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Sensing a *Bacillus anthracis* biomarker with well-known OLED emitter EuTta₃Phen†‡

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Dipicolinic acid (DPA) is a distinctive biomarker for bacterial spores. Here, we present the successful demonstration of dramatic Switch-OFF sensing of DPA using an easily synthesised Eu(III) phosphor applied primarily in light-emitting devices. The sensor in the presence of water and phosphate is also demonstrated to be effective.

Luminescent lanthanide (Ln) complexes are an important class of emitter with applications in organic light-emitting diodes (OLEDs),¹ sensors,² and cellular imaging³ due to their distinct and narrow f-f emission bands, long lifetimes and broad Stokes shifts.⁴ Quite often, however, the emission intensity of lanthanide compounds can be weak due to their poor molar absorptivity, due to Laporte forbidden f-f transitions. This is especially true with weakly coordinated Ln aqua cations.⁵ Introduction of appropriate ligands (usually with extended π -systems) to the Ln³⁺ trication can activate highly efficient emission profiles by sensitization of the Ln center with a more absorptive ligand 'antenna'. This process has been termed Absorption-Energy Transfer-Emission (AETE).^{6,7} One distinct advantage to employing sensitised Ln³⁺ emitters is their low sensitivity to molecular oxygen, which can easily quench transition metal-centred phosphorescence.⁸ This is a crucial design consideration as it relates to practical small molecule sensors. Of particular interest, is the detection of 2,6-pyridine-dicarboxylic acid or dipicolinic acid (DPA; Fig. 1).^{9–12} DPA is a unique bio-marker for bacterial spore detection, with bacterial endospores comprising up to 10% by mass of DPA,¹³ the *Bacillus anthracis* spore, for instance, is the endospore vessel used in the transmission of the anthrax bioweapon.¹⁴

A very nice example of such a Ln³⁺-based sensor was first introduced in 2007 by Ponce and co-workers, outlining a simple synthetic route to a coordinated terbium(III)-based Turn-ON

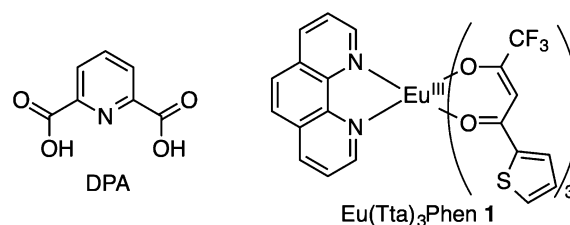


Fig. 1 Chemical structures of 2,6-pyridine dicarboxylic acid (dipicolinic acid or DPA; analyte) and EuTta₃Phen (**1**; sensor).

sensor (using cyclen derivative DO2A),⁹ where the detection of DPA sensitises the Ln-centre, activating the very efficient AETE pathway. The authors later demonstrated that europium(III) was a much less effective Turn-ON sensor than Tb³⁺, likely due to a larger energy gap (thus weakly coupled) between the DPA triplet state and the Eu³⁺ ⁵D₀ excited state.¹² With this consideration in mind, we set out to exploit this gap by alternatively identifying a strongly emissive Eu(III) complex that could act as a dramatic and sensitive Turn-OFF sensor.

In this communication we outline the use of well known Eu³⁺ complex as a potential DPA sensing material. As noted above, Eu³⁺ is not effectively sensitised by DPA, therefore, our approach is to use an easily synthesised emitter (one-pot) that is highly sensitised with the expectation that emission will be quenched in the presence of DPA. Europium³⁺(thenoyltrifluoroacetato)₃(1,10-phenanthroline), or EuTta₃Phen (**1**; first reported in 1995; Fig. 1),^{15,16} is still one of the brightest red emitters, and has been widely used in red-emitting OLED materials.¹ We postulated that with a quantum yield of approximately $\Phi = 0.48$ in dichloromethane ($\Phi = 0.72$ in the solid-state),¹⁷ interference by a competitive tridentate ligand such as DPA would have a visually noticeable effect on the emission intensity.

Given the surrounding environment associated with bacterial spore analysis, the effects of water, phosphate, and DPA were assessed against **1** separately and also together. Respective UV-Vis titrations with both water and *N,N,N,N*-tetrabutylammonium phosphate monobasic (TBA-H₂PO₄) into solutions of THF containing Eu(Tta)₃Phen **1** resulted in diminished absorbance with little

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† This submission is dedicated to the late Prof. Richard F. Langler for his contributions to undergraduate research. Fittingly, the experimental work in this account was fully carried out by final year undergraduate project students.

‡ Electronic supplementary information (ESI) available: Supporting UV-Vis, fluorescence emission, and NMR spectra. See DOI: 10.1039/c5tb02469b



Fig. 2 UV-Vis titration experiments where $\text{Eu}(\text{Tta})_3\text{Phen}$ (1×10^{-5} M; THF) is titrated with (a) DPA (2×10^{-4} M; THF) and (b) DPA (2×10^{-4} M; THF) in the presence of 2% water and 0.5 equivalents of $\text{TBA} \cdot \text{H}_2\text{PO}_4$ (2×10^{-4} M) at 298 K.

indication of any distinct reactive process or ligand exchange (refer to Fig. S1 and S2 of ESI†). However, titration with DPA exhibited a clear interaction with the appearance of an isosbestic point at 326 nm, altering the absorbance of the parent species (Fig. 2a). This was further reinforced by titration of **1** in 2% H_2O and phosphate (1×10^{-4} M) with DPA, exhibiting a near identical profile with additional $\text{H}_2\text{O}/\text{H}_2\text{PO}_4^-$ equilibria (Fig. 2b), indicating a change in the substrate composition.

As eluded above, we then performed emission studies to further test the efficacy of this system by observing how the introduction of DPA alters the fluorescence spectra of lanthanide complex **1**. In its unperturbed state, **1** exhibits characteristic $\text{Eu}(\text{III})$ $^5\text{D}_0 \rightarrow ^7\text{F}_n$ ($n = 0-4$) transitions. Upon addition of water (0–20% v/v; Fig. S4, ESI†) complete quenching of **1** is exhibited above 10% water content in THF. Emission loss is likely a result of a combination of proton transfer and vibrational decay processes (through the ninth coordination site). This is supported by minimal perturbations in the corresponding UV-Vis titration displaying little evidence of any ligand exchange. Surprisingly, titration of **1** with $\text{TBA} \cdot \text{H}_2\text{PO}_4$ (Fig. S5, ESI†) showed an increase in emission intensity by approximately 20%. Repeating the experiment with a solvent mixture of 2% water in THF, however, exhibited a modest decrease in emission intensity (Fig. S6, ESI†), though far from a complete emission quench.

Introduction of DPA to a THF solution of **1** showed a remarkable decrease in emission (Fig. 3a), particularly the strong

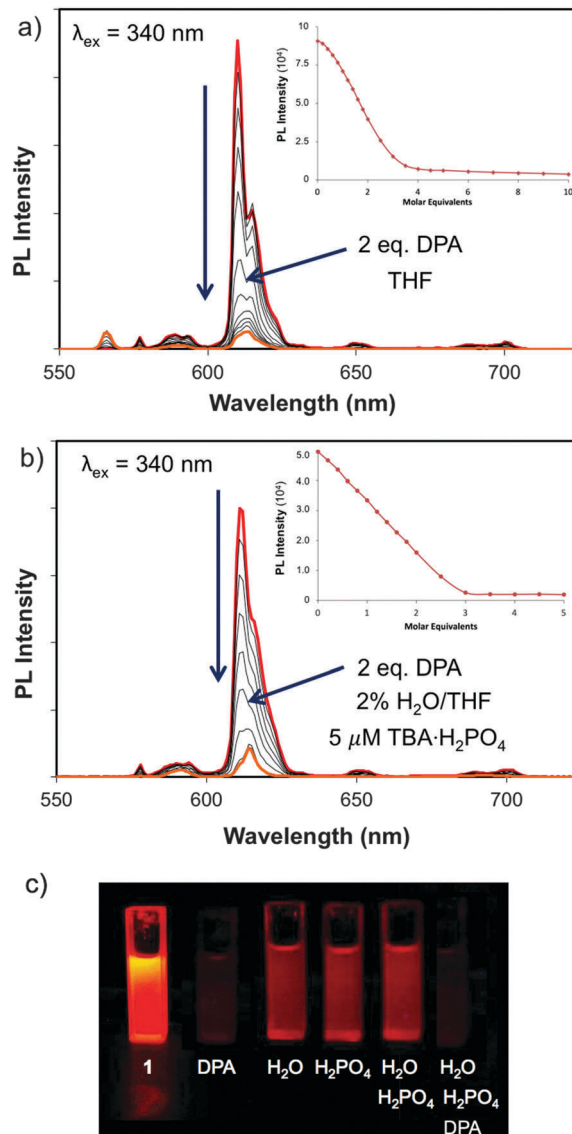


Fig. 3 Fluorescence emission titration experiments where **1** (1×10^{-5} M; THF) is titrated with (a) DPA (2×10^{-4} M; THF) and (b) DPA (2×10^{-4} M; THF) in the presence of 2% water and 0.5 equivalents of $\text{TBA} \cdot \text{H}_2\text{PO}_4$ (5×10^{-6} M) at 298 K ($\lambda_{\text{ex}} = 340$ nm). Reaction profiles are illustrated in ESI†. (c) Visual comparison of emission properties of **1** upon exposure to DPA, H_2O , and/or $\text{TBA} \cdot \text{H}_2\text{PO}_4$. Insets in (a) and (b): emission response as a function of increasing DPA concentration.

$^5\text{D}_0 \rightarrow ^7\text{F}_2$ transition. From the reaction profile (Fig. S7, ESI†) it is evident that emission is fully quenched upon full saturation of the Eu^{3+} coordination sphere with three molar equivalents of DPA. The robustness of the system was also successfully assessed with titrations of **1** with DPA in a 2% water/THF mixture (Fig. S8, ESI†), and moreover in a 2% water/THF mixture with a background concentration of $\text{TBA} \cdot \text{H}_2\text{PO}_4$ (5×10^{-6} M; Fig. 3b). These experiments demonstrate that with modest water and phosphate concentrations present, sensitivity and selectivity of **1** as a DPA sensor remain intact, and that this small allowance for water and phosphate content should translate to **1** performing as an effective practical in-the-field sensor. Fig. 3c illustrates how **1**

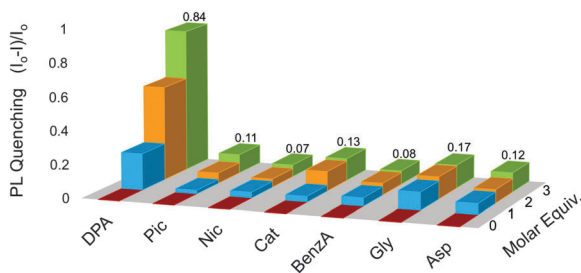


Fig. 4 Selectivity study by fluorescence emission spectrophotometry assessing impact of model analytes on emission intensity of **1** (1×10^{-5} M; $\lambda_{\text{ex}} = 340$ nm; $\lambda_{\text{em}} = 609$ nm; THF; 298 K). DPA: dipicolinic acid; Pic: 2-picolinic acid; Nic: nicotinic acid; Cat: catechol; BenzA: benzoic acid; Gly: glycine; Asp: L-aspartic acid.

performs as visual sensor for DPA detection in the absence and presence of water and phosphate. In order to probe the selective utility of **1** we further undertook a small selectivity study that included to amino acids (glycine and L-aspartic acid) as well as a representative selection of various competitive protic/bidentate substrates (benzoic acid, catechol, nicotinic acid, 2-picolinic acid) and evaluated these results against those of DPA (Fig. 4). This analysis demonstrated that a number of biologically or structurally relevant small molecule analytes have little to no effect on the emission profile of **1** compared to the extent of quenching imparted by DPA. Lastly, we assessed the sensitivity of the system, and can demonstrate that DPA does enact a spectroscopically measurable emission quench on **1** at nanomolar concentrations (Fig. S9 (ESI †); titrating **1** (10 nM) with small aliquots of DPA (2×10^2 nM)). It has been established that there is approximately 1×10^8 molecules of DPA per bacterial spore,¹⁰ and thus based on crude mathematics, this corresponds to the detection of as few as 100 spores by employing **1** as a sensor.

Mechanistically, there are two initial pathways for ligand exchange: displacement of 1,10-phenanthroline (Phen) for DPA; displacement of thenoyltrifluoroacetone (TtaH) for DPA. Exchange of the ET ligand, Phen, for DPA would exhibit a complete loss of emission after one molar equivalent of added DPA,^{1,18} destroying the AETE process. The non-equimolar stoichiometry observed in the fluorescence emission spectra leads one to believe that the ligand acting as the 'antenna' is being displaced first, reducing the efficiency of AETE, but not eliminating it until higher concentrations of DPA are reached. Assessment of the quenching mechanism, which will enable future improvements to the sensor, led to a series of ^1H NMR studies. Initially, we wanted to see if a ligand exchange process could be identified from THF- d_8 solutions of **1** titrated with DPA, however due to the paramagnetic characteristics of Eu^{3+} , broadening of the signals made the protons indistinguishable (Fig. S10, ESI †). A diamagnetic lanthanum(III) analogue was then synthesised, $\text{LaTta}_3\text{Phen}$ (**2**),¹⁹ and treated with DPA. ^1H NMR analysis confirmed that upon addition of DPA to a sample of **2** (the La^{3+} -DPA product precipitates from solution due to higher NMR concentrations), unreacted **2** and freshly displaced TtaH are observed in the spectrum (Fig. S11, ESI †). This study gives us insight into the behaviour of **1** when exposed to DPA, it does

confirm the reaction stoichiometry and why the loss of emission is not immediate.

In summary, we have identified that the easily synthesised **1** can act as an effective Turn-OFF sensor in the detection of 2,6-pyridinedicarboxylic acid (DPA), a prevalent biomarker in bacterial spores. We further demonstrated that this lanthanide-based emitter can continue to function effectively in the presence of small quantities of both water and phosphate, is selective for DPA, and is detectable down to nanomolar concentrations. The mechanism for emission quenching was also revealed to proceed first *via* ligand displacement of the thenoyltrifluoroacetone-ligand by DPA. We are currently developing analogues of **1** to improve stability for incorporation into various polymer matrices.

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Notes and references

- 1 J. Wang, R. Wang, J. Yang, Z. Zheng, M. D. Carducci, T. Cayou, N. Peyghambarian and G. E. Jabbour, *J. Am. Chem. Soc.*, 2001, **123**, 11378.
- 2 L. F. Smith, B. A. Blight, H.-J. Park and S. Wang, *Inorg. Chem.*, 2014, **53**, 8036.
- 3 A. Thibon and V. C. Pierre, *Anal. Bioanal. Chem.*, 2009, **394**, 107.
- 4 H. C. Aspinall, *Chemistry of the f-Block Elements*, Gordon and Breach Science Publishers, Amsterdam, 2001.
- 5 J. Georges, *Analyst*, 1993, **118**, 1481.
- 6 M. L. Cable, D. J. Levine, J. P. Kirby, H. B. Gray and A. Ponce, *Advances in Inorganic Chemistry*, Elsevier, San Diego, 2011, vol. 63, pp. 1–45.
- 7 K. Binnemans, *Chem. Rev.*, 2009, **109**, 4283.
- 8 N. Kaltsoyannis and P. Scott, *The f-elements*, Oxford University Press Inc., New York, 1999.
- 9 M. L. Cable, J. P. Kirby, K. Sorasaene, H. B. Gray and A. Ponce, *J. Am. Chem. Soc.*, 2007, **129**, 1474.
- 10 M. L. Cable, J. P. Kirby, D. J. Levine, M. J. Manary, H. B. Gray and A. Ponce, *J. Am. Chem. Soc.*, 2009, **131**, 9562.
- 11 I. Lee, W.-K. Oh and J. Jang, *J. Hazard. Mater.*, 2103, **252–253**, 186.
- 12 H. Tan, C. Ma, L. Chen, F. Xu, S. Chen and L. Wang, *Sens. Actuators, B*, 2014, **190**, 621.
- 13 W. G. Murrel, *The Bacterial Spore*, Academic Press, New York, 1969.
- 14 J. A. Jernigan, D. S. Stephens, D. A. Ashford, C. Omenaca, M. S. Topiel, M. Galbraith, M. Tapper, T. L. Fisk, S. Zaki, T. Popovic, R. F. Meyer, C. P. Quinn, S. A. Harper, S. K. Fridkin, J. J. Sejar and C. W. Shepard, *Emerging Infect. Dis.*, 2001, **7**, 933.
- 15 T. Sano, M. Fujita, T. Fujii, Y. Hamada, K. Shibata and K. Kuroki, *Jpn. J. Appl. Phys., Part 1*, 1995, **34**, 1883.
- 16 L. Liu, Z. Hong, J. Peng, X. Liu, C. Liang, Z. Liu, J. Yu and D. Zhao, *Synth. Met.*, 1997, **91**, 267.
- 17 C. Freund, W. Porzio, L. Giovannella, F. Vignali, M. Pasini and S. Destri, *Inorg. Chem.*, 2011, **50**, 5417.
- 18 J.-C. G. Bünzli and C. Piquet, *Chem. Soc. Rev.*, 2005, **34**, 1048.
- 19 G. Zhong, Y. Feng, K. Yang and G. Zhu, *Chem. Lett.*, 1996, 775.

