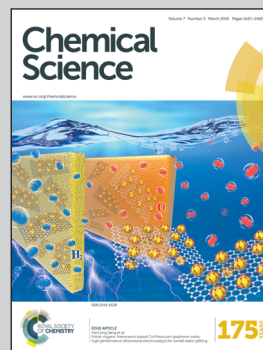


Showcasing research from Jonathan Nitschke's laboratory at the University of Cambridge, United Kingdom.

Title: Dual stimuli-induced formation of a  $\mu$ -hydroxido bridged  $[\text{Zn}_3\text{L}_5(\mu\text{-OH})_6]^{12+}$  half-pipe

A low-symmetry metal-organic architecture was prepared that contains a central guest-binding groove. The formation of this 'half-pipe' requires a base and the right kind of template ion – perchlorate, perrhenate, or a sulfonate – in addition to  $\text{Zn}^{\text{II}}$  and a rigid tritopic organic ligand.

As featured in:



See Jonathan R. Nitschke et al., *Chem. Sci.*, 2016, 7, 1702.



[www.rsc.org/chemicalscience](http://www.rsc.org/chemicalscience)

Registered charity number: 207890



CrossMark  
click for updates

Cite this: *Chem. Sci.*, 2016, 7, 1702

## Dual stimuli-induced formation of a $\mu$ -hydroxido bridged $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$ half-pipe†

Christopher S. Wood, Tanya K. Ronson, Anna J. McConnell, Derrick A. Roberts and Jonathan R. Nitschke\*

Low-symmetry metal–organic architectures that feature unusual binding motifs are useful for exploring new modes of guest recognition. Such structures remain difficult to create using current rational design principles. One approach to constructing such architectures is to employ ligands with coordination vectors oriented to preclude the formation of simple, low nuclearity molecular assemblies upon complexation to metal ions. Here we report two new supramolecular assemblies generated from such a ligand: a simple metastable  $[\text{Zn}_3\text{L}_3]^{6+}$  assembly, which was observed to convert to a more complex  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  twisted half-pipe architecture. Two chemically distinct stimuli—an anionic template and a base—must be applied for the conversion to occur. Perchlorate, perrhenate, trifluoromethanesulfonate and 2-naphthalenesulfonate were found to act as competent templates for the  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  structure.

Received 16th October 2015  
Accepted 11th November 2015

DOI: 10.1039/c5sc03926f

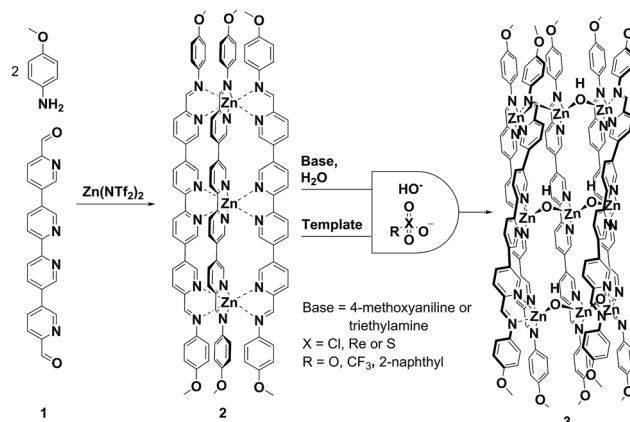
www.rsc.org/chemicalscience

### Introduction

Complex, responsive molecular networks constitute living systems and are thus subjects of intense study.<sup>1–4</sup> Recently, the development of synthetic chemical networks that respond in complex ways to stimuli has provided a platform to study and emulate the fundamental processes of natural systems.<sup>5–8</sup> A promising strategy is to exploit the dynamic nature of organic<sup>9–14</sup> and metal–organic self-assembled complexes<sup>15–19</sup> to generate systems that undergo rearrangement upon application of a chemical signal. Previously reported systems have registered a single response to a single signal, *e.g.*, a template,<sup>20–23</sup> ligand<sup>24–27</sup> or light.<sup>28–35</sup> More complex behaviour can be designed when systems exhibit differential responses to multiple signals, acting as logic gates.<sup>36–44</sup> In the present case of an AND gate,<sup>45,46</sup> two distinct stimuli must be present together in order to elicit a functional response.

Subcomponent self-assembly<sup>47–54</sup> has proved to be a powerful tool for generating metallosupramolecular assemblies.<sup>55–62</sup> Many of these assemblies, by virtue of the dynamic interactions that hold them together, exhibit complex responses to external stimuli.<sup>63,64</sup> Increasing the number and type of ligands that comprise the system can increase the complexity of its response, as different components respond to stimuli in distinct ways.

We have recently developed a strategy to favour the formation of heteroleptic assemblies containing multitopic pyridyl-imine ligands.<sup>65,66</sup> Quaterpyridine dialdehyde subcomponent **1** (Scheme 1) was designed to form ligands with three rigidly parallel coordination vectors following metal-templated Schiff-base formation. The orientations of these vectors precluded the formation of discrete structures upon complexation with iron(II) salts, instead favouring the formation of mixed-ligand structures when other complimentary ligands were employed in the self-assembly reaction.<sup>65</sup> With zinc(II), the geometry of **1**,



**Scheme 1** Schematic representation of the synthesis of  $[\text{Zn}_3\text{L}_3]^{6+}$  assembly **2** and its conversion to a  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  twisted half-pipe structure **3**. Its formation requires the presence of both a base that forms the six  $\mu\text{-OH}$  bridges in the structure and template anions that reside in the hydroxide lined groove of the structure. All reactions were carried out in acetonitrile at 70 °C for 12 h.

Department of Chemistry, University of Cambridge, Lensfield Road, CB2 1EW, UK.  
E-mail: jrn34@cam.ac.uk

† Electronic supplementary information (ESI) available. CCDC 1410155–1410158. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5sc03926f



together with *in-situ*-generated  $\mu$ -hydroxide complimentary ligands and specific anionic templates, gave rise to a singular twisted 'half-pipe' architecture, **3**.

## Results and discussion

When **1** (1 equiv.), 4-methoxyaniline (2 equiv.) and zinc(II) bis(trifluoromethylsulfonyl)imide (triflimide, 2 equiv.) were mixed in acetonitrile, assembly **2** was observed to form. Its  $^1\text{H}$  NMR spectrum was consistent with a symmetric structure possessing a single ligand environment (ESI, Fig. S1 $^\dagger$ ) and the ESI mass spectrum was consistent with a  $[\text{Zn}_3\text{L}_3]^{6+}$  formulation (ESI, Fig. S5 $^\dagger$ ). By contrast, no well-defined species were observed to form in the analogous reaction involving iron(II) in place of zinc(II).<sup>65</sup> We hypothesize that **2** forms with zinc(II) due to its larger ionic radius (compared to  $\text{Fe}^{II}$ ) and its lack of a stereoelectronic preference for a regular octahedral coordination environment.<sup>67</sup> The ground-state geometry of **2** may be either helical<sup>68</sup> ( $D_3$  point symmetry) or a 'mesocate'<sup>69–72</sup> ( $D_{3h}$  symmetry).

Upon addition of a suitable template anion and a source of hydroxide, **2** reorganised to give the  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  twisted half-pipe architecture **3**, featuring a longitudinal groove into which template anions were observed to bind (Scheme 1). The conversion of **2** into **3** was monitored by  $^1\text{H}$  NMR spectroscopy following addition of tetrabutylammonium perchlorate (4 equiv.) to an acetonitrile solution of **2**, 4-methoxyaniline (6 equiv.) and  $\text{H}_2\text{O}$  (ca. 50 equiv.) at room temperature. After two minutes the resonances associated with **2** decreased in intensity and were replaced with a complicated series of broad resonances. Over a period of 12 h at  $70^\circ\text{C}$  the resonances associated with **3** appeared and grew in intensity (ESI, Section 1.7 $^\dagger$ ). ESI-MS of the resulting solution revealed signals with  $m/z$  values corresponding to  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  with both perchlorate and triflimide anions. After 12 h at  $70^\circ\text{C}$  there were no further changes to the  $^1\text{H}$  NMR spectrum of the reaction mixture. Resonances corresponding to **2** could, however, still be observed, which is attributed to the stoichiometric mismatch between the  $\text{M}_3\text{L}_3$  and  $\text{M}_9\text{L}_5$  architectures.

Half-pipe **3** could also be prepared directly from its precursors. When zinc(II) perchlorate hexahydrate (6 equiv.), **1** (5 equiv.) and 4-methoxyaniline (16 equiv.) were mixed in acetonitrile for 12 h at  $70^\circ\text{C}$ , a turbid yellow suspension resulted. Following filtration and precipitation with diethyl ether the resulting product possessed a  $^1\text{H}$  NMR spectrum corresponding to a species with five distinct ligand environments (ESI, Fig. S9 and S10 $^\dagger$ ). Slow diffusion of diethyl ether into an acetonitrile solution of the product afforded X-ray quality crystals. X-ray crystallography revealed the structure of complex **3**, which contains nine zinc(II) centres linked by five ligands (Fig. 1). Three ligands pack tightly side-by-side with their coordination vectors parallel and aromatic stacking interactions occurring between neighbouring ligands. The remaining two ligands lie opposite, with their coordination vectors antiparallel to the first three. This arrangement produces a longitudinal groove between the two outermost ligands, within which four perchlorate anions bind (Fig. 1a and c). These anions hydrogen

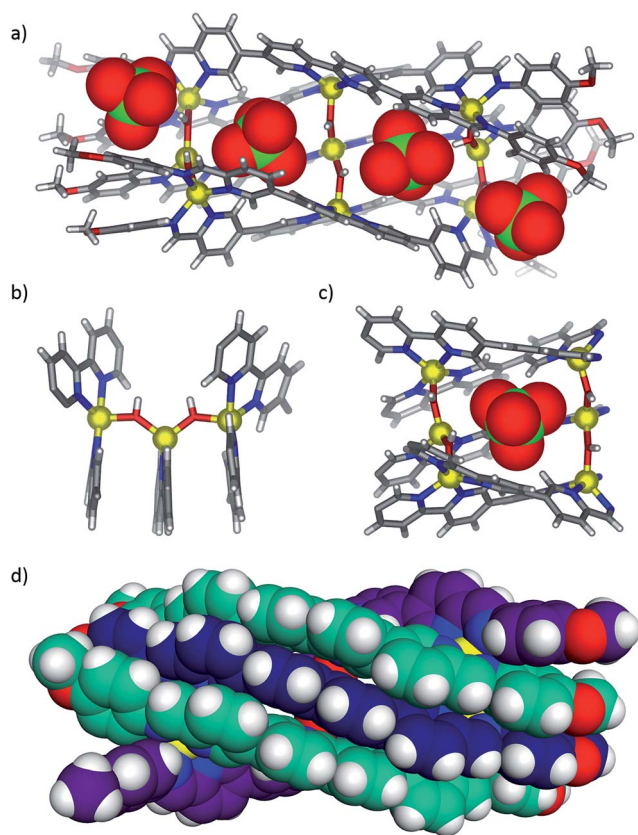


Fig. 1 X-ray crystal structure of  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6](\text{ClO}_4)_{12}$  twisted half-pipe **3**. (a) View down the pseudo  $C_2$  axis of symmetry. (b) Cutaway view showing two  $\mu\text{-OH}$  bridging anions coordinating to three adjacent zinc centres. (c) Cutaway showing the perchlorate anion templates inside the groove (C, grey; N, blue; Zn, yellow; O, red; Cl, green; H, white). (d) Space filling representation of **3** showing the three tightly packed ligands and helical twist. Uniform colours represent individual ligands. Disorder, non-coordinating anions and solvent are omitted for clarity.

bond to six bridging hydroxide anions which line the base of the groove, each of which bridges two zinc centres bound by adjacent ligands (Fig. 1b). Thus, the three inner zinc(II) centres adopt a tetrahedral coordination sphere defined by one bidentate and two hydroxide ligands, and the six outer zinc(II) centres are five-coordinate, each bound by two bidentate ligands and a single hydroxide. Both ESI-MS and NMR spectroscopy (ESI, Section 1.3 $^\dagger$ ) provided results consistent with the observed solid-state structure.

The hydroxide ( $\mu\text{-OH}$ ) bridges in **3** are inferred to form when 4-methoxyaniline deprotonates water present in the reaction mixture. Indeed, we found that the cleanest formation of **3** could be observed when six extra equivalents of either 4-methoxyaniline or triethylamine were included in a reaction mixture of the correct stoichiometry. However, the direct addition of six equivalents of tetramethylammonium hydroxide caused the formation of intractable precipitates in the presence of perchlorate, possibly due to the kinetic precipitation of zinc oxide or hydroxide in the presence of free  $\text{HO}^-$ . The perchlorate anions, which were observed to bind in the groove of the structure, are also necessary but not sufficient to cause the





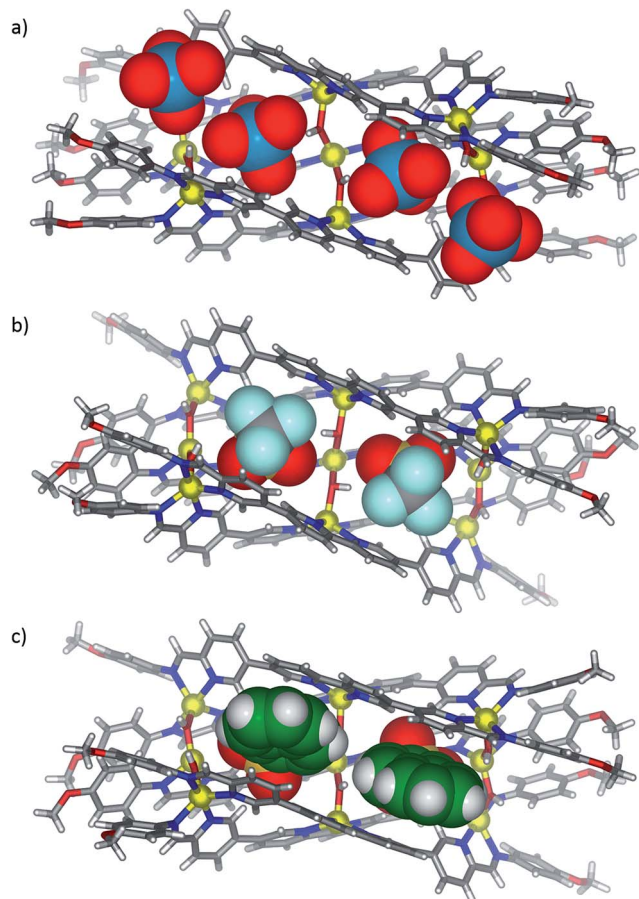


Fig. 2 X-ray crystal structures of **3** templated by (a) perrhenate, (b) triflate and (c) 2-naphthalenesulfonate. The templating anions are shown in space-filling mode. Disorder, non-coordinating anions and solvent are omitted for clarity.

transformation of **2** into **3**. Complex **2** is stable for 12 h at 70 °C in the presence of either four equivalents of tetrabutylammonium perchlorate or six equivalents of 4-methoxyaniline and water. In order to convert complex **2** into **3** both a templating anion and a source of hydroxide were thus necessary.

Perrhenate, trifluoromethanesulfonate (triflate) and 2-naphthalenesulfonate anions were also found to serve as competent templates for the  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]$  motif. Mixtures of zinc(II) triflimide (9 equiv.), **1** (5 equiv.), 4-methoxyaniline (16 equiv.) and a suitable salt of the template anion (4 equiv.) were heated in acetonitrile at 70 °C for 12 h. The resulting turbid yellow suspensions were filtered, and the products isolated by precipitation with diethyl ether.  $^1\text{H}$  NMR spectroscopy and ESI-MS analyses confirmed that both **2** and **3** were present in solution in all three cases. The proportion of **3** to **2** could be increased by slow crystallization and subsequent isolation of the crystalline solid by filtration. However, attempts to increase the concentration of **3** relative to **2** by increasing the concentration of the templating anion were found to increase the formation of insoluble precipitates. The solubility of the Zn(II) salts of these templating anions may be limiting the degree to which the equilibrium between these two species can be influenced through template addition.

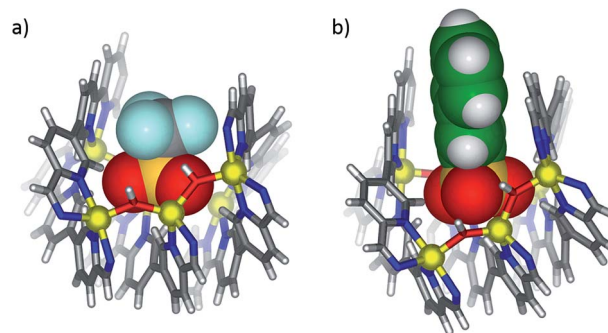


Fig. 3 View down the length of the groove in cutaway X-ray crystal structures of **3** templated by (a) triflate and (b) 2-naphthalenesulfonate. The templating anions are shown in space-filling mode.

X-ray quality crystals were obtained for all three anion-templated structures by slow diffusion of diethyl ether into acetonitrile solutions. Solution and solid-state analyses indicated all of these assemblies share the same overall **3** architecture. In the case of the perrhenate anion, which is a close structural analogue of perchlorate, four anions were observed to bind within the groove in similar positions as perchlorate (Fig. 2a); the outer two perrhenates were, however, disordered and modelled with partial occupancy. By contrast, triflate and 2-naphthalenesulfonate were observed to occupy only the two inner binding sites of the groove (Fig. 2b and c). In each case, the  $\text{SO}_3^-$  group was oriented into the groove while the  $\text{CF}_3$  and naphthalene groups pointed away from the groove between the two outer ligands. The planar aromatic naphthalene moieties underwent  $\pi$ -stacking interactions with the two outer ligands, with centroid-centroid distances of 3.7–4.0 Å (Fig. 3).

These additional  $\pi$ -stacking interactions are postulated to engender a greater binding affinity for 2-naphthalenesulfonate than was observed for the other templating anions.<sup>73</sup> Two equivalents of 2-naphthalenesulfonate sufficed to displace the perchlorate, perrhenate or triflate anions from **3**; complex mixtures of products were observed when only one equivalent of 2-naphthalenesulfonate was added (ESI, Section 1.12<sup>†</sup>). A more comprehensive hierarchy of binding affinities could not be determined due to the similarity and complexity of the spectra obtained in other anion displacement experiments (ESI Section 1.11<sup>†</sup>).

## Conclusions

In conclusion, we demonstrate that the combination of a metal template without stereoelectronic coordination preferences – zinc(II) – with a subcomponent designed to disfavour the formation of simple low nuclearity structures – **1** – produces an assembly that is susceptible to multi-stimuli-induced structural transformations. Here a simple  $[\text{Zn}_3\text{L}_3]^{6+}$  assembly was found to rearrange to a more complex  $[\text{Zn}_9\text{L}_5(\mu\text{-OH})_6]^{12+}$  twisted half-pipe assembly only in the presence of both a base, necessary for the generation of the secondary hydroxide ligands, and a templating anion. The cationic, chiral groove of **3**, and other congeners of this new structure type, may prove useful in recognising and



binding more complex anionic substrates, perhaps inducing new chemical transformations within its highly-charged, half-encapsulated inner environment.

## Acknowledgements

This work was supported by the UK Engineering and Physical Sciences Research Council (EPSRC). DAR acknowledges the Gates Cambridge Scholarship for PhD funding. We thank Diamond Light Source (UK) for synchrotron beam time on I19 (MT8464).

## Notes and references

- 1 E. Schrodinger, *What is Life?: With Mind and Matter and Autobiographical Sketches*, Cambridge University Press, Cambridge, 2012.
- 2 D. Bray, *Science*, 2003, **301**, 1864–1865.
- 3 V. Spirin and L. A. Mirny, *Proc. Natl. Acad. Sci. U. S. A.*, 2003, **100**, 12123–12128.
- 4 L. H. Hartwell, J. J. Hopfield, S. Leibler and A. W. Murray, *Nature*, 1999, **402**, C47–C52.
- 5 B. Lewandowski, G. de Bo, J. W. Ward, M. Papmeyer, S. Kuschel, M. J. Aldegunde, P. M. E. Gramlich, D. Heckmann, S. M. Goldup, D. M. D'Souza, A. E. Fernandes and D. A. Leigh, *Science*, 2013, **339**, 189–193.
- 6 S. N. Semenov, A. S. Y. Wong, R. M. van der Made, S. G. J. Postma, J. Groen, H. W. H. van Roekel, T. F. A. de Greef and W. T. S. Huck, *Nat. Chem.*, 2015, **7**, 160–165.
- 7 T. le Saux, R. Plasson and L. Jullien, *Chem. Commun.*, 2014, **50**, 6189–6195.
- 8 M. G. Cowan, J. Olguin, S. Narayanaswamy, J. L. Tallon and S. Brooker, *J. Am. Chem. Soc.*, 2012, **134**, 2892–2894.
- 9 P. Mukhopadhyay, P. Y. Zavalij and L. Isaacs, *J. Am. Chem. Soc.*, 2006, **128**, 14093–14102.
- 10 S. Otto, *Acc. Chem. Res.*, 2012, **45**, 2200–2210.
- 11 R. B. P. Elmes, N. Busschaert, D. D. Czech, P. A. Gale and K. A. Jolliffe, *Chem. Commun.*, 2015, **51**, 10107–10110.
- 12 J. Kim, K. Baek, D. Shetty, N. Selvapalam, G. Yun, N. H. Kim, Y. H. Ko, K. M. Park, I. Hwang and K. Kim, *Angew. Chem., Int. Ed.*, 2015, **54**, 2693–2697.
- 13 R. T. Woodward, L. A. Stevens, R. Dawson, M. Vijayaraghavan, T. Hasell, I. P. Silverwood, A. V. Ewing, T. Ratvijitvech, J. D. Exley, S. Y. Chong, F. Blanc, D. J. Adams, S. G. Kazarian, C. E. Snape, T. C. Drage and A. I. Cooper, *J. Am. Chem. Soc.*, 2014, **136**, 9028–9035.
- 14 G. Zhang and M. Mastalerz, *Chem. Soc. Rev.*, 2014, **43**, 1934–1947.
- 15 J. T. Foy, D. Ray and I. Aprahamian, *Chem. Sci.*, 2015, **6**, 209–213.
- 16 A. J. McConnell, C. S. Wood, P. P. Neelakandan and J. R. Nitschke, *Chem. Rev.*, 2015, **115**, 7729–7793.
- 17 S. Tashiro and M. Shionoya, *Chem. Lett.*, 2013, **42**, 456–462.
- 18 J. B. Beck and S. J. Rowan, *J. Am. Chem. Soc.*, 2003, **125**, 13922–13923.
- 19 W. Cullen, C. A. Hunter and M. D. Ward, *Inorg. Chem.*, 2015, **54**, 2626–2637.
- 20 M. Scherer, D. L. Caulder, D. W. Johnson and K. N. Raymond, *Angew. Chem., Int. Ed.*, 1999, **38**, 1587–1592.
- 21 S. Freye, R. Michel, D. Stalke, M. Pawliczek, H. Frauendorf and G. H. Clever, *J. Am. Chem. Soc.*, 2013, **135**, 8476–8479.
- 22 K. Hagiwara, M. Otsuki, M. Akita and M. Yoshizawa, *Chem. Commun.*, 2015, **51**, 10451–10454.
- 23 B. Hasenknopf, J.-M. Lehn, N. Boumediene, A. Dupont-Gervais, A. van Dorsselaer, B. Kneisel and D. Fenske, *J. Am. Chem. Soc.*, 1997, **119**, 10956–10962.
- 24 C. G. Oliveri, P. A. Ulmann, M. J. Wiester and C. A. Mirkin, *Acc. Chem. Res.*, 2008, **41**, 1618–1629.
- 25 J.-R. Li and H.-C. Zhou, *Nat. Chem.*, 2010, **2**, 893–898.
- 26 Q.-Q. Wang, R. Ara Begum, V. W. Day and K. Bowman-James, *Inorg. Chem.*, 2012, **51**, 760–762.
- 27 D. Samanta and P. S. Mukherjee, *Chem.–Eur. J.*, 2014, **20**, 12483–12492.
- 28 M. Han, R. Michel, B. He, Y.-S. Chen, D. Stalke, M. John and G. H. Clever, *Angew. Chem., Int. Ed.*, 2013, **52**, 1319–1323.
- 29 S.-S. Sun, J. A. Anspach and A. J. Lees, *Inorg. Chem.*, 2002, **41**, 1862–1869.
- 30 T. Murase, S. Sato and M. Fujita, *Angew. Chem., Int. Ed.*, 2007, **46**, 5133–5136.
- 31 P. Mobian, J.-M. Kern and J.-P. Sauvage, *Angew. Chem., Int. Ed.*, 2004, **43**, 2392–2395.
- 32 J. Baram, E. Shirman, N. Ben-Shitrit, A. Ustinov, H. Weissman, I. Pinkas, S. G. Wolf and B. Rybtchinski, *J. Am. Chem. Soc.*, 2008, **130**, 14966–14967.
- 33 L.-Y. Yao and V. W.-W. Yam, *J. Am. Chem. Soc.*, 2015, **137**, 3506–3509.
- 34 Y.-F. Han, G.-X. Jin and F. E. Hahn, *J. Am. Chem. Soc.*, 2013, **135**, 9263–9266.
- 35 Y.-F. Han, G.-X. Jin, C. G. Daniliuc and F. E. Hahn, *Angew. Chem., Int. Ed.*, 2015, **54**, 4958–4962.
- 36 A. P. de Silva and N. D. McClenaghan, *Chem.–Eur. J.*, 2004, **10**, 574–586.
- 37 H. Li, J.-N. Zhang, W. Zhou, H. Zhang, Q. Zhang, D.-H. Qu and H. Tian, *Org. Lett.*, 2013, **15**, 3070–3073.
- 38 A. Credi, V. Balzani, S. J. Langford and J. F. Stoddart, *J. Am. Chem. Soc.*, 1997, **119**, 2679–2681.
- 39 G. McSkimming, J. H. R. Tucker, H. Bouas-Laurent and J.-P. Desvergne, *Angew. Chem., Int. Ed.*, 2000, **39**, 2167–2169.
- 40 G. de Ruiter and M. E. van der Boom, *Angew. Chem., Int. Ed.*, 2012, **51**, 8598–8601.
- 41 Y. Liu, A. Offenhäusser and D. Mayer, *Angew. Chem., Int. Ed.*, 2010, **49**, 2595–2598.
- 42 T. Gunnlaugsson, D. A. Mac Donail and D. Parker, *Chem. Commun.*, 2000, 93–94.
- 43 Z. Qi, P. Malo de Molina, W. Jiang, Q. Wang, K. Nowosinski, A. Schulz, M. Gradzielski and C. A. Schalley, *Chem. Sci.*, 2012, **3**, 2073–2082.
- 44 M. A. Saeed, H. T. M. Le and O. S. Miljanic, *Acc. Chem. Res.*, 2014, **47**, 2074–2083.
- 45 A. P. de Silva, N. H. Q. Gunaratne and C. P. McCoy, *Nature*, 1993, **364**, 42–44.
- 46 A. Romieu, *Org. Biomol. Chem.*, 2015, **13**, 1294–1306.
- 47 H. Bunzen, Nonappa, E. Kalenius, S. Hietala and E. Kolehmainen, *Chem.–Eur. J.*, 2013, **19**, 12978–12981.



- 48 V. E. Campbell, R. Guillot, E. Riviere, P.-T. Brun, W. Wernsdorfer and T. Mallah, *Inorg. Chem.*, 2013, **52**, 5194–5200.
- 49 Y. Wu, X.-P. Zhou, J.-R. Yang and D. Li, *Chem. Commun.*, 2013, **49**, 3413–3415.
- 50 D. H. Ren, D. Qiu, C. Y. Pang, Z. Li and Z. G. Gu, *Chem. Commun.*, 2015, **51**, 788–791.
- 51 X. Wu, N. Xu, Z. Zhu, Y. Cai, Y. Zhao and D. Wang, *Polym. Chem.*, 2014, **5**, 1202–1209.
- 52 P. D. Frischmann, V. Kunz and F. Würthner, *Angew. Chem., Int. Ed.*, 2015, **54**, 7285–7289.
- 53 D. Lewing, H. Koppetz and F. E. Hahn, *Inorg. Chem.*, 2015, **54**, 7653–7659.
- 54 T. K. Ronson, S. Zarra, S. P. Black and J. R. Nitschke, *Chem. Commun.*, 2013, **49**, 2476–2490.
- 55 P. D. Frischmann and M. J. MacLachlan, *Chem. Soc. Rev.*, 2013, **42**, 871–890.
- 56 J. Dömer, J. C. Slootweg, F. Hupka, K. Lammertsma and F. E. Hahn, *Angew. Chem., Int. Ed.*, 2010, **49**, 6430–6433.
- 57 A. E. Kaifer, S. Yi, V. Brega and B. Captain, *Chem. Commun.*, 2012, **48**, 10295–10297.
- 58 X. P. Zhou, J. Liu, S. Z. Zhan, J. R. Yang, D. Li, K. M. Ng, R. W. Y. Sun and C. M. Che, *J. Am. Chem. Soc.*, 2012, **134**, 8042–8045.
- 59 K.-C. Sham, S.-M. Yiu and H.-L. Kwong, *Inorg. Chem.*, 2013, **52**, 5648–5650.
- 60 X.-P. Zhou, Y. Wu and D. Li, *J. Am. Chem. Soc.*, 2013, **135**, 16062–16065.
- 61 P. D. Frischmann, V. Kunz, V. Stepanenko and F. Würthner, *Chem.–Eur. J.*, 2015, **21**, 2766–2769.
- 62 D.-H. Ren, D. Qiu, C.-Y. Pang, Z. Li and Z.-G. Gu, *Chem. Commun.*, 2015, **51**, 788–791.
- 63 D. M. Wood, W. Meng, T. K. Ronson, A. R. Stefankiewicz, J. K. M. Sanders and J. R. Nitschke, *Angew. Chem., Int. Ed.*, 2015, **54**, 3988–3992.
- 64 V. E. Campbell, X. de Hatten, N. Delsuc, B. Kauffmann, I. Huc and J. R. Nitschke, *Nat. Chem.*, 2010, **2**, 684–687.
- 65 C. S. Wood, T. K. Ronson, A. M. Belenguer, J. J. Holstein and J. R. Nitschke, *Nat. Chem.*, 2015, **7**, 354–358.
- 66 C. Browne, W. J. Ramsay, T. K. Ronson, J. Medley-Hallam and J. R. Nitschke, *Angew. Chem., Int. Ed.*, 2015, **54**, 11122–11127.
- 67 I. A. Riddell, Y. R. Hristova, J. K. Clegg, C. S. Wood, B. Breiner and J. R. Nitschke, *J. Am. Chem. Soc.*, 2013, **135**, 2723–2733.
- 68 R. Krämer, J. M. Lehn and A. Marquis-Rigault, *Proc. Natl. Acad. Sci. U. S. A.*, 1993, **90**, 5394–5398.
- 69 J. Xu, T. N. Parac and K. N. Raymond, *Angew. Chem., Int. Ed.*, 1999, **38**, 2878–2882.
- 70 S. E. Howson, A. Bolhuis, V. Brabec, G. J. Clarkson, J. Malina, A. Rodger and P. Scott, *Nat. Chem.*, 2012, **4**, 31–36.
- 71 C. R. K. Glasson, G. V. Meehan, J. K. Clegg, L. F. Lindoy, J. A. Smith, F. R. Keene and C. Motti, *Chem.–Eur. J.*, 2008, **14**, 10535–10538.
- 72 M. Albrecht, *Chem. Rev.*, 2001, **101**, 3457–3498.
- 73 R. Custelcean, *Chem. Soc. Rev.*, 2014, **43**, 1813–1824.

