ni(TAP)₂ + H⁺], [Ni₂(µ₂-Ind)(TAP)₂ + H⁺], and [Ni₂(µ₂-Ind)₂(TAP)₂ + H⁺] ions (Fig. S16, ESI†). Therefore, although the formation of indazolate-based oligomers takes place under the reaction conditions, all our attempts to isolate individual products were, unfortunately, unsuccessful.

Dimer 2 was converted to trimer 4 when a mixture of 2 and NiCl₂ (in a molar ratio of 1 : 1) was heated with 3 equiv. NCNMe₂ and excess HMe₂Pz at 110 ºC in DMF for 1 d (yield 52%; Scheme 1, e). Formation of 4 was confirmed using HRESI-MS and powder XRD, while HRESI-MS analysis of the mother liquor demonstrates the presence of unreacted 2.

These interconversions of 1–4 explain the relatively low yields of 2–4 and the strong dependence of the outcome of the reaction on the molar ratio of the reactants. Analysis of the reaction mixtures using HRESI-MS indicates the co-existence of several mono-, di-, and trimeric (azolate)₂Ni²⁺ species, while their corresponding tetramers and nickel-containing species of higher nuclearities were not detected even upon prolonged heating of the reaction mixtures at 50 ºC in DMF or at 140–160 ºC in HOCH₂CH₂OH.

Conclusions

The results of this work can be considered from the following main perspectives. We synthesized two triazapentadienyl-terminated di- and two trinuclear Ni²⁺ complexes featuring the bridging azolate ligands, [Ni₂(µ₂-azolate)₂(TAP)₂] and [Ni₂(µ₂-azolate)₄(TAP)₄] (TAP = HN—C(O)N(C(O)Me)—NH). We observed that these dimers and trimers and also [Ni(TAP)₂] could undergo interconversions depending on the reactants involved and the reaction conditions (Scheme 1).

The nuclearity of the complexes depends on the amount of azoles (or, in contrast, acid) in the system, steric hindrance of the substituents in the ring systems, and the temperature of the reaction. A higher amount of azoles leads to higher nuclearity systems, whereas the addition of an acid promotes conversion of trimers to dimers and then to [Ni(TAP)₂]. As it is more sterically demanding, due to the restricted rotational freedom in the complexes, H(3,5-Ph₂)Pz forms only dimer 3, whereas H(3,5-Me₂)Pz gives dimeric 2 and trimeric 4 species. Higher temperature favors the generation of trimeric 4 rather than dimeric 2.

All these observations open up a route to control both the nuclearity of nickel(II)-azole complexes and allows the synthesis of well-defined (azolate)₂Ni²⁺ species. Partially our observations are in agreement with those reported previously for azolate species bearing the terminal CpNi fragment. Although in these two studies the dependence of nuclearity on the amount ofazole/acid was not established, the authors observed the dependence of the nuclearity of azolateNi²⁺ species on both the steric hindrance of azoles and temperature. We expect that our...
data might be useful in the generation of nickel(II) systems of higher nuclearity in, e.g. solvothermal syntheses.41

Using XRD we detected unconventional nickel(ii)···nickel(ii) metallophilic interactions in 3, and provided the first case where the availability of these metallophilic interactions was confirmed theoretically using a DFT method. Furthermore, we succeeded in getting an estimate of the energy for these non-covalent contacts (3.5–4.4 kcal mol−1)—that is lower than those for Pd(ii)···Pd(ii) (4.3–6.0 kcal mol−1, ref. 37) and Pt(ii)···Pt(ii) (3.9–11.7 kcal mol−1, ref. 36a-g) systems—and it is believed that the existence of these Ni(ii)···Ni(ii) interactions is mainly determined by the packing effects. It has now been proven that the strength of the metallophilic interactions for compounds of group 10 elements increases down through the group according to their ability to overlap d2 and pπ orbitals of the neighboring metal centers.

Experimental

Materials and instrumentation

Solvents, dialkylcyanamides, and nickel(ii) salts NiBr2·3H2O, NiI2·6H2O, and Ni(OTf)2 were obtained from commercial sources and used as received. Anhydrous nickel chloride was prepared according to a previously reported procedure.42 FTIR spectra (4000–400 cm−1) were recorded on a Shimadzu IR Prestige-21 instrument in KBr pellets (Fig. S9–S12 of the ESI†). Microanalyses (C, H, N) were carried out on a Euro EA3028-HT Prestige-21 instrument in KBr pellets (Fig. S9–S12 of the ESI†). GC-MS identification of the by-products from the synthesis of 1 was performed on a Shimadzu GCMS-QP2010 Ultra instrument equipped with a Stabilwax 30 m × 0.32 × 0.50 μm column. The temperature program was set at 70–230 °C with a linear rate of 10 °C min−1 and the column was held at 230 °C for 75 min. 1H and 13C[1H] NMR (400.13 and 100.613 MHz, respectively) spectra were measured on a Bruker Avance 400 spectrometer in Me2SO-d6 (for 1) and CDCl3 (for 2–4) at ambient temperature; residual solvent signals were used as the internal standard (Fig. S17–S24, ESI†). The TG-DTA study (Fig. S28–S31, ESI†) was performed using a NETZSCH TG 209 F1 Libra thermoanalyzer and MnO2 powder was used as a standard. The initial weights of the samples were in the range of 0.9–1.4 mg. The experiments were run in an open aluminum crucible in a stream of argon at a heating rate of 10 K min−1. The final temperature was 610 °C. Processing of the thermal data was performed using Proteus analysis software.44 TLC was performed on Merck Silica gel 60G F254 plates.

X-ray crystal structure determination

For single-crystal XRD experiments all crystals were fixed on a micro mount and placed on an Agilent Technologies SuperNova diffractometer using CuKα monochromated radiation (1, 3, 5) and on a Xcalibur, Eos diffractometer using MoKα monochromated radiation (2, 4); the crystals were measured at a temperature of 100 K. The structures were solved using the direct methods on a SHELXS45 structure solution program (2, 3) and Superflip46 structure solution program (1 and 4). All structures were refined by means of the SHELX program45 incorporated into the OLEX2 program package.47 The crystallographic data and some refinement parameters are given in Table S3 (ESI†). The carbon-bound and nitrogen-bound H atoms were placed in the calculated positions and were included in the refinement in the ‘riding’ model approximation, with Uiso(H) set to 1.5Ueq(C) and C–H 0.96 Å for the CH2 groups, with Uiso(H) set to 1.2Ueq(C) and C–H 0.93 Å for the CH groups, and with Uiso(H) set to 1.2Ueq(N) and N–H 0.86 Å for the NH groups. Empirical absorption correction was applied in the CrystalsPro program complex218 using spherical harmonics, implemented in the SCALE3 ABSPACK scaling algorithm. The unit cell of 3 also contains disordered molecules of the solvent that have been treated as a diffuse contribution to the overall scattering without specific atom positions using SQUEEZE/PLATON.48 Crystallographic data for this paper were deposited at the Cambridge Crystallographic Data Centre and the CCDC numbers are 1440113, 1440114, 1440621, 1442880, and 1453875. Polycrystalline samples were studied in the 2θ range of 5–45° with a Bruker ‘‘D2 Phaser’’ powder diffractometer equipped with CuKα source monochromated radiation.

Computational details

The single point calculations for 2–4 have been carried out at the DFT level of theory using the M06 functional (this functional describes reasonably weak dispersion forces and non-covalent interactions)49 with the help of the Gaussian-09 program package.50 The experimental X-ray geometries were taken as starting points. Two approaches were used in the present work, viz. (i) the Douglas–Kroll–Hess 2nd order scalar relativistic calculations51 requested relativistic core Hamiltonian were carried out using the DZP-DKH basis sets52 for all atoms; (ii) the non-relativistic calculations were carried out using the standard 6-311+G* basis sets for all atoms. The topological analysis of the electron density distribution with the help of the atoms in molecules (AIM) method developed by Bader has been performed by using the Multiwfn program (version 3.3.4).53 The Cartesian atomic coordinates for 2–4 are given in Table S2 (ESI†).

Syntheses and characterization

Generation of [Ni(TAP)2] (1). Any one of the nickel(ii) salts NiX2·nH2O (X = Cl, OTf, n = 0; Br, n = 3; I, n = 6; 0.5 mmol) was dissolved in MeOH (3 mL), and NCNR2 (R2 = Me2 or C2H4OC2H4; 2 mmol) was added to the solution. The green homogeneous reaction mixture was left to stand in air at RT for 2 weeks (or at 50 °C for 3 d) in a closed flask. The color of the reaction mixture gradually turned to bright-orange (X = Br, I) and a light-green precipitate was formed. In the case of X = Cl, the reaction mixture remained homogenous. The colorless crystals of [M2N]X (X = Cl, Br, I) and/or [Me2NC2H4O]X (X = Cl, Br, I) were
precipitated from the reaction mixtures obtained at 50 °C. The reaction mixture (X = Cl; RT) or mother liquor, separated via filtration from the residue, was left to stand to undergo slow evaporation in air and orange crystals were formed after the complete evaporation of the solvent (2 d, RT). The crystalline product was dissolved in diethyl ether (3 mL) and the obtained solution was passed through a column filled with Merck silica gel 60 (eluent EtOAc : hexane, 10 : 1, v/v). The red fraction was collected and evaporated to dryness at RT in air giving a crystalline product, which was dried over CaCl₂ in a desiccator at 20–25 °C.

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\text{Ni}^{2+}\text{[TAP]}_2(1)\] (57.5 mg, 35% (at RT), 63 mg, 39% (at 50 °C)). \(R_2 = 0.60\) (acetone : hexane, 10 : 1, v/v). The complex sublimes at 175 °C upon heating. Calcd. for \(\text{C}_{567.1231}\), \(817.1779\) ([M + Ni + 2Me₂Pz + H] +, requires 817.1763).

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\text{MeOH (5 mL), and NCNR}_2 (R_2 = \text{Me}_2\text{C}_2\text{H}_4\text{OC}_2\text{H}_4; 2 \text{ mmol}) and 13.5 (dried over CaCl₂ in a desiccator at 20–25 °C).}
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163.8 (C–H); 1614 vs. (C–N). 1H NMR (CDCl₃, 400 MHz): 6.38 (s, 1H, C–H(Pz)), 5.38 (s, 6H, OCH₃), 3.82 (s, 3H, CH₃), 3.78 (s, 6H, OCH₃), 2.46 (s, 6H, C–CH₃). 13C{1H} NMR (CDCl₃, 100 MHz): 163.8 (C–N), 148.0 (CH(Pz)), 105.4 (C(Pz)), 54.3 (CH₂O), 13.5 (CH₃).
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\text{Ni}^{2+}\text{[Me₂Pz]}_2(1)\] (3). 71 mg, 34%. Mp: 295 °C (dec.). Anal. calcd. for \(\text{C}_{413.0839}\text{N}_{10}\text{O}_{12}\text{Ni}_2; C, 38.07; H, 5.32; N, 24.66. Found: C, 37.94; H, 5.33; N, 23.96%. HRESI-MS, \(m/z\): 319.0669 \([\text{Ni}[\text{TAP}]_2 + H]^+\), requires 319.0670. 567.1173 ([M + 2H] +, requires 567.1231). 1H NMR (CDCl₃, 500 MHz): 6.39 (s, 1H, C–H(Pz)), 5.38 (s, 6H, OCH₃), 3.82 (s, 3H, CH₃), 3.78 (s, 6H, OCH₃), 2.46 (s, 6H, C–CH₃). 13C{1H} NMR (CDCl₃, 125 MHz): 163.8 (C–N), 148.0 (CH(Pz)), 105.4 (C(Pz)), 54.3 (CH₂O), 13.5 (CH₃).
...requires 319.0670). 0.015 mmol) and NiCl₂ (1.9 mg, 0.015 mmol) were dissolved in RT in a mixture of DMF (1.5 mL) and NCNMe₂ (3.2 mg, 0.046 mmol), whereupon HMe₂Pz (18.4 mg, 0.19 mmol) was added. The homogeneous yellow reaction mixture formed was left to stand at 50 °C in a closed flask. After 1 d the yellow-orange precipitate released, was decanted, washed with three 1.5 mL portions of MeOH and with three 1.5 mL portions of Et₂O. The crystalline product dried over CaCl₂ in a desiccator at RT in a mixture of DMF (1.5 mL) and NCNMe₂ (3.2 mg, 0.046 mmol), whereupon glacial acetic acid (0.8 mg, 0.013 mmol) were used. The yield of conversion of 4 to 1.

**Conversion of 2 to 1.** Complex 2 (14.2 mg, 0.025 mmol) was dissolved in CHCl₃ (1 mL) at RT, whereupon glacial acetic acid (1.2 mg, 0.02 mmol) was added. The bright yellow homogeneous solution formed was left to stand at 50 °C in a closed flask. After 2 d the color of the reaction mixture turned to pale yellow and a pale blue amorphous precipitate was released, which was centrifuged and the mother liquor was evaporated at RT to dryness. Pink crystals were grown from the mother liquor. The yield of 1 is 1.5 mg, 19% based on Ni. HRESI+, m/z: 567.1253 ([Ni₂(Me₂Pz)₂(TAP)₂]+ + H⁺), requires 567.1231), 817.1793 ([M + H]+, requires 817.1763). The crystalline phase of [Ni₃((μ₂-Me₂Pz)2(TAP))₂]²⁺ identified by powder XRD (Fig. S27, ESI).

**Conversion of 4 to 1.** This conversion was conducted similarly to the previous method. Complex 4 (11.3 mg, 0.014 mmol) and glacial acetic acid (0.8 mg, 0.013 mmol) were used. The yield of 1 is 1.0 mg, 22% based on Ni. HRESI+, m/z: 319.0660 ([M + H]+), requires 319.0670).

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