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Rh-catalyzed direct synthesis of 2,2'-dihydroxybenzophenones and xanthenes†

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An efficient rhodium-catalyzed direct synthesis of 2,2'-dihydroxybenzophenones and xanthenes was developed from functionalized salicylaldehydes. This approach provides an easy access to various functionalized 2,2'-dihydroxybenzophenone and xanthone core skeletons. This study also revealed the crucial role of the hydroxy group in the reductive homo-coupling process to generate 2,2'-dihydroxybenzophenones. Overall the outcome of the reaction course was also found to be influenced by the electronics of the substituent groups in salicylaldehydes.

Introduction

Dihydroxybenzophenone is a medicinally important structural motif with various biological activities.^{1,2} Some of the derivatives have been used as sunscreen additives³ and as intermediates in the synthesis of chiral ligands.⁴ Furthermore, symmetrically substituted xanthone derivatives⁵ have been used as precursors in the synthesis of fluorophores^{5a,b} and in fluorescent dyes.^{5c} Additionally, xanthone such as persulfated 3,6-(*O*- β -glucopyranosyl)xanthone (**E**) was known as dual inhibitor of FXa (anticoagulant agent) and with antiplatelet effect.⁶ Some of the relevant 2,2'-dihydroxybenzophenone and xanthone skeletons are given in Fig. 1.

In this context, methods involving the synthesis of 2,2'-dihydroxybenzophenones are noteworthy and rely on the Fries rearrangement of *O*-acylphenols,⁷ Lewis acid promoted arylation of phenols⁸ and hydrolysis of xanthone derivatives.^{1a} In addition, some of the metal-catalyzed approaches known for the synthesis of 2,2'-dihydroxybenzophenones and xanthenes are given in Scheme 1. In literature, the direct hydroxylation of *ortho*-selective C–H oxygenations involving Pd, Cu, Ru or Rh catalysts were reported for the synthesis of 2,2'-dihydroxybenzophenones.^{9–11}

Li and co-workers reported one-step synthesis of xanthone derivatives from aryloxybenzaldehydes involving cross-dehydrogenative coupling process.¹² Furthermore, Peng and co-workers reported the synthesis of xanthenes from 2-hydroxybenzaldehyde and 1,2-dibromoarene under Pd-catalyzed conditions.¹³ Further, 2-hydroxybenzophenone derivatives with suitable functional groups such as bromo, methoxy and nitro groups were used for the synthesis of xanthenes in the

presence of copper catalyst.¹⁴ Further, Li group also showed the role of various phosphine ligands in the oxidative decarbonylation of aromatic aldehydes to give biaryls and arylketones.¹⁵ However, this reactivity with *ortho*-substituted aromatic aldehydes was found to be non-selective. In addition, the study of aromatic aldehydes with *ortho*-hydroxy substitution *i.e.* salicylaldehydes is conspicuously missing in the literature. However, in some instances *ortho*-hydroxy group was known to retard decarbonylation of aromatic aldehydes under rhodium-catalyzed conditions^{16–19} and this could be one reason for the lack of study with salicylaldehydes in decarbonylative couplings. These substrates were effectively used in

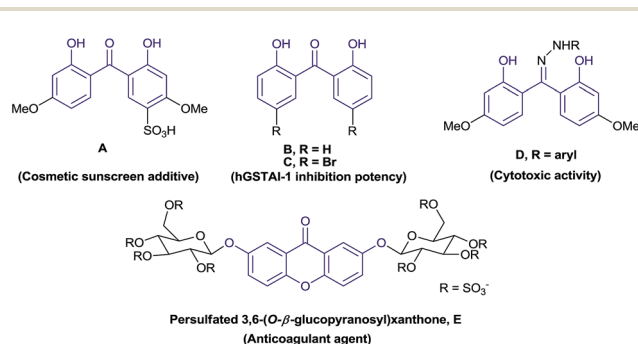
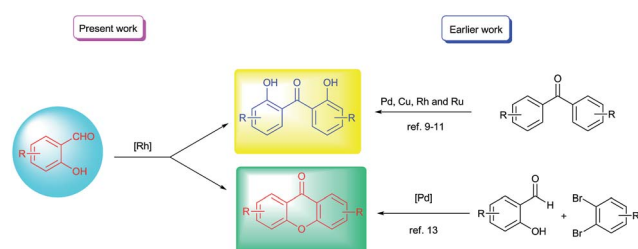


Fig. 1 Biologically active molecules.



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hydroacylations of activated alkenes^{16,18} and alkynes.¹⁷ Against this background, in our recent study with salicylaldehydes under rhodium-catalyzed conditions, we found the novel reductive decarbonylative couplings of salicylaldehydes to give symmetrically substituted 2,2'-dihydroxybenzophenones and xanthenes selectively under different reaction conditions (Scheme 1). It is our pleasure to report these results.

Results and discussion

In our study using salicylaldehyde (**1a**) under rhodium-catalyzed conditions (Table 1), initially we observed the formation of 2,2'-dihydroxybenzophenone (**2.1**) in 52% yield (entry 1). This prompted us to standardize the condition for the direct synthesis of 2,2'-dihydroxybenzophenone from salicylaldehyde.

To achieve this, we carried out a thorough systematic screening. For example, this reaction in *N,N*-dimethylacetamide (DMA) and *N*-methyl-2-pyrrolidone (NMP) as solvents furnished lower yields of 2,2'-dihydroxybenzophenone (entries 2 and 3). Further screening was done with change in base employing K₃PO₄ or K₂CO₃ and these reactions gave 28% and 43% yields respectively (entries 4 and 5). Additional use of triphenylphosphine did not improve the yield (entry 6). A few more additives were screened (entries 7–9) out of which, Cu(OAc)₂·H₂O (1 equiv.) was found to be relatively more effective giving 67% yield (entry 9). Encouragingly, this yield was further increased to 80% with 2 equiv. of Cu(OAc)₂·H₂O (entry 10). At this stage, various Rh-catalysts were also screened to test their efficacy (entries 11–13) and this effort proved to be ineffective.

Table 1 Screening conditions^{ab}

Entry	Catalyst	Base	Additive (equiv.)	Solvent	Yield 2.1 (%)
1	Rh(CO) ₂ (acac)	Na ₂ CO ₃	—	DMF	52
2	Rh(CO) ₂ (acac)	Na ₂ CO ₃	—	DMA	28
3	Rh(CO) ₂ (acac)	Na ₂ CO ₃	—	NMP	39
4	Rh(CO) ₂ (acac)	K ₃ PO ₄	—	DMF	28
5	Rh(CO) ₂ (acac)	K ₂ CO ₃	—	DMF	43
6	Rh(CO) ₂ (acac)	Na ₂ CO ₃	PPh ₃ (0.1)	DMF	26
7	Rh(CO) ₂ (acac)	Na ₂ CO ₃	Et ₃ N (1)	DMF	48
8	Rh(CO) ₂ (acac)	Na ₂ CO ₃	TBAB (1)	DMF	32
9	Rh(CO) ₂ (acac)	Na ₂ CO ₃	Cu(OAc) ₂ (1)	DMF	67
10	Rh(CO) ₂ (acac)	Na ₂ CO ₃	Cu(OAc) ₂ (2)	DMF	80
11	RhCl(PPh ₃) ₃	Na ₂ CO ₃	Cu(OAc) ₂ (2)	DMF	63
12	{RhCl(cod)} ₂	Na ₂ CO ₃	Cu(OAc) ₂ (2)	DMF	56
13	RhCl ₃ ·3H ₂ O	Na ₂ CO ₃	Cu(OAc) ₂ (2)	DMF	67
14	Rh(CO) ₂ (acac)	—	Cu(OAc) ₂ (2)	DMF	63
15	Rh(CO) ₂ (acac)	—	—	DMF	24
16	—	Na ₂ CO ₃	Cu(OAc) ₂ (2)	DMF	—

^a Conditions: salicylaldehyde (0.5 mmol, 1 equiv.), catalyst (0.025 mmol, 0.05 equiv.), base (1 mmol, 2 equiv.), additive, 100 °C, solvent (3 mL), 16 h. ^b Isolated yields.

Table 2 Synthesis of symmetrical 2,2'-dihydroxybenzophenones^{ab}

Entry	Salicylaldehyde	Product	Yield (%)
1			80 ^c
2			76 ^c
3			68 ^c
4			73 ^d
5			76
6			65
7			51
8			46 ^e
9			33 (26)
10			30 (20) ^f
11			43 (23) ^f

^a Conditions: salicylaldehyde (0.5 mmol, 1 equiv.), Rh(CO)₂(acac) (0.025 mmol, 0.05 equiv.), Na₂CO₃ (1 mmol, 2 equiv.), Cu(OAc)₂·H₂O (1 mmol, 2 equiv.), 100 °C, DMF (3 mL), 24 h. ^b Isolated yields are given. Xanthone product yields are given in parenthesis. ^c 16 h. ^d At 120 °C. ^e 2,4-Dichlorophenol was formed with 10% yield. ^f Without using Na₂CO₃.



A control reaction carried out without base provided lowered yield (entry 14). Additional controls without base/additive or catalyst gave either poor or no product formation (entries 15 and 16). From this study, it was found that salicylaldehyde (1 equiv.), Rh(CO)₂(acac) (0.05 equiv.), Na₂CO₃ (2 equiv.), Cu(OAc)₂·H₂O (2 equiv.) at 100 °C in DMF as an effective protocol for the decarbonylative coupling of salicylaldehyde (**1a**) to give 2,2'-dihydroxybenzophenone (**2.1**).

Using this optimized protocol, further scope and reactivity of differently substituted salicylaldehydes was studied (Table 2). This study employing electron-rich salicylaldehydes (**1b–1f**) was found to be high yielding in comparison to the corresponding reactivity obtained with electron-deficient aryl aldehydes (**1g–1k**). Thus, salicylaldehyde and other derivatives of it with 4-methoxy, 3-methoxy, 4-*N,N*-diethylamino, 5-methyl and 3,4,5-trimethoxy groups gave the corresponding functionalized 2,2'-dihydroxybenzophenones in 65–80% yields (**2.1–2.6**). As given in Fig. 2, the product **2.2** was also confirmed by X-ray analysis.

Moreover, salicylaldehydes substituted with 5-bromo, 5-chloro, 3,5-dichloro, 5-fluoro and 5-acetyl groups gave respective 2,2'-dihydroxybenzophenones in 30–51% yields (**2.7–2.11**). In some of these cases, we observed the formation of xanthone as a cyclized product in minor amounts (given in the parenthesis, Table 2).

This prompted us to explore whether it is possible to obtain xanthone as a major product directly from salicylaldehydes in a one-pot operation. Hence, we performed this reaction at 100 or 120 °C. These results are summarized in Table 3. The formation of 2,2'-dihydroxybenzophenone and its *in situ* cyclization to give xanthone was found to be effective under the conditions studied to give moderate yields (**3.1–3.10**). This effort gave various functionalized xanthenes containing formyl, bromo, chloro, fluoro, methyl, acetyl, nitro, hydroxy and trimethoxy groups. One of the positive factors to be highlighted is that salicylaldehydes substituted with electron-withdrawing groups reacted at 100 °C to give the corresponding xanthenes in moderate yields. Thus, salicylaldehyde substrates at C-5 substitution with chloro, acetyl, formyl and nitro groups got cyclized to xanthenes at 100 °C (**3.6–3.9**). Other substrates with 5-methyl, 4-hydroxy, 5-fluoro and 3,4,5-trimethoxy groups were cyclized at 120 °C to give xanthone products (**3.2–3.4** and **3.10**). Further, the obtained reactivity was found to be site-selective as we obtained xanthone **3.8** with formyl group intact using **1l** substrate. This indicated the preferential reaction of formyl group *ortho* to hydroxy during decarbonylative homo-coupling process. The presence of formyl and halo groups in xanthone products endowed them with additional synthetic advantage for further transformations.

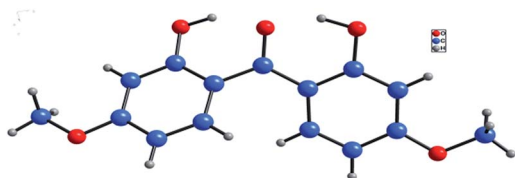


Fig. 2 X-ray structure of **2.2**.

Table 3 Synthesis of symmetrically substituted xanthenes^{ab}

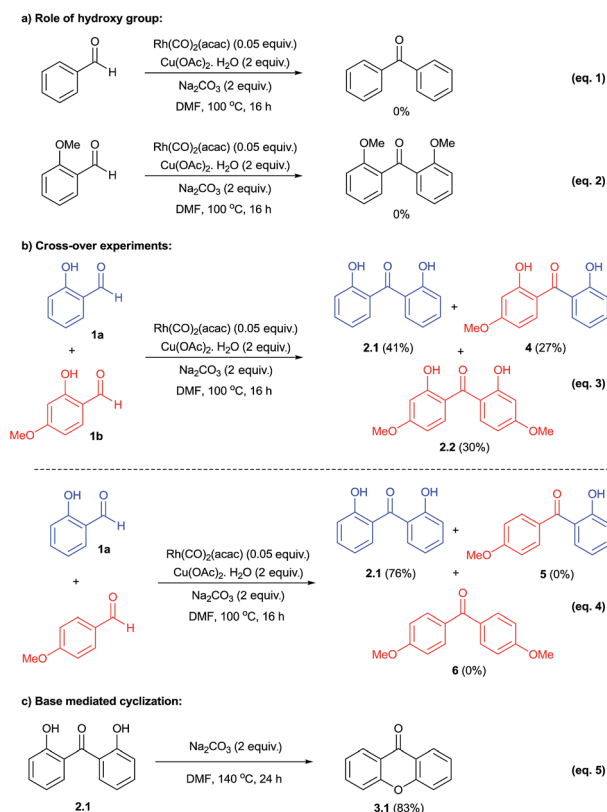
Entry	Salicylaldehyde	Product	Yield (%)
1			47 ^c
2			39 ^d
3			35 ^d
4			45 ^c
5			46
6			50 (54)
7			54 (50)
8			46 (54)
9			59 (57)
10			49 ^c

^a Conditions: salicylaldehyde (0.5 mmol, 1 equiv.), Rh(CO)₂(acac) (0.025 mmol, 0.05 equiv.), Na₂CO₃ (1 mmol, 2 equiv.), Cu(OAc)₂·H₂O (1 mmol, 2 equiv.), 120 °C, DMF (3 mL), 24 h. ^b Isolated yields are given. Yields obtained with 100 °C heating condition are given in parenthesis. ^c 16 h. ^d Decarbonylated product phenol formed in minor amount.

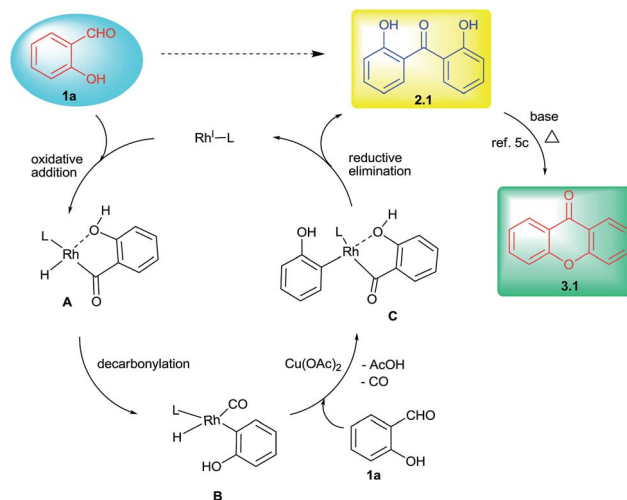


To probe the mechanism, we performed a series of control experiments as given in Scheme 2. Firstly, the reactivity of benzaldehyde and 2-methoxybenzaldehyde was examined under the established conditions and in these cases the corresponding decarbonylative homo-coupled products were not obtained (eqn (1) and (2)). This observation in conjunction with our results with salicylaldehydes indicated that the hydroxy group is crucial for the desired product formation. Next, the reactivity of cross-decarbonylative homo-couplings was studied using two different salicylaldehydes. This led to the formation of a mixture containing both homo- and cross-coupled products (eqn (3)) indicating the involvement of two aldehydic C–H activations at two different stages during the catalytic cycle. Another cross-over experiment performed with salicylaldehyde and *p*-anisaldehyde gave only homo-coupled product (2.1) from salicylaldehyde in 76% yield (eqn (4)). In this reaction, we have not observed the formation of any cross-coupled product signifying the non-participation of *p*-anisaldehyde during the reaction course. Thus the presence of –OH group in salicylaldehyde has a definitive role during aldehydic C–H activation.

Further, formation of xanthone was examined with the direct use of 2,2'-dihydroxybenzophenone in the presence of base under heating conditions. This reaction gave xanthone product (3.1) in 83% yield (eqn (5)) and it clearly indicated the important role of base in the cyclization of 2,2'-dihydroxybenzophenone during the reaction course.^{5c}



Scheme 2 Control experiments.



Scheme 3 Proposed mechanistic cycle.

Based on these investigations, the following mechanistic cycle was proposed (Scheme 3). The initial oxidative addition of salicylaldehyde with Rh(I) was expected to form the intermediate A. This would undergo decarbonylation¹⁵ to generate arylrhodium B. This in turn involves in the C–H activation of second molecule of salicylaldehyde followed by copper acetate mediated transformation provides arylrhodium C species.¹⁹ This intermediate would then undergo reductive elimination to give 2,2'-dihydroxybenzophenone (2.1) as decarbonylative homo-coupled product.¹⁸ The minor formation of phenol in some cases as decarbonylated product also supports the involvement of arylrhodium B. The proposed second involvement of aldehydic C–H with arylrhodium B also explains the cross-over experiment using two different salicylaldehydes giving a mixture of products (Scheme 2, eqn (3)). The prominent role of –OH indicated by the control reaction carried out by using salicylaldehyde and aryl aldehyde (Scheme 2, eqn (4)) in which case no cross-coupled product (either 5 or 6) was obtained. Further base mediated cyclization of 2,2'-dihydroxybenzophenone directly provided xanthone product under heating conditions (Scheme 2, eqn (5)).^{5c}

Conclusions

In conclusion, we have developed an efficient methodology for the decarbonylative homo-coupling of functionalized salicylaldehydes to give 2,2'-dihydroxybenzophenones. This study was further extended to the preparation of xanthone derivatives directly from functionalized salicylaldehydes in a one-pot operation. Another highlight is that, this study showed the important and crucial role of hydroxy group in the reductive homo-coupling of salicylaldehydes. The overall outcome of the reaction course was also influenced by the electronics of the substituent groups in salicylaldehydes. This study thus unravelled the direct formation of 2,2'-dihydroxybenzophenones and xanthenes from functionalized salicylaldehydes under rhodium-catalyzed conditions.



Experimental

General

All decarbonylative couplings were performed in dry Schlenk tubes under nitrogen atmosphere conditions. Salicylaldehyde and 3-methoxysalicylaldehyde were purchased from Spectrochem. 5-Bromosalicylaldehyde and 4-*N,N*-diethylamino-salicylaldehyde were purchased from Sigma-Aldrich. Literature procedures were followed for the preparation of other functionalized salicylaldehydes.^{20–22} All solvents were dried according to standard procedures. Purification of the products was carried out by silica-gel column chromatography using ethyl acetate/hexane as eluent. ¹H and ¹³C NMR spectra were recorded with JEOL ECS 400 (400 MHz) and JEOL ECX 500 (500 MHz). HRMS was recorded using Electron Ionization (EI) and Electro-Spray Ionization (ESI) techniques with Waters CAB155 GCT Premier analyzer and Waters HAB 213 Q-TOF Premier analyzer. IR spectra were recorded on PerkinElmer FT-IR. X-ray data was recorded on Bruker SMART APEX-II CCD diffractometer. Melting points reported are uncorrected.

Representative coupling procedure for Table 2

This reaction was carried out by charging a dry Schlenk tube with salicylaldehyde (61 mg, 0.5 mmol, 1 equiv.), Rh(CO)₂(acac) (6.4 mg, 0.025 mmol, 0.05 equiv.), Na₂CO₃ (106 mg, 1 mmol, 2 equiv.), Cu(OAc)₂·H₂O (200 mg, 1 mmol, 2 equiv.) and DMF (3 mL). The resultant mixture was stirred at 100 °C in an oil bath for 24 h. After that, contents were cooled to rt, quenched with dil. HCl and extracted with ethyl acetate (30 mL). The organic portion was washed with water (15 mL), brine (15 mL), dried over anhydrous MgSO₄ and concentrated. The crude product was subjected to silica-gel column chromatography using ethyl acetate/hexane as eluent. The desired product **2.1** was obtained as yellow liquid (43 mg, 80%).

Representative coupling procedure for Table 3

The above procedure given for Table 2 was followed with heating at 120 °C. The desired product **3.1** was obtained as white solid (23 mg, 47%).

2.1.^{23a} Yellow liquid (43 mg, 80%); *R*_f (3% EtOAc/hexane) 0.55. ¹H NMR (400 MHz, CDCl₃) δ 10.60 (s, 2H, –OH), 7.62 (dd, *J* = 8.0, 1.7 Hz, 2H, Ar–H), 7.54–7.49 (m, 2H, Ar–H), 7.09 (dd, *J* = 8.4, 1.1 Hz, 2H, Ar–H), 6.94 (ddd, *J* = 8.0, 7.2, 1.1 Hz, 2H, Ar–H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 202.54, 161.88, 136.08, 133.22, 119.97, 119.02, 118.74 ppm. IR (neat, cm^{–1}): 3235, 1624, 1590, 1484, 1256, 1234, 1154, 939, 756, 648. HRMS (EI⁺): calcd for C₁₃H₁₀O₃ [M]⁺ 214.0630; found 214.0630.

2.2.^{23b} White solid (52 mg, 76%); mp 128–130 °C, *R*_f (10% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 11.34 (s, 2H, –OH), 7.55 (d, *J* = 8.9 Hz, 2H, Ar–H), 6.53 (d, *J* = 2.5 Hz, 2H, Ar–H), 6.47 (dd, *J* = 8.9, 2.6 Hz, 2H, Ar–H), 3.87 (s, 6H, –OMe) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 199.24, 165.70, 164.81, 134.55, 113.57, 107.29, 101.68, 55.76 ppm. IR (KBr, cm^{–1}): 2977, 1616, 1590, 1358, 1243, 1202, 1119, 963, 845, 593. HRMS (EI⁺): calcd for C₁₅H₁₄O₅ [M]⁺ 274.0841; found 274.0840.

2.3. Yellow liquid (47 mg, 68%); *R*_f (10% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 10.01 (s, 2H, –OH), 7.12 (d, *J* = 8.0 Hz, 2H, Ar–H), 7.07 (d, *J* = 8.0 Hz, 2H, Ar–H), 6.87 (t, *J* = 8.0 Hz, 2H, Ar–H), 3.94 (s, 6H, –OMe) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 202.13, 150.22, 148.52, 123.91, 121.43, 118.69, 116.03, 56.40 ppm. IR (neat, cm^{–1}): 3381, 2942, 2842, 1628, 1586, 1456, 1439, 1340, 1254, 996, 833, 756, 744, 733. HRMS (EI⁺): calcd for C₁₅H₁₄O₅ [M]⁺ 274.0841; found 274.0841.

2.4. Yellow solid (65 mg, 73%); mp 112–114 °C, *R*_f (5% EtOAc/hexane) 0.56. ¹H NMR (500 MHz, CDCl₃) δ 11.91 (s, 2H, –OH), 7.51 (d, *J* = 8.8 Hz, 2H, Ar–H), 6.20–6.18 (m, 4H, Ar–H), 3.41 (q, *J* = 7.1 Hz, 8H, NCH₂CH₃), 1.21 (t, *J* = 7.1 Hz, 12H, NCH₂CH₃) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 196.15, 164.67, 152.92, 134.67, 109.31, 103.12, 98.02, 44.71, 12.83 ppm. IR (KBr, cm^{–1}): 2973, 2931, 1614, 1556, 1526, 1349, 1295, 1263, 1137, 1110, 1077, 829, 788, 702. HRMS (ESI⁺): calcd for C₂₁H₂₉N₂O₃ [M + H]⁺ 357.2178; found 357.2173.

2.5. Yellow solid (46 mg, 76%); mp 100–102 °C, *R*_f (3% EtOAc/hexane) 0.36. ¹H NMR (400 MHz, CDCl₃) δ 10.38 (s, 2H, –OH), 7.39 (d, *J* = 1.5 Hz, 2H, Ar–H), 7.33 (dd, *J* = 8.4, 2.2 Hz, 2H, Ar–H), 6.99 (d, *J* = 8.4 Hz, 2H, Ar–H), 2.31 (s, 6H, –CH₃) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 202.49, 159.70, 136.89, 132.86, 128.13, 119.87, 118.45, 20.69 ppm. IR (KBr, cm^{–1}): 3207, 3038, 2924, 1622, 1582, 1481, 1387, 1342, 1288, 1236, 1197, 968, 822, 797, 734. HRMS (EI⁺): calcd for C₁₅H₁₄O₃ [M]⁺ 242.0943; found 242.0947.

2.6. Yellow liquid (64 mg, 65%); *R*_f (50% EtOAc/hexane) 0.65. ¹H NMR (400 MHz, CDCl₃) δ 10.27 (s, 2H, –OH), 6.85 (s, 2H, Ar–H), 4.05 (s, 6H, –OMe), 3.96 (s, 6H, –OMe), 3.77 (s, 6H, –OMe) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 199.64, 151.02, 148.83, 145.28, 141.74, 114.98, 109.64, 61.42, 61.31, 56.73 ppm. IR (KBr, cm^{–1}): 3403, 2940, 2837, 1608, 1490, 1464, 1368, 1275, 1210, 1159, 1076. HRMS (EI⁺): calcd for C₁₉H₂₂O₉ [M]⁺ 394.1264; found 394.1269.

2.7. White solid (32 mg, 51%); mp 140–142 °C, *R*_f (3% EtOAc/hexane) 0.25. ¹H NMR (400 MHz, CDCl₃) δ 10.18 (s, 2H, –OH), 7.33–7.27 (m, 4H, Ar–H), 7.08 (ddd, *J* = 8.8, 4.5, 0.7 Hz, 2H, Ar–H) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 200.37, 158.00, 154.94 (d, *J*_{C–F} = 238.7 Hz), 123.94 (d, *J*_{C–F} = 23.4 Hz), 120.36 (d, *J*_{C–F} = 7.3 Hz), 119.32 (d, *J*_{C–F} = 6.2 Hz), 117.71 (d, *J*_{C–F} = 24.2 Hz) ppm. IR (KBr, cm^{–1}): 3334, 3079, 1624, 1613, 1508, 1479, 1429, 1246, 1194, 1176, 1134, 838, 694. HRMS (EI⁺): calcd for C₁₃H₈F₂O₃ [M]⁺ 250.0442; found 250.0446.

2.8. Yellow solid (41 mg, 46%); mp 154–156 °C, *R*_f (5% EtOAc/hexane) 0.16. ¹H NMR (400 MHz, CDCl₃) δ 9.94 (s, 2H, –OH), 7.63 (d, *J* = 2.5 Hz, 2H, Ar–H), 7.37 (d, *J* = 2.5 Hz, 2H, Ar–H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 198.81, 154.61, 135.52, 129.79, 124.70, 124.20, 121.88 ppm. IR (KBr, cm^{–1}): 3486, 3081, 2918, 2850, 1622, 1462, 1428, 1325, 1256, 1219, 1169, 871, 799, 736, 718. HRMS (EI⁺): calcd for C₁₃H₆Cl₄O₃ [M]⁺ 349.9071; found 349.9078.

2.9. Yellow solid (31 mg, 33%); mp 120–122 °C, *R*_f (5% EtOAc/hexane) 0.4. ¹H NMR (400 MHz, CDCl₃) δ 10.26 (s, 2H, –OH), 7.69 (s, 2H, Ar–H), 7.62 (d, *J* = 8.9 Hz, 2H, Ar–H), 7.01 (d, *J* = 8.9 Hz, 2H, Ar–H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 200.07, 160.76, 139.16, 134.59, 120.97, 120.92, 111.15 ppm. IR (KBr,



cm⁻¹): 3412, 2923, 1736, 1624, 1608, 1585, 1466, 1337, 1226, 1182, 826, 516. HRMS (ESI⁻): calcd for C₁₃H₇Br₂O₃ [M - H]⁻ 368.8762; found 368.8767.

2.10.^{23c} White solid (22 mg, 30%); mp 146–148 °C, R_f (3% EtOAc/hexane) 0.36. ¹H NMR (400 MHz, CDCl₃) δ 10.29 (s, 2H, -OH), 7.56 (d, *J* = 2.6 Hz, 2H, Ar-H), 7.50 (dd, *J* = 8.9, 2.6 Hz, 2H, Ar-H), 7.07 (d, *J* = 8.9 Hz, 2H, Ar-H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 200.27, 160.31, 136.38, 131.59, 124.32, 120.56, 120.32 ppm. IR (KBr, cm⁻¹): 3263, 2926, 1625, 1611, 1585, 1465, 1339, 1224, 1185, 960, 830, 718. HRMS (ESI⁻): calcd for C₁₃H₇Cl₂O₃ [M - H]⁻ 280.9772; found 280.9776.

2.11. White solid (32 mg, 43%); mp 178–180 °C, R_f (20% EtOAc/hexane) 0.18. ¹H NMR (400 MHz, CDCl₃) δ 10.94 (s, 2H, -OH), 8.30 (t, *J* = 2.4 Hz, 2H, Ar-H), 8.17 (dt, *J* = 8.8, 2.2 Hz, 2H, Ar-H), 7.19–7.16 (m, 2H, Ar-H), 2.56 (s, 6H, -COCH₃) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 201.63, 195.58, 165.59, 136.22, 134.08, 129.03, 119.28, 118.94, 26.40 ppm. IR (KBr, cm⁻¹): 3141, 2925, 1678, 1627, 1588, 1482, 1360, 1272, 1214, 836. HRMS (ESI⁺): calcd for C₁₇H₁₈NO₅ [M + NH₄]⁺ 316.1185; found 316.1184.

3.1.^{23d} White solid (23 mg, 47%); mp 162–164 °C, R_f (3% EtOAc/hexane) 0.5. ¹H NMR (400 MHz, CDCl₃) δ 8.35 (dd, *J* = 8.4, 1.3 Hz, 2H, -OH), 7.75–7.71 (m, 2H, Ar-H), 7.51–7.48 (m, 2H, Ar-H), 7.41–7.36 (m, 2H, Ar-H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 177.42, 156.34, 134.98, 126.89, 124.07, 122.00, 118.13 ppm. IR (KBr, cm⁻¹): 2922, 1656, 1608, 1481, 1459, 1346, 1332, 752. HRMS (ESI⁺): calcd for C₁₃H₉O₂ [M + H]⁺ 197.0603; found 197.0604.

3.2.^{24a} White solid (22 mg, 39%); mp 128–130 °C, R_f (3% EtOAc/hexane) 0.18. ¹H NMR (400 MHz, CDCl₃) δ 8.11 (d, *J* = 1.9 Hz, 2H, Ar-H), 7.51 (dd, *J* = 8.3, 2.1 Hz, 2H, Ar-H), 7.37 (d, *J* = 8.5 Hz, 2H, Ar-H), 2.46 (s, 6H, -Me) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 177.56, 154.60, 136.09, 133.63, 126.18, 121.60, 117.87, 20.99 ppm. IR (KBr, cm⁻¹): 3466, 3057, 1660, 1619, 1481, 1308, 1218, 1142, 824, 809, 794, 539. HRMS (EI⁺): calcd for C₁₅H₁₂O₂ [M]⁺ 224.0837; found 224.0835.

3.3. White solid (33 mg, 35%); mp 152–154 °C, R_f (40% EtOAc/hexane) 0.32. ¹H NMR (400 MHz, CDCl₃) δ 7.49 (s, 2H, Ar-H), 4.12 (s, 6H, -OMe), 4.07 (s, 6H, -OMe), 3.98 (s, 6H, -OMe) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 175.68, 150.27, 148.20, 146.06, 141.82, 116.99, 100.94, 62.02, 61.66, 56.47 ppm. IR (KBr, cm⁻¹): 2940, 2852, 1649, 1597, 1470, 1424, 1383, 1208, 1127, 1104, 1058, 860, 766. HRMS (EI⁺): calcd for C₁₉H₂₀O₈ [M]⁺ 376.1158; found 376.1151.

3.4.^{24a} White solid (26 mg, 45%); mp 158–160 °C, R_f (5% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 7.95 (dd, *J* = 8.3, 2.9 Hz, 2H, Ar-H), 7.54–7.45 (m, 4H, Ar-H) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 176.04, 158.97 (d, *J*_{C-F} = 244.4 Hz), 152.51, 123.46 (d, *J*_{C-F} = 25.5 Hz), 122.01 (d, *J*_{C-F} = 7.1 Hz), 120.21 (d, *J*_{C-F} = 8.0 Hz), 111.52 (d, *J*_{C-F} = 23.4 Hz) ppm. IR (KBr, cm⁻¹): 3055, 1671, 1626, 1475, 1291, 1166, 1134, 819, 776, 544. HRMS (EI⁺): calcd for C₁₃H₆F₂O₂ [M]⁺ 232.0336; found 232.0330.

3.5.^{24a} White solid (41 mg, 46%); mp 192–194 °C, R_f (5% EtOAc/hexane) 0.2. ¹H NMR (400 MHz, CDCl₃) δ 8.43–8.42 (m, 2H, Ar-H), 7.81 (ddd, *J* = 8.9, 2.5, 0.6 Hz, 2H, Ar-H), 7.39 (dd, *J* = 8.9, 0.4 Hz, 2H, Ar-H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 174.94, 154.97, 138.21, 129.48, 122.95, 120.21, 117.65 ppm. IR (KBr, cm⁻¹): 1661, 1608, 1458, 1289, 1127, 813. HRMS (EI⁺): calcd for C₁₃H₆Br₂O₂ [M]⁺ 351.8735; found 351.8735.

3.6.^{24a} White solid (33 mg, 50%); mp 210–212 °C, R_f (3% EtOAc/hexane) 0.5. ¹H NMR (400 MHz, CDCl₃) δ 8.29 (d, *J* = 2.6 Hz, 2H, Ar-H), 7.69 (dd, *J* = 8.8, 2.7 Hz, 2H, Ar-H), 7.47 (d, *J* = 9.0 Hz, 2H, Ar-H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 175.24, 154.56, 135.45, 130.32, 126.28, 122.50, 119.96 ppm. IR (KBr, cm⁻¹): 3084, 1669, 1610, 1462, 1289, 825, 816. HRMS (EI⁺): calcd for C₁₃H₆Cl₂O₂ [M]⁺ 263.9745; found 263.9742.

3.7.^{24b} White solid (38 mg, 54%); mp 202–204 °C, R_f (20% EtOAc/hexane) 0.2. ¹H NMR (400 MHz, CDCl₃) δ 8.91 (d, *J* = 2.2 Hz, 2H, Ar-H), 8.41 (dd, *J* = 8.8, 2.3 Hz, 2H, Ar-H), 7.62 (d, *J* = 8.8 Hz, 2H, Ar-H), 2.74 (s, 6H, -COCH₃) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 196.38, 176.42, 158.78, 134.58, 133.68, 128.41, 121.30, 119.14, 26.85 ppm. IR (KBr, cm⁻¹): 2925, 2854, 1679, 1664, 1602, 1480, 1361, 1256, 1245, 1120, 830, 587. HRMS (EI⁺): calcd for C₁₇H₁₂O₄ [M]⁺ 280.0736; found 280.0733.

3.8. White solid (34 mg, 54%); mp 228–230 °C, R_f (20% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 10.15 (s, 2H, -CHO), 8.86 (d, *J* = 1.9 Hz, 2H, Ar-H), 8.34 (dd, *J* = 8.7, 2.0 Hz, 2H, Ar-H), 7.69 (d, *J* = 8.7 Hz, 2H, Ar-H) ppm. ¹³C NMR (125 MHz, CDCl₃) δ 190.16, 175.79, 159.48, 133.90, 133.22, 131.64, 122.02, 119.83 ppm. IR (KBr, cm⁻¹): 3438, 3075, 2851, 1699, 1673, 1608, 1478, 1452, 1239, 1168, 1115, 832. HRMS (EI⁺): calcd for C₁₅H₈O₄ [M]⁺ 252.0423; found 252.0425.

3.9.^{25a} Yellow solid (41 mg, 57%); mp 230–232 °C, R_f (20% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 9.11–9.07 (m, 2H, Ar-H), 8.56–8.52 (m, 2H, Ar-H), 7.68–7.64 (m, 2H, Ar-H) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 174.29, 158.60, 144.48, 129.94, 123.31, 121.18, 120.06 ppm. IR (KBr, cm⁻¹): 2917, 1669, 1610, 1533, 1464, 1345, 1284, 1083, 834, 747, 665. HRMS (EI⁺): calcd for C₁₃H₆N₂O₆ [M]⁺ 286.0226; found 286.0220.

3.10.^{24b} White solid (28 mg, 49%); mp 316–318 °C, R_f (30% EtOAc/hexane) 0.25. ¹H NMR (400 MHz, DMSO-*d*₆) δ 11.28 (s, 2H, -OH), 7.18 (d, *J* = 9.2 Hz, 2H, Ar-H), 6.33–6.30 (m, 4H, Ar-H) ppm. ¹³C NMR (125 MHz, DMSO-*d*₆) δ 199.34, 162.87, 160.72, 133.58, 115.30, 107.41, 102.48 ppm. IR (KBr, cm⁻¹): 3401, 2255, 2128, 1651, 1049, 1026, 1004, 826, 764, 631. HRMS (ESI⁻): calcd for C₁₃H₇O₄ [M - H]⁻ 227.0344; found 227.0341.

4.^{25b} White solid (33 mg, 27%); mp 74–76 °C, R_f (3% EtOAc/hexane) 0.3. ¹H NMR (400 MHz, CDCl₃) δ 11.63 (s, 1H, -OH), 10.29 (s, 1H, -OH), 7.60–7.55 (m, 2H, Ar-H), 7.50–7.46 (m, 1H, Ar-H), 7.07 (dd, *J* = 8.4, 1.1 Hz, 1H, Ar-H), 6.95–6.91 (m, 1H, Ar-H), 6.54 (d, *J* = 2.5 Hz, 1H, Ar-H), 6.48 (dd, *J* = 9.0, 2.5 Hz, 1H, Ar-H), 3.88 (s, 3H, -OMe) ppm. ¹³C NMR (100 MHz, CDCl₃) δ 200.63, 166.35, 165.80, 160.98, 135.25, 135.18, 132.56, 120.45, 118.93, 118.56, 113.36, 107.68, 101.61, 55.83 ppm. IR (KBr, cm⁻¹): 2936, 1615, 1585, 1482, 1347, 1269, 1237, 1151, 924, 759. HRMS (ESI⁺): calcd for C₁₄H₁₃O₄ [M + H]⁺ 245.0814; found 245.0811.

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