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Palladium-catalyzed regioselective and stereo-invertive ring-opening borylation of 2-arylaziridines with bis(pinacolato)diboron: experimental and computational studies†

Youhei Takeda,^{*a} Akinobu Kuroda,^a W. M. C. Sameera,^b Keiji Morokuma^{*b} and Satoshi Minakata^{*a}

A palladium catalyzed regioselective borylative ring opening reaction of 2-arylaziridines to give β -amino- β -arylethylborates was developed. The reaction reported herein represents the first example of ring-opening borylation of non-vinyl aziridines and direct borylative C(sp³)-N bond cleavage of neutral organic substrates. NMR studies and density functional theory (DFT) calculations suggested that the active intermediate for the reaction is a PdL₂ complex [L = P(*t*-Bu)₂Me]. The multi-component artificial force-induced reaction method (MC-AFIR) located the transition states for the regioselectivity-determining aziridine ring opening that proceeds in an S_N2 fashion, and explained the selectivity of the reaction. The full catalytic cycle consists of a selectivity-determining aziridine ring opening (oxidative addition), a proton transfer, phosphine ligand dissociation from the catalyst, boron-boron bond cleavage, and reductive elimination. Water is important to drive the transmetalation step. The calculated overall mechanism and selectivity are consistent with the experimental results.

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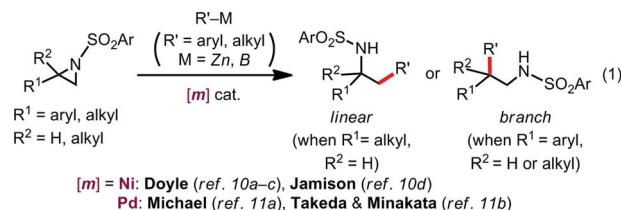
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Introduction

Aziridines, three-membered strained azaheterocycles, serve as versatile building blocks in modern organic synthesis.¹ Regioselective ring opening of aziridines with nucleophiles, which is driven by release of its ring strain, is one of the most useful transformations of aziridines into ubiquitous β -amino-functionalized motifs.² Aziridines undergo oxidative addition to low-valent late transition metal complexes at the terminal C-N bond in an S_N2 fashion, generating the corresponding oxidative adducts, azametallacyclobutanes.³⁻⁵ In light of this reactivity, diverse transition metal-catalyzed regio-, stereo-, and/or chemoselective transformations of aziridines are feasible,⁶⁻⁹ where the oxidative addition is smoothly coupled with subsequent elementary processes, such as transmetalation, migratory insertion, and reductive elimination; regioselective Rh- and Co-catalyzed carbonylative ring expansion of aziridines, pioneered by Alper and co-workers, represent the embodiment of this idea.⁶

Over the past few years, numerous advances have been made in transition metal-catalyzed ring-opening cross coupling of aziridines, where aziridines can be utilized as a non-classical alkyl electrophilic partner [eqn (1)].^{10,11} In 2012, Doyle reported the first Ni-catalyzed regioselective Negishi alkylation of 2-aryl-*N*-tosylaziridines that leads to branch-type products *via* the regioselective cleavage of a benzylic C-N bond.^{10a} Since this report, several groups including us^{11b} have developed Ni-¹⁰ and Pd-catalyzed¹¹ regioselective ring-opening cross coupling reactions of aziridines to form C(sp³)-C(sp²)/C(sp³) bonds, with the regioselectivity seemingly governed by the aziridine substrate rather than the catalyst [eqn (1)]. Despite the above-mentioned successes in C-C cross couplings, other catalytic ring-opening C-E (E \neq C) bond forming reactions of aziridines based on a cross-coupling mechanism are also unexploited.¹²



^aDepartment of Applied Chemistry, Graduate School of Engineering, Osaka University, Yamadaoka 2-1, Suita, Osaka 565-0871, Japan. E-mail: takeda@chem.eng.osaka-u.ac.jp; minakata@chem.eng.osaka-u.ac.jp

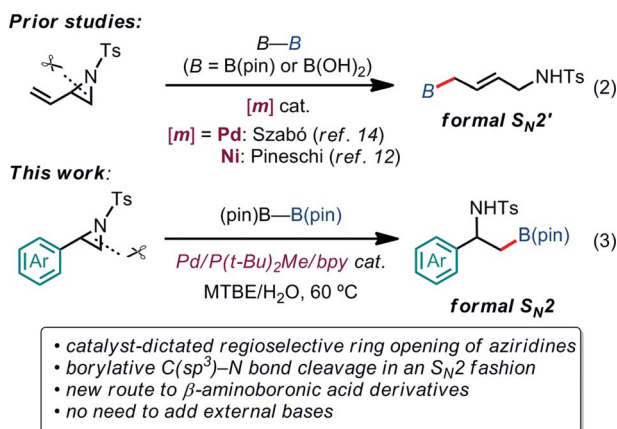
^bFukui Institute for Fundamental Chemistry, Kyoto University, Takano-Nishihiraki-cho 34-4, Sakyo-ku, Kyoto 606-8103, Japan. E-mail: morokuma.keiji.3a@kyoto-u.ac.jp

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Our latest findings^{11b} led us to seek new catalytic ring-opening C-E coupling systems of aziridines. Moreover, we became intrigued with the C-B coupling of aziridines to give β -amino-functionalized alkylboronates; alkylboronic acid

derivatives serve not only as useful building blocks in organic synthesis, but are also biologically interesting motifs in medicinal chemistry.¹³ Only a few precedents in transition-metal-catalyzed borylative substitution of vinylaziridines *via* C(sp³)-N bond cleavage have been reported. Szabó and Pineschi disclosed a Pd(II) pincer complex¹⁴ and a Ni(0)/BINAP¹² system as catalysts for the borylative ring opening of vinylaziridines with diboron reagents *via* the cleavage of an allylic C-N bond to give γ -amino alkylboronic acid derivatives (formal S_N2'-type reaction),¹⁵ respectively [eqn (2)]. Nevertheless, to the best of our knowledge, the catalytic direct displacement of the C(sp³)-N bond of aziridines with a C(sp³)-B bond on the same carbon (formal S_N2-type reaction) has never been described to date.

We herein report a regioselective ring-opening C-B cross-coupling reaction of 2-arylaziridines that is realized using a Pd/P(*t*-Bu)₂Me/bpy catalytic system to give β -amino-alkylboronates [eqn (3)], which can serve as versatile building blocks to synthesize amino-functionalized compounds and biologically relevant β -amino acid surrogates.^{16,17} Notably, our reaction features the opposite regioselectivity in C-N bond cleavage (at the 3-position) to those previously reported for C-C cross couplings using the same type of aziridine substrate (at the 2-position).^{10a,c,11b} Also, our C-B coupling system represents the first example of a formal S_N2-type borylative C(sp³)-N bond cleavage of neutral compounds.¹⁸ It should be noted that the present C-B coupling proceeds smoothly under neutral conditions as in the cases of borylative substitution of allyl carboxylates¹⁹ and carbonates.^{19a,20} In these cases, the leaving group-derived oxy anions (RCO₂⁻ and RO⁻) serve as the internal bases, while the metal catalyzed borylative substitution of alkyl halides usually requires the addition of stoichiometric external base or activator to promote transmetalation.²¹ In addition to experimental elaboration, we have performed theoretical calculations, applying density functional theory (DFT) and the multi-component artificial force-induced reaction (MC-AFIR) method to determine the mechanism of the reaction and explain the origin of the selectivity, and to clarify why external base is not required in this reaction system.



Results and discussion

Experimental part

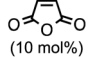
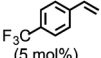
Development of Pd-catalyzed borylative ring-opening reaction of 2-arylaziridines. To identify the reaction conditions for the borylative ring-opening reaction, optimization studies applying racemic 2-phenyl-*N*-tosyl-aziridines (**1a**) as a model substrate with bis(pinacolato)diboron B₂(pin)₂ (**2**) were performed (Tables S1–S13 in the ESI†).²² The executive summary of the optimization study is shown in Table 1.

The highest yield (81%, isolated yield: 71%) of borylated product **3a** was obtained when **1a** was treated with 1.2 equiv. of B₂(pin)₂ at 60 °C in the presence of catalytic amounts of Cp(allyl)Pd (2 mol%), P(*t*-Bu)₂Me²³ (1 mol%), and 2,2'-bipyridine (bpy, 20 mol%) in a mixed solvent of methyl-*tert*-butylether (MTBE) and H₂O (entry 1, "standard conditions"). Most importantly, the regioselectivity of this ring opening and C-B coupling was opposite to those observed with the C-C coupling of the same substrate,^{10a,c,11b} implying that the oxidative addition occurred at the C-N bond on the terminal carbon (the 3-position of the aziridine). The regioisomer of **3a** was not detected in the crude ¹H NMR spectra, indicating that the regioselectivity in the ring-opening of **1a** should be almost perfect. In fact, the regioselective oxidative addition dictated by the interactions between the substrate and the Pd(0) catalyst was supported by theoretical calculations (*vide infra*). The L : Pd ratio [L = P(*t*-Bu)₂Me] was found to have a significant impact on the product distribution (entries 2 and 3). As the L : Pd ratio was increased from 0.5 (entry 1) to 2 to 3, the yields of **3a** decreased to 74% and 45%, respectively (entries 2 and 3). This irregular ratio of L : Pd implies that there might be a complex equilibrium of PdL_{*n*} species generated *in situ*, some of which are specifically active in the catalytic cycle. In fact, this speculation was partly supported with ³¹P NMR experiments and theoretical calculations (*vide infra*). Addition of electron deficient unsaturated compounds, which can coordinate to metal complexes, is an effective strategy to suppress undesired β -hydride elimination or to accelerate reductive elimination, thereby often leading to better results in alkyl cross coupling reactions.²⁴ Indeed, the effect of adding bpy was significantly effective for suppressing the formation of byproduct **6** as well as **5**, which was presumably derived from the hydrolysis of **6** with H₂O (entries 4–7). The effect of the ligands was also significant (entries 8–10). *N*-heterocyclic carbene (NHC) ligands, SIPr for instance,^{11b} promoted the consumption of **1a** and produced undesired products **4** and **5** (entry 8). The results obtained from sterically demanding trialkylphosphine/Pd catalysts (entries 9 and 10) suggest a sluggish oxidative addition step, which is in good agreement with Wolfe's report.⁹ Furthermore, as mentioned in the introduction section, the ring opening borylation disclosed herein does not require the addition of any external base, which is usually used to activate the transmetalation step for the borylative substitution of alkyl halides. In connection with this, addition of water is crucial. Moreover, in the absence of H₂O, no conversion of **1a** was observed (entry 11), while at least 1 equiv. of H₂O allowed the borylation (Table S11 in the ESI†).



Table 1 Evaluation of reaction parameters

Reaction scheme: (±)-1a reacts with $B_2(\text{pin})_2$ (2, 1.2 equiv), $\text{Cp}(\text{allyl})\text{Pd}$ (2 mol%), $\text{P}(t\text{-Bu})_2\text{Me}$ (1 mol%), and bpy (20 mol%) in $\text{MTBE}/\text{H}_2\text{O}$ ($v/v = 12:1$) at 60°C for 3 h to yield products 3a, 4, 5, and 6 under "standard conditions".

Entry	Difference from the "standard conditions"	Yield ^a (%)				Recovery of 1a ^a (%)
		3a	4	5	6	
1	None	81	0	0	13	0
2	$\text{Cp}(\text{allyl})\text{Pd}$ (1 mol%)/ $\text{P}(t\text{-Bu})_2\text{Me}$ (2 mol%)	74	0	11	12	3
3	$\text{Cp}(\text{allyl})\text{Pd}$ (1 mol%)/ $\text{P}(t\text{-Bu})_2\text{Me}$ (3 mol%)	45	0	10	14	0
4 ^b	Without bpy	52	0	14	25	0
5 ^b	$\text{MeO}_2\text{C}-\text{CH}=\text{CH}-\text{CO}_2\text{Me}$ (10 mol%) instead of bpy	0	0	10	0	90
6 ^b	 (10 mol%) instead of bpy	Trace	0	3	0	96
7 ^b	 (5 mol%) instead of bpy	64	0	8	16	0
8	$\text{Pd}_2(\text{dba})_3/\text{SIPr}$ was used as catalyst	0	5	37	0	48
9	$\text{Pd}_2(\text{dba})_3/\text{PCy}_3$ was used as catalyst	5	10	0	0	85
10	$\text{Pd}[\text{P}(t\text{-Bu})_3]_2$ was used as catalyst	0	0	0	0	97
11 ^c	Without H_2O	0	0	0	7	84

^a Determined with GC or ^1H NMR. ^b 1 mol% of $\text{Cp}(\text{allyl})\text{Pd}$ and 2 mol% of ligand were used. ^c 1 mol% of $\text{Cp}(\text{allyl})\text{Pd}$ and 0.5 mol% of ligand were used.

Theoretical calculations implied that H_2O serves as a proton source (H^+) as well as a source of internal base ($\text{Pd}-\text{OH}$) to promote the four-membered transmetalation between $B_2(\text{pin})_2$ and Pd (*vide infra*). To investigate the fate of the other $B(\text{pin})$ moiety of $B_2(\text{pin})_2$ in the reaction, the ^{11}B NMR spectrum of the crude product was acquired in benzene- d_6 . The only peak other than the remaining $B_2(\text{pin})_2$ was detected at $\delta = 22.6$ ppm as a slightly broad singlet. The comparison with the reported ^{11}B resonances of $(\text{pin})\text{B}-\text{OH}$ ($\delta = 22.5$ ppm in benzene- d_6)^{25a} and $(\text{pin})\text{B}-\text{O}-\text{B}(\text{pin})$ ($\delta = 21.6$ ppm in benzene- d_6)^{25b} would suggest that one of the $B(\text{pin})$ moieties of the starting diboron reagent was converted to $(\text{pin})\text{B}-\text{OH}$ during the reaction.

Substrate scope and synthetic utility. After establishing the optimized conditions, we explored the substrate scope of the borylative substitution reaction (Table 2). In terms of 2-arylaziridines, *p*-, *m*-, and *o*-tolyl-substituted aziridines **2b**, **2c**, and **2d** smoothly underwent borylative ring opening in a regioselective manner to give the corresponding β -amide- β -tolethylboronates **3b**, **3c**, and **3d** in good yields. Aziridine bearing a *p*-fluorophenyl substituent at the 2-position was efficiently borylated, giving the corresponding alkylboronate **3e** in a high yield. It is important to note that the *p*-Cl functionality on the aryl ring of aziridine **2f** survived the reaction conditions to provide the borylative product **3f** in a high yield. Nevertheless, the reaction with *p*-bromophenyl-substituted aziridine did not proceed at all. An invaluable functional group in medicinal chemistry, a CF_3 -installed aziridine **2g** gave the corresponding

CF_3 -bearing alkylboronate **3g**. Notably, this ring opening borylation was applicable to the arylaziridines bearing a strong electron withdrawing group (NO_2 , OAc , and CO_2Me), and kept these functionalities intact. This highlights an advantage of using neutral conditions.

Furthermore, the borylation was scalable, and **3a** was prepared on a gram scale (2.89 g, 72% yield) from 10 mmol of **1a** (pS17 of the ESI†). With respect to the limitation of substrate, 2-pyridyl- (**1k**) and 2-alkylaziridines (**1l**, **1m**, and **1n**) were not consumed under the reaction conditions. Furthermore, this reaction is very sensitive to the steric hindrance around the 3-positioned carbon, where oxidative addition to $\text{Pd}(0)$ has occurred. For example, 2-methyl-2-phenylaziridine (**1o**) and *trans*-2-phenyl-3-methylaziridine (**1p**) were not consumed at all under the standard conditions. As for the scope of the *N*-functional groups of aziridines, *p*- $\text{MeOC}_6\text{H}_4\text{SO}_2$ and *t*- BuSO_2 -substituted 2-phenylaziridines smoothly underwent borylation in a regioselective manner (at the 3-position) to produce the corresponding β -amino- β -phenylethylboronates **3r** (65%) and **3s** (52%), respectively. On the other hand, *N*-*p*-nosyl-2-phenylaziridine did not give borylated product (Table S12 in the ESI†). The use of an *N*-acylated aziridine, *N*-Boc-2-phenylaziridine for instance, did not give borylated products at a noticeable level, highlighting the efficacy of the *N*-tosyl group for the borylation. Regarding the diboron reagents, the use of bis(neopentyl glycolato)diboron gave the corresponding aminoboronate product (**3a'**) in 39% yield, while the borylation using bis(hexylene



Table 2 Scope and limitation of the borylation^{a,b}

Reaction scheme showing the borylation of aziridine **1** with boronic ester **2** to form product **3**.

Reagents and conditions:

- Cp(allyl)Pd (2 mol%)
- P(*t*-Bu)₂Me (1 mol%)
- bpy (20 mol%)
- MTBE/H₂O (v/v 12:1)
- 60 °C, 3 h

Structure **1**: Aziridine with substituent **I**Ar.

Structure **2**: Boronic ester, (pin)B-B(pin), (1.2 equiv).

Structure **3**: Product, 1-(N-tosylamino)-2-(B(pin))-1-**I**Ar-ethane.

Yields for various substituted products **3b** through **3j**:

- 3b** (Me): 83% (71%)
- 3c** (Me): 77% (71%)
- 3d** (Me): 86% (81%)
- 3e** (F): 88% (75%)
- 3f** (Cl): 90% (81%)
- 3g** (F₃C): 85% (74%)
- 3h^c** (O₂N): 87% (81%)
- 3i** (AcO): 71% (58%)
- 3j** (MeO₂C): 83% (77%)

No Reaction with

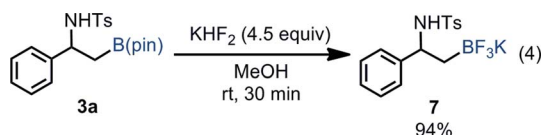
Structures of aziridines that do not react:

- 1k**: Aziridine with a pyridine ring.
- 1l**: Aziridine with substituent **R**, where **R** = *n*-Bu (**1l**), Bn (**1m**).
- 1n**: Aziridine with an allyl group.
- 1o**: Aziridine with a 1-methyl-2-phenylethyl group.
- 1p**: Aziridine with a 1-methyl-2-phenylethyl group.

^a Reaction conditions: **1** (0.50 mmol), **2** (0.60 mmol), Cp(allyl)Pd (10 μmol), P(*t*-Bu)₂Me (5 μmol), bpy (10 μmol) in MTBE/H₂O (1.5 mL, v/v 12 : 1) at 60 °C under N₂ atmosphere for 3 h. ^b The values outside and inside parentheses indicate ¹H NMR and isolated yields, respectively. ^c The reaction was conducted at 80 °C.

glycolato)diboron produced the corresponding borylated products in a low yield as an inseparable diastereomeric mixture, presumably due to the existence of two chiral centers on the benzylic and glycolate carbons (Table S13 in the ESI†).

Borylated product **3a** was successfully transformed into β-amino-β-aryl-substituted alkyltrifluoroborate **7** in excellent yield, which can serve as a versatile building block to synthesize β-amino-β-arylethanes through Pd-catalyzed cross coupling reactions [eqn (4)].²⁶

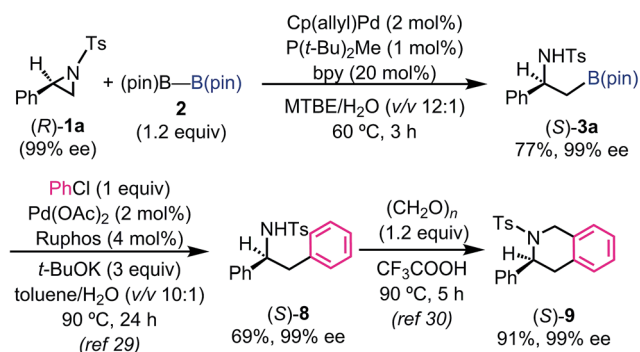


In order to understand the stereochemical outcome of the secondary stereogenic center of the borylated product and to demonstrate the utility of the alkylboronate, an enantiopure aziridine (*R*)-**1a** (99% ee from chiral HPLC analysis) was subjected to the standard reaction conditions (Scheme 1). HPLC analysis of the borylated product **3a** revealed that the stereochemical information was completely retained (99% ee, *S*)

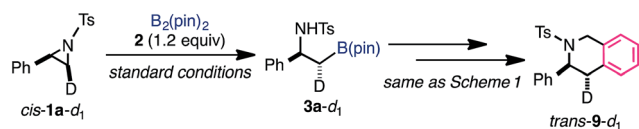
through the reaction. This result excludes the possibility of the reaction pathway of oxidative addition/β-hydride elimination/re-insertion of [Pd-H] species leading to **3a**, although it cannot perfectly exclude such a pathway.²⁷ To demonstrate the synthetic utility of the enantiopure β-amino-alkylboronate (*S*)-**3a**, the asymmetric synthesis of 3-phenyl-1,2,3,4-tetrahydroisoquinoline derivative was conducted (Scheme 1). The tetrahydroisoquinoline skeleton is a ubiquitous motif in alkaloid natural products and constitutes a biologically important aza-heterocycle.²⁸ Enantiopure alkylboronate (*S*)-**3a** was subjected to the Pd-catalyzed Suzuki-Miyaura cross coupling²⁹ with chlorobenzene to give the optically active 1,2-diphenylamine derivative (*S*)-**8** in 69% yield, while keeping the enantiopurity intact (99% ee, chiral HPLC). The following Pictet-Spengler reaction³⁰ successfully gave the enantiopure tetrahydroisoquinoline product (*S*)-**9** in excellent yield.

Mechanistic aspect. To obtain the stereochemical course in the oxidative addition step, a similar synthetic sequence using deuterated aziridine *cis*-**1a**-*d*₁ as a substrate was conducted (Scheme 2). The stereochemistry on the terminal carbon (the 3-position) was completely inverted (for details, see the ESI†). This result agrees with the S_N2 nature of the oxidative addition process of aziridine toward Pd(0) complexes reported by our group^{11b} and others.^{5,11a} Moreover, this stereo-invertive process is supported with computational studies (*vide infra*).

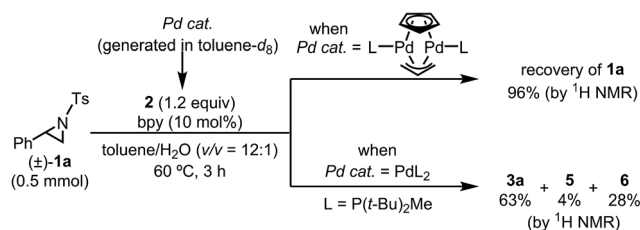
To identify the Pd species generated by mixing P(*t*-Bu)₂Me and Cp(allyl)Pd, we monitored the ³¹P{¹H} NMR spectra of these mixtures in THF-*d*₈ at 60 °C (Fig. 1). Upon addition of 2 equiv. of Cp(allyl)Pd to L (*i.e.*, L : Pd = 0.5), the sharp resonance corresponding to free L at δ 15.3 ppm (Fig. 1a, lit. δ 12.5 ppm at room temperature in toluene-*d*₈),³¹ completely disappeared and three new peaks appeared at δ 37.0 (sharp, strong), 44.7 (broad, weak), and 62.0 (broad, weak) ppm (Fig. 1b). According to the data reported by the Baird group,³² these peaks are assignable to a dinuclear Pd(I) complex [Pd₂L₂(μ-Cp)(μ-allyl)] (lit. δ 35.4 ppm at 65 °C in toluene-*d*₈), a bisphosphine Pd(0) complex PdL₂ (lit. δ 41.0 ppm at 65 °C in toluene-*d*₈), and an [(η⁵-Cp)Pd(η¹-allyl)L] species ([[(η⁵-Cp)Pd(η¹-PhC₃H₄)L] lit. δ 61.7 ppm at room temperature in toluene-*d*₈), respectively, as illustrated in Fig. 1. After stirring the mixture at 60 °C for 1 h, the broadened weak peak at δ 62.0 ppm corresponding to [(η⁵-Cp)Pd(η¹-allyl)L]



Scheme 1 Preparation of enantiopure β-aminoboronate and its synthetic application.



Scheme 2 Borylation of deuterated aziridine.



Scheme 3 Verification test of the catalytically active species.

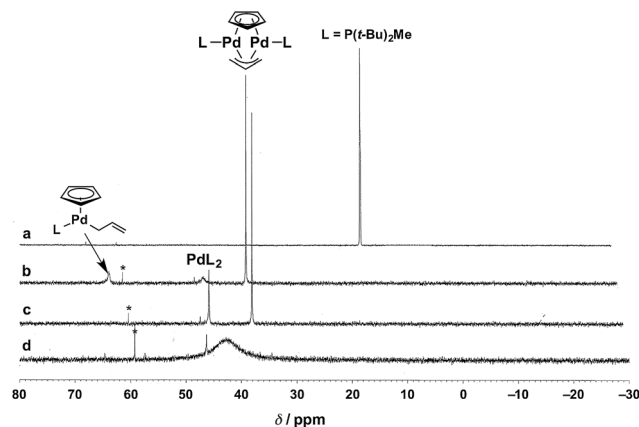


Fig. 1 Stacked $^{31}\text{P}\{^1\text{H}\}$ NMR spectra acquired at 60 °C in $\text{THF-}d_8$. (a) Free L [$\text{L} = \text{P}(\text{t-Bu})_2\text{Me}$]; (b, c) the reaction of $\text{Cp}(\text{allyl})\text{Pd}$ with 0.5 equiv. of L (b) upon mixing, (c) 1 h; (d) the reaction of $\text{Cp}(\text{allyl})\text{Pd}$ with 1 equiv. of L after 1 h. The peaks with asterisks are attributed to phosphine oxide.

disappeared, and the intensity of the peak at δ 44.7 ppm (PdL_2) increased as well (the integration ratio of $[\text{Pd}_2\text{L}_2(\mu\text{-Cp})(\mu\text{-allyl})] : \text{PdL}_2 = 3 : 2$, Fig. 1c). These peaks were not changed when bpy (10 equiv.) was added (Fig. S1b in the ESI†). This result would indicate that bpy does not coordinate to these Pd species, which was also supported by our computational results (*vide infra*). Moreover, upon addition of another equivalent of L (*i.e.*, $\text{L} : \text{Pd} = 1$), the strongest sharp peak at δ 37.0 ppm (dinuclear Pd complex) disappeared, and a broad peak around δ 42.6 ppm appeared (Fig. 1d). Furthermore, addition of one more equivalent of L ($\text{L} : \text{Pd} = 2$) gave a significantly broadened peak ($\delta \sim 30$ ppm, Fig. S1a in the ESI†), indicating significant exchange among several PdL_n species. Our observations could indicate that PdL_2 undergoes rapid exchange with PdL_n in $\text{THF-}d_8$ when the $\text{L} : \text{Pd}$ ratio is greater than 0.5. Therefore, the same would be the case with the real reaction system.

Based on the NMR experiments, $[\text{Pd}_2\text{L}_2(\mu\text{-Cp})(\mu\text{-allyl})]$ or PdL_2 would be the catalytically active Pd species (*i.e.*, the starting point of the catalytic cycle). To determine which is the active species in the catalytic reaction, both complexes were generated separately as the sole components according to Baird's conditions (*i.e.*, 1 : 1 and 2 : 1 mixtures of L and $\text{Cp}(\text{allyl})\text{Pd}$ were mixed in toluene- d_8 and stirred at 77 °C for 1 h),³² and the resulting Pd complexes were used as a catalyst in the borylation reaction in toluene (Schemes 3). As a result, while the dinuclear $\text{Pd}(\text{I})$ complex was found to be totally inert for the borylation reaction, the bisphosphine complex $\text{Pd}(\text{O})\text{L}_2$ gave almost the same result as the real reaction system. Therefore, we conclude that $\text{Pd}(\text{O})\text{L}_2$ would be the starting active species in the catalytic

cycle, and this conclusion was strongly supported by theoretical calculations (*vide infra*).

Computational part

Computational methods. Geometry optimizations were performed with the B3LYP-D3^{33–36} functional as implemented in the Gaussian09 program.³⁷ The polarizable continuum model (PCM) was used as the implicit solvent model with a dielectric constant (ϵ) of 2.3714.³⁸ The SDD basis set and associated effective core potential were applied for Pd,^{39,40} and 6-31G(d) basis sets were used for the other atoms (BS1).^{41–44} All stationary points, local minima (LMs) or transition states (TSs), were optimized without any constraint. Vibrational frequency calculations were performed to confirm the nature of the stationary points. Pseudo-IRC calculations were performed to confirm the connectivity between the TSs and LMs. The final potential energies of the optimized structures were calculated as single-point energies on the optimized structures, where a SDD basis set was used for Pd, and cc-pVTZ basis sets were used for the other atoms (BS2).^{45–47} In the results section, we report both the relative Gibbs free energy (ΔG) at 298.15 K and 1 atm and the relative electronic energy with the zero point energy correction (ΔE).

The aziridine ring opening step of the mechanism takes place *via* many different TSs, and therefore a proper sampling is very important. An automatic exploration of all important reaction pathways was accomplished using the multi-component artificial force induced reaction (MC-AFIR) method, as implemented in the global reaction route mapping (GRRM) strategy.^{48–50} An artificial force parameter (γ) of 300 kJ mol^{-1} was applied, and this is suitable for finding TSs within 300 kJ mol^{-1} . The MC-AFIR calculations were terminated when no new AFIR LM was found for 10 consecutive AFIR paths (NFault = 10). In AFIR calculations, the energy and derivatives were obtained using the ONIOM(B3LYP-D3:PM6-D3) method.^{51–57} Partitioning of the molecular system is shown in Fig. 2a. A model catalyst was used for MC-AFIR calculations (Fig. 2). SDD basis sets and the associated effective core potentials were used for palladium, and 3-21G basis sets (BS3) were applied for the high-level region of ONIOM calculations.^{58–60} All AFIR paths were inspected and approximate TSs were identified. Then, the real phosphine ligand was introduced to all approximate TSs, and the TS structures were fully optimized (without artificial force) with B3LYP-D3/BS1 method. A Boltzmann distribution of the transition states was used to calculate the regioselectivity.

Energy decomposition analysis (EDA)^{61,62} was performed for the key TSs leading to the desired and the undesired products. B3LYP-D3/BS2 level and PCM were used for the EDA. In this analysis, a TS structure was divided into **A** (catalyst) and **B** (substrate) (Fig. 3). INT_{AB} is the interaction energy between **A** and **B** at their optimized TS structure. The deformation energy (DEF) is the energy of **A** and **B** at the optimized TS, relative to the optimized structures of isolated **A** and **B** (denoted as A_0 and B_0). The energy difference ($\Delta\Delta E$) between the optimized transition states, $(\text{AB})_1$ and $(\text{AB})_2$, is the sum of the internal energy difference ($\Delta\text{INT}_{\text{AB}}$) and the deformation energy difference (ΔDEF).

Thermodynamically stable complexes in solution. Our starting point was to explore thermodynamically stable complexes in solution. First, we checked the relative stability of PdL , PdL_2 , PdL_3 , and PdL_4 complexes [$\text{L} = \text{P}(\text{t-Bu})_2\text{Me}$]. Among them, the PdL_2 complex (**I**) is the most stable complex, and this is in agreement with our NMR studies. The PdL_3 and PdL_4 complexes are 3.4 and 26.0 kcal mol^{-1} higher in energy, respectively. Despite several attempts, no PdL_4 complex was found, as one of the four ligands (**L**) always dissociated due to steric repulsion. We used PdL_2 (**I**) as a candidate to generate other possible complexes in solution. Moreover, the two vacant sites of the PdL_2 complex, **A** and **B**, can be filled by the potential ligands in solution, specifically H_2O , solvent (denoted as **Sol**), and additive (denoted as **bpy**). Table 3 summarizes the possible complexes and energies relative to the PdL_2 complex (**I**).

Starting from the PdL_2 complex (**I**), coordination of one H_2O , **L**, and **bpy** on **I** gives rise to $\text{PdL}_2(\text{H}_2\text{O})$ (**Iw**, $-0.5 \text{ kcal mol}^{-1}$), PdL_3 (3.4 kcal mol^{-1}), and $\text{PdL}_2(\text{bpy})$ (4.0 kcal mol^{-1}) complexes, respectively. Among the three-coordinate complexes, **Iw** is the thermodynamically most stable complex, and is only 0.5 kcal mol^{-1} more stable than **I**. Therefore, both **I**

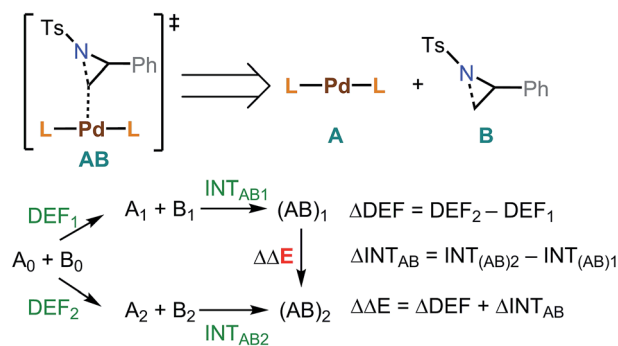


Fig. 3 EDA between two optimized transition states leading to the desired and undesired products.

and **Iw** can be formed in solution. In **Iw**, an H_2O molecule coordinates to the metal with one of the two hydrogen atoms. It is important to note that the additive (**bpy**) can coordinate to the metal, and the resulting complex, $\text{PdL}_2(\text{bpy})$, is, however, 4.5 kcal mol^{-1} higher than the most stable complex, **Iw**. All four-coordinate complexes found in the calculations, $\text{PdL}_2(\text{bpy})(\text{H}_2\text{O})$ (2.8 kcal mol^{-1}), $\text{PdL}_2(\text{H}_2\text{O})_2$ (4.7 kcal mol^{-1}), and $\text{PdL}_3(\text{H}_2\text{O})$ (7.9 kcal mol^{-1}) are relatively higher in energy. Coordination of solvent (**Sol**) is not possible due to steric repulsion between the bulky groups of the solvent molecules (**MTBE**) and **I**. We concluded that the thermodynamically most stable complexes in solution are **Iw** and **I**. The general rules that lead to the formation of the thermodynamically stable complexes in solution are: (1) coordination of two ligands on PdL_2 , which gives rise to four-coordinate complexes, is not favorable; (2) in terms of making three-coordinate complexes, the binding preference of the third ligand follows the order $\text{H}_2\text{O} > \text{L} > \text{bpy}$; (3) solvent **MTBE** would not coordinate to the PdL_2 complex due to steric repulsion; (4) coordination of the additive (**bpy**) is not favorable. According to Table 3, additive coordination on PdL_2 is not favorable as the subsequent complex $\text{PdL}_2(\text{bpy})$ is 4.0 kcal mol^{-1} higher than PdL_2 , while the $\text{PdL}_2(\text{bpy})_2$ complex is not stable. At the same time, it is important to note that aziridine binding on PdL_2 is relatively easier (*vide infra*), and the corresponding adduct is only 2.6 kcal mol^{-1} higher than PdL_2 . Therefore, the additive would not coordinate to the PdL_2 complex before the aziridine substrate binding. In our present mechanistic study, we do not consider the role of the additive in

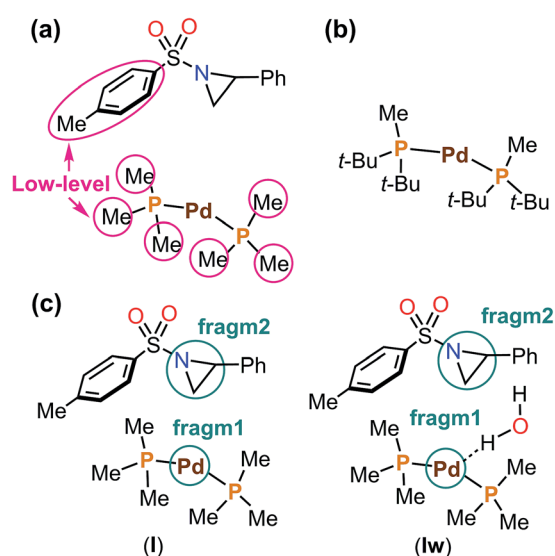


Fig. 2 (a) ONIOM partitioning of the model complex into high- and low-levels, (b) the real complex, (c) the artificial force ($\gamma = 300 \text{ kJ mol}^{-1}$) was applied between **fragm1** (catalyst) and **fragm2** (substrate) without or with an explicit water molecule.

Table 3 Possible complexes $\text{Pd}(\text{L})_2(\text{A})(\text{B})$ [$\text{L} = \text{P}(\text{t-Bu})_2\text{Me}$, **Sol** = **MTBE**] and their relative ΔG (ΔE in parentheses) in kcal mol^{-1a}

A	B				
	Empty	H_2O	bpy	L	Sol
Empty	0.0, (0.0)	-0.5, (-8.2)	4.0, (-7.4)	3.4, (-10.9)	X, (Sol)
H_2O	—	4.7, (-14.0)	2.8, (-17.9)	7.9, (-15.1)	X, (Sol)
bpy	—	—	X, (bpy)	X, (L)	X, (Sol)
L	—	—	—	X, (L)	X, (Sol)
Sol	—	—	—	—	X, (Sol)

^a The symbol "X" indicates that the ligand in parentheses dissociates upon structure optimization.



the mechanism because the reaction works even in the absence of additive (Table 1, entry 4).

Aziridine ring-opening step. Our next task was the exploration of the aziridine ring-opening step. For this purpose, we used the most stable **I** and **Iw** complexes as the active intermediates for the reaction, and 2-phenyl-*N*-tosyl-aziridine (**1**) was used as the substrate. Therefore, two independent MC-AFIR calculations, with and without an explicit water molecule, were performed as shown in Fig. 2c. The fully optimized TS structures, 73 TSs from **Iw** and 33 TSs from **I**, were categorized into 15 groups: **TSI-II_A** through **TSI-II_J** from **Iw**, and **TSI-II_K** through **TSI-II_O** from **I**, based on structural similarities (Fig. 4). The lowest energy TS of each group, their relative energies, and their existence probability are depicted in Table 4 (see Table S14 in the ESI† for a full description of all the TSs).

Among the calculated TSs (Table 4), **TS1-t-Iw** (12.6 kcal mol⁻¹) which belongs to the **TSI-II_K** group is the lowest energy TS of this step. In this TS, the aziridine ring opening takes place at the less hindered carbon in an S_N2 fashion, leading to the desired product forming. The same product can be obtained through **TS2-t-Iw** (12.9 kcal mol⁻¹), **TS4-t-I** (13.3 kcal mol⁻¹), **TS5-t-Iw** (13.4 kcal mol⁻¹), **TS6-t-I** (13.5 kcal mol⁻¹), **TS7-t-I** (13.6 kcal mol⁻¹), **TS8-t-I** (13.7 kcal mol⁻¹), and **TS9-t-Iw** (13.9 kcal mol⁻¹). On the other hand, at **TS3-b-Iw** (13.2 kcal mol⁻¹) of the group **TSI-II_L**, aziridine ring opening occurs at the hindered carbon (the 2-position), and this is the lowest TS leading to the undesired product. Based on a Boltzmann distribution over the calculated TSs, the regioselectivity is calculated to be 89 : 11, which is qualitatively in agreement with the experimental results (99 : 1).

In the calculated TSs for **TSI-II_E** (31.5 kcal mol⁻¹), **TSI-II_F** (30.9 kcal mol⁻¹), and **TSI-II_M** (24.7 kcal mol⁻¹), aziridine ring opening occurs at the less hindered carbon (Fig. 4). However,

these TSs are relatively higher in energy due to the fact that the *N*-*p*-toluenesulfonyl and 2-phenyl substituents of the substrate are much closer to the bulky groups of **Iw** or **I**. Similarly, TSs in the group of **TSI-II_G** (31.3 kcal mol⁻¹), **TSI-II_H** (31.9 kcal mol⁻¹), and **TSI-II_N** (28.1 kcal mol⁻¹) show high energy barriers, where aziridine ring opening occurs at the hindered carbon. TSs in the group of **TSI-II_I** (31.8 kcal mol⁻¹), **TSI-II_J** (31.8 kcal mol⁻¹), and **TSI-II_O** (28.1 kcal mol⁻¹) represent the aziridine ring opening through the carbon-carbon bond, which is, however, not possible due to very large reaction barriers. Furthermore, we checked the effect of H₂O molecules on aziridine ring opening. Starting from the lowest energy TS for this step (**TS1-t-Iw**, 12.6 kcal mol⁻¹), up to four H₂O molecules were introduced, and the corresponding TSs were calculated. The calculated ΔG of **TS1-t-Iw** with one H₂O molecule (**TS1-t-Iw-w**, 16.1 kcal mol⁻¹), two H₂O molecules (**TS1-t-Iw-w₂**, 14.3 kcal mol⁻¹), three H₂O molecules (**TS1-t-Iw-w₃**, 15.6 kcal mol⁻¹), and four H₂O molecules (**TS1-t-Iw-w₄**, 16.4 kcal mol⁻¹) suggested that explicit H₂O molecules in the system would not stabilize the **TS1-t-Iw**.

TS1-t-Iw is a major contributor to the aziridine ring opening at the less hindered carbon, while **TS3-b-Iw** is the lowest TS for aziridine ring opening at the hindered carbon (Fig. 5). The calculated Gibbs free energy difference between **TS1-t-Iw** and **TS3-b-Iw** is 0.6 kcal mol⁻¹, while the potential energy difference is 3.2 kcal mol⁻¹. When we do not consider the zero point energy corrections, the potential energy difference is 3.5 kcal mol⁻¹, and we have used this value for the EDA. According to the EDA (Table 5), the origin of this difference comes from $\Delta\text{INT}_{\text{AB}}$ (3.9 kcal mol⁻¹), indicating better interactions at **TS1-t-Iw**. This is due to the fact that the aziridine substrate can approach closer to the metal in **TS1-t-Iw** (Pd-C = 2.47 Å) (vs. Pd-C = 2.61 Å at **TS3-b-Iw**) (Fig. 5) than in **TS3-b-Iw** due to the lower steric repulsion.

The free energy profile for the early stages of the mechanism is shown in Fig. 6. The reaction starts from the thermodynamically most stable complex, **Iw**. Therefore, we report energies relative to **Iw**. Then, aziridine coordination on **I** leads to a complex **II** (2.6 kcal mol⁻¹), and the subsequent aziridine ring opening occurs through **TS1-t-Iw** (**TSII-III**, 12.6 kcal mol⁻¹). Aziridine ring opening is also possible from **Iw** (not shown in Fig. 6) with an overall barrier of 13.3 kcal mol⁻¹ (**TS4-t-Iw**), which is, however, 0.7 kcal mol⁻¹ higher than **TS1-t-Iw**. Beyond **TS1-t-Iw**, an intermediate, **III** (8.4 kcal mol⁻¹) is formed.

Proton transfer to anionic amine. The next step of the mechanism would be a proton transfer from H₂O molecules in solution to the anionic nitrogen atom in the water-coordinated intermediate **III**. This proton transfer process takes place smoothly from **IIIw₃** (0.3 kcal mol⁻¹), which is 8.1 kcal mol⁻¹ more stable than **III**. Furthermore, three H₂O molecules make a hydrogen bonding network starting from the anionic amine that is connected to the metal through a water oxygen coordination. The TS for this proton transfer process is 12.0 kcal mol⁻¹ (**TSIII-IVw₃**). Proton transfer is also possible from the analogous intermediate with four H₂O molecules, **IIIw₄** (11.3 kcal mol⁻¹), and the subsequent transition state is, however, 1.3 kcal mol⁻¹ higher than **TSIII-IVw₃** (not shown in Fig. 6). We were unable to locate TSs starting from the analogous

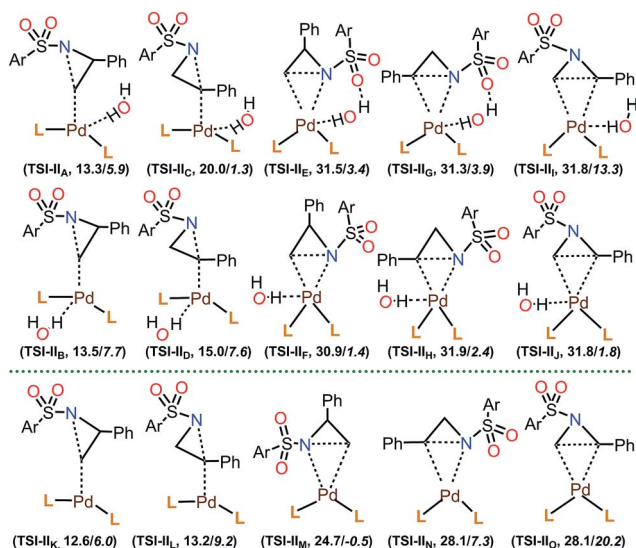


Fig. 4 Groups of transition states for the aziridine ring opening, starting from **Iw** (**TSI-II_A** through **TSI-II_J**) and **I** (**TSI-II_K** through **TSI-II_O**), and energies of the lowest energy TS of each group relative to the PdL₂(H₂O) (**Iw**) complex (Ar = *p*-tolyl; L = P(*t*-Bu)₂Me). ΔG values are in plain text, and ΔE values are in italics.



Table 4 Low energy TSs for the aziridine ring opening starting from **I** and **Iw**^a

TS	Group/regioselectivity	$\Delta\Delta G$	ΔG (ΔE)	Existence probability (%)
TS1-t-lw	K/terminal	0.0	12.6 (6.0)	27.6
TS2-t-lw	K/terminal	0.3	12.9 (5.4)	17.9
TS3-b-lw	L/benzylic	0.6	13.2 (9.2)	10.0
TS4-t-l	A/terminal	0.7	13.3 (5.9)	9.6
TS5-t-lw	K/terminal	0.8	13.4 (6.9)	7.1
TS6-t-l	B/terminal	0.9	13.5 (7.7)	6.1
TS7-t-l	B/terminal	1.0	13.6 (7.7)	4.8
TS8-t-l	A/terminal	1.1	13.7 (6.8)	4.6
TS9-t-lw	K/terminal	1.3	13.9 (7.6)	2.9
TS10-t-lw	K/terminal	2.3	14.9 (7.6)	0.6
TS11-b-l	D/benzylic	2.4	15.0 (7.6)	0.5
TS12-t-lw	K/terminal	2.6	15.2 (−1.2)	0.4
TS13-t-lw	K/terminal	2.6	15.2 (8.3)	0.3
TS14-t-lw	K/terminal	2.7	15.3 (7.5)	0.3
TS15-t-lw	K/terminal	3.3	15.9 (9.0)	0.1
TS16-t-lw	K/terminal	3.5	16.1 (8.3)	0.1
TS17-t-lw	K/terminal	3.6	16.2 (7.9)	0.1
TS18-t-l	B/terminal	3.7	16.3 (9.2)	0.1
TS19-t-l	A/terminal	3.7	16.3 (−0.8)	0.1

^a ΔG and ΔE values are indicated in kcal mol^{−1}, relative to **Iw**. “t” and “b” indicate the terminal (3-position) and benzylic carbons (2-position), respectively.

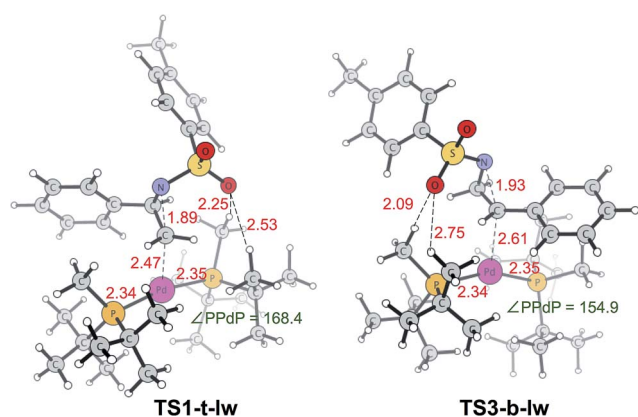


Fig. 5 Lowest energy TSs leading to the desired product (TS1-t-lw) and undesired product (TS3-b-lw). Bond lengths (red) are in Å and bond angles (green) are in degrees.

Table 5 EDA for the potential energy difference (kcal mol^{−1}) between the key transition states, TS1-t-lw and TS3-b-lw

TS	DEF (DEF _A , DEF _B)	INT _{AB}	ΔE
TS1-t-lw	23.2 (1.2, 22.0)	−26.5	−3.3
TS3-b-lw	22.8 (2.7, 20.1)	−22.6	0.2
	ΔDEF	ΔINT_{AB}	$\Delta\Delta E$
	−0.4 (−1.5, 1.1)	3.9	3.5

intermediates with two H₂O molecules (**IIIw**₂, 1.1 kcal mol^{−1}) and one H₂O molecule (**IIIw**, 2.3 kcal mol^{−1}), where the potential energy surface for the proton transfer processes was found to be repulsive (Fig. S2 and S3†). In **IIIw**₃, three water molecules form a complete head-to-tail chain between the anionic amines and Pd centres. As a result, the synchronous three proton

transfer via **TSIII–IVw**₃ takes place and gives a protonated N, two H₂O molecules, and a hydroxyl group bound to the metal center. Based on our analysis, we conclude that three H₂O molecules are required for the proton transfer process, leading to an intermediate **IVw**₃ (2.0 kcal mol^{−1}). Once the proton is transferred, the chain of water molecules is no longer required, and the corresponding intermediate without H₂O, **IV**, is 2.7 kcal mol^{−1} more stable than **IVw**₃.

B–B bond coordination and cleavage. Now, the B(pin)–B(pin) species binds to **IV** and the subsequent intermediate **V** (0.5 kcal mol^{−1}) is formed. Starting from **V**, inner-sphere B–B bond cleavage does not take place due to steric repulsion between the bulky alkyl groups in **L** and B(pin)–B(pin). One of the two phosphine ligands of the catalyst must dissociate, and the resulting intermediate, **VI**, is only 1.8 kcal mol^{−1} higher than **V**. Then, B–B bond cleavage takes place in an inner sphere fashion with a barrier of 10.8 kcal mol^{−1} (**TSVI–VII**), giving rise to **VII** (−2.1 kcal mol^{−1}). We have checked the possibility of ligand exchange between the (pin)B–OH and a stronger phosphine ligand **L** on intermediate **VII**, and the resulting complex, **VIII**, is 24.5 kcal mol^{−1} more stable than **VII**.

cis–trans isomerization and reductive elimination. The energy profiles for the final stages of the reaction are shown in Fig. 7. In both the intermediates **VII** and **VIII**, B(pin) is *trans* to the alkyl group derived from the aziridine substrate, and therefore *cis/trans* isomerization must take place before the reductive elimination. The *cis* isomer **VIII'** is 2.5 kcal mol^{−1} more stable than the *trans* isomer **VIII**. This step may occur through ligand dissociation processes. Since we expect this isomerization to be a low energy process, we did not study this in depth. The subsequent reductive elimination from **VIII'** occurs through a barrier of 0.1 kcal mol^{−1} (**TSVIII'–IX**). In the resulting intermediate **IX**, the product (**P**) is still at the metal



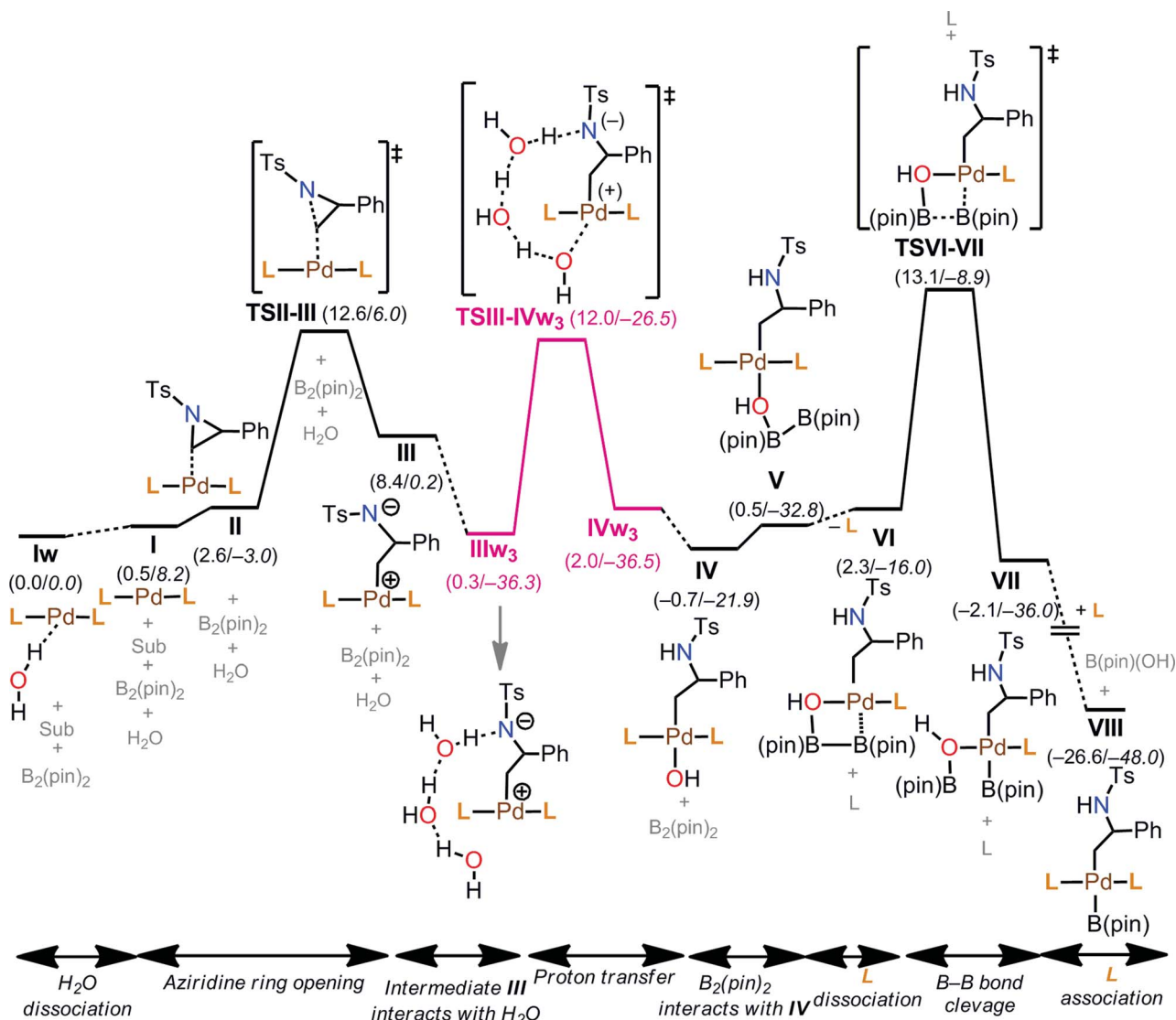


Fig. 6 Free energy profile for the reaction mechanism (aziridine ring opening, proton transfer, and B–B bond cleavage processes). Energies are indicated in kcal mol^{-1} . ΔG values are in plain text, and ΔE values are in italics.

coordination sphere, and **P** can be removed easily to recover the active form of the catalyst, PdL_2 (**I**), for the next catalytic cycle. Similarly, the formation of the *cis* isomer of **VII**, the less stable intermediate **VII'** ($-19.4 \text{ kcal mol}^{-1}$), and the subsequent reductive elimination can occur easily with a barrier of $5.7 \text{ kcal mol}^{-1}$ (**TSVII'–X**). Then, a phosphine ligand (**L**) can coordinate to the resulting intermediate **X**, which ultimately yields the product, **P**.

Side product formation

Under the optimized reaction conditions, imine **6** is formed as a side product (Table 1). We have explored the mechanism for the side reaction as shown in Fig. S4 in the ESI.†

Catalytic cycle

Putting together the present results from the experimental and theoretical studies, the proposed catalytic cycle for the bor-ylation is shown in Fig. 8 (black cycle). $Pd(0)L_2$ is generated

through the sequential reduction of $Cp(allyl)Pd$. Then, $Pd(0)L_2$ attacks the less hindered carbon of the aziridine in an S_N2 fashion (a) to give a stereo-inverted oxidative adduct. A hydrogen bonded chain of H_2O molecules plays two roles in the following steps: (i) as a proton source to quench TsN^- (b) and (ii) as an internal base to form the $[Pd(OH)]$ species. This intermediate activates the B–B bond of $B_2(pin)_2$ (c). The dissociation of a phosphine ligand (d) facilitates the transmetalation (e). The free phosphine ligand participates in the catalytic cycle again to form the $(alkyl)PdL_2B(pin)$ complex (f). The *trans*–*cis* isomerization (g) followed by reductive elimination (h) leads to the C–B cross-coupled product and the completion of the catalytic cycle. A side reaction may occur after the oxidation addition (brown cycle): the oxidative adduct of aziridine undergoes hydrogen transfer (i) followed by reductive elimination (j) to produce an imine byproduct. Our computational study suggests that the aziridine ring-opening step for substrate **1a** has barriers of $12.6 \text{ kcal mol}^{-1}$ (terminal position) and

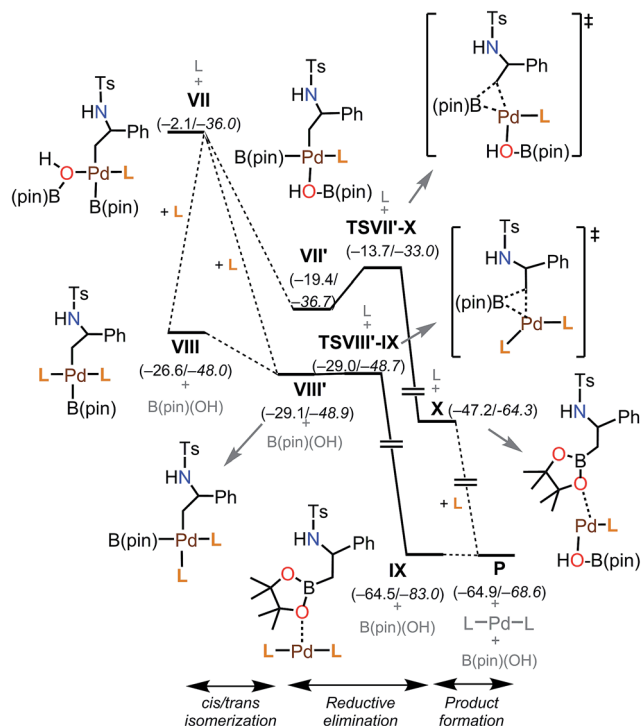


Fig. 7 Free energy profile for the later steps of the mechanism. Energies are indicated in kcal mol⁻¹, ΔG values are in plain text, and ΔE values are in italics.

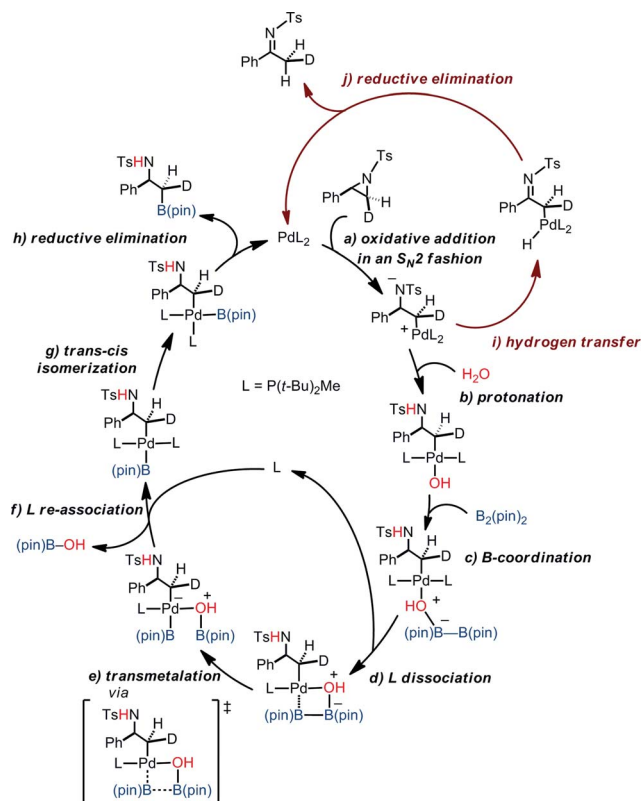


Fig. 8 A proposed catalytic cycle of the borylation.

13.2 kcal mol⁻¹ (benzylic position) (Table 4). For an alkyl aziridine, in particular **11**, the aziridine ring-opening step showed barriers of 15.4 kcal mol⁻¹ (2-position) and 16.3 kcal mol⁻¹ (3-position). In the absence of the aryl group, the aziridine ring opening of **11** is 2.8 kcal mol⁻¹ higher than **1**, where the reaction is difficult with **11**. This is qualitatively in agreement with our experimental results, where the reaction does not proceed with **11**.

Conclusions

In conclusion, we have developed a Pd-catalyzed regioselective borylative ring opening reaction of 2-arylaziridines to give β -aminoalkylboronates that are otherwise difficult to synthesize using existing methodologies. Importantly, the regioselectivity of the ring opening is controlled by the interactions between the catalyst and the substrate. The S_N2 nature of the oxidative addition of aziridine was verified using deuterated aziridine and computational studies. Furthermore, the borylative reaction is applicable under neutral conditions that allow for high functional compatibility. The mechanism of the full catalytic cycle was proposed using DFT and MC-AFIR methods. The aziridine ring opening is initiated by the active species PdL₂. TSs were systematically determined for the selectivity-determining aziridine ring opening step, and the calculations reasonably reproduced the experimental regioselectivity. The next step of the mechanism is a proton transfer that is facilitated by a H₂O hydrogen bond chain. The resulting Pd-hydroxo species activates the transmetalation step, where inner sphere boron-boron bond cleavage occurs, and leads to the final reductive elimination step. These experimental and theoretical findings open up an avenue to the further development of transition metal-catalyzed ring-opening C–E bond forming cross couplings of aziridines.

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Notes and references

- 1 *Aziridines and Epoxides in Organic Synthesis*, ed. A. K. Yudin, Wiley-VCH, Weinheim, 2006.
- 2 For reviews on ring opening of aziridines, see: (a) D. Tanner, *Angew. Chem., Int. Ed. Engl.*, 1994, **33**, 599–619; (b) X. E. Hu, *Tetrahedron*, 2004, **60**, 2701–2743; (c) P. Lu, *Tetrahedron*, 2010, **66**, 2549–2560; (d) S. Stanković, M. D'hooghe, S. Catak, H. Eum, M. Waroquier, V. V. Speybroeck, N. De Kimpe and H. J. Ha, *Chem. Soc. Rev.*, 2011, **41**, 643–665.
- 3 For a review on the preparation and reactivities of azametallacyclobutanes, see: A. Dauth and J. A. Love, *Dalton Trans.*, 2012, **41**, 7782–7791.
- 4 B. L. Lin, C. R. Clough and G. L. Hillhouse, *J. Am. Chem. Soc.*, 2002, **124**, 2890–2891.
- 5 J. E. Ney and J. P. Wolfe, *J. Am. Chem. Soc.*, 2006, **128**, 15415–15422.
- 6 (a) H. Alper, F. Urso and D. J. H. Smith, *J. Am. Chem. Soc.*, 1983, **105**, 6737–6738; (b) S. Calet, F. Urso and H. Alper, *J. Am. Chem. Soc.*, 1989, **111**, 931–934; (c) M. E. Piotti and H. Alper, *J. Am. Chem. Soc.*, 1996, **118**, 111–116; (d) P. Davoli, I. Moretti, F. Prati and H. Alper, *J. Org. Chem.*, 1999, **64**, 518–521.
- 7 For a review on the transition metal-catalyzed transformations of small ring compounds including aziridines, see: C.-Y. Huang and A. G. Doyle, *Chem. Rev.*, 2014, **114**, 8153–8198.
- 8 A wide variety of ring expansions of vinylaziridines via π -allylmetal species and ring-opening S_N2' substitutions have been developed. For reviews, see: (a) E. A. Ilardi and J. T. Njardarson, *J. Org. Chem.*, 2013, **78**, 9533–9540; (b) H. Ohno, *Chem. Rev.*, 2014, **114**, 7784–7814.
- 9 Wolfe and coworkers reported a Pd/PR₃-catalyzed isomerization of aziridines to ketimines via a tandem process consisting of oxidative addition/ β -hydride elimination/tautomerization, see: J. P. Wolfe and J. E. Ney, *Org. Lett.*, 2003, **5**, 4607–4610.
- 10 (a) C.-Y. Huang and A. G. Doyle, *J. Am. Chem. Soc.*, 2012, **134**, 9541–9544; (b) D. K. Nielsen, C.-Y. Huang and A. G. Doyle, *J. Am. Chem. Soc.*, 2013, **135**, 13605–13609; (c) C.-Y. Huang and A. G. Doyle, *J. Am. Chem. Soc.*, 2015, **137**, 5638–5641; (d) K. L. Jensen, E. A. Standley and T. F. Jamison, *J. Am. Chem. Soc.*, 2014, **136**, 11145–11152.
- 11 (a) M. L. Duda and F. E. Michael, *J. Am. Chem. Soc.*, 2013, **135**, 18347–18349; (b) Y. Takeda, Y. Ikeda, A. Kuroda, S. Tanaka and S. Minakata, *J. Am. Chem. Soc.*, 2014, **136**, 8544–8547.
- 12 The Pineschi group reported a Ni(0)-catalyzed borylative ring opening of vinylaziridines that proceeds in a formal S_N2' fashion. Apparently, this reaction proceeds through an oxidative addition of vinylaziridines to Ni(0) to generate a π -allyl Ni(II) species. However, little is described about its mechanism: S. Crotti, F. Bertolini, F. Macchia and M. Pineschi, *Org. Lett.*, 2009, **11**, 3762–3765.
- 13 (a) *Boronic Acids: Preparation and Applications in Organic Synthesis, Medicine and Materials*, ed. D. G. Hall, Wiley-VCH, Weinheim, 2nd edn, 2011; (b) C. M. Crudden, B. W. Glasspoole and C. J. Lata, *Chem. Commun.*, 2009, 6704–6716; (c) H. K. Scott and V. K. Aggarwal, *Chem.-Eur. J.*, 2011, **17**, 13124–13132.
- 14 S. Sebelius, V. J. Olsson and K. J. Szabó, *J. Am. Chem. Soc.*, 2005, **127**, 10478–10479.
- 15 As a relevant transformation, a metal-free borylative ring opening of vinyl epoxides and aziridines has been reported: X. Sanz, G. M. Lee, C. Pubill-Ulldemolins, A. Bornet, H. Gulyás, S. A. Westcott, C. Bo and E. Fernández, *Org. Biomol. Chem.*, 2013, **11**, 7004–7010.
- 16 For selected examples of synthetic methods of β -aminoboronates, see: (a) D. N. Butler and A. H. Soloway, *J. Am. Chem. Soc.*, 1964, **86**, 2961; (b) D. H. Kinder and M. M. Ames, *J. Org. Chem.*, 1987, **52**, 2452–2454; (c) C. Morin, *Tetrahedron*, 1994, **50**, 12521–12569; (d) G. A. Molander and F. Vargas, *Org. Lett.*, 2007, **9**, 203–206; (e) N. Matsuda, K. Hirano, T. Satoh and M. Miura, *J. Am. Chem. Soc.*, 2013, **135**, 4934–4937; (f) Q. Li, C. W. Liskey and J. F. Hartwig, *J. Am. Chem. Soc.*, 2014, **136**, 8755–8765; (g) N. Ursinyova, R. B. Bedford and T. Gallagher, *Eur. J. Org. Chem.*, 2016, 673–677.
- 17 For reports on the biological activities of β -aminoethylboronates, see: (a) A. S. Gorovoy, O. Gozhina, J.-S. Svendsen, G. V. Tetz, A. Domorad, V. V. Tetz and T. Lejon, *J. Pept. Sci.*, 2013, **19**, 613–618; (b) A. S. Gorovoy, O. V. Gozhina, J. S. Svendsen, A. A. Domorad, G. V. Tetz, V. V. Tetz and T. Lejon, *Chem. Biol. Drug Des.*, 2013, **81**, 408–413.
- 18 Recently, Ni-catalyzed benzylic C(sp³)-N borylation of benzylammonium triflates has been reported: (a) H. Zhang, S. Hagihara and K. Itami, *Chem.-Eur. J.*, 2015, **21**, 16796–16800; (b) C. H. Basch, K. M. Cobb and M. P. Watson, *Org. Lett.*, 2016, **18**, 136–139; (c) J. Hu, H. Sun, W. Cai, X. Pu, Y. Zhang and Z. Shi, *J. Org. Chem.*, 2016, **81**, 14–24.
- 19 (a) T. Ishiyama, T. Ahiko and N. Miyaara, *Tetrahedron Lett.*, 1996, **37**, 6889–6892; (b) T. Ahiko, T. Ishiyama and N. Miyaara, *Chem. Lett.*, 1997, **26**, 811–812; (c) S. Sebelius, O. A. Wallner and K. J. Szabó, *Org. Lett.*, 2003, **5**, 3065–3068.
- 20 (a) H. Ito, C. Kawakami and M. Sawamura, *J. Am. Chem. Soc.*, 2005, **127**, 16034–16035; (b) H. Ito, S. Ito, Y. Sasaki, K. Matsuura and M. Sawamura, *J. Am. Chem. Soc.*, 2007, **129**, 14856–14857.
- 21 (a) C.-T. Yang, Z.-Q. Zhang, H. Tajuddin, C.-C. Wu, J. Liang, J.-H. Liu, Y. Fu, M. Czyżewska, P. G. Steel, T. B. Marder and L. Liu, *Angew. Chem., Int. Ed.*, 2012, **51**, 528–532; (b) H. Ito and K. Kubota, *Org. Lett.*, 2012, **14**, 890–893; (c) A. S. Dudnik and G. C. Fu, *J. Am. Chem. Soc.*, 2012, **134**, 10693–10697; (d) S. K. Bose, K. Fucke, L. Liu, P. G. Steel and T. B. Marder, *Angew. Chem., Int. Ed.*, 2014, **53**, 1799–1803; (e) T. C. Atack, R. M. Lecker and S. P. Cook, *J. Am. Chem. Soc.*, 2014, **136**, 9521–9523; (f) R. B. Bedford, P. B. Brenner, E. Carter, T. Gallagher, D. M. Murphy and D. R. Pye, *Organometallics*, 2014, **33**, 5940–5943.
- 22 For detailed information about the effects of the reaction conditions on yields and product distributions, see the ESI†



- 23 (a) M. R. Netherton and G. C. Fu, *Angew. Chem., Int. Ed.*, 2002, **41**, 3910–3912; (b) J. H. Kirchhoff, M. R. Netherton, I. D. Hills and G. C. Fu, *J. Am. Chem. Soc.*, 2002, **124**, 13662–13663; (c) I. D. Hills, M. R. Netherton and G. C. Fu, *Angew. Chem., Int. Ed.*, 2003, **42**, 5749–5752.
- 24 J. Johnson and T. Rovis, *Angew. Chem., Int. Ed.*, 2008, **47**, 840–871.
- 25 (a) H. F. Bettinger, M. Filthaus, H. Bornemann and I. M. Oppel, *Angew. Chem., Int. Ed.*, 2008, **47**, 4744–4747; (b) S. Bagherzadeh and N. P. Mankad, *J. Am. Chem. Soc.*, 2015, **137**, 10898–10901.
- 26 For a review on the cross coupling of alkyltrifluoroborates, see: G. A. Molander, *J. Org. Chem.*, 2015, **80**, 7837–7848.
- 27 T.-S. Mei, H. H. Patel and M. S. Sigman, *Nature*, 2014, **508**, 340–344.
- 28 (a) J. D. Scott and R. M. Williams, *Chem. Rev.*, 2002, **102**, 1669–1730; (b) K. W. Bentley, *Nat. Prod. Rep.*, 2006, **23**, 444–463.
- 29 L. Zhang, Z. Zuo, X. Leng and Z. Huang, *Angew. Chem., Int. Ed.*, 2014, **53**, 2696–2700.
- 30 L. K. Lukanov, A. P. Venkov and N. M. Mollov, *Synthesis*, 1987, 204–206.
- 31 D. M. Norton, E. A. Mitchell, N. R. Botros, P. G. Jessop and M. C. Baird, *J. Org. Chem.*, 2009, **74**, 6674–6680.
- 32 E. A. Mitchell and M. C. Baird, *Organometallics*, 2007, **26**, 5230–5238.
- 33 A. D. Becke, *Phys. Rev. A*, 1988, **38**, 3098–3100.
- 34 C. T. Lee, W. Yang and R. G. Parr, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1988, **37**, 785–789.
- 35 B. Miehlich, A. Savin, H. Stoll and H. Preuss, *Chem. Phys. Lett.*, 1989, **157**, 200–206.
- 36 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, **132**, 154104.
- 37 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr, J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski and D. J. Fox, *Gaussian 09, Revision D.01*, Gaussian, Inc, Wallingford, CT, 2009.
- 38 A. V. Marenich, C. J. Cramer and D. G. Truhlar, *J. Phys. Chem. B*, 2009, **113**, 6378–6396.
- 39 P. Fuentealba, H. Preuss, H. Stoll and L. Von Szentpály, *Chem. Phys. Lett.*, 1982, **89**, 418–422.
- 40 T. H. Dunning Jr and P. J. Hay, in *Modern Theoretical Chemistry: Methods of Electronic Structure Theory*, ed. H. F. Schaefer, Plenum, New York, 1977, vol. 3, pp. 1–28.
- 41 R. Ditchfield, W. J. Hehre and J. A. Pople, *J. Chem. Phys.*, 1971, **54**, 724–728.
- 42 W. J. Hehre, R. Ditchfield and J. A. Pople, *J. Chem. Phys.*, 1972, **56**, 2257–2261.
- 43 P. C. Hariharan and J. A. Pople, *Theor. Chim. Acta*, 1973, **28**, 213–222.
- 44 M. M. Francl, W. J. Pietro, W. J. Hehre, J. S. Binkley, M. S. Gordon, D. J. DeFrees and J. A. Pople, *J. Chem. Phys.*, 1982, **77**, 3654–3665.
- 45 T. H. Dunning Jr, *J. Chem. Phys.*, 1989, **90**, 1007–1023.
- 46 R. A. Kendall, T. H. Dunning Jr and R. J. Harrison, *J. Chem. Phys.*, 1992, **96**, 6796–6806.
- 47 D. E. Woon and T. H. Dunning Jr, *J. Chem. Phys.*, 1993, **98**, 1358–1371.
- 48 S. Maeda, T. Taketsugu and K. Morokuma, *J. Am. Chem. Soc.*, 2015, **137**, 3433–3445.
- 49 S. Maeda and K. Morokuma, *J. Chem. Phys.*, 2010, **132**, 241102.
- 50 S. Maeda and K. Morokuma, *J. Chem. Theory Comput.*, 2011, **7**, 2335–2345.
- 51 S. Maeda, K. Ohno and K. Morokuma, *Phys. Chem. Chem. Phys.*, 2013, **15**, 3683–3701.
- 52 S. Dapprich, I. Komáromi, K. S. Byun, K. Morokuma and M. J. Frisch, *J. Mol. Struct.: THEOCHEM*, 1999, **461–462**, 1–21.
- 53 T. Vreven, K. S. Byun, I. Komáromi, S. Dapprich, J. A. Montgomery, K. Morokuma and M. J. Frisch, *J. Chem. Theory Comput.*, 2006, **2**, 815–826.
- 54 T. Vreven and K. Morokuma, *J. Comput. Chem.*, 2000, **21**, 1419–1432.
- 55 F. Maseras and K. Morokuma, *J. Comput. Chem.*, 1995, **16**, 1170–1179.
- 56 T. Matsubara, S. Sieber and K. Morokuma, *Int. J. Quantum Chem.*, 1996, **60**, 1101–1109.
- 57 M. Svensson, S. Humbel, R. D. J. Froese, T. Matsubara, S. Sieber and K. Morokuma, *J. Phys. Chem.*, 1996, **100**, 19357–19363.
- 58 P. J. Hay and W. R. Wadt, *J. Chem. Phys.*, 1985, **82**, 270–283.
- 59 W. R. Wadt and P. J. Hay, *J. Chem. Phys.*, 1985, **82**, 284–298.
- 60 P. J. Hay and W. R. Wadt, *J. Chem. Phys.*, 1985, **82**, 299–310.
- 61 K. Morokuma, *J. Chem. Phys.*, 1971, **55**, 1236–1244.
- 62 K. Kitaura and K. Morokuma, *Int. J. Quantum Chem.*, 1976, **10**, 325–340.

