

CrossMark  
click for updatesCite this: *Chem. Sci.*, 2016, 7, 4582

## A room temperature phosphorescence encoding [2]rotaxane molecular shuttle†

Xiang Ma,\* Jing Zhang, Jingjing Cao, Xuyang Yao, Tiantian Cao, Yifan Gong, Chunchang Zhao and He Tian

A novel bistable molecular shuttle, composed of a Pt(II) porphyrin-containing dibenzo[24]crown-8 macrocycle threaded onto two different recognition sites (secondary dialkylammonium ( $\text{NH}_2^+$ ) and 4,4'-bipyridinium ( $\text{Bpym}^{2+}$ ) units), and the anthracene (Anth) moiety as one terminal stopper, was synthesized by click chemistry. Its acid–base switchable shuttling could be addressed by both the room temperature phosphorescence (RTP) emission signals of the Pt(II) porphyrin moiety and the fluorescence emission of the Anth unit, as well as their lifetime changes. When the macrocycle was switched to be located on the  $\text{NH}_2^+$  site close to Anth, the Pt(II) porphyrin moiety exhibited strong RTP emission, excited in the Anth band at 370 nm. This was due to the distance-dependent efficient singlet energy transfer between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to a triplet state in Pt(II) porphyrin, while its RTP emission dramatically decreased when located on the  $\text{Bpym}^{2+}$  site far from the Anth unit. When excited in the porphyrin band at 402 nm, the RTP emission lifetimes changed obviously. This is the first rotaxane-type molecular shuttle whose shuttling has been encoded by RTP signals.

Received 19th February 2016

Accepted 25th March 2016

DOI: 10.1039/c6sc00769d

www.rsc.org/chemicalscience

## Introduction

Rotaxanes,<sup>1</sup> as one of the challenging topics of molecular machines, have attracted continued interest and attention. A variety of appealing rotaxanes have been fabricated over the recent years, with various functions or applications in such areas as switches,<sup>2</sup> logic gates,<sup>3</sup> artificial muscles,<sup>4</sup> molecular elevators,<sup>5</sup> molecular motors,<sup>6</sup> and so forth. The output signals of most rotaxane shuttles were observed by NMR, UV-Vis absorption, cyclic voltammetry, circular dichroism, fluorescence emission, and so forth.<sup>7</sup> Room temperature phosphorescence (RTP), with the advantages of longer emission wavelengths and lifetimes, larger 'Stokes' shifts, higher signal-to-noise ratios, quick response and low-cost detection *etc.*,<sup>8</sup> is another excellent method which is expected to be employed to address the shuttling and conformation of rotaxane shuttles.<sup>9</sup> Up to now, few rotaxane examples employing phosphorescence signal changes as the output have been reported.<sup>10</sup>

Porphyrins are important compounds in relation to electron and energy transfer processes. Stoddart<sup>11</sup> and Sauvage<sup>12</sup> first introduced the porphyrin moiety into rotaxane systems.

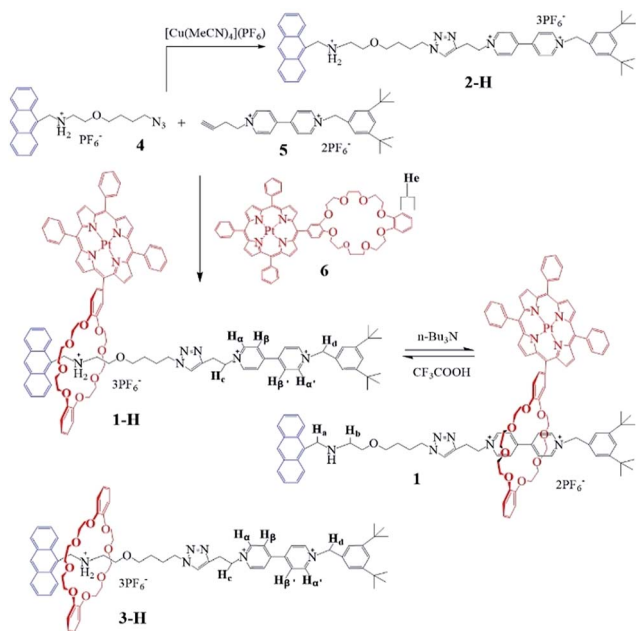
Subsequently, several rotaxanes and catenanes containing porphyrins have been investigated.<sup>13</sup> Particularly, Pt(II) porphyrins have been reported to have characteristic long wave RTP emissions.<sup>14</sup> Therefore, incorporating a Pt(II) porphyrin moiety into a rotaxane might engender new signals, *i.e.* RTP emission and lifetime changes, to address the shuttling and conformation of rotaxanes.

Herein, we designed and synthesized novel acid–base controlled molecular shuttles<sup>15</sup> based on switchable [2]rotaxanes **1** (Scheme 1), with Pt(II) porphyrin-based dibenzo[24]crown-8 as the host macrocycle threaded onto two different recognition sites (secondary dialkylammonium ( $\text{NH}_2^+$ ) and 4,4'-bipyridinium ( $\text{Bpym}^{2+}$ ) units), and the anthracene (Anth) moiety as one terminal stopper. In this [2]rotaxane, reversible motions of the macrocycle between the two stations could be controlled by the protonation and deprotonation of the  $\text{NH}_2^+$  recognition site, as previously observed in the related system.<sup>16</sup> These motions affect the interactions between the porphyrin moiety on the macrocycle and the Anth stopper. Upon protonation of the  $\text{NH}_2^+$  recognition site, the macrocycle is located at said site, and the porphyrin unit nearby could also interact with the Anth unit by energy transfer. Deprotonation of the  $\text{NH}_2^+$  center could engender the displacement of the macrocycle to the  $\text{Bpym}^{2+}$  station, and destroy the energy transfer process due to the elongated distance between the Anth and porphyrin units. These switching processes could be investigated using conventional methods, such as <sup>1</sup>H NMR spectroscopy and fluorescence spectroscopy. Gratifyingly, after the metallization

Key Laboratory for Advanced Materials and Institute of Fine Chemicals, East China University of Science & Technology, Shanghai 200237, P. R. China. E-mail: maxiang@ecust.edu.cn

† Electronic supplementary information (ESI) available: Materials, general procedures and synthesis; additional characterization data of compounds, control experiments and corresponding discussions. See DOI: 10.1039/c6sc00769d





Scheme 1 Chemical structures of the macrocycle **6**, and the [2]rotaxanes **1-H**, **1** and **3-H**.

of the porphyrin core with Pt(II) salt, the shuttling of rotaxane **1-H** accompanied by the macrocycle **6** staying between the two recognition sites could also be monitored by obvious RTP signal changes, which makes this the first [2]rotaxane molecular shuttle with switchable RTP output addresses so far.

## Results and discussions

The [2]rotaxane **1-H** was synthesized conveniently *via* a copper(I) catalyzed azide-alkyne 'click' cycloaddition. As shown in Scheme 1, treating equal equivalents of **4** and **5** with 2.0 equivalents of macrocycle **6** in the absence of light and oxygen led to the formation of the [2]rotaxane **1-H** with a relatively low yield of 13%, which might be due to the relatively lower binding constant of the DB24C8 decorated with Pt(II) porphyrin (macrocycle **6**) with the  $\text{NH}_2^+$  center of compound **4** and steric hindrance, besides the obvious product loss during purification by chromatography on silica gel. The triazole ring and the oxygen atom of the dumbbell were used as long linkers not only to increase the distance between the two recognition sites, preventing energy transfer from the anthracene moiety to the 4,4'-bipyridium moiety as much as possible, but also to ensure the stability of the interlocking states by avoiding adverse factors like steric hindrance. Another reference [2]rotaxane **3-H** was constructed by replacing macrocycle **6** with porphyrin-free dibenzo[24]crown-8 (DB24C8). The addition of 1.0 equiv. of  $n\text{-Bu}_3\text{N}^{16,17}$  to the solution of [2]rotaxane **1-H** results in the deprotonated form of **1**. Reprotonation can be performed by the addition of a slight molar excess of  $\text{CF}_3\text{COOH}$ .

The [2]rotaxane formation and its reversible switching process were monitored by  $^1\text{H}$  NMR spectroscopy. As shown in Fig. 1, the dumbbell and macrocycle components of **1-H/1** are

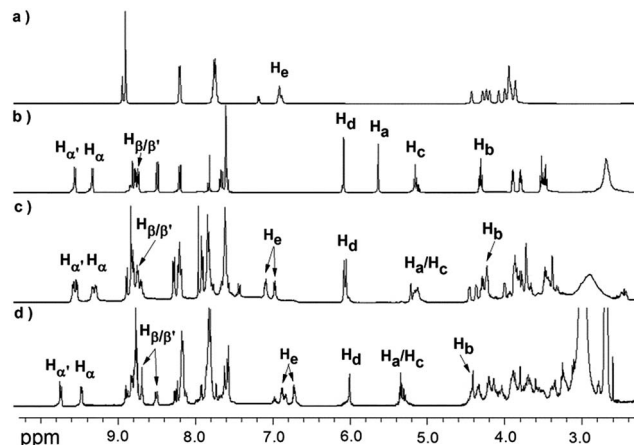


Fig. 1 Partial  $^1\text{H}$  NMR spectra (500 MHz in  $\text{CD}_3\text{COCD}_3$  at 298 K) of: (a) the macrocycle **6**; (b) the dumbbell compound **2-H**; (c) the [2]rotaxane **1-H**; and (d) the deprotonated [2]rotaxane **1**.

present in equimolar amounts and are interlocked with each other. The spectrum (Fig. 1c) demonstrates that the macrocycle **6** binds selectively with the  $\text{NH}_2^+$  center, since the characteristic resonances  $\text{H}_\alpha$  and  $\text{H}_\beta$  of the dumbbell **2-H**, at  $\delta$  5.59 and 4.39 ppm (Fig. 1b), for the four  $\text{CH}_2$  protons adjacent to the  $\text{NH}_2^+$  group experienced upfield shifts,  $\text{H}_\alpha$  at  $\delta$  5.24 ppm ( $\Delta\delta = 0.35$  ppm) and  $\text{H}_\beta$  at 4.22 ppm ( $\Delta\delta = 0.17$  ppm), respectively, upon rotaxane (**1-H**) formation. Furthermore, the aromatic protons  $\text{H}_\epsilon$  of the macrocycle **6** were split into two peaks ( $\delta$  7.15 and 7.13 ppm) after complexation.

After the addition of  $n\text{-Bu}_3\text{N}$ , the  $^1\text{H}$  NMR spectrum of **1** shows dramatic chemical shift changes with respect to that of **1-H**. The resonances ( $\delta$  5.24 and 4.22 ppm) of the  $\text{CH}_2\text{-NH}_2^+\text{-CH}_2$  protons disappeared and new signals ( $\delta$  5.38 and 4.40 ppm, Fig. 1d), assigned to  $\text{CH}_2\text{-NH-CH}_2$  protons, appeared. Moreover, the macrocycle **6** glides towards the  $\text{Bpym}^{2+}$  unit, resulting in the obvious changes of the proton resonance values of  $\text{Bpym}^{2+}$  ( $\delta$  [ $\Delta\delta$ ] for  $\text{H}_{\alpha'}$ ,  $\text{H}_\alpha$ ,  $\text{H}_{\beta'}$  and  $\text{H}_\beta$  (ppm): 9.56 [+0.22], 9.28 [+0.18], 8.70 [−0.10] and 8.51 [−0.28], respectively). Upon addition of  $\text{CF}_3\text{COOH}$ , the original  $^1\text{H}$  NMR spectrum was almost regenerated, demonstrating that the switching process was fully reversible *via* de-/reprotonation (Fig. S5†).

The photophysical and photochemical properties of the shuttling processes of rotaxane **1-H/1** were thoroughly investigated (Table S1†). Pt(II) porphyrin-based dibenzo[24]crown-8 **6** displayed relatively weak fluorescence emission bands at 650 and 715 nm in air equilibrated  $\text{CH}_3\text{CN}$  (Fig. S7†). Notably, no RTP was exhibited even in a deaerated organic solution, but strong RTP signals were observed in the medium containing the anionic surfactant sodium dodecyl sulfate (SDS) and  $\text{Na}_2\text{SO}_3$  for removing oxygen.<sup>18</sup> The characteristic RTP emission peaks of **6** were exhibited at 667 and 736 nm (inset, Fig. S7†) when excited in the porphyrin band at 402 nm, and experienced a redshift compared with the fluorescence emission peak (Fig. S7†). Upon forming the [2]rotaxane **1-H**, the fluorescence emission was almost quenched,<sup>19</sup> while its RTP emission showed a slight increase due to the movement restriction caused by the



macrocycle **6** in the rotaxane system when excited at 402 nm (Fig. S7 and S8†).

The switching of the [2]rotaxanes between **1** and **1-H** was accompanied by obvious RTP signal changes. As shown in Fig. 2, both RTP spectra of the [2]rotaxanes **1-H** and **1** showed intense RTP emissions in the region above 600 nm, upon excitation in the porphyrin band at 402 nm, which were characteristic of the Pt(II) porphyrin moieties. Deprotonation induced very small changes in this band, and the intensity decrease of the deprotonated [2]rotaxane **1** could be due to the quenching effect of the Bpym<sup>2+</sup> unit through charge transfer. When the excitation was performed in the Anth band at 370 nm, the RTP spectra exhibited dramatic changes (emission intensities increased by more than 1200%, Fig. 2). Such a porphyrin-based phosphorescence band was still present in **1-H**, but with obvious decreases in intensity of **1**. The RTP lifetime of **1-H** excited at 370 nm was found to be about 57.1 μs (Fig. S12†). The reversibility of the switching process of the rotaxanes **1-H/1** was proven to be favourable *via* testing the RTP intensity change upon addition of *n*-Bu<sub>3</sub>N and CF<sub>3</sub>COOH alternately (Fig. 2, inset). Moreover, Table 1 lists the RTP lifetimes of all of the Pt(II) porphyrin-containing species. The [2]rotaxane **1** exhibited a shorter RTP lifetime than **1-H**, which resulted from the electron transfer to the Bpym<sup>2+</sup> moiety in [2]rotaxane **1** upon deprotonation.

Electrochemical measurements of **1-H**, **1** and the reference [2]rotaxanes **3-H** and **3** using the cyclic voltammetry technique were carried out to confirm the possible electron transfer from

the Pt(II) porphyrin unit to the bipyridinium moiety in [2]rotaxane **1** (Fig. 3). For the DB24C8-based [2]rotaxanes **3-H** and **3**, it was clear that the reduction waves of the bipyridinium unit were displaced to more negative positions as a result of being engaged in a charge transfer interaction with the electron donating macrocycle.<sup>20</sup> However, for **1-H** and **1**, the positions of the reduction waves of the bipyridinium unit were almost unaltered although it was proven by NMR that the macrocycle also moved towards the bipyridinium unit when **1-H** was turned into **1**. For unperturbed bipyridinium reduction potentials, there are few electrochemical experiments in the literature, although several similar systems containing Zn-porphyrin have been reported.<sup>20b</sup> A probable reason for the unaltered reduction potentials of bipyridinium was the electron transfer from the Pt(II) porphyrin unit to the bipyridinium moiety, since the presence of an electron-donor moiety usually leads to a less negative potential value for the first and second reductions of the bipyridium unit.<sup>20a</sup>

Addition of the base (*n*-Bu<sub>3</sub>N) made the RTP emission band of [2]rotaxane **1** almost quenched (Fig. 2). This behaviour cannot be simply assigned to the electron transfer to the Bpym<sup>2+</sup> unit; it should be due to the interaction between the Anth unit and the nearby porphyrin moiety of the macrocycle. As in the control experiment, both the 1 : 1 mixture solutions of the macrocycle **6** with dumbbell **2-H** and of **6** with **2** (after the addition of *n*-Bu<sub>3</sub>N) showed no obvious RTP emission bands

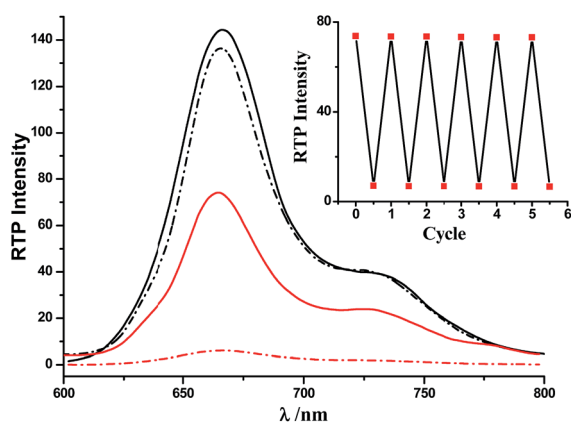


Fig. 2 RTP emission spectra of **1-H** (solid line) and **1** (dashed line). Excitation was performed in the Anth band at 370 nm (red, the RTP lifetime of **1-H** was found to be about 57.1 μs, Fig. S12†) and the Pt(II) porphyrin band at 402 nm (black, for the lifetimes see Table 1).  $4.0 \times 10^{-6}$  mol L<sup>-1</sup>. Inset: RTP intensity changes at 667 nm after addition of *n*-Bu<sub>3</sub>N and CF<sub>3</sub>COOH to **1-H** and **1**, over six cycles ( $\lambda_{\text{ex}}$  = 370 nm).

Table 1 RTP emission bands and lifetimes of all of the Pt(II) porphyrin-containing species

Compound	<b>6</b>	<b>1-H</b>	<b>1</b>
$\lambda_{\text{ex}}/\lambda_{\text{em}}$ (nm)	408/667	403/665	403/666
$I_{\text{RTP}}$	169.18	144.40	136.30
$\tau$ (μs)	54.5	48.6	42.9

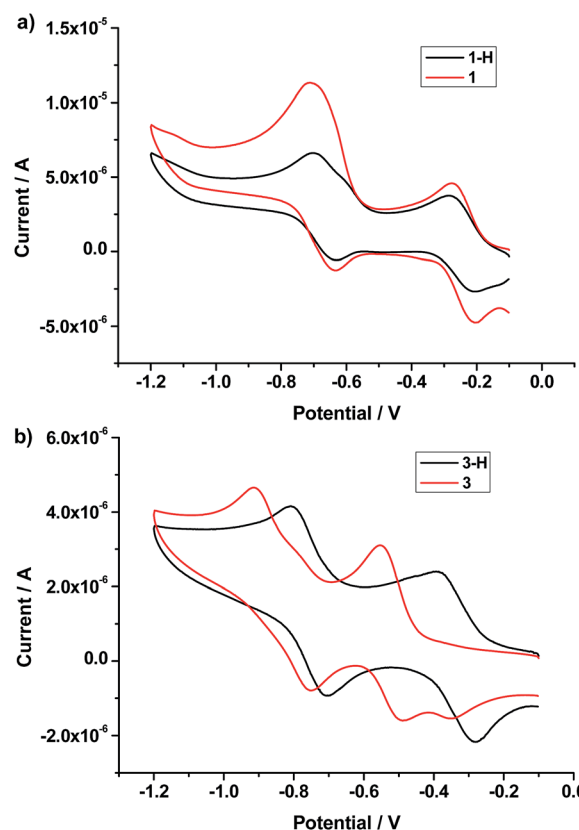


Fig. 3 Cyclic voltammograms of (a) **1-H** and **1** (after the addition of tributylamine); and (b) DB24C8-based [2]rotaxanes **3-H** and **3** (after the addition of tributylamine) in acetonitrile at 298 K.



when excited in the Anth band by light at 370 nm (Fig. S17 and S18†). Upon deprotonation, the intensity of the fluorescence emission band of the Anth unit of the dumbbell **2**, shown in Fig. 4, decreases. This is attributed to the electron transfer interaction between the Anth moiety and the amino group.<sup>20</sup> Conversely, the intensity of the Anth bands for the [2]rotaxane **1** experiences a dramatic increase with respect to that of **1-H**. Consequently, both the increase of the porphyrin based RTP emission and the decrease of the Anth based fluorescence emission for **1-H** compared with those of its deprotonated species should be attributed to the efficient singlet energy transfer process between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to triplet state in Pt(II) porphyrin (Fig. 5).

The quantum yields for the various RTP emission processes were also investigated, as shown in Table 2. The maximum absorption peak of the Pt(II) porphyrin-containing macrocycle **6** is at about 402 nm while the absorption at around 370 nm is very weak (Fig. S6†). Thus the quantum yields of **6**, **6** + base and [2]rotaxane **1-H** excited at 402 nm were relatively higher than the ones excited at 370 nm. After forming [2]rotaxane, the quantum yield of **1-H** (0.16) was slightly increased with respect to that of the host macrocycle **6** due to the efficient singlet energy transfer process between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to triplet state in Pt(II) porphyrin. While the quantum yields of [2]rotaxane **1** (**1-H** + base) excited at both 402 nm and 370 nm were too low to quantify, since the 4,4'-bipyridinium (Bpym<sup>2+</sup>) units in **1** had some quenching effects on the emission of the porphyrin moiety of **6**. Also, the distance dependent energy transfer efficiency from the Anth to porphyrin units decreased when the macrocycle **6** was located on the Bpym<sup>2+</sup> site far from the Anth moiety.

The fluorescence lifetime of the excited Anth unit in the [2]rotaxane **1-H** was found to be 2.79 ns (Fig. S13†). It was shorter than that of the corresponding dumbbell **2-H** (4.13 ns,

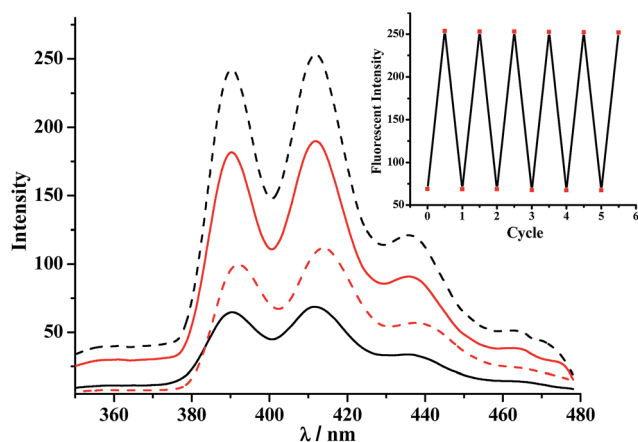


Fig. 4 Fluorescence spectra of the [2]rotaxanes **1-H/1** (black, solid/dash lines) and the dumbbell **2-H/2** (red, solid/dash lines). Excitation was performed in the Anth band at 370 nm,  $1.0 \times 10^{-5}$  mol L<sup>-1</sup>. Inset: fluorescence intensity changes at 412 nm after addition of *n*-Bu<sub>3</sub>N and CF<sub>3</sub>COOH to **1-H** and **1**, over six cycles ( $\lambda_{\text{ex}}$  = 370 nm).

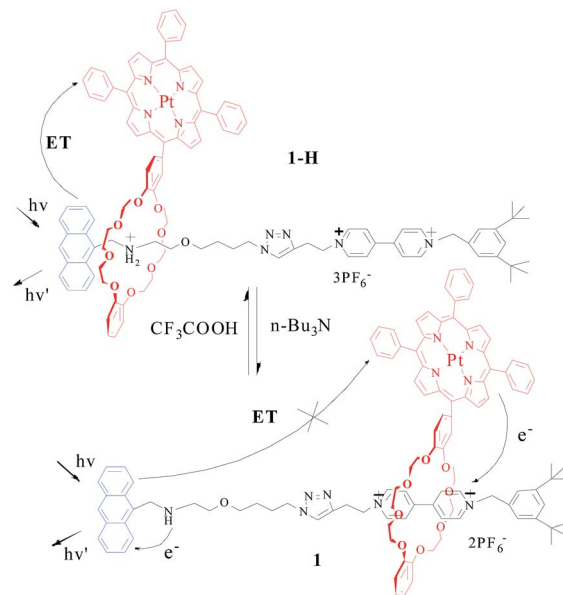


Fig. 5 Photoinduced charge and energy transfer processes in [2]rotaxanes **1-H/1** (singlet energy transfer process between the Anth and porphyrin moieties followed by intersystem crossing from a singlet to triplet state in Pt(II) porphyrin).

Table 2 RTP quantum yields for the processes involving the Pt(II) porphyrin-containing macrocycle **6**

Compound	Quantum yields (excited at 370 nm)	Quantum yields (excited at 402 nm)
<b>6</b>	0.14	2.1
<b>6</b> + base	0.10	2.2
<b>1-H</b>	0.16	1.2
<b>1</b> ( <b>1-H</b> + base) base: tributylamine	N/A	N/A

Fig. S14†), which was attributed to the energy transfer from Anth to Pt(II) porphyrin moieties in **1-H** (Fig. S19–S21† and discussion therein). Upon the addition of base, the lifetime of the excited Anth unit in **1** increased compared with **1-H** and reached up to 3.79 ns (Fig. S13†), which was due to the break up of the energy transfer process from the Anth to Pt(II) porphyrin units owing to the longer distance between them (Fig. S20†). The emission spectrum showed similarities with that of the original [2]rotaxane **1-H** upon addition of an equivalent of CF<sub>3</sub>COOH to its solution, indicating that the macrocycle was situated on the NH<sub>2</sub><sup>+</sup> center once again. The fluorescence emission intensity changes of [2]rotaxane **1** (at 412 nm) also showed good reversibility after alternating protonation/deprotonation for several cycles, indicating that its shuttling was totally reversible (Fig. 4, inset).

## Conclusions

In summary, a novel bistable molecular shuttle, based on an acid–base switchable [2]rotaxane in which the Pt(II) porphyrin–



containing macrocycle could be stimulated to shuttle back and forth between the two recognition sites, was constructed, whose shuttling process was accompanied by reversible changes both in the RTP emission (both intensities and lifetimes) of the Pt(II) porphyrin moiety and the fluorescence emission of the Anth unit. It should be noticed that this is the first example of a rotaxane-type molecular shuttle whose shuttling could be encoded by RTP signals. The concept and advance in this work, employing RTP signals to address the shuttling of a rotaxane, supply a novel method to encode a molecular shuttle and pave the way for designing molecular machines with multi-responses and multi-functions.

## Experimental section

### Synthesis of 1-H

A solution of the azide **4** (56 mg, 0.11 mmol) and the ring **6** (200 mg, 0.17 mmol) in 10 mL of  $\text{CHCl}_3$  was stirred for 24 h at room temperature under Ar. Then the BIPY<sup>2+</sup> derivative **5** (71 mg, 0.1 mmol) and  $\text{Cu}(\text{MeCN})_4\text{PF}_6$  (20 mg, 0.05 mmol) in 10 mL of dry  $\text{Me}_2\text{CO}$  were added, and the whole solution was stirred for another 3 days. After the solvent was evaporated, the resulting reddish solid was washed with  $\text{H}_2\text{O}$ , air-dried and purified by column chromatography on silica gel ( $\text{Me}_2\text{CO}$  to 2%  $\text{NH}_4\text{PF}_6/\text{Me}_2\text{CO}$ ) to obtain the [2]rotaxane **1-H** (31 mg, 13%). <sup>1</sup>H NMR (500 MHz,  $(\text{CD}_3)_2\text{CO}$ ):  $\delta_{\text{H}}$  (ppm) 9.56 (d, 2H, <sup>3</sup>*J* = 5.9 Hz), 9.31 (d, 2H, <sup>3</sup>*J* = 5.5 Hz), 8.86 (m, 8H), 8.74 (d, 2H, <sup>3</sup>*J* = 6.2 Hz), 8.50 (d, 2H, <sup>3</sup>*J* = 6.5 Hz), 8.58 (d, 3H, <sup>3</sup>*J* = 9.0 Hz), 8.23 (t, 6H, <sup>3</sup>*J* = 7.3 Hz), 8.11 (d, 2H, <sup>3</sup>*J* = 8.1 Hz), 7.87 (m, 11H), 7.62 (s, 3H), 7.53 (m, 4H), 7.15 (d, 2H, <sup>3</sup>*J* = 6.0 Hz), 7.13 (d, 2H, <sup>3</sup>*J* = 6.0 Hz), 6.09 (s, 2H), 5.59 (s, 2H), 5.24 (t, 2H, <sup>3</sup>*J* = 6.5 Hz), 4.70–3.70 (m, 29H), 3.60 (t, 3H, <sup>3</sup>*J* = 6.0 Hz), 3.39 (m, 4H), 2.64 (s, 2H), 2.15 (s, 2H), 1.32 (s, 18H). <sup>13</sup>C NMR (100 MHz,  $\text{CD}_3\text{OD}$ ):  $\delta_{\text{C}}$  (ppm), 153.92, 151.77, 151.50, 147.39, 146.85, 134.40, 133.58, 132.80, 132.24, 131.65, 130.62, 128.85, 128.61, 128.24, 127.72, 126.67, 125.36, 124.96, 124.32, 78.92, 75.18, 71.47, 67.46, 66.56, 62.30, 61.23, 58.31, 56.32, 51.31, 35.94, 31.76, 28.10, 26.02, 23.71, 21.87, 18.38, 14.44. HRMS (ESI): *m/z* [**1-H-PF<sub>6</sub>**]<sup>+</sup> calc. for  $[\text{C}_{112}\text{H}_{115}\text{N}_{10}\text{F}_{12}\text{O}_9\text{P}_2\text{Pt}]^+$  2228.8, found 2229.2. MALDI-TOF-MS: *m/z* [**1-H-3PF<sub>6</sub>**]<sup>3+</sup> calc. for  $\text{C}_{112}\text{H}_{115}\text{N}_{10}\text{O}_9\text{Pt}$ , 1938.85, found 1938.78.

### Synthesis of 2-H

A solution of the azide **4** (197.6 mg, 0.4 mmol), the BIPY<sup>2+</sup> derivative **5** (140 mg, 0.2 mmol), and  $\text{Cu}(\text{MeCN})_4\text{PF}_6$  (33.2 mg, 0.1 mmol) in dry  $\text{Me}_2\text{CO}$  (10 mL) was stirred for 48 h at room temperature under Ar. After the solvent was evaporated, the resulting reddish solid was washed with  $\text{H}_2\text{O}$ , air dried and purified by column chromatography on silica gel ( $\text{Me}_2\text{CO}$  to 2%  $\text{NH}_4\text{PF}_6/\text{Me}_2\text{CO}$ ) to obtain the thread **2-H** (73 mg, 30%). <sup>1</sup>H NMR (500 MHz,  $(\text{CD}_3)_2\text{CO}$ ):  $\delta_{\text{H}}$  (ppm) 9.49 (d, 2H, <sup>3</sup>*J* = 5.9 Hz), 9.31 (d, 2H, <sup>3</sup>*J* = 5.5 Hz), 8.74 (d, 2H, <sup>3</sup>*J* = 6.2 Hz), 8.50 (d, 2H, <sup>3</sup>*J* = 6.5 Hz), 8.58 (d, 3H, <sup>3</sup>*J* = 9.0 Hz), 8.11 (d, 2H, <sup>3</sup>*J* = 8.1 Hz), 7.62 (s, 3H), 7.53 (m, 4H), 6.09 (s, 2H), 5.59 (s, 2H), 5.24 (t, 2H, <sup>3</sup>*J* = 6.5 Hz), 4.39 (s, 2H), 3.84 (s, 1H), 3.60 (t, 3H, <sup>3</sup>*J* = 6.0 Hz), 3.39 (m, 4H), 2.64 (s, 2H), 2.15 (s, 2H), 1.32 (s, 18H). <sup>13</sup>C NMR (100 MHz,  $\text{CD}_3\text{OD}$ ):  $\delta_{\text{C}}$  (ppm) 151.62, 148.78, 146.05, 145.58, 141.63,

133.50, 130.85, 130.73, 129.13, 127.18, 127.06, 126.46, 125.58, 124.37, 123.58, 123.34, 123.17, 69.86, 65.45, 64.06, 59.96, 49.25, 42.37, 34.79, 31.22, 26.71, 26.07. HRMS (ESI): *m/z* [**2-H-PF<sub>6</sub>**]<sup>+</sup> calc. for  $[\text{C}_{50}\text{H}_{61}\text{F}_{12}\text{OP}_2]^+$  1051.4, found 1051.3.

### Synthesis of 3-H

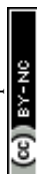
A solution of the azide **4** (200 mg, 0.4 mmol) and DB24C8 (225 mg, 0.5 mmol) in  $\text{CHCl}_3$  (30 mL) was stirred for 48 h at room temperature under argon. Then the BIPY<sup>2+</sup> derivative **5** (140 mg, 0.2 mmol) and the catalyst  $\text{Cu}(\text{CH}_3\text{CN})_4\text{PF}_6$  (150 mg, 0.4 mmol) were added and the whole solution was stirred for another 2 days. After the solvent was removed under reduced pressure, the resulting reddish solid was washed with  $\text{H}_2\text{O}$ , air dried and purified by column chromatography on silica gel ( $\text{Me}_2\text{CO}$  to 2%  $\text{NH}_4\text{PF}_6/\text{Me}_2\text{CO}$ ) to afford the DB24C8 based [2]rotaxane **3-H** (62 mg, 19%). <sup>1</sup>H NMR (400 MHz,  $\text{DMSO-d}_6$ ):  $\delta_{\text{H}}$  (ppm) 9.61 (d, 2H, *J* = 6.0 Hz), 9.42 (d, 2H, *J* = 6.0 Hz), 8.77 (s, 1H), 8.57–8.48 (m, 4H), 8.36 (d, 2H, *J* = 6.0 Hz), 8.17 (d, 2H, *J* = 8.0 Hz), 7.73 (s, 1H), 7.69–7.54 (m, 6H), 7.46 (s, 1H), 6.75–6.61 (m, 4H), 6.51–6.31 (m, 4H), 5.89 (s, 2H), 5.21 (s, 2H), 5.07 (t, 2H, *J* = 8.0 Hz), 4.32 (t, 2H, *J* = 6.4 Hz), 4.11–3.99 (m, 4H), 3.96 (m, 4H), 3.83–3.66 (m, 10H), 3.64–3.41 (m, 12H), 3.16 (d, 2H, *J* = 4.8 Hz), 1.93–1.78 (m, 2H), 1.59–1.48 (m, 2H), 1.30 (s, 18H). <sup>13</sup>C NMR (100 MHz,  $\text{DMSO-d}_6$ ):  $\delta_{\text{C}}$  (ppm) 151.63, 146.77, 146.54, 145.24, 142.89, 133.78, 130.95, 130.69, 129.13, 127.00, 126.59, 125.54, 124.39, 123.81, 123.41, 123.11, 122.25, 120.50, 111.98, 70.22, 69.89, 69.47, 67.78, 49.14, 46.86, 42.56, 34.80, 31.22, 26.71, 26.14, 25.07, 22.55. HRMS (ESI): *m/z* [**3-H-3PF<sub>6</sub>**]<sup>3+</sup> calc. for  $[\text{C}_{74}\text{H}_{93}\text{N}_6\text{O}_9]^+$  1209.69, found 1209.64.

## Acknowledgements

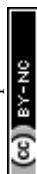
We acknowledge the financial support from the Programme of Introducing Talents of Discipline to Universities (B16017), NSFC (21421004, 21476075 & 21272072), the National Basic Research 973 Program and the Fundamental Research Funds for the Central Universities. We also thank Profs D. Qu and Q. Wang for their helpful discussion.

## Notes and references

- (a) D. Leigh, M. A. F. Morales, E. M. Pérez, J. K. Y. Wong, C. G. Saiz, A. M. Z. Slawin, A. J. Carmichael, D. M. Haddleton, A. M. Brouwer, W. J. Buma, G. W. H. Worpel, S. León and F. Zerbetto, *Angew. Chem., Int. Ed.*, 2005, **44**, 3062; (b) H. Tian and Q.-C. Wang, *Chem. Soc. Rev.*, 2006, **35**, 361; (c) E. R. Kay, D. A. Leigh and F. Zerbetto, *Angew. Chem., Int. Ed.*, 2007, **46**, 72; (d) V. Balzani, A. Credi and M. Venturi, *Chem. Soc. Rev.*, 2009, **38**, 1542; (e) X. Ma and H. Tian, *Chem. Soc. Rev.*, 2010, **39**, 70; (f) Z.-J. Zhang, H.-Y. Zhang, H. Wang and Y. Liu, *Angew. Chem., Int. Ed.*, 2011, **50**, 10834; (g) N. D. Suhan, S. J. Loeb and S. H. Eichhorn, *J. Am. Chem. Soc.*, 2013, **135**, 400; (h) C. Ke, R. A. Smaldone, T. Kikuchi, H. Li, A. P. Davis and J. F. Stoddart, *Angew. Chem., Int. Ed.*, 2013, **52**, 381; (i) J. Winn, A. Pinczewski and S. M. Goldup, *J. Am. Chem.*



- Soc.*, 2013, **135**, 13318; (j) Y.-H. Lin, C.-C. Lai, Y.-H. Liu, S.-M. Peng and S.-H. Chiu, *Angew. Chem., Int. Ed.*, 2013, **52**, 10231; (k) M. J. Langton and P. D. Beer, *Acc. Chem. Res.*, 2014, **47**, 1935; (l) S. Chao, C. Romuald, K. F. Marotte, C. Clavel and F. Coutrot, *Angew. Chem., Int. Ed.*, 2014, **27**, 6914; (m) Y.-D. Yang, C.-C. Fan, B. M. Rambo, H.-Y. Gong, L.-J. Xu, J.-F. Xiang and J. L. Sessler, *J. Am. Chem. Soc.*, 2015, **137**, 12966; (n) A. Fernandez, E. M. Pineda, C. A. Muryn, S. Sproules, F. Moro, G. A. Timco, E. J. L. McInnes and R. E. P. Winpenny, *Angew. Chem., Int. Ed.*, 2015, **127**, 11008; (o) M. Xue, Y. Yang, X. Chi, X. Yan and F. Huang, *Chem. Rev.*, 2015, **115**, 7398.
- 2 (a) C. P. Collier, G. Mattersteig, E. W. Wang, Y. Luo, K. Beverly, J. Sampaio, F. M. Raymo, J. F. Stoddart and J. R. Heath, *Science*, 2000, **289**, 1172; (b) V. Blanco, D. A. Leigh, U. Lewandowska, B. Lewandowski and V. Marcos, *J. Am. Chem. Soc.*, 2014, **136**, 15775; (c) A. Caballero, L. Swan, F. Zapata and P. D. Beer, *Angew. Chem., Int. Ed.*, 2014, **53**, 11854; (d) M. Woźny, J. Pawłowska, A. Osior, P. Świder, R. Bilewicz and B. Korybut-Daszkiewicz, *Chem. Sci.*, 2014, **5**, 2836; (e) P. G. Young, K. Hirose and Y. Tobe, *J. Am. Chem. Soc.*, 2014, **136**, 7899; (f) E. Moulin and N. Giuseppone, *Nat. Nanotechnol.*, 2014, **9**, 331.
- 3 (a) D.-H. Qu, Q.-C. Wang and H. Tian, *Angew. Chem., Int. Ed.*, 2005, **44**, 5296; (b) D.-H. Qu, Q.-C. Wang and H. Tian, *Adv. Mater.*, 2006, **18**, 2035; (c) S. Kou, H. N. Lee, D. van Noort, K. M. K. Swamy, S. H. Kim, J. H. Soh, K.-M. Lee, S.-W. Nam, J. Yoon and S. Park, *Angew. Chem., Int. Ed.*, 2008, **47**, 872; (d) F. Lohmann, J. Weigandt, J. Valero and M. Famulok, *Angew. Chem., Int. Ed.*, 2014, **53**, 10372; (e) A. Wolf, E. Moulin, J.-J. Cid, A. Goujon, G. Du, E. Busseron, G. Fuks and N. Giuseppone, *Chem. Commun.*, 2015, **51**, 4212.
- 4 (a) Y. Liu, A. H. Flood, P. A. Bonvallet, S. A. Vigon and J. F. Stoddart, *J. Am. Chem. Soc.*, 2005, **127**, 9745; (b) R. E. Dawson, S. F. Lincoln and C. J. Easton, *Chem. Commun.*, 2008, 3980; (c) C. Gao, X. Ma, Q. Zhang, Q. Wang, D.-H. Qu and H. Tian, *Org. Biomol. Chem.*, 2011, **9**, 1126; (d) C. J. Bruns and J. F. Stoddart, *Nat. Nanotechnol.*, 2013, **8**, 794; (e) C. J. Bruns and J. F. Stoddart, *Acc. Chem. Res.*, 2014, **47**, 2186; (f) C. J. Bruns, M. Frascioni, J. Iehl, K. J. Hartlieb, S. T. Schneckeli, C. Cheng, S. I. Stupp and J. F. Stoddart, *J. Am. Chem. Soc.*, 2014, **136**, 4714; (g) L. Gao, Z. Zhang, B. Zheng and F. Huang, *Polym. Chem.*, 2014, **5**, 5734.
- 5 (a) J. D. Badjic, V. Balzani, A. Credi, S. Silvi and J. F. Stoddart, *Science*, 2004, **303**, 1845; (b) J. D. Badjic, C. M. Ronconi, J. F. Stoddart, V. Balzani, S. Silvi and A. Credi, *J. Am. Chem. Soc.*, 2006, **128**, 1489; (c) Z.-J. Zhang, M. Han, H.-Y. Zhang and Y. Liu, *Org. Lett.*, 2013, **15**, 1698.
- 6 (a) M. Schliwa, *Molecular Motors*, Wiley-VCH, Weinheim, 2003; (b) G. S. Kottas, L. I. Clarke, D. Horinek and J. Michl, *Chem. Rev.*, 2005, **105**, 1281; (c) S. P. Fletcher, F. Dumur, M. M. Pollard and B. L. Feringa, *Science*, 2005, **310**, 80; (d) E. R. Kay and D. A. Leigh, *Pure Appl. Chem.*, 2008, **80**, 17; (e) S. Lee and W. Lu, *Nanotechnology*, 2011, **22**, 205501.
- 7 (a) C. A. Stanier, S. J. Alderman, T. D. W. Claridge and H. L. Anderson, *Angew. Chem., Int. Ed.*, 2002, **41**, 1769; (b) T. Iijima, S. A. Vignon, H.-R. Tseng, T. Jarroson, J. K. M. Sanders, F. Marchioni, M. Venturi, E. Apostoli, V. Balzani and J. F. Stoddart, *Chem.-Eur. J.*, 2004, **10**, 6375; (c) A. Altieri, F. G. Gatti, E. R. Kay, D. A. Leigh, D. Martel, F. Paolucci, A. M. Z. Slawin and J. K. Y. Wong, *J. Am. Chem. Soc.*, 2003, **125**, 8644; (d) G. Bottari, D. A. Leigh and E. M. Pérez, *J. Am. Chem. Soc.*, 2003, **125**, 13360; (e) X. Ma, Q.-C. Wang, D.-H. Qu, Y. Xu, F.-Y. Ji and H. Tian, *Adv. Funct. Mater.*, 2007, **17**, 829; (f) Q.-C. Wang, D.-H. Qu and H. Tian, *Angew. Chem., Int. Ed.*, 2004, **43**, 2661; (g) E. M. Pérez, D. T. F. Dryden, D. A. Leigh, T. Gilberto and Z. Francesco, *J. Am. Chem. Soc.*, 2004, **126**, 12210.
- 8 (a) Y. Wei, S. Shuang, C. Dong, W. Jin, C. Gao, Y. Wei and C. Liu, *Trends Appl. Spectrosc.*, 2004, **5**, 195; (b) I. Sánchez-Barragán, J. M. Costa-Fernández, M. Valledor, J. C. Campo and A. Sanz-Medel, *Trends Anal. Chem.*, 2006, **25**, 958.
- 9 (a) R. A. Paynter, S. L. Wellons and J. D. Winefordner, *Anal. Chem.*, 1974, **46**, 736; (b) L. J. C. Love, M. Skrilec and J. G. Habarta, *Anal. Chem.*, 1980, **52**, 754.
- 10 (a) X. Ma, J. Cao, Q. Wang and H. Tian, *Chem. Commun.*, 2011, **47**, 3559; (b) J. Cao, X. Ma, M. Min, T. Cao, S. Wu and H. Tian, *Chem. Commun.*, 2014, **50**, 3224; (c) H. Masai, J. Terao, S. Makuta, Y. Tachibana, T. Fujihara and Y. Tsuji, *J. Am. Chem. Soc.*, 2014, **136**, 14714; (d) H. Chen, X. Ma, S. Wu and H. Tian, *Angew. Chem., Int. Ed.*, 2014, **53**, 14149.
- 11 P. R. Ashton, M. R. Johnston, J. F. Stoddart, M. S. Tolley and J. W. Wheeler, *J. Chem. Soc., Chem. Commun.*, 1992, 1128.
- 12 J. C. Chambron, V. Heitz and J. P. Sauvage, *J. Chem. Soc., Chem. Commun.*, 1992, 1131.
- 13 (a) M. J. Blanco, M. C. Jiménez, J. C. Chambron, V. Heitz, M. Linke and J. P. Sauvage, *Chem. Soc. Rev.*, 1999, **28**, 293; (b) H. Sasabe, N. Kihara, Y. Furusho, K. Mizuno, A. Ogawa and T. Takata, *Org. Lett.*, 2004, **6**, 3957; (c) K. Li, D. I. Shuster, D. M. Guldi, M.-A. Herranz and L. Echegoyen, *J. Am. Chem. Soc.*, 2004, **126**, 3388; (d) D. I. Schster, K. Li, D. M. Guldi and J. Ramey, *Org. Lett.*, 2004, **6**, 1919; (e) J.-S. Marois, K. Cantin, A. Desmarais and J.-F. Morin, *Org. Lett.*, 2008, **10**, 33; (f) J. P. Collin, J. Frey, V. Heitz, J. P. Sauvage, C. Tock and L. Allouche, *J. Am. Chem. Soc.*, 2009, **131**, 5609; (g) M. J. Langton, J. D. Matichak, A. L. Thompson and H. L. Anderson, *Chem. Sci.*, 2011, **2**, 1897; (h) Y. Yamada, N. Mihara, S. Shibano, K. Sugimoto and K. Tanaka, *J. Am. Chem. Soc.*, 2013, **135**, 11505.
- 14 (a) D. B. Papkovsky, G. V. Ponomarev, W. Trettnak and P. O'Leary, *Anal. Chem.*, 1995, **67**, 4112; (b) P. Hartmann and W. Trettnak, *Anal. Chem.*, 1996, **68**, 2615; (c) A. Mills and A. Lepre, *Anal. Chem.*, 1997, **69**, 4653; (d) M. A. Baldo, D. F. O'Brien, Y. You, A. Shoustikov, S. Sibley, M. E. Thompson and S. R. Forrest, *Nature*, 1998, **395**, 151; (e) C.-M. Che, Y.-J. Hou, M. C. W. Chan, J. H. Guo, Y. Liu and Y. Wang, *J. Mater. Chem.*, 2003, **13**, 1362; (f) T. Mani, M. Tanabe, S. Yamauchi, N. V. Tkachenko and S. A. Vinogradov, *J. Phys. Chem. Lett.*, 2012, **3**, 3115.



- 15 (a) A. M. Brouwer, C. Frochot, F. G. Gatti, D. A. Leigh, L. Mottier, F. Paolucci, S. Roffia and G. W. H. Worpel, *Science*, 2001, **291**, 2124; (b) J. Wu, K. C.-F. Leung, D. Benítez, J.-Y. Han, S. J. Cantrill, L. Fang and J. F. Stoddart, *Angew. Chem., Int. Ed.*, 2008, **39**, 7470; (c) S.-Y. Hsueh, C.-T. Kuo, T.-W. Lu, C.-C. Lai, Y.-H. Liu, H.-F. Hsu, S.-M. Peng, C.-H. Chen and S.-H. Chiu, *Angew. Chem., Int. Ed.*, 2010, **48**, 9170; (d) Z. Zhang, C. Han, G. Yu and F. Huang, *Chem. Sci.*, 2012, **3**, 3026; (e) Z. Meng, J.-F. Xiang and C.-F. Chen, *Chem. Sci.*, 2014, **5**, 1520; (f) H. Li, X. Li, A. Hans and D.-H. Qu, *Org. Lett.*, 2014, **16**, 4940.
- 16 (a) M.-V. Martínez-Díaz, N. Spencer and J. F. Stoddart, *Angew. Chem., Int. Ed.*, 1997, **36**, 1904; (b) P. R. Ashton, R. Ballardini, V. Balzani, I. Baxter, A. Credi, M. C. T. Fyfe, M. T. Gandolfi, M. Gómez-López, M.-V. Martínez-Díaz, A. Piersanti, N. Spencer, J. F. Stoddart, M. Venturi, A. J. P. White and D. J. Williams, *J. Am. Chem. Soc.*, 1998, **120**, 11932.
- 17 Since the Bpym<sup>2+</sup> unit is very sensitive to the nucleophilic base, tributylamine (Bu<sub>3</sub>N) was chosen as the ideal base, which was strong enough to deprotonate the NH<sup>2+</sup> center and behaved as a relatively weak (or unreactive) nucleophile towards Bpym<sup>2+</sup> units. See ref. 16 and I. Eléna, C. Alberto, B. Vincenzo, S. Francesco and M. Luigi, *Chem.-Eur. J.*, 1999, **5**, 984.
- 18 (a) M. Roza-Fernández, M. J. Valencia-González and M. E. Díaz-Gacía, *Anal. Chem.*, 1997, **69**, 2406; (b) J. A. M. Pulgarín, A. A. Molina and P. F. López, *Analyst*, 1997, **122**, 253; (c) A. S. Carretero, C. C. Blanco, B. C. Díaz and A. F. Gutiérrez, *Analyst*, 1998, **123**, 1069; (d) Y.-T. Wang and W.-J. Jin, *Spectrochim. Acta, Part A*, 2008, **70**, 871.
- 19 (a) H. Onagi and J. Rebek, Jr., *Chem. Commun.*, 2005, **36**, 4604; (b) W.-Y. Wong, K. C.-F. Leung and J. F. Stoddart, *Org. Biomol. Chem.*, 2010, **8**, 2332; (c) H. Zhang, B. Zhou, H. Li, D.-H. Qu and H. Tian, *J. Org. Chem.*, 2013, **78**, 2091; (d) H. Zhang, J. Hu and D.-H. Qu, *Org. Lett.*, 2012, **14**, 2334; (e) C.-S. Kwan, A. S. C. Chan and K. C.-F. Leung, *Org. Lett.*, 2016, **18**, 976.
- 20 (a) P. R. Ashton, R. Ballardini, V. Balzani, M. Bělohradský, M. T. Gandol, D. Philp, L. Prodi, F. M. Raymo, M. V. Reddington, N. Spencer, J. F. Stoddart, M. Venturi and D. J. Williams, *J. Am. Chem. Soc.*, 1996, **118**, 4931; (b) M. Fathalla, J. C. Barnes, R. M. Young, K. J. Hartlieb, S. M. Dyar, S. W. Eaton, A. A. Sarjeant, D. T. Co, M. R. Wasielewski and J. F. Stoddart, *Chem.-Eur. J.*, 2014, **20**, 14690.

