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Perfluoroalkyl-substituted conjugated polymers as electron acceptors for all-polymer solar cells: the effect of diiodoperfluoroalkane additives†

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A series of six diketopyrrolopyrrole (DPP) based conjugated polymers with a varying content of solubilizing perfluoroalkyl chains were synthesized. Based on a systematic investigation of the influence of the solvent on the photovoltaic performance, it is found that 1,6-diiodoperfluorohexane (IC₆F₁₂I) is an effective solvent additive to enhance the power conversion efficiency (PCE) of DPP polymers with perfluoroalkyl side chains. The polymers consist of thiazole-flanked DPP units that alternate along the main chain with varying ratios of thiophene (T) and perfluoroalkyl benzodithiophene (FBDT) units. The polymers possess high molecular weights, narrow band gaps and good crystalline properties. The DPP polymers were used as electron acceptors in bulk heterojunction solar cells with another DPP polymer as the electron donor. A solvent mixture of CHCl₃ : 1-chloronaphthalene (1-CN) is found to provide the best PCE of 2.9% in non-fluorine based DPP polymer solar cells, but yields a low PCE of 0.52% for perfluoroalkyl-containing polymer solar cells. Perfluoroalkyl-containing polymer solar cells fabricated from CHCl₃ with IC₆F₁₂I as the processing additive show a significantly improved PCE of 2.1%. The morphology analysis of the blend films reveals that IC₆F₁₂I as an additive improves the micro-phase separation between the polymer donor and acceptor, which results in enhanced charge generation.

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Introduction

There has been increasing interest in non-fullerene conjugated materials to replace fullerene derivatives as electron acceptors in organic solar cells in recent years.^{1–3} A significant number of conjugated small molecules^{4–20} and polymers^{21–35} with excellent electron transport properties and aligned energy levels have been designed and synthesized, which are well-suited as electron acceptors in organic photovoltaic devices. Currently, perylenediimide and naphthalenediimide based materials are considered as the most promising non-fullerene acceptors,^{36,37} since these materials have shown high electron mobilities of about 1 cm² V⁻¹ s⁻¹ (ref. 38) and deep lowest unoccupied

molecular orbital (LUMO) levels which are similar to those of fullerene derivatives. As a consequence, perylenediimide- and naphthalenediimide-based small molecules and conjugated polymers as electron acceptors in organic solar cells show impressively high power conversion efficiencies (PCEs) above 7%,^{4–6,39} approaching the performance of fullerene-based solar cells with PCEs over 10%.^{40,41}

In terms of the large possible variation in materials, conjugated polymers possess an advantage compared to fullerene derivatives, such as [6,6]-phenyl-C₆₁-butyric acid methyl ester (PCBM), for which variations in the optical band gap and LUMO energies are more restricted. In recent years, conjugated polymers have been widely exploited as electron donors in polymer solar cells (PSCs),^{42,43} and design principles from these studies can also be utilized for the development of new acceptor polymers. In particular, when introducing electron-rich and electron-deficient moieties into a conjugated backbone, the so-called donor–acceptor polymers exhibit distinct variations in frontier orbital energy levels, absorption spectra, and charge-carrier mobilities, which allow the conjugated polymers to be used as electron donors or acceptors in solar cell devices. For instance, the diketopyrrolopyrrole (DPP) unit is widely used as an electron-deficient building block to construct narrow band gap polymers with near-infrared absorption.^{44,45} DPP polymers also exhibit high hole mobilities, exceeding 10 cm² V⁻¹ s⁻¹ in favourable cases,⁴⁶ and impressive crystallinity,^{47,48} such that

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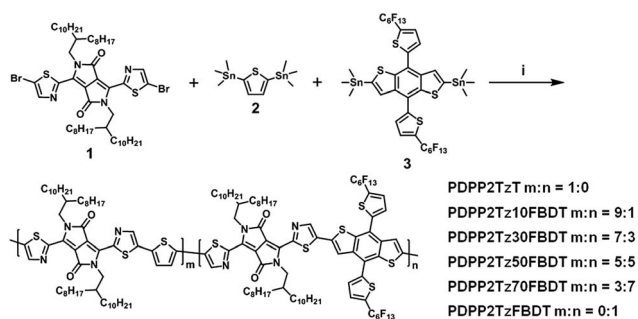
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PSCs based on DPP polymers reach PCEs above 8%.^{49–51} Meanwhile, organic field-effect transistors (FETs) based on DPP polymers also present high electron mobilities above $5 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$,^{46,52} which are higher than those of fullerene derivatives,⁵³ indicating their potential for use as electron acceptors in PSCs. One recent example is the design of a polymer in which the DPP is flanked by two thiazole rings to effectively lower the LUMO level.³⁴ The resulting DPP polymer can act as an electron acceptor in combination with another DPP polymer as the electron donor in PSCs to reach a PCE of 2.9%. At present this represents the highest PCE when using DPP polymers as electron acceptors. It will be of importance to further explore new DPP polymer acceptors for efficient photovoltaic devices.

Recently, we designed and synthesized a DPP polymer acceptor PDPP2TzFBBDT (Scheme 1) that has similar frontier energy levels to those of PCBM and can be potentially used as a universal acceptor for PSCs.⁵⁴ However, our initial attempt to use PDPP2TzFBBDT as an electron acceptor with PDPP5T³⁴ as an electron donor was not successful and resulted in very poor performance with PCEs of 0.19%. PDPP2TzFBBDT has long perfluoroalkyl chains that are partially responsible for the deep LUMO levels, but the lipophobic perfluoroalkyl chains also cause poor miscibility with PDPP5T, resulting in large phase separation in blends of these two polymers, poor charge generation, and consequently low PCEs. Similar behaviour was also reported in donor polymer : fullerene systems, in which the donor polymers that bear long perfluoroalkyl chains showed large phase-separated domains and hence provide poor PCEs in PSCs.^{55,56} Therefore, it is important to find a way to improve the micro-phase separation in these blend films in order to apply perfluoroalkyl-based conjugated polymers in organic photovoltaic devices.

In this work, we explore the use of diiodoperfluoroalkanes as additives in solution-processed PSCs based on perfluoroalkyl DPP acceptor polymers. Based on the similarity principle, we assume that polymers with perfluoroalkyl units have better solubility in perfluoroalkanes as additives, which would prevent the fast precipitation and therefore reduce the domain size of the polymers. A similar function of *o*-dichlorobenzene (*o*-DCB) as an additive in DPP polymer solar cells was also observed.⁵⁷



Scheme 1 Chemical structures of the DPP polymers and their synthetic routes. (i) Stille polymerization by using $\text{Pd}_2(\text{dba})_3/\text{PPH}_3$ in toluene/DMF (10 : 1, v/v) at 115°C . Note: $m : n$ was determined from the feed ratio of the monomers 2 and 3.

We synthesized several DPP acceptor polymers, in which the thiazole-flanked DPP segment was linked with thiophene (T) and perfluoroalkyl benzodithiophene (FBBDT) units in different ratios (Scheme 1). The polymers were found to have small optical band gaps (E_g) and good crystalline properties, depending on the ratio of T and FBBDT units. The resulting polymers were used as electron acceptors in polymer–polymer solar cells with PDPP5T as the electron donor, which were solution-processed from chloroform (CHCl_3) with 1-chloronaphthalene (1-CN) or -diiodoperfluoroalkanes as processing additives. The solar cells based on PDPP5T : PDPP2TzFBBDT fabricated from CHCl_3 : $\text{IC}_6\text{F}_{12}\text{I}$ exhibited a much improved PCE of 2.1% compared to cells from CHCl_3 or CHCl_3 : 1-CN due to their better micro-phase separation. The results reveal that diiodoperfluoroalkanes can be used to effectively tune the morphology of perfluoroalkyl-based polymer solar cells. This encourages the design of perfluoroalkyl-based conjugated polymers for efficient solar devices.

Experimental

Materials and measurements

All synthetic procedures were performed under an argon atmosphere. Commercial chemicals were used as received.

THF and toluene were distilled from sodium under a nitrogen atmosphere. The monomers 3,6-bis(5-bromothiazol-2-yl)-2,5-bis(2-octyldodecyl)-2,5-dihydropyrrolo[3,4-*c*]pyrrole-1,4-dione (1)⁵⁴ and (4,8-bis(5-perfluorohexylthiophen-2-yl)benzo[1,2-*b*:4,5-*b'*]dithiophene-2,6-diyl)bis(trimethylstannane) (3),⁵⁴ and the polymers PDPP2TzT,³⁵ PDPP2TzFBBDT⁵⁴ and PDPP5T⁵⁸ were synthesized according to the literature procedures. 2,5-Bis(trimethylstannyl)thiophene (2) was purchased from Aldrich Chemical Co. and recrystallized from methanol before use in polymerization reactions.

$^1\text{H-NMR}$ spectra of the polymers were recorded on a Bruker AVIII 500WB NMR Spectrometer at 100°C with 1,1,2,2-tetrachloroethane- d_2 as the solvent and tetramethylsilane (TMS) as the internal standard. The molecular weight was determined by gel permeation chromatography (GPC) at 140°C on a PL-GPC 220 system (Agilent Technologies with a Knauer PDA detector) using a PLgel 10 μm MIXED-B LS column and *o*-DCB as the eluent against polystyrene standards. A low concentration of 0.1 mg mL^{-1} polymer in *o*-DCB was applied to reduce aggregation. Optical absorption spectra were recorded on a JASCO V-570 spectrometer with a slit width of 2.0 nm and a scan speed of 1000 nm min^{-1} . Cyclic voltammetry (CV) was performed under an inert atmosphere at a scan rate of 0.1 V s^{-1} and 1 M tetrabutylammonium hexafluorophosphate in acetonitrile as the electrolyte. An ITO glass slide covered with a thin layer polymer (approx. 20 nm) was used as the working electrode. The counter and reference electrodes were a Pt wire and Ag/AgCl, respectively. Atomic force microscopy (AFM) images were recorded using a Digital Instruments Nanoscope IIIa multimode atomic force microscope in the tapping mode under ambient conditions. The tips were purchased from Bruker (Model: SCANA-SYST-AIR with one cantilever, $T = 650 \text{ nm}$, $L = 115 \mu\text{m}$, $W = 25 \mu\text{m}$ and spring constant of 0.4 N m^{-1}). 2D grazing-incidence



wide angle X-ray scattering (2D-GIWAXS) measurements were performed by using a Xenocs WAXS/SAXS system, with an X-ray wavelength of 1.5418 Å. The incident angle was 0.2°. The sample-to-detector distance was 127.5 mm. The scattered X-rays were detected by using a Dectris Pilatus 100k counting detector. The counting time was 4 h. All film samples were prepared by spin-coating solutions on Si/SiO₂ substrates. Steady state fluorescence spectra were recorded at room temperature using an Edinburgh Instruments FLS980 double-monochromator luminescence spectrometer equipped with a nitrogen-cooled near-IR sensitive photomultiplier (Hamamatsu).

Photovoltaic devices with an inverted configuration were made by spin coating a ZnO sol-gel at 4000 rpm for 60 s onto pre-cleaned, patterned ITO substrates. The photoactive layer was deposited by spin coating a chloroform solution containing PDPP5T and thiazole-based DPP polymers and an appropriate amount of processing additives such as 1-CN, 1,4-diodoperfluorobutane (IC₄F₈I), 1,6-diodoperfluorohexane (IC₆F₁₂I) or 1,8-diodoperfluorooctane (IC₈F₁₆I) in air. MoO₃ (10 nm) and Ag (100 nm) were deposited by vacuum evaporation at *ca.* 4 × 10⁻⁵ Pa as the back electrode.

The active area of the cells was 0.04 cm². The *J-V* characteristics were measured by using a Keithley 2400 source meter unit under AM1.5G spectrum from a solar simulator (Enlitech model SS-F5-3A). The illumination intensity was determined at 100 mW cm⁻² using a monocrystalline silicon reference cell with a KG5 filter. The short-circuit current density (*J*_{sc}) under AM1.5G conditions was estimated from the spectral response and convolution with the solar spectrum. The external quantum efficiency was measured by using a Solar Cell Spectral Response Measurement System QE-R3011 (Enli Technology Co., Ltd.). The thickness of the active layers in the photovoltaic devices was measured on a Veeco Dektak XT profilometer.

PDPP2Tz10FBDT

To a degassed solution of monomer 1 (83.66 mg, 0.082 mmol), monomer 2 (30.21 mg, 0.074 mmol) and monomer 3 (10.78 mg, 0.008 mmol) in toluene (2 mL) and DMF (0.2 mL), tris(dibenzylidene-acetone)dipalladium(0) (2.25 mg, 2.5 μmol) and triphenylphosphine (2.58 mg, 9.8 μmol) were added. The mixture was stirred at 115 °C for 24 h, after which the reaction mixture was precipitated in methanol and filtered through a Soxhlet thimble. The polymer was extracted with acetone, hexane and chloroform. The chloroform fraction was reduced and the polymer was precipitated in acetone. The polymer was collected by filtering over a 0.45 μm PTFE membrane filter and dried in a vacuum oven to yield PDPP2Tz10FBDT (80 mg, 94%) as a dark solid. GPC (*o*-DCB, 140 °C): *M*_n = 77.2 kg mol⁻¹, PDI = 2.05.

PDPP2Tz30FBDT

The same procedure as for PDPP2Tz10FBDT was used, but now 1 (49.68 mg, 0.049 mmol), 2 (13.95 mg, 0.034 mmol) and 3 (19.21 mg, 0.015 mmol) were used as the monomers. Yield: 55 mg (93%). GPC (*o*-DCB, 140 °C): *M*_n = 65.1 kg mol⁻¹, PDI = 2.05.

PDPP2Tz50FBDT

The same procedure as for PDPP2Tz10FBDT was used, but now 1 (42.12 mg, 0.041 mmol), 2 (8.45 mg, 0.021 mmol) and 3 (27.14 mg, 0.021 mmol) were used as the monomers. Yield: 53 mg (93%). GPC (*o*-DCB, 140 °C): *M*_n = 66.0 kg mol⁻¹, PDI = 2.72.

PDPP2Tz70FBDT

The same procedure as for PDPP2Tz10FBDT was used, but now 1 (53.60 mg, 0.052 mmol), 2 (6.46 mg, 0.016 mmol) and 3 (48.36 mg, 0.037 mmol) were used as the monomers. Yield: 71 mg (90%). GPC (*o*-DCB, 140 °C): *M*_n = 61.9 kg mol⁻¹, PDI = 2.38.

Results and discussion

Synthesis and characterization

Perfluoroalkyl-DPP based polymers were prepared by Stille polymerization (Scheme 1). The synthesis of two derivatives, PDPP2TzT³⁵ and PDPP2TzFBDT,⁵⁴ has been described previously. The copolymers PDPP2Tz10FBDT, PDPP2Tz30FBDT, PDPP2Tz50FBDT, and PDPP2Tz70FBDT were synthesized by copolymerizing the dibromo-DPP monomer 1 with the bis-stannyl monomers of T (2) and FBDT (3), in which the ratio of the co-monomers 2 and 3 was adjusted to prepare polymers with different contents of perfluoroalkyl units. Stille polymerizations were performed under identical conditions using Pd₂(dba)₃/PPh₃ as the catalyst system and toluene/DMF as the solvent at 115 °C.

The polymers show different ¹H-NMR spectra (ESI, Fig. S1†), but it is difficult to determine the ratio of the polymer segments. Therefore, the *m* : *n* ratio is denoted according to the feed ratio of the co-monomers 2 and 3 (Scheme 1) and does not necessarily represent the actual ratio of T and FBDT units in the polymer backbone. The molecular weight of these polymers has been determined by GPC using *o*-DCB as the eluent. As shown in Table 1 and Fig. S2 (ESI†), most polymers possess a similarly high molecular weight between 60 and 80 kg mol⁻¹. The solubility of PDPP2TzFBDT is very poor in *o*-DCB such that only a soluble small molecular weight fraction can be measured by GPC. The similar molecular weight of these DPP copolymers is also beneficial for investigation of solvent influence on photovoltaic devices.

Optical and electrochemical properties

The optical absorption spectra of the DPP polymers in CHCl₃ solution and in solid state thin films are shown in Fig. 1 and the parameters are summarized in Table 1. All polymers exhibit near-infrared absorption with absorption onsets ranging from 772 to 805 nm in CHCl₃ solution and red-shifted absorptions in thin films with onsets between 787 and 861 nm. PDPP2TzT in which the thiazole-flanked DPP unit is alternating with T as the aromatic units has the lowest *E*_g of 1.44 eV, while PDPP2TzFBDT where the thiazole-flanked DPP unit is alternating with FBDT units shows the largest *E*_g of 1.58 eV among these polymers. When using a ratio of T and FBDT as 9 : 1 in PDPP2Tz10FBDT, the *E*_g slightly increases to 1.46 eV. Interestingly, as the ratio of T and FBDT is further changed to 7 : 3, 5 : 5, and 3 : 7, the polymers



Table 1 Molecular weight and optical and electrochemical properties of the thiazole-based DPP polymers

Polymer	M_n^a (kg mol ⁻¹)	PDI ^b	Solution			Film				
			λ_{peak} (nm)	λ_{onset} (nm)	E_g (eV)	λ_{peak} (nm)	λ_{onset} (nm)	E_g (eV)	E_{LUMO}^c (eV)	E_{HOMO}^d (eV)
PDPP2TzT	74.0	2.80	691, 757	805	1.54	710, 774	861	1.44	-4.03	-5.47
PDPP2Tz10FBDT	77.2	2.05	691, 753	800	1.55	701, 760	848	1.46	-4.39	-5.85
PDPP2Tz30FBDT	65.1	2.05	689, 745	799	1.55	688, 744	823	1.51	-4.27	-5.78
PDPP2Tz50FBDT	66.0	2.72	689, 734	799	1.55	683, 728	831	1.49	-4.30	-5.79
PDPP2Tz70FBDT	61.9	2.38	680, 729	806	1.54	679, 731	822	1.51	-4.24	-5.75
PDPP2TzFBDT ^e	13.1	1.73	672, 712	772	1.61	672, 719	787	1.58	-4.56	-6.14

^a Determined with GPC at 140 °C using *o*-DCB as the eluent. ^b PDI is the polydispersity index. ^c $E_{\text{LUMO}} = -5.23 - E_{\text{red}}$. ^d $E_{\text{HOMO}} = E_{\text{LUMO}} - E_g$. ^e The molecular weight was determined by using a low molecular weight fraction that dissolved in *o*-DCB at 140 °C.

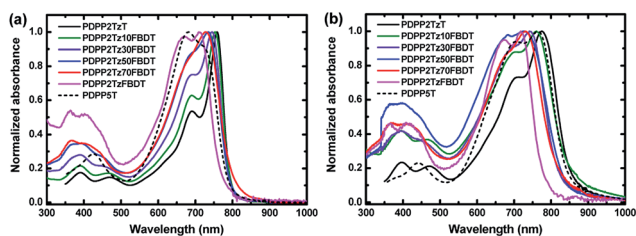


Fig. 1 Optical absorption spectra of the DPP polymers (a) in CHCl₃ solution and (b) in solid state thin films.

PDPP2Tz30FBDT, PDPP2Tz50FBDT, and PDPP2Tz70FBDT have very similar E_g of 1.49 to 1.51 eV in thin films, which are red-shifted compared to that of the copolymer PDPP2TzFBDT. The absorption spectra of the polymers in CHCl₃ solution with a concentration of 0.001 g L⁻¹ are shown in Fig. S3 (ESI[†]). Interestingly, when increasing the content of FBDT units, the intensity of the absorption at 700 nm decreases, indicating a lower absorption coefficient for FBDT-based polymers.

The electrochemical properties of the DPP polymers were determined by CV measurements (ESI[†], Fig. S4 and Table 1). The highest occupied molecular orbital (HOMO) level of PDPP2TzT is at -5.47 eV and the LUMO level is at -4.03 eV. When the T unit is replaced by the FBDT unit, the resulting PDPP2TzFBDT polymer exhibits deeper HOMO and LUMO levels at -6.14 and -4.56 eV, respectively. The copolymers with T and FBDT show HOMO levels around -5.75 and -5.85 eV and LUMO levels around -4.24 and -4.39 eV. The results illustrate that FBDT units can effectively lower the energy levels of electron acceptors. It is also noted that the LUMO levels were not linearly reduced with increasing FBDT content, which is possibly from the different aggregation of the polymers in *o*-DCB and the deviation when exacting reduction potential from CV curves in Fig. S4, ESI[†].

Crystalline properties

In order to investigate the crystalline properties and molecular packing in solid state films, 2D-GIWAXS⁵⁹ was applied on thin films spin coated from CHCl₃ on Si/SiO₂ substrates and the results are shown in Fig. 2 and Table 2. The q values are directly

extracted from the peaks of GIWAXS line cuts. PDPP2TzT shows a distinct (100) diffraction peak in the in-plane direction, which correlates with a lamellar packing distance of 22.4 Å of the 2-octyldecyl side chains. The (010) diffraction peak in the out-of-plane direction for PDPP2TzT is related to the π - π stacking distance of 3.67 Å of the conjugated backbone. These results reveal that PDPP2TzT is a semi-crystalline polymer with a distinct “face on” orientation of the polymer chains, which is beneficial for charge transport in the vertical direction in bulk heterojunction solar cells.⁴⁰ The polymers incorporating both T and FBDT into the main chain also exhibit the (100) and (010)

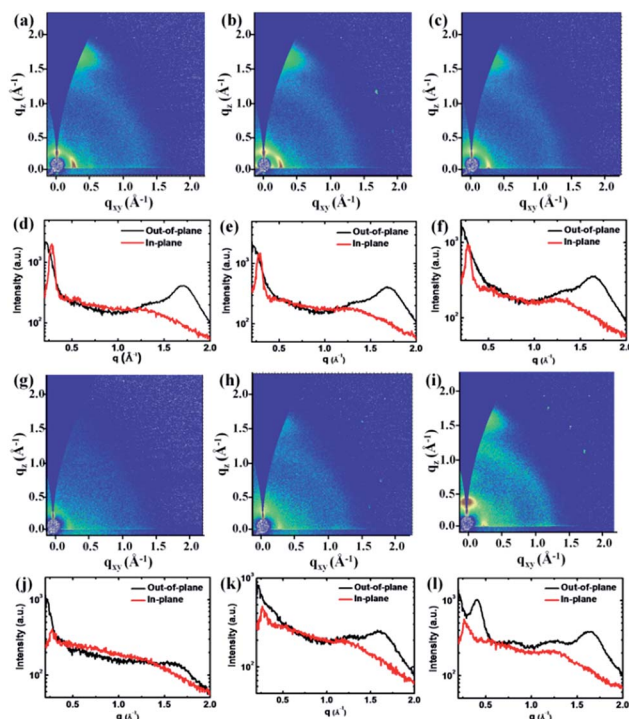


Fig. 2 (a)–(c) and (g)–(i) 2D-GIWAXS patterns of the polymer thin films spin coated from CHCl₃. (d)–(f) and (j)–(l) The out-of-plane and in-plane cuts of the corresponding 2D-GIWAXS patterns. (a) and (d) PDPP2TzT; (b) and (e) PDPP2Tz10FBDT; (c) and (f) PDPP2Tz30FBDT; (g) and (j) PDPP2Tz50FBDT; (h) and (k) PDPP2Tz70FBDT; (i) and (l) PDPP2TzFBDT.



Table 2 Crystallographic parameters of the polymer thin films from 2D-GIWAXS measurements

Polymer	Lamellar spacing		π - π spacing	
	q (\AA^{-1})	d (\AA)	q (\AA^{-1})	d (\AA)
PDPP2TzT	0.28	22.4	1.71	3.67
PDPP2Tz10FBDT	0.28	22.4	1.70	3.70
PDPP2Tz30FBDT	0.28	22.4	1.64	3.83
PDPP2Tz50FBDT	0.27	23.3	1.63	3.85
PDPP2Tz70FBDT	0.27	23.3	1.61	3.90
PDPP2TzFBDT ^a	0.27	23.3	1.64	3.83
PDPP5T	0.33	19.0	1.64	3.83

^a Additional diffraction peak was present: $q_z = 0.41 \text{ \AA}^{-1}$ and $d = 15.3 \text{ \AA}$.

diffraction peaks indicating a “face on” orientation for the 9 : 1 and 7 : 3 ratios, but the intensity of the diffractions is strongly reduced for the 5 : 5 and 3 : 7 co-monomer ratios. This is not unexpected because the introduction of the differently sized T and FBDT units in a random fashion will destroy the translational symmetry along the chain, which precludes obtaining highly ordered polymer domains. For PDPP2TzFBDT the diffraction peaks in the 2D-GIWAXS are restored, demonstrating that it is not the FBDT unit itself that causes the reduced crystallinity of the DPP polymers that were made with both co-monomers. PDPP2TzFBDT exhibits both (100) and (010) diffractions peaks in the out-of-plane direction, while in the in-plane direction the diffraction peaks are much less pronounced. We tentatively attribute this behaviour to the absence of a clear preference for the “face-on” or “edge-on” orientation of the PDPP2TzFBDT polymer chains on the surface. Apparently both types of domains are present, but the orientation of the crystallites is random.

All thiazole-based DPP polymers presented here exhibit a similar d -spacing of about 23.0 \AA for the (100) diffraction peak, originating from the lamellar packing distance induced by the 2-octyldecyl side chains. The π - π stacking distances gradually increased from 3.67 \AA to 3.90 \AA for the polymers PDPP2TzT to PDPP2Tz70FBDT due to the increasing number of sterically demanding 5-perfluorohexylthiophene-2-yl substituents, but then decreased to 3.83 \AA for PDPP2TzFBDT possibly because of a more regular polymer structure (Table 2). A closer π - π stacking distance provides an improved wave function overlap between neighbouring chains which is beneficial for charge transport. It is also interesting to note that PDPP2TzFBDT shows a new diffraction peak at $q_z = 0.41 \text{ \AA}^{-1}$ with a distance of $d = 15.3 \text{ \AA}$, which is possibly induced by the lamellar stacking of perfluoroalkyl units (ESI, Fig. S5[†]).

Finally, as a comparison and because it is used as an electron donor, PDPP5T with five flanking thiophene units shows a (100) diffraction peak in the in-plane direction with $d = 19.04 \text{ \AA}$ and an (010) peak in the out-of-plane direction with $d = 3.83 \text{ \AA}$, indicating a “face on” orientation (ESI, Fig. S6[†]).

In summary, the 2D-GIWAXS data reveal that PDPP2TzT and PDPP2TzFBDT are more crystalline than their T/FBDT mixed copolymers and that the absence of the 5-perfluorohexylthiophene-2-yl substituents in PDPP2TzT enables a closer π - π stacking (3.67

\AA) than in PDPP2TzFBDT (3.83 \AA), which can enhance charge transport for the former.

Polymer-polymer solar cell performance

The thiazole-flanked DPP polymers were applied as electron acceptors using PDPP5T as the electron donor in polymer-polymer photovoltaic devices with an inverted polarity configuration, in which ITO/ZnO and MoO₃/Ag were used as electron and hole extracting contacts, respectively. PDPP5T has a similar absorption spectrum to that of the DPP acceptor polymers (Fig. 1).

When the photoactive layers were spin coated from CHCl₃ without processing additives low PCEs of 0.11–0.39% (ESI, Table S1[†]) were obtained. The low performance is attributed to the large micro-phase separation between the donor and acceptor polymers as inferred from the corresponding AFM images (ESI, Fig. S7[†]). This behaviour is similar to that of PDPP5T : PCBM solar cells when solution processed from CHCl₃ solution.⁵⁷

To improve the blend morphology, the photoactive layers of the PDPP5T : DPP-polymer (1 : 1 w/w) blends were spin coated from CHCl₃ solutions using 1-CN, IC₄F₈I or IC₆F₁₂I as high boiling point solvent additives.

When the active layer was spin coated from CHCl₃ with 3% 1-CN as the additive, PDPP5T : PDPP2TzT cells resulted in a PCE of 2.8% with a J_{sc} of 7.1 mA cm⁻², an open-circuit voltage (V_{oc}) of 0.81 V, and a fill factor (FF) of 0.49. Using the same solvent mixture, the solar cells based on PDPP5T with PDPP2Tz10FBDT, PDPP2Tz30FBDT, PDPP2Tz50FBDT, and PDPP2Tz70FBDT as electron acceptors showed gradually decreasing PCEs from 2.3% down to 1.5%, 0.52%, and 0.26%. The lower PCEs were mainly caused by a reduction of J_{sc} from 5.7 to 0.82 mA cm⁻². PDPP5T : PDPP2TzFBDT cells had a slightly increased PCE of 0.43% when spin coated from CHCl₃ : 1-CN solution, but is still much lower than that of PDPP5T : PDPP2TzT cells (Table 3).

We then tested diiodoperfluoroalkanes as processing additives. For PDPP5T : PDPP2TzFBDT cells we used IC₄F₈I, IC₆F₁₂I and IC₈F₁₆I and found that both IC₄F₈I and IC₆F₁₂I significantly increase the PCE. With IC₈F₁₆I as the additive, the solar cells showed very poor J - V characteristics with large leakage current. This may be due to the high melting point ($\sim 75 \text{ }^\circ\text{C}$) of IC₈F₁₆I, producing holes in the active layers when it was removed *via* a high vacuum process. Above 5%, the volume ratio of IC₆F₁₂I in CHCl₃ has little influence on the device performance of PDPP5T : PDPP2TzFBDT cells (ESI, Table S3[†]). After optimization, the PCE of PDPP5T : PDPP2TzFBDT cells spin coated from CHCl₃ : IC₆F₁₂I was found to be 2.0%, with $J_{sc} = 7.4 \text{ mA cm}^{-2}$, $V_{oc} = 0.67$ and FF = 0.41 at a film thickness of 75 nm (Fig. 3a). The external quantum efficiency (EQE) of the optimized PDPP5T : PDPP2TzFBDT cells shows a broad photoresponse from 300 to 850 nm with a maximum EQE over 0.3 in the near-infrared spectral region where the polymer absorbs light (Fig. 3b). PDPP5T : PDPP2TzFBDT cells fabricated from CHCl₃ with 20% IC₄F₈I also provided PCEs up to 1.8% (ESI, Table S3[†]).

We then optimized the other PDPP5T : DPP-polymer solar cells using CHCl₃ : IC₆F₁₂I (9 : 1) as the solvent mixture (Fig. 3



Table 3 Characteristics of optimized solar cells of the PDPP5T : DPP-polymer. The hole and electron mobilities from SCLC measurements are also included

Acceptor	Solvent	J_{sc}^a (mA cm ⁻²)	V_{oc} (V)	FF	PCE ^a (%)	PCE _{best} (%)	μ_h (cm ² V ⁻¹ s ⁻¹)	μ_e (cm ² V ⁻¹ s ⁻¹)	μ_h/μ_e
PDPP2TzT	CHCl ₃ : 1-CN 3%	7.1 ± 0.3	0.81 ± 0.00	0.49 ± 0.03	2.8 ± 0.04	2.9	2.0 × 10 ⁻⁴	1.3 × 10 ⁻⁴	2
	CHCl ₃ : IC ₆ F ₁₂ I 10% ^b	—	—	—	—	—	—	—	—
PDPP2Tz10FBDT	CHCl ₃ : 1-CN 3%	5.7 ± 0.5	0.79 ± 0.08	0.49 ± 0.02	2.3 ± 0.4	2.4	1.7 × 10 ⁻⁴	6.7 × 10 ⁻⁵	3
	CHCl ₃ : IC ₆ F ₁₂ I 10% ^b	—	—	—	—	—	—	—	—
PDPP2Tz30FBDT	CHCl ₃ : 1-CN 3%	4.1 ± 0.4	0.75 ± 0.01	0.46 ± 0.02	1.5 ± 0.2	1.7	9.7 × 10 ⁻⁵	1.6 × 10 ⁻⁵	6
	CHCl ₃ : IC ₆ F ₁₂ I 10%	1.3 ± 0.1	0.78 ± 0.01	0.37 ± 0.01	0.38 ± 0.04	0.42	7.0 × 10 ⁻⁵	4.0 × 10 ⁻⁵	2
PDPP2Tz50FBDT	CHCl ₃ : 1-CN 3%	1.6 ± 0.2	0.75 ± 0.01	0.43 ± 0.03	0.52 ± 0.08	0.68	1.9 × 10 ⁻⁴	5.6 × 10 ⁻⁶	33
	CHCl ₃ : IC ₆ F ₁₂ I 10%	1.9 ± 0.3	0.75 ± 0.01	0.36 ± 0.09	0.50 ± 0.08	0.56	4.4 × 10 ⁻⁵	1.1 × 10 ⁻⁵	4
PDPP2Tz70FBDT	CHCl ₃ : 1-CN 3%	0.8 ± 0.1	0.78 ± 0.02	0.40 ± 0.04	0.26 ± 0.04	0.32	2.0 × 10 ⁻⁴	1.6 × 10 ⁻⁶	125
	CHCl ₃ : IC ₆ F ₁₂ I 10%	1.8 ± 0.1	0.76 ± 0.05	0.42 ± 0.03	0.48 ± 0.04	0.54	1.5 × 10 ⁻⁴	2.4 × 10 ⁻⁶	65
PDPP2TzFBDT	CHCl ₃ : 1-CN 3%	1.3 ± 0.2	0.68 ± 0.03	0.43 ± 0.03	0.43 ± 0.09	0.52	4.9 × 10 ⁻⁴	1.9 × 10 ⁻⁶	265
	CHCl ₃ : IC ₆ F ₁₂ I 10%	7.4 ± 0.3	0.67 ± 0.06	0.41 ± 0.02	2.03 ± 0.02	2.1	9.2 × 10 ⁻⁵	3.7 × 10 ⁻⁶	25

^a J_{sc} was calculated by integrating the EQE spectrum with the AM1.5G spectrum. ^b The poor wettability of the mixed solution on the surface of ITO/ZnO prevents forming continuous thin films for solar devices. The weight ratio of PDPP5T to the acceptor DPP-polymers is 1 : 1. The thickness of the active layers is around 75 nm.

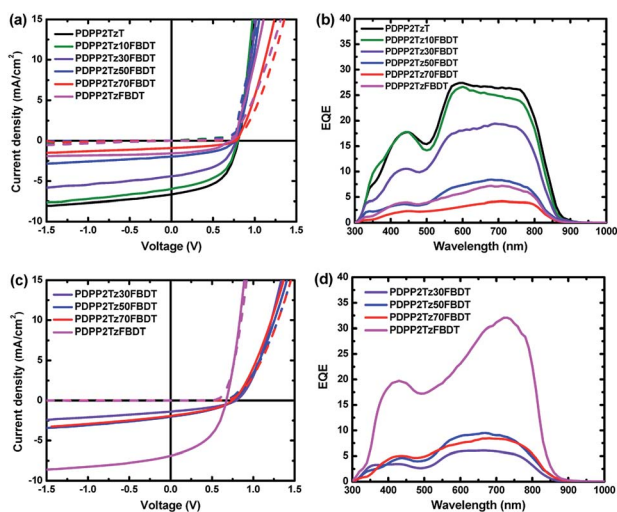


Fig. 3 (a) and (c) J - V characteristics in the dark (dashed lines) and under white light illumination (solid lines). (b) and (d) EQE of the optimized PDPP5T : DPP-polymer solar cells. (a) and (b) The active layers fabricated from CHCl₃ solution with 3% 1-CN. (c) and (d) The active layers fabricated from CHCl₃ solution with 10% IC₆F₁₂I.

and Table 3). We failed to fabricate solar cells for PDPP2TzT and PDPP2Tz10FBDT with PDPP5T using CHCl₃ : IC₆F₁₂I as solvent due to the poor wettability of the solutions on the ITO/ZnO surface, but for the polymers with an increased number of FBDT units, working devices were obtained. The current density–voltage (J - V) characteristics and EQE of the optimized solar cells are shown in Fig. 3 and Table 3. PDPP5T : PDPP2Tz30FBDT cells fabricated from CHCl₃ : IC₆F₁₂I provided a low PCE of 0.38% with $J_{sc} = 1.3$ mA cm⁻² compared to the PCE of 1.5% for the same blend processed from CHCl₃ : 1-CN. For the PDPP2Tz50FBDT and PDPP2Tz70FBDT based cells with a higher content of FBDT units, the PCEs obtained with IC₆F₁₂I as the co-solvent were similar to the PCEs obtained with 1-CN.

The effect of perfluoroalkyl-based additives on the device performance was further investigated by using different donor polymers, PTB7-Th⁶⁰ and PDPP2T-DTP⁶¹ (ESI, Fig. S8†). Solar cells based on PTB7-Th or PDPP2T-DTP as the donor and PDPP2TzFBDT as the acceptor show high performance with PCEs of 2.5% and 1.0% when fabricated from CHCl₃ with IC₆F₁₂I as additives. In comparison, the PCEs of the cells processed from CHCl₃ with 1-CN are 2.1% and 0.6% (ESI, Table S4 and Fig. S9†). The results confirm that IC₆F₁₂I as the additive can enhance the photovoltaic performance based on acceptor polymers bearing perfluoroalkyl side units such as PDPP2TzFBDT. However, all cells show relatively low absolute PCE compared to other high efficiency non-fullerene solar cells, indicating a suboptimal morphology.

Charge transport in the blends

To study the influence of different additives on the charge transport in the blend films, we determined hole and electron mobilities from space charge limited current (SCLC) measurements using a device configuration consisting of ITO/MoO₃/active layer/Au for hole-only devices and ITO/ZnO/active layer/LiF/Al for electron-only devices. PDPP5T : PDPP2TzT layers fabricated from CHCl₃ : 1-CN have well balanced hole and electron mobilities around 10⁻⁴ cm² V⁻¹ s⁻¹ (Table 3). With the same solvent mixture, the hole mobilities of other layers were similar to those of the PDPP2TzT cell, but the electron mobilities gradually decreased to 10⁻⁶ cm² V⁻¹ s⁻¹ with increasing content of FBDT in acceptors. The unbalanced hole and electron mobilities could explain the low PCEs in FBDT-based polymer solar cells from CHCl₃ : 1-CN. When using IC₆F₁₂I as the additive, the electron mobilities were slightly improved compared to those from CHCl₃ : 1-CN, but the hole mobilities were reduced (Table 3). As a consequence, more balanced hole and electron mobilities can be achieved. As an example the μ_h/μ_e ratio of the PDPP5T : PDPP2TzFBDT blend decreases from



265 to 25 when replacing 1-CN by IC₆F₁₂I. The more balanced charge transport is accompanied by an enhancement of the PCE from 0.52% to 2.1%.

Morphology investigation

Because the efficiency of solar cells is intimately related to the morphology, the photoactive layers were further investigated by AFM, 2D-GIWAXS and photoluminescence (PL). When the films were fabricated from CHCl₃ without additives, large domains were observed in AFM images (ESI, Fig. S7†). These disappeared in thin films processed from CHCl₃ with additives (Fig. 4). Fig. 4 shows that for PDPP5T : PDPP2Tz and PDPP5T : PDPP2Tz10FBDT blends processed from CHCl₃ : 1-CN the lateral dimensions and the height of the surface corrugation are smaller (Fig. 4a and b) than for the blends with a higher FBDT content (Fig. 4c–f), suggesting that the latter blends have a coarser micro-morphology with larger domains.

When the same blends were fabricated from CHCl₃ : IC₆F₁₂I the AFM images showed a strongly reduced surface roughness (Fig. 4g–j) compared to those from CHCl₃ : 1-CN (Fig. 4c–f) and smaller lateral dimensions, strongly suggesting an enhanced mixing of the donor and acceptor polymers in the bulk-heterojunction systems. For PDPP5T : PDPP2TzFBDT (Fig. 4j), the surface corrugation is enhanced compared to the other blends.

The blended thin films were further analysed using 2D-GIWAXS measurements. A PDPP5T : PDPP2TzFBDT film spin coated from CHCl₃ : IC₆F₁₂I showed a lower intensity for the in-plane (100) and out-of-plane (010) diffraction peaks (Fig. 5b and

d) as compared to the diffraction of the same blend spin coated from CHCl₃ : 1-CN (Fig. 5a and c). Hence, the higher mixing induced by IC₆F₁₂I reduces the crystallization of the polymers.

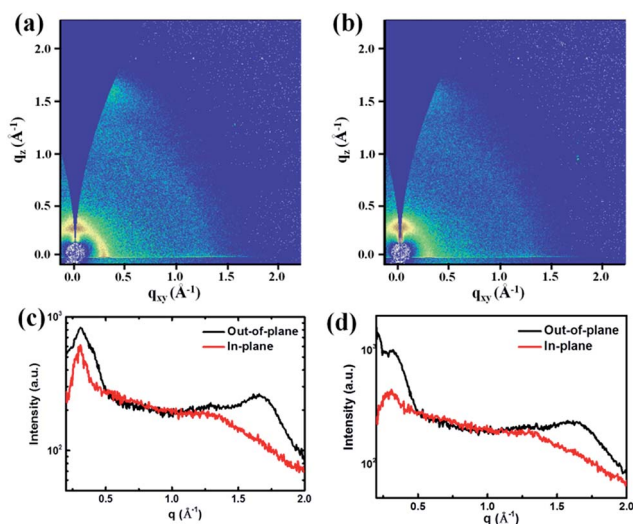


Fig. 5 (a) and (b) 2D-GIWAXS patterns of PDPP5T : PDPP2TzBDT (1 : 1) thin films spin coated from different solutions. (c) and (d) The out-of-plane and in-plane cuts of the corresponding 2D-GIWAXS patterns. (a) and (c) Solution processed from CHCl₃ with 3% 1-CN as the additive. (b) and (d) Solution processed from CHCl₃ with 10% IC₆F₁₂I as the additive.

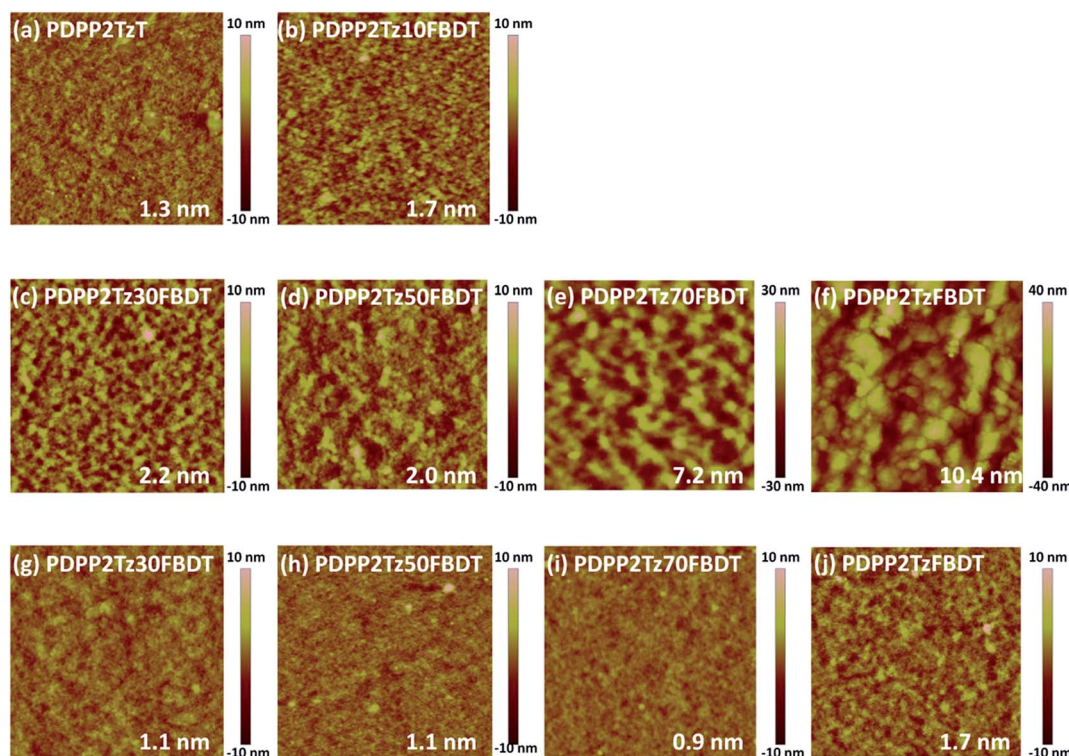


Fig. 4 AFM height images ($3 \times 3 \mu\text{m}^2$) of PDPP5T : DPP-polymer (1 : 1) blends spin coated from (a)–(f) CHCl₃ containing 3 vol% 1-CN and (g)–(j) CHCl₃ containing 10 vol% IC₆F₁₂I. The root mean square (RMS) roughness R_q is also included in the panels.



Steady state PL was applied to further study the morphology difference originating from the spin coating solvent (Fig. 6). The pure polymer and blend thin films were excited at 760 nm and show fluorescence between 800 and 1200 nm. The PL spectra were corrected for the fraction of absorbed photons at the excitation wavelength using the absorption spectra of the same films (ESI, Fig. S10†). Fig. 6 shows that the PL intensity of blend films is significantly reduced compared to that of pure films, indicating the charge transfer from PDPP5T to thiazole-bridged DPP polymers. When processed from CHCl_3 : $\text{IC}_6\text{F}_{12}\text{I}$, the blend films have generally a lower PL intensity compared to the PL of films from CHCl_3 : 1-CN. In particular, the PL intensity of PDPP5T : PDPP2TzFBTD spin coated from CHCl_3 : $\text{IC}_6\text{F}_{12}\text{I}$ was greatly quenched (Fig. 6f). In this case the luminescence of the PDPP5T : PDPP2TzFBTD blend is almost the same as that of pure PDPP5T when the film is spin coated from CHCl_3 : 1-CN. A higher PL quenching indicates a better charge generation which is expected when the mixing of the two components is enhanced.

Role of diiodoperfluoroalkane additives

The use of perfluoroalkyl based processing additives to CHCl_3 resulted in more intimately blended films for perfluoroalkyl-substituted DPP acceptor polymers with donor polymers such as PDPP5T than with 1-CN as the additive. The PCE, however, was only significantly enhanced for PDPP2TzFBTD. For the other co-polymers, with varying ratios of T and FBDT units, the blends become better mixed using $\text{IC}_6\text{F}_{12}\text{I}$ but the PCE is not

really improved. As a result of their irregular structure the mixed T/FBDT co-polymers are less crystalline and possibly the resulting morphology is too well-mixed.

Although diiodoperfluoroalkane additives effectively enhance the PCE of PDPP5T : PDPP2TzFBTD solar cells, the PCE is still lower than that of the PDPP5T : PDPP2Tz blends without the perfluoro substituents on the acceptor polymer. The main reason is the lower V_{oc} and lower FF. The latter is likely related to the lower electron mobility of PDPP5T : PDPP2TzFBTD films compared to PDPP5T : PDPP2Tz films (Table 3). Hence, designing perfluoroalkyl-conjugated polymers with high electron mobilities and optimizing the morphology by looking for new solvents will be the routes to further improve the PCEs of these cells.

Conclusions

A series of conjugated polymers based on thiazole-flanked DPP units were synthesized by incorporating different ratios of T and FBDT units into the main chain to tailor the perfluoroalkyl content of the polymers. The perfluoroalkyl based polymers were applied as electron acceptors in polymer solar cells with PDPP5T as the electron donor. While a CHCl_3 : 1-CN solvent mixture provides a PCE of 2.9% for PDPP5T : PDPP2TzT layers without perfluoroalkyl units, this solvent combination only provides a PCE of 0.52% for PDPP5T : PDPP2TzFBTD with perfluoroalkyl units. We showed that the PCE of PDPP5T : PDPP2TzFBTD can be dramatically enhanced to 2.1% by using a CHCl_3 : $\text{IC}_6\text{F}_{12}\text{I}$ solvent mixture. Detailed analysis using AFM, 2D-GIWAXS and PL measurements reveals that $\text{IC}_6\text{F}_{12}\text{I}$ as the additive is helpful to enhance the mixing of the donor and perfluoroalkyl-based acceptor in the blended films that enhance the charge generation. The results demonstrate that diiodoperfluoroalkane solvents can be efficient processing additives to improve the performance of solution processed perfluoroalkyl based polymer solar cells.

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Notes and references

- 1 A. Facchetti, *Mater. Today*, 2013, **16**, 123–132.
- 2 S. M. McAfee, J. M. Topple, I. G. Hill and G. C. Welch, *J. Mater. Chem. A*, 2015, **3**, 16393–16408.
- 3 Y. Lin and X. Zhan, *Mater. Horiz.*, 2014, **1**, 470–488.

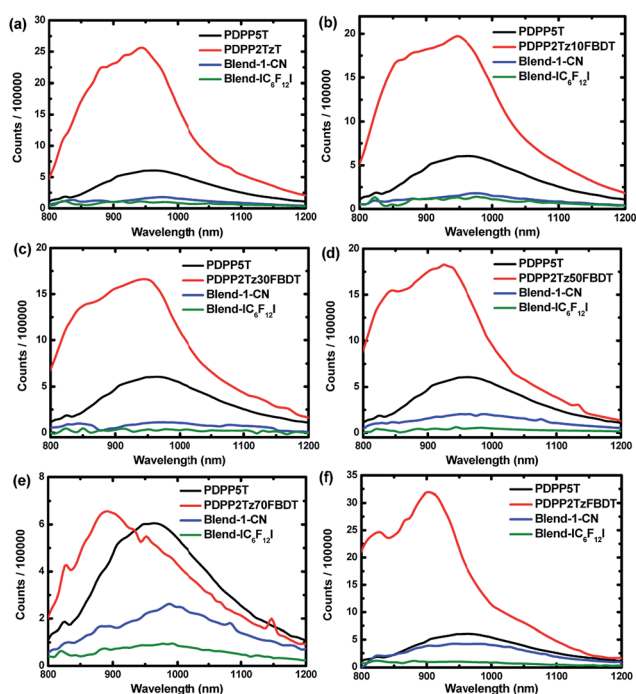


Fig. 6 Photoluminescence spectra of the PDPP5T, thiazole based DPP polymers and blends of PDPP5T : DPP polymers (1 : 1) fabricated from CHCl_3 with 3% 1-CN or 10% $\text{IC}_6\text{F}_{12}\text{I}$. The thin films have a similar thickness around 80–95 nm and were excited at 760 nm for measurements.



- 4 D. Sun, D. Meng, Y. Cai, B. Fan, Y. Li, W. Jiang, L. Huo, Y. Sun and Z. Wang, *J. Am. Chem. Soc.*, 2015, **137**, 11156–11162.
- 5 Y.-J. Hwang, H. Li, B. A. E. Courtright, S. Subramanian and S. A. Jenekhe, *Adv. Mater.*, 2016, **28**, 124–131.
- 6 Y. Zhong, M. T. Trinh, R. Chen, G. E. Purdum, P. P. Khlyabich, M. Sezen, S. Oh, H. Zhu, B. Fowler, B. Zhang, W. Wang, C.-Y. Nam, M. Y. Sfeir, C. T. Black, M. L. Steigerwald, Y.-L. Loo, F. Ng, X. Y. Zhu and C. Nuckolls, *Nat. Commun.*, 2015, **6**, 8242/1–8.
- 7 R. Shivanna, S. Shoaee, S. Dimitrov, S. K. Kandappa, S. Rajaram, J. R. Durrant and K. S. Narayan, *Energy Environ. Sci.*, 2014, **7**, 435–441.
- 8 W. Jiang, L. Ye, X. Li, C. Xiao, F. Tan, W. Zhao, J. Hou and Z. Wang, *Chem. Commun.*, 2014, **50**, 1024–1026.
- 9 Y. Zang, C.-Z. Li, C.-C. Chueh, S. T. Williams, W. Jiang, Z.-H. Wang, J.-S. Yu and A. K. Y. Jen, *Adv. Mater.*, 2014, **26**, 5708–5714.
- 10 A. Sharenko, C. M. Proctor, T. S. van der Poll, Z. B. Henson, T.-Q. Nguyen and G. C. Bazan, *Adv. Mater.*, 2013, **25**, 4403–4406.
- 11 P. E. Hartnett, A. Timalina, H. S. S. R. Matte, N. Zhou, X. Guo, W. Zhao, A. Facchetti, R. P. H. Chang, M. C. Hersam, M. R. Wasielewski and T. J. Marks, *J. Am. Chem. Soc.*, 2014, **136**, 16345–16356.
- 12 O. K. Kwon, J.-H. Park, D. W. Kim, S. K. Park and S. Y. Park, *Adv. Mater.*, 2015, **27**, 1951–1956.
- 13 L. Y. Liu, J. Y. L. Lai, S. Chen, Y. Li, K. Jiang, J. Zhao, Z. Li, H. Hu, T. Ma, H. Lin, J. Liu, J. Zhang, F. Huang, D. Yu and H. Yan, *J. Mater. Chem. A*, 2015, **3**, 13632–13636.
- 14 B. Jiang, X. Zhang, C. Zhan, Z. Lu, J. Huang, X. Ding, S. He and J. Yao, *Polym. Chem.*, 2013, **4**, 4631–4638.
- 15 Y. Liu, L. Zhang, H. Lee, H.-W. Wang, A. Santala, F. Liu, Y. Diao, A. L. Briseno and T. P. Russell, *Adv. Energy Mater.*, 2015, **5**, 1500195.
- 16 Y. Lin, J. Wang, Z.-G. Zhang, H. Bai, Y. Li, D. Zhu and X. Zhan, *Adv. Mater.*, 2015, **27**, 1170–1174.
- 17 X. Zhang, J. Zhang, H. Lu, J. Wu, G. Li, C. Li, S. Li and Z. Bo, *J. Mater. Chem. C*, 2015, **3**, 6979–6985.
- 18 X.-F. Wu, W.-F. Fu, Z. Xu, M. Shi, F. Liu, H.-Z. Chen, J.-H. Wan and T. P. Russell, *Adv. Funct. Mater.*, 2015, **25**, 5954–5966.
- 19 J. J. W. Jung and W. H. Jo, *Chem. Mater.*, 2015, **27**, 6038–6043.
- 20 O. K. Kwon, J.-H. Park, S. K. Park and S. Y. Park, *Adv. Energy Mater.*, 2015, **5**, 1400929.
- 21 J. J. M. Halls, C. A. Walsh, N. C. Greenham, E. A. Marseglia, R. H. Friend, S. C. Moratti and A. B. Holmes, *Nature*, 1995, **376**, 498–500.
- 22 E. Zhou, J. Cong, Q. Wei, K. Tajima, C. Yang and K. Hashimoto, *Angew. Chem., Int. Ed.*, 2011, **50**, 2799–2803.
- 23 D. Mori, H. Benten, H. Ohkita, S. Ito and K. Miyake, *ACS Appl. Mater. Interfaces*, 2012, **4**, 3325–3329.
- 24 R. Stalder, J. Mei, J. Subbiah, C. Grand, L. A. Estrada, F. So and J. R. Reynolds, *Macromolecules*, 2011, **44**, 6303–6310.
- 25 Y.-J. Hwang, B. A. E. Courtright, A. S. Ferreira, S. H. Tolbert and S. A. Jenekhe, *Adv. Mater.*, 2015, **27**, 4578–4584.
- 26 J. W. Jung, J. W. Jo, C.-C. Chueh, F. Liu, W. H. Jo, T. P. Russell and A. K. Y. Jen, *Adv. Mater.*, 2015, **27**, 3310–3317.
- 27 Y.-J. Hwang, T. Earmme, B. A. E. Courtright, F. N. Eberle and S. A. Jenekhe, *J. Am. Chem. Soc.*, 2015, **137**, 4424–4434.
- 28 C. Lee, H. Kang, W. Lee, T. Kim, K.-H. Kim, H. Y. Woo, C. Wang and B. J. Kim, *Adv. Mater.*, 2015, **27**, 2466–2471.
- 29 H. Li, Y.-J. Hwang, T. Earmme, R. C. Huber, B. A. E. Courtright, C. O'Brien, S. H. Tolbert and S. A. Jenekhe, *Macromolecules*, 2015, **48**, 1759–1766.
- 30 I. H. Jung, W.-Y. Lo, J. Jang, W. Chen, D. Zhao, E. S. Landry, L. Lu, D. V. Talapin and L. Yu, *Chem. Mater.*, 2014, **26**, 3450–3459.
- 31 S. Subramanian, T. Earmme, N. M. Murari and S. A. Jenekhe, *Polym. Chem.*, 2014, **5**, 5707–5715.
- 32 D. Mori, H. Benten, I. Okada, H. Ohkita and S. Ito, *Energy Environ. Sci.*, 2014, **7**, 2939–2943.
- 33 Y. Zhou, T. Kurosawa, W. Ma, Y. Guo, L. Fang, K. Vandewal, Y. Diao, C. Wang, Q. Yan, J. Reinspach, J. Mei, A. L. Appleton, G. I. Koleilat, Y. Gao, S. C. B. Mannsfeld, A. Salleo, H. Ade, D. Zhao and Z. Bao, *Adv. Mater.*, 2014, **26**, 3767–3772.
- 34 W. Li, W. S. C. Roelofs, M. Turbiez, M. M. Wienk and R. A. J. Janssen, *Adv. Mater.*, 2014, **26**, 3304–3309.
- 35 W. Li, Y. An, M. M. Wienk and R. A. J. Janssen, *J. Mater. Chem. A*, 2015, **3**, 6756–6760.
- 36 X. Zhan, Z. a. Tan, B. Domercq, Z. An, X. Zhang, S. Barlow, Y. Li, D. Zhu, B. Kippelen and S. R. Marder, *J. Am. Chem. Soc.*, 2007, **129**, 7246–7247.
- 37 X. Guo, A. Facchetti and T. J. Marks, *Chem. Rev.*, 2014, **114**, 8943–9021.
- 38 H. Yan, Z. Chen, Y. Zheng, C. Newman, J. R. Quinn, F. Dotz, M. Kastler and A. Facchetti, *Nature*, 2009, **457**, 679–686.
- 39 L. Gao, Z.-G. Zhang, L. Xue, J. Min, J. Zhang, Z. Wei and Y. Li, *Adv. Mater.*, 2015, **28**, 1884–1890.
- 40 Y. Liu, J. Zhao, Z. Li, C. Mu, W. Ma, H. Hu, K. Jiang, H. Lin, H. Ade and H. Yan, *Nat. Commun.*, 2014, **5**, 5293/1–8.
- 41 S. Zhang, L. Ye, W. Zhao, B. Yang, Q. Wang and J. Hou, *Sci. China: Chem.*, 2015, 248–256.
- 42 Y. J. Cheng, S. H. Yang and C. S. Hsu, *Chem. Rev.*, 2009, **109**, 5868–5923.
- 43 L. Dou, Y. Liu, Z. Hong, G. Li and Y. Yang, *Chem. Rev.*, 2015, **115**, 12633–12665.
- 44 S. Y. Qu and H. Tian, *Chem. Commun.*, 2012, **48**, 3039–3051.
- 45 K. H. Hendriks, W. Li, M. M. Wienk and R. A. J. Janssen, *J. Am. Chem. Soc.*, 2014, **136**, 12130–12136.
- 46 Z. Yi, S. Wang and Y. Liu, *Adv. Mater.*, 2015, **27**, 3589–3606.
- 47 C. Xiao, G. Zhao, A. Zhang, W. Jiang, R. A. J. Janssen, W. Li, W. Hu and Z. Wang, *Adv. Mater.*, 2015, **27**, 4963–4968.
- 48 H. A. Um, D. H. Lee, D. U. Heo, D. S. Yang, J. Shin, H. Baik, M. J. Cho and D. H. Choi, *ACS Nano*, 2015, **9**, 5264–5274.
- 49 H. Choi, S.-J. Ko, T. Kim, P.-O. Morin, B. Walker, B. H. Lee, M. Leclerc, J. Y. Kim and A. J. Heeger, *Adv. Mater.*, 2015, **27**, 3318–3324.
- 50 R. S. Ashraf, I. Meager, M. Nikolka, M. Kirkus, M. Planells, B. C. Schroeder, S. Holliday, M. Hurhangee, C. B. Nielsen, H. Sirringhaus and I. McCulloch, *J. Am. Chem. Soc.*, 2014, **137**, 1314–1321.
- 51 K. H. Hendriks, G. H. L. Heintges, V. S. Gevaerts, M. M. Wienk and R. A. J. Janssen, *Angew. Chem., Int. Ed.*, 2013, **52**, 8341–8344.



- 52 B. Sun, W. Hong, Z. Yan, H. Aziz and Y. Li, *Adv. Mater.*, 2014, **26**, 2636–2642.
- 53 T. B. Singh, N. Marjanović, P. Stadler, M. Auinger, G. J. Matt, S. Günes, N. S. Sariciftci, R. Schwödiauer and S. Bauer, *J. Appl. Phys.*, 2005, **97**, 083714.
- 54 A. Zhang, C. Xiao, D. Meng, Q. Wang, X. Zhang, W. Hu, X. Zhan, Z. Wang, R. A. J. Janssen and W. Li, *J. Mater. Chem. C*, 2015, **3**, 8255–8261.
- 55 P. Homyak, Y. Liu, S. Ferdous, F. Liu, T. P. Russell and E. B. Coughlin, *Chem. Mater.*, 2015, **27**, 443–449.
- 56 G. He, X. Wan, Z. Li, Q. Zhang, G. Long, Y. Liu, Y. Hou, M. Zhang and Y. Chen, *J. Mater. Chem. C*, 2014, **2**, 1337–1345.
- 57 S. Kouijzer, J. J. Michels, M. van den Berg, V. S. Gevaerts, M. Turbiez, M. M. Wienk and R. A. J. Janssen, *J. Am. Chem. Soc.*, 2013, **135**, 12057–12067.
- 58 W. Li, W. S. C. Roelofs, M. M. Wienk and R. A. J. Janssen, *J. Am. Chem. Soc.*, 2012, **134**, 13787–13795.
- 59 P. M. Buschbaum, *Adv. Mater.*, 2014, **26**, 7692–7709.
- 60 J.-D. Chen, C. Cui, Y.-Q. Li, L. Zhou, Q.-D. Ou, C. Li, Y. Li and J.-X. Tang, *Adv. Mater.*, 2015, **27**, 1035–1041.
- 61 W. Li, K. H. Hendriks, A. Furlan, A. Zhang, M. M. Wienk and R. A. J. Janssen, *Chem. Commun.*, 2015, **51**, 4290–4293.

