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# Crystalline pyrazine-2-amidoxime isolated by diffusion method and its structural and behavioral analysis in the context of crystal engineering and microbiological activity†

 Agnieszka Chylewska,<sup>\*a</sup> Małgorzata Ogryzek,<sup>a</sup> Angelika Głębocka,<sup>a</sup> Artur Sikorski,<sup>a</sup> Katarzyna Turecka,<sup>b</sup> Ewa. D. Raczynska<sup>c</sup> and Mariusz Makowski<sup>\*a</sup>

Pyrazine-2-amidoxime (PAOX) is a structural analogue of a popular drug, *i.e.* pyrazine-2-carboxamide (PZA). The crystalline PAOX was obtained by diffusion as a method of crystallization. Various types of intermolecular interactions between the H-bond donors (CH, NH, and OH) and H-bond acceptors (hydroxyl O, imino N, aza N) were found between PAOX molecules in X-ray diffraction studies. It was observed in the crystal structure that PAOX forms not only dimers but also stable helical-like polymers, stabilized by intermolecular interactions between two neighbouring molecules. Their geometric, energetic and spectroscopic properties were also characterised by DFT methods. Thermal decomposition of PAOX was examined with the use of a TG-IR analysis (20–1000 °C), and the results were further resolved. Electrochemical behaviour of the compound studied in acetonitrile in the absence or presence of methanol was described in detail, and mechanisms of the anodic oxidation and cathodic reduction were proposed. The complexometric properties of PAOX were examined against selected d-block metal ions in acetonitrile, as well as in aqueous solution. Biological assay of PAOX was performed to determine the antimicrobial activity and potential pharmaceutical applications. The minimum inhibitory (MIC) and minimal bactericidal (or fungicidal) concentrations (MBC/MFC) for PAOX were determined against six microorganisms.

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## Introduction

The discovery of the biological activity of pyrazine derivatives<sup>1–3</sup> led to the systematic search for potential drugs within the discussed class of compounds. The pyrazine ring system has proven to be an interesting class in heterocyclic chemistry but it has received little attention in the literature so far. Some of its derivatives are important as anticancer agents<sup>4</sup> with low toxicity and blood platelet aggregation inhibitors.<sup>5</sup> They also show anti-

inflammatory, antifungal, antiparasitic and antioxidant activities.<sup>6–11</sup> In addition, pyrazine derivatives are used as bone metabolism improvers, as adenosine antagonists, and as controlling herbicides.<sup>12</sup>

Amidoximes retain a complementary position in drug design and discovery, being individually considered as pharmacophores and playing a role as parent compounds in various transformations.<sup>13</sup> Moreover, amidoximes can serve as starting materials for the synthesis of multiple heterocyclic compounds and could be useful as building blocks for a number of heteroarenes.<sup>14,15</sup> It has also been reported that many drugs containing the amidoxime group are biologically active and display antihypertensive, antibacterial, antitrypanocidal and cytostatic properties.<sup>16,17</sup> They were found as NO generators *in vivo*, which had neuromodulatory and neurotransmitter effects.<sup>18</sup> The cyclic amidoximes were also reported to possess anamnestic, antihypoxic and hypotensive activities.<sup>19,20</sup> Homo- and heterometallic oximate complexes have played a considerable role in coordination chemistry and crystal engineering, because of their fascinating topological geometries and interesting properties.<sup>21</sup> One of the most appropriate examples may be pyrazine-2-amidoxime (PAOX), which structure corresponds to the structure of pyrazine-2-carboxamide (PZA) (Fig. 1).

<sup>a</sup>Faculty of Chemistry, University of Gdansk, W. Stwosza 63, 80-308 Gdansk, Poland.  
E-mail: [agnieszka.chylewska@ug.edu.pl](mailto:agnieszka.chylewska@ug.edu.pl); [mariusz.makowski@ug.edu.pl](mailto:mariusz.makowski@ug.edu.pl); Tel: +48 58 523 50 52; +48 58 523 50 55

<sup>b</sup>Faculty of Pharmacy with Subfaculty of Laboratory Medicine, Al. Hallera 107, 80-416 Gdansk, Poland

<sup>c</sup>Department of Chemistry, Warsaw University of Life Sciences, Nowoursynowska 159c, 02-776 Warszawa, Poland

† Electronic supplementary information (ESI) available: File includes both experimental, and theoretical (calculated) IR, Raman <sup>1</sup>H NMR spectra, experimental protocols details (X-ray, microbiological tests) and titration spectra, as well as atom Cartesian coordinates, and formation free energies of three dimers D1–D3, and a tetramer T1. CCDC 1474463 full crystallographic details, excluding structure features, have been deposited for pyrazine-2-amidoxime. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6ra10537h



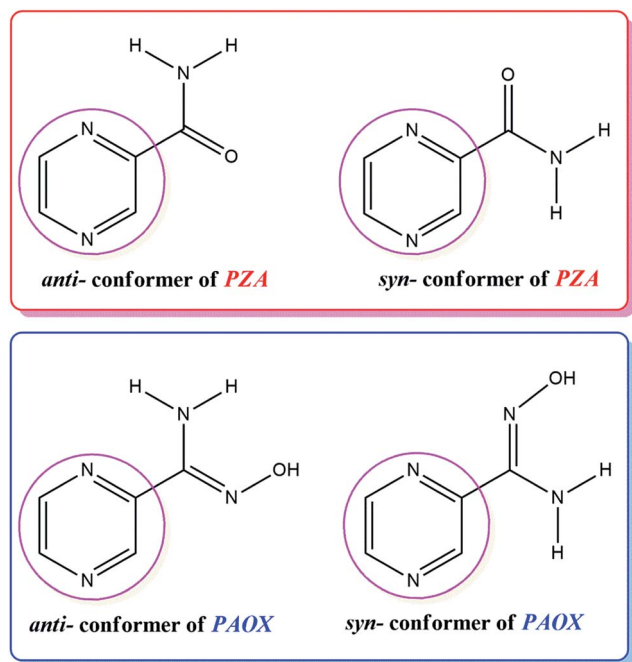


Fig. 1 Molecule structure of pyrazine-2-amidoxime (PAOX) as analogue of known pyrazine-2-carboxamide (PZA).

The choice of **PAOX** for our studies was based on the following reasons. The ability of this molecule to coordinate metal ions through one or both N and O donor atoms of the amidoxime group as well as one or both N donor atoms of the pyrazine ring.<sup>22</sup> In this case, **PAOX** can act as a chelating terminal or bridging ligand (Fig. 2), and shows the interesting mode of self-activation by metal ions, which could be used in synthetic inorganic chemistry.<sup>23</sup> It is worth pointing out here that the structure of pyrazine-2-amidoxime as a ligand should directly influence the structure of complexes in the presence of

such an organic ligand in the coordination sphere; the formation of clusters or two-dimensional molecular brick-wall complexes with Ni(II) or Fe(II) ions has already been examined.<sup>24,25</sup> The same relation can be possible in the case of its magnetic or other properties.<sup>24,25</sup> Because solubility, polymorphism, crystallinity and crystal habits are solvent-dependent, thus solvent screening is fundamental and essential, especially in the pharmaceutical industry.<sup>24,25</sup> That is why the study in solution is a very important assay and allows the determination of optimal conditions of crystallization and complexometric titrations.

The designing of chelators with high selectivity for metal ions are very important because of their potential analytical applications; they are significant for the removal, separation and increasing concentration of metallic species. They can be potentially applied in the recycling of resources in metallurgy or in the waste water treatment. Transition metal ions such as Cd(II), Pb(II) or Hg(II) are recognized as highly toxic and their presence in the environmental waters or soils is undesirable.<sup>24,25</sup> Cu(II), Ni(II) and Co(II) are bio-essential and there are numerous metalloenzymes containing these elements. However, even essential metal ions can be toxic when present in excess.<sup>26,27</sup> Therefore, separation of these trace metals is necessary due to their potential health and ecological hazard. The complexation and absorption process is one of the ways which can be used for the removal of metals from waste water. For this purpose, **PAOX** has been examined and its complexometric properties investigated in aqueous and acetonitrile solutions.

Some oxime derivatives were indicated as antioxidants with scavenging activity towards different forms of reactive species<sup>28–30</sup> (e.g. hydrogen peroxide, hydroxyl or nitric oxide radicals), which was also related to their metal chelating properties.<sup>31</sup> The investigation of the redox behaviour of **PAOX**, using electrochemical technique, leads to better understanding of its biological redox reactions or generally – for this class of compounds. The mechanism of pyridinium oxime reduction was investigated and published about twenty years ago.<sup>32–34</sup> Although this area for pyridinium oximes is well understood in organic electrochemistry, no experimental data are available on the redox properties of **PAOX**.

The structural properties of pyrazine-2-amidoxime were investigated here in the solid state and in solution as well as in the gas phase by performing density functional theory (DFT) calculations. Our study of inter- and intramolecular interactions between side chain of **PAOX** individual units resulted in the new knowledge about the three-dimensional helix-like formation. Due to the above, **PAOX** can play the role of biomimetic foldamer, which is generally an artificial model for molecular folding and could express other functions (catalyst, receptor). The crystal forms of each foldamer need to be resolved at first, in order to provide the information for the following step of studies about its interaction properties or functional abilities. Here, we investigate co-crystallization structural studies of other pyrazine derivatives.<sup>35,36</sup> Although, we focused here on untypical properties, such as aromatic amidoxime as foldamer, its conformation, thermal, redox and complexometric properties.

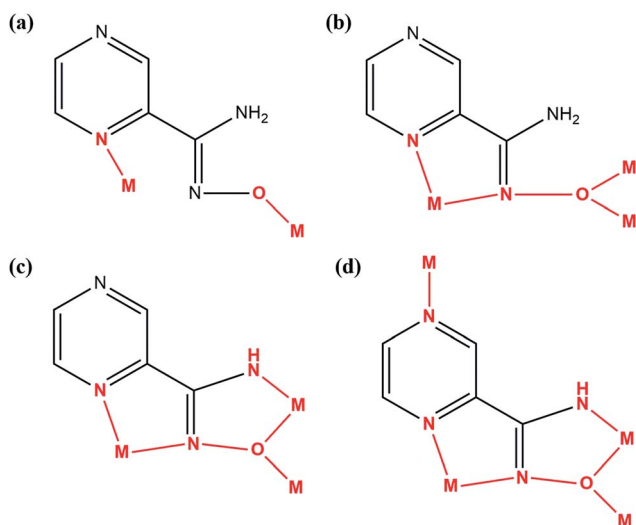


Fig. 2 Coordination modes of metal – **PAOX** in different types of complexes: (a) and (b) dinuclear; (c) trinuclear; (d) tetranuclear.<sup>25</sup>



## Results and discussion

### Crystallographic studies

Single-crystal X-ray diffraction measurements showed that pyrazine-2-amidoxime crystallizes in the monoclinic  $P2_1$  space group with two **PAOX** molecules (denote as A and B) in the asymmetric unit (Fig. 3, Table S1†).

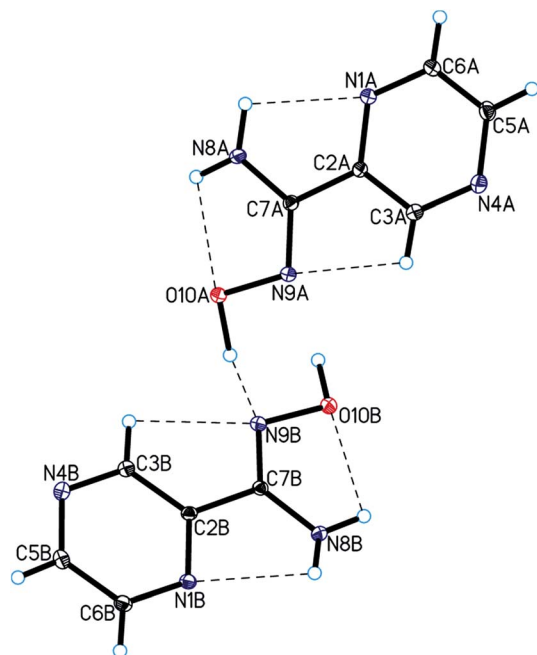


Fig. 3 Molecular structure of the title compound showing the atom-labelling scheme. Displacement ellipsoids are drawn at the 25% probability level and H atoms are shown as small spheres of arbitrary radius (intramolecular hydrogen bonds are represented by dashed lines).

The bond lengths, valence and torsion angles (Tables S2–S4†) characterizing the geometry of the pyrazine ring and amidoxime moieties are consistent with the geometric parameters observed for these groups of compounds.<sup>37,38</sup> In the **PAOX** crystals, pyrazine rings are nearly planar with an average deviation from planarity of 0.001(2) and 0.012(2)° for molecules A and B, respectively, and form a dihedral angle of 10.2(2) and 3.6(2)° with the amidoxime group in molecules A and B, respectively. In the crystal structure of pyrazine-2-amidoxime, intramolecular N–H⋯O, N–H⋯N, and C–H⋯N and intermolecular O–H⋯N, N–H⋯N, and C–H⋯O hydrogen bonds stabilize the structure (Fig. 3, Table 1). An analysis of the crystal packing of the title compound showed that adjacent molecules of **PAOX** are linked by O10–H10⋯N9 hydrogen bonds and interact by  $\pi_{(\text{pyrazine})} - \pi_{(\text{pyrazine})}$  interactions (Fig. 4, Table 1).

As a result, we can observe a helical-like arrangement of the molecules of **PAOX** in the crystals. The neighbouring helical-like structures are linked by N8–H8⋯N1 hydrogen bonds (dimer D2) to produce blocks along the *a*-axis (Fig. 5, Table 1). The adjacent blocks are linked by N8–H8⋯N4 and C5–H5⋯O10 hydrogen bonds to form a three-dimensional framework structure (Fig. 6, Table 1).

Because of the existence of several possible interaction sites in the structure of **PAOX**<sup>36</sup> by itself and crystallographic results included in this work formation energies of five possible dimers (D1 in Fig. 3 and D2–D5 in Fig. 5, respectively) and one tetramer (T1) to model helical-like turn (Fig. 4) were considered. The T1 form was chosen as a smallest fragment of the helical-like structure to in order check its stability in comparison to dimers. The frequencies were calculated, first to prove that the structures correspond to the energy minima, and then to estimate the corresponding zero-point energies and to assess their theoretical IR, Raman, and UV spectra for all the structures discussed in this paper. These calculations allowed to assess the

Table 1 Geometry of hydrogen bonding interactions in **PAOX** crystal structure. Data obtained from single-crystal X-ray diffraction measurements. The atom labels and numbers in agreement with those in Fig. 4–6<sup>a</sup>

D–H⋯A	<i>d</i> (D–H) (Å)	<i>d</i> (H⋯A) (Å)	<i>d</i> (D⋯A) (Å)	<D–H⋯A (°)
<b>Intramolecular</b>				
N8A–H8A⋯O10A	0.92(3)	2.18(3)	2.535(2)	102(2)
N8B–H8B⋯O10B	0.89(3)	2.31(3)	2.599(2)	99(2)
N8A–H8C⋯N1A	0.96(3)	2.33(3)	2.717(2)	103(2)
N8B–H8D⋯N1B	0.87(3)	2.35(2)	2.730(3)	106(2)
C3A–H3A⋯N9A	0.93	2.52	2.817(3)	99
C3B–H3B⋯N9B	0.93	2.47	2.781(3)	100
<b>Intermolecular</b>				
O10A–H10A⋯N9B	0.92(3)	1.82(3)	2.734(2)	170(3)
O10B–H10B⋯N9A <sup>i</sup>	0.99(3)	1.86(3)	2.849(2)	174(2)
N8A–H8A⋯N4B <sup>ii</sup>	0.92(3)	2.42(3)	3.316(3)	165(3)
N8B–H8B⋯N4A <sup>iii</sup>	0.89(3)	2.45(3)	3.266(3)	153(2)
N8A–H8C⋯N1B <sup>iv</sup>	0.96(3)	2.22(3)	3.054(3)	144(2)
N8B–H8D⋯N1A <sup>v</sup>	0.87(3)	2.48(3)	3.275(3)	151(2)
C5A–H5A⋯O10B <sup>vi</sup>	0.93	2.51	3.273(3)	139
C5B–H5B⋯O10A <sup>vii</sup>	0.93	2.50	3.374(3)	156

<sup>a</sup> Symmetry code: (i)  $-1 + x, y, z$ ; (ii)  $-x, -1/2 + y, -z$ ; (iii)  $1 - x, 1/2 + y, 1 - z$ ; (iv)  $x, -1 + y, z$ ; (v)  $x, 1 + y, z$ ; (vi)  $1 - x, -1/2 + y, 1 - z$ ; (vii)  $1 - x, 1/2 + y, -z$ .



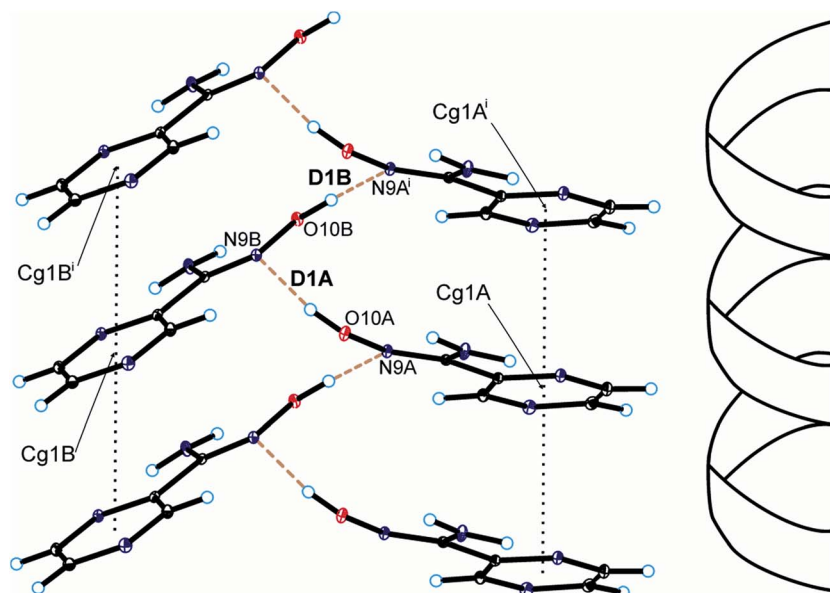


Fig. 4 Helical arrangement of molecules of pyrazine-2-amidoxime in the crystal packing of the title compound (hydrogen bonds are represented by dashed lines and  $\pi$ - $\pi$  interactions by dotted lines; symmetry codes are the same as in Table S2†).

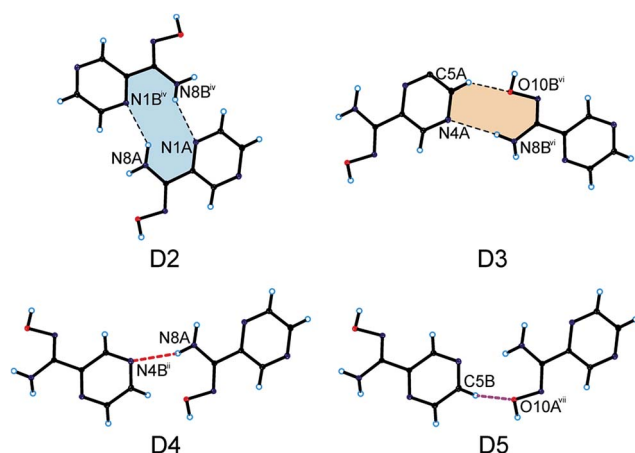


Fig. 5 Dimers formed by PAOX molecules in the crystal packing of the title compound (hydrogen bonds are represented by dashed lines; symmetry codes are the same as in Table S2†).

most stable and probable organizations of **PAOX** molecules in its crystal structure. According to the results of calculations performed in this work dimers D2, D4, and D5 are alike. For this reason, only one structure D2 was considered in the following discussion. Free energies and atom coordinates of optimized structures of polymers (D1, D2, D3, and T1), obtained by using the DFT method (B3LYP/6-311+G\*\*) *in vacuo* were collected in Table S5 in the ESI.† The counterpoise procedure was considered in the energy calculations. The formation energies of polymers were collected in Table 2.

The crystallographic data (coordinates) obtained in this work were used in the calculations as starting points and allowed to assess the formation energies of the systems studied. Based on the results in Table 2 it was found that the most stable, and

probable interaction system in the crystal structures of **PAOX** is D1 with the complexation energy  $-11.31 \text{ kcal mol}^{-1}$ . This result is in agreement with the crystal analysis. As shown in Table 1, the shortest hydrogen bonds were found in the case when two **PAOX** molecules were planar, interacted through amidoxime group (Fig. 3), and were stabilized by two hydrogen bonds  $\text{N9A} \cdots \text{O10B}$ , and  $\text{O10A} \cdots \text{N9B}$ , with the distances experimentally found at about 2.85 and 2.73 Å, respectively. The second energetically preferred system found in the crystal structure of **PAOX** was its tetramer, modelling the helical-like fragment, with the formation energy  $-7.63 \text{ kcal mol}^{-1}$ .

The experimental results showed that four **PAOX** molecules in the tetramer (T1) could form a helical-like turn. All possible hydrogen bond distances were longer for T1 than those in the case of D1 but shorter than those for D2, respectively. In terms of complexation energy D3 was slightly more stable than D2. For D2 and D3, the formation energies were about  $-4.00 \text{ kcal mol}^{-1}$ . The knowledge of the **PAOX**'s crystal structure may have a cognitive importance for supramolecular chemistry and can be implemented from the crystal engineering's point of view in the designing of complexes containing this compound.

### Spectroscopic characterization

The infrared and Raman spectroscopies together with quantum-chemical calculations performed in parallel are important techniques of analysis used for confirming the structure of compounds. The vibrational band assignments of **PAOX** have been made in accordance with their position, magnitude, nature and relative intensities. Also the assignments have been made in analogy with the structurally related molecules. On this basis, characteristic bands of aromatic and aliphatic groups of the C-H-N-O scaffold were identified. The results of the analyses of the theoretical and experimental FT-IR





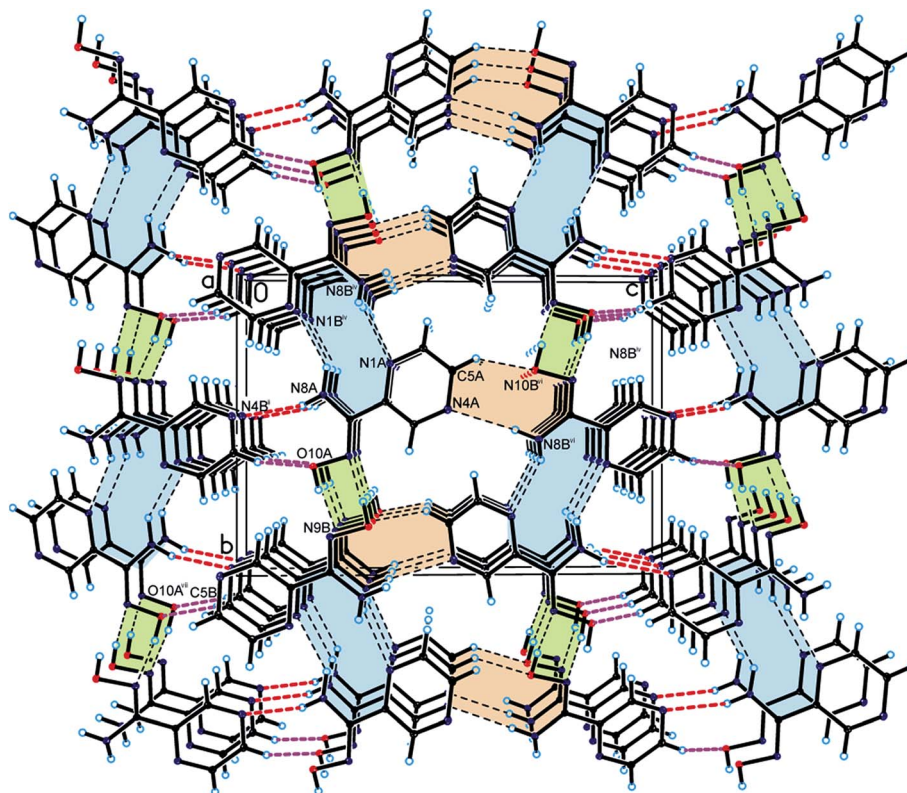


Fig. 6 Crystal packing of the title compound viewed along the *a*-axis (helical arrangement is highlighted in green; hydrogen bonds are represented by dashed lines).

**Table 2** The free energy values of complexation of systems studied in the crystal structure of PAOX. The energy values are in [kcal mol<sup>-1</sup>] after inclusion of the BSSE. Calculations were obtained at DFT(B3LYP) level with the use of 6-311+G\*\* basis set *in vacuo*

System <sup>a</sup>	kcal mol <sup>-1</sup>
D1	-11.31
D2	-4.00
D3	-4.28
T1	-7.63

<sup>a</sup> D1 – denotes a dimer in Fig. 4, D2-D3 – denote dimers in Fig. 6, and T1 denotes a tetramer in Fig. 5. W tekście jest kcal mol<sup>-1</sup>.

and Raman spectra for monomer, the most energetically stable dimer D1, and solid pyrazine-2-amidoxime are collected in Table S6 and shown in Fig. S1 and S2, respectively, in the ESI.† The vibrational frequencies obtained by DFT calculation are also shown in Table S6† for comparison.

The comparison of infrared and Raman spectra of PAOX and structural data provide specific vibrations. In the experimental FT-IR range 3437–3331 cm<sup>-1</sup>, the peaks generated *via* symmetric and asymmetric stretching modes of NH<sub>2</sub> group are observed as strong and sharp for solid PAOX. They correspond to multiple interactions possible in the solid state. The calculated FT-IR frequencies for monomeric PAOX and the dimer D1 are equal to 3587, 3414 (monomer), and 3531, 3410 cm<sup>-1</sup> (dimer) respectively. Similarly, the same signals can be found in

the Raman spectra with their values being of 3537 and 3425 cm<sup>-1</sup> for monomeric PAOX, and 3532 and 3410 cm<sup>-1</sup> for D1. The associated O–H stretching vibration generated in all cases studied, *i.e.* experimental broad band at 3144, and calculated for monomer and D1 band at 3657 and 3294 cm<sup>-1</sup>, respectively, in their infrared spectra, and 3060, 3657, 3294 cm<sup>-1</sup> in their Raman spectra. The C–H stretching vibrations of hetero aromatic structure occur in the region 3100–3000 cm<sup>-1</sup> for asymmetric and 2990–2900 cm<sup>-1</sup> for symmetric stretching modes of vibration.<sup>39,40</sup> In many cases, they are too weak for detection. Hence in the present study, the experimentally observed C–H symmetric stretching peaks were found at 2816 cm<sup>-1</sup>, and those calculated at 3016, and 3017 cm<sup>-1</sup> for isolated D1.

The band observed at 3059 cm<sup>-1</sup> (calculated 3037–2999, and 3017 cm<sup>-1</sup>, respectively) is associated with the asymmetric stretching vibration C–H in the FT-Raman spectrum. The C–H out-of-plane bending modes occur in the spectral range<sup>41</sup> 900–667 cm<sup>-1</sup>. The bands observed in the region 1153–1018 cm<sup>-1</sup> are assigned to <sup>TM</sup>(CH) in-plane vibrations in experimental FT-IR spectrum. The <sup>TM</sup>(CH) out-of-plane modes generated sharp bands 862 and 844 cm<sup>-1</sup>. The corresponding peaks appeared at 1188–1056 cm<sup>-1</sup> and 845 cm<sup>-1</sup> in the FT-Raman spectrum. Based on the DFT calculated vibrational spectra, the C–H stretching vibrations are assigned to the calculated values 1128–1043 (1130 for D1), 825–805 (823 for D1), and 1134–1020 (1130 for D1) cm<sup>-1</sup> in the case of FT-IR and FT-Raman spectra, respectively.



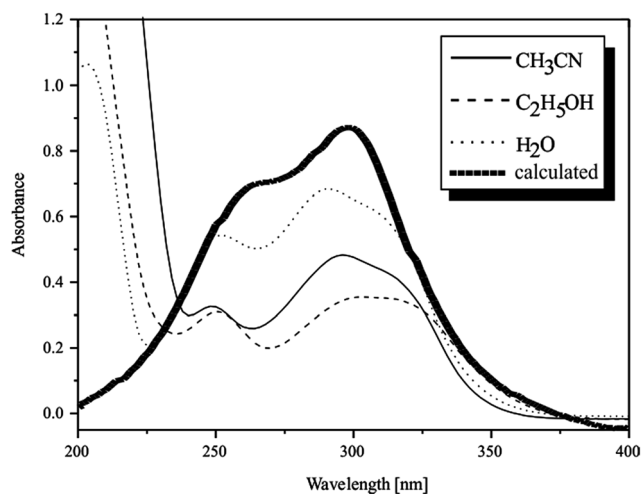


Fig. 7 UV spectra of pyrazine-2-amidoxime obtained in three solvents: acetonitrile (MeCN), ethanol (EtOH), water (H<sub>2</sub>O), and from calculation (B3LYP, 6-311+G\*\*).

In the present investigations, a very strong, sharp band was found at 1659 cm<sup>-1</sup> in the experimental FT-IR along with 1662 cm<sup>-1</sup> in the experimental FT-Raman. These two signals are associated with C=N stretching vibration of oxime group. In our case the experimentally observed values of peak due to pyrazine ring C=N stretching vibrations were found at 1590 cm<sup>-1</sup> (FT-IR) and 1576 cm<sup>-1</sup> (Raman). The theoretically calculated values of C=N stretching vibration of oxime group were found at 1619 (monomer), 1627 (D1) in both types of spectra.

The absorption peaks in the range 1470–1525 cm<sup>-1</sup> in the heterocyclic compounds are ascribed to carbon vibrations (C–C).<sup>42</sup> In the present investigation, the FT-IR bands appeared at 1520, 1483 cm<sup>-1</sup> (calculated for monomer at 1510 and 1472, and D1 at 1426 cm<sup>-1</sup>) and FT-Raman bands were observed at 1521

and 1491 cm<sup>-1</sup> [calculated 1493 (isolated), 1495 (D1), and 1432 (isolated) cm<sup>-1</sup>] were caused by asymmetric and symmetric carbon–carbon ring stretching mode, respectively.

The experimental IR spectrum of **PAOX** presents characteristic band which confirms the presence of oxime group in the structure of examined pyrazine derivative.<sup>43</sup> The sharp, strong peak found at 953 cm<sup>-1</sup> in the experimental FT-IR spectrum can be attributed to the stretching mode<sup>44</sup> of  $\nu(\text{N-O})_{\text{oxime}}$ . Based on the DFT calculated vibrational spectra, this band appeared in the region of 914 (monomer) and 940 cm<sup>-1</sup> (D1). This indicates the presence of oxime group, in accordance with the crystallographic data.

The electronic spectrum of the **PAOX** shows in water two absorption bands at 252 nm and 290 nm which are ascribed to  $\pi \rightarrow \pi^*$  and  $n \rightarrow \pi^*$  transitions, respectively (Fig. 7). Dissolution of **PAOX** in a different solvent leads to a change in the position of the maxima of absorption. Changing the solvent from water to acetonitrile (at constant **PAOX** concentration,  $6 \times 10^{-5}$  M) causes blue shift of  $\pi \rightarrow \pi^*$  absorption band to shorter wavelengths – hypsochromic effect, with significant change in its intensity ( $\Delta A$  about  $-0.24$  – hypsochromic effect). Moreover, the value of molar absorption coefficient ( $\epsilon$ ) is nearly double reduced when the solvent was changed from water ( $\epsilon_{\text{PAOX}} = 9500 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ) to acetonitrile ( $\epsilon_{\text{PAOX}} = 5500 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ). The  $n \rightarrow \pi^*$  absorption band of **PAOX** aqueous solution (290 nm) moves in MeCN to longer wavelengths – red shift and changes intensity from 0.68 in water to 0.48 in acetonitrile. The value of  $\epsilon$  decreased from  $11333 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$  (H<sub>2</sub>O) to  $8000 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$  (MeCN).

Maintaining the constant concentration of **PAOX** ( $6 \times 10^{-5}$  M) and changing the solvent from water to ethanol changes the values of absorption and wavelength of the examined compound. In the case of maximum assigned to  $\pi \rightarrow \pi^*$  transition, the slightly hypsochromic effect of absorption band with a significant decrease in intensity ( $\Delta A$  about  $-0.26$ ) was observed. The value of **PAOX** molar absorption coefficient in

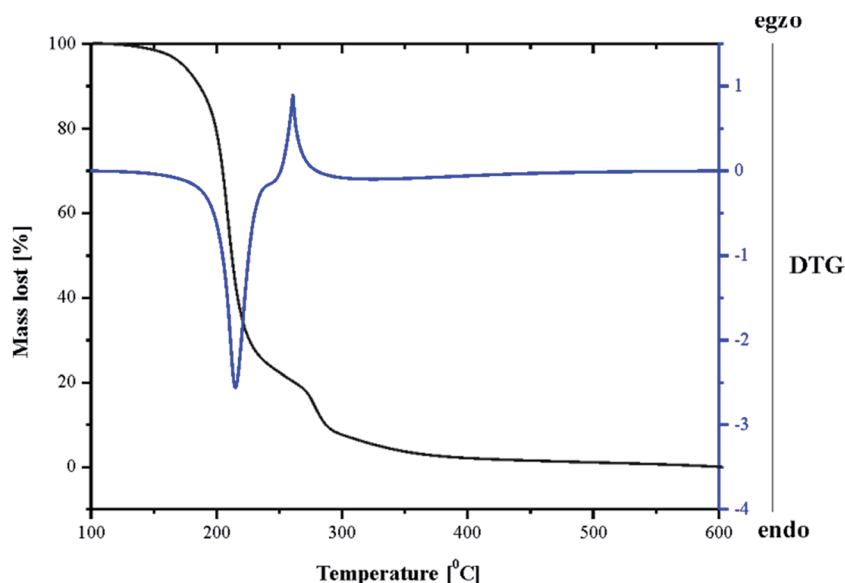


Fig. 8 TG and DTG curves obtained during thermal decomposition of pyrazine-2-amidoxime.



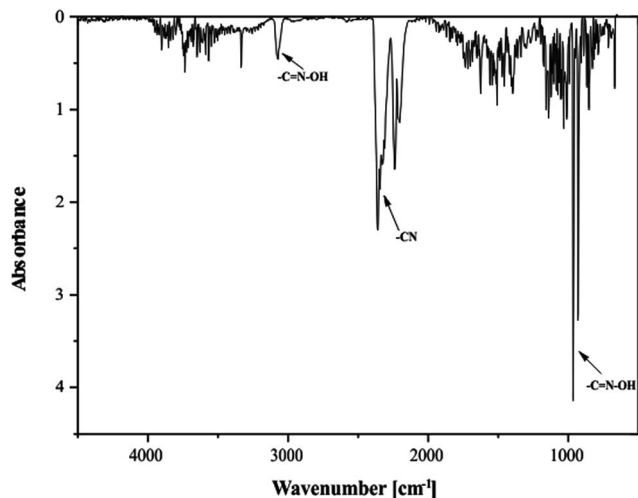


Fig. 9 IR spectrum for decomposition of pyrazine-2-amidoxime, temperature 148 °C, after 35 min.

ethanolic solution is about 60% lower than the value of  $\epsilon_{\text{PAOX}}$  in aqueous solvent ( $\epsilon_{\text{PAOX}} = 9500 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$  in water;  $\epsilon_{\text{PAOX}} = 5167 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$  in ethanol). The second maximum ascribed to  $n \rightarrow \pi^*$  transition in the spectrum of ethanolic solution of **PAOX** appeared at 308 nm. The bathochromic shift to longer wavelength with a significant decrease in intensity ( $\Delta\lambda$  about  $-0.33$ ) could be observed when the solvent was changed from water to ethanol (hypochromic effect). Furthermore,  $\epsilon$  is nearly twice lower in ethanol ( $\epsilon_{\text{PAOX}} = 5833 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ) as compared to water ( $\epsilon_{\text{PAOX}} = 11\,333 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ).

### Thermogravimetry investigations

Thermal analyses of solid pyrazine-2-amidoxime were carried out by the TG and DTG techniques. The experimental results revealed that the degradation occurred in two stages (see Fig. 8). The study of thermal decomposition of **PAOX** showed an increased heat endurance of this compound, which did not contain crystal water and generally decomposed at temperatures exceeding 170–400 °C.

The first step between 153 and 270 °C (the black line in Fig. 8) was accompanied by an 82.04% mass loss, attributed to

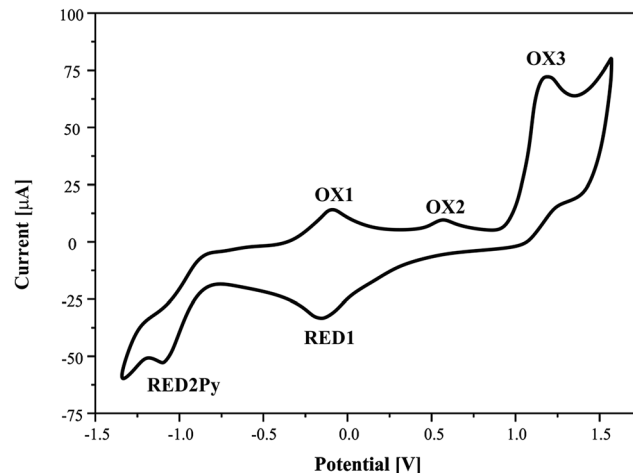


Fig. 10 Cyclic voltammograms of pyrazine-2-amidoxime (1 mM) in acetonitrile solution, 0.1 M TBP. Scan rate =  $100 \text{ mV s}^{-1}$ .

the loss of a pyrazine molecule, which was identified as IR spectra of products (Fig. 9). The second step (271–316 °C) was accompanied by a 13.22% mass loss, which might be due to the decomposition of amidoxime group. The residual mass of 4.74% above 320 °C was probably caused by simultaneous formation of nitrogen species: nitrogen ( $\text{N}_2$ ), nitrous oxide ( $\text{N}_2\text{O}$ ) and ammonia ( $\text{NH}_3$ ), continued to the moment of physical capacity of the TG curve registration (range of heating). The DTA curve (the blue line in Fig. 9) for **PAOX** shows two peaks. The first one is endo- and the second is exothermic at 153–270 °C and 271–316 °C, respectively.

### Electrochemical studies

The redox behaviour of the pyrazine-2-amidoxime was studied in acetonitrile solution at a platinum working electrode. The relevant electrochemical data are summarized in Table 3 and cyclic voltammogram (CV) is shown in Fig. 10.

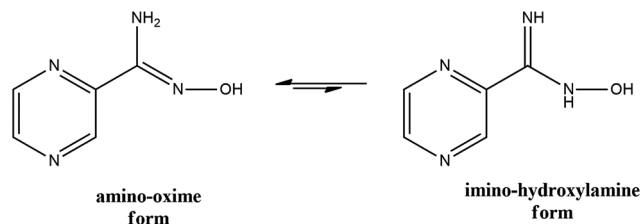
It can be seen that CV in acetonitrile solvent showed the occurrence of three consecutive anodic reactions: the first, reversible reaction of amine group oxidation **OX1** and **RED1**, the second, main and well-defined peak at **OX3**, and a broad and ill-defined peak **OX2**. Due to electrochemical data, we suggest that

Table 3 Electrochemical data of pyrazine-2-amidoxime in acetonitrile (TBAP 0.1 M, scan rate  $100 \text{ mV s}^{-1}$ ,  $20 \pm 0.1$  °C)

Symbol of the reaction	Electrode nature	Potential value $E_{\text{pa}}$ or $E_{\text{pc}}$ [V]	CV reversible nature			CV irreversible nature
			$\Delta E_p$ [V]	$E_{1/2}$	$I_a/I_c$	$I_a$ [ $\mu\text{A}$ ]
<b>OX1</b>	Anodic	−0.058	0.067	−0.092	0.603	—
<b>RED1</b>	cathodic	−0.125	—	—	—	—
<b>OX2</b>	Anodic	0.619	—	—	—	8.35
<b>OX3</b>	Anodic	1.231	—	—	—	76.59
<b>RED2Py</b>	cathodic	−1.060	0.072	−1.096	0.688	−55.28
<b>OX2Py<sup>a</sup></b>	Anodic	−1.132	—	—	—	—

<sup>a</sup> The presented representative potential was observed after an addition the first portion of methanol to **PAOX** acetonitrile solution (1 mM); reference electrode – SCE;  $E_{1/2} = 0.5(E_{\text{pa}} + E_{\text{pc}})$  where  $E_{\text{pa}}$  and  $E_{\text{pc}}$  are anodic and cathodic peak potential, respectively;  $\Delta E_p$  as mathematical subtract both electrode potential values.





Scheme 1 The crucial tautomers of pyrazine-2-amidoxime.

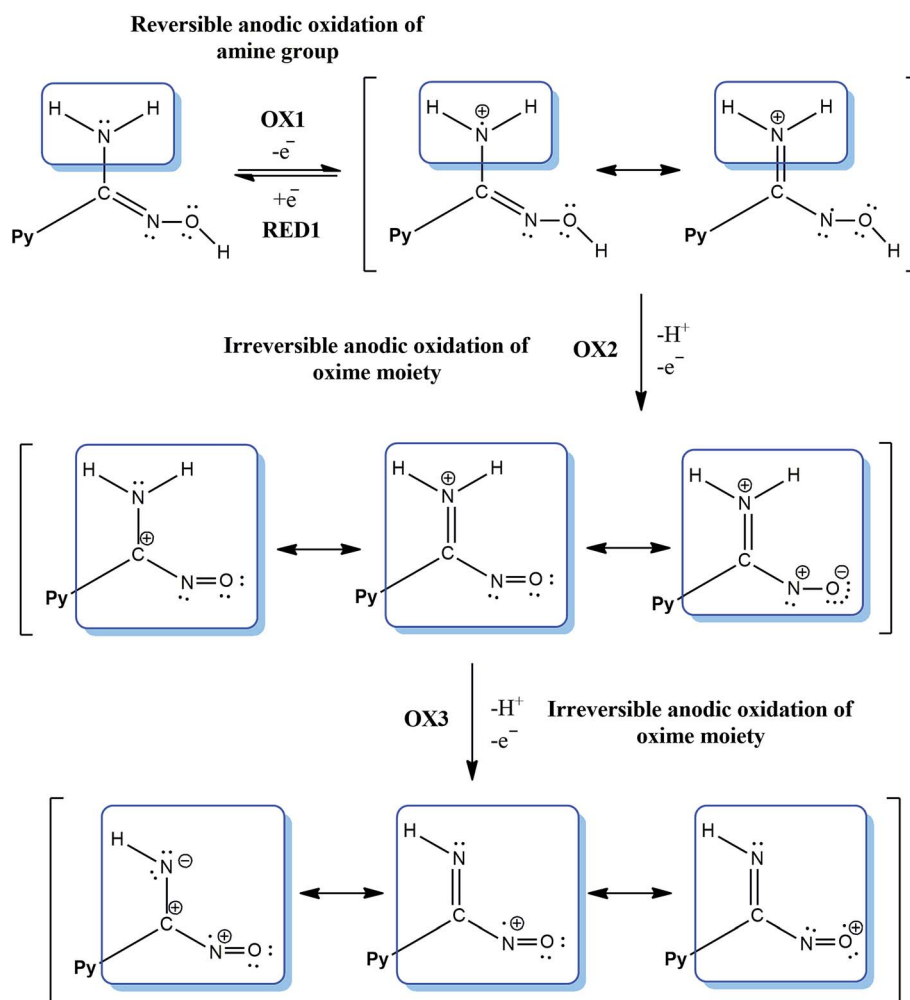
oxidation of reaction (positive potentials), involving a proton and a single-electron abstractions, forms the anion oxime group, leading to the formation of iminoxy radical (see Schemes 1 and 2). It is known that a large number of biologically active compounds exhibit reduction potentials greater than  $-0.5$  V, like **RED2Py** for irreversible reaction of **PAOX**, which can permit electron acceptance from *in vivo* donors.<sup>45</sup>

The electrochemical process in negative potential region is due to the reduction of pyrazine ring – Py (potential marked as **RED1Py** in Fig. 10) in **PAOX**. Pyrazine decomposition occurs simultaneously with hydrogen evolution, and a reduction peak **RED1Py**

can be observed at  $-1.060$  V (vs. SCE). After electrochemical reduction of Py in **PAOX**, the potential sweep was switched to the positive direction. No oxidation peak was observed for this reaction, due to the above it can be concluded that the reduction of pyrazine ring as part of **PAOX** is an irreversible process.

The follow-up chemical reaction is directly related to the nucleophilic attack of the molecule presented in the solution under study, e.g. methanol (Scheme 3). The addition of methanol to the acetonitrile solutions of the investigated **PAOX** resulted in substantial change in CVs (Fig. 11). With an increase in methanol concentration (from  $0.001$  M to  $0.5$  M), the potentials of reduction (**RED2Py**) and oxidation (**OX3**) shifted towards more positive values. **RED2Py** potential values increased initially and then decreased, but in the case of **OX3** potential values – at the beginning they decreased and subsequently increased (all obtained changes were related with successive, additional portions of methanol).

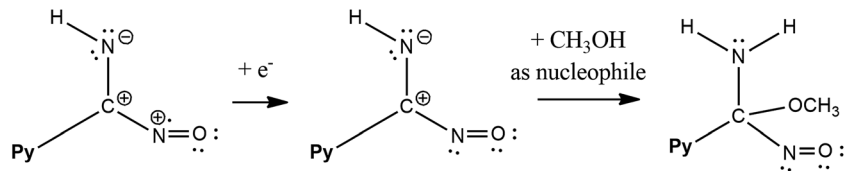
The methanol is a nucleophile, therefore the shift of the potential **OX3** and disappearance of the potential **OX2** with the increase in the methanol concentration indicate that the chemical reaction preceding the second electrode stage is accompanied by the proton abstraction as an irreversible



Scheme 2 The electrochemical equations for amidoxime moiety oxidation.







Scheme 3 The proposed mechanism of nucleophilic attack.

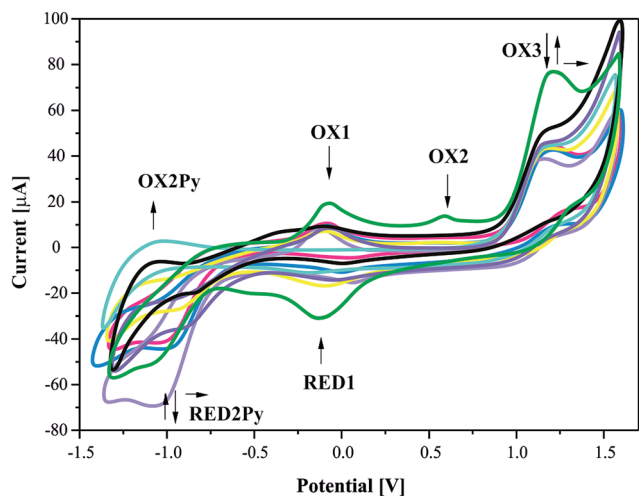


Fig. 11 Cyclic voltammograms of 1 mM PAOX acetonitrile solution (0.1 M TBP as electrolyte) on a Pt electrode in the presence of methanol, M (green line) 0, (pink line) 0.001, (violet line) 0.005, (yellow line) 0.01, (light-violet line) 0.05, (blue line) 0.1, (black line) 0.2, (navy line) 0.5. Scan rate = 20 mV s<sup>-1</sup>.

reaction (see Scheme 2). Moreover, a new very weak oxidation peak **OX2Py** was observed (Fig. 11), which described the re-oxidation reaction of **DHPAOX**<sup>+</sup> during the addition of methanol (see Scheme 4). All the proposed mechanisms are modifications taking into account the data obtained from our own research, they are based, however, on our best scientific knowledge.<sup>46–49</sup>

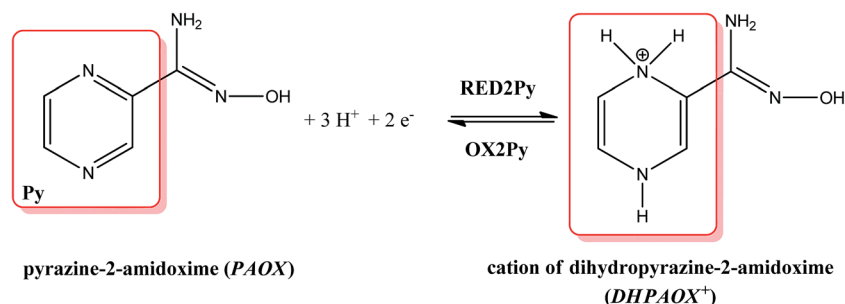
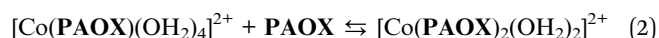
### Complexometric properties

The presence of chromophore groups in a molecule of a ligand, whose absorption spectra change upon complexation, allows for the use of spectroscopy methods to study the chelation

process. On the basis of the results from spectroscopy measurements it is possible to show a quantity of equilibria existing in a particular solution. Furthermore, this investigation can provide complete information about species formed during the titration process.

Changes in the position and intensity of the absorption peak after addition of metal ion to pyrazine-2-amidoxime solution show which ions form the most stable complex in the selected solution. Consequently, the influence of the metal ions Co(II), Mn(II), Fe(II) and Cr(III) on the electronic spectrum of **PAOX** in acetonitrile and in water was examined. Significant changes in the position and intensity of bands were observed for all analyzed ions mainly in acetonitrile (see ESI, Fig. S3–S6 present the result for MeCN, and Fig. S7–S9† for H<sub>2</sub>O, respectively).

Based on the analysis of the spectrophotometric titration curves, the number of equilibria and stoichiometry of complexes formed with cobalt(II) ion were determined for both solvent systems (Fig. 12a and 13a). The intensities of absorption bands gradually increase and slight bathochromic shift can be observed at MeCN. At the same time two isosbestic points appeared at 285 nm and 328 nm, respectively. These effects suggest the existence of metal ion interactions with donor atom of **PAOX** in the system studied. The stoichiometric diagram (the absorbance at 300 nm as a function of Co<sup>2+</sup>–ligand molar ratio) was plotted to illustrate specific quantity stoichiometry in the system studied (Fig. 12b – MeCN and Fig. 13b – water). The proposed complexing equilibria models (eqn (1) and (2)) correspond to the formation of two metal ion–ligand complexes in aqueous solution.



Scheme 4 The proposed reversible electrochemical equation for pyrazine ring of **PAOX** (observed in negative potentials).

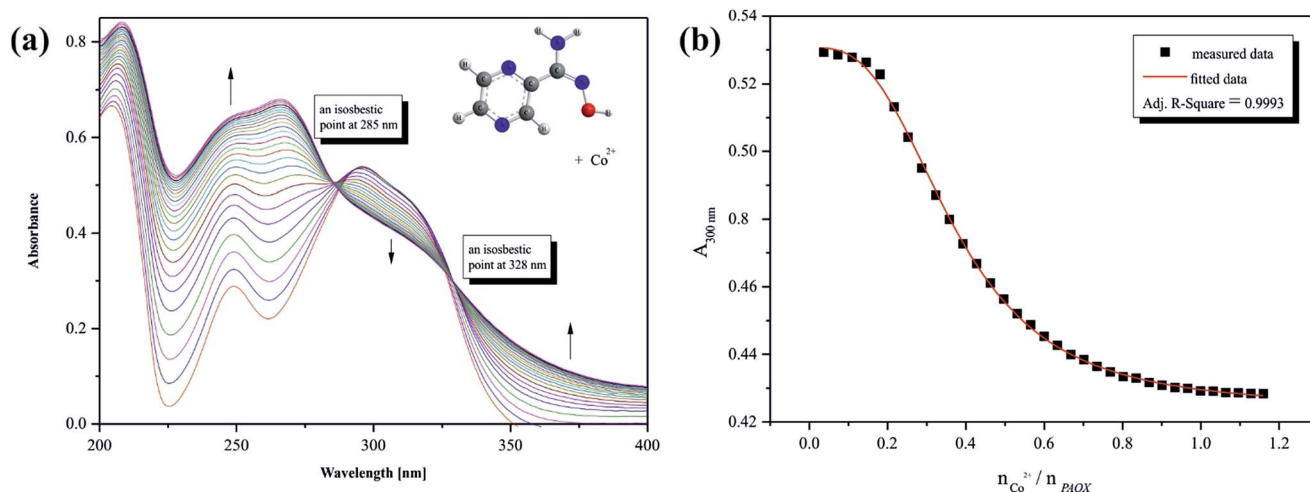
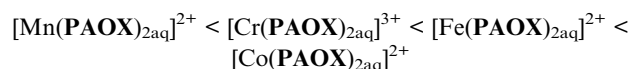


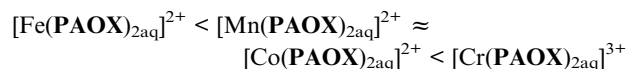
Fig. 12 (a) Representative titration spectra for PAOX in acetonitrile with Co(II): (a) full titration curves, (b) absorption as a function of metal–ligand molar ratio at 300 nm. For titration conditions, see ‘Analytical instrumentations’. Arrows indicate the change in absorbance during the titration.

On the basis of the results from spectrophotometric measurements, the values of gradual and cumulative formation constants for selected metal ion complexes were determined in acetonitrile and water systems (Table 4). It must be pointed out that acetonitrile is a better solvent for the observation of the complexometric process between d-block metal ions and PAOX. It corresponds to PAOX better solubility in non-aqueous solvents due to the absence of weak intermolecular interactions with the solvent.

Complexometric properties of pyrazine-2-amidoxime were also reported in order to obtain the full information about the behaviour of the compound studied as a bio-ligand towards d-block metal ions. The complexes formed in acetonitrile solution can be put to order according to their increasing stability, which is consistent with the Irving–Williams theoretical predictions:



The order of the value stability constants of complexes under study can be established also for aqueous solution:



### Microbiological assay

The activity of pyrazine-2-amidoxime was examined for the reference strains, provided by the Institute of Immunology and Experimental Therapy of the Polish Academy of Sciences, Wrocław, Poland. The selected reference strains include: Gram-positive bacteria: *Enterococcus hirae* ATCC 8043 and *Staphylococcus aureus* ATCC 6538; Gram-negative bacteria such as *Escherichia coli* ATCC 8739, *Proteus vulgaris* ATCC 6896,

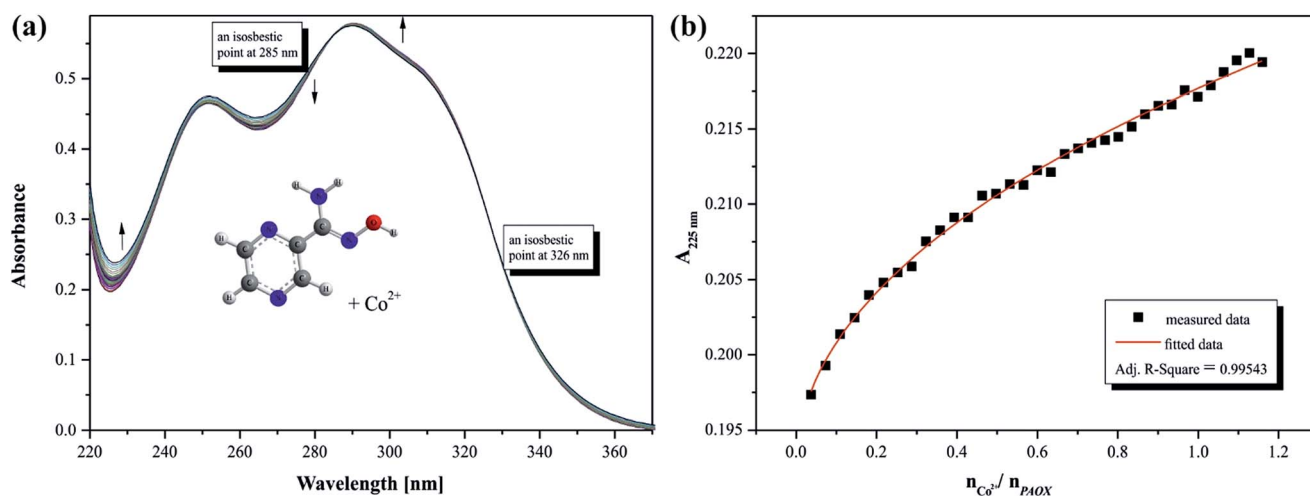


Fig. 13 (a) Representative titration spectra for PAOX in water with Co(II): (a) full titration curves, (b) absorption as a function of metal–ligand molar ratio at 225 nm. For titration conditions, see section ‘Analytical instrumentations’. Arrows indicate the change in absorbance during the titration.



**Table 4** Values of complex stability constants for **PAOX** with selected metal ions in acetonitrile (MeCN) and in water (H<sub>2</sub>O) determined by spectrophotometric method. All results constants were determined at 22 ± 0.1 °C

Metal ion	Molar ratio	Spectrophotometric technique					
		MeCN			H <sub>2</sub> O		
		log K <sub>1</sub>	log K <sub>2</sub>	log β	log K <sub>1</sub>	log K <sub>2</sub>	log β
Cr <sup>3+</sup>	1 : 2	3.51 (±0.02)	4.91 (±0.01)	8.42 (±0.01)	5.75 (±0.18)	7.67 (±0.35)	13.42 (±0.26)
Co <sup>2+</sup>	1 : 2	3.65 (±0.03)	6.74 (±0.03)	10.39 (±0.03)	3.42 (±0.09)	4.74 (±0.09)	8.16 (±0.13)
Fe <sup>2+</sup>	1 : 2 (MeCN)/1 : 1 (H <sub>2</sub> O)	3.98 (±0.04)	4.99 (±0.04)	8.97 (±0.04)	4.79 (±0.12)	—	4.79 (±0.12)
Mn <sup>2+</sup>	1 : 2	4.10 (±0.03)	3.97 (±0.05)	8.07 (±0.04)	3.79 (±0.14)	4.35 (±0.07)	8.14 (±0.11)



**Fig. 14** Crystals of pyrazine-2-amidoxime obtained by diffusion crystallization (microscopic image, 40× zoom).

*Pseudomonas aeruginosa* ATCC 9027 and yeast *Candida albicans* ATCC 4635. The values of the Minimal Bactericidal (fungicidal) Concentration (MBC) and Minimal Inhibitory Concentration (MIC) were determined. In the microbiological studies quantitative methods, such as microdilution in liquid (MIC) and solid (MBC) culture media, were applied respectively.

The results of biological assays show that the investigated compound exhibited bactericidal activity against both Gram negative and Gram positive bacteria strains. Pyrazine-2-amidoxime showed killing and inhibiting properties at a concentration of 5.79 mM. Moreover, studied pyrazine derivative had about 10-fold higher inhibitory activity against *C. albicans* than other investigated strains. The **PAOX** inhibited growth of selected yeast at the concentration of 0.58 mM. The present studies will indicate the possibility of using **PAOX** as a new antimicrobial agent. Since *C. albicans* is the major fungal pathogen in humans (carried by over 50% of the population), in this context further investigations in this field are reasonable.

## Conclusions

On the basis of our research, **PAOX** was found to be stable in the experiment conditions at a room temperature. Its decomposition at higher temperatures (>150 °C) was observed. The structural studies in the solid state, solution,<sup>36</sup> and gas phase<sup>36</sup> showed the same tautomeric preferences, *anti*-conformation on

the C2–C7 single bond, and *Z*-configuration on the amidoxine C=N double bond.<sup>36</sup>

There was no possibility of observing the self-association of **PAOX** in solution with the use of spectroscopic methods in our earlier studies<sup>36</sup> because of the low solubility of **PAOX** in nonpolar solvents. Finally, the crystallographic methods showed what interaction types occurred and what groups were involved in it. These interactions were also found in the experimental IR and Raman spectra, recorded for solid **PAOX**, and confirmed computationally at the DFT level for selected polymers of **PAOX**. Two *anti*-conformers were observed in the crystal structure of **PAOX** (hereinafter referred to as A and B molecules). They differ between each other only by an angle in the pyrazine ring and  $-(\text{NH}_2)\text{C}=(\text{NH}-\text{OH})$  group (molecule A showed larger value of this angle than B); both with nearly planar pyrazine ring. The controlled use of the H-bonds formed between **PAOX**-A molecules chain and **PAOX**-B molecules chain, and  $\pi$ - $\pi$  stacking between the same types of **PAOX** molecules in two chains ( $\pi_{\text{A}}-\pi_{\text{A}}$  or  $\pi_{\text{B}}-\pi_{\text{B}}$ ) as a component in foldamers is challenging. The **PAOX** ability to create channel structures can be used better as a co-crystal host or antimicrobial agent.

The electrochemical properties of pyrazine-2-amidoxime were discussed, as well as the mechanisms of cathodic reduction of pyrazine part and anodic oxidations of amidoxime moieties (in the absence or presence of methanol) were proposed in detail, based on our own research and knowledge acquired from scientific literature. It should be pointed out that the acquired knowledge about electrochemical behaviour of **PAOX** can implicate its subsequent application (as a biologically active compound). Our biological assay indicated the antifungal activity of pyrazine-2-amidoxime against *C. albicans*. **PAOX** is active against both Gram-positive and Gram-negative bacteria. The **PAOX**'s binding ability to d-block metal ions was examined in aqueous and acetonitrile solutions. The obtained results showed that pyrazine-2-amidoxime behaves like a bidentate ligand and it coordinates selected metal ions by azomethine and amine nitrogen atoms, respectively.

## Experimental protocols

### Materials and methods

All starting materials were commercially available and used as such unless otherwise noted. Pyrazine-2-amidoxime of analytical purity grade (≥97%) was purchased from a commercial



source (Sigma Aldrich). Twice distilled water (Hydrolab-Reference purified) with conductivity not exceeding  $0.09 \mu\text{S cm}^{-1}$  was used in the preparation of all the solutions under study.

### X-ray crystallography

X-ray quality crystals of pyrazine-2-amidoxime (Fig. 14) were obtained from diffusion crystallization. 2 mg of pyrazine-2-amidoxime was dissolved in 1 mL of acetonitrile (MeCN), and 2.5 mL of tetrahydropyran (THP) was used as an antisolvent in a closed container. The sample was left to crystallize at room temperature for a period of over two weeks. All solvents used during crystallization were purified before by standard procedures. The elemental analysis confirmed the composition of the obtained PAOX (as white crystals) with  $\pm 0.3\%$  of the theoretical values {calcd for  $\text{C}_5\text{H}_6\text{N}_4\text{O}$  (%) C, 43.48; H, 4.38; N, 40.56 and found C, 43.50; H, 4.41; N, 40.59}.

The X-ray diffraction (XRD) measurement was carried out on an Oxford Diffraction Gemini R ULTRA Ruby CCD diffractometer with  $\text{CuK}\alpha$  ( $\lambda = 1.54184$ ) radiation source. Data were reduced using CrysAlis RED software.<sup>50</sup> The structural resolution procedure was carried out using the SHELX package (for details see: ESI†).<sup>51</sup> All interactions were determined using the PLATON program.<sup>52</sup> The ORTEP<sup>53</sup> and Mercury<sup>54</sup> programs were used to prepare the molecular graphics.

### Computational details

Geometries of five dimers and one tetramer were fully optimized with the use of the DFT method along with B3LYP functional,<sup>55–57</sup> and the 6-311+G(d,p) basis set.<sup>58</sup> Basis set superposition error (BSSE) was also accounted for in the calculations of energy formations for dimers and a tetramer.<sup>59</sup> All calculations were performed *in vacuo* without symmetry constraints, according to the procedures included in the Gaussian-03 series of programs.<sup>60</sup>

### Analytical instrumentations

The percentage composition of the elements of the examined compounds was determined with the use of an element analyzer Carlo Erba EA 1108 CHNS. The infrared spectra of pyrazine-2-amidoxime was recorded as potassium bromide (KBr) pellets using Brüker Infrared Spectrometer in the range of  $4500\text{--}500 \text{ cm}^{-1}$ . The electronic spectra were recorded on a Perkin Elmer Lambda 650 spectrophotometer in the range of  $220\text{--}400 \text{ nm}$  with a spectral band width  $2 \text{ nm}$ . The concentrations of compounds in acetonitrile solution considered in spectrophotometric measurements were around  $10^{-6} \text{ M}$ . Raman FRA 106 – laser NdYAG ( $1060 \text{ nm}$ ) with MCT detector was used to register Raman's spectra of solid state PAOX under study.

The thermogravimetric and IR spectra were measured by means of a thermal equaliser TG209 Nietzsche equipment with FT-IR. All experiments were carried out in an argon atmosphere. The analyzer was equipped with a programmed temperature controller, which automatically maintains constant temperature during thermal events. The TG weight-loss measurements were made at  $20\text{--}1000^\circ\text{C}$  at a heating rate of  $15^\circ\text{C min}^{-1}$ . The

infrared spectra were registered in Nujol mulls using a Brüker IFS 66 spectrophotometer. All measurements were verified at least twice.

CV experiments were done on an Gamry Potentiostat/Galvanostat (model 600). Data were collected and were analysed by the model Framework Analysis System Software on a PC computer. Platinum wire electrodes were used as both counter and working electrodes, and combined calomel electrode was used as a reference electrode (acetonitrile with an electrolyte inside). Solution cyclic voltammetry was carried out in a acetonitrile as a solvent containing tetrabutylammonium perchlorate TBP ( $0.1 \text{ M}$ ) as an electrolyte for PAOX ( $10^{-3} \text{ M}$ ). All solutions in the cell were purged with ultra-purity argon (Ar) for  $10\text{--}15$  minutes before each experiment, and a blanket of Ar was used during the experiment.

### Minimal bactericidal or fungicidal concentrations

The procedure of microbiological tests was described in our previous paper.<sup>35</sup> For details see ESI.†

## Conflict of interest

The authors declare no competing financial interest.

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