



CrossMark
 click for updates

Cite this: *RSC Adv.*, 2016, 6, 16150

Dye-sensitized solar cell based on an inclusion complex of a cyclic porphyrin dimer bearing four 4-pyridyl groups and fullerene C₆₀[†]

Yousuke Ooyama,^{*a} Koji Uenaka,^a Takuya Kamimura,^b Shuwa Ozako,^b Masahiro Kanda,^a Taro Koide^b and Fumito Tani^{*b}

Cyclic free-base porphyrin dimers (H₄-C₄-CPD_{Py}(TEO) and H₄-Ptz-CPD_{Py}(TEO)) linked by butadiyne or phenothiazine bearing four 4-pyridyl groups and their inclusion complexes (C₆₀⊂H₄-C₄-CPD_{Py}(TEO) and C₆₀⊂H₄-Ptz-CPD_{Py}(TEO)) with fullerene C₆₀ have been applied to dye-sensitized solar cells (DSSCs) as a new class of porphyrin dye sensitizers with pyridyl anchoring groups for attachment on a TiO₂ electrode. The FTIR spectra of the porphyrin dimers adsorbed on TiO₂ nanoparticles demonstrated that these porphyrin dimers are adsorbed on the TiO₂ surface through the formation of hydrogen bonding of pyridyl groups and/or pyridinium ions at Brønsted acid sites on the TiO₂ surface. The adsorption amount of the porphyrin dimers adsorbed on the TiO₂ electrode is 2.0 × 10¹⁷ molecules per cm², that is, the adsorption amount of the porphyrin unit is 4.0 × 10¹⁷ cm⁻², which is higher than that of dye sensitizers with pyridyl groups reported so far. The photovoltaic performance of DSSCs based on phenothiazine-bridged cyclic porphyrin dimer H₄-Ptz-CPD_{Py}(TEO) is higher than that of DSSCs based on butadiyne-linked cyclic porphyrin dimer H₄-C₄-CPD_{Py}(TEO). Moreover, the photovoltaic performances of DSSCs based on cyclic free-base porphyrin dimers are higher than those of DSSCs based on their C₆₀ inclusion complexes C₆₀⊂H₄-C₄-CPD_{Py}(TEO) and C₆₀⊂H₄-Ptz-CPD_{Py}(TEO). On the basis of the electrochemical measurements (voltammetry and electrochemical impedance spectroscopy) and the transient absorption spectroscopy, the differences in the photovoltaic performances among these cyclic free-base porphyrin dimers are discussed from kinetic and thermodynamic considerations concerning the electron transfer processes in DSSCs.

Received 14th January 2016
 Accepted 1st February 2016

DOI: 10.1039/c6ra01131d

www.rsc.org/advances

Introduction

Dye-sensitized solar cells (DSSCs) employing dye-adsorbed TiO₂ electrodes are one of the most promising new renewable photovoltaic cells utilizing the sun as a free and inexhaustible energy source because of their interesting construction and operational principles, and low cost of production, since Grätzel and co-workers produced high-performance DSSCs based on a Ru-complex dye, which showed a solar energy-to-electricity conversion yield (η) of 11%.¹ To further improve the photovoltaic performance of DSSCs, many kinds of ruthenium (Ru) dyes, porphyrin dyes, phthalocyanine dyes and organic dyes bearing carboxyl groups as anchoring groups, which are adsorbed on

the TiO₂ electrode through the bidentate bridging linkage between the carboxyl group of the dye and Brønsted acid sites (surface-bound hydroxyl groups, Ti-OH) on the TiO₂ surface, have been developed as dye sensitizers during the last two decades.²⁻⁹ In particular, porphyrin dyes have been regarded as promising candidates for photosensitizers as a result of their strong Soret (400–500 nm) and moderate Q band (500–700 nm) absorption properties, as well as their electrochemical, photochemical and thermal stabilities. Much effort in molecular design and development of porphyrin dye sensitizers having carboxyl group have been made to further improve the photovoltaic performances of DSSCs so far.¹⁰⁻¹⁵ Consequently, DSSCs based on donor- π -acceptor (D- π -A) porphyrin dyes bearing the diarylamino group as an electron donor and the benzothiadiazole-benzoic acid moiety as an electron acceptor, which exhibited good absorption features (bathochromic shift and broadening of the Soret and Q bands), have achieved η value of up to ca. 13%.^{11d,e}

On the other hand, we have reported that a new type of D- π -A dye sensitizers bearing pyridyl group as electron-withdrawing anchoring group were predominantly adsorbed on the TiO₂ electrode through coordinate bonding between the pyridyl

^aDepartment of Applied Chemistry, Graduate School of Engineering, Hiroshima University, Higashi-Hiroshima 739-8527, Japan. E-mail: yooyama@hiroshima-u.ac.jp; Fax: +81-82-424-5494

^bInstitute for Materials Chemistry and Engineering, Kyushu University, 744 Motooka, Nishi-ku, Fukuoka 819-0395, Japan. E-mail: tanif@ms.ifoc.kyushu-u.ac.jp; Fax: +81-92-802-6224

[†] Electronic supplementary information (ESI) available: ¹H NMR and ¹³C NMR for compounds 1, 2 and H₄-C₄-CPD_{Py}(TEO). See DOI: 10.1039/c6ra01131d



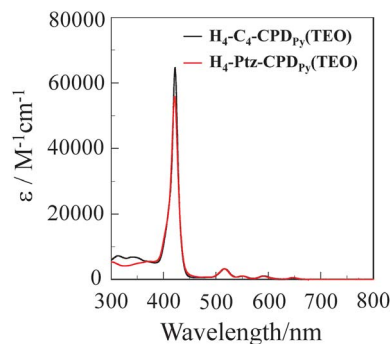


Fig. 1 UV/vis absorption spectra of $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$ and $\text{H}_4\text{-Ptz-CPDPy(TEO)}$ in benzonitrile.

obtain their exact absorption spectra because the 1 : 1 complex of $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$ or $\text{H}_4\text{-Ptz-CPDPy(TEO)}$ with C_{60} is in dissociation equilibrium in solution of 10^{-5} to 10^{-6} M concentration which is suitable for the measurement of absorption spectra of porphyrins. In our previous work, however, we have demonstrated that upon addition of C_{60} to the benzonitrile solution of the cyclic porphyrin dimers, their Soret bands were redshifted with a decrease in intensity, whereas their Q bands were slightly redshifted but increased in intensity.²⁰

The UV/vis absorption spectra of cyclic free-base porphyrin dimers and their C_{60} inclusion complexes adsorbed on TiO_2 film are shown in Fig. 2. It is worth mentioning here that the adsorption amount of the porphyrin dimers adsorbed on TiO_2 film is 2.0×10^{17} molecules per cm^2 , that is, the adsorption amount of porphyrin unit is 4.0×10^{17} cm^{-2} , which is higher than those ($<2.0 \times 10^{17}$ molecules per cm^2) of porphyrin dye sensitizers and D- π -A dye sensitizer bearing pyridyl groups reported so far.^{16,18} The Soret bands of $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$ and $\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$ were redshifted compared to those of $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$ and $\text{H}_4\text{-Ptz-CPDPy(TEO)}$, although there is little difference in the Q band between the cyclic free-base porphyrin dimers and their C_{60} inclusion complexes. Thus, the UV/vis absorption spectra of these porphyrin dimers-adsorbed TiO_2 films are in good agreement with those in benzonitrile.

Electrochemical and photochemical properties of cyclic free-base porphyrin dimers and their inclusion complexes with fullerene C_{60}

The electrochemical properties of cyclic free-base porphyrin dimers and their C_{60} inclusion complexes were determined by

cyclic voltammetry (CV) and differential pulse voltammetry (DPV) (see Fig. S7–S10 in ESI†).²⁰ The oxidation potential of the porphyrin unit in $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$ (0.74 V vs. Fc/Fc^+) showed anodic shift by 0.03 V, compared with that of $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$ (0.71 V) (Table 1). The oxidation potential of the porphyrin unit in $\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$ (0.75 V) also showed anodic shift by 0.03 V, compared with that of $\text{H}_4\text{-Ptz-CPDPy(TEO)}$ (0.72 V). The reduction potentials corresponding to the reduction of the fullerene entity of $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$ and $\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$ were observed at -0.94 V and -0.96 V, respectively, which is cathodically shifted by 0.02 V and 0.04 V, respectively, compared to pristine C_{60} (-0.92 V).^{20e} The small anodic shift of the oxidation potential of the porphyrin and the small cathodic shift of the reduction potential of C_{60} compared with their reference compounds is indicative of the charge transfer interaction between the porphyrins and C_{60} . The HOMO and LUMO energy levels were evaluated from the oxidation wave and the singlet excited energy of these porphyrin dimers (1.90 eV) based on the Q absorption band and fluorescence band in benzonitrile (*ca.* 650 nm), respectively. The HOMO and LUMO energy levels of these porphyrin dimers was *ca.* -5.5 eV and *ca.* -3.6 eV, respectively (Table 1). Thus, this result shows that the HOMO energy levels are more positive than the I_3^-/I^- redox potential (-4.9 eV), and thus this indicates that an efficient regeneration of the oxidized porphyrin dimers by electron transfer from the I_3^-/I^- redox couple in the electrolyte is thermodynamically feasible. Evidently, the LUMO energy levels of these porphyrin dimers are higher than the energy level (E_{cb}) of the CB of TiO_2 (-4.0 eV), suggesting that an electron injection to the CB of TiO_2 is thermodynamically feasible (some researchers have proposed that an energy gap of over 0.2–0.3 eV is necessary for efficient electron injection).^{2–7}

Time-resolved absorption spectroscopy for cyclic free-base porphyrin dimers and their C_{60} inclusion complexes was performed by femtosecond laser flash photolysis after photoexcitation.²⁰ The photodynamics of these porphyrin dimers is summarized in Fig. 3. The decay of photoexcited state of $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$ has two steps: the first step has a lifetime of 18 ps, which corresponds to the disappearance of the singlet excited state of the $^1\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}^*$ (*ca.* -3.6 eV), that is, the $^1\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}^*$ undergoes intrasupramolecular electron transfer to give a completely charge-separated state $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}^{\cdot+}$ (*ca.* -3.7 eV). The $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}^{\cdot+}$ decays with a lifetime of 470 ps to the ground state. The decay of photoexcited state of $\text{C}_{60}\text{-H}_4\text{-Ptz-}$

Table 1 Optical, electrochemical data and HOMO and LUMO energy levels of $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$, $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$, $\text{H}_4\text{-Ptz-CPDPy(TEO)}$, and $\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$

Porphyrin dimers	$\lambda_{\text{max}}^{\text{abs}}/\text{nm}$ ($\epsilon/\text{M}^{-1}\text{cm}^{-1}$) ^a	$E_{\text{pa}}^{\text{oxb}}/\text{V}$ vs. Fc/Fc^+	HOMO ^c /eV	LUMO ^c /eV
$\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$	422 (646 000), 517 (30 800), 551 (9800), 590 (9600), 647 (4300)	0.71	-5.51	-3.61
$\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$	—	0.74	-5.54	-3.64
$\text{H}_4\text{-Ptz-CPDPy(TEO)}$	421 (55 800), 516 (32 100), 550 (10 500), 590 (9500), 646 (5600)	0.72	-5.52	-3.62
$\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$	—	0.75	-5.55	-3.65

^a In benzonitrile. ^b Anodic (E_{pa}) peak potentials for oxidation were recorded in benzonitrile/ Bu_4NPF_6 (0.1 M) solution for the porphyrin unit in $\text{H}_4\text{-C}_4\text{-CPDPy(TEO)}$, $\text{C}_{60}\text{-H}_4\text{-C}_4\text{-CPDPy(TEO)}$, $\text{H}_4\text{-Ptz-CPDPy(TEO)}$ and $\text{C}_{60}\text{-H}_4\text{-Ptz-CPDPy(TEO)}$. ^c HOMO and LUMO energy levels of the porphyrin unit.



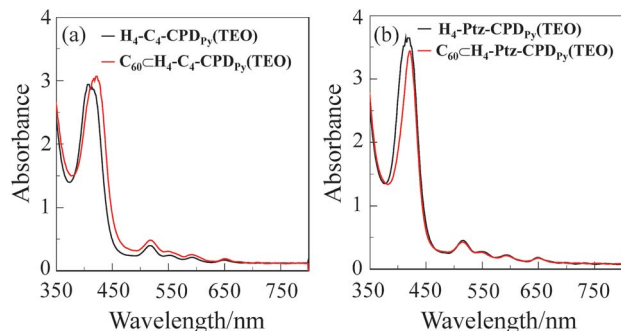


Fig. 2 UV/vis absorption spectra of (a) H₄-C₄-CPDPy(TEO) and C₆₀-H₄-C₄-CPDPy(TEO) and (b) H₄-Ptz-CPDPy(TEO) and C₆₀-H₄-Ptz-CPDPy(TEO) adsorbed on TiO₂ film (3 μm).

CPDPy(TEO) also includes the charge-separated states C₆₀⁻-H₄-Ptz-CPDPy(TEO)⁺ (ca. -3.8 eV), but has multiple steps to the ground state *via* the triplet charge separated state ³(C₆₀⁻-H₄-Ptz⁺-CPDPy(TEO)) (ca. -4.3 eV) with a lifetime of 0.71 ms: the first step has a lifetime of 20 ps, which is similar to that (18 ps) of C₆₀-H₄-C₄-CPDPy(TEO). On the other hand, the energy diagrams for the photochemical events in these porphyrin dimers revealed that the energy levels of charge-separated states is very close to or lower than the E_{cb} of the CB of TiO₂ electrode. This suggests that the electron injection from the C₆₀⁻ in the charge-separated states to the CB of TiO₂ electrode is thermodynamically difficult.

FTIR spectra of cyclic free-base porphyrin dimers and their inclusion complexes with fullerene C₆₀

To elucidate the adsorption states of cyclic free-base porphyrin dimers and their C₆₀ inclusion complexes on TiO₂ nanoparticles, we measured the FTIR spectra of the porphyrin dimer powders and the porphyrin dimers adsorbed on TiO₂ nanoparticles (Fig. 4). For the powders of all the four porphyrin dimers H₄-C₄-CPDPy(TEO), H₄-Ptz-CPDPy(TEO), C₆₀-H₄-C₄-CPDPy(TEO), and C₆₀-H₄-Ptz-CPDPy(TEO), the C=N stretching band of pyridyl group was clearly observed at around 1590 cm⁻¹. Interestingly, when these porphyrin dimers were adsorbed on the TiO₂ surface, a new band appeared at around 1650 cm⁻¹, which indicates the formation of a pyridinium ion with Brønsted acid sites on TiO₂ surface.²¹ In addition, the C=N

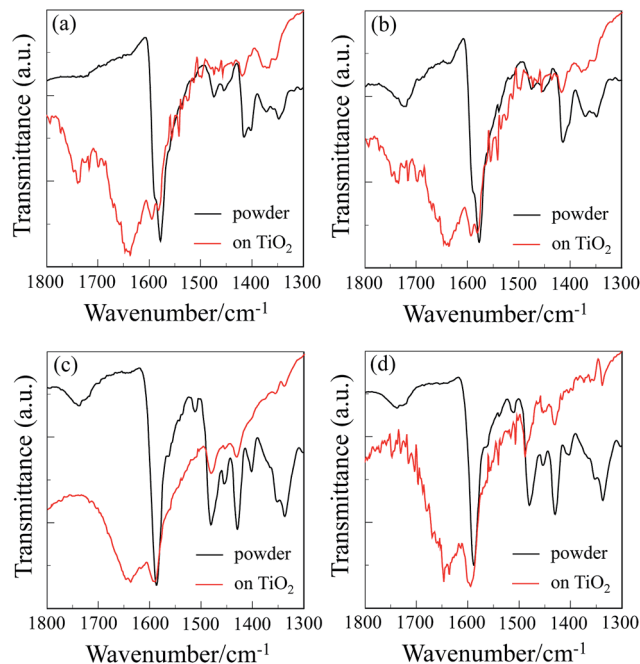


Fig. 4 FTIR spectra of the porphyrin dimer powders and the porphyrin dimers adsorbed on TiO₂ nanoparticles for (a) H₄-C₄-CPDPy(TEO), (b) C₆₀-H₄-C₄-CPDPy(TEO), (c) H₄-Ptz-CPDPy(TEO), and (d) C₆₀-H₄-Ptz-CPDPy(TEO).

stretching band at around 1590 cm⁻¹ is shifted by 3–5 cm⁻¹ to higher wavenumber, that is, the resulting band can be assigned to the hydrogen-bonded pyridyl group to Brønsted acid sites on the TiO₂ surface. Consequently, these observations demonstrate that these porphyrin dimers are adsorbed on the TiO₂ surface through the formations of hydrogen bonding of pyridyl groups and/or pyridinium ion at Brønsted acid sites on the TiO₂ surface. Thus, the UV/vis absorption and the FTIR spectra of the porphyrin dimers adsorbed on TiO₂ film indicate that high dye loading and high surface coverage of the TiO₂ electrode for DSSCs based on porphyrin dye sensitizers bearing pyridyl group are achieved by employing the cyclic porphyrin dimers bearing four pyridyl anchoring groups possessing the bonding ability to the two points on Brønsted acid sites on TiO₂ surface.

Photovoltaic performances of DSSCs based on cyclic free-base porphyrin dimers and their inclusion complexes with fullerene C₆₀

The DSSC was prepared using the dye-adsorbed TiO₂ electrode (9 μm), Pt-coated glass as a counter electrode, and an acetonitrile solution with iodine (0.05 M), lithium iodide (0.1 M), and 1,2-dimethyl-3-propylimidazolium iodide (0.6 M) as an electrolyte. The photocurrent-voltage (*I*-*V*) characteristics were measured under simulated solar light (AM 1.5, 100 mW cm⁻²). The incident photon-to-current conversion efficiency (IPCE) spectra and the *I*-*V* curves are shown in Fig. 5. The photovoltaic performance parameters are collected in Table 2. The IPCE values corresponding to the Soret band (29% at 424 nm) and the Q band (6% at 520 nm) for the phenothiazine-bridged cyclic

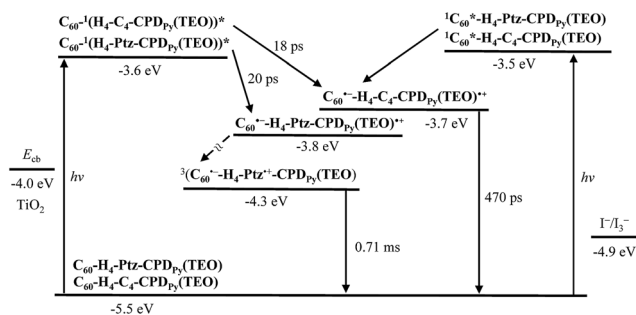


Fig. 3 Energy diagrams for the photochemical events in C₆₀-H₄-C₄-CPDPy(TEO) and C₆₀-H₄-Ptz-CPDPy(TEO).



porphyrin dimer $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ are higher than those (15% at 420 nm for the Soret band and 2% at 522 nm for the Q band) of the butadiyne-linked cyclic porphyrin dimer $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (Fig. 5a). The I - V curves show that the short-circuit photocurrent density (J_{sc}) and η values of $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (1.40 mA cm^{-2} and 0.35%) are higher than those of $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (0.79 mA cm^{-2} and 0.19%) (Fig. 5b). The higher photovoltaic performance for DSSC based on $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ may be attributed to the efficient electron injection from the photoexcited dye ($^1\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})^*$) to the CB of TiO_2 electrode because the phenothiazine unit possessing electron donating ability can provide a unidirectional flow of electrons toward the pyridyl anchoring group upon photoexcitation of the porphyrin. Interestingly, the photovoltaic performances of DSSC based on the C_{60} inclusion complexes $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ are lower than those of $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ and $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$. The electron injection ($\approx 100 \text{ ps}$) from the photoexcited porphyrin to the CB of TiO_2 electrode is in kinetically competition with the intrasupramolecular electron transfer (18–20 ps), that is, the formation of charge-separated state $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})^{\cdot+}$ and $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})^{\cdot+}$ (Fig. 3). Therefore, the lower photovoltaic performances for DSSC based on the C_{60} inclusion complexes would be attributed to the formation of charge-separated state, leading to low electron-injection efficiency from the photoexcited porphyrin to CB of TiO_2 electrode. On the other hand, the photovoltaic performance of DSSC based on $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ is slightly lower than that of $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, which may be attributed to the thermodynamically difficult electron-injection due to the lower triplet charge separated state $^3(\text{C}_{60}^{\cdot-}\text{-H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO}))$.

Moreover, it is worth mentioning here that the open-circuit photovoltage (V_{oc}) values of the C_{60} inclusion complexes $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (315 mV) and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (304 mV) are lower than those of $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (376 mV) and $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (396 mV). Thus, electrochemical impedance spectroscopy (EIS) analysis was performed to study the electron recombination process in DSSCs based on these porphyrin dimers in the dark under a forward bias of -0.60 V with a frequency range of 10 mHz to 100 kHz. The large

Table 2 DSSC performance parameters of $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$, and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$

Porphyrin dimers ^a	$J_{\text{sc}}/\text{mA cm}^{-2b}$	$V_{\text{oc}}^b/\text{mV}$	ff^b	η^b (%)
$\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$	0.79	376	0.64	0.19
$\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$	0.27	315	0.69	0.06
$\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$	1.40	396	0.63	0.35
$\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$	0.24	304	0.54	0.04

^a The adsorption amount of the porphyrin dimers adsorbed on TiO_2 electrode is 2.0×10^{17} molecules per cm^2 under the adsorption condition of 0.1 mM porphyrin dimer solution in chloroform. The amount of adsorbed dye on TiO_2 electrode was determined from the calibration curve by absorption spectral measurement of the concentration change of the dye solution before and after adsorption.

^b Under a simulated solar light (AM 1.5, 100 mW cm^{-2}).

semicircle in the Nyquist plot (Fig. 6a), which corresponds to the midfrequency peaks in the Bode phase plots, represents the charge recombination between the injected electrons in TiO_2 and I_3^- ions in the electrolyte, that is, the charge-transfer resistances at the $\text{TiO}_2/\text{dye}/\text{electrolyte}$ interface. The Nyquist plots show that the resistance values of the large semicircle for $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (13 Ω) and $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (14 Ω) are higher than those of $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (5 Ω) and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (5 Ω), indicating that the electron recombination resistances for $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ and $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ are larger than those of $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$. The electron recombination lifetimes (τ_e) expressing the electron recombination between the injected electrons in TiO_2 and I_3^- ions in the electrolyte, extracted from the angular frequency (ω_{rec}) at the midfrequency peak in the Bode phase plot (Fig. 6b) using $\tau_e = 1/\omega_{\text{rec}}$, increase in the order of $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (0.1 ms) < $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (0.1 ms) < $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$ (2.2 ms) < $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$ (2.5 ms), respectively, which are consistent with the sequence of V_{oc} values in the DSSCs. Consequently, this is attributed to the fact that the formation of charge-separated state $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})^{\cdot+}$ and $\text{C}_{60}^{\cdot-}\text{-H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})^{\cdot+}$ facilitates the approach of I_3^- ions to the TiO_2 surface because of complexation between the charge-separated state and I_3^-

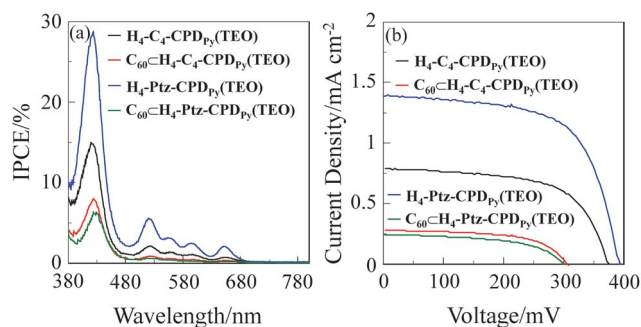


Fig. 5 (a) IPCE spectra and (b) I - V curves of DSSCs based on $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$, and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$.

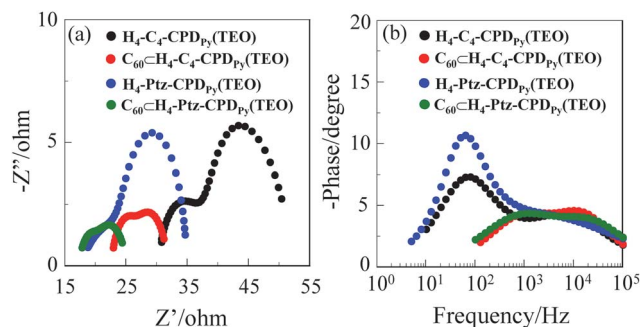


Fig. 6 (a) Nyquist plots and (b) Bode phase plots of DSSCs based on $\text{H}_4\text{-C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{C}_{60}\text{C}_4\text{-CPD}_{\text{Py}}(\text{TEO})$, $\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$, and $\text{C}_{60}\text{H}_4\text{-Ptz-CPD}_{\text{Py}}(\text{TEO})$.



ions, leading to faster charge recombination and thus resulting in a lower V_{oc} value.^{3,6}

Conclusions

In this work, to seek a direction in molecular design toward creating a new class of porphyrin dye sensitizers bearing pyridyl anchoring groups for achieving high dye loading and high surface coverage of the TiO_2 electrode for dye-sensitized solar cells (DSSCs), cyclic free-base porphyrin dimers ($H_4-C_4-CPD_{Py}(TEO)$ and $H_4-Ptz-CPD_{Py}(TEO)$) linked by butadiyne or phenothiazine bearing 4-pyridyl groups and their C_{60} inclusion complexes ($C_{60} \subset H_4-C_4-CPD_{Py}(TEO)$ and $C_{60} \subset H_4-Ptz-CPD_{Py}(TEO)$) have been applied to DSSCs as the dye sensitizers. The Soret bands of $C_{60} \subset H_4-C_4-CPD_{Py}(TEO)$ and $C_{60} \subset H_4-Ptz-CPD_{Py}(TEO)$ were redshifted with a decrease in intensity, whereas their Q bands were slightly redshifted but increased in intensity, compared to $H_4-C_4-CPD_{Py}(TEO)$ and $H_4-Ptz-CPD_{Py}(TEO)$. These porphyrin dimers are adsorbed on the TiO_2 surface through the formations of hydrogen bonding of pyridyl groups and/or pyridinium ion at Brønsted acid sites on the TiO_2 surface. It was found that the adsorption amount of the porphyrin dimers adsorbed on TiO_2 film is 2.0×10^{17} molecules per cm^2 , that is, the adsorption amount of porphyrin unit is $4.0 \times 10^{17} cm^{-2}$, which is higher than those ($<2.0 \times 10^{17}$ molecules per cm^2) of porphyrin dye sensitizers and D- π -A dye sensitizer bearing pyridyl groups reported so far. The higher photovoltaic performance for DSSC based on $H_4-Ptz-CPD_{Py}(TEO)$ relative to $H_4-C_4-CPD_{Py}(TEO)$ may be attributed to the efficient electron injection from the photoexcited dye ($^1H_4-Ptz-CPD_{Py}(TEO)^*$) to the CB of TiO_2 because the phenothiazine unit possessing electron donating ability can provide a unidirectional flow of electrons toward the pyridyl anchoring group upon photoexcitation of the porphyrin. On the other hand, the photovoltaic performances of DSSC based on $C_{60} \subset H_4-C_4-CPD_{Py}(TEO)$ and $C_{60} \subset H_4-Ptz-CPD_{Py}(TEO)$ are lower than those of $H_4-C_4-CPD_{Py}(TEO)$ and $H_4-Ptz-CPD_{Py}(TEO)$. The transient absorption spectroscopy and electrochemical measurements revealed that the lower photovoltaic performances for DSSC based on the C_{60} inclusion complexes would be attributed to the formation of charge-separated state between the photoexcited porphyrin and C_{60} , leading to low electron-injection efficiency from the photoexcited porphyrin to TiO_2 electrode. Consequently, this work provides that the cyclic porphyrin dimers with four pyridyl anchoring groups possessing bonding ability to the two points on Brønsted acid sites on TiO_2 surface, which can achieve the high dye loading and the high surface coverage of the TiO_2 electrode, would be expected to be one of the most promising classes of porphyrin dye sensitizers with pyridyl group.

Experimental

General

Infrared (IR) absorption spectra were recorded on a Perkin Elmer Spectrum One FT-IR spectrometer by ATR method and a Bio-Rad FTS6000 spectrometer by KBr disk method. Ultraviolet/visible (UV/vis) absorption spectra were observed

with Shimadzu UV-3150 and UV-2500PC spectrophotometers. Nuclear magnetic resonance (NMR) spectra were recorded on a Bruker AVANCE III 600 spectrometer. The resonance frequencies are 600 MHz for 1H and 151 MHz for ^{13}C . Chemical shifts were reported as δ values in ppm relative to tetramethylsilane. High-resolution fast atom bombardment mass spectra (HR-FAB-MS) were measured with 3-nitrobenzyl alcohol (NBA) as a matrix and recorded on a JEOL LMS-HX-110 spectrometer. Melting point (Mp) measurements were carried out on a Yanaco MP-S3 melting point meter. Cyclic voltammetry (CV) and differential pulse voltammetry (DPV) were carried out on an ALS630C (BAS, Japan) electrochemical analyzer in deaerated benzonitrile containing 0.1 M tetra-*n*-butylammonium hexafluorophosphate (Bu_4NPF_6) as a supporting electrolyte at room temperature under an Ar atmosphere. A conventional three-electrode cell was used with a platinum working electrode (surface area of $0.3 mm^2$) and a platinum wire as a counter electrode. The measured potentials were recorded with respect to an Ag/AgNO₃ (0.01 M) reference electrode. All potentials (*vs.* Ag⁺/Ag) were converted to values with respect to that of the Fc⁺/Fc redox couple, which was measured under the same conditions. The highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO) energy levels of porphyrin dimers were evaluated from the spectral analyses and the DPV data. The HOMO energy level was evaluated from the oxidation wave for porphyrin dimers. The LUMO energy level was estimated from the singlet excited energy of these porphyrin dimers (1.90 eV) based on the Q absorption band nm and fluorescence band in benzonitrile.

Synthesis

5,15-Bis[3-[2-(2-ethoxyethoxy)ethoxy]-5-[(trimethylsilyl)ethynyl]phenyl]-10,20-di[4-pyridyl]porphine (1). A solution of 3-[2-(2-ethoxyethoxy)ethoxy]-5-[(trimethylsilyl)ethynyl]benzaldehyde²² (4.63 g, 13.8 mmol) and *meso*-(4-pyridyl)-dipyrrromethane²³ (3.08 g, 13.9 mmol) in CH_2Cl_2 (1.4 L) was degassed with a stream of N_2 for 30 min. To this solution, trifluoroacetic acid (TFA) (12.3 mL, 166 mmol) was added and stirred under a N_2 atmosphere. After 30 min, the resultant mixture was neutralized with NEt_3 (23 mL, 0.17 mol) and then 2,3-dichloro-5,6-dicyano-1,4-benzoquinone (DDQ) (4.7 g, 21 mmol) dissolved in benzene (250 mL) was added to the mixture. After an overnight stirring, the resultant mixture was purified by column chromatography (silica-gel, $CHCl_3/EtOH = 50/1$) then recrystallized from $CHCl_3/i$ -PrOH to give **1** as a purple solid (1.27 g, 18%). 1H NMR ($CDCl_3$, 600 MHz): δ -2.91 (br-s, 2H, -NH), 0.27 (s, 18H, -Si(CH_3)₃), 1.18 (t, $J = 7.2$ Hz, 6H, -O(CH_2CH_2O)₂CH₂CH₃), 3.52 (q, $J = 7.2$ Hz, 4H, -O(CH_2CH_2O)₂CH₂CH₃), 3.63 (dd, $J = 4.8$ and 6.0 Hz, 4H, -O(CH_2)₂OCH₂CH₂OC₂H₅), 3.76 (t, $J = 4.8$ Hz, 4H, -O(CH_2)₂OCH₂CH₂OC₂H₅), 3.96 (t, $J = 4.8$ Hz, 4H, -OCH₂CH₂O(CH_2)₂OC₂H₅), 4.36 (t, $J = 4.8$ Hz, 4H, -OCH₂CH₂O(CH_2)OC₂H₅), 7.47 (br-s, 2H, Ar-H), 7.75 (br-s, 2H, Ar-H), 7.93 (br-s, 2H, Ar-H), 8.17 (d, $J = 5.4$ Hz, 4H, Ar-H), 8.81 (d, $J = 4.2$ Hz, 4H, pyrrole β -H), 8.92 (d, $J = 4.8$ Hz, 4H, pyrrole β -H), 9.06 (d, $J = 6.0$ Hz, 4H, Ar-H); ^{13}C NMR ($CDCl_3$, 151 MHz): δ 0.1, 15.3, 66.9, 68.1, 69.9, 70.0, 71.2, 94.7, 104.9, 117.2, 117.4, 119.7, 122.6, 122.6, 129.6, 131.1,



143.1, 148.5, 150.3, 157.2; HR-FAB-MS (NBA): m/z calcd for $C_{64}H_{68}N_6O_6Si_2$: 1072.4739; found: 1072.4718; UV/vis ($CHCl_3$): λ_{max} (ϵ , $cm^{-1} M^{-1}$) 421 (392 000), 515 (18 200), 549 (5400), 588 (5400), 645 (2100); IR (KBr): $\nu = 3439, 3317, 3092, 3022, 2956, 2928, 2867, 2714, 2614, 2533, 2156, 1591, 1582, 1476, 1420, 1404, 352, 1323, 1297, 1247, 1196, 1173, 1161, 1114, 1069, 976, 927, 847, 800, 788, 760, 730, 697, 660, 563, 411 cm^{-1}$; Mp: $>300^\circ C$.

{5,15-Di[3-[2-(2-ethoxyethoxy)ethoxy]-5-ethynyl]phenyl}-10,20-di[4-pyridyl]porphinato}zinc(II) (2). Zn(OAc) $_2 \cdot 2H_2O$ (2.4 g, 11 mmol) dissolved in MeOH (60 mL) was added to a solution of **1** (1.16 g, 1.08 mmol) in CH_2Cl_2 (1.1 L). The resultant mixture was refluxed overnight under N_2 atmosphere and then cooled to room temperature. The reaction mixture was neutralized with saturated aqueous solution of $NaHCO_3$ (200 mL), washed with water (200 mL, 2 times), dried over Na_2SO_4 and evaporated. The residue and $KF \cdot 2H_2O$ (0.63 g, 6.7 mmol) were dissolved in DMF (40 mL) and the mixture was stirred overnight under a N_2 atmosphere. After the stirring, the resultant mixture was diluted with $CHCl_3$ (100 mL) and washed with water (500 mL, 2 times). The organic layer was dried over Na_2SO_4 and evaporated to give the product as a purple solid (1.00 g, 99%). 1H NMR (pyridine- d_5 , 600 MHz): δ 1.00 (t, $J = 7.2$ Hz, 6H, $-O(CH_2CH_2O)_2CH_2CH_3$), 3.33 (q, $J = 6.6$ Hz, 4H, $-O(CH_2CH_2O)_2CH_2CH_3$), 3.52 (t, $J = 4.8$ Hz, 4H, $-O(CH_2)_2OCH_2CH_2OC_2H_5$), 3.67 (t, $J = 4.8$ Hz, 4H, $-O(CH_2)_2OCH_2CH_2OC_2H_5$), 3.87 (t, $J = 4.8$ Hz, 4H, $-OCH_2CH_2-O(CH_2)_2OC_2H_5$), 4.15 (s, 2H, $-C\equiv CH$), 4.31 (t, $J = 4.8$ Hz, 4H, $-OCH_2CH_2O(CH_2)OC_2H_5$), 7.68 (d, $J = 0.6$ Hz, 2H, Ar-H), 8.05 (br-s, 2H, Ar-H), 8.23 (d, $J = 4.2$ Hz, 2H, Ar-H), 8.26 (d, $J = 4.2$ Hz, 4H, Ar-H), 9.02 (d, $J = 4.8$ Hz, 4H, pyrrole β -H), 9.11 (d, $J = 5.4$ Hz, 4H, Ar-H), 9.15 (d, $J = 4.8$ Hz, 4H, pyrrole β -H); ^{13}C NMR (pyridine- d_5 , 151 MHz): δ 15.5, 66.5, 68.5, 70.0, 70.3, 71.2, 79.6, 84.5, 117.6, 118.7, 120.6, 122.0, 130.1, 131.7, 132.2, 132.9, 145.2, 148.7, 149.8, 150.3, 150.7, 151.4, 157.7; HR-FAB-MS (NBA): m/z calcd for $C_{58}H_{50}N_6O_6Zn$: 990.3083; found: 990.3082; UV/vis ($CHCl_3$): λ_{max} (ϵ , $cm^{-1} M^{-1}$) 425 (418 000), 555 (16 200), 595 (2400); IR (KBr): $\nu = 3433, 3283, 2972, 2928, 2872, 1592, 1581, 1540, 1523, 1488, 1456, 1419, 1344, 1321, 1298, 1283, 1246, 1204, 1181, 1113, 1071, 1048, 998, 938, 874, 819, 796, 763, 719, 675, 665, 569, 545, 436, 419 cm^{-1}$; Mp: $>300^\circ C$.

H_4 -C $_4$ -CPDP $_2$ (TEO). The free-base dimer was prepared from **2** (397 mg, 0.4 mmol) with CuCl (8.0 g, 80 mmol) in pyridine (1.2 L) according to the reported procedure.^{20a,d} The crude product was purified by flash column chromatography (silica-gel, $CHCl_3/EtOH = 100/1$ to $10/1$) and recrystallized from $CHCl_3/i$ -PrOH to give the product as a purple powder (48 mg, 13%). 1H NMR ($CDCl_3$, 600 MHz): δ -3.04 (br-s, 4H, -NH), 1.21 (t, $J = 7.2$ Hz, 12H, $-O(CH_2CH_2O)_2CH_2CH_3$), 3.55 (q, $J = 7.2$ Hz, 8H, $-O(CH_2CH_2O)_2CH_2CH_3$), 3.67 (dd, $J = 4.8$ and 6.0 Hz, 8H, $-O(CH_2)_2OCH_2CH_2OC_2H_5$), 3.81 (dd, $J = 4.8$ and 6.0 Hz, 8H, $-O(CH_2)_2OCH_2CH_2OC_2H_5$), 4.03 (t, $J = 4.8$ Hz, 8H, $-OCH_2CH_2-O(CH_2)_2OC_2H_5$), 4.42 (t, $J = 4.8$ Hz, 8H, $-OCH_2CH_2O(CH_2)OC_2H_5$), 6.88 (s, 4H, Ar-H), 7.25 (dd, $J = 0.6$ and 1.8 Hz, 4H, Ar-H), 7.94 (br-s, 8H, Ar-H), 8.22 (t, $J = 1.8$ Hz, 4H, Ar-H), 8.62 (d, $J = 4.2$ Hz, 8H, pyrrole β -H), 8.70 (d, $J = 4.2$ Hz, 8H, pyrrole β -H), 8.96 (br-s, 8H, Ar-H); ^{13}C NMR ($CDCl_3$, 151 MHz): δ 15.3, 66.9, 68.3, 70.0, 70.1, 71.3, 74.5, 83.1, 114.9, 117.1, 118.6, 120.4, 120.8, 129.3, 135.3, 143.0, 148.4, 149.9, 157.8; HR-FAB-MS

(NBA): m/z calcd for $C_{116}H_{100}N_{12}O_{12}$: 1852.7584; found: 1852.7545; UV/vis (PhCN): λ_{max} (ϵ , $cm^{-1} M^{-1}$); 422 (646 000), 517 (30 800), 551 (9800), 590 (9600), 647 (4300); IR (ATR): $\nu = 3312, 2971, 2920, 2865, 1578, 1474, 1455, 1415, 1372, 1348, 1285, 1254, 1221, 1195, 1172, 1108, 1062, 993, 974, 930, 871, 849, 797, 728, 692, 659, 563, 540, 517, 467, 411 cm^{-1}$; Mp: $>300^\circ C$.

Preparation of DSSCs

The TiO_2 paste (JGC Catalysts and Chemicals Ltd., PST-18NR) was deposited on a fluorine-doped-tin-oxide (FTO) substrate by doctor-blading, and sintered for 50 min at $450^\circ C$. The 9 μm thick TiO_2 electrode was immersed into 0.1 mM porphyrin dimer solution in chloroform for 15 hours enough to adsorb the dye sensitizers. The DSSCs were fabricated by using the TiO_2 electrode ($0.5 \times 0.5 cm^2$ in photoactive area) thus prepared, Pt-coated glass as a counter electrode, and a solution of 0.05 M iodine, 0.1 M lithium iodide, and 0.6 M 1,2-dimethyl-3-propylimidazolium iodide in acetonitrile as electrolyte. The photocurrent-voltage characteristics were measured using a potentiostat under a simulated solar light (AM 1.5, 100 $mW cm^{-2}$). IPCE spectra were measured under monochromatic irradiation with a tungsten-halogen lamp and a monochromator. The amount of adsorbed dye on TiO_2 nanoparticles was determined from the calibration curve by absorption spectral measurement of the concentration change of the porphyrin dye solution before and after adsorption. Absorption spectra of the dyes adsorbed on TiO_2 nanoparticles were recorded on the porphyrin dyes-adsorbed TiO_2 film (thickness of 3 μm) in the transmission mode with a calibrated integrating sphere system. Electrochemical impedance spectroscopy (EIS) for DSSCs in the dark under a forward bias of $-0.60 V$ with a frequency range of 10 mHz to 100 kHz was measured with a AMETEK Versa STAT 3.

Acknowledgements

This work was supported by Grants-in-Aid (No. 20108009 and 15K05432 to F. T. and No. 15H03859 to Y. O.) from Ministry of Education, Culture, Sports, Science and Technology of Japan and Research Grants to F. T. from Tokuyama Science and Technology Foundation and Iketani Science and Technology Foundation.

Notes and references

- 1 B. O'Regan and M. Grätzel, *Nature*, 1991, **353**, 737.
- 2 A. Hagfeldt, G. Boschloo, L. Sun, L. Klöö and H. Pettersson, *Chem. Rev.*, 2010, **110**, 6595.
- 3 (a) Z. Ning and H. Tian, *Chem. Commun.*, 2009, 5483; (b) Z. Ning, Y. Fu and H. Tian, *Energy Environ. Sci.*, 2010, **3**, 1170.
- 4 A. Mishra, M. K. R. Fischer and P. Bäuerle, *Angew. Chem., Int. Ed.*, 2009, **48**, 2474.
- 5 Y. Ooyama and Y. Harima, *Eur. J. Org. Chem.*, 2009, **18**, 2903.
- 6 Y. Ooyama and Y. Harima, *ChemPhysChem*, 2012, **13**, 4032.
- 7 N. Manfredi, B. Ceconi and A. Abboto, *Eur. J. Org. Chem.*, 2014, 7069.



- 8 (a) X. Wang, J. Yang, H. Yu, F. Li, L. Fan, W. Sun, Y. Liu, Z. Y. Koh, J. Pan, W.-L. Yim, L. Yan and Q. Wang, *Chem. Commun.*, 2014, **50**, 3965; (b) S.-G. Li, K.-J. Jiang, J.-H. Huang, L.-M. Yang and Y.-L. Song, *Chem. Commun.*, 2014, **50**, 4309; (c) D. K. Panda, F. S. Goodson, S. Ray and S. Saha, *Chem. Commun.*, 2014, **50**, 5358; (d) K. Kakiage, Y. Aoyama, T. Yano, T. Otsuka, T. Kyomen, M. Unno and M. Hanaya, *Chem. Commun.*, 2014, **50**, 6379; (e) A. Amacher, C. Yi, J. yang, M. P. Bircher, Y. Fu, M. Cascella, M. Grätzel, S. Decurtins and S.-X. Liu, *Chem. Commun.*, 2014, **50**, 6540; (f) J. Yang, P. Ganesan, J. Teuscher, T. Moehl, Y. J. Kim, C. Yi, P. Comte, K. Pei, T. W. Holcombe, M. K. Nazeeruddin, J. Hua, S. K. Zakeeruddin, H. Tian and M. Grätzel, *J. Am. Chem. Soc.*, 2014, **136**, 5772.
- 9 (a) T. Ikeuchi, H. Nomoto, N. Masaki, M. J. Griffith, S. Mori and M. Kimura, *Chem. Commun.*, 2014, **50**, 1941; (b) R. Agosta, R. Grisorio, L. De Marco, G. Romanazzi, G. P. Suranna, G. Gigli and M. Manca, *Chem. Commun.*, 2014, **50**, 9451; (c) A. Dessi, M. Calamante, A. Mordini, M. Peruzzin, A. Sinicropi, R. Basosi, F. F. de Biani, M. Taddei, D. Colonna, A. D. Carlo, G. Reginato and L. Zani, *Chem. Commun.*, 2014, **50**, 13952; (d) X. Sun, Y. Wang, X. Li, H. Ågren, W. Zhu, H. Tian and Y. Xie, *Chem. Commun.*, 2014, **50**, 15609.
- 10 (a) H. Imahori, T. Umeyama and S. Ito, *Acc. Chem. Res.*, 2009, **42**, 1809; (b) S. Eu, S. Hayashi, T. Umeyama, Y. Matano, Y. Araki and H. Imahori, *J. Phys. Chem. C*, 2008, **112**, 4396; (c) A. Kira, Y. Matsubara, H. Iijima, T. Umeyama, Y. Matano, S. Ito, M. Niemi, N. V. Tkachenko, H. Lemmetyinen and H. Imahori, *J. Phys. Chem. C*, 2010, **114**, 11293; (d) H. Imahori, H. Iijima, H. Hayashi, Y. Toude, T. Umeyama, Y. Matano and S. Ito, *ChemSusChem*, 2011, **4**, 797; (e) K. Kurotobi, Y. Toude, K. Kawamoto, Y. Fujimori, S. Ito, P. Chabera, V. Sundström and H. Imahori, *Chem.–Eur. J.*, 2013, **19**, 17075; (f) T. Higashino and H. Imahori, *Dalton Trans.*, 2015, **44**, 448.
- 11 (a) T. Bessho, S. M. Zakeeruddin, C.-Y. Yeh, E. W.-G. Diau and M. Grätzel, *Angew. Chem., Int. Ed.*, 2010, **49**, 6646; (b) A. Yella, H.-W. Lee, H. N. Tsao, C. Yi, A. K. Chandiran, M. K. Nazeeruddin, E. W.-G. Diau, C.-Y. Yeh, S. M. Zakeeruddin and M. Grätzel, *Science*, 2011, **334**, 629; (c) L.-L. Li and E. W.-G. Diau, *Chem. Soc. Rev.*, 2013, **42**, 291; (d) S. Mathew, A. Yella, P. Gao, R. Humphry-Baker, B. F. E. Curchod, N. Ashari-Astani, I. Tavernelli, U. Rothlisberger, M. K. Nazeeruddin and M. Grätzel, *Nat. Chem.*, 2014, **6**, 242; (e) A. Yella, C.-L. Mai, S. M. Zakeeruddin, S.-N. Chang, C.-H. Hsieh, C.-Y. Yeh and M. Grätzel, *Angew. Chem., Int. Ed.*, 2014, **53**, 2973.
- 12 (a) M. V. Martínez-Díaz, G. de la Torre and T. Torres, *Chem. Commun.*, 2010, **46**, 7090; (b) M. J. Griffith, K. Sunahara, P. Wagner, K. Wagner, G. G. Wallace, D. L. Officer, A. Furube, R. Katoh, S. Mori and A. J. Mozer, *Chem. Commun.*, 2012, **48**, 4145; (c) K. Ladomenou, T. N. Kitsopoulos, G. D. Sharma and A. G. Coutsolelos, *RSC Adv.*, 2014, **4**, 21379.
- 13 (a) Y.-C. Chang, C.-L. Wang, T.-Y. Pan, S.-H. Hong, C.-M. Lan, H.-H. Kuo, C.-F. Lo, H.-Y. Hsu, C.-Y. Lin and E. W.-G. Diau, *Chem. Commun.*, 2011, **47**, 8910; (b) M. Ishida, D. Hwang, Y. B. Koo, J. Sung, D. Y. Kim, J. L. Sessler and D. Kim, *Chem. Commun.*, 2013, **49**, 9164; (c) C. Lo, S. Hsu, C. Wang, Y. Cheng, H. Lu, E. W. Diau and C. Lin, *J. Phys. Chem. C*, 2010, **114**, 12018; (d) J. Lu, X. Xu, K. Cao, J. Cui, Y. Zhang, Y. Shen, X. Shi, L. Liao, Y. Cheng and M. Wang, *J. Mater. Chem. A*, 2013, **1**, 10008; (e) R. B. Ambre, G.-F. Chang, M. R. Zanwar, C.-F. Yao, E. W.-G. Diau and C.-H. Hung, *Chem.–Asian J.*, 2013, **8**, 2144; (f) R. B. Ambre, G.-F. Chang and C.-H. Hung, *Chem. Commun.*, 2014, **50**, 725.
- 14 (a) C.-L. Wang, Y.-C. Chang, C.-M. Lan, C.-F. Lo, E. Wei-Guang Diau and C.-Y. Lin, *Energy Environ. Sci.*, 2011, **4**, 1788; (b) C.-H. Wu, T.-Y. Pan, S.-H. Hong, C.-L. Wang, H.-H. Kuo, Y.-Y. Chu, E. W.-G. Diau and C.-Y. Lin, *Chem. Commun.*, 2012, **48**, 4329; (c) H.-P. Lu, C.-L. Mai, C.-Y. Tsia, S.-J. Hsu, C.-P. Hsieh, C.-L. Chiu, C.-Y. Yeh and E. W.-G. Diau, *Phys. Chem. Chem. Phys.*, 2009, **11**, 10270; (d) C.-H. Wu, M.-C. Chen, P.-C. Su, H.-H. Kuo, C.-L. Wang, C.-Y. Lu, C.-H. Tsai, C.-C. Wu and C.-Y. Lin, *J. Mater. Chem. A*, 2014, **2**, 991; (e) J. M. Ball, N. K. S. Davis, J. D. Wilkinson, J. Kirkpatrick, J. Teuscher, R. Gunning, H. L. Anderson and H. J. Snaith, *RSC Adv.*, 2012, **2**, 6846; (f) C.-L. Mai, W.-K. Huang, H.-P. Lu, C.-W. Lee, C.-L. Chiu, Y.-R. Liang, E. W.-G. Diau and C.-Y. Yeh, *Chem. Commun.*, 2010, **46**, 809.
- 15 (a) K. Sunahara, M. J. Griffith, T. Uchiyama, P. Wagner, D. L. Officer, G. G. Wallace, A. J. Mozer and S. Mori, *ACS Appl. Mater. Interfaces*, 2013, **5**, 10824; (b) T. Hamamura, J. T. Dy, K. Tamaki, J. Nakazaki, S. Uchida, T. Kubo and H. Segawa, *Phys. Chem. Chem. Phys.*, 2014, **16**, 4551; (c) H.-P. Wu, Z.-W. Ou, T.-Y. Pan, C.-M. Lan, W.-K. Huang, H.-W. Lee, N. M. Reddy, C.-T. Chen, W.-S. Chao, C.-Y. Yeh and E. W.-G. Diau, *Energy Environ. Sci.*, 2012, **5**, 9843; (d) C.-L. Wang, J.-Y. Hu, C.-H. Wu, H.-H. Kuo, Y.-C. Chang, Z.-J. Lan, H.-P. Wu, E. Wei-Guang Diau and C.-Y. Lin, *Energy Environ. Sci.*, 2014, **7**, 1392; (e) J. Luo, M. Xu, R. Li, K.-W. Huang, C. Jiang, Q. Qi, W. Zeng, J. Zhang, C. Chi, P. Wang and J. Wu, *J. Am. Chem. Soc.*, 2014, **136**, 265.
- 16 (a) Y. Ooyama, S. Inoue, R. Asada, G. Ito, K. Kushimoto, K. Komaguchi, I. Imae and Y. Harima, *Eur. J. Org. Chem.*, 2010, **92**; (b) Y. Ooyama, S. Inoue, T. Nagano, K. Kushimoto, J. Ohshita, I. Imae, K. Komaguchi and Y. Harima, *Angew. Chem., Int. Ed.*, 2011, **50**, 7429; (c) Y. Ooyama, T. Nagano, S. Inoue, I. Imae, K. Komaguchi, J. Ohshita and Y. Harima, *Chem.–Eur. J.*, 2011, **17**, 14837; (d) Y. Ooyama, N. Yamaguchi, I. Imae, K. Komaguchi, J. Ohshita and Y. Harima, *Chem. Commun.*, 2013, **49**, 2548; (e) Y. Ooyama, Y. Hagiwara, T. Mizumo, Y. Harima and J. Ohshita, *New J. Chem.*, 2013, **37**, 2479; (f) Y. Ooyama, T. Sato, Y. Harima and J. Ohshita, *J. Mater. Chem. A*, 2014, **2**, 3293.
- 17 (a) Y. Harima, T. Fujita, Y. Kano, I. Imae, K. Komaguchi, Y. Ooyama and J. Ohshita, *J. Phys. Chem. C*, 2013, **117**, 16364; (b) N. Shibayama, H. Ozawa, M. Abe, Y. Ooyama and H. Arakawa, *Chem. Commun.*, 2014, **50**, 6398.



- 18 (a) J. Lu, X. Xu, Z. Li, k. Cao, J. Cui, Y. Zhang, Y. Shen, Y. Li, J. Zhu, S. Dai, W. Chjen, Y. Cheng and M. Wang, *Chem.–Asian J.*, 2013, **8**, 956; (b) T. Sakurada, Y. Arai and H. Segawa, *RSC Adv.*, 2014, **4**, 13201; (c) D. Daphnomili, G. Landrou, P. Singh, A. Thomas, K. Yesudas, B. K. G. D. Sharma and A. G. Goutsolelos, *RSC Adv.*, 2012, **2**, 12899; (d) D. Daphnomili, G. D. Sharma, S. Biswas, T. K. R. Justin and A. G. Goutsolelos, *J. Photochem. Photobiol., A*, 2013, **253**, 88; (e) C. Stangel, A. Bagaki, P. A. Angaridis, G. Charalambidis, G. D. Sharma and A. G. Coutsolelos, *Inorg. Chem.*, 2014, **53**, 11871.
- 19 (a) M.-D. Zhang, H.-X. Xie, X.-H. Ju, L. Qin, Q.-X. Yang, H.-G. Zheng and X.-F. Zhou, *Phys. Chem. Chem. Phys.*, 2013, **15**, 634; (b) L. Wang, X. yang, S. Li, M. Cheng and L. Sun, *RSC Adv.*, 2013, **3**, 13677; (c) J. Mao, D. Wang, S.-H. Liu, Y. Hang, Y. Xu, Q. Zhang, W. Wu, P.-T. Chou and J. Hua, *Asian J. Org. Chem.*, 2014, **3**, 153; (d) J. Massin, L. Ducasse, T. Toupance and C. Olivier, *J. Phys. Chem. C*, 2014, **118**, 10677; (e) L. Zhang, J. M. Cole and C. Dai, *ACS Appl. Mater. Interfaces*, 2014, **6**, 7535.
- 20 (a) H. Nobukuni, Y. Shimazaki, H. Uno, Y. Naruta, K. Okubo, T. Kojima, S. Fukuzumi, S. Seki, H. Sakai, T. Hasobe and F. Tani, *Chem.–Eur. J.*, 2010, **16**, 11611; (b) H. Nobukuni, T. Kamimura, H. Uno, Y. Shimazaki, Y. Naruta and F. Tani, *Bull. Chem. Soc. Jpn.*, 2012, **85**, 862; (c) K. Sakaguchi, T. Kamimura, H. Uno, S. Mori, S. Ozako, H. Nobukuni, M. Ishida and F. Tani, *J. Org. Chem.*, 2014, **79**, 2980; (d) T. Kamimura, K. Ohkubo, Y. Kawashima, H. Nobukuni, Y. Naruta, F. Tani and S. Fukuzumi, *Chem. Sci.*, 2013, **4**, 1451; (e) T. Kamimura, K. Ohkubo, Y. Kawashima, S. Ozako, K. Sakaguchi, S. Fukuzumi and F. Tani, *J. Phys. Chem. C*, 2015, **119**, 25634.
- 21 (a) T. J. Dines, L. D. MacGregor and C. H. Rochester, *Phys. Chem. Chem. Phys.*, 2001, **3**, 2676; (b) M. I. Zaki, M. A. Hasan, F. A. Al-Sagheer and L. Pasupulety, *Colloids Surf., A*, 2001, **190**, 261; (c) O. Kasende and T. Zeegers-Huyskens, *J. Phys. Chem.*, 1984, **88**, 2132; (d) H. Takahashi, K. Mamola and E. K. Plyler, *J. Mol. Spectrosc.*, 1966, **21**, 217; (e) M. A. Montañez, I. L. Tocón, J. C. Otero and J. I. Marcos, *J. Mol. Struct.*, 1999, **482–483**, 201.
- 22 T. Kamimura, M. Komura, H. Komiyama, T. Iyoda and F. Tani, *Chem. Commun.*, 2015, **51**, 1685.
- 23 C. Ruzié, L. Michaudet and B. Boitrel, *Tetrahedron Lett.*, 2002, **43**, 7423.

