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# Putting an ultrahigh concentration of amine groups into a metal–organic framework for CO<sub>2</sub> capture at low pressures†

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Tremendous efforts have been devoted to increasing the CO<sub>2</sub> capture performance of porous materials, especially for low CO<sub>2</sub> concentration environments. Here, we report that hydrazine can be used as a diamine short enough to functionalize the small-pore metal–organic framework [Mg<sub>2</sub>(dobdc)] (H<sub>4</sub>dobdc = 2,5-dihydroxy-1,4-benzenedicarboxylic acid). By virtue of the ultrahigh concentration of free amine groups (6.01 mmol g<sup>-1</sup> or 7.08 mmol cm<sup>-3</sup>) capable of reversible carbamic acid formation, the new material [Mg<sub>2</sub>(dobdc)(N<sub>2</sub>H<sub>4</sub>)<sub>1.8</sub>] achieves a series of new records for CO<sub>2</sub> capture, such as single-component isotherm uptakes of 3.89 mmol g<sup>-1</sup> or 4.58 mmol cm<sup>-3</sup> at the atmospheric CO<sub>2</sub> concentration of 0.4 mbar at 298 K and 1.04 mmol g<sup>-1</sup> or 1.22 mmol cm<sup>-3</sup> at 328 K, as well as more than a 4.2 mmol g<sup>-1</sup> or 4.9 mmol cm<sup>-3</sup> adsorption/desorption working capacity under dynamic mixed-gas conditions with CO<sub>2</sub> concentrations similar to those in flue gases and ambient air.

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## Introduction

Applications of porous materials for capturing CO<sub>2</sub> from flue gas (*ca.* 150 mbar) and from air (*ca.* 0.4 mbar) have become an increasingly attractive area of research due to their importance in many applications related to the environment, energy and health.<sup>1–7</sup> Compared with post-combustion capture, direct air capture (DAC) is more challenging, since the CO<sub>2</sub> adsorption ability of most porous materials is low at trace CO<sub>2</sub> concentrations.<sup>8</sup>

Among the various types of adsorbents, porous coordination polymers (PCPs) or metal–organic frameworks (MOFs) with designable and modifiable pore surfaces are very attractive for CO<sub>2</sub> capture.<sup>9–50</sup> To increase the CO<sub>2</sub> uptake at extremely low pressures, incorporating alkylamine groups on the pore surface is the most effective strategy so far,<sup>46–50</sup> because of the chemisorption of CO<sub>2</sub> associated with the formation of zwitterionic carbamate or carbamic acid. For instance, [Mg<sub>2</sub>(dobpdc)(eda)<sub>1.6</sub>] (eda-Mg<sub>2</sub>(dobpdc), eda = ethylenediamine, H<sub>4</sub>dobpdc = 4,4'-dihydroxy-(1,1'-biphenyl)-3,3'-dicarboxylic acid)<sup>48</sup> holds the records for gravimetric and volumetric CO<sub>2</sub>

adsorption capacities with 2.85 mmol g<sup>-1</sup> and 2.72 mmol cm<sup>-3</sup>, respectively, at 298 K and 0.4 mbar, this is because it has a high concentration of free amine groups at 3.86 mmol g<sup>-1</sup> or 3.69 mmol cm<sup>-3</sup>. Obviously, the CO<sub>2</sub> adsorption does not reach saturation at this condition, which is ascribed to the strong intermolecular hydrogen bonds between two adjacent amine groups that need to be opened by higher pressure CO<sub>2</sub>.<sup>47,51</sup> As a consequence, its CO<sub>2</sub> uptake at 0.4 mbar sharply drops to 0.120 mmol g<sup>-1</sup> or 0.115 mmol cm<sup>-3</sup> at a slightly higher temperature of 323 K. To overcome this problem, replacing the eda molecule with an even shorter diamine can be a good strategy, which not only reduces the intermolecular hydrogen bonding locks but also reduces the adsorbent weight.

Hydrazine (N<sub>2</sub>H<sub>4</sub>) can be considered as the smallest/shortest diamine, and its basicity is only slightly weaker than that of alkylamines (Table S1†).<sup>52</sup> It should be noted that the long/large length/size of eda and its derivatives is the main reason for choosing the large-pore [Mg<sub>2</sub>(dobpdc)] instead of the simpler small-pore [Mg<sub>2</sub>(dobdc)] (H<sub>4</sub>dobdc = 2,5-dihydroxy-1,4-benzenedicarboxylic acid) as a host framework.<sup>49</sup> Nevertheless, to the best of our knowledge, an N<sub>2</sub>H<sub>4</sub>-grafted PCP has not been reported so far. Herein, we show that N<sub>2</sub>H<sub>4</sub> can be used to functionalize the prototypical small-pore PCP [Mg<sub>2</sub>(dobdc)] without the formation of the intermolecular amine–amine interaction, achieving an ultrahigh concentration of useful amine groups in both the gravimetric and volumetric points of view, as well as a series of new benchmark CO<sub>2</sub> capture performances, ranging from air to flue-gas conditions.

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## Results and discussion

The hypothetical structure of  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$ , *i.e.*  $[\text{Mg}_2(\text{dobdc})]$  fully functionalized by one  $\text{N}_2\text{H}_4$  molecule per  $\text{Mg}(\text{II})$  ion, was simulated using molecular mechanics (MM) combined with periodic density functional theory (PFT). During structural optimization, the host framework and  $\text{N}_2\text{H}_4$  molecules were considered as rigid and flexible, respectively. The simulation results showed that the void ratio of  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$  remains at 36.5%, and the smallest aperture diameter is *ca.* 4.8 Å (Fig. 1a), which should be enough for the diffusion and accommodation of  $\text{CO}_2$  molecules. Furthermore, the shortest possible intermolecular  $\text{N}\cdots\text{N}$  separations between pairs of adjacent amine groups on the *ab*-plane and along the *c*-axis were simulated as 3.81 and 4.30 Å, respectively (Fig. 1b and c), which are much longer than the values for typical  $\text{N}-\text{H}\cdots\text{N}$  hydrogen bonds. For comparison, we also simulated the structure of  $\text{eda}-\text{Mg}_2(\text{dobpdc})$ , giving relatively short  $\text{N}\cdots\text{N}$  separations of 3.06 and 3.07 Å on the *ab*-plane and along the *c*-axis, respectively (Fig. 1d and e), which can be assigned as a kind of weak hydrogen bonding.<sup>53</sup>

Yellow microcrystalline  $[\text{Mg}_2(\text{dobdc})]$  (**1**) was suspended in anhydrous toluene under  $\text{N}_2$  atmosphere at room temperature, and upon the addition of a toluene solution of anhydrous  $\text{N}_2\text{H}_4$ , the colour of the solid instantaneously changed to light yellow. The light yellow product was washed with toluene and then successively soaked in hexane and methanol to remove the

excess  $\text{N}_2\text{H}_4$ . Elemental analysis gave a chemical formula of  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_{1.8}]$  (**1a**) for the light yellow product. Pawley refinements of the powder X-ray diffraction (PXRD) patterns showed similar unit-cell parameters for **1** and **1a** and confirmed the retention of framework integrity after  $\text{N}_2\text{H}_4$  grafting (Fig. S1†). PXRD and thermogravimetry-mass spectrometry analyses showed that **1a** can be fully activated at 403 K (Fig. S1 and S2†), with negligible release of the  $\text{N}_2\text{H}_4$  molecules.  $\text{N}_2$  sorption isotherms measured at 77 K gave apparent Langmuir surface areas of 1925 and 1012  $\text{m}^2 \text{g}^{-1}$  and pore volumes of 0.67 and 0.30  $\text{cm}^3 \text{g}^{-1}$ , respectively, for **1** and **1a** (Fig. S3†). Considering that their crystallographic pore volumes are 0.66 and 0.31  $\text{cm}^3 \text{g}^{-1}$ , respectively, the  $\text{N}_2$  sorption data further confirmed the good sample crystallinity and purity, as well as the chemical formula of **1a** derived from elemental analysis. **1a** possesses a high concentration of surface-appended  $\text{N}_2\text{H}_4$  molecules of 6.01  $\text{mmol g}^{-1}$  or 7.08  $\text{mmol cm}^{-3}$ .

At 298 K and 1 bar (Fig. 2a), **1a** showed a  $\text{CO}_2$  uptake of 5.51  $\text{mmol g}^{-1}$  or 6.49  $\text{mmol cm}^{-3}$ , corresponding to 1.65  $\text{CO}_2$  molecules per formula unit or 0.919  $\text{CO}_2$  molecule per  $\text{N}_2\text{H}_4$  molecule, which are lower than the values for **1** (8.04  $\text{mmol g}^{-1}$  or 7.40  $\text{mmol cm}^{-3}$ ) (Table S2†), this is because the relatively large pore volume of **1** can accommodate additional physically adsorbed  $\text{CO}_2$ . At 0.15 bar, the gravimetric uptake of **1a**



Fig. 1 (a) Comparison of the structures of  $[\text{Mg}_2(\text{dobdc})]$  (drawn from the reported crystal structure) and  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$  (simulated from PFT). (b–e) Comparison of the shortest possible intermolecular hydrogen bonding in  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$  and  $[\text{Mg}_2(\text{dobpdc})(\text{eda})_2]$  (obtained using a PFT simulation, in which the flexibility of diamines was considered). (b)  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$  on the *a*–*b* plane, (c)  $[\text{Mg}_2(\text{dobdc})(\text{N}_2\text{H}_4)_2]$  along the *c*-axis, (d)  $[\text{Mg}_2(\text{dobpdc})(\text{eda})_2]$  on the *a*–*b* plane, and (e)  $[\text{Mg}_2(\text{dobpdc})(\text{eda})_2]$  along the *c*-axis.

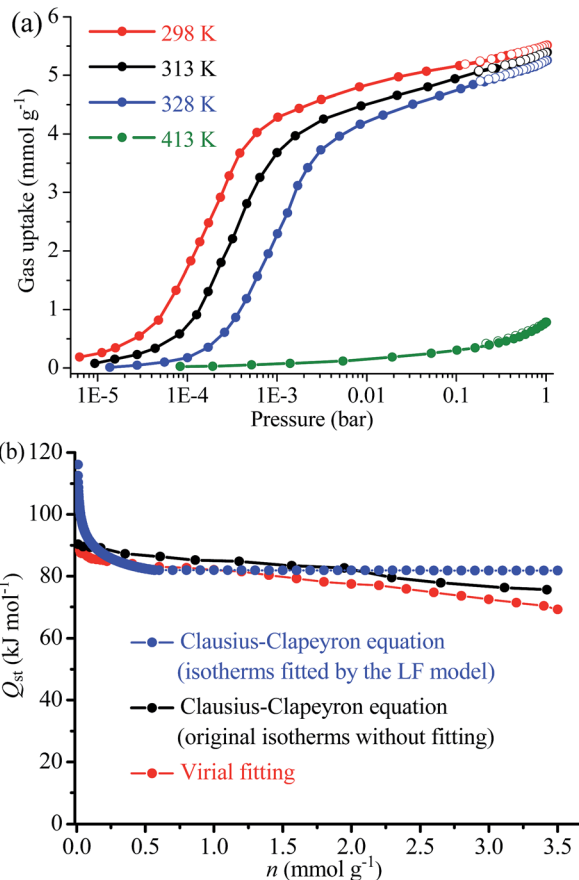


Fig. 2 (a)  $\text{CO}_2$  adsorption (solid) and desorption (open) isotherms measured at 298, 313, 328 and 413 K and (b) the coverage-dependent  $\text{CO}_2$  adsorption enthalpy of **1a**.



(5.18 mmol g<sup>-1</sup>) is still lower than that of **1** (5.71 mmol g<sup>-1</sup>). However, the volumetric uptake of **1a** (6.10 mmol cm<sup>-3</sup>) is 17% higher than that of **1** (5.25 mmol cm<sup>-3</sup>), illustrating the stronger CO<sub>2</sub> binding affinity of the amine groups. At a much lower pressure of 0.4 mbar, corresponding to the CO<sub>2</sub> concentration in air, **1a** showed an exceptionally high CO<sub>2</sub> uptake of 3.89 mmol g<sup>-1</sup> or 4.58 mmol cm<sup>-3</sup>, which is not only a drastic improvement over **1** (0.09 mmol g<sup>-1</sup> or 0.08 mmol cm<sup>-3</sup>), but also obviously higher than the previous records (Table S2†). The CO<sub>2</sub> sorption kinetics of **1a** were also evaluated using thermogravimetric analysis under dry air at 298 K. A CO<sub>2</sub> uptake of 3.66 mmol g<sup>-1</sup> or 16.1 wt% (3.71 mmol g<sup>-1</sup> or 16.3 wt% for the single-component CO<sub>2</sub> adsorption isotherm at 0.4 mbar and 298 K) was observed after *ca.* 950 min by switching the atmosphere from pure N<sub>2</sub> at 403 K to dry air at 298 K. The equilibrium time is much longer than that at higher CO<sub>2</sub> partial pressures, being similar to known examples such as eda-Mg<sub>2</sub>(dobpdc) (1000 min, Fig. S4†).<sup>12,48</sup> Interestingly, even at a higher temperature of 328 K, **1a** still exhibits a relatively high CO<sub>2</sub> uptake of 1.04 mmol g<sup>-1</sup> or 1.22 mmol cm<sup>-3</sup> at 0.4 mbar, which is 2–3 times higher than the previous records (Table S3†).

The coverage-dependent CO<sub>2</sub> adsorption enthalpy ( $Q_{st}$ ) of **1a** was calculated using the Clausius–Clapeyron equation and the Virial fitting method using isotherms measured at 298, 313, and 328 K (Fig. 2b, S5 and S6†). Among various isotherm models, only the Langmuir–Freundlich (LF) one gave a fair fitting. Based on the Clausius–Clapeyron equation, the Langmuir–Freundlich isotherms gave a very high  $Q_{st}$  of 118 kJ mol<sup>-1</sup> at zero-coverage, while the original isotherms (without data fitting) gave near-zero-coverage  $Q_{st}$  of 90 kJ mol<sup>-1</sup>. These  $Q_{st}$  values are higher than the values of most reported materials (Table S2†) and indicative of chemisorption. On the other hand, the Virial fitting gave similar results as the Clausius–Clapeyron equation using original isotherms. *In situ* infrared (IR) spectra of **1a** showed that the primary amine peaks at 3330 and 3287 cm<sup>-1</sup> weakened under CO<sub>2</sub> atmosphere. The residue primary amine groups can be assigned to those anchoring on the Mg(II) sites. Meanwhile, two new absorption peaks appeared at 3397 and 1250 cm<sup>-1</sup>. The former new peak may arise from the stretching band of the N–H bond and/or carboxylic O–H bond, while the latter can be assigned to the stretching band of the C–N bond of carbamic acid. The peak for the asymmetric deformation of the ammonium group at ~2200 cm<sup>-1</sup> was invisible, indicating the absence of proton transfer, which is consistent with the relatively long intermolecular separation of hydrazine molecules (Fig. S7†).<sup>54,55</sup> After being heated under vacuum at 403 K, this peak disappeared, indicating the chemisorption of CO<sub>2</sub> is reversible. The formation of carbamic acid was further confirmed *via* <sup>13</sup>C NMR measurements with the observation of a broad resonance peak centered at  $\delta = 162$  ppm (Fig. 3).<sup>22</sup>

The CO<sub>2</sub> adsorption and desorption behaviours of **1a** under mixed-gas and kinetic conditions were analyzed *via* thermogravimetry (Fig. 4), in which the sample was blown repeatedly under conditions between a 15 : 85 CO<sub>2</sub>/N<sub>2</sub> (v/v) mixture at 313 K (the typical flue-gas environment) and a pure N<sub>2</sub> flow (a typical regeneration method for temperature-vacuum swing adsorption (TVSA) like process) at 403 K. It should be noted that



Fig. 3 Solid-state <sup>13</sup>C NMR spectra of **1a** with and without adsorbed CO<sub>2</sub>. An asterisk (\*) marks the carbon atom of CO<sub>2</sub>.

the regeneration temperature is lower than for other alkylamine functionalized PCPs. By switching from pure N<sub>2</sub> at 403 K to the mixed gas and allowing the temperature to automatically decrease to 313 K after 17 min, a weight increase of 17.0% was observed, corresponding to volumetric CO<sub>2</sub> uptake of 3.86 mmol g<sup>-1</sup> or 4.55 mmol cm<sup>-3</sup>, which is 78% of the single-component isotherm uptake at 313 K and 0.15 bar. The

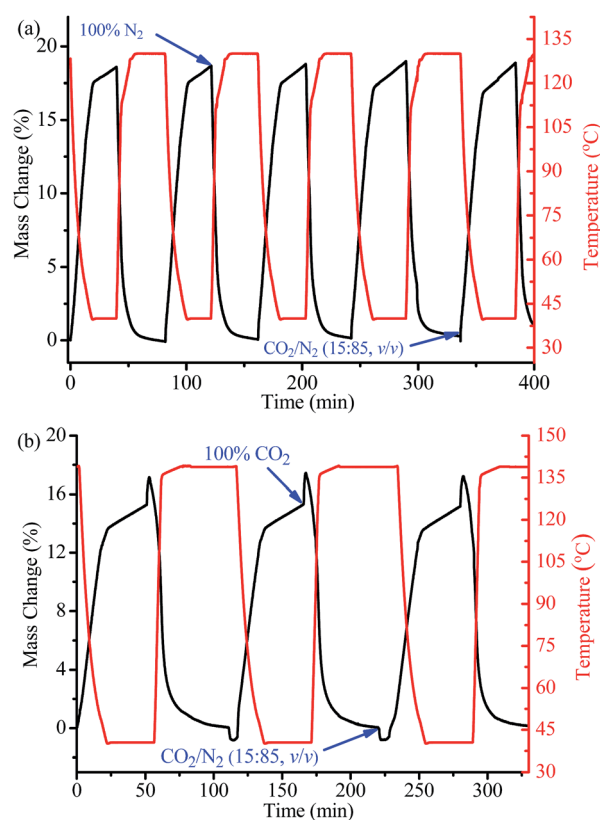


Fig. 4 Repeated adsorption–desorption kinetics for **1a** (a) between a 15 : 85 CO<sub>2</sub>/N<sub>2</sub> (v/v) flow at 313 K and a pure N<sub>2</sub> flow at 403 K and (b) between a 15 : 85 CO<sub>2</sub>/N<sub>2</sub> (v/v) flow at 313 K and a pure CO<sub>2</sub> flow at 413 K.





CO<sub>2</sub> uptake further increased to 18.7% (4.25 mmol g<sup>-1</sup> or 5.01 mmol cm<sup>-3</sup>, *i.e.*, 87% of the isotherm value) by maintaining the temperature at 313 K for an additional 23 min. Although the adsorption and desorption processes for **1a** were not allowed to reach equilibrium, as its relatively small pores are not beneficial for fast guest diffusion, the CO<sub>2</sub> adsorption/desorption working capacities obtained here are much higher than all the reported values for other materials under similar conditions and time scales (Table S4†). After 5 such adsorption–desorption cycles, there was nearly no change in the adsorption capacity, illustrating the high thermal stability and good CO<sub>2</sub> adsorption–desorption reversibility of **1a**, as indicated by the thermogravimetry-mass spectrometry analyses (Fig. S2†).

A pure temperature swing adsorption (TSA) process for **1a** was further carried out, and a working capacity of 3.52 mmol g<sup>-1</sup> or 4.15 mmol cm<sup>-3</sup> was obtained between the adsorption in a 15 : 85 CO<sub>2</sub>/N<sub>2</sub> (v/v) mixture at 313 K for 48 min and the desorption in a pure CO<sub>2</sub> flow at 413 K (optimized) for 54 min. It should be noted that, even though it was obtained in a relatively short time which did not allow the adsorption/desorption to reach equilibration, this working capacity is higher than previously reported values obtained in similar and longer adsorption/desorption times (Table S4†). To evaluate the regeneration energy for CO<sub>2</sub> desorption, the heat capacity of **1a** was quantified *via* differential scanning calorimetry. About -200 J g<sup>-1</sup> was evolved as the material was cooled from 413 to 313 K (Fig. S8†). With these data (see calculation method in the ESI†), approximately 3.02 MJ of energy would be required to regenerate 1 kg of CO<sub>2</sub> from **1a** (Fig. 4b).<sup>49,56</sup> The regeneration energy for CO<sub>2</sub> desorption from **1a** (3.02 MJ kg<sup>-1</sup>) is higher than those of mmeda-Mg<sub>2</sub>(dobpdc) (2.34 MJ kg<sup>-1</sup>) and state-of-the-art amine-based solutions (2.6 MJ kg<sup>-1</sup>), but lower than that of monoethanolamine (3.5 MJ kg<sup>-1</sup>).<sup>49</sup>

To simulate more practical CO<sub>2</sub> capture applications, we also measured the column breakthrough curves of **1a** using a binary 10 : 90 CO<sub>2</sub>/N<sub>2</sub> (v/v) mixture at 313 K and 1 bar. As shown in Fig. 5a and S9a,† the breakthrough point (outlet concentration > detection limit) of CO<sub>2</sub> was observed at a mixture injection amount of 44.3 mmol g<sup>-1</sup> or 52.2 mmol cm<sup>-3</sup>, exceeding all the reported values even measured at a lower temperature of 298 K (Table S5,† note that a lower temperature generally leads to a higher adsorption amount, at the same pressure), including 42.2 mmol g<sup>-1</sup> or 38.8 mmol cm<sup>-3</sup> for **1** and 25.5 mmol g<sup>-1</sup> or 50.2 mmol cm<sup>-3</sup> for the recently discovered microporous copper silicate SGU-29.<sup>7</sup> After that, the outlet CO<sub>2</sub> concentration rose gradually to the inlet value at 51.6 mmol g<sup>-1</sup> or 60.8 mmol cm<sup>-3</sup>. The capability of **1a** for trace CO<sub>2</sub> capture was investigated by column breakthrough tests using a 1 : 999 CO<sub>2</sub>/N<sub>2</sub> (*i.e.* 1000 ppm CO<sub>2</sub>) mixture at 298 K (Fig. 5b). The CO<sub>2</sub> breakthrough and saturation points were observed at 4210 and 4400 mmol g<sup>-1</sup>, or 4960 and 5180 mmol cm<sup>-3</sup>, respectively.

The CO<sub>2</sub> uptake of the adsorbent **1a** in the columns can be calculated (see ESI† for calculation method) as 4.89 and 4.22 mmol g<sup>-1</sup>, or 5.76 and 4.97 mmol cm<sup>-3</sup>, for the 10 : 90 and 1 : 999 CO<sub>2</sub>/N<sub>2</sub> mixtures, respectively, which are both *ca.* 98% of the values obtained from the single-component adsorption isotherms (4.97 mmol g<sup>-1</sup> or 5.85 mmol cm<sup>-3</sup> at 313 K and

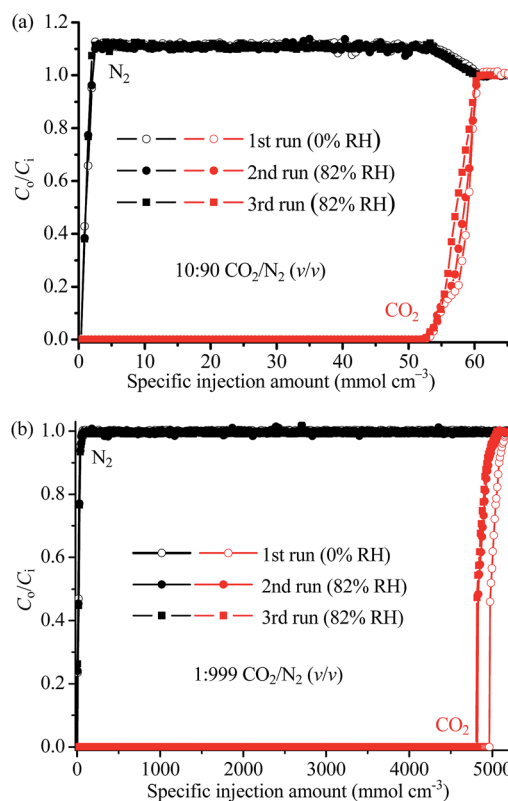


Fig. 5 Repeated breakthrough curves of the **1a** column operated at 1 bar. (a) 10 : 90 CO<sub>2</sub>/N<sub>2</sub> (v/v) at 313 K. (b) 1 : 999 CO<sub>2</sub>/N<sub>2</sub> (v/v) mixture at 298 K. Lines are drawn to guide eyes. C<sub>i</sub> and C<sub>o</sub> are the concentrations of each gas at the inlet and outlet, respectively.

0.1 bar, 4.30 mmol g<sup>-1</sup> or 5.07 mmol cm<sup>-3</sup> at 298 K and 1 mbar). Suffering from weak CO<sub>2</sub>/N<sub>2</sub> selectivity and/or low adsorption rate, many adsorbents show a much lower CO<sub>2</sub> uptake under mixed-gas conditions compared to single-component adsorption isotherms (Table S5†). The exceptionally high adsorption efficiency of **1a** implies that its CO<sub>2</sub> diffusion rate and CO<sub>2</sub>/N<sub>2</sub> selectivity are high enough under such mixed-gas breakthrough conditions.

We also tested the CO<sub>2</sub> capture performances and stabilities of the columns under high humidity.<sup>9–13</sup> After the breakthrough curves reached equilibrium under the above mentioned dry conditions (0% RH, RH = relative humidity), the columns were activated under He flow at 403 K, cooled down to the measurement temperature of 313 K, saturated using humid He (82% RH), and then switched to humid CO<sub>2</sub>/N<sub>2</sub> mixtures (82% RH, no change to other parameters) to start a new breakthrough experiment. Remarkably, the breakthrough curve for the humid 10 : 90 CO<sub>2</sub>/N<sub>2</sub> mixture almost overlapped with that measured at the dry condition (Fig. 5a). For the humid 1 : 999 CO<sub>2</sub>/N<sub>2</sub> mixture, the CO<sub>2</sub> breakthrough point slightly (2.6% compared with that observed for the dry mixture) shortened to 4100 mmol g<sup>-1</sup> or 4830 mmol cm<sup>-3</sup> (Fig. 5b and S9b†). These observations indicated that water has little effect on the CO<sub>2</sub> adsorption capacity of **1a**, even when the CO<sub>2</sub> concentration is extremely low. It should be noted that H<sub>2</sub>O can strongly compete with CO<sub>2</sub>



for the adsorption sites in most porous materials, and only a handful of examples have demonstrated such a high water-proof CO<sub>2</sub> capture performance (Table S5†).

As shown in Fig. 5, the same activation/breakthrough procedures were applied to the columns once again, and the breakthrough curves almost overlapped with each other, indicating the high stability of the adsorbent **1a** and the column even under high humidity. This also indicates that **1a** is particularly selective for CO<sub>2</sub> over water, and confirms that the grafted N<sub>2</sub>H<sub>4</sub> molecules were not replaced by water molecules.

## Conclusions

In summary, we demonstrated that the classical small-pore PCP [Mg<sub>2</sub>(dobdc)] can be modified by the shortest diamine N<sub>2</sub>H<sub>4</sub> to obtain a very powerful CO<sub>2</sub> adsorbent [Mg<sub>2</sub>(dobdc)(N<sub>2</sub>H<sub>4</sub>)<sub>1.8</sub>]. Thanks to the ultrahigh concentration of amine groups and the chemisorption of CO<sub>2</sub> associated with carbamic acid formation, this material exhibited exceptionally high CO<sub>2</sub> adsorption capacity at low pressures (from 0.4 mbar to 150 mbar) and a wide range of temperatures (from 298 to 328 K), well exceeding previous records. More importantly, these CO<sub>2</sub> adsorption performances can be maintained under mixed-gas kinetic conditions, even in the presence of high humidity.

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## Notes and references

- S. U. Rege, R. T. Yang and M. A. Buzanowski, *Chem. Eng. Sci.*, 2000, **55**, 4827–4838.
- S. U. Rege, R. T. Yang, K. Y. Qian and M. A. Buzanowski, *Chem. Eng. Sci.*, 2001, **56**, 2745–2759.
- J. C. Santos, F. D. Magalhaes and A. Mendes, *Ind. Eng. Chem. Res.*, 2008, **47**, 6197–6203.
- Y. Belmabkhout, R. Serna-Guerrero and A. Sayari, *Chem. Eng. Sci.*, 2010, **65**, 3695–3698.
- S. Sung and M. P. Suh, *J. Mater. Chem. A*, 2014, **2**, 13245–13249.
- Y. Lin, Q. Yan, C. Kong and L. Chen, *Sci. Rep.*, 2013, **3**, 1859.
- S. J. Datta, C. Khumnoon, Z. H. Lee, W. K. Moon, S. Doco, T. H. Nguyen, I. C. Hwang, D. Moon, P. Oleynikov, O. Terasaki and K. B. Yoon, *Science*, 2015, **350**, 302–306.
- A. Goepfert, M. Czaun, G. K. Surya Prakash and G. A. Olah, *Energy Environ. Sci.*, 2012, **5**, 7833–7853.
- A. Kumar, D. G. Madden, M. Lusi, K.-J. Chen, E. A. Daniels, T. Curtin, J. J. Perry and M. J. Zaworotko, *Angew. Chem., Int. Ed.*, 2015, **54**, 14372–14377.
- A. C. Kizzie, A. G. Wong-Foy and A. J. Matzger, *Langmuir*, 2011, **27**, 6368–6373.
- J. A. Mason, T. M. McDonald, T.-H. Bae, J. E. Bachman, K. Sumida, J. J. Dutton, S. S. Kaye and J. R. Long, *J. Am. Chem. Soc.*, 2015, **137**, 4787–4803.
- W. R. Lee, H. Jo, L.-M. Yang, H. Lee, D. W. Ryu, K. S. Lim, J. H. Song, D. Y. Min, S. S. Han, J. G. Seo, Y. K. Park, D. Moon and C. S. Hong, *Chem. Sci.*, 2015, **6**, 3697–3705.
- O. Shekhah, Y. Belmabkhout, Z. Chen, V. Guillerm, A. Cairns, K. Adil and M. Eddaoudi, *Nat. Commun.*, 2014, **5**, 4228.
- N. T. T. Nguyen, H. Furukawa, F. Gándara, H. T. Nguyen, K. E. Cordova and O. M. Yaghi, *Angew. Chem., Int. Ed.*, 2014, **53**, 10645–10648.
- S. Choi, T. Watanabe, T.-H. Bae, D. S. Sholl and C. W. Jones, *J. Phys. Chem. Lett.*, 2012, **3**, 1136–1141.
- P.-Q. Liao, A.-X. Zhu, W.-X. Zhang, J.-P. Zhang and X.-M. Chen, *Nat. Commun.*, 2015, **6**, 6350.
- P.-Q. Liao, W.-X. Zhang, J.-P. Zhang and X.-M. Chen, *Nat. Commun.*, 2015, **6**, 8697.
- P.-Q. Liao, H. Chen, D.-D. Zhou, S.-Y. Liu, C.-T. He, Z. Rui, H. Ji, J.-P. Zhang and X.-M. Chen, *Energy Environ. Sci.*, 2015, **8**, 1011–1016.
- P. K. Thallapally, R. K. Motkuri, C. A. Fernandez, B. P. McGrail and G. S. Behrooz, *Inorg. Chem.*, 2010, **49**, 4909–4915.
- J. Liu, J. Tian, P. K. Thallapally and B. P. McGrail, *J. Phys. Chem. C*, 2012, **116**, 9575–9581.
- J. Liu, P. K. Thallapally, B. P. McGrail, D. R. Brown and J. Liu, *Chem. Soc. Rev.*, 2012, **41**, 2308–2322.
- A. M. Fracaroli, H. Furukawa, M. Suzuki, M. Dodd, S. Okajima, F. Gandara, J. A. Reimer and O. M. Yaghi, *J. Am. Chem. Soc.*, 2014, **136**, 8863–8866.
- S. R. Caskey, A. G. Wong-Foy and A. J. Matzger, *J. Am. Chem. Soc.*, 2008, **130**, 10870–10871.
- B. Li, Z. Zhang, Y. Li, K. Yao, Y. Zhu, Z. Deng, F. Yang, X. Zhou, G. Li, H. Wu, N. Nijem, Y. J. Chabal, Z. Lai, Y. Han, Z. Shi, S. Feng and J. Li, *Angew. Chem., Int. Ed.*, 2012, **51**, 1412–1415.
- J.-R. Li, J. Yu, W. Lu, L.-B. Sun, J. Sculley, P. B. Balbuena and H.-C. Zhou, *Nat. Commun.*, 2013, **4**, 1538.
- S. M. Cohen, *Chem. Rev.*, 2012, **112**, 970–1000.
- Y. He, W. Zhou, R. Krishna and B. Chen, *Chem. Commun.*, 2012, **48**, 11813–11831.
- T. Fukushima, S. Horike, Y. Inubushi, K. Nakagawa, Y. Kubota, M. Takata and S. Kitagawa, *Angew. Chem., Int. Ed.*, 2010, **49**, 4820–4824.
- B. Zheng, J. Bai, J. Duan, L. Wojtas and M. J. Zaworotko, *J. Am. Chem. Soc.*, 2010, **133**, 748–751.
- R. Vaidhyanathan, S. S. Iremonger, G. K. Shimizu, P. G. Boyd, S. Alavi and T. K. Woo, *Science*, 2010, **330**, 650–653.
- M. Wriedt, J. P. Sculley, A. A. Yakovenko, Y. G. Ma, G. J. Halder, P. B. Balbuena and H. C. Zhou, *Angew. Chem., Int. Ed.*, 2012, **51**, 9804–9808.
- P.-Q. Liao, D.-D. Zhou, A.-X. Zhu, L. Jiang, R.-B. Lin, J.-P. Zhang and X.-M. Chen, *J. Am. Chem. Soc.*, 2012, **134**, 17380–17383.
- J. A. Johnson, S. Chen, T. C. Reeson, Y. S. Chen, X. C. Zeng and J. Zhang, *Chem.–Eur. J.*, 2014, **20**, 7632–7637.
- J. An, S. J. Geib and N. L. Rosi, *J. Am. Chem. Soc.*, 2009, **132**, 38–39.



- 35 T. Panda, P. Pachfule, Y. Chen, J. Jiang and R. Banerjee, *Chem. Commun.*, 2011, **47**, 2011–2013.
- 36 T. Devic, F. Salles, S. Bourrelly, B. Moulin, G. Maurin, P. Horcajada, C. Serre, A. Vimont, J.-C. Lavalley, H. Leclerc, G. Clet, M. Daturi, P. L. Llewellyn, Y. Filinchuk and G. Férey, *J. Mater. Chem.*, 2012, **22**, 10266–10273.
- 37 C. R. Wade and M. Dincă, *Dalton Trans.*, 2012, **41**, 7931–7938.
- 38 C. Montoro, E. Garcia, S. Calero, M. A. Perez-Fernandez, A. L. Lopez, E. Barea and J. A. R. Navarro, *J. Mater. Chem.*, 2012, **22**, 10155–10158.
- 39 H. J. Park and M. P. Suh, *Chem. Sci.*, 2013, **4**, 685–690.
- 40 L.-H. Xie and M. P. Suh, *Chem.–Eur. J.*, 2013, **19**, 11590–11597.
- 41 S. Xiang, Y. He, Z. Zhang, H. Wu, W. Zhou, R. Krishna and B. Chen, *Nat. Commun.*, 2012, **3**, 954.
- 42 P. Wang, Y. Lu Yi, Q. Liu and W.-Y. Sun, *Sci. Sin.: Chim.*, 2013, **43**, 1288–1296.
- 43 S. S. Nagarkar, A. K. Chaudhari and S. K. Ghosh, *Inorg. Chem.*, 2011, **51**, 572–576.
- 44 A. Khutia and C. Janiak, *Dalton Trans.*, 2014, **43**, 1338–1347.
- 45 L.-J. Li, P.-Q. Liao, C.-T. He, Y.-S. Wei, H.-L. Zhou, J.-M. Lin, X.-Y. Li and J.-P. Zhang, *J. Mater. Chem. A*, 2015, **3**, 21849–21855.
- 46 T. M. McDonald, D. M. D'Alessandro, R. Krishna and J. R. Long, *Chem. Sci.*, 2011, **2**, 2022–2028.
- 47 T. M. McDonald, J. A. Mason, X. Kong, E. D. Bloch, D. Gygi, A. Dani, V. Crocella, F. Giordanino, S. O. Odoh, W. S. Drisdell, B. Vlasisavljevich, A. L. Dzubak, R. Poloni, S. K. Schnell, N. Planas, K. Lee, T. Pascal, L. F. Wan, D. Prendergast, J. B. Neaton, B. Smit, J. B. Kortright, L. Gagliardi, S. Bordiga, J. A. Reimer and J. R. Long, *Nature*, 2015, **519**, 303–308.
- 48 W. R. ChangLee, S. Y. Hwang, D. W. Ryu, K. S. Lim, S. S. Han, D. Moon, J. Choi and C. S. Hong, *Energy Environ. Sci.*, 2014, **7**, 744–751.
- 49 T. M. McDonald, W. R. Lee, J. A. Mason, B. M. Wiers, C. S. Hong and J. R. Long, *J. Am. Chem. Soc.*, 2012, **134**, 7056–7065.
- 50 A. Demessence, D. M. D'Alessandro, M. L. Foo and J. R. Long, *J. Am. Chem. Soc.*, 2009, **131**, 8784–8786.
- 51 N. Planas, A. L. Dzubak, R. Poloni, L.-C. Lin, A. McManus, T. M. McDonald, J. B. Neaton, J. R. Long, B. Smit and L. Gagliardi, *J. Am. Chem. Soc.*, 2013, **135**, 7402–7405.
- 52 B. Lee, S. H. Kang, D. Kang, K. H. Lee, J. Cho, W. Nam, O. H. Han and N. H. Hur, *Chem. Commun.*, 2011, **47**, 11219–11221.
- 53 J. Emsley, *Chem. Soc. Rev.*, 1980, **9**, 91–124.
- 54 N. B. Cramer and C. N. Bowman, *J. Polym. Sci., Part A: Polym. Chem.*, 2001, **39**, 3311–3319.
- 55 A. Danon, P. C. Stair and E. Weitz, *J. Phys. Chem. C*, 2011, **115**, 11540–11549.
- 56 J. A. Mason, K. Sumida, Z. R. Herm, R. Krishna and J. R. Long, *Energy Environ. Sci.*, 2011, **4**, 3030–3040.

