Supramolecular aggregates of single-molecule magnets: exchange-biased quantum tunneling of magnetization in a rectangular [Mn₃]₄ tetramer†

Tu N. Nguyen, Wolfgang Wernsdorfer, Muhandis Shiddiq, Khalil A. Abboud, Stephen Hill and George Christou

The syntheses and properties of four magnetically-supramolecular oligomers of triangular Mn₃ units are reported: dimeric [Mn₆O₂(O₂CMe)₆(CH₂OH)₂(pdpd)₂] (3) and [Mn₆O₂(O₂CMe)₆(py)₂(pdpd)₂][ClO₄]₂ (4), and tetrameric [Mn₉O₄(O₂CMe)₁₂(pdpd)₁₂][ClO₄]₄ (R = Me (5), Bu (6)). They were all obtained employing 3-phenyl-1,5-dipyrindin-2-ylpentane-1,5-dione dioxide (pdpdH₂), either in direct synthesis reactions involving oxidation of Mn¹¹⁺ salts or in metathesis reactions with the preformed complex [Mn₆O₂(O₂CMe)₆(py)₂][ClO₄] (4); complex 6 was then obtained by carboxylate substitution on complex 5. Complexes 3 and 4 contain two [Mn₁₆̅₄(Mn₁₃̅₂-O)]⁺ and [Mn₁₆₄(Mn₁₃-O)]⁺ units, respectively, linked by two pdpd⁻² groups. Complexes 5 and 6 contain four [Mn₁₆̅₄(Mn₁₃-O)]⁺ units linked by six pdpd⁻² groups into a rectangular tetramer [Mn₆₄]₄. Solid-state dc magnetic susceptibility studies showed that the Mn₃ subunits in 3 and 4 have a ground-state spin of S = 3/2 and S = 2, respectively, while the Mn₃ subunits in 5 and 6 possess an S = 6 ground state. Complexes 5 and 6 exhibit frequency-dependent out-of-phase (x′′) ac susceptibility signals indicating 5 and 6 to be tetramers of Mn₃ single-molecule magnets (SMMs). High-frequency EPR studies of a microcrystalline powder sample of 5·2CH₂Cl₂ provided precise spin Hamiltonian parameters of D = −0.33 cm⁻¹, E = 0.03 cm⁻¹, B₄₂ = −8.0 × 10⁻⁵ cm⁻¹, and g = 2.0. Magnetization vs. dc field sweeps on a single crystal of 5·xCH₂Cl₂ gave hysteresis loops below 1 K that exhibit exchange-biased quantum tunneling of magnetization (QTM) steps with a bias field of 0.19 T. Simulation of the loops determined that each Mn₃ unit is exchange-coupled to the two neighbors linked to it by the pdpd⁻² linkers, with an antiferromagnetic inter-Mn₃ exchange interaction of J/K₈ = −0.011 K (J = −2JS₁S₂ convention). The work demonstrates a rational approach to synthesizing magnetically-supramolecular aggregates of SMMs as potential multi-qubit systems for quantum computing.

Introduction

Single-molecule magnets (SMMs) are individual molecules that function as single-domain nanoscale magnetic particles, and they bring the advantages of molecular chemistry to the field of nanoscale magnetism, such as crystallinity, monodispersity, solubility, and a shell of organic ligands that can be readily varied. Different from molecule-based magnets, which require intermolecular interactions and long-range ordering, SMMs are a class of metal clusters that display superparamagnetic behavior down to a characteristic blocking temperature, Tₘ, below which they exhibit slow magnetization dynamics. For Mn and most 3d transition metal SMMs, this behavior arises from the combination of a large ground-state spin (S) and Ising-type magnetoanisotropy (negative zero-field splitting parameter, D), which leads to a significant energy barrier to the thermal equilibration of the molecular magnetic moment. Experimentally, an SMM exhibits frequency-dependent out-of-phase ac magnetic susceptibility signals, and hysteresis in a plot of magnetization vs. applied dc magnetic field. The significant advantage of the single-molecule approach to nanomagnetism is the ability to investigate “particles” that are naturally identical and have well-defined sizes and electronic levels. Moreover, the possibility to modify the ligand sphere by standard solution chemistry methods allows for exquisite tuning of structures and magnetic properties of the molecules. Various families of SMMs have been discovered; the majority of which are Mn-based complexes with Mn₆₄ being the largest among them. SMMs have also been shown to display interesting quantum phenomena such as quantum tunneling of magnetization...
Our group has therefore initiated a program to link covalently two or more SMMs to give non-polymeric, magnetically-supermolecular aggregates-of-SMMs’ exhibiting very weak inter-SMM exchange interactions. The objective is to obtain materials exhibiting exchange-biased QTM and, in some cases, allow detailed study of quantum superposition states and related phenomena. For these reasons we are targeting small oligomers at this stage, in order to better detect the signatures of these nebulous and fragile effects. In the present work, we report the synthesis and properties of four oligomeric aggregates of Mn₆ units; two [Mn₃]₂ dimers and two [Mn₃]₄ tetramers, all containing dioximate inter-M₃₄ linker groups. We shall describe their structures and magnetic properties and, for one of the [Mn₃]₄ tetramers, the detailed analysis of high-frequency EPR and micro-SQUID magnetic data. Some preliminary results have been previously communicated.²⁷

Experimental section

Syntheses

All manipulations were performed under aerobic conditions using chemicals and solvents as received unless otherwise stated. [(Mn₆O₅(O₂CMe)₈(pdpd)₂(MeOH)₂)] (3)

A solution of Mn(O₂CMe)₂ · 4H₂O (0.25 g, 1.0 mmol) in 15 mL of MeCN/MeOH (2 : 1 v/v) was treated with NEt₃ (0.14 mL, 1.0 mmol) and pdpH₂ (0.18 g, 0.50 mmol). The solution was stirred for 1 hour at room temperature, filtered and the dark brown filtrate left undisturbed to concentrate slowly by evaporation. X-ray quality crystals of 3 · xMeCN · yMeCN slowly formed over one week. These were collected by filtration, washed with Et₂O, and dried under vacuum. The yield was 75% based on Mn. Anal. calcld (found) % for 3 · 4H₂O (Mn₆O₅C₁₈H₁₇O₁₈N₈): C 42.72 (42.39); H 4.54 (4.18); N 6.64 (6.25). Selected IR data (KBr, cm⁻¹): 1600 (s), 1474 (m), 1156 (w), 1173 |

[Mn₆O₅(O₂CMe)₈(pdpd)₂(MeOH)₂] [ClO₄]₂ (4)

A brown solution of complex 1 (0.44 g, 0.50 mmol) in 15 mL of MeCN/MeOH (2 : 1 v/v) was treated with pdpH₂ (0.18 g, 0.50 mmol). The solution was stirred for 1 hour at room temperature and then evaporated under vacuum to obtain black powder. The powder was dissolved in CH₂Cl₂ (25 mL) and the resulting solution was layered with Et₂O (50 mL). X-ray quality crystals of 4 · xEt₂O · yCH₂Cl₂ slowly formed over 2 days. The crystals were collected by filtration, washed with Et₂O and dried under vacuum. The yield was 80% based on Mn. Anal. calcld (found) % for 4 · 3H₂O (Mn₆O₅C₁₈H₁₇O₁₈N₈Cl₂): C 41.63 (41.79); H 3.90 (3.62); N 7.14 (6.77). Selected IR data (KBr, cm⁻¹): 1600 (s), 1474
A brown solution of complex 1 (0.044 g, 0.050 mmol) in 25 mL of CH₂Cl₂ was treated with pdpdH₂ (0.036 g, 0.10 mmol). The solution was stirred for 1 hour at room temperature, filtered, and the black filtrate left undisturbed. X-ray quality crystals of 5 · CH₂Cl₂ slowly formed over 2 days. The crystals were collected by filtration, washed with CH₂Cl₂ and dried under vacuum. The yield was 50% based on Mn. Anal. calcd (% for \(\text{Mn}_12\text{O}_{4\text{(O}_2\text{CMe)12(pdpd)6(ClO}_4)_4}\): C 43.99; H 3.59; N 8.10. Selected IR data (KBr, cm⁻¹): 1601 (s), 1584 (s), 1561 (s), 1516 (w), 1478 (s), 1440 (s), 1400 (s), 1384 (m), 1348 (s), 1219 (s), 1188 (m), 1165 (w), 1096 (s), 783 (m), 756 (w), 700 (m), 663 (m), 622 (s), 449 (m).

**Table 1** Crystal data and structure refinement parameters for complexes 3–6

<table>
<thead>
<tr>
<th>Parameter</th>
<th>3 · xMeOH·yMeCN</th>
<th>4 · xC₆H₄H₂O·yCH₂Cl₂</th>
<th>5 · xCH₂Cl₂</th>
<th>6 · xCH₂Cl₂</th>
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<tr>
<td>Formula</td>
<td>C₆H₄N₃Mn₆N₁₀O₂₆</td>
<td>C₆H₄Cl₁₀N₆M₆N₁₀O₁₂</td>
<td>C₁₅₀₂H₂₅₂Cl₄₄M₆N₂₄O₄₈</td>
<td>C₆H₅Cl₁₄M₆N₂₄O₄₈</td>
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<td>FW g mol⁻¹</td>
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<td>2419.84</td>
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<td>Crystal system</td>
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<td>Triclinic</td>
<td>Monoclinic</td>
<td>Monoclinic</td>
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<tr>
<td>Space group</td>
<td>C₄₆₃</td>
<td>P</td>
<td>C₁₅₀₂</td>
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<td>16.7979(18)</td>
<td>42.112(3)</td>
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<tr>
<td>b, Å</td>
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<tr>
<td>c, Å</td>
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<td>20.538(4)</td>
<td>41.03(4)</td>
<td>41.03(4)</td>
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<td>90</td>
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<td>(\beta), °</td>
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<td>(\gamma), °</td>
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<td>116.263(2)</td>
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<tr>
<td>V, Å³</td>
<td>8027.2(15)</td>
<td>4948.4(11)</td>
<td>46 413(9)</td>
<td>46 413(9)</td>
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<td>Z</td>
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<td>2</td>
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<td>4</td>
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<td>100(2)</td>
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<tr>
<td>(\lambda), Å⁻¹</td>
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<tr>
<td>(\mu), mm⁻¹</td>
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<td>1.145</td>
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<td>(R₁), %</td>
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<td>0.0732</td>
<td>0.0740</td>
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<tr>
<td>(wR₂), %</td>
<td>0.1276</td>
<td>0.2028</td>
<td>0.1617</td>
<td>0.1218</td>
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\[R₁ = \sum \left| F₁ - |F₁| \right|/\sum |F₁|, wR₂ = \left[ \sum \left| \left(F₁ - F₁\text{calc}ight)^2 \right|/\sum \left| F₁ \right|^2 \right]^{1/2}\]

\[w = 1/\left( \sigma^2(F₁) + (m \times p)^2 + n \times p \right)\]

\[^{a}\text{Including solvate molecules.}
[^{b}\text{Graphite monochromator.}
[^{c}\text{I} > 2\sigma(I).}
[^{d}\text{R₁ = }\sum \left| F₁ - |F₁| \right|/\sum |F₁|, wR₂ = \left[ \sum \left| \left(F₁ - F₁\text{calc}ight)^2 \right|/\sum \left| F₁ \right|^2 \right]^{1/2}\]
correction, which was subtracted from the experimental susceptibility to give the molar paramagnetic susceptibility ($\chi_M$). Low-temperature (−1.8 K) hysteresis loop and dc relaxation measurements were performed at Institut Néel using an array of micro-SQUIDS.78 The high sensitivity of this magnetometer allows the study of single crystals of SMMs of the order of 10–500 µm. The field can be applied in any direction by separately driving three orthogonal coils. High-frequency EPR (HF-EPR) powder spectra were collected at the Electron Magnetic Resonance facility of the U.S. National High Magnetic Field Laboratory using a broadband (10 to 700 GHz) transmission probe in which microwaves are propagated into and out of a 17 T superconducting magnet through cylindrical light pipes.79 Tunable microwave radiation was generated using a phase-locked Virginia Diodes solid-state source operating at 13 ± 1 GHz, followed by a chain of solid-state multipliers and amplifiers. Field modulation was employed in conjunction with a phase-sensitive homodyne detection scheme using a cold bolometer, yielding derivative mode spectra. Temperature control was achieved using helium gas flow cryostats.

Results and discussion

Syntheses

The synthetic strategies employed in this work are based on the report in 2007 that the use of the mono-oxime methyl-2-pyridyl ketone oxime (mpkoH; Fig. 1) in Mn$^{III}$ carboxylate chemistry gives the triangular complex $[\text{Mn}_6\text{O}_2\text{O}_2\text{CMe}]_3($mpko$)_3][\text{ClO}_4]$ (2), which has an S = 6 ground state and is an SMM.80 The cation of 2 has $C_3$ symmetry with the acetate and mpko$^-$ ligands on opposite sides of the Mn$_3$ plane (Fig. 1). This tripodal arrangement of the three oximates made this Mn$_3$ SMM attractive to us for our present objectives because it suggested that their replacement with dioximates to link multiple Mn$_3$ units together might give molecular oligomers rather than polymers.

After some planning, we settled on the dioxime 3-phenyl-1,5-dio(2-pyridin-2-yl)pentane-1,5-dione dioxime (pdpdH$_2$) for this work. It was designed taking into consideration the established binding modes of mpkoH, which has been well investigated in transition metal chemistry by many groups,81−83 and the decision to have a three-carbon linker between the two ends to impart sufficient length and flexibility to yield any potential oligomeric product without steric problems. The Ph ring at the central carbon atom was included to facilitate the synthesis, which comprised the reaction of two equivalents of 2-acetylpyridine with benzaldehyde to give the dione84 followed by a standard conversion to the dioxime with hydroxylamine. We recognized that the overall flexibility of pdpdH$_2$ meant that it was not ‘pre-programmed’ to yield any particular degree of oligomerization or product topology, but this was as desired at this initial stage in order to explore various possibilities.

Two procedures to 2 had been used$^{85}$ that could now be employed with pdpdH$_2$: (i) direct synthesis with simple reagents involving comproportionation between Mn$^{II}$ acetate and Mn$^{IV}$ in a 3 : 1 ratio in the presence of MeCO$_2$H and mpkoH, the Mn$^{III}$ : Mn$^{II}$ ratio of 3 : 1 giving a Mn$^{3.25+}$ average; and (ii) reaction of mpkoH with the non-SMM $[\text{Mn}_6\text{O}_2\text{O}_2\text{CMe}]_3($py$)_3][\text{ClO}_4]$ (1) in a 3 : 1 ratio (eqn (1)). Analogues of 2 with other carboxylates were also available from these methods using the corresponding starting materials.

$$[\text{Mn}_3\text{O}_2\text{CMe}]_3($py$)_3]^+ + 3\text{mpkoH} \rightarrow [\text{Mn}_3\text{O}_2\text{CMe}]_3($mpko$)_3]^+ + 3\text{RCO}_2\text{H} + 3\text{py} \quad (1)$$

The reactions between pdpdH$_2$ and a series of manganese salts were screened under a variety of conditions: solvent, stoichiometry, presence or absence of base, etc. Reaction of Mn$_2$O$_2$CMe$_2$, pdpdH$_2$ and NEt$_3$ in a 2 : 1 : 2 molar ratio in MeCN : MeOH (2 : 1 v/v) led to $[\text{Mn}_6\text{O}_2\text{O}_2\text{CMe}]_3($pdpd$_2$)$_2($MeOH)$_3][\text{ClO}_4]$ (3) (eqn (2)).

$$6\text{MnO}_2\text{CMe}_2 + 2\text{pdpdH}_2 + 2\text{MeOH} + 4\text{NEt}_3 + \text{O}_2 \rightarrow [\text{Mn}_6\text{O}_2\text{O}_2\text{CMe}]_3($pdpd$_2$)$_2($MeOH)$_3] + 4\text{MeCO}_2\text{H} + 4\text{Et}_3\text{NH}^+ + 4\text{H}_2\text{O} \quad (2)$$

3 is mixed-valence (2Mn$^{III}$, 4Mn$^{II}$) indicating incomplete oxidation of Mn$^{III}$ under these conditions. To explore whether an all-Mn$^{III}$ analogue was obtainable, we used the preformed Mn$^{III}$ complex 1. Indeed, the 1 : 1 reaction of pdpdH$_2$ with 1 in MeCN : MeOH (2 : 1 v/v) successfully gave $[\text{Mn}_6\text{O}_2\text{CMe}]_3($pdpd$_2$)$_2($py$_3$)$_2][\text{ClO}_4]$ (4), the all-Mn$^{III}$ version of 3 (eqn (3)). In both 3 and 4, each pdpd$^{2-}$ oxide group clips onto the same edge of each of the two Mn$_3$ triangular units, the other edges being bridged by only acetate groups (vide infra).86
2[Mn_{6}O(O_{2}CMe)_{14}(py)_{6}]^{3+} + 2pdpdH_{2} \rightarrow \\
[Mn_{6}O_{2}(O_{2}CMe)_{14}(py)_{2}(pdpd)]^{3+} + 4MeCO_{2}^{-} + 4pyH^{+} \ (3)

We now turned our attention to obtaining analogues of 2

with dioximate linkers bound to all edges of the triangular core

rather than mpko$. Given our observations with 3 and 4, we
decided to employ preformed 1 rather than Mn$^{II}$ salts. Thus, the

1 : 2 reaction of 1 and pdpdH_{2} in CH_{2}Cl_{2} was explored and this

led to [Mn_{3}O_{4}(O_{2}CMe)_{12}(pdpd)_{6}]^{4+} \ (eqn (4)).

4[Mn_{6}O(O_{2}CMe)_{14}(py)_{6}]^{3+} + 6pdpdH_{2} \rightarrow \\
[Mn_{6}O_{2}(O_{2}CMe)_{14}(py)_{2}(pdpd)]^{4+} + 12MeCO_{2}^{-} + 12pyH^{+} \ (4)

The excess of pdpdH_{2} was beneficial, the stoichiometric 2 : 3

ratio of eqn (4) giving a much lower yield. Ratios higher than

1 : 2 did not improve the yield further. The CH_{2}Cl_{2} solvent was

crucial for obtaining good crystals, the MeCN : MeOH (2 : 1 v/v)

medium, as used for 3 and 4, also giving 5 but as crystals unsuitable for crystallography.

With 5 attained and characterized (vide infra), it was of

importance for the magnetic studies to determine if its ligation

could be modified in a controlled fashion. We have previously

shown that the acetate groups of, for example, the [Mn_{12}O_{12}

(O_{2}CMe)_{12}(H_{2}O)_{12}] SMM can be substituted with numerous

other carboxylates,\footnote{K. Y. W. Ng, S. N. M. H. S. B. Rentziou,

M. J. Spinks, A. J. Bell, Chem. Sci. 2015, 6, 9740.} the substitution being driven to completion

by removal of MeCO_{2}H as its toluene azetrope under vacuum.

This controlled modification has been crucial for many studies

on SMMs. We have now found that a similar carboxylate

substitution using pivalic acid is possible for 5, without

breaking any of the dioximate linkages, to give 6 (eqn (5)).

This mild carboxylate substitution augurs well that controlled

modification of supramolecular aggregates of many types will

be possible in the future to optimize them for particular studies

or attachment to functional units or surfaces.

[Mn_{12}O_{4}(O_{2}CMe)_{12}(pdpd)_{6}]^{4+} + 12BuCO_{2}H \rightarrow \\
[Mn_{12}O_{2}(O_{2}CMe)_{12}(pdpd)]^{4+} + 12MeCO_{2}H \ (5)

Description of structures

The structures of 3 and the cation of 4 are shown in Fig. 2 and 3,

and selected interatomic distances and angles are listed in

ESL.$^{76}$ The compounds are essentially isostructural. Their cores

(Fig. 3) consist of two [Mn$^{II}$ Mn$^{III}$ Mn$^{III}$ (1)]$^{6+}$ (3) or [Mn$^{III}$ Mn$^{III}$ Mn$^{II}$ (1)]$^{4+}$ (4)

triangular subunits linked by two pdpd$^{2-}$ groups to give

[Mn_{3}]$^{0/2-}$ dimers. The O$^{2-}$ ions lie only slightly out of the Mn_{3}

planes in both complexes (d = 0.002–0.017 Å). Interestingly,

both pdpd$^{2-}$ linkers attach to the same Mn_{3} edge so that these

edges are bridged by two pyridyloximate groups whereas the

other edges are bridged by two acetate groups each. This is

distinctly different from 2 (Fig. 1) where each edge is bridged by

one mpko$^{-}$ and one acetate. The ligation is completed by

terminal MeOH groups on the Mn$^{II}$ atoms (Mn_{3}, Mn_{3}) in 3 and

pyridine groups on the Mn$^{III}$ ions at the corresponding positions

in 4. The Mn oxidation states were deduced from charge

balance considerations and the Mn$^{III}$ Jahn–Teller distortions

(Fig. 3), and confirmed by bond valence sum (BVS) calculations.

BVS calculations also confirmed that the $\mu_{3}$-O atoms are O$^{2-}$

ions.$^{76}$

The Mn_{3} triangles are far from equilateral, with Mn⋯Mn

separations and Mn–O$^{2-}$–Mn angles for the edges bridged by

two oximate groups being noticeably smaller than those for

edges bridged by two acetates. This is a characteristic of a diatomic N–O oximate bridge vs. a triatomic O–C–O carboxylate

bridge. The triangles are crystallographically scalene but

isosceles within the 3σ criterion. Overall, both dimers have C_{2v}

symmetry, either crystallographic (3) or virtual (4). Inspection of

short contacts reveals that each [Mn_{3}]$^{2-}$ dimer molecule of 3

connects to two neighbors by two hydrogen-bonds to each

(Oacet$^{-}$–OMeOH = 2.752(4) Å) to form a {[Mn_{3}]$_{2}$}$_{n}$ zig-zag chain

(Fig. S1†).$^{76}$

Fig. 2 Complete structure of 3 (top) and the cation of 4 (bottom) with H atoms omitted for clarity. Colour code: Mn$^{III}$ green; Mn$^{II}$ sky blue; O red; N blue; C grey.

Fig. 3 The cores of 3 (top) and the cation of 4 (bottom) emphasizing the inter-Mn3 connectivity, the Mn3 planes (grey triangles), and Jahn–Teller axes (green bonds). Colour code: Mn$^{III}$ green; Mn$^{II}$ sky blue; O red; N blue; C grey.
The structure of the cation of 5 is shown in Fig. 4. The cation of 6 (Fig. S2†) is very similar and will not be described separately. Selected \([\text{Mn}_3(\mu_3-O)]\) core distances and angles for the cation of 5 are listed in Table 2, and fuller listings for both cations are available in ESL.†‡ The asymmetric unit of 5 consists of two essentially superimposable \([\text{Mn}_3]_4^{4+}\) cations (only the one shown in Fig. 4 will be discussed further), eight \(\text{ClO}_4\) anions, and large amounts of disordered \(\text{CH}_2\text{Cl}_2\) solvent; the asymmetric unit of 6 contains only one cation. The cation consists of four \([\text{Mn}_3(\mu_3-O)]_7^{4+}\) units linked by six \(\text{pdpd}_2\) groups to give a rectangular \([\text{Mn}_3]_4\) supramolecule with each \(\text{Mn}_3\) unit at a corner. One of the two \(\mu_\text{-MeCO}_2\) ligands bridging each edge of 1 has been replaced by a bridging oximate from a \(\text{pdpd}^2\) group. The long and short sides of the rectangle are bridged by one and two \(\text{pdpd}^2\) groups, respectively (Fig. 4). In addition, the \(\text{pdpd}^2\) pyridyl groups have replaced the terminal pyridine group. The long and short sides of the rectangle are bridged by \(\text{pdpd}^2\) groups to give a molecular tetramer rather than a polymer. The flexibility of \(\text{pdpd}^2\), however, is likely why a rectangular rather than tetrahedral \([\text{Mn}_3]_4\) topology, as might be expected from four tripodal units and stiff linkers, has been obtained. The elemental analysis data for vacuum-dried 5 and 6 both indicate two \(\text{CH}_2\text{Cl}_2\) molecules, and this is consistent with the observation of two \(\text{CH}_2\text{Cl}_2\) molecules encapsulated within the central cavity of the \(\text{Mn}_{12}\) cations, with no opening big enough for them to escape (Fig. S3†).†‡

As seen also for 2, the \(\text{Mn}---\text{Mn}\) separations and \(\text{Mn}---\text{O}^{2-}---\text{Mn}\) angles span a small range and lead to each triangle being virtually isosceles within the 3σ criterion. Interestingly, both types of distortion from equilateral to isosceles are seen, i.e., short/long/long and short/short/long \(\text{Mn}---\text{Mn}\) separations, and corresponding small/large/large and small/small/large \(\text{Mn}---\text{O}---\text{Mn}\) angles. There are two of each type of \(\text{Mn}_3\) unit in the cation in Fig. 4 and Table 2. The \(\text{Mn}^{III}\) oxidation states of 4 at 5.0 K. Interactions established in the \(\text{Mn}_{12}\) tetramer? Variable-temperature, direct current (dc) magnetic susceptibility measurements were performed on vacuum-dried polycrystalline samples of 3–6 in an applied field of 1000 G (0.10 T) and in the 5.0–300 K temperature range. The samples were restrained in eicosane to prevent torquing. The data for 3·4\(\text{H}_2\text{O}\) and 4·3\(\text{H}_2\text{O}\) are shown as \(\chi_M T\) versus \(T\) in Fig. 5. \(\chi_M T\) for 3·4\(\text{H}_2\text{O}\) is 15.51 cm\(^3\) K mol\(^{-1}\) at 300 K, and steadily decreases with decreasing temperature to 3.32 cm\(^3\) K mol\(^{-1}\) at 5.0 K. 4·3\(\text{H}_2\text{O}\) displays a similar behavior, with a \(\chi_M T\) of 19.76 cm\(^3\) K mol\(^{-1}\) at 300 K, decreasing to 5.31 cm\(^3\) K mol\(^{-1}\) at 5.0 K. Interactions between the \(\text{Mn}_3\) subunits are expected to be very weak, and thus the \(\chi_M T\) at 300 K for 3·4\(\text{H}_2\text{O}\), which is much smaller than the spin-only (\(g = 2\)) value for four \(\text{Mn}^{III}\) and two \(\text{Mn}^{II}\) units.

![Fig. 4](image)

**Fig. 4** (Top) complete molecular structure of the cation of 5 with H atoms omitted for clarity; (bottom) the core emphasizing the connectivity, orientation of the \(\text{Mn}_3\) planes (blue triangles), and the Jahn–Teller axes (green bonds). Colour code: \(\text{Mn}^{III}\) green; O red; N blue; C grey.

**Table 2** Selected \([\text{Mn}_3(\mu_3-O)]\) core distances (Å) and angles (°) in the cation of 5.

<table>
<thead>
<tr>
<th>Distance (Å)</th>
<th>Angle (°)</th>
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<tbody>
<tr>
<td>Mn1⋯Mn3</td>
<td>3.184(2)</td>
</tr>
<tr>
<td>Mn1⋯Mn2</td>
<td>3.212(2)</td>
</tr>
<tr>
<td>Mn2⋯Mn3</td>
<td>3.201(2)</td>
</tr>
<tr>
<td>Mn3⋯Mn5</td>
<td>3.189(2)</td>
</tr>
<tr>
<td>Mn4⋯Mn6</td>
<td>3.173(2)</td>
</tr>
<tr>
<td>Mn5⋯Mn6</td>
<td>3.216(2)</td>
</tr>
<tr>
<td>Mn7⋯Mn8</td>
<td>3.180(2)</td>
</tr>
<tr>
<td>Mn7⋯Mn9</td>
<td>3.206(2)</td>
</tr>
<tr>
<td>Mn8⋯Mn9</td>
<td>3.205(2)</td>
</tr>
<tr>
<td>Mn10⋯Mn11</td>
<td>3.185(2)</td>
</tr>
<tr>
<td>Mn10⋯Mn12</td>
<td>3.197(2)</td>
</tr>
<tr>
<td>Mn11⋯Mn12</td>
<td>3.218(2)</td>
</tr>
</tbody>
</table>

**Magnetoochemistry**

**Direct current magnetic susceptibility studies.** In addition to general magnetic characterization of 3–6, there were two crucial questions targeted at 5 (and 6) that needed to be addressed: (i) is each \(\text{Mn}_3\) subunit still an SMM with \(S = 6\), or has the aggregation altered the properties vs. 2? and (ii) is there any evidence for weak inter-\(\text{Mn}_3\) interactions within the \([\text{Mn}_3]_2\) tetramer? Variable-temperature, direct current (dc) magnetic susceptibility (\(\chi_M\)) measurements were performed on vacuum-dried polycrystalline samples of 3–6 in an applied field of 1000 G (0.10 T) and in the 5.0–300 K temperature range. The samples were restrained in eicosane to prevent torquing. The data for 3·4\(\text{H}_2\text{O}\) and 4·3\(\text{H}_2\text{O}\) are shown as \(\chi_M T\) versus \(T\) in Fig. 5. \(\chi_M T\) for 3·4\(\text{H}_2\text{O}\) is 15.51 cm\(^3\) K mol\(^{-1}\) at 300 K, and steadily decreases with decreasing temperature to 3.32 cm\(^3\) K mol\(^{-1}\) at 5.0 K. 4·3\(\text{H}_2\text{O}\) displays a similar behavior, with a \(\chi_M T\) of 19.76 cm\(^3\) K mol\(^{-1}\) at 300 K, decreasing to 5.31 cm\(^3\) K mol\(^{-1}\) at 5.0 K. Interactions between the \(\text{Mn}_3\) subunits are expected to be very weak, and thus the \(\chi_M T\) at 300 K for 3·4\(\text{H}_2\text{O}\), which is much smaller than the spin-only (\(g = 2\)) value for four \(\text{Mn}^{III}\) and two \(\text{Mn}^{II}\) units.
Table 3  Selected structural parameters for the [Mn₃(µ₃-O)] triangles of 5 and 6

<table>
<thead>
<tr>
<th>Complex 5</th>
<th>Triangle</th>
<th>$d^a$</th>
<th>Torsion angle$^b$</th>
<th>$\phi$ (°)</th>
<th>Complex 6</th>
<th>Triangle</th>
<th>$d^a$</th>
<th>Torsion angle$^b$</th>
<th>$\phi$ (°)</th>
</tr>
</thead>
<tbody>
<tr>
<td>M₃N₂M₃</td>
<td>0.31</td>
<td>M₃N₁⁻O₀M₃</td>
<td>17.98</td>
<td></td>
<td>M₃N₃M₃</td>
<td>0.31</td>
<td>M₃N₁⁻O₃M₃</td>
<td>14.38</td>
<td></td>
</tr>
<tr>
<td>M₂N₁−O₁M₁</td>
<td>14.49</td>
<td>M₃N₂⁻O₁M₂</td>
<td>10.90</td>
<td></td>
<td>M₂N₄⁻O₄M₁</td>
<td>16.60</td>
<td>M₃N₆⁻O₅M₂</td>
<td>13.30</td>
<td></td>
</tr>
<tr>
<td>M₃N₂⁻O₁M₂</td>
<td>10.90</td>
<td>M₃N₅⁻O₅M₅</td>
<td>19.82</td>
<td></td>
<td>M₅N₁⁻O₇M₅</td>
<td>18.31</td>
<td>M₅N₁⁻O₇M₄</td>
<td>18.31</td>
<td></td>
</tr>
<tr>
<td>M₅N₁⁻O₇M₄</td>
<td>18.28</td>
<td>M₅N₅⁻O₈M₆</td>
<td>10.75</td>
<td></td>
<td>M₅N₆⁻O₈M₅</td>
<td>12.64</td>
<td>M₅N₆⁻O₈M₄</td>
<td>12.64</td>
<td></td>
</tr>
</tbody>
</table>

$^a$ Displacement (Å) of µ₃-oxide atom from the Mn₃ plane. $^b$ Mn–N–O–Mn torsion angle (°).

non-interacting ions (20.75 cm³ K mol⁻¹), indicates dominant or wholly antiferromagnetic (AF) interactions within the Mn₃ subunits. Complex 4, however, has $\chi_M^T$ value at 300 K slightly higher than the spin-only (g = 2) value for six Mn²⁺ non-interacting ions ($\chi_M^T = 18.00$ cm³ K mol⁻¹), which together with the overall profile of the plot indicates both ferromagnetic (F) and AF interactions. The 5.0 K values are consistent with dimers of essentially non-interacting S = 3/2 and S = 2 subunits for 3 and 4, respectively, with g slightly less than 2.0; the spin-only values would be 3.75 cm³ K mol⁻¹ and 6.00 cm³ K mol⁻¹, respectively. In both cases, any inter-Mn₃ interactions would only contribute to the $\chi_M^T$ decreases barely visible at the lowest T.

The $\chi_M^T$ vs. T plots for 5·2CH₂Cl₂ and 6·2CH₂Cl₂ are essentially identical (Fig. 6), as expected from their near-identical structures. $\chi_M^T$ for 5·2CH₂Cl₂ increases from 48.25 cm³ K mol⁻¹ at 300 K to a plateau value of 76.55 cm³ K mol⁻¹ at 20 K, and then decreases slightly to 70.62 cm³ K mol⁻¹ at 5.0 K. For 6·2CH₂Cl₂, $\chi_M^T$ is 46.34 cm³ K mol⁻¹ at 300 K, increasing to 76.57 cm³ K mol⁻¹ at 25 K before slightly decreasing to 70.97 cm³ K mol⁻¹ at 5.0 K. The overall profiles are extremely similar to that of monomer 2, indicating each of the four Mn₃ units of 5 and 6 to still be ferromagnetically coupled with S = 6 ground states. Indeed, the 300 K values of 5·2CH₂Cl₂ and 6·2CH₂Cl₂ are much larger than the spin-only (g = 2) value for twelve Mn²⁺ atoms (36 cm³ K mol⁻¹), and the peak values are as expected for four ostensibly non-interacting S = 6 units with g slightly less than 2.0 (spin-only value 84 cm³ K mol⁻¹). The decreases below 20 K can be attributed to zero-field splitting (ZFS) within the S = 6 ground states, weak inter-Mn₃ interactions, and any Zeeman effects from the applied field.

The data were fit to the theoretical $\chi_M^T$ vs. T expressions to obtain the exchange interaction parameters ($J_{ij}$) between Mn₁-Mn₃ pairs. We assumed the interaction through the bridging pdpd⁻¹ to be too weak to affect the data except at the lowest temperatures, and thus fit the data assuming non-interacting Mn₃ units; however, the H-bonding between different dimers of 3 should give somewhat stronger inter-Mn₃ interactions (vide infra). For complexes 3/4 and 5/6, the theoretical $\chi_M^T$ is the sum for two and four independent Mn₃ units, respectively. In accord
with the mixed-valence nature of 3 and the Mn₃ structural parameters for 4–6, we used an isosceles triangle model for each Mn₃ unit. For convenience, we label this triangle in a general way as Mn1Mn2Mn2', with Mn2Mn2' being the unique edge. The Heisenberg–Dirac–van Vleck (HDVV) spin Hamiltonian is then given by eqn (6), where $J = J_{12} = J_{12}'$ and $J = J_{22}'$. Using the

$$\hat{H} = -2J(\hat{S}_1 \cdot \hat{S}_2 + \hat{S}_1 \cdot \hat{S}_2) - 2J\hat{S}_2 \cdot \hat{S}_2 \tag{6}$$

substitutions $\hat{S}_A = \hat{S}_2 + \hat{S}_2$ and $\hat{S}_T = \hat{S}_1 + \hat{S}_1$, where $\hat{S}_T$ is the total spin of the Mn₃, the spin Hamiltonian becomes that in eqn (7).

$$E(S_T, S_A) = -J[S_T(S_T + 1) - S_A(S_A + 1)] - J[S_A(S_A + 1)] \tag{7}$$

The eigenvalues are given in eqn (8), where $E(S_T, S_A)$ is the energy of state $S_T$ arising

$$\hat{H} = -J[S_T^2 - S^2 - S^2] - J[S_A^2 - S^2 - S^2]$$

from $S_A$. For 3, $S_1 = 5/2$, $S_2 = 2$, and $S_T$ ranges from 1/2 to 13/2. For 4–6, $S_1 = S_2 = S = 2$, and $S_T$ ranges from 0 to 6. A theoretical $\chi_MT$ vs. $T$ expression was derived from the use of the van Vleck equation$^{78,76}$ and addition of a temperature-independent paramagnetism (TIP) term, which was kept constant at $100 \times 10^{-6}$ cm$^3$ K mol$^{-1}$ per Mn. The data at the lowest $T$ were omitted because they were affected not by factors included in eqn (6), particularly the inter-dimer H-bonding in 3 (vide supra). The obtained fits ($R^2 > 0.99$) are shown as solid lines in Fig. 5 and 6, and the fit parameters are collected in Table 4. For 4, 5 and 6, there are at least two crystallographically inequivalent Mn₃ subunits, and the $J$ and $J'$ are to be taken as average values.

For both 3 and 4, $J$ and $J'$ are AF and F, respectively. The $J$ vs. $f$ error surfaces for the fits, obtained with the program GRID,$^{78}$ showed them to be unique (Fig. S5 and S6†). The obtained $J$ and $J'$ reveal 3 to have an $S = 3/2$ ground state with an $S = 5/2$ first excited state at 49.7 cm$^{-1}$. 4 has an $S = 2$ ground state and an $S = 3$ first excited state at 29.6 cm$^{-1}$. The presence of both $J$ and $F$ interactions is as expected from the structures of 3 and 4. The bis-acetate-bridged edges correspond to the AF $J$ interactions, while the bis-oximate-bridged edge corresponds to the F $J'$ interaction. This is consistent with previous studies of (i) [Mn₃(O₃CR)₆L₉] complex (2Mn III/Mn II and 3Mn II),$^{77}$ where all couplings are AF, and (ii) mono-oximate bridged 2, where all couplings are F.$^{80}$ In 3 and 4, two oximates now bridge the $J'$ (Mn III) edge and the coupling is again F, consistent with the non-zero oximate Mn–O–N–Mn torsion angles in 3 (av 17.1°) and 4 (av 15.7°), one of the recognized effects that promote F coupling between Mn III atoms.$^{100,101}$ Interestingly, on

conversion of the Mn II atoms in 3 to Mn III in 4, and the shortening of the resulting Mn–O bond, the $J'$ coupling increases by nearly a factor of three even though its Mn–O–N–Mn torsion angles decrease to 15.7°. This emphasizes the importance of not just the torsion angles but also of perturbations to coupling pathways through the central O$^-$ ion to the overall F $J'$ coupling in 3 and 4, and other O$^-$-bridged Mn₃ triangles, as we have discussed elsewhere.$^{102}$

For 5 and 6, both $J$ and $J'$ are F. However, the error surface for 5 in Fig. 7 (and for 6 in Fig. S8†) show two minima of near-equivalent quality located either side of the $J$ vs. $J'$ line. These two minima are assigned to the two different distortions from equilateral to isosceles, as discussed above. Since the crystal structure of the cation of 5 shows that both distortions are present, we conclude that both $J$/$J'$ pairs are present and contribute to the observed magnetic properties. In our previous study of monomer 2 and its other carbonylate analogues,$^{86}$ only one type of isosceles triangle was observed crystallographically and therefore only the best fit $J$/$J'$ pairs were reported ($J = 12.1$–$18.6$ cm$^{-1}$ and $J' = 1.5$–$6.7$ cm$^{-1}$), which are very similar to one of the fits in the present work.$^{80}$ Further, the two fits for 5 and 6 each have the same average within the fit uncertainties: $J_{av} = 11.8 \pm 0.1$ cm$^{-1}$ for 5, and $J_{av} = 10.3 \pm 0.1$ cm$^{-1}$ for 6. In all cases, the Mn₃ units have an $S = 6$ ground state, with an $S = 5$ first excited state. The first excited states lie 79.5 and 84.1 cm$^{-1}$ above the ground states of 5 and 6, respectively, and the ground states are thus well-isolated.

The ground states of the Mn₃ units in 5 and 6 were confirmed by fits of magnetization ($M$) data collected in the 0.1–7 T and 1.8–10 K ranges. The data, plotted as $M$/$N\mu_B$ vs. $H/T$ in Fig. S9† ($N$ is Avogadro’s number and $\mu_B$ is the Bohr magneton) were fit, as described for 2,$^{86}$ using the program MAGNET.$^{77}$ by diagonalization of the spin Hamiltonian matrix assuming only the ground state is populated, incorporating axial anisotropy ($D_3S^2$) and Zeeman terms, and employing a full powder average.$^{103,104}$ We assumed each Mn₃ to be acting independently, i.e., that any

Table 4  Fit parameters for the fits of $\chi_MT$ vs. $T$ data for 3–6

<table>
<thead>
<tr>
<th>Complex</th>
<th>$J$ (cm$^{-1}$)</th>
<th>$J'$ (cm$^{-1}$)</th>
<th>$g$</th>
<th>$S$</th>
</tr>
</thead>
<tbody>
<tr>
<td>3-H₂O</td>
<td>-9.9 ± 0.5</td>
<td>10.4 ± 3.9</td>
<td>1.96 ± 0.04</td>
<td>3/2</td>
</tr>
<tr>
<td>4-H₂O</td>
<td>-4.9 ± 0.1</td>
<td>28.4 ± 2.0</td>
<td>1.99 ± 0.01</td>
<td>2</td>
</tr>
<tr>
<td>5-2CH₃Cl₂</td>
<td>16.8 ± 0.6</td>
<td>1.5 ± 0.7</td>
<td>1.91 ± 0.01</td>
<td>6</td>
</tr>
<tr>
<td>6-2CH₃Cl₂</td>
<td>7.2 ± 0.7</td>
<td>21.4 ± 0.6</td>
<td>1.91 ± 0.01</td>
<td>6</td>
</tr>
<tr>
<td>7.2 ± 0.8</td>
<td>17.6 ± 0.8</td>
<td>1.91 ± 0.01</td>
<td>6</td>
<td></td>
</tr>
</tbody>
</table>

Fig. 7  2-D contour plot of the error surface for the $J$ vs. $J'$ fit of the dc $\chi_MT$ vs. $T$ data for complex 5. The darkest-brown region has the lowest error, where the two fits (asterisks) of comparable quality are located, and the blue region has the highest error. See Table 4 for the fit parameters.
inter-Mn$_3$ interactions through pdpd$^{2-}$ linkers were too weak to have a noticeable effect at these fields and temperatures. In support of the latter, we obtained excellent fits with $S = 6$, $g = 1.92(1)$, $D = -0.30(2)$ cm$^{-1}$ for 5 and $S = 6$, $g = 1.89(1)$, $D = -0.30(2)$ cm$^{-1}$ for 6. These are very similar to those for 2 ($S = 6$, $D = -0.34$ cm$^{-1}$, and $g = 1.92$). The $g$ vs. $D$ error surfaces (Fig. S10†) revealed alternative fits with positive $D$ but of much poorer quality.

**Alternating current (ac) magnetic susceptibility studies.** Ac magnetic susceptibility measurements were performed in the 1.8–15 K range in a 3.5 G ac field oscillating at 5–1000 Hz. The in-phase ($\chi'_{M}$, plotted as $\chi_{MT}$) and out-of-phase ($\chi''_{M}$) ac susceptibility data for 5 are shown in Fig. 8; those for 3, 4 and 6 are in Fig. S11–S13. In each case, the data at >10 K or so are in agreement with the ground state $S$ values deduced from the dc fits, providing independent support for those conclusions. For 5, for example, the near-plateau $\chi'_{MT}$ of $\sim$75 cm$^3$ K mol$^{-1}$ is in good agreement with the dc peak value and is as expected for four $S = 6$ Mn$_3$ units with $g < 2.0$ (spin-only value $4 \times 21$ cm$^3$ K mol$^{-1}$). At lower $T$, $\chi'_{MT}$ drops due to zero-field splitting and any inter-Mn$_3$ interactions. Below 2.5 K, a characteristic frequency-dependent drop due to slow magnetization relaxation of a SMM is observed, and out-of-phase $\chi''_{M}$ signals are seen. This is very similar to the behaviour previously observed for monomer 2. For 3 and 4, extrapolation of the $\chi'_{MT}$ data from above 8 K to 0 K gives $\sim$3.6 and $\sim$5.3 cm$^3$ K mol$^{-1}$, respectively, consistent with the values expected for two independent $S = 3/2$ and $S = 2$ units (spin-only values 3.75 and 6.00 cm$^3$ K mol$^{-1}$, respectively). No $\chi''_{M}$ signals were observed, and thus 3 and 4 are not SMMs. The out-of-phase $\chi''_{M}$ signals of complexes 5 and 6 are tails of peaks lying below 1.8 K, suggesting 5 and 6 might be tetramers of Mn$_3$ SMMs.

**High-frequency electron paramagnetic resonance (HF-EPR) studies.** In order to more precisely determine the spin-Hamiltonian parameters of 5, HF-EPR data were collected for a finely ground sample of 5-CH$_2$Cl$_2$ that was incorporated into a KBr pellet in order to avoid field-alignment of the micro-crystallites within the powder. Measurements were performed at multiple high frequencies in the 52.4 to 422.4 GHz range, and at temperatures from 3.5 to 20 K.

Temperature dependent EPR spectra collected at 217.6 GHz are presented in Fig. 9. The two very sharp features (indicated by asterisks) at 3.2 and 5.5 T have been well characterized and are attributed to paramagnetic oxygen impurities trapped in the KBr pellet. The sharp feature observed around the isotropic $g = 2.00$ position (marked by ”*” at 7.77 T) corresponds to an impurity phase within the powder, possibly containing Mn$^{3+}$; this is often found to be the case in Mn-containing polynuclear clusters. The spectra were recorded in field-derivative mode (dI/dB, where $I$ denotes the absorption intensity), making it relatively easy to determine which peaks correspond to which components of the spectrum. The inset to Fig. 9a displays

![Fig. 9](image_url)

**Fig. 9** Temperature dependence of HF-EPR spectra collected on a microcrystalline sample of complex 5-CH$_2$Cl$_2$ (experimental (a) and simulated (b)) restrained in KBr; the spectra were recorded in field derivative mode at 217.6 GHz in the 3.5 to 20 K range. The features in (a) are labeled according to the scheme described in the main text. The top inset in (a) depicts a typical derivative mode powder spectrum for a biaxial system, illustrating the features expected for the $x$-, $y$-, and $z$-components of the spectrum. The features in (a) marked by * and $\times$ are attributed to known impurities.
a trivial example of a $T = 0$ derivative mode powder spectrum for a biaxial $S = 1/2$ system with three distinct $g$-tensor components; the onset of absorption appears as a peak in the derivative mode spectrum, and corresponds to the easy-axis of magnetization (largest $g$ component); the cessation of absorption thus appears as a dip in the derivative, corresponding to the hard-axis (smallest $g$ component); meanwhile, the intermediate component of the $g$-tensor occurs at the maximum in absorption and, therefore, looks like the derivative of the $z$-component. The same general behavior is expected for high-spin systems that experience zero-field splitting (ZFS) interactions, and examination of the resonances on the low (high) field side of $g = 2.00$ in Fig. 9a indicates that they are indeed peaks (dips). In the case of a uniaxial ZFS tensor, the parallel ($B||z$) component of the powder spectrum typically extends about twice as far from the $g = 2.00$ position, compared to the perpendicular ($B \perp z$) component, with the extent of the deviation from $g = 2.00$ being directly proportional to the magnitude of the axial $D$ parameter. On this basis, the multiple low-field features in Fig. 9a can be attributed to resolved parallel ($B||z$) excitations. Careful inspection of the 20 K spectrum reveals six such peaks ($z_1$ to $z_6$) on the low field side of $g = 2.00$. This is an indication of an $S = 6$ ground state for the Mn$_3$ units. The peaks correspond to the following fine-structure transitions within the ground $S = 6$ spin multiplet: $m_S = -6 \rightarrow -5 (z_6), m_S = -5 \rightarrow -4 (z_5), \ldots$, and $m_S = -1 \rightarrow 0 (z_1)$. The fact that the spectral weight associated with the parallel spectrum shifts to the low field, $m_S = -6 \rightarrow -5 (z_6)$, transition upon cooling provides confirmation of the negative sign of $D$, i.e., the $m_S = -6$ state is lowest in energy when $B||z$.

In contrast to the parallel spectrum, the fine structure transitions within the perpendicular spectrum are not resolved. There are several possible reasons for this. First of all, the relevant peaks should be much more closely spaced (two perpendicular components, each with half the separation of the parallel spectrum) and, therefore, more difficult to resolve. Nevertheless, this alone cannot account for the observed spectrum. Another complicating factor is the existence of the two distinct isosceles triangles (with short/short/long and long/long/short edges) within the tetramer (vide supra), which one would expect to have different ZFS parameters, with the difference likely being more pronounced for the transverse 2nd-order anisotropy parameter, $E$. The transverse anisotropy results from a projection of the anisotropies of the individual Mn$^{III}$ ions onto the plane of the triangle. The 2nd-order projection is strictly zero ($E = 0$) for the perfectly equilateral case, i.e., there is an exact cancelation. However, for the isosceles case, the projections onto the plane of the triangle do not cancel, and the value of $E$ can be expected to be extremely sensitive to the distortion away from a perfect equilateral triangle. This contrasts the case for the axial molecular ZFS parameter, $D$, which results from an addition of the single-ion anisotropies. The effect of the two triangular topologies is therefore likely to be quite similar to a disorder that disproportionately affects the transverse anisotropy parameter, $E$. This type of disorder is well known in other Mn$^{III}$ containing clusters, including the Mn$_{12}$-acetate SMM, where it has been shown to significantly impact the perpendicular spectrum without affecting the parallel components significantly. These influences can be taken into consideration when simulating spectra using a program such as EasySpin.108 We note that the temperature dependent studies (Fig. 9) do not provide any strong indications that there are low-lying states with $S = 5, 4, \ldots$, i.e., no additional peaks emerge upon raising the temperature that cannot be accounted for via simulations that assume an isolated $S = 6$ ground state.

Fig. 9b displays the simulation obtained from the spin Hamiltonian in eqn (9), where $\hat{S}$ and $\hat{S}_i$ ($i = x, y, z$) are spin operators,

$$\hat{H} = \mu_B \vec{B} \cdot \vec{S} + D(S_z^2 - \frac{1}{3} S(S+1)) + B_z^0 \hat{O}_z$$

where $\vec{g}$ is the Landé $g$ tensor, $D$ and $E$ are the second-order axial and rhombic ZFS parameters, respectively, and the final term represents an axial fourth-order ZFS interaction.108 The parameters used for the simulation are $S = 6, D = 0.33 \text{ cm}^{-1}, |E| = 0.030 \text{ cm}^{-1}, B_z^0 = -8.0 \times 10^{-5} \text{ cm}^{-1}$, and $g_x = g_y = g_z = 2.00$. As can be seen, the relevant parallel and perpendicular portions of the simulations agree with the experiments very well, and the obtained $D$ value is consistent with that deduced from the magnetic studies. Gaussian distributions of the $D$ and $E$ parameters ($\sigma_D = 8\%$, $\sigma_E = 40\%$) were employed in order to reproduce the observed line broadening. As expected, the strain in $E$ is much more pronounced than that of $D$, which explains the almost complete loss of resolution of the transverse part of the EPR spectrum.

In order to obtain tighter constraints on the spin Hamiltonian parameters for complex SMMs, the positions of parallel component peaks obtained at multiple frequencies are displayed on a 2D frequency vs. field plot in Fig. 10. The combined data set was then simulated using the Hamiltonian of eqn (9); the solid lines in Fig. 10 represent the best simulation of the spectral features observed for 5-xCH$_2$Cl$_2$ at multiple frequencies in the 52.4 to 422.4 GHz range. Simulated transition curves employing eqn (9) are included for $B\parallel(x, blue), y \parallel (green)$ and $z \parallel (red)$. See the text for the simulation parameters.

**Fig. 10** 2-D plot of frequency vs. the field locations of powder EPR spectral features observed for 5-xCH$_2$Cl$_2$ at multiple frequencies in the 52.4 to 422.4 GHz range. Simulated transition curves employing eqn (9) are included for $B\parallel(x, blue), y \parallel (green)$ and $z \parallel (red)$. See the text for the simulation parameters.
simulations, accounting for the uneven field spacing of the resonances. One can estimate a kinetic barrier ($U$) to magnetization relaxation of $\sim 12$ cm$^{-1}$ from the $S$ and $D$ values or $\sim 15$ cm$^{-1}$ if the effect of the significant $B_0^2$ is included.

We conclude this section by discussing the interaction between the four Mn$_3$ units in the tetramer. EPR spectra for the powder sample suggest that the Mn$_3$ units are essentially uncoupled, or the coupling is too weak to detect by EPR. Hysteresis loop data suggest an antiferromagnetic coupling of $-0.0076$ cm$^{-1}$ (vide infra) between adjacent Mn$_3$ units within the tetramer. This coupling is smaller than the $D$ and $E$ strains ($\sigma_D = 0.026$ cm$^{-1}$ and $\sigma_E = 0.012$ cm$^{-1}$), which means that the spectral broadening (FWHM = 2.35σ) is likely to far exceed any additional splittings of the EPR spectra induced by the inter-SMM exchange. Consequently, it cannot be detected by EPR.

The picture that emerges from the combined magnetic and HF-EPR studies on 5 and 6 described above is that magnetically-supramolecular aggregates of the Mn$_3$ SMMs with $S = 6$ can be obtained using the dioximate pdp$_2^-$, and that these do indeed contain Mn$_3$ units that still retain the structural and magnetic properties of the monomeric analogue 2. Any inter-Mn$_3$ exchange interactions through the dioximate groups are either zero or too weak to affect the magnetic data collected above 1.8 K. The tetramers 5 and 6 therefore behave as an assembly of independent Mn$_3$ units and their data can be fit as such to give fit parameters that are essentially identical with those previously obtained for 2. The crucial question now needing to be addressed is whether there are indeed any inter-Mn$_3$ interactions present, and if so what is their effect? The answers to these questions required studies at even lower temperatures.

**Magnetization hysteresis and decay studies.** Magnetization ($M$) vs. dc field scans on a single crystal of 5-xCH$_2$Cl$_2$ were carried out on a micro-SQUID using field sweep rates in the 0.001–0.280 T s$^{-1}$ range.$^{74}$ Hysteresis loops were observed below $\sim 1.0$ K (Fig. 11) whose coercivities increase with decreasing temperature and increasing field sweep rate, as expected for SMMs. The loops also exhibit a number of steps characteristic of quantum tunnelling of magnetization (QTM). The blocking temperature is $\sim 1.0$ K, above which M relaxes faster than the time scale of the hysteresis measurement. We therefore conclude that each of the four Mn$_3$ units in 5-xCH$_2$Cl$_2$ is an SMM, as previously found for monomer 2, which they closely resemble.

Before analysing the QTM step pattern, we describe the determination of the effective relaxation barrier ($U_{\text{eff}}$) from magnetization decay measurements. A large dc magnetic field was applied to the crystal at $\sim 5$ K to saturate the magnetization in one direction, and $T$ decreased to a chosen value in the 0.03–1.0 K range. The field was then removed and $M$ measured as a function of time (Fig. 12, inset). From these data, relaxation time ($\tau$) vs. $T$ data were extracted and used to construct an Arrhenius plot based on eqn (10) (Fig. 12), where $k$ is the Boltzmann constant. Below $\sim 0.3$ K, $\tau$ is temperature-independent

$$\ln(\tau) = \ln(\tau_0) + \frac{U_{\text{eff}}}{kT}$$

and all relaxation is by QTM via the lowest energy $m_s = \pm 6$ levels of the ground state $S = 6$ manifold. Above $\sim 0.3$ K, the relaxation is thermally activated, and a fit of the linear region at higher $T$ (dashed line in Fig. 12) gave $U_{\text{eff}} = 8.5$ cm$^{-1}$ ($U_{\text{eff}}/k_B = 12.2$ K) and $\tau_0 = 5 \times 10^{-9}$ s. The $U_{\text{eff}}$ is smaller than the upper limit to the barrier ($U = S|D| ~ 11.9$ cm$^{-1}$ ($U/k_B = 17.1$ K) calculated using $D$ from the EPR data. This is as expected for QTM via
a thermally (phonon) assisted pathway involving higher-energy \( m_S \) levels of the \( S = 6 \) ground state.

We now address whether the four Mn\(_3\) SMM units in 5·xCH\(_2\)Cl\(_2\) are weakly interacting with each other. The answer is clearly yes because the hysteresis loops show an exchange-bias (shift) of the QTM steps. The first step in the hysteresis loop of an SMM on scanning from negative to positive fields is normally at zero field, where the \( m_S \) levels on either side of the anisotropy barrier are in resonance and QTM can occur, reversing the orientation of the magnetization vector. The presence of an AF exchange-coupled neighbor provides a bias field that shifts the QTM step (resonance tunneling) to a new field position before zero. This was first seen for the hydrogen-bonded [Mn\(_4\)]\(_2\) dimer of \( S = 9/2 \) SMMs, and a related explanation is valid for 5, except that each Mn\(_3\) SMM will be seen to be exchange-coupled to two neighboring Mn\(_3\) SMMs via the bridging pdpd\(^{2-}\) groups.

The loops for 5·xCH\(_2\)Cl\(_2\) in Fig. 11 show the first QTM step before zero field, confirming weak AF interactions between the Mn\(_3\) subunits. However, there is still a step at zero field and at a slow sweep rate of 0.001 T s\(^{-1}\), a third step is evident after zero field giving a three-step pattern with equal step separations that is symmetric about zero field. This is unlike anything seen previously in exchange-biased dimers. At faster sweep rates, the step patterns change markedly, the symmetric three-step pattern being lost and new steps appearing. The overall loop profiles and their sweep-rate dependence point to a more complicated exchange-coupled system than any encountered previously and which needed to be carefully explained.

To obtain a simplified spin Hamiltonian to model the [Mn\(_3\)]\(_4\) tetramer, each Mn\(_3\) SMM can be treated as a giant spin of \( S = 6 \) with easy-axis anisotropy. The spin Hamiltonian (\( \hat{H} \)) for each Mn\(_3\) is given by eqn (11), where \( i = 1-4 \) refers to the four Mn\(_3\) SMMs of the tetramer (Fig. 13), \( \mu_0 \) is the vacuum permeability,

\[
\hat{H}_i = D\hat{S}_{z,i}^2 + \hat{H}_{\text{trans},i} + g\mu_B\mu_0\hat{S}_{z,i}H_z
\]

and \( H_z \) is the applied longitudinal field. The last term is the Zeeman energy associated with the applied field. The exact form of the transverse anisotropy \( \hat{H}_{\text{trans},i} \) is not important in this discussion. The general Hamiltonian for the tetramer is then given by eqn (12), where \( J_{ij} \) are exchange coupling parameters between subunits \( i \) and \( j, S_i = S_j = 6 \), and the exchange interactions are assumed to be isotropic (\( J_{ij} = J_{ji} \)).

Since 5 is a rectangular [Mn\(_3\)]\(_4\) aggregate, and assuming the interactions are superexchange couplings through the pdpd\(^{2-}\) linkers (this assumption will be confirmed later), there should be two different inter-Mn\(_3\) interactions, \( J_1 \) and \( J_2 \), which are likely to be comparable but not identical in magnitude; the diagonal interaction (\( J_{11} \)) should be essentially zero and can be ignored. Therefore, the spin Hamiltonian is given in eqn (13).

\[
\hat{H} = \sum\hat{H}_i - 2J_{12}(\hat{S}_1\cdot\hat{S}_2 + \hat{S}_3\cdot\hat{S}_4)
\]

Tunnelling among the \( (2S + 1)^4 = 28 \) 561 energy states is allowed by the small transverse anisotropy \( \hat{H}_{\text{trans},i} \) and the transverse coupling terms containing \( \hat{S}_{xi} \) and \( \hat{S}_{yi} \) operators. Calculating the energy states of [Mn\(_3\)]\(_4\) by exact diagonalization would be a major undertaking, and for convenience we instead neglect all transverse anisotropy terms. The spin Hamiltonian in eqn (13) can then be rewritten as eqn (14).

\[
\hat{H} = \hat{H}_1 + \hat{H}_2 + \hat{H}_3 + \hat{H}_4 - 2J_1(\hat{S}_1\cdot\hat{S}_2 + \hat{S}_3\cdot\hat{S}_4) - 2J_2(\hat{S}_1\cdot\hat{S}_4 + \hat{S}_2\cdot\hat{S}_3)
\]

Since \( J_1 = J_3 \), we shall assume for now that \( J_1 = J_2 = J \). The eigenvalues of eqn (14) are then given by eqn (15).

\[
E = D(m_{S1}^2 + m_{S2}^2 + m_{S3}^2 + m_{S4}^2) + g\mu_B\mu_0H_z(m_{S1} + m_{S2} + m_{S3} + m_{S4}) - 2J(m_{S1}m_{S2} + m_{S3}m_{S4} + m_{S1}m_{S4} + m_{S2}m_{S3})
\]
rectangular topology of 5. Coupling to one or three neighbors should exhibit no steps at zero field since there will always be a net bias present from the odd number of neighbors; only for two neighbors can the net bias be zero – namely, in the case when their two effects cancel out, as will be described below.

The spin-state energy vs. applied dc field was calculated using eqn (15) for all the states of 5. It was then adjusted by varying the magnitude of $J$ and $D$ to simulate the experimental hysteresis loop at 0.001 T s$^{-1}$. The resulting plot (Fig. 14, top) was obtained with $J = -0.0076 \text{ cm}^{-1}$ ($\hat{H}/k_\theta = -0.011 \text{ K}$) ($\hat{H} = -2J\hat{S}_i\hat{S}_j$ convention) and $D = -0.33 \text{ cm}^{-1}$. When the spin Hamiltonian contains transverse terms ($\hat{H}_{\text{trans}}$), the level crossings can be ‘avoided level crossings’. The spin $S$ is ‘in resonance’ between two states when the local longitudinal field is close to an avoided level crossing, and tunneling may occur. The plot shows all avoided level crossings at which QTM transitions could be seen, and we can now explain the three-step hysteresis loop at 0.001 T s$^{-1}$ in Fig. 11, bottom.

At the lowest $T$ of 0.04 K, only the lowest energy states involving $m_S = \pm 6$ states will be populated (red lines in Fig. 14). As the field is scanned from $-1$ T, where the four Mn$_3$ magnetization vectors will have all been polarized into the $m_S = -6$ orientation (Fig. 13a), the following is seen:

(a) the first QTM step corresponds to tunneling of a Mn$_3$ vector from the $m_S = -6$ to the $m_S = +6$ state; this occurs at $-0.19$ T, which thus represents the total bias field from two $m_S = -6$ neighbors. In the format $(m_{S1}, m_{S2}, m_{S3}, m_{S4})$, where $i = 1-4$ are the four Mn$_3$ SMMs within 5, the $-0.19$ T step corresponds to the $(-6, -6, -6, -6)$ to $(-6, +6, -6, -6)$ tunneling transition (Fig. 13b); for convenience, other possible equivalent product states, $(+6, -6, -6, -6), (-6, -6, +6, +6),$ and $(-6, -6, -6, +6),$ are not shown;

(b) the second step at zero field is assigned to tunneling of molecules with $-6, +6$ neighbors ($S_1$ and $S_3$ in Fig. 13b), which therefore experience a zero bias (if $J_1 = J_2$, or a small bias related to $|J_1 - J_2|$ if $J_1 \neq J_2$). The step at zero field is thus the $(-6, +6, -6, -6)$ to $(-6, +6, +6, -6)$ tunneling transition, which will then have some probability of undergoing a spin flip-flop relaxation to the $(-6, +6, +6, +6)$ ground state since $J$ is AF (Fig. 13c);

(c) If enough vectors have tunneled in the first two steps to $+6$ (i.e., if the sweep rate is slow enough), then a third step is expected for Mn$_3$ vectors tunneling in the presence of a $+6, +6$ bias field ($S_1$ and $S_3$ in Fig. 13c). This would be symmetrically located with respect to the first step, and should thus occur at $+0.19$ T.

The three-step pattern at 0.001 T s$^{-1}$ sweep rate (in Fig. 11, bottom) is thus as predicted at this slow scan rate, which is required to see situation (c). Note that the avoided level crossings at positions marked with asterisks might at first glance be predicted to give additional QTM steps, but these involve multiple (2, 3, or 4) tunneling transitions at the same time, e.g. $(-6, -6, -6, -6)$ to $(-6, +6, +6, +6)$, $(-6, -6, -6, -6)$ to $(-6, +6, +6, +6)$, etc. They thus have a very low probability of occurring and in fact we see no evidence for their occurrence. Note that double tunneling transitions also occur at $\pm 0.19$ T and 0 T, i.e., at the avoided level crossings where the single transitions are seen as steps in the loops.

At faster scan rates, the three-step hysteresis loop changes markedly. On scanning from $-1$ T at 0.28 T s$^{-1}$, the first step has now decreased in size, as has the second step at zero field, and the third step has disappeared completely. This is as expected from standard Landau–Zener theory,$^{109,110}$ which predicts the tunneling probability decreases with increasing field scan rate, i.e., fewer molecules are in resonance long enough for their spins to tunnel. As these steps decrease or disappear, new steps appear in the 0.0–0.5 T range as the magnetization relaxes via alternative tunnelling pathways; some steps appear and disappear as the scan rate increases (e.g. at 0.33 T), and by 0.28 T s$^{-1}$, two broad features at $\pm 0.4$ T and $\pm 0.6$ T are evident. We assign these tunneling transitions to levels involving one $m_S = \pm 5$ state, i.e., at positions involving avoided level crossings of red and blue lines in Fig. 14. These include $(-6, -6, -6, -6)$ to $(-6, +6, -6, -5)$ at $\pm 0.1$ T, $(-6, +6, -6, -6)$ to $(-6, +6, +6, -5)$ at $\pm 0.2$ T, $(-6, +6, -6, -6)$ to $(-6, +6, +5, -5)$ at $\pm 0.3$ T, $(-6, +6, +5, +5)$ at $\pm 0.6$ T, and others.

It should be noted that all the steps in Fig. 11 are rather broad compared with, e.g., those for monomer 2 (Fig. S14†).
This is due to the two independent cations in the asymmetric unit having different orientations, and the four Mn₃ planes within each cation not being coplanar (Fig. 4, bottom). As a result, the applied field will be at a range of angles to the z-axes (easy-axes) of the eight Mn₃, leading to step broadening. This is unfortunate because had the steps been sharper, we might have been able to resolve the different tunneling transitions that occur at different sweep rates in the hysteresis loops. We did attempt to collect micro-SQUID data on single crystals of 6·xCH₂Cl₂ but they were more fragile than those of 5·xCH₂Cl₂ and we could not get satisfactory data.

As a further test of the validity of the above model and Fig. 14, we carried out hysteresis minor-loop sweeps from −1 T to various fields up to zero and then reversed the sweep direction. The objective was to scan through the avoided level crossing centered at −0.19 T to allow some molecules to tunnel from (−6, −6, −6, −6) to (−6, +6, −6, −6), i.e., step (a) above, and then on reversing the sweep to see the (−6, +6, −6, −6) to (−6, −6, −6, −6) QTM transition at just under −0.55 T (Fig. 15, inset). The separation between these two QTM steps is proportional to the energy difference between the (−6, −6, −6, −6) and (−6, −5, −6, −6) states, i.e., it depends on D. The desired step was observed, although broad due to the factors discussed above. Nevertheless, a field separation of ΔH = 0.4 T could be estimated. Using eqn (16) gives D = −0.37 cm⁻¹ with g = 2.0, in satisfying agreement with values obtained from the magnetization and HF-EPR fits, and thus supporting the overall validity of the model and the analysis of the hysteresis loops. The value of D suggests the population of Mn₃ units sampled in the minor-loop scans had an average angle between their z-axes and the applied field of ~28°.

$$|D| = g\mu_B\Delta H$$  \hspace{1cm} (16)

**Validity of assumptions.** The above discussion explains the sweep rate dependence of the hysteresis loops for 5·xCH₂Cl₂, and the general conclusion reached is that each Mn₃ SMM is interacting with its two neighbors in the rectangle to which it is linked by the bridging pdpd²⁻ linkers. We now address two important follow-up questions: (i) is it safe to assume that the two neighbors with which each Mn₃ is interacting are those connected to it by the pdpd²⁻ linkers? And (ii) is the assumption that $J_1 = J_2$ valid, and how would we know if it is not?

It could be argued that the pdpd²⁻ linkers are not the superexchange pathways through which the Mn₃ interact. Might the interactions instead be dipolar between a Mn₃ and nearby neighbors (intra- or inter-tetramer), or via inter-tetramer π-stacking or weak H-bonds? However, 5·xCH₂Cl₂ contains no significant inter-tetramer contacts, other than some C–H···π and π···π contacts involving the central Ph ring of the pdpd²⁻ group, and many C–H···Cl H-bonds to the solvent molecules. Nevertheless, we compared inter-tetramer distances and contacts in 5·xCH₂Cl₂ crystals with those in 2·3CH₂Cl₂. Shortest inter-Mn₃ Mn–Mn distances are comparable (Fig. S4†), ~6.0 and ~6.6 Å in 5 and 2, respectively, and similar C–H···π, π···π, and C–H···Cl contacts are present in both. In fact, 2 contains π···π stacking between the pyridyl rings on neighboring Mn₃ units (3 does not) and thus might be expected to give stronger inter-Mn₃ exchange couplings than in 5. Thus, if the inter-Mn₃ couplings that give the exchange-biased QTM in 5 were dipolar in origin or superexchange through any of the mentioned contacts, we would expect to see similar (or even larger) shifts of the QTM steps in 2. In fact, the hysteresis loops of 2·3CH₂Cl₂ show no discernible shift of the zero field step (Fig. S14†). We thus conclude that the only significant inter-Mn₃ coupling leading to the exchange-biased QTM in 5·xCH₂Cl₂ is indeed superexchange through the pdpd²⁻ linkers.

The rectangular structure of 5 suggests $J_1 \neq J_2$, although both should be very weak. We therefore calculated what might be expected in this case, and whether we might be able to identify this difference experimentally. A representative energy plot calculated with a reasonable $J_1/J_2 = 0.7$ ratio is shown in Fig. S15,† where it is also compared with the $J_1 = J_2$ plot. The only change to the three-step hysteresis loop of the $J_1 = J_2$ situation is the step at zero-field splitting into two, with splitting proportional to $|J_1 - J_2|$. The other steps still occur symmetrically about zero at ±($J_1 + J_2$). With better step resolution, we might have been able to pick up a splitting in the zero-field peak, but the broadness of the steps due to the range of magnetic field alignments, as described above, precludes this. We thus conclude that even though it is likely that $J_1 \neq J_2$, the difference must be too small to have a noticeable effect on the hysteresis loops within the experimental step broadness.

**Conclusions.**

We summarize below the main results and conclusions of this study:

1. Employment of a designed dioxime group has allowed the isolation from straightforward procedures of two non-SMM [Mn₃]₃, and two SMM [Mn₃]₄ aggregates; the latter serve to
achieve the overall objective of this work, to obtain covalently-linked oligomers of SMMs to extend previous work beyond the dimeric systems studied in the past. The dioxime was not programmed to give any particular oligomer topology, and a rectangular one was obtained.

(2) Each Mn₃ SMM subunit in [Mn₃]₄ tetramers has retained the structure and magnetic properties of the monomeric analogue, encouraging us that the oligomerization methodology has the potential for general application to a range of linkers without damage to the units or their SMM properties.

(3) Carboxylate substitution on preformed [Mn₃]₄ tetramer 5 has been accomplished without rupture of the oligomer linkages, providing a convenient means for post-synthetic modification of oligomers for comparative studies or certain applications.

(4) Lower dioxime : Mn ratios afford [Mn₃]₂ dimers at two oxidation levels, [MnⁿMnˢ]ₓ and [Mnⁿ]ₓ, with asymmetric ligation not previously observed, leading to a mix of F and AF exchange interactions in both complexes.

(5) The magnetic analysis confirms that each Mn₃ unit is still ferromagnetically-coupled with an S = 6 ground state, and there is no evidence in the data at 1.8 K and above for any inter-Mn₃ exchange-interactions. In addition, both types of distortion of the Mn₃ triangles from equilateral to isosceles have been seen and magnetically analyzed, characterizing two best-fit minima in the error surface and the degree of difference caused to the constituent J values.

(6) Detailed multi-frequency HF-EPR spectra at a range of temperatures, and their simulations, have provided the intra-Mn₃ magnetic properties, and addressing the di complications from dipolar or other inter-oligomer pathways.

(7) Hysteresis loops obtained on a single crystal of 5 -xCH₂Cl₂ established that the four Mn₃ SMMs are each weakly coupled to two neighbors, and this interaction is manifested as an exchange-bias of the QTM steps, with the exact positions of the latter determined by the spin alignments at these two neighbors. The exchange-biased hysteresis loops at 0.04 K and 0.001 T s⁻¹ scan rates have been explained by identification of the ground state quantum transitions taking place.

(8) Equally important, the large changes in QTM step locations at higher scan rates have been explained as a function of decreasing tunneling probability with increasing scan rate, resulting in relaxation via QTM transitions involving higher energy levels.

(9) Using the hysteresis loop of the analogous Mn₃ monomer as a control, a crucial conclusion has been reached that the inter-Mn₃ interaction is via superexchange through the bridging dioximate linkers. This shows that intra-oligomer couplings and quantum effects can be studied in the solid state without complications from dipolar or other inter-oligomer pathways.

(10) The validity of the overall model used to describe the quantum behavior has been tested by simulations of the 1- or 2J coupling model, and also by hysteresis minor-loop studies that have been shown to give a Mn₃ D value comparable to those from the magnetization fits and HF-EPR studies.

Overall, the described work represents a rational approach to obtaining covalently-connected multiple Mn₃ SMMs that is now being extended to other oligomers with various structural topologies and quantum mechanical couplings in order to assess their potential for use as multi-qubit assemblies and other systems based on SMMs.

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References


51 The term ‘magnetically-supramolecular’ is based on the general definition of supramolecular chemistry as “…chemical systems made up of a discrete number of assembled molecular subunits or components. The forces responsible for the spatial organization may vary from weak (intermolecular forces, electrostatic or hydrogen bonding) to strong (covalent bonding), provided that the degree of electronic coupling between the molecular components remains small with respect to relevant energy parameters of the component.”22 In the present situation, ‘magnetically-supramolecular’ means covalently-linked magnetic molecular components (SMMs) with inter-component exchange interactions that are very weak with respect to the intra-SMM interactions.


Although almost all lattice CH2Cl2 molecules are disordered, the two encapsulated ones are ordered and...


