A room temperature phosphorescence encoding [2]rotaxane molecular shuttle†

Xiang Ma,∗ Jing Zhang, Jingjing Cao, Xuyang Yao, Tiantian Cao, Yifan Gong, Chunchang Zhao and He Tian

A novel bistable molecular shuttle, composed of a Pt(II) porphyrin-containing dibenzo[24]crown-8 macrocycle threaded onto two different recognition sites (secondary dialkylammonium (NH2)+ and 4,4′-bipyridinium (Bpym2+) units), and the anthracene (Anth) moiety as one terminal stopper, was synthesized by click chemistry. Its acid–base switchable shuttling could be addressed by both the room temperature phosphorescence (RTP) emission signals of the Pt(II) porphyrin moiety and the fluorescence emission of the Anth unit, as well as their lifetime changes. When the macrocycle was switched to be located on the NH2+ site close to Anth, the Pt(II) porphyrin moiety exhibited strong RTP emission, excited in the Anth band at 370 nm. This was due to the distance-dependent efficient singlet energy transfer between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to a triplet state in Pt(II) porphyrin, while its RTP emission dramatically decreased when located on the Bpym2+ site far from the Anth unit. When excited in the porphyrin band at 402 nm, the RTP emission lifetimes changed obviously. This is the first rotaxane-type molecular shuttle whose shuttling has been encoded by RTP signals.

Introduction

Rotaxanes,† as one of the challenging topics of molecular machines, have attracted continued interest and attention. A variety of appealing rotaxanes have been fabricated over the recent years, with various functions or applications in such areas as switches,5 logic gates,6 artificial muscles,7 molecular elevators,8 molecular motors,9 and so forth. The output signals of most rotaxane shuttles were observed by NMR, UV-Vis absorption, cyclic voltammetry, circular dichroism, fluorescence emission, and so forth.7 Room temperature phosphorescence (RTP), with the advantages of longer emission wavelengths and lifetimes, larger ‘Stokes’ shifts, higher signal-to-noise ratios, quick response and low-cost detection etc.,8 is another excellent method which is expected to be employed to address the shuttling and conformation of rotaxane shuttles.9 Up to now, few rotaxane examples employing phosphorescence signal changes as the output have been reported.10

Porphyrins are important compounds in relation to electron and energy transfer processes. Stoddart11 and Sauvage12 first introduced the porphyrin moiety into rotaxane systems. Subsequently, several rotaxanes and catenanes containing porphyrins have been investigated.13 Particularly, Pt(II) porphyrins have been reported to have characteristic long wave RTP emissions.14 Therefore, incorporating a Pt(II) porphyrin moiety into a rotaxane might engender new signals, i.e. RTP emission and lifetime changes, to address the shuttling and conformation of rotaxanes.

Herein, we designed and synthesized novel acid–base controlled molecular shuttles15 based on switchable [2]rotaxanes 1 (Scheme 1), with Pt(II) porphyrin-based dibenzo[24]crown-8 as the host macrocycle threaded onto two different recognition sites (secondary dialkylammonium (NH2+) and 4,4′-bipyridinium (Bpym2+) units), and the anthracene (Anth) moiety as one terminal stopper. In this [2]rotaxane, reversible motions of the macrocycle between the two stations could be controlled by the protonation and deprotonation of the NH2+ recognition site, as previously observed in the related system.16 These motions affect the interactions between the porphyrin moiety on the macrocycle and the Anth stopper. Upon protonation of the NH2+ recognition site, the macrocycle is located at said site, and the porphyrin unit nearby could also interact with the Anth unit by energy transfer. Deprotonation of the NH2+ center could engender the displacement of the macrocycle to the Bpym2+ station, and destroy the energy transfer process due to the elongated distance between the Anth and porphyrin units. These switching processes could be investigated using conventional methods, such as 1H NMR spectroscopy and fluorescence spectroscopy. Gratifyingly, after the metallization...
of the porphyrin core with Ptn salt, the shuttling of rotaxane 1-H accompanied by the macrocycle 6 staying between the two recognition sites could also be monitored by obvious RTP signal changes, which makes this the first [2]rotaxane molecular shuttle with switchable RTP output addresses so far.

Results and discussions

The [2]rotaxane 1-H was synthesized conveniently via a copper(i) catalyzed azide–alkyne ‘click’ cycloaddition. As shown in Scheme 1, treating equal equivalents of 4 and 5 with 2.0 equivalents of macrocycle 6 in the absence of light and oxygen led to the formation of the [2]rotaxane 1-H with a relatively low yield of 13%, which might be due to the relatively lower binding constant of the DB24C8 decorated with Ptn porphyrin (macrocycle 6) with the NH3+ center of compound 4 and steric hindrance, besides the obvious product loss during purification by chromatography on silica gel. The triazole ring and the oxygen atom of the dumbbell were used as long linkers not only to increase the distance between the two recognition sites, preventing energy transfer from the anthracene moiety to the 4,4′-bipyridium moiety as much as possible, but also to ensure the stability of the interlocking states by avoiding adverse factors like steric hindrance. Another reference [2]rotaxane 3-H was constructed by replacing macrocycle 6 with porphyrin-free dibenzo[24]crown-8 (DB24C8). The addition of 1.0 equiv. of n-Bu3N to the solution of [2]rotaxane 1-H results in the deprotonated form of 1. Reprotonation can be performed by the addition of a slight molar excess of CF3COOH.

The [2]rotaxane formation and its reversible switching process were monitored by 1H NMR spectroscopy. As shown in Fig. 1, the dumbbell and macrocycle components of 1-H/1 are present in equimolar amounts and are interlocked with each other. The spectrum (Fig. 1c) demonstrates that the macrocycle 6 binds selectively with the NH3+ center, since the characteristic resonances H4 and H5 of the dumbbell 2-H, at $\delta$ 5.59 and 4.39 ppm (Fig. 1b), for the four CH2 protons adjacent to the NH3+ group experienced upfield shifts, H4 at $\delta$ 5.24 ppm ($\Delta\delta = 0.35$ ppm) and H5 at 4.22 ppm ($\Delta\delta = 0.17$ ppm), respectively, upon rotaxane (1-H) formation. Furthermore, the aromatic protons H6 of the macrocycle 6 were split into two peaks ($\delta$ 7.15 and 7.13 ppm) after complexation.

After the addition of n-Bu3N, the 1H NMR spectrum of 1 shows dramatic chemical shift changes with respect to that of 1-H. The resonances ($\delta$ 5.24 and 4.22 ppm) of the CH2–NH3–CH2 protons disappeared and new signals ($\delta$ 5.38 and 4.40 ppm, Fig. 1d), assigned to CH2–NH–CH2 protons, appeared. Moreover, the macrocycle 6 glides towards the Bpym2+ unit, resulting in the obvious changes of the proton resonance values of Bpym3+ ($\delta$ [A$\delta$] for H$\alpha$, H$\beta$, H$\gamma$, and H$\delta$ ppm): 9.56 [0.22], 9.28 [0.18], 8.70 [0.10] and 8.51 [0.28], respectively. Upon addition of CF3COOH, the original 1H NMR spectrum was almost regenerated, demonstrating that the switching process was fully reversible via de-reprotonation (Fig. S5†).

The photophysical and photochemical properties of the shuttling processes of rotaxane 1-H/1 were thoroughly investigated (Table S1†). Ptn porphyrin-based dibenzo[24]crown-8 6 displayed relatively weak fluorescence emission bands at 650 and 715 nm in air equilibrated CH3CN (Fig. S7†). Notably, no RTP was exhibited even in a deaerated organic solution, but strong RTP signals were observed in the medium containing the anionic surfactant sodium dodecyl sulfate (SDS) and Na2SO3 for removing oxygen.18 The characteristic RTP emission peaks of 6 were exhibited at 667 and 736 nm (inset, Fig. S7†) when excited in the porphyrin band at 402 nm, and experienced a redshift compared with the fluorescence emission peak (Fig. S7†). Upon forming the [2]rotaxane 1-H, the fluorescence emission was almost quenched,19 while its RTP emission showed a slight increase due to the movement restriction caused by the

Scheme 1 Chemical structures of the macrocycle 6, and the [2]rotaxanes 1-H, 1 and 3-H.
[Extracted text from the image]
when excited in the Anth band by light at 370 nm (Fig. S17 and S18†). Upon deprotonation, the intensity of the fluorescence emission band of the Anth unit of the dumbbell 2, shown in Fig. 4, decreases. This is attributed to the electron transfer interaction between the Anth moiety and the amino group.20 Conversely, the intensity of the Anth bands for the [2]rotaxane 1 experiences a dramatic increase with respect to that of 1-H. Consequently, both the increase of the porphyrin based RTP emission and the decrease of the Anth based fluorescence emission for 1-H compared with those of its deprotonated species should be attributed to the efficient singlet energy transfer process between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to triplet state in Pt(II) porphyrin (Fig. 5).

The quantum yields for the various RTP emission processes were also investigated, as shown in Table 2. The maximum absorption peak of the Pt(II) porphyrin-containing macrocycle 6 is at about 402 nm while the absorption at around 370 nm is very weak (Fig. S6†). Thus the quantum yields of 6, 6 + base and [2]rotaxane 1-H excited at 402 nm were relatively higher than the ones excited at 370 nm. After forming [2]rotaxane, the quantum yield of 1-H (0.16) was slightly increased with respect to that of the host macrocycle 6 due to the efficient singlet energy transfer process between the Anth unit and the porphyrin moiety, followed by intersystem crossing from a singlet to triplet state in Pt(II) porphyrin. While the quantum yields of [2]rotaxane 1 (1-H + base) excited at both 402 nm and 370 nm were too low to quantify, since the 4,4’-bipyridinium (Bpym2+) units in 1 had some quenching effects on the emission of the porphyrin moiety of 6. Also, the distance dependent energy transfer efficiency from the Anth to porphyrin units decreased when the macrocycle 6 was located on the Bpym2+ site far from the Anth moiety.

The fluorescence lifetime of the excited Anth unit in the [2]rotaxane 1-H was found to be 2.79 ns (Fig. S13†). It was shorter than that of the corresponding dumbbell 2-H (4.13 ns, Fig. S14†), which was attributed to the energy transfer from Anth to Pt(II) porphyrin moieties in 1-H (Fig. S19–S21† and discussion therein). Upon the addition of base, the lifetime of the excited Anth unit in 1 increased compared with 1-H and reached up to 3.79 ns (Fig. S13†), which was due to the break up of the energy transfer process from the Anth to Pt(II) porphyrin units owing to the longer distance between them (Fig. S20†). The emission spectrum showed similarities with that of the original [2]rotaxane 1-H upon addition of an equivalent of CF3COOH to its solution, indicating that the macrocycle was situated on the NH2 center once again. The fluorescence emission intensity changes of [2]rotaxane 1 (at 412 nm) also showed good reversibility after alternating protonation/deprotonation for several cycles, indicating that its shuttling was totally reversible (Fig. 4, inset).

Conclusions

In summary, a novel bistable molecular shuttle, based on an acid–base switchable [2]rotaxane in which the Pt(II) porphyrin-
containing macrocycle could be stimulated to shuttle back and forth between the two recognition sites, was constructed, whose shuttling process was accompanied by reversible changes both in the RTP emission (both intensities and lifetimes) of the Pt(n) porphyrin moiety and the fluorescence emission of the Anth unit. It should be noticed that this is the first example of a rotaxane-type molecular shuttle whose shuttling could be encoded by RTP signals. The concept and advance in this work, employing RTP signals to address the shuttling of a rotaxane, supply a novel method to encode a molecular shuttle and pave the way for designing molecular machines with multi-responses and multi-functions.

**Experimental section**

**Synthesis of 1-H**

A solution of the azide 4 (56 mg, 0.11 mmol) and the ring 6 (200 mg, 0.17 mmol) in 10 mL of CHCl₃ was stirred for 24 h at room temperature under Ar. The BIPY²⁺ derivative 5 (71 mg, 0.1 mmol) and Cu(MeCN)₄PF₆ (92 mg, 0.05 mmol) in 10 mL of dry Me₆CO were added, and the whole solution was stirred for another 3 days. After the solvent was evaporated, the resulting redish solid was washed with H₂O, air dried and purified by column chromatography on silica gel (Me₆CO to 2% NH₄PF₆/Me₆CO) to obtain the [2]rotaxane 1-H (31 mg, 13%).¹H NMR (500 MHz, CDCl₃): δH (ppm) 9.56 (d, 2H, J = 5.9 Hz), 9.31 (d, 2H, J = 5.5 Hz), 8.86 (m, 8H), 8.74 (d, 2H, J = 6.2 Hz), 8.50 (d, 2H, J = 6.5 Hz), 8.58 (d, 3H, J = 9.0 Hz), 8.23 (t, 6H, J = 7.3 Hz), 8.11 (d, 2H, J = 8.1 Hz), 7.87 (m, 11H), 7.62 (s, 3H), 7.53 (m, 4H), 7.15 (d, 2H, J = 6.0 Hz), 7.13 (d, 2H, J = 6.0 Hz), 6.09 (s, 2H), 5.59 (s, 2H), 5.24 (t, 2H, J = 6.5 Hz), 4.70–3.70 (m, 29H), 3.60 (t, 3H, J = 6.0 Hz), 3.39 (m, 4H), 2.64 (s, 2H), 2.15 (s, 2H), 1.32 (s, 18H).¹C NMR (100 MHz, CDCl₃): δC (ppm) 153.92, 151.77, 151.50, 147.39, 146.85, 134.40, 133.58, 132.80, 132.24, 131.63, 130.62, 128.85, 128.61, 128.24, 127.72, 126.67, 125.36, 124.96, 124.32, 78.92, 75.18, 71.47, 67.46, 66.56, 62.30, 61.23, 58.31, 56.32, 51.31, 35.94, 31.76, 28.10, 26.02, 23.71, 21.87, 18.38, 14.44. HRMS (ESI): m/z [1-H-PF₆]⁺ calc. for [C₁₁₂H₁₁₅N₁₀O₉Pt]F₁₂O₂P₂Pt⁺ 2228.88, found 2229.2. MALDI-TOF-MS: m/z [1-H-PF₆]⁺ calc. for C₁₁₂H₁₁₅N₁₀O₂Pt 1938.85, found 1938.78.

**Synthesis of 2-H**

A solution of the azide 4 (197.6 mg, 0.44 mmol), and the BIPY²⁺ derivative 5 (140 mg, 0.2 mmol), and Cu(MeCN)₄PF₆ (33.2 mg, 0.1 mmol) in dry Me₆CO (10 mL) was stirred for 48 h at room temperature under Ar. After the solvent was evaporated, the resulting redish solid was washed with H₂O, air dried and purified by column chromatography on silica gel (Me₆CO to 2% NH₄PF₆/Me₆CO) to afford the compound 2-H (32 mg, 13%).¹H NMR (500 MHz, CDCl₃): δH (ppm) 9.42 (d, 2H, J = 6.0 Hz), 8.77 (s, 1H), 8.57–8.48 (m, 4H), 8.36 (d, 2H, J = 6.0 Hz), 8.17 (d, 2H, J = 8.0 Hz), 7.73 (s, 1H), 7.69–7.54 (m, 6H), 7.46 (s, 1H), 6.75–6.61 (m, 4H), 6.51–6.31 (m, 4H), 5.89 (s, 2H), 5.21 (s, 2H), 5.07 (t, 2H, J = 8.0 Hz), 4.32 (t, 2H, J = 6.4 Hz), 4.11–3.99 (m, 4H), 3.96 (m, 4H), 3.83–3.66 (m, 10H), 3.64–3.41 (m, 12H), 3.16 (d, 2H, J = 4.8 Hz), 1.93–1.78 (m, 2H), 1.59–1.48 (m, 2H), 1.30 (s, 18H).¹C NMR (100 MHz, CDCl₃): δC (ppm) 151.63, 146.77, 146.54, 145.24, 142.89, 133.78, 130.95, 130.69, 129.13, 127.00, 126.59, 125.54, 124.39, 123.81, 123.41, 123.11, 122.25, 120.50, 111.98, 70.22, 69.89, 69.67, 67.78, 49.14, 46.86, 42.56, 34.80, 31.22, 26.71, 26.14, 25.07, 22.55. HRMS (ESI): m/z [3-H-PF₆]⁺ calc. for [C₇₄H₇₅N₂O₂P₂F₁₂O₄]⁺ 1209.64, found 1209.64.

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**Notes and references**


17 Since the Bpym$_2^+$ unit is very sensitive to the nucleophilic base, tributylamine (Bu$_3$N) was chosen as the ideal base, which was strong enough to deprotonate the NH$_2^+$ center and behaved as a relatively weak (or unreactive) nucleophile towards Bpym$_2^+$ units. See ref. 16 and I. Eléna, C. Alberto, B. Vincenzo, S. Francesco and M. Luigi, *Chem.–Eur. J.*, 1999, **5**, 984.

