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Crystal structure and proton conductivity of $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$: insights from neutron powder diffraction and solid-state NMR spectroscopy†

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The solid-state synthesis and structural characterisation of perovskite $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ ($x = 0.0, 0.1, 0.2, 0.3, 0.4$) and its corresponding hydrated ceramics are reported. Powder and neutron X-ray diffractions reveal the presence of cubic perovskites (space group $Pm\bar{3}m$) with an increasing cell parameter as a function of scandium concentration along with some indication of phase segregation. ^{119}Sn and ^{45}Sc solid-state NMR spectroscopy data highlight the existence of oxygen vacancies in the dry materials, and their filling upon hydrothermal treatment with D_2O . It also indicates that the Sn^{4+} and Sc^{3+} local distribution at the B-site of the perovskite is inhomogeneous and suggests that the oxygen vacancies are located in the scandium dopant coordination shell at low concentrations ($x \leq 0.2$) and in the tin coordination shell at high concentrations ($x \geq 0.3$). ^{17}O NMR spectra on ^{17}O enriched $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ materials show the existence of Sn–O–Sn, Sn–O–Sc and Sc–O–Sc bridging oxygen environments. A further room temperature neutron powder diffraction study on deuterated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ refines the deuteron position at the 24k crystallographic site ($x, y, 0$) with $x = 0.579(3)$ and $y = 0.217(3)$ which leads to an O–D bond distance of 0.96(1) Å and suggests tilting of the proton towards the next nearest oxygen. Proton conduction was found to dominate in wet argon below 700 °C with total conductivity values in the range 1.8×10^{-4} to 1.1×10^{-3} S cm^{-1} between 300 and 600 °C. Electron holes govern the conduction process in dry oxidizing conditions, whilst in wet oxygen they compete with protonic defects leading to a wide mixed conduction region in the 200 to 600 °C temperature region, and a suppression of the conductivity at higher temperature.

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1. Introduction

During the past three decades proton conducting ceramics have been widely studied due to their high ionic conductivities in the intermediate temperature region of 300–600 °C. In particular these materials have been proposed to be used as proton conducting electrolytes in protonic ceramic fuel cells (PCFC). The lower operating temperature of PCFCs would provide substantial advantages over solid oxide fuel cells (SOFC) based on oxide ion-conducting electrolytes. The higher operating temperatures of

SOFCs, usually 700–900 °C, have limited their technological development due to high system costs, performance degradation rates, slow start-up and shutdown cycles. For example, in the 700–900 °C temperature range, the use of chromium containing interconnector steels might cause chromium poisoning of the electrodes¹ and shorten the lifetime of the cell. Reduced start-up times and relaxed matching of the thermal expansion coefficients of the various fuel cell components are additional benefits that accompany the lowering of the operating temperature.

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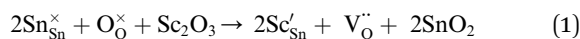
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† Electronic supplementary information (ESI) available: Rietveld fit of dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ sample (Fig. S1), ^{119}Sn (Fig. S2), ^{45}Sc (Fig. S3–S6) and ^{17}O (Fig. S7) spectra of all materials as a function of Sc doping concentration, ^{45}Sc MQMAS of deuterated $\text{BaSn}_{0.9}\text{Sc}_{0.1}\text{O}_{3-\delta}$ (Fig. S4), ^{45}Sc MQMAS of dry and deuterated $\text{BaSn}_{0.8}\text{Sc}_{0.2}\text{O}_{3-\delta}$ (Fig. S5), ^{45}Sc MQMAS of dry and deuterated $\text{BaSn}_{0.7}\text{Sc}_{0.3}\text{O}_{3-\delta}$ (Fig. S6), ^{17}O MQMAS of ^{17}O enriched $\text{BaSn}_{0.8}\text{Sc}_{0.2}\text{O}_{3-\delta}$ and $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ (Fig. S8). See DOI: 10.1039/c5ta09744d

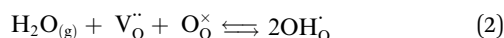
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Acceptor doped perovskites provide many eligible systems for proton conducting electrolytes, *e.g.* BaZr_{1-x}Y_xO_{3-δ},²⁻⁴ BaCe_{1-x}Y_xO_{3-δ},² SrCe_{1-x}Y_xO_{3-δ},⁵ BaZr_{1-x}Yb_xO_{3-δ} (ref. 6) all with 0 ≤ x ≤ 0.2, and BaZr_{1-x}In_xO_{3-δ} (ref. 7) with 0 ≤ x ≤ 1. Below 700 °C, BaZr_{1-x}Y_xO_{3-δ} possesses a bulk proton conductivity greater than the best oxide ion conductors.² Proton incorporation is reliant on the formation of oxygen vacancies in a process commonly referred to as acceptor doping in the A²⁺B⁴⁺O₃²⁻ type perovskites. Here, a portion of the tetravalent cations at the B-site is substituted by trivalent dopant cations resulting in the formation of charge compensating oxygen vacancies in the system. This process can be described using Kröger-Vink notation for Sc-doped BaSnO₃ with Sc³⁺ doped on the Sn⁴⁺ site as:

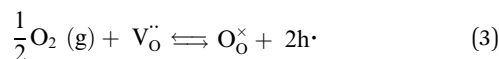


with Sc'_{Sn} corresponding to a Sc³⁺ ion sitting on a Sn site with a negative charge, and V_O^{··} to an oxygen vacancy with two positive charges. When in contact with a H₂O bearing gas the oxygen vacancies V_O^{··} are filled *via* the following reaction:

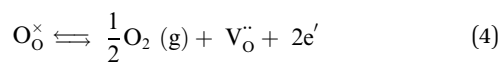


with OH_O[·] corresponding to a OH⁻ ion sitting on a O lattice site with a positive charge.

However, in oxidizing conditions and in some systems, electronic holes can instead compensate for the vacancies *via* the following equation leading to p-type hole (h[·]) conduction:



Under low oxygen partial pressures the following mechanism can occur yielding n-type electronic conduction:



Significant proton conduction has been reported in substituted stannate phases such as BaIn_{0.5}Sn_{0.5}O_{2.75},⁸ Ba₂YSnO_{5.5},⁹ BaSn_{1-x}M_xO_{3-δ} with M = Sc, Y, In and Gd, x = 0.125 (ref. 10) and x = 0.25 (ref. 11) and BaSn_{1-x}Y_xO_{3-δ} (0 ≤ x ≤ 0.5).¹² More recently, Li and Nino reported on proton conductivity of BaSn_{0.9}M_{0.1}O_{3-δ} (M = In, Lu, Er and Y) in oxidising and reducing conditions,¹³ whilst Bévilion *et al.*¹⁴ used a density functional theory approach to probe the energy landscape of the proton in substituted BaSn_{1-x}M_xO_{3-x/2}.

In this study, BaSn_{1-x}Sc_xO_{3-δ} was selected as the system of interest as the recent studies highlighted above have established acceptor doped BaSnO₃ as a promising alternative candidate to the more widely studied BaZrO₃ and BaCeO₃ systems. Scandium was chosen here as the dopant as it has an ionic radius that is only slightly larger than that of tin (0.745 Å and 0.69 Å for Sc³⁺ and Sn⁴⁺ respectively in 6-fold coordination).¹⁵ We report the preparation and characterisation of the BaSn_{1-x}Sc_xO_{3-δ} series with 0 ≤ x ≤ 0.4 *via* PXRD and solid-state NMR techniques, and a more detailed study of the highest acceptor doped sample, BaSn_{0.6}Sc_{0.4}O_{3-δ}. While the location of

the oxygen vacancies was determined by ¹¹⁹Sn, ⁴⁵Sc, and ¹⁷O multinuclear solid-state NMR spectroscopy by investigating the presence of Sn and Sc cations with various coordination numbers, the position of the deuterium ions in D₂O treated BaSn_{0.6}Sc_{0.4}O_{3-δ} was found by neutron powder diffraction (NPD). Finally, the electrical conductivity was studied using electrochemical impedance spectroscopy (EIS) recorded under different atmospheres to reveal the temperature dependence of the dominating charge carriers.

2. Experimental

2.1. Synthesis

BaSn_{1-x}Sc_xO_{3-δ} with x = 0, 0.1, 0.2, 0.3 and 0.4 were synthesized by a solid state reaction using stoichiometric amounts of BaCO₃ (Merck 99%), SnO₂ (Sigma-Aldrich 99.9%), and Sc₂O₃ (Sigma-Aldrich 99.9%). The reactants were weighed and finely mixed to a paste using a mortar and pestle and ethanol before heating at 1000 °C for 8 h. The powders were then ball milled to a fine powder for 8 h in a Teflon milling house with ethanol using a planetary ball mill and zirconium milling balls. The powders were then dried and pressed into pellets, and subsequently reacted at 1200 °C for 72 h before being ball milled, pelletized and heated again at 1455 °C for 24 h. The sintered pellets were thereafter milled into a fine powder to give the as-prepared samples. All the heating steps were performed under an oxygen gas flow.

Hydration of BaSn_{1-x}Sc_xO_{3-δ} was performed by heating the powders with a stoichiometric amount of D₂O, calculated to correspond to the complete filling of oxygen vacancies, in a hydrothermal bomb at 225 °C for 12 h. Drying of samples for NMR and NPD measurements was performed by treating at 900 °C for 8 h under vacuum. ¹⁷O NMR data were collected on samples that have been enriched in ¹⁷O by heating the freshly dried samples (1 h at 950 °C under vacuum) under 50% ¹⁷O enriched O₂ gas (Isotec, 99%) for 2 days at 950 °C.

Conductivity measurements on a sample of BaSn_{0.6}Sc_{0.4}O_{3-δ} were performed on a 16 mm diameter, 72% dense pellet (made by uni-axially pressing of powders at 8 tons) which was sintered at 1455 °C for 24 h. The pellet was then coated on both faces with platinum paste, heated for 2 h at 1000 °C to remove the organic component of the paste, and finally, treated for 7 days in a furnace at 300 °C with a vapour saturated N₂ gas flow (*p*(H₂O) ≈ 0.40 atm) to give a pre-hydrated sample.

2.2. X-ray powder diffraction (PXRD)

PXRD data for the as-prepared samples were collected on a Bruker AXS D8 ADVANCE VARIO X-ray powder diffractometer (CuK_{α1} = 1.54058 Å) equipped with a LynxEye detector and a germanium (111) primary monochromator. The step size used was 0.050° with a collection time of 0.7 s per step in the 27° to 72° 2-theta range.

2.3. Neutron powder diffraction (NPD)

NPD data were collected at room temperature on dried BaSnO₃, dried BaSn_{0.6}Sc_{0.4}O_{3-δ} and D₂O treated BaSn_{0.6}Sc_{0.4}O_{3-δ} samples



using the Polaris¹⁶ instrument at the ISIS neutron facility, and subsequently analysed using the GSAS^{17,18} software package. Data from two detector banks were used for the structure refinements, *i.e.* the backscattering detector bank covering scattering angles of $130^\circ < 2\theta < 160^\circ$, and a *d*-spacing range of $0.2 < d(\text{\AA}) < 3.2$, with a resolution of $\Delta d/d \sim 5 \times 10^{-3}$, and the 90° detector bank ($85^\circ < 2\theta < 95^\circ$; $0.3 < d(\text{\AA}) < 4.1$; $\Delta d/d \sim 7 \times 10^{-3}$). Data were collected for approximately 9 h for the D₂O treated BaSn_{0.6}Sc_{0.4}O_{3- δ} sample and 1 h for the dried samples.

Rietveld refinements^{18–20} included the following parameters: a scale factor, the cubic lattice parameter *a*, background parameters describing a reciprocal interplanar function, isotropic thermal vibration parameters for the cation sites, u_{Ba} , $u_{\text{Sn/Sc}}$, and anisotropic parameters for the oxygen site, u_{11} , $u_{22} = u_{33}$ and 4 profile parameters describing Gaussian and Lorentzian contributions to the Bragg peak profiles in the cubic space group *Pm* $\bar{3}$ *m*. Ba was set at 1b ($\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$), Sn/Sc at 1a (0, 0, 0) and O at 3d ($\frac{1}{2}, 0, 0$). Ahmed *et al.*²¹ reported the likelihood of the deuteron being located at the 24k (0.55, 0.20, 0) crystallographic site for BaZr_{0.5}In_{0.5}O_{2.5}(OD)_{0.5} and this was used as a starting point in the analysis of the data from the deuterium containing sample.

2.4. Solid-state NMR

¹¹⁹Sn NMR spectra were acquired at 11.7 T on a wide bore Oxford 500 MHz Varian Infinity Plus spectrometer using a 3.2 mm HX Chemagnetics probehead tuned to 186.26 MHz. The BaSnO₃ (dried) and BaSn_{1-x}Sc_xO_{3- δ} samples (in vacuum dried and D₂O treated forms) were packed under nitrogen gas atmosphere in 3.2 mm zirconia rotors, which were then spun at a spinning frequency $\nu_r = 20$ kHz. ¹¹⁹Sn single pulse experiments were carried out using a $\pi/2$ pulse width of 2 μ s (*i.e.* at an rf field amplitude of $\nu_1^{\text{Sn}} = 125$ kHz) and a recycle delay of 70 s allowing full relaxation of the ¹¹⁹Sn spins. Chemical shifts were externally referenced to SnO₂ at -604.3 ppm.

High field ⁴⁵Sc NMR experiments were performed at 19.6 T on an ultra-narrow bore Bruker DRX 830 MHz spectrometer at the National High Magnetic Field Laboratory, Tallahassee, Florida, USA using a home-built 1.8 mm single channel probe²² tuned to 202.44 MHz. All samples were packed inside 1.8 mm rotors, spun at a spinning frequency ν_r of 33.333 kHz, and short recycle delays of 0.2 s allowing full relaxation of the ⁴⁵Sc spins were used for the 1D spectra. *t*₁ rotor synchronized two-dimensional (2D) triple-quantum MAS (TQMAS) experiments^{23–25} were performed using a shifted-echo pulse sequence and the Soft-Pulse-Added-Mixing (SPAM) enhancement pulse.²⁶ Hard and soft pulses were performed at radio-frequency (rf) field amplitudes of $\nu_1^{\text{Sc}} = 150$ kHz and approximately $\nu_1^{\text{Sc}} = 20$ kHz, respectively. Chemical shifts were externally referenced to a 1 M solution of Sc(NO₃)₃ in water at 0.0 ppm.

¹⁷O NMR experiments were carried out on a 17.6 T wide bore Bruker Avance 750 MHz spectrometer equipped with a 4 mm HXY (in double resonance mode) probehead and operating at 101.72 MHz. All samples were packed inside 4 mm rotors and spun at a spinning frequency ν_r of 15 kHz. ¹⁷O one-dimensional spectra were recorded using a one pulse sequence with selective

pulse widths of $\pi/6 = 0.6 \mu$ s and at an rf field amplitude of $\nu_1^{\text{O}} = 50$ kHz. *t*₁ rotor synchronized two-dimensional (2D) TQMAS experiments were performed using the *z*-filtered pulse sequence.²⁷ Hard and soft pulses were performed at rf field amplitudes of $\nu_1^{\text{O}} = 50$ kHz and approximately $\nu_1^{\text{O}} = 10$ kHz, respectively. The recycle delays were set to 5 s for all experiments. Chemical shifts were externally referenced to water at 0.0 ppm.

All data were processed with MatLab and MatNMR.²⁸

2.5. Impedance spectroscopy

A ProboStatTM (NorECs AS, Norway) cell coupled to a Solartron 1260 frequency response analyser in standalone mode was used to collect electrochemical impedance data. Data collection was between 1 Hz and 1 MHz at 1 V rms amplitude between 150 and 1000 °C in steps of 50 °C with an equilibration time of 30 minutes before data collection. Data was collected for BaSn_{0.6}Sc_{0.4}O_{3- δ} in the following sequence: pre-hydrated sample heating and cooling in dry Ar gas, wet (humidified) Ar gas cooling, wet O₂ gas on cooling, and finally dry O₂ gas on cooling. Two silica tubes, one inside the other, were used to cover the cell, and two P₂O₅ gas traps before the cell were used to ensure dry gas conditions within the cell. A dense mullite–alumina tube was used in conjunction with a water bubbler at ambient temperature to provide wet gas ($p(\text{H}_2\text{O}) \approx 0.025$ atm) within the cell.

3. Results

3.1. X-ray diffraction

Fig. 1 shows the PXRD pattern for all dried BaSn_{1-x}Sc_xO_{3- δ} samples ($x = 0.0, 0.1, 0.2, 0.3$ and 0.4). These data reveal that all samples are highly crystalline and the patterns indicate that the phases adopt a cubic perovskite structure (space group *Pm* $\bar{3}$ *m*) across the range of compositions. The cell parameters, obtained from profile fitting using Jana2006,²⁹ increased with increase of dopant fraction in agreement with the Sc³⁺ ionic radius (0.745 Å) being larger than Sn⁴⁺ (0.69 Å).¹⁵ Close inspection of the data revealed evidence of peak shoulders at $2\theta \approx 44^\circ, 54^\circ$ for the $x = 0.1, 0.2$ and 0.3 samples. This behaviour was rationalised in terms of phase segregation into BaSnO₃ and a BaSn_{1-x}Sc_xO_{3- δ} phase comparatively rich in scandium in order to preserve the overall stoichiometry of the initial sample reactants. This behaviour was not apparent for the $x = 0.4$ sample, and its cell parameter of 4.1367(1) Å showed a significant enlargement compared to the value of 4.1156(1) Å determined for un-doped BaSnO₃. BaSn_{0.6}Sc_{0.4}O_{3- δ} , which showed the highest incorporation of scandium based on the PXRD results, was therefore selected for further study *via* neutron diffraction and impedance measurements.

3.2. Neutron diffraction

The NPD data for BaSn_{0.6}Sc_{0.4}O_{3- δ} presented in Fig. 2 revealed that a minor Sc₂O₃ impurity phase was present in the vacuum dried sample. The large neutron scattering lengths of scandium (12.29 fm) and oxygen (5.803 fm) compared to their relatively



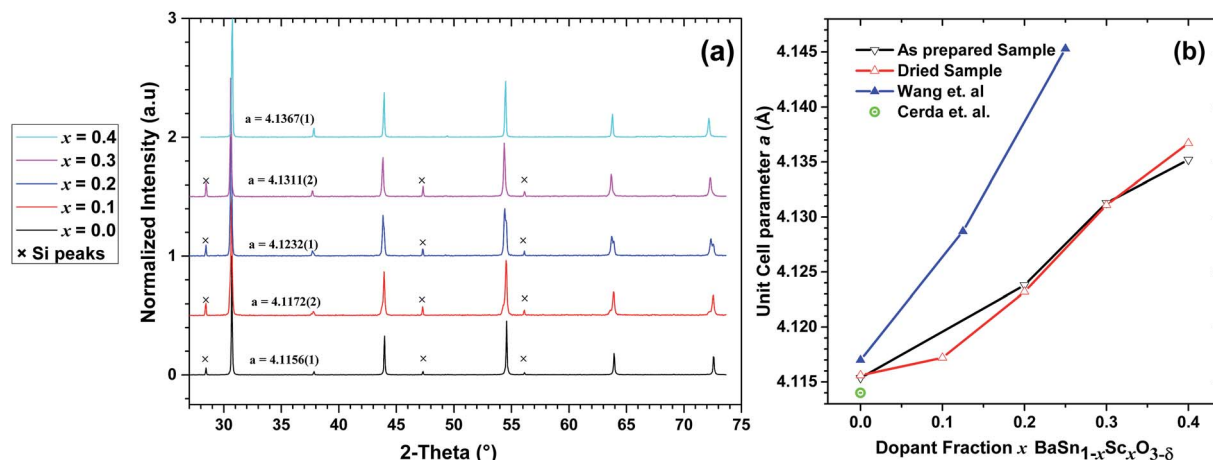


Fig. 1 (a) PXRD patterns collected on dried $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ samples with the indicated cell parameters in Ångströms. (b) Comparison of the cell parameters for the dried and as prepared $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ samples with values reported by Wang *et al.*^{10,11} and Cerda *et al.*³⁰

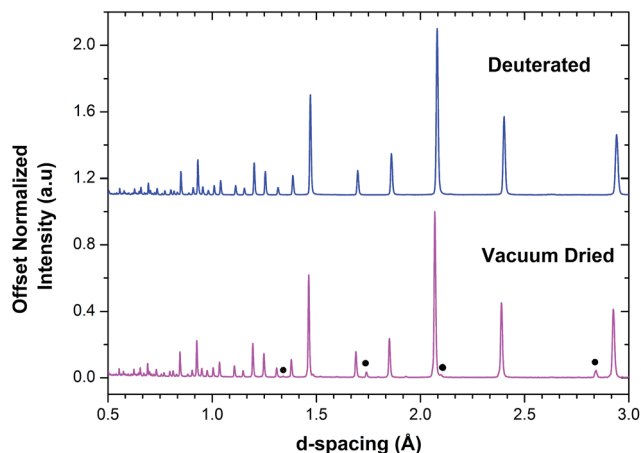


Fig. 2 NPD patterns of vacuum dried and deuterated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$. Black filled circles indicate reflections arising from a small amount of Sc_2O_3 .

weaker X-ray scattering powers could explain why this minor phase was detected in the neutron pattern (Fig. 2) but was not visible in the PXRD patterns (Fig. 1). The Sc_2O_3 peaks are not visible in the NPD of deuterated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ where only a single, deuterium containing, perovskite phase is present.

3.2.1. Vacuum dried $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$. As starting models three phases namely, $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{2.8}$ and BaSnO_3 , both modelled using the cubic $Pm\bar{3}m$ crystal system, and Sc_2O_3 were included into the Rietveld analysis of the dried sample, phase 1, 2 and 3 respectively. The weight fractions of these three phases obtained from the refinement were 95.55(1) wt%, 3.34(8) wt% and 1.11(1) wt%, respectively (Fig. S1†). The Sn : Sc site occupancy in the dominant perovskite phase 1 was reciprocally linked and refined to give a small increase in the Sn : Sc ratio, with 0.666(2) and 0.334(2) site occupancies for Sn and Sc, respectively. The overall sample stoichiometry was consistent with the initial 0.6 Sn and 0.4 Sc molar fractions. Modelling the oxygen atoms with an anisotropic displacement parameter

(ADP) significantly reduced the values of the χ^2 goodness of fit parameter from 15.66 to 10.86. Simultaneous refinement of the oxygen ADP and occupancy was not deemed reliable due to the high degree of correlation between these two variables. Hence the occupancy of the oxygen site in phase 1 was set to 0.944 as would be expected for a dried sample with a $\text{BaSn}_{0.666}\text{Sc}_{0.334}\text{O}_{3-\delta}$ composition with a final χ^2 value of 6.372. The final agreement to the data is shown in the supplementary data (Fig. S1†). Note that for simplicity we continue to refer to this sample as $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$, despite the slightly lower Sc content of the perovskite phase.

3.2.2. Hydrothermally D_2O treated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$. Given the absence of additional impurity reflections, the Sn^{4+} and Sc^{3+} occupancies used in the Rietveld fit to the NPD data were fixed to the nominal values of 0.6 and 0.4, respectively, for D_2O treated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$. The occupancy of the oxygen and barium sites was permitted to vary and both favoured a value slightly above unity and were therefore set to one. This is consistent with complete filling of oxygen vacancies by OD groups as per eqn (2) above during the hydrothermal treatment with D_2O . The fit improved significantly by allowing the oxygen ADP factor to vary anisotropically. The deuterium position was investigated by Rietveld analysis and the use of Fourier difference maps as described previously.²¹ Missing positive scattering was observed near fractional coordinates $x = 0.55$, $y = 0.2$ and $z = 0.0$, *i.e.* the crystallographic 24k site. The deuterium site occupancy was calculated from the number of filled oxygen vacancies with respect to the refined oxygen occupancy, *e.g.* $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{2.6}(\text{OD})_{0.4}$, which corresponds to a 24k site occupancy of ~ 0.017 . The isotropic ADP parameter was then set free to refine together with the atomic coordinates x and y of the deuterium at the 24k site. This resulted in a significant reduction in the standard uncertainties of the refined parameters and a small reduction in the goodness of fit parameters. The deuterium positional coordinates ($x, y, 0$) refined to (0.579(3), 0.217(3), 0). Results of the Rietveld analysis of the NPD data are listed in Table 1, and the final Rietveld fit achieved is shown in Fig. 3.



Table 1 Refined parameters from neutron powder diffraction of dried BaSnO₃ and both dried and deuterated BaSn_{0.6}Sc_{0.4}O_{3-δ}

Refinement parameters	BaSn _{0.6} Sc _{0.4} O _{3-δ}		
	Dry BaSnO ₃	Dry	Deuterated
<i>a</i> (Å)	4.11588(2)	4.13549(1)	4.15716(2)
<i>U</i> _{iso} (Å ²) × 100			
Ba	0.416(4)	0.614(5)	1.200(5)
Sn/Sc	0.231(3)	0.712(6)	1.057(4)
O <i>U</i> ₁₁	0.294(8)	0.334(11)	0.498(8)
O <i>U</i> _{22-<i>U</i>₃₃}	0.923(6)	1.201(7)	1.261(5)
D (<i>x</i> , <i>y</i> , 0)	—	—	0.579(3), 0.217(3), 0
D (<i>x</i> , <i>y</i> , 0) <i>U</i> _{iso}	—	—	11.3(4)
Occupancy			
Ba (0.5, 0.5, 0)	1.0 ^a	1.0 ^a	1.0 ^a
Sn (0, 0, 0)	1.0 ^a	0.666(2)	0.6
Sc (0, 0, 0)	—	0.334(2)	0.4
O (0.5, 0, 0)	1.0	0.944 ^b	1.0 ^a
D (<i>x</i> , <i>y</i> , 0)	—	—	0.0167 ^c
Bond distances (Å)			
12 × Ba–O	2.91037(1)	2.92423(1)	2.93956(1)
6 × Sn/Sc–O	2.05794(1)	2.06774(1)	2.07858(1)
1 × O–D	—	—	0.959(12)
1 × O–D ^(1-<i>y</i>, <i>x</i>, <i>z</i>)	—	—	2.109(7)
1 × O–D ^(-<i>y</i>, <i>x</i>, <i>z</i>)	—	—	2.680(6)
2 × O–D ^(1+<i>z</i>, <i>y</i>, -<i>x</i>)	—	—	2.862(9)
χ ²	48.16	6.37	12.84
<i>R</i> _{wp} (%)	0.0314	0.0283	0.0124
<i>R</i> _p (%)	0.0360	0.0441	0.0169
Phases (wt%)			
Main	100	95.55(1)	100
Sc ₂ O ₃	—	1.11(1)	—
BaSnO ₃	—	3.34(8) ^d	—
Variables	—	26	47

^a Occupancies refined to slightly larger than 1, and therefore fixed to 1.

^b Due to correlation between the oxygen site ADP and occupancy this value was not refined in final stages, instead it was fixed to the value determined by the amount of refined scandium $x = 0.334$ (oxygen fraction = $(3 - (0.334/2))/3$). ^c Occupancy of the deuterium site was fixed to reflect a deuterium content consistent with complete filling of the oxygen vacancies. ^d The refined unit cell parameter was 4.11457(17) Å.

3.3. Solid-state NMR

3.3.1. ¹¹⁹Sn NMR. The ¹¹⁹Sn magic angle spinning (MAS) NMR spectrum of dry BaSnO₃ (Fig. S2a in the ESI†) shows a sharp resonance at -679 ppm corresponding to tin in a six-fold symmetrical environment.³¹ This environment corresponds to Sn surrounded by 6 tin atoms in its first cation coordination shell, giving rise to a Sn(OSn)₆ local environment,³¹ the only chemical environment in the undoped BaSnO₃ material.

In dry Sc-substituted BaSnO₃, a new set of ¹¹⁹Sn resonances with intensity proportional to Sc concentration appears at around -640 ppm (Fig. 4a). This feature is assigned to tin in six fold environments surrounded by at least one scandium cation based on previous NMR studies of the related Y-doped BaSnO₃,

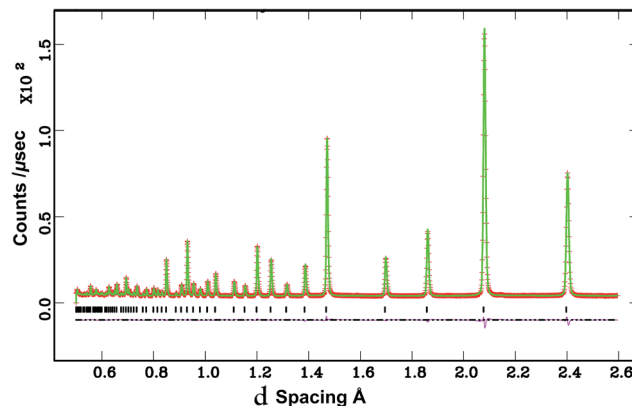


Fig. 3 Rietveld fit to the data for BaSn_{0.6}Sc_{0.4}O_{2.6}(OD)_{0.4} with the deuteron modelled at the 24k site. Both the observed (red crosses) and calculated (continuous green line) profiles are plotted and the position of the reflections is marked with vertical black bars. The difference curve lies at the bottom in purple.

materials.³² In this system, the six-coordinated Sn cations with various numbers of Y ions in their first cationic coordination shells, *i.e.* Sn(OSn)₅(OY), Sn(OSn)₄(OY)₂, Sn(OSn)₃(OY)₃, *etc.* could be individually observed, the ¹¹⁹Sn resonance shifting by +27 to +34 ppm per added Y ion. Here the ¹¹⁹Sn spectra of Sc-doped BaSnO₃ lacks such sharp, resolved features, most likely due to a smaller frequency shift per Sc ion added to the vicinity of the Sn nuclei. Indeed the ionic radius of Sc³⁺ in 6-fold coordination (0.745 Å) is much closer to the one of Sn⁴⁺ (0.69 Å) in comparison to Y³⁺ (0.90 Å) leading to smaller local distortions in the case of Sc substitution and therefore smaller frequency shifts. A weak but sharp BaSnO₃ resonance is seen in all four samples ($x = 0.1, 0.2, 0.3$ and 0.4) at -679 ppm most likely due to a separate BaSnO₃ impurity phase; the weakest BaSnO₃ resonance was seen for $x = 0.4$, consistent with the low phase fraction obtained in the NPD refinement of this phase (Fig. 2 and S1†).

A second broad resonance, centered at -450 ppm (Fig. 4a), is assigned to five-coordinated Sn environments in line with a shift to higher frequency going from six to five-fold Sn coordination,^{31,32} a trend generally observed for a number of nuclei.³³

On hydrothermal D₂O treatment of dry BaSn_{1-x}Sc_xO_{3-δ} (Fig. 4b), the ¹¹⁹Sn NMR resonance at -450 ppm, associated with the five-coordinated Sn environments, totally disappears, which is consistent with its assignment, and shows complete reaction of the oxygen vacancies \dot{V}_O with D₂O during hydration to form six-coordinated Sn environments (experimentally observed at -636 ppm). The sharp -679 ppm resonance is also seen, providing evidence for BaSnO₃ environments although no sign of an impurity phase was detected in the NPD data, presumably due to its low concentration.

3.3.2. ⁴⁵Sc NMR. Fig. 5 and S3 (in the ESI†) show the one-dimensional ⁴⁵Sc spectra of dry and deuterated BaSn_{1-x}Sc_xO_{3-δ} as a function of Sc doping level obtained at a high magnetic field of 19.6 T under fast MAS. The spectra of the dry samples contain a main resonance centred at around 110 ppm in addition to a much broader resonance in the 200–100 ppm region, which disappears upon hydrothermal D₂O treatment. Solid-state NMR



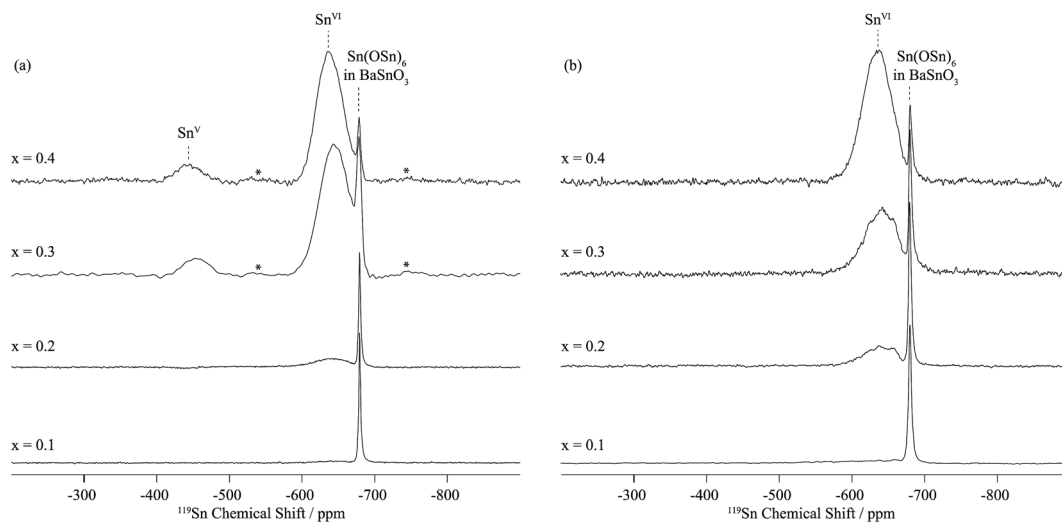


Fig. 4 ^{119}Sn MAS NMR single pulse spectra of (a) dry $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ and (b) deuterated $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ as a function of Sc doping level x were obtained at 11.7 T and under a MAS frequency of 20 kHz. Sn^{VI} and Sn^{V} denote six and five coordinated tin environments. Asterisks (*) indicate spinning side bands.

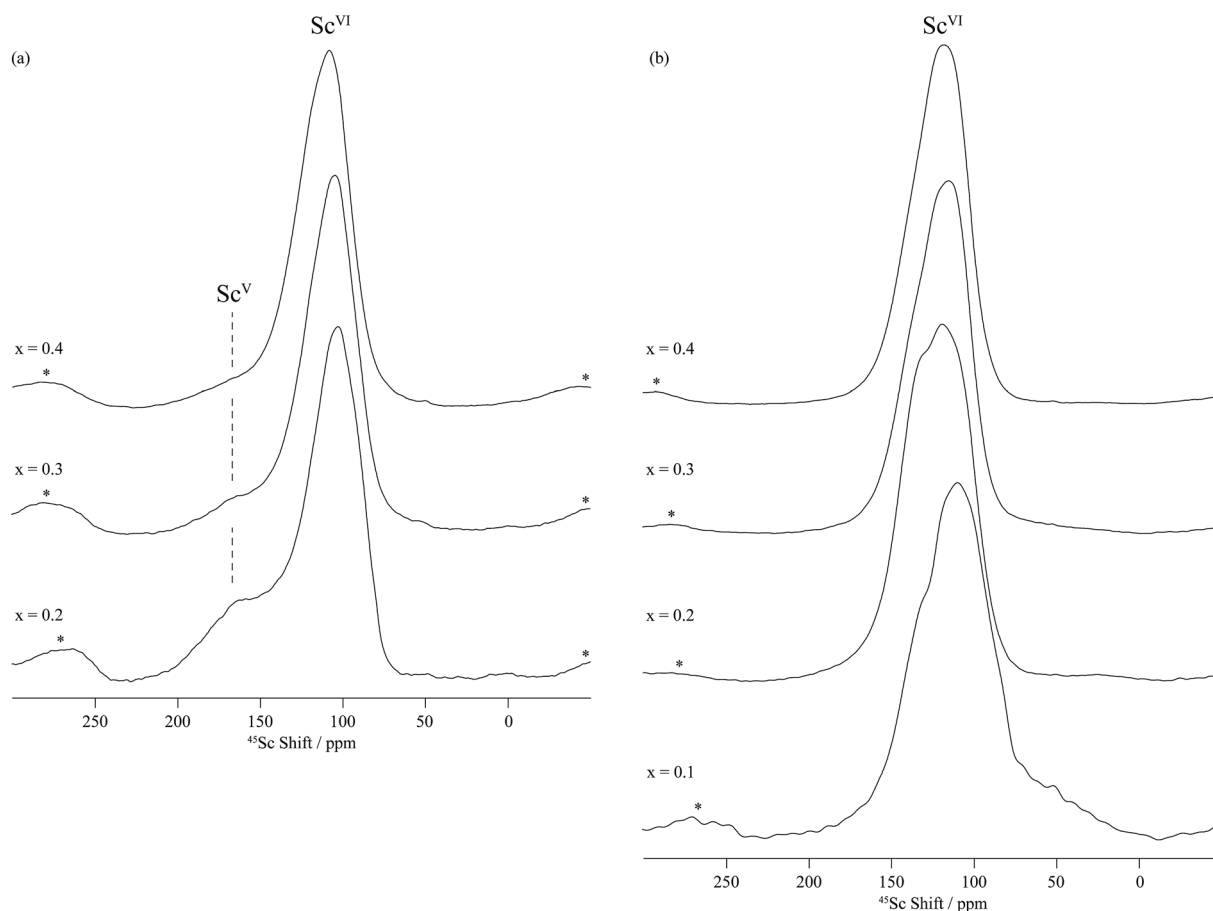


Fig. 5 Central transition ^{45}Sc MAS NMR spectra of (a) dry $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ and (b) deuterated $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ as a function of Sc doping level x . The spectra were obtained at 19.6 T and under MAS frequency of 33.33 kHz. The ^{45}Sc MAS NMR spectra of dry $\text{BaSn}_{0.9}\text{Sc}_{0.1}\text{O}_{3-\delta}$ was not recorded. Sc^{VI} and Sc^{V} denote six and five coordinated scandium environments. Asterisks (*) indicate spinning side bands.



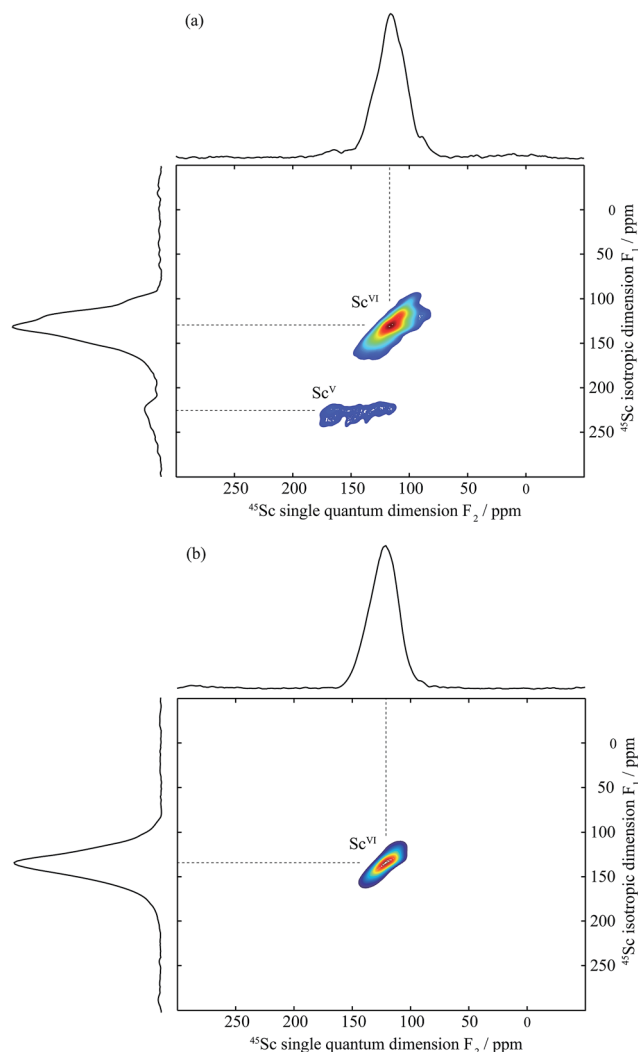


Fig. 6 Two-dimensional sheared triple-quantum ^{45}Sc MAS spectra of (a) dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ and (b) deuterated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ obtained at 19.6 T with a MAS frequency of 33.33 kHz. 9600 transients were accumulated for each of the 24 (for (a)) and 16 (for (b)) t_1 increments at a recycle delay of 0.4 s. Top: anisotropic skyline projection (the ^{45}Sc MAS NMR single pulse spectra are given in Fig. 5 and S3†). Left: isotropic skyline projection of the TQMAS spectra. Sc^{VI} and Sc^{V} denote six and five coordinated scandium environments, respectively.

spectra of quadrupolar nuclei such as ^{45}Sc (spin = 7/2) are often broad even under MAS because of residual second-order quadrupolar interactions. This can be removed by performing a two-dimensional triple-quantum MAS experiment (TQMAS)^{23–25} whose vertical projection along the F_1 dimension yields a one-dimensional isotropic spectrum free of second-order broadening. Such experiments have been recorded for the dry and deuterated $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ materials (Fig. 6, S4–S6†) and are discussed below.

Two sets of resonances are clearly observed in the F_1 vertically projected spectra of the dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ sample at shifts of approximately 125 and 225 ppm (Fig. 6) demonstrating the presence of two different scandium environments. Extraction of the shifts of these resonances in the horizontal F_2 dimension

Table 2 Experimental ^{119}Sn , ^{45}Sc and ^{17}O NMR parameters dry BaSnO_3 , dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ and deuterated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ ^a

Site	Environment	$\delta_{\text{iso}}/\text{ppm}$	C_Q/MHz	η_Q
Dry BaSnO_3				
Sn	$\text{Sn}^{\text{VI}}(\text{OSn})_6$	–679	— ^b	
Sc	— ^b			
O ^c	$\text{Sn}^{\text{VI}}\text{–O–Sn}^{\text{VI}}$	152	6.1	0.0
Dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$				
Sn	$\text{Sn}^{\text{VI}}(\text{OSn})_6$ in BaSnO_3	–679(1) ³¹	— ^b	
	Sn^{VI}	–640(10)		
	$\text{Sn}^{\text{V}}(\text{V}\ddot{\text{O}})$	–450(10)		
Sc	Sc^{VI}	120(5)	7(2)	0.7(3)
	Sc^{V}	199(7)	20(2)	0.0(1)
O	Sn–O–Sn	203(5)	8(1)	— ^d
	Sn–O–Sc	248(10)	3(1)	— ^d
	Sn–O–Sc	261(8)	3(1)	— ^d
	Sc–O–Sc	420	5(1)	— ^d
Hydrothermally D_2O treated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$				
Sn	$\text{Sn}^{\text{VI}}(\text{OSn})_6$ in BaSnO_3	–679(1) ³¹	— ^b	
	Sn^{VI}	–636(10)		
Sc	Sc^{VI}	122(5)	7(2)	0.8(1)
O	— ^d			

^a ^{45}Sc and ^{17}O NMR parameters were determined from the peak positions in the TQMAS spectra at 19.6 and 17.6 T, respectively (see ESI).²⁵ ^b Not applicable. ^c Data from ref. 32. ^d Not measured experimentally.

allow isotropic chemical shifts values of around 120 and 200 ppm to be extracted (see Table 2) and assigned to 6- and 5-coordinated scandium environments, respectively, based on previous study by Stebbins *et al.*,³⁴ our previous work on the $\text{BaZr}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ series,³⁵ and Takamura *et al.*'s recent hydration study of 10% mol Sc-substituted BaZrO_3 .³⁶ The 5-coordinated Sc environment which has a very large linewidth (leading to a quadrupolar coupling of around 20 MHz) is ascribed to the presence of an oxygen vacancy in the 1st coordination shell of a Sc atom.

Upon hydrothermal D_2O treatment, the oxygen vacancies are filled by protonic (deuterons) and OD defects resulting in the loss of the 5-coordinated Sc as revealed by the ^{45}Sc NMR spectra given in Fig. 5b and S4–S6b.† These spectra are now dominated by resonances centered at around 125 ppm and corresponding to 6-coordinated Sc only (Table 2). In fact, more than one 6-coordinated Sc environments are often visible in the ^{45}Sc MQMAS spectra of deuterated $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$, and are assigned to ScO_6 (as in the dry samples) and $\text{ScO}_5(\text{OD})$ environments (*i.e.* 6-coordinated scandium in the vicinity of a protonic defect).

Note that the Sc_2O_3 impurity seen by NPD was not observed by ^{45}Sc NMR of dry $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ (Fig. 6a). This is attributed to the very small amount of Sc_2O_3 (1.1%, Table 1), which is probably below the NMR detection limit, and to the fact that the two 6-coordinated Sc sites in Sc_2O_3 have isotropic chemical shifts of 108 and 128 ppm,³⁴ very close to the value for 6-coordinated Sc in this sample. The lack of resolution might therefore also prevent its observation.

3.3.3. ^{17}O NMR. ^{17}O NMR is usually very challenging due to the very low natural abundance of ^{17}O (approximately 0.037%),



often requiring isotopic enrichment. This is routinely performed in these materials *via* a gas–solid exchange reaction with ^{17}O enriched O_2 gas^{37,38} at elevated temperatures (see the Experimental section for further details). All the one-dimensional ^{17}O MAS NMR spectra of ^{17}O enriched $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ (Fig. 7 and S7†) show three sets of resonances, at around

150 ppm, in the 200–300 ppm region, and centred at around 420 ppm (Table 2). The resonance around 150 ppm region consists of a single oxygen environment (see the ^{17}O MQMAS spectra, Fig. S8†) and is assigned to a bridging oxygen bound to two tin cations, *i.e.*, Sn–O–Sn, based on our previous work on $\text{BaSn}_{1-x}\text{Y}_x\text{O}_{3-\delta}$.³² As the Sc content is increased, this resonance

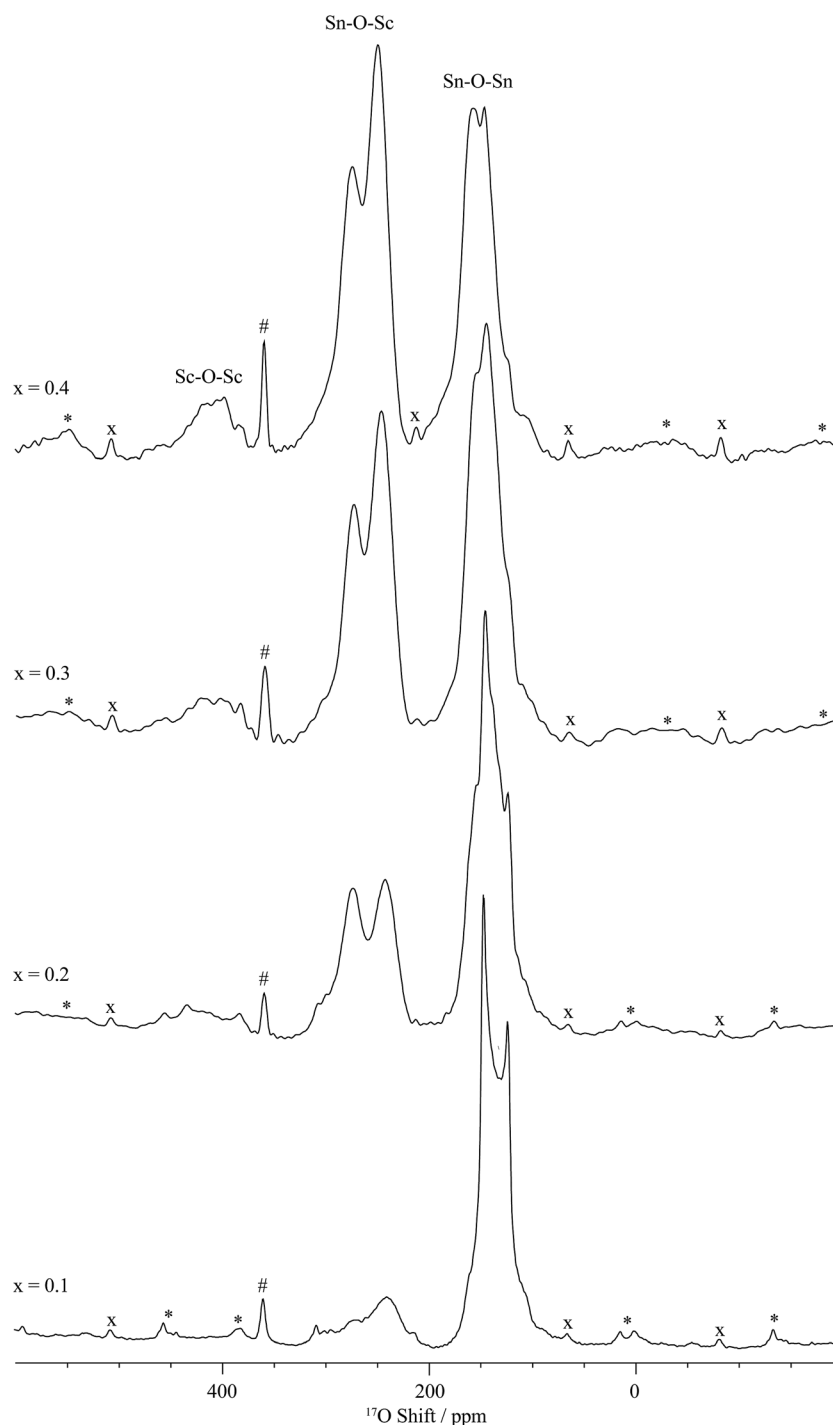


Fig. 7 Central transition ^{17}O MAS NMR single pulse spectra of ^{17}O enriched $\text{BaSn}_{1-x}\text{Sc}_x\text{O}_{3-\delta}$ as a function of the Sc doping level x . The spectra were obtained at 17.6 T and under MAS frequency of 15 kHz. The asterisks (*), dash (#) and crosses (x) denote spinning side bands, the ^{17}O signal of the ZrO_2 rotor and its spinning side bands, respectively.



remains strongly present (see Fig. 7 for $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$) indicating the lack of significant Sn/Sc ordering in this material. Two peaks are clearly observed in the 200–300 ppm region, their intensities increasing with Sc concentration (relative to those for Sn–O–Sn); they are both tentatively assigned to Sc–O–Sn environments, the presence of two resonances possibly arising from Sc and Sn being five or six coordinated. A third broad resonance centred at around 420 ppm is also observed and is assigned to Sc–O–Sc oxygen environments based on previous work on the related Sc doped BaZrO_3 cubic perovskite.³⁵

3.4. Impedance spectroscopy

Fig. 8 shows the complex plane plot of the pre-hydrated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ sample on heating at 100 °C in dry Ar. Two time constants are observed, including one in the high frequency region near the origin; the data is modelled using two (RQ) elements, representing a resistor and constant phase

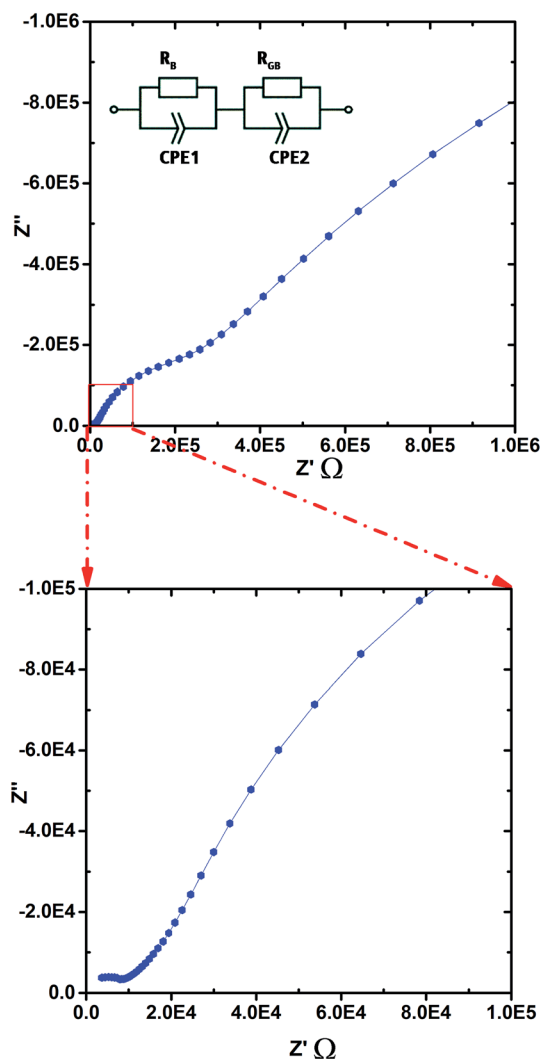


Fig. 8 Typical Nyquist plot of the impedance showing the full 1 Hz to 1 MHz range for pre-hydrated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ on heating in a dry Ar atmosphere at 100 °C, the blue lines representing visual guides. A magnified view at low Z' is shown in the insert.

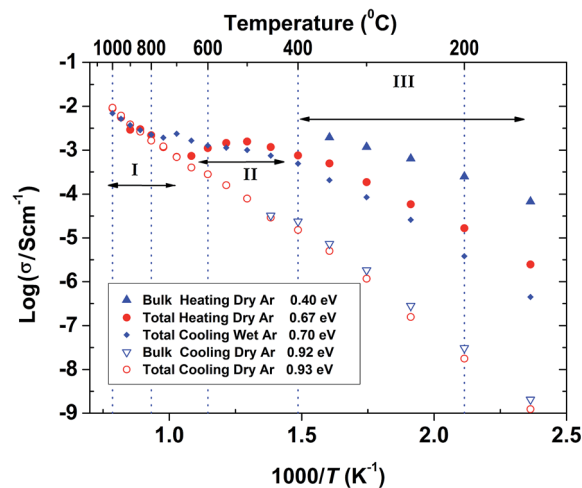


Fig. 9 Arrhenius plot of the conductivity of an initially hydrated $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ sample under dry or wet Ar gas atmospheres. Three conductivity regimes (I, II and III) are observed, as highlighted, when the sample contains protons. The activation energies in the temperature regime III (below 400 °C) are indicated in the inserts.

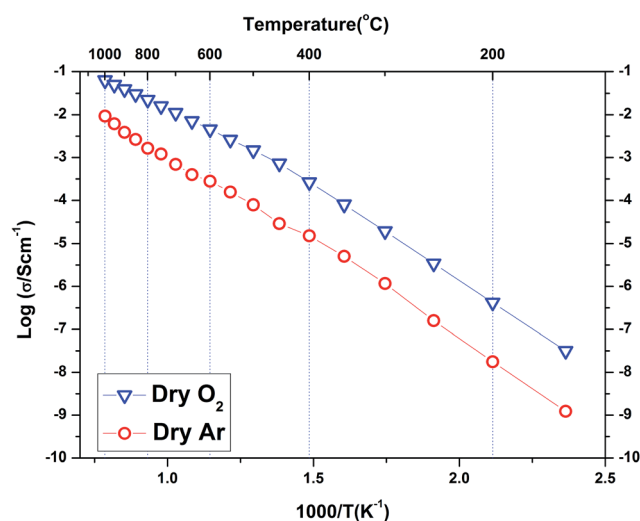


Fig. 10 A comparison of the total conductivity of $\text{BaSn}_{0.6}\text{Sc}_{0.4}\text{O}_{3-\delta}$ under dry Ar gas and dry O_2 gas. The activation energies below 400 °C were 0.95 eV and 0.90 eV respectively.

element in parallel, connected in series. The derived capacitances were $1.17 \times 10^{-11} \text{ F cm}^{-2}$ and $8.14 \times 10^{-9} \text{ F cm}^{-2}$ consistent with bulk and grain boundary processes, respectively. The feature at the lowest frequencies is attributed to electrode processes. For the initial heating and cooling data, it was possible to separate bulk and grain boundary conductivity in this manner at temperatures below approximately 400 °C. At higher temperatures, and for the other atmospheres and thermal protocols, the data were analysed using a similar approach but here the distinction between bulk and grain boundary was not as clear and only the total conductivity (bulk + grain boundary) could be extracted.



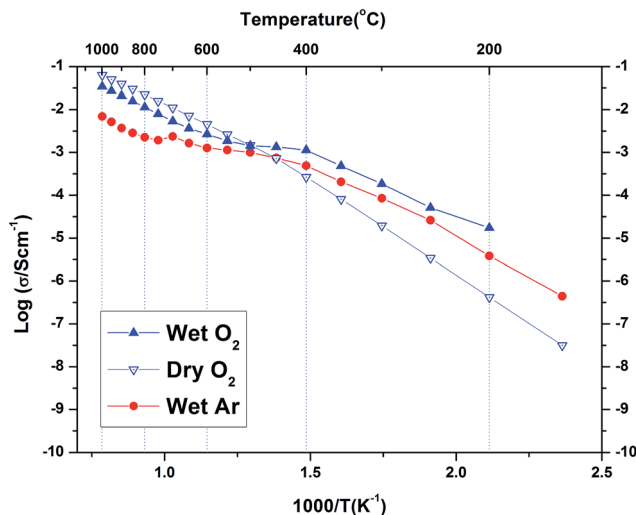


Fig. 11 Total conductivity under dry O₂, wet O₂ and wet Ar gas conditions. The activation energy for the conduction processes were 0.61 eV and 0.64 eV in dry and wet O₂ respectively in the region above 600 °C while below 400 °C they were 0.90 eV and 0.63 eV respectively.

The conductivity data collected for BaSn_{0.6}Sc_{0.4}O_{3-δ} under Ar consisted of three regions (see Fig. 9). Region I between 800–1000 °C has O²⁻ anions or possibly electron holes as the dominant charge carriers, while region II, between 400 and 800 °C is characteristic of the growing influence of protons and displays a characteristic plateau³⁹ that reflects the simultaneously varying proton concentration and proton mobility. Region III, at $T \leq 400$ °C, is dominated by proton charge carriers. Comparison of the conductivity under dry Ar vs. dry O₂ conditions reveals that the sample possesses significantly higher conductivity, approximately one order of magnitude greater, under oxidizing conditions throughout the entire temperature interval as evident in Fig. 10. Conductivity under wet oxygen and above 450 °C was found, unexpectedly, to be lower than that in dry oxygen (Fig. 11).

4. Discussion

4.1. Phase formation

The cell parameters extracted from the PXRD results of the as-synthesized BaSn_{1-x}Sc_xO_{3-δ} ($x = 0.0, 0.1, 0.2, 0.3, 0.4$) samples were found to increase with the level of Sc³⁺ dopant content. This expansion is expected as it reflects the greater ionic radius of Sc³⁺ (0.745 Å, 6-fold coordination) in comparison to Sn⁴⁺ (0.69 Å).¹⁵ Additionally, samples with dopant concentrations in the range $x = 0.1-0.3$, contained peak shoulders in their XRD patterns, which were attributed to a level of phase segregation. These results were in agreement with the ¹¹⁹Sn NMR data which revealed the presence of sharp Sn(OSn)₆ environments similar to that observed in BaSnO₃ (Fig. S2†).

The cell parameters obtained in this paper for BaSn_{1-x}Sc_xO_{3-δ} are visibly lower than those reported by Wang and co-workers,^{10,11} (Fig. 1b) but for BaSnO₃ our reported value is still higher than the value of 4.1140 Å reported by both Roth *et al.*⁴⁰

and Cerda *et al.*³⁰ Anomalous behaviour has been reported for Sc³⁺ doped BaZrO₃ (ref. 41) system with samples sintered at a lower temperature having a larger unit cell parameter when compared to samples sintered at higher temperatures. Hiraiwa *et al.*⁴¹ demonstrate this behaviour is unique to the Sc³⁺ dopant and is contrary to the behaviour of other dopants for BaZrO₃. However no hypothesis exists yet to explain this behaviour.

Although BaSn_{0.6}Sc_{0.4}O_{3-δ} was initially found to be phase pure by PXRD, subsequent NPD data revealed the presence of some Sc₂O₃, and hence the possible presence of BaSnO₃ in the dried BaSn_{0.6}Sc_{0.4}O_{3-δ} sample. Indeed, very weak intensities, seen as shoulders on the main perovskite peaks, were visible in the NPD pattern (Fig. S1†) and the refined cell parameter obtained for the minor BaSnO₃ component of 4.11457(17) Å showed good agreement with that of the dried BaSnO₃ sample ($a = 4.11588(2)$ Å). This was confirmed by the presence of the typical Sn(OSn)₆ resonance at -679 ppm in the ¹¹⁹Sn NMR data.

On hydrothermal treatment of BaSn_{0.6}Sc_{0.4}O_{3-δ} with D₂O, the minor Sc₂O₃ impurity was not observed in the NPD data, possibly indicating the complete solubility of the Sc₂O₃ into the perovskite structure is obtained under these conditions; a small signal characteristic of the local Sn environment present in BaSnO₃ was still observed by ¹¹⁹Sn NMR (Fig. S2†). The significant difference in melting points of SnO₂ (1630 °C) and Sc₂O₃ (2485 °C) reactants suggests that different cation diffusion rates are likely to be a contributing factor for the observed sample inhomogeneity across the BaSn_{1-x}Sc_xO_{3-δ} series. The solution based synthesis approaches utilised by Wang *et al.*¹⁰ and Buannic *et al.*³² in the preparation of BaSn_{1-x}Y_xO_{3-δ} ($0.0 \leq x \leq 0.5$) may be expected to help overcome this issue, although we note that the presence of undoped BaSnO₃ was also reported for BaSn_{0.9}Y_{0.1}O_{3-δ}.³²

4.2. Deuteron site

The deuteron site of the hydrothermally D₂O treated BaSn_{0.6}Sc_{0.4}O_{3-δ} sample was successfully refined by Rietveld analysis. The deuteron atomic coordinates at the 24k site were $x = 0.579(3)$ and $y = 0.217(3)$ resulting in an average O–D bond distance of

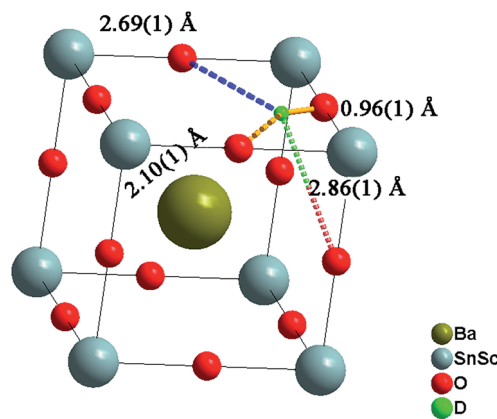


Fig. 12 Illustration of the BaSn_{0.6}Sc_{0.4}O_{2.6}(OD)_{0.4} unit cell showing the location of the deuteron site and its associated bond lengths to the nearest oxygen ions. The spheres represent in decreasing size Ba > Sn/Sc > O > D.



0.96(1) Å, in good agreement with literature values.^{21,42} The local environment around a deuteron occupying the 24k site is illustrated in Fig. 12 and shows three O–D interatomic distances of relevance for the proton transfer step towards acceptor oxygen ions, the closest being at 2.11(1) Å. It is clear that the local proton configuration is highly anisotropic and these results are in good agreement with experimental studies of related, highly substituted, perovskites, BaZr_{0.5}In_{0.5}O_{2.75} (ref. 21) and BaSn_{0.5}In_{0.5}O_{2.75}.⁴³ The presence of a similar next nearest O–D interaction at ~2.15 Å was also recently found for BaTi_{0.5}In_{0.5}O_{2.53}(OD)_{0.44} from reverse Monte Carlo analysis of total scattering neutron diffraction data.⁴⁴ These findings are in line with theoretical^{45–47} and experimental^{48,49} studies that revealed a clear tendency for protons to relax towards the dopant ions. The proximity of the second nearest oxygen atom indicates a tendency for enhanced hydrogen bonding interactions that will influence the proton diffusion. Whilst hydrogen bonding is expected to increase the likelihood of success for a proton transfer between neighbouring oxygen ions, the re-orientation step necessary for long range diffusion, involves breaking of the same H-bonds. It is therefore presently unclear what the full implications of the deuteron site are for migrating protons.

Both the D₂O treated and the vacuum dried samples showed large and highly anisotropic ADPs for the oxygen site. This indicates significant static disorder of the oxygen ions as previously found²¹ and it is important to stress that the refined structural models will represent a long range, time averaged, picture.

The diffraction data for the deuterated BaSn_{0.6}Sc_{0.4}O_{2.8} sample revealed an increase in the cell parameter, *a*, compared to the as-prepared sample (Table 1). This lattice expansion is due to the filling of oxygen vacancies by larger hydroxyl (OD) groups and was similar in magnitude to that reported in related perovskites.^{39,50} The hydration process was also clearly reflected in the solid-state NMR data. The presence of a five-coordinate Sn and Sc peaks in the spectra of the dry BaSn_{0.6}Sc_{0.4}O_{2.8} sample confirmed the existence of oxygen vacancies (*V*_O) and the loss of these signals upon hydrothermal D₂O treatment confirmed their subsequent filling.

4.3. Local Sn and Sc environments

The fine structure observed for the contribution of each Y dopant cation in Sn(OSn)_{6–z}(OY)_z (0 ≤ *z* ≤ 0.4) environment in the ¹¹⁹Sn NMR spectra of the related BaSn_{1–x}Y_xO_{3–δ} phases^{31,32} is not observed here for BaSn_{0.6}Sc_{0.4}O_{3–δ} (Fig. 4 and S2†). Additionally, a signal associated with Sn(OSn)₆ environments persists for the highly substituted Sc sample whereas it is absent for BaSn_{0.6}Y_{0.4}O_{2.8} and BaSn_{0.5}Y_{0.5}O_{2.75}.³² This suggests that there is greater disorder at the B-site and/or smaller changes in the specific local environments due to the much closer ionic radii of 6-fold coordinated Sn⁴⁺ (0.69 Å) and Sc³⁺ (0.745 Å) vs. 6-fold Y³⁺ (0.96 Å).¹⁵ More importantly, the presence of noticeable concentrations of Sn–O–Sn and Sc–O–Sc environments, as revealed by ¹⁷O NMR on ¹⁷O enriched BaSn_{0.6}Sc_{0.4}O_{3–δ}, confirms the absence of significant Sn/Sc ordering; for strict ordering and *x* = 0.5, only the Sc–O–Sn environment should be present, analogous to the behaviour of BaSn_{0.5}Y_{0.5}O_{2.75} (Ba₂SnYO_{5.5}).³² For *x* < 0.5 this environment

dominates, with lower concentrations of Sn–O–Sn environments being present, their concentration increasing with decreasing Sc content. For *x* = 0.1 and 0.2, the ratio of Sn(OSn)_{6–n}(Osc)_n with *n* > 0 to Sn(OSn)₆ sites is low but increases dramatically for *x* = 0.3 and 0.4, the amount of segregated BaSnO₃ being minimum for the latter. There is a possibility that Sn, if hosting Sc in its vicinity, has a preference for hosting a high number of Sc, *i.e.* Sn(OSn)_{6–n}(Osc)_n with *n* ≥ 3. In such case, the concentration of Sn(OSn)_{6–n}(Osc)_n with *n* ≥ 3 would be small for *x* = 0.1 and 0.2 and, combined to a broad resonance (as seen for *x* = 0.3 and 0.4), would yield to a very weak signal. This hypothesis is corroborated by the recurring segregation of a non-negligible amount of BaSnO₃, by the presence of a fair number of Sn–O–Sc bonds, and by the limited amount of Sc–O–Sc linkages as observed by ¹⁷O (Fig. 7) for *x* = 0.1 and 0.2. The greater size difference between Y³⁺ and Sn⁴⁺ drives a stronger tendency for ordering of the B-site cations that becomes nearly perfect with alternating Sn–O–Y–O–Sn linkages in BaSn_{0.5}Y_{0.5}O_{2.75} (Ba₂SnYO_{5.5}) as demonstrated by the existence a single main resonance at 259 ppm for Sn–O–Y moieties in the ¹⁷O NMR spectra.³² While some preferential cationic arrangement is possibly occurring in BaSn_{1–x}Sc_xO_{3–δ}, it is not as predominant as in BaSn_{1–x}Y_xO_{3–δ}.

The one-dimensional ⁴⁵Sc NMR spectra of dry BaSn_{0.8}Sc_{0.2}O_{3–δ} (Fig. 5) reveals a clear signal for five coordinated Sc, whilst the ¹¹⁹Sn data show no signal of 5-coordinated Sn for the *x* = 0.1 and 0.2 samples. Taken together, this strongly implies that oxygen vacancies are preferentially found in between or near Sc cations at low doping levels (*x* ≤ 0.2). These findings agree with the results of Buannic *et al.*,³⁵ Oikawa *et al.*^{36,51} (for *x* ≤ 0.1) on the related BaZr_{1–x}Sc_xO_{3–δ} system that suggested a tendency for the amount of 5 coordinated Sc to increase with doping level for *x* ≤ 0.2. Interestingly, the one-dimensional ⁴⁵Sc NMR spectra reveal that the ratio of Sc^V to Sc^{VI} decreases as the Sc content increases above *x* = 0.2. A notable difference between the BaSn_{1–x}Sc_xO_{3–δ} and BaZr_{1–x}Sc_xO_{3–δ} systems is the level of dopant solubility which reaches a maximum for BaZr_{1–x}Sc_xO_{3–δ} at *x* ≈ 0.2,³⁵ whereas the scandium incorporation level reaches *x* ≈ 0.35 for the nominal *x* = 0.4 BaSn_{1–x}Sc_xO_{3–δ} sample based on the Sc site occupancy refined from neutron diffraction analysis for the dry material (Table 1). Buannic *et al.*,³⁵ speculated earlier that the avoidance of energetically unfavourable Sc–O–Sc linkages, that are expected to become more numerous in systems where B-site ordering does not occur, may be a driving force for phase segregation into Sc₂O₃ and BaZr_{1–x}Sc_xO_{3–δ} with lower *x*. The present NMR findings for BaSn_{0.6}Sc_{0.4}O_{3–δ}, showing coexistence of Sn(OSn)₆ and SnO₅ coordinations and low levels of ScO₅ (in comparison to the *x* = 0.2 sample), therefore points towards a relative abundance of Sn – *V*_O – Sn local environments in comparison with a purely statistical cation and vacancy distribution. Possibly the Sn⁴⁺ ion is more flexible with respect to oxygen vacancies than Zr⁴⁺, and this plays a role in facilitating a relatively higher scandium incorporation into the perovskite matrix.

In summary, consideration of all the NMR data reveals an intricate picture in relation to the local B-site environments. The picture that emerges is nonetheless consistent with the formation of increasing levels of oxygen vacancies upon acceptor doping and



Table 3 Comparison of conductivity parameters with literature values of acceptor doped, BaSnO₃ and BaTiO₃ perovskite systems

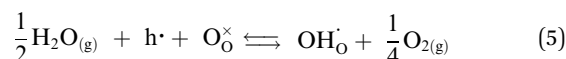
Material	E_a protonic (total)/ eV	E_a protonic (bulk)/ eV	Total conductivity (protonic) in wet gas/S cm ⁻¹	Ref.
BaSn _{0.9} In _{0.1} O _{3-δ}	—	0.54	1.3×10^{-3} (4% H ₂ /96% N ₂ 500 °C)	11
BaSn _{0.875} Sc _{0.125} O _{3-δ}	0.87	0.52	4×10^{-4} (Ar, 500 °C)	10
BaSn _{0.75} Sc _{0.25} O _{3-δ}	0.73	0.52	8×10^{-4} (Ar, 500 °C)	56
BaSn _{0.6} Sc _{0.4} O _{3-δ}	0.7–0.67	0.40	1.07×10^{-3} (Ar, 600 °C)	This work
BaSn _{0.9} Y _{0.1} O _{3-δ}	—	0.49	4×10^{-5} (4% H ₂ /96% N ₂ 500 °C)	13
BaSn _{0.875} Y _{0.125} O _{3-δ}	0.72	0.38	2.3×10^{-4} (Ar, 500 °C)	10
BaSn _{0.75} Y _{0.25} O _{3-δ}	0.66	0.34	7×10^{-4} (Ar, 500 °C)	12
BaSn _{0.5} Y _{0.5} O _{2.75} (Ba ₂ YSnO _{5.5} double perovskite structure)	—	0.51	1.3×10^{-3} (bulk) (N ₂ , 600–400 °C)	9
BaTi _{0.8} Sc _{0.2} O _{3-δ} (6 H hexagonal structure)	0.80	0.77	1×10^{-5} (Ar, 400 °C)	58
BaTi _{0.5} Sc _{0.5} O _{3-δ}	0.46	0.22	2.89×10^{-4} (Ar, 550 °C)	52
BaTi _{0.3} Sc _{0.7} O _{3-δ}	0.48	—	2×10^{-3} (Ar, 600 °C)	58
BaTi _{0.5} In _{0.5} O _{3-δ}	0.48	—	2.1×10^{-4} (Ar, 500 °C)	59
BaTi _{0.2} In _{0.8} O _{3-δ}	—	0.42	1.1×10^{-3} (450–600 °C)	57

the tendency for partial phase segregation, probably on nanometric-length scales, observed during the phase formation of the BaSn_{1-x}Sc_xO_{3-δ} series as discussed above.

4.4. Conductivity

The conductivity of BaSn_{0.6}Sc_{0.4}O_{3-δ} reveals a complex dependency on pO_2 and pH_2O . In dry, proton free, conditions the material reveals a major p-type contribution that is similar in magnitude to that reported recently for a related, highly acceptor doped, BaTi_{0.5}Sc_{0.5}O_{3-δ}.⁵² The behaviour is also comparable to Ba₂SnYO_{5.5},⁹ BaSn_{1-x}Y_xO_{3-δ} (ref. 12) and BaZr_{0.8}Y_{0.2}O_{3-δ} (ref. 53) for which hole conduction dominates at high $p(O_2)$ and high T . This enhanced conductivity can be rationalised through the partial filling of oxygen vacancies resulting in the formation of mobile electron holes as described in eqn (3) above.

Compared with the dry gas conditions, the conductivity under wet gas conditions (Ar or O₂), and that obtained from the initial heating run in dry Ar on the pre-hydrated sample, were significantly higher in the intermediate temperature region of 150–650 °C (200–450 °C in wet O₂) due to protons acting as the main charge carriers. Above 650 °C the conductivity obtained under wet and dry Ar were very similar as the material had probably dehydrated. This is characteristic of several related proton-conducting systems^{9,54} and reflects a transition to predominant oxide ion as protonic defects become unstable at higher T . Remarkably, at $T > 450$ °C under oxidizing conditions (Fig. 10), the conductivity is lower under wet gas compared to dry gas, although overall it still remains significantly higher than under wet Ar. This can be rationalised on the basis of the strong p-type character of the material that leads to a competition between holes and protons in wet oxidising conditions. A combination of the two reactions described by eqn (2) and (3) above would result in the consumption of holes in the presence of water vapour as per eqn (5) below.



This consumption of holes depletes the number of available $h\cdot$ charge carriers and hence lowers the total conductivity at $T > 450$ °C to values below those of the dry oxygen condition. This kind of behaviour has been observed for BaZr_{0.9-x}Pr_xGd_{0.1}O_{3-δ} where highly mobile electron holes ($h\cdot$) were found to dominate conductivity even to very low temperatures and in wet conditions.⁵⁵ It is clear that within the T range ~400–600 °C, BaSn_{0.6}Sc_{0.4}O_{3-δ} displays significant mixed proton and electron conduction in wet oxygen, indicating potential suitability as a cathode material for PCFCs.

Table 3 lists a summary of obtained conductivity parameters for a number of acceptor doped BaSnO₃ and BaTiO₃ systems reported in the literature. For BaSn_{0.6}Sc_{0.4}O_{3-δ}, it was possible to separate a bulk contribution for the initial heating run of the pre-hydrated sample at temperatures up to 350 °C. As apparent from the impedance data shown in Fig. 8 and the Arrhenius plot presented in Fig. 9, the total protonic conductivity is dominated by the highly resistive grain boundaries at relatively low temperatures. The bulk proton conduction in fact reaches a very high value of $\sim 2 \times 10^{-3}$ S cm⁻¹ at 350 °C (Fig. 9). The activation energy of bulk proton conductivity of BaSn_{0.6}Sc_{0.4}O_{3-δ}, estimated from the heating cycle of the pre-hydrated sample in the T range 100 to 250 °C, was 0.40(1) eV. This shows reasonable agreement with the 0.52 eV reported for both of the more lightly Sc substituted BaSnO₃ phases^{10,56} but is closer to the 0.38 eV and 0.34 eV reported for bulk proton conduction in BaSn_{0.875}Y_{0.125}O_{3-δ} (ref. 10) and BaSn_{0.75}Y_{0.25}O_{3-δ},⁵⁶ respectively. The bulk activation energies reported by Wang *et al.* for BaSn_{1-x}Y_xO_{3-δ} ($0.05 \leq x \leq 0.375$)¹² are approximately 0.1 eV lower than the 0.51 eV reported for bulk proton conductivity in Ba₂YSnO_{5.5} (ref. 9) suggesting that the long range B-site ordering found in this phase is not beneficial to proton migration. Our present findings for BaSn_{0.6}Sc_{0.4}O_{3-δ} support this trend in as much as



a low (0.40 eV) activation energy for bulk proton mobility is observed. Direct comparison of bulk activation energies is, however, not straightforward and values extracted from impedance data may also reflect partial contributions from, for example, defect formation enthalpies, dopant to proton trapping and the effects of grain boundaries. Therefore, although the trend from the data on acceptor doped BaSnO₃ (Table 3) seemingly supports more facile bulk proton diffusion in disordered systems we avoid drawing wider conclusions in regards to the impact of B-site ordering on proton mobility in perovskites.

The activation energies for the total protonic conduction of BaSn_{0.6}Sc_{0.4}O_{3-δ} lie in the range 0.67–0.70 eV in wet Ar which is lower than the 0.87 eV recently reported for BaSn_{0.875}Sc_{0.125}O_{3-δ},¹⁰ and closer to the 0.73 eV obtained for BaSn_{0.75}Sc_{0.25}O_{3-δ}.⁵⁶ The total proton conductivity of 1.07×10^{-3} S cm⁻¹ obtained for BaSn_{0.6}Sc_{0.4}O_{3-δ} in wet Ar at 600 °C is similar to that of BaIn_{0.8}Ti_{0.2}O_{2.6} (ref. 57) and BaTi_{0.3}Sc_{0.7}O_{3-δ},⁵⁸ and is significantly higher than that reported previously for BaSn_{1-x}Sc_xO_{3-δ} with lower scandium contents.^{10,11} This behaviour probably reflects the greater proton concentration in the more highly doped system. A trend of increasing proton conduction with increasing dopant concentration might be emerging from Table 3. Ultimately, however, it is the proton mobility, and understanding how it is influenced by factors such as the level of B-site cation ordering and the chemical nature of the ions, that is critical in order to obtain new materials with significantly enhanced proton conductivity.

5. Conclusions

Scandium substitution of the tin site within BaSnO₃ has been achieved by solid-state synthesis. Some degree of phase segregation was observed in the dry materials but it has largely disappeared in the BaSn_{0.6}Sc_{0.4}O_{2.8} sample after D₂O treatment. Analysis of X-ray and neutron diffraction data has indicated an average cubic symmetry of space group *Pm* $\bar{3}$ *m* and the deuteron position was successfully located at the 24k site (0.579(3), 0.217(3), 0) from Rietveld analysis.¹¹⁹ Sn solid-state NMR revealed a series of local tin environments consistent with 6 and 5 coordinate Sn environments and BaSnO₃ impurities. The resonance from the 6-coordinate site is broad indicating a wide range of Sn environments differing in the number of Sn and Sc cation in the 1st B-site cation coordination shell. This behaviour is very different to the structure of BaSn_{1-x}Y_xO_{3-δ} with a high concentration of yttrium, in which Y–O–Sn ordering occurs. The five-coordinated Sn is observed at high Sc doping levels ($x \geq 0.4$), confirming the presence of oxygen vacancies nearby tin. Conversely, the ⁴⁵Sc NMR data showed the existence of intense peaks for five-coordinated Sc, the relative fraction of five to six-coordinated Sc increasing with decreasing Sc content, suggesting preferential trapping of oxygen vacancies in between or near Sc cations at lower Sc concentrations. For all compositions, the five coordinated Sc and Sn environments vanished after hydration as OH groups filled the available oxygen vacancies.

BaSn_{0.6}Sc_{0.4}O_{3-δ} was found to be predominantly a p-type conductor under oxidizing atmospheres with proton conduction dominating at lower temperatures. The competition

between holes and protons results in a suppression of the conductivity at $T > 450$ °C in wet oxidizing conditions in comparison to dry oxygen. This mixed ionic and electron hole conduction means that the material could be utilized in gas separation membrane applications or cathodes of proton conducting fuel cells. The current study showed that highly scandium substituted BaSnO₃ supports very high bulk proton conductivity, comparable to that reported for Y-doped BaZrO₃ and the double perovskite Ba₂YSnO_{5.5}. It is suggested that further work on the material should focus on the growth of large grained samples in order to reduce grain boundary resistance, and investigate the chemical stability of the material under CO₂ atmospheres. Given the present findings, solution synthesis routes aiming to create a more homogeneous mixing of the B-site ions at the atomic level are also of potential interest.

Authors' contribution

I. A., S. G. E., S. T. N. conceptualized and planned the project, F. G. K., I. A. synthesized the samples. F. G. K. did EIS experiments and EIS data analysed with C. S. K., F. G. K., I. A., S. T. N., S. H., S. G. E. performed neutron data collection and structural analysis. L. B. prepared the ¹⁷O enriched materials. L. B., I. H., Z. G. and F. B. carried out the NMR experiments. L. B., F. B. and C. P. G. performed analysis of the NMR data. The manuscript was written with contribution from all co-authors.

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