



Cite this: *Photochem. Photobiol. Sci.*, 2016, **15**, 731

Received 11th April 2016,

Accepted 29th April 2016

DOI: 10.1039/c6pp00102e

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## Photooxygenation of alkanes by dioxygen with *p*-benzoquinone derivatives with high quantum yields†

Kei Ohkubo,<sup>a,\*</sup> Kensaku Hirose<sup>b</sup> and Shunichi Fukuzumi<sup>a,c,d</sup>

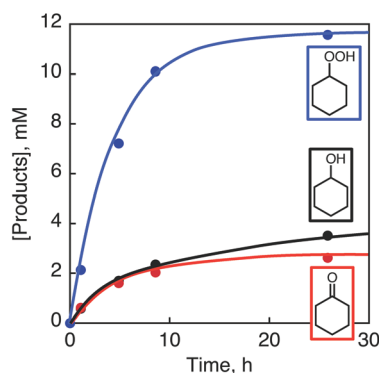
Alkanes were oxygenated by dioxygen with *p*-benzoquinone derivatives such as *p*-xyloquinone in alkanes which are used as solvents to yield the corresponding alkyl hydroperoxides, alcohols and ketones under visible light irradiation with high quantum yields ( $\Phi = 1000, 1600\%$ ). The photooxygenation is started by hydrogen atom abstraction from alkanes by the triplet excited states of *p*-benzoquinone derivatives as revealed by laser-induced transient absorption spectral measurements.

Oxidation of alkanes, which are saturated hydrocarbons, to produce basic raw materials in chemical industry is quite difficult and requires harsh conditions. For example, the conventional industrial process for oxidation of cyclohexane to a mixture of cyclohexanol and cyclohexanone known as KA-oil, which are important intermediates for the manufacture of nylon-6 and nylon-6,6, requires the use of soluble cobalt or manganese salts as homogeneous catalysts at 140–180 °C and 0.8–2 MPa (7.9–19.8 atm) in air.<sup>1,2</sup> So far extensive efforts have been devoted to develop catalytic oxidation of cyclohexane using hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>),<sup>3–6</sup> *tert*-butyl hydroperoxide (TBHP),<sup>6,7</sup> iodosylbenzene (PhIO),<sup>8</sup> ozone (O<sub>3</sub>),<sup>9</sup> and dioxygen (O<sub>2</sub>).<sup>3,10–12</sup> as oxidants. Among various oxidants, O<sub>2</sub> is an ideal oxidant because it is ubiquitous and the reduced product (H<sub>2</sub>O or H<sub>2</sub>O<sub>2</sub>) is safe. Oxygenation of cyclohexane by O<sub>2</sub> was made possible by using the 9-mesityl-10-methylacridinium ion as an organic photocatalyst combined with HCl in O<sub>2</sub>-saturated acetonitrile (MeCN) at room temperature under visible light irradiation.<sup>13</sup> The photochemical oxygenation of cyclohexane

is initiated by electron transfer from Cl<sup>−</sup> to the mesitylene radical cation moiety of the electron-transfer state of the 9-mesityl-10-methylacridinium ion to produce Cl<sup>•</sup> that abstracts a hydrogen atom from cyclohexane to generate a cyclohexyl radical, which is readily oxygenated by O<sub>2</sub>.<sup>13</sup> There has been no report on the photochemical oxygenation of cyclohexane with an organic photooxidant and without an inorganic compound such as HCl.

We report herein inorganic-free oxygenation of cyclohexane by O<sub>2</sub> using *p*-benzoquinone derivatives as the only organic photocatalyst in O<sub>2</sub>-saturated cyclohexane under visible light irradiation. Other alkanes were also oxygenated by O<sub>2</sub> with *p*-benzoquinone derivatives in O<sub>2</sub>-saturated alkanes under visible light irradiation. The photooxygenation mechanism is clarified by laser-induced transient absorption measurements.

Photoirradiation of an O<sub>2</sub>-saturated cyclohexane solution (3.0 cm<sup>3</sup>) containing *p*-xyloquinone (PXQ, 2.0 mM) for 26 h with a xenon lamp ( $\lambda > 390$  nm) resulted in the production of oxygenated products (Fig. 1), which were detected by



**Fig. 1** Reaction time profiles for the formation of cyclohexyl hydroperoxide (blue), cyclohexanol (black) and cyclohexanone (red) in the photochemical oxygenation of an oxygen-saturated cyclohexane solution containing PXQ (2.0 mM) under photoirradiation by the use of a xenon lamp (500 W;  $\lambda > 390$  nm) at 298 K. The GC charts for the determination of the yields are shown in Fig. S1 (ESI†).

<sup>a</sup>Department of Applied Chemistry, Graduate School of Engineering, Osaka University, Suita, Osaka 565-0871, Japan. E-mail: ookubo@chem.eng.osaka-u.ac.jp

<sup>b</sup>Department of Material and Life Science, Graduate School of Engineering, Osaka University, ALCA and SENTAN, Japan Science and Technology Agency (JST), Suita, Osaka 565-0871, Japan

<sup>c</sup>Department of Chemistry and Nano Science, Ewha Womans University, Seoul 120-750, Korea. E-mail: fukuzumi@chem.eng.osaka-u.ac.jp

<sup>d</sup>Faculty of Science and Technology, Meijo University, ALCA and SENTAN, Japan Science and Technology Agency (JST), Nagoya, Aichi 468-8502, Japan

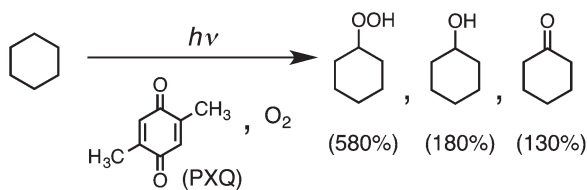
† Electronic supplementary information (ESI) available: Experimental details and GC data for the product analyses. See DOI: 10.1039/c6pp00102e



GC-MS (Fig. S1 in the ESI†).<sup>14</sup> The yields based on PXQ were 580% for cyclohexyl hydroperoxide, 180% for cyclohexanol and 130% for cyclohexanone (Scheme 1). Cyclohexyl hydroperoxide was also detected by the titrations of sodium iodide<sup>15</sup> and triphenylphosphine<sup>12</sup> (see the Experimental section in the ESI†). The conversion of neat cyclohexane was 22%. The quantum yield of the photooxygenation was determined to be as high as 100% from the initial rate using monochromatised light ( $\lambda = 420$  nm) (see the Experimental section in the ESI†).

When PXQ was replaced by *p*-benzoquinone (Q), the yields were decreased to 16% for cyclohexanol and 14% for cyclohexanone (Fig. S2a in the ESI†). In the case of tetramethyl-*p*-benzoquinone, no oxygenated product was formed under the otherwise same experimental conditions. No photoreactivity of tetramethyl-*p*-benzoquinone may result from the steric effect of the methyl group.

Photooxygenation of branched alkanes such as 3-methylpentane with PXQ occurred to produce the corresponding alcohol and ketone products (Table 1, entry 2) with the total yield of 1000% based on the initial concentration of PXQ (2.0 mM). In the case of normal-chain alkanes such as *n*-hexane and *n*-pentane, the efficient catalytic oxygenation also took place with the high quantum yields (510% for *n*-hexane and 140% for *n*-pentane) as shown in Table 1 (GC data are shown in Fig. S3–S5 in the ESI†).



Scheme 1

Nanosecond laser flash photolysis was employed to observe the hydrogen-transfer dynamics of the triplet excited state of PXQ ( $^3\text{PXQ}^*$ ; \* denotes the excited state) in cyclohexane. The transient absorption spectra were recorded at 0.1  $\mu\text{s}$  after nanosecond pulse irradiation at 355 nm of a deaerated cyclohexane solution (spectrum in black) and acetonitrile solution (red) containing PXQ as shown in Fig. 2a. The transient absorption bands appearing at 440 nm and 405 nm are assigned to  $^3\text{PXQ}^*$ <sup>16</sup> and  $\text{PXQH}^*$ ,<sup>17</sup> respectively. The decay at 440 nm obeyed pseudo-first-order kinetics as shown in Fig. 2b. The decay rate constants ( $k_{\text{obs}}$ ) of  $^3\text{PXQ}^*$  increased with increasing concentrations of cyclohexane (Fig. 2c). The rate constant of hydrogen transfer from cyclohexane to  $^3\text{PXQ}^*$  ( $k_{\text{H}}$ ) was determined to be  $1.3 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$  from the slope of the plot of  $k_{\text{obs}}$  vs. [cyclohexane] (Fig. 2c). The value in the case of cyclohexane-*d*<sub>12</sub> ( $k_{\text{D}}$ ) was also determined to be  $4.3 \times 10^6 \text{ M}^{-1} \text{ s}^{-1}$  (Fig. 2c; red plot). The deuterium kinetic isotope effect ( $\text{KIE} = k_{\text{H}}/k_{\text{D}}$ ) was determined to be 3.0. Such a large KIE value indicates that the rate-determining step of the photochemical oxygenation is hydrogen-atom transfer from cyclohexane to the triplet excited state of PXQ ( $^3\text{PXQ}^*$ ), which produces a cyclohexyl radical and  $\text{PXQH}^*$ . The hydrogen-atom abstraction from various substrates by the triplet excited states of ketones has been known for a long time.<sup>18–21</sup> However, the hydrogen-atom abstraction from cyclohexane has been difficult, because cyclohexane has been used as an inert solvent for transient absorption measurements of the triplet excited states of ketones.<sup>22</sup>

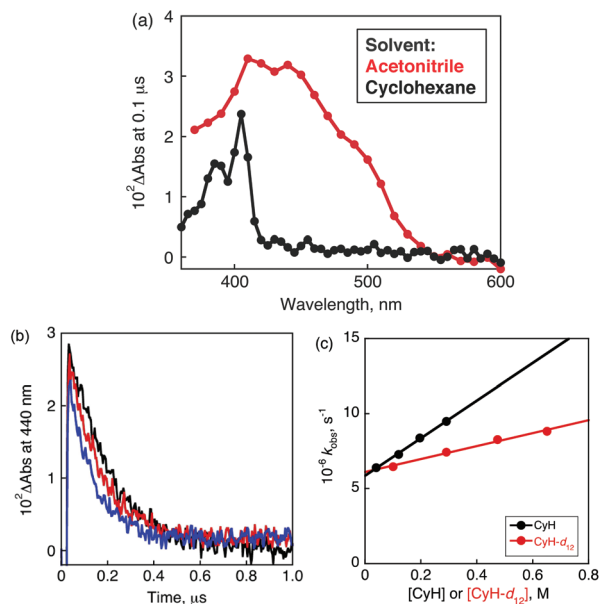
The photooxygenation of cyclohexane (CyH) is initiated by hydrogen abstraction from cyclohexane by  $^3\text{PXQ}^*$  to form the radical pair of a cyclohexyl radical ( $\text{Cy}^\bullet$ ) and a PXQ semiquinone radical ( $\text{PXQH}^\bullet$ ) (Scheme 2). The addition of  $\text{O}_2$  to  $\text{Cy}^\bullet$  affords a cyclohexylperoxyl radical ( $\text{CyOO}^\bullet$ ), which abstracts a hydrogen atom from cyclohexane to produce the cyclohexyl radical ( $\text{Cy}^\bullet$ ) and cyclohexyl hydroperoxide ( $\text{CyOOH}$ ). Cyclohexanol ( $\text{CyOH}$ ) and cyclohexanone ( $\text{Cy}=\text{O}$ ) as final oxygenated products are formed *via* the disproportionation of  $\text{CyOO}^\bullet$ .<sup>13,23</sup>

**Table 1** Product yields and quantum yields ( $\Phi$ ) for the photooxygenation of alkanes with PXQ in  $\text{O}_2$ -saturated alkanes under visible light irradiation

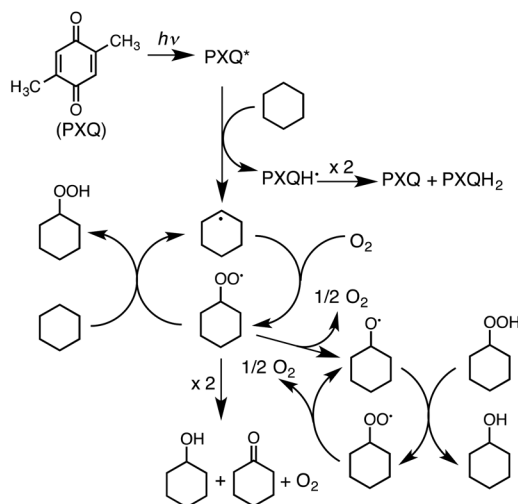
Entry	Substrate	Product (yields vs. $[\text{PXQ}]_0$ ) <sup>a</sup>	$\Phi^b$ %
1		(180%)    (130%)    (580%)	1000
2		(31%)    (39%)    (220%) (76%)    (39%)    (670%)	1600
3		(53%)    (38%)    (210%) (63%)    (49%)    (280%)	510
4		(78%)    (66%)    (310%) (35%)    (35%)    (310%)	140

<sup>a</sup>  $[\text{PXQ}]_0 = 2.0 \text{ mM}$ ,  $T = 298 \text{ K}$ . <sup>b</sup> Total quantum yield for the formation of oxygenated products. Excitation wavelength: 420 nm.





**Fig. 2** (a) Transient absorption spectra of a deaerated acetonitrile solution (red) or a deaerated cyclohexane solution (black) containing PXQ (12 mM) recorded at 0.1  $\mu$ s after nanosecond laser excitation at 355 nm. (b) Decay time profiles at 440 nm due to  $^3\text{PXQ}^*$  with cyclohexane [0 mM (black), 150 mM (red), 300 mM (blue)]. (c) Plots of  $k_{\text{obs}}$  for the decay at 440 nm vs. [cyclohexane] (black) and [cyclohexane- $d_{12}$ ] (red) in acetonitrile.



**Scheme 2** Proposed reaction mechanism for the photooxygenation of cyclohexane by dioxygen.

$\text{PXQH}^\bullet$  also disproportionates to  $\text{PXQ}$  and  $\text{PXQH}_2$ . However, the yield of  $\text{CyOH}$  (180%) is higher than that of  $\text{Cy=O}$  (130%) (Table 1, entry 1). The additional  $\text{CyOH}$  may be formed by decomposition of  $\text{CyOOH}$  to  $\text{CyOH}$  via the radical chain reactions, where the cyclohexyloxy radical ( $\text{CyO}^\bullet$ ) produced by disproportionation of  $\text{CyOO}^\bullet$  abstracts hydrogen from  $\text{CyOOH}$  to produce  $\text{CyOH}$ , accompanied by regeneration of  $\text{CyOO}^\bullet$ .

The extremely high quantum yields of the photooxygenation such as 1600% for 3-methylpentane and 1000% for cyclohexane result from these autooxidation radical chain processes.

In conclusion, efficient oxygenation reactions of alkanes are initiated by hydrogen-atom transfer from alkanes to the triplet excited state of *p*-xyloquinone acting as a photocatalyst under the solvent-free and ambient conditions, where molecular oxygen acts as an oxidant. Radical chain autooxidation processes with high concentrations of substrates (solvent) have enabled us to obtain the high quantum yields. Thus, the present study provides an environmentally benign way for photooxygenation of alkanes by  $\text{O}_2$  with unusually high quantum yields.

This work was supported by Grants-in-Aid (no. 16K13964, 26620154 and 26288037 to K. O.) from the Ministry of Education, Culture, Sports, Science and Technology (MEXT); ALCA and SENTAN projects from JST, Japan (to S. F.).

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