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## Efficient access to conjugated 4,4'-bipyridinium oligomers using the Zincke reaction: synthesis, spectroscopic and electrochemical properties†

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The cyclocondensation reaction between rigid, electron-rich aromatic diamines and 1,1'-bis(2,4-dinitrophenyl)-4,4'-bipyridinium (Zincke) salts has been harnessed to produce a series of conjugated oligomers containing up to twelve aromatic/heterocyclic residues. These oligomers exhibit discrete, multiple redox processes accompanied by dramatic changes in electronic absorption spectra.

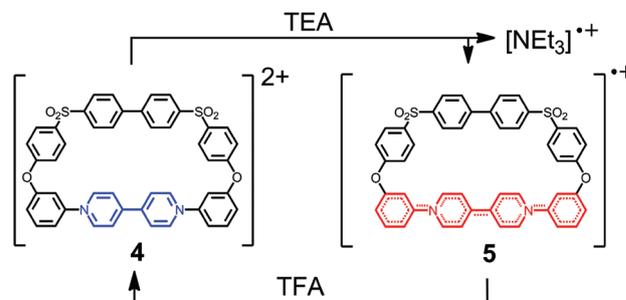
### Introduction

Bipyridinium residues are amongst the most intensively studied building blocks in supramolecular chemistry.<sup>1–3</sup> They have found widespread application in molecular systems that exhibit controlled switching,<sup>4–8</sup> rotational motion<sup>9–11</sup> and potential for data storage.<sup>12–15</sup> Many of these complex nano-systems harness the controllable electrochemical properties of 4,4'-bipyridinium ions (“viologens”) in either their synthesis or application.<sup>16</sup> The reversible one- and two-electron reductions of 4,4'-bipyridinium derivatives (see ESI – Scheme 1†) are accompanied by a dramatic change in the UV-vis absorption spectra of the materials, an effect which has been widely studied in the context of electrochromic devices.<sup>4,17–19</sup> The colour of cationic bipyridinium residues is highly dependent on the electronic ground state energy of the molecule, which can be finely tuned by varying the substituents.<sup>20–24</sup>

There have been a number of studies concerning the synthesis of oligomers and polymeric materials containing 4,4'-bipyridine units in the backbone. Typically, the electroactive 4,4'-bipyridinium groups are separated by at least one methylene residue.<sup>25–28</sup> This structural arrangement precludes electronic communication between the bipyridinium groups.<sup>29</sup> In addition, there have been several reports of the synthesis of so-called “extended viologens” whereby two pyridinium residues

are separated by increasing numbers of aromatic rings.<sup>30–40</sup> These have rapidly found widespread application in the construction of supramolecular complexes.<sup>41–46</sup> There are also examples of (di)phenyl viologen structures in which the conjugation length has been extended by the addition of an aromatic residue to the *N*-termini of the viologen.<sup>29,47–53</sup> Small molecules containing 4,4'- or 3,3'-bipyridine units,<sup>36,54–57</sup> and polymers<sup>58</sup> with 3,3'-bipyridine residues have been shown to exhibit interesting conductivity properties in the solid state.

Our own work in this area has previously concentrated on the synthesis of 4,4'-bipyridinium-containing macrocycles such as **4** (Scheme 1). We synthesized a family of related macrocycles by exploiting the Zincke reaction,<sup>59</sup> in which the condensation of nucleophilic amines with *N*-(2,4-dinitrophenyl)pyridinium salts resulted in efficient access to *N*-substituted pyridinium residues.<sup>60</sup> These macrocycles underwent one-electron reduction both chemically [*e.g.*, by triethylamine (TEA)] and electrochemically, to yield the corresponding radical cations such as **5**.<sup>60</sup> The unpaired spin density was found to be delocalized over four aromatic/heterocyclic rings,



**Scheme 1** Reduction of dicationic macrocycle **4** to radical monocation **5** in which the unpaired electron is delocalized over the four aromatic rings highlighted in red.

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but was prevented from more extensive delocalization by the break in conjugation as a consequence of the ether linkages in these macrocyclic systems.

Here we report the syntheses and spectroelectrochemical analyses of a series of redox-active oligomers containing progressively increasing numbers of 4,4'-bipyridinium units, culminating in the synthesis of a hexa-cationic trimer containing twelve aromatic/heterocyclic residues.

## Results and discussion

The syntheses of the new family of 4,4'-bipyridinium oligomers are shown in Scheme 2. The route to the targeted oligomers started from 1,1'-bis(2,4-dinitrophenyl)-[4,4'-bipyridin]-1,1'-dium dichloride (**6**), which can be readily prepared on a multi-gram scale.<sup>61</sup> As an example of the efficiency of the Zincke reaction, the reaction of Zincke salt **6** with dimethyl 5-aminoisophthalate **7** afforded unimer **1** in 93% (isolated) yield.

Higher molecular-weight oligomers in the series were accessed through intermediate **8** that was produced in 59% yield in two steps from known precursor **9**,<sup>62</sup> using simple extraction and crystallization procedures for purification. Condensation of two equivalents of **8** with the electron-rich aromatic diamine **10** resulted in the formation of dimer **2**, in 77% yield. In contrast, reaction of three equivalents of **10** with **6** gave diamine **11** (95%) that was converted smoothly to trimer **3** by reaction with two equivalents of precursor **8** (72.5%),

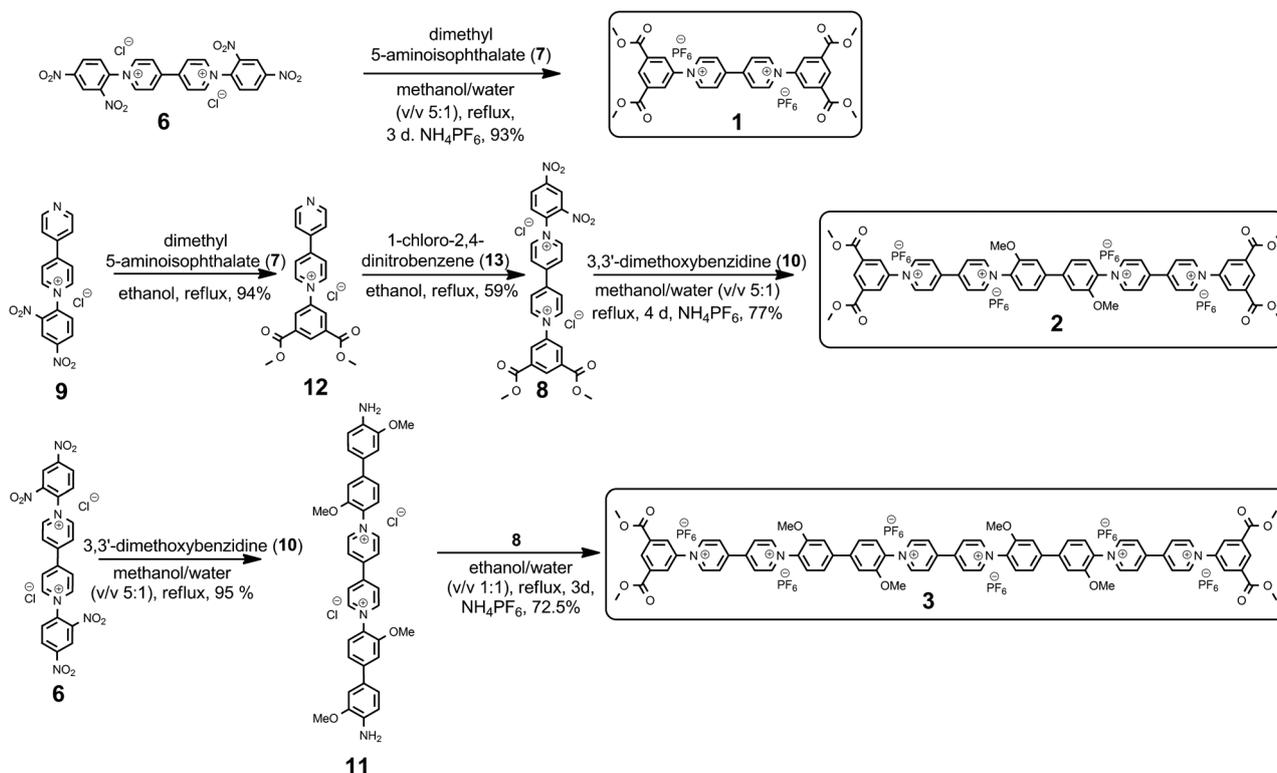
which was isolated as its hexafluorophosphate salt. Significant quantities of all three oligomers (*e.g.*, *ca.* 1 g of trimer **3**) could be obtained using these protocols, without recourse to column chromatography until the final step.

Electronic absorption spectra of **1**, **2** and **3** in dimethylformamide (DMF) are presented in Fig. 1. The main difference between the spectra of **1** and **2** is the new absorption band of the dimer at *ca.* 380 nm, tailing into the visible region. Its intensity increases markedly with oligomer length. At the same time the dominant UV absorption at *ca.* 300 nm shifts slightly to higher energy.

We next studied the chemically reversible redox process using TEA and trifluoroacetic acid (TFA), as described previously for the macrocyclic systems.<sup>60</sup> Fig. 2(A) shows the <sup>1</sup>H NMR spectrum of an orange solution of trimer **3**, which exhibits clearly resolved signals in the aromatic region (7.5 to 10.0 ppm), as well as two distinct resonances from the methoxy-ether and methoxy-ester residues (3.8 and 4.0 ppm).

The protons of the pyridinium residues at the termini of the oligomer, closest to the electron withdrawing ester groups, resonate at lower field (*H<sub>a</sub>* at 9.80 ppm) compared to the signals from the bipyridinium residue in the central section (protons *H<sub>e</sub>*, at 9.52 ppm). Thus, the three bipyridinium residues are in two distinct environments: two at the termini of the oligomer and one in the centre.

Upon addition of excess triethylamine, all the signals corresponding to the protons of the aromatic and heterocyclic rings disappear, as a consequence of the formation of a para-



Scheme 2 Synthesis of conjugated 4,4'-bipyridinium oligomers **1**, **2** and **3**.



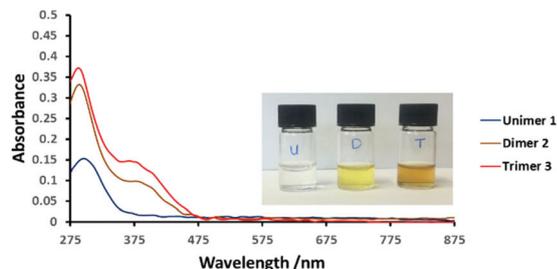


Fig. 1 UV-Vis spectra of compounds **1**, **2** and **3** in dimethylformamide (DMF, 0.2 mM) at 25 °C. Inset: Photograph of the compounds in DMF solution (0.2 mM).

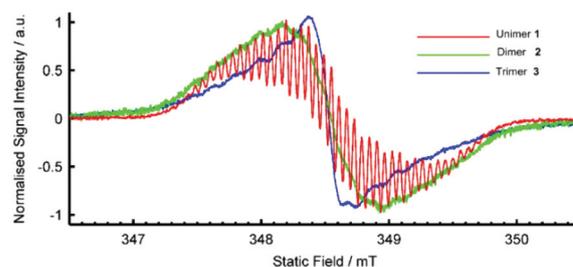


Fig. 3 EPR Spectra of radical species generated from **1**, **2** and **3** (1 mM) by addition of excess TEA in deaerated acetone solution at room temperature.

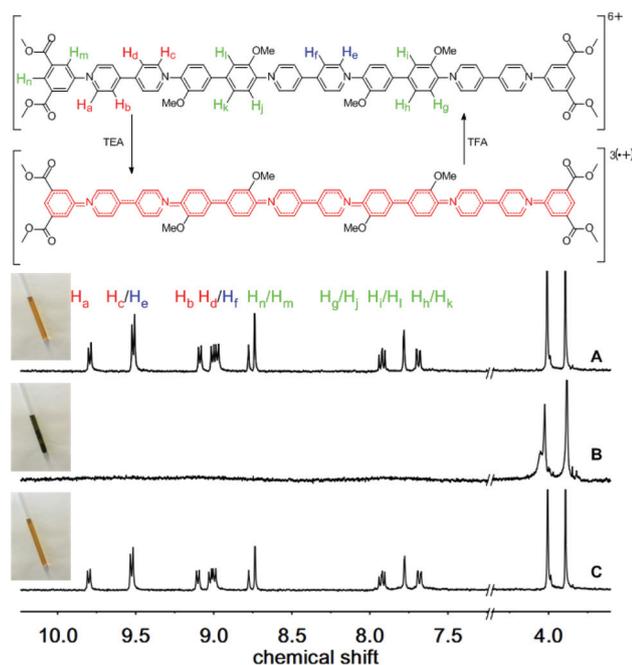


Fig. 2 Chemical reduction (3× one-electron) and reoxidation of trimer **3** using TEA and TFA, respectively. The aromatic and heterocyclic rings highlighted in red indicate the region of unpaired spin density (not bond order).  $^1\text{H}$  NMR spectra and photographs of the samples: (A) trimer **3** in acetone- $d_6$ , (B) after addition of 20 equiv. of TEA and (C) subsequent to addition of an excess of TFA to B.

magnetic tris(radical cation) (Fig. 2B). This change in the  $^1\text{H}$  NMR spectrum is accompanied by a transformation in the color of the solution to a deep green, which is characteristic of the formation of radical cationic chromophores of this type.<sup>63</sup> However, the  $^1\text{H}$  NMR signals corresponding to the non-conjugated methoxy groups remain visible at *ca.* 3.8 and 4.0 ppm. Addition of TFA to the tris(radical cation) not only regenerates the original color of the solution but also restores the missing signals in the aromatic region of the  $^1\text{H}$  NMR spectrum (Fig. 2C; see also the ESI – Fig. S1 and S2† for **1** and **2**, respectively).

The radical cationic redox states of compounds **1**, **2** and **3** were investigated by electron paramagnetic resonance (EPR)

spectroscopy in acetone at room temperature. The EPR spectrum of  $1^{+\cdot}$  centred at  $g = 2.0034(3)$  is characteristic of a viologen radical cation, showing well-resolved hyperfine splittings arising from a pair of equivalent  $^{14}\text{N}$  nuclei and multiple groups of equivalent  $^1\text{H}$  nuclei (Fig. 3).<sup>64,65</sup> The bis(radical cationic) form of dimer **2** and the tris(radical cationic) form of trimer **3** are also EPR-active. This indicates that the electron spins of the dimer biradical do not exclusively pair to form a diamagnetic singlet state, though a half-field transition characteristic of a triplet state was not observed. The EPR signal of  $2^{2(+\cdot)}$  exhibits the same broad envelope as  $1^{+\cdot}$ , but the hyperfine structure is apparent only through weak shoulders. The EPR spectrum of  $3^{3(+\cdot)}$  also has unresolved hyperfine shoulders, though the peak-to-peak linewidth has narrowed. The changes in the spectral shape from unimer to dimer to trimer are characteristic of successive broadening of the hyperfine spectrum as may arise from increasingly rapid Heisenberg spin exchange or electron transfer processes.<sup>66</sup>

The electrochemical reduction of **1**, **2** and **3** was studied by cyclic voltammetry (CV) at a polished glassy carbon disc electrode, using anhydrous DMF as solvent, containing 0.1 M tetrabutylammonium hexafluorophosphate (TBAPF<sub>6</sub>) as supporting electrolyte. Ferrocene was used as internal reference. Cyclic voltammograms of compounds **1**, **2** and **3** are shown in Fig. 4 and Table 1. Two reversible one-electron cathodic waves for unimer **1** occur at  $E_{1/2} = -0.63$  and  $-0.84$  V vs.  $\text{Fc}/\text{Fc}^+$ . For

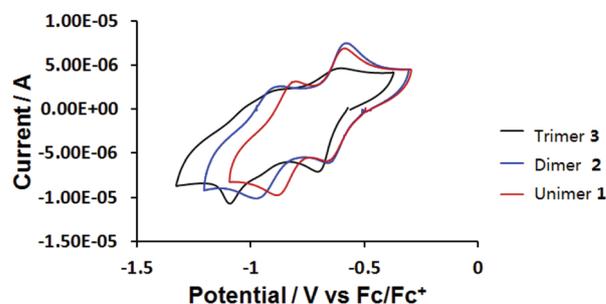


Fig. 4 Cyclic voltammograms of 0.2 mM compounds **1**, **2** and **3** at a glassy carbon disc ( $d = 2$  mm) electrode in anhydrous DMF at  $v = 500$  mV  $\text{s}^{-1}$ .



**Table 1** Cyclic voltammetric half-wave potentials observed on a glassy carbon disc ( $d = 2$  mm) electrode in anhydrous DMF at  $\nu = 500$  mV s<sup>-1</sup>

	CV = 0.5 V s <sup>-1</sup> $E_{1/2}$ (V) vs. Fc/Fc <sup>+</sup>	CV = 0.5 V s <sup>-1</sup> $E_{1/2}^2$ (V) vs. Fc/Fc <sup>+</sup>
Unimer 1	-0.63	-0.84
Dimer 2	-0.63	-0.92
Trimer 3	-0.66	-1.06

**Table 2** Square wave voltammetric peak and cyclic voltammetric half-wave potentials of trimer 3 observed on a glassy carbon disc ( $d = 2$  mm) electrode in anhydrous DMF

SQW, $f = 10$ Hz, $t_p = 25$ mV $E_p$ (V) vs. Fc/Fc <sup>+</sup>	CV, $\nu = 0.5$ V s <sup>-1</sup> $E_{1/2}$ (V) vs. Fc/Fc <sup>+</sup>
-0.65	-0.66
-0.93	Unresolved
-1.05	-1.06

dimer 2, these waves are found at  $E_{1/2} = -0.63$  and  $-0.92$  V. The data indicate that the two viologen moieties in tetracationic dimer 2 are electronically independent, as revealed by the identical first reduction potentials of 2 and dicationic compound 1, resulting from lack of conjugation through the twisted, central biphenyl rings. The second reduction potential of 2 is shifted to a more negative value compared to that of 1. This difference can be attributed to increased electronic conjugation in the more planar bis(radical cationic) form of 2 generated at the first cathodic wave.

Hexacationic compound 3 is reduced at only a slightly more negative potential than 1 or 2 ( $E_{1/2} = -0.66$  V,  $\Delta E_p = 100$  mV). As observed for 2, all three viologen units in 3 are reduced to their corresponding radical cations at the same electrode potential, suggesting that the viologens are twisted with respect to the chain direction and are therefore electronically independent. The cyclic voltammogram of 3 indicates a strong influence of adsorption phenomena at this scan rate (resolved cathodic waves and diminished anodic counter-waves). Although the resolution of the CV improved slightly at higher scan rates ( $\nu = 3$  V s<sup>-1</sup>; Fig. S8 in the ESI†), the rather unusual second cathodic step which ultimately converts the tris(radical cationic) form of 3 to the neutral species, remains indistinct.

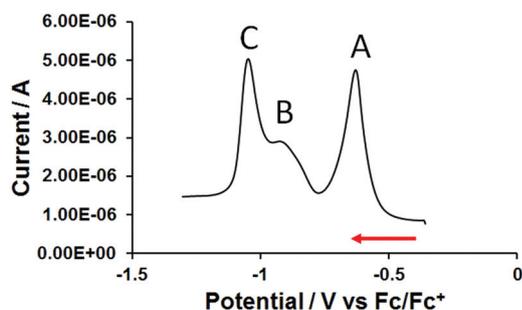
Thus, to further investigate the electrochemical properties of 3, a square wave voltammogram was acquired. Under these conditions three distinct reduction events were resolved (signals A, B and C in Fig. 5 and Table 2). Although the peak current of square wave voltammetry (SWV) signals can be influenced by solubility effects and molecular re-organisation during reduction, it would appear that reduction steps for the single central (signal B, Fig. 5) and two terminal bipyridinium

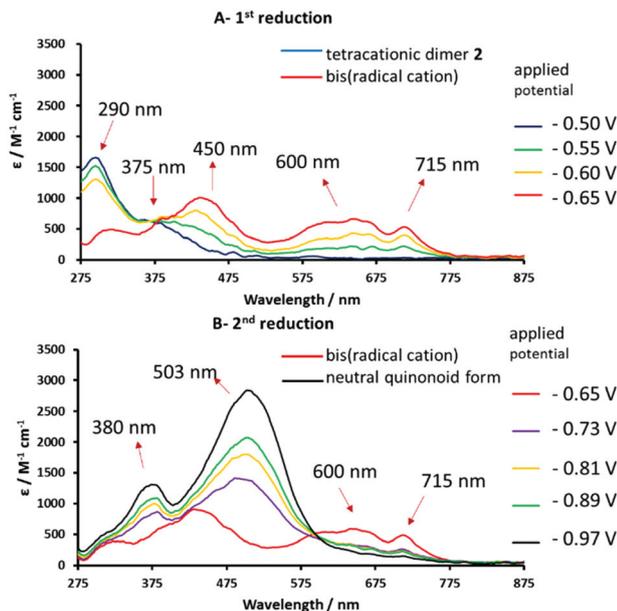
units (signal C, Fig. 5) to the neutral, quinoidal species can, nevertheless, be resolved during this experiment.

Infra-red spectroelectrochemistry (IR-SEC) was carried out on unimer 1 within an OTTLE cell<sup>67</sup> (Fig. S6 in the ESI†) to gain insight into the changes in the conjugated structure throughout the redox cycle. Formation of the viologen radical cation was accompanied by the appearance of a strong  $\nu(\text{C}=\text{C})$  band at 1639 cm<sup>-1</sup> associated with this species. There was, however, negligible change in the wavenumber of the  $\nu(\text{C}=\text{O})$  band at 1734 cm<sup>-1</sup> arising from the terminal ester groups. This observation suggests that the ester groups are not significantly conjugated with the aromatic system, so that their influence on the reduction potential will only be minor. Unfortunately, IR-SEC experiments with 2 and 3 were precluded by poor solubility of the reduced species in DMF at the high concentrations required to give reasonable absorbance values.

In order to study the redox-induced changes in electronic spectra of these systems, thin-layer ultraviolet-visible spectroelectrochemical (UV-vis SEC) measurements were carried out at 293 K with 0.2 mM 1, 2 or 3 in DMF/0.1 M TBAPF<sub>6</sub> (Fig. S11 in ESI†, 6 and 7). The observed spectral changes for all three species exhibit isosbestic points, excluding the possibility of side-reactions on the timescale of the experiment. All cathodic steps were fully reversible and parent electronic absorption spectra were recovered upon reoxidation. Fig. 6 show the UV-vis spectral monitoring of the electrochemical reduction of unimer 1 and dimer 2 at the two well-defined cathodic waves shown in Fig. 4. The tetracationic species shows two absorption bands at 290 and 375 nm (Fig. 6A). After completion of the first reduction step, the new absorption bands at 450, 600 and 715 nm are indicative of the bis(radical cation) (Fig. 6A and B).<sup>18</sup> Continuation of the cathodic sweep results in a second transformation in the electronic absorption spectrum due to formation of the neutral redox form absorbing at 380 and 503 nm (Fig. 6B; see also the ESI – Fig. S11† for 1).

Spectroelectrochemical (UV-vis) data for the four stable redox forms of trimeric species 3 were recorded at different reduction potentials corresponding to the three cathodic waves measured previously (Table 2). Conversion of the hexacationic species to the tris(radical cation) results in a spectrum similar to that observed for the analogous radical cation of 1 and bis(radical cation) of 2, with absorption maxima at 450, 600, 650 and 715 nm (Fig. 7A). Continued decrease in the applied cathodic potential results in the formation of a second stable species associated with absorption bands at 384 and 508 nm,

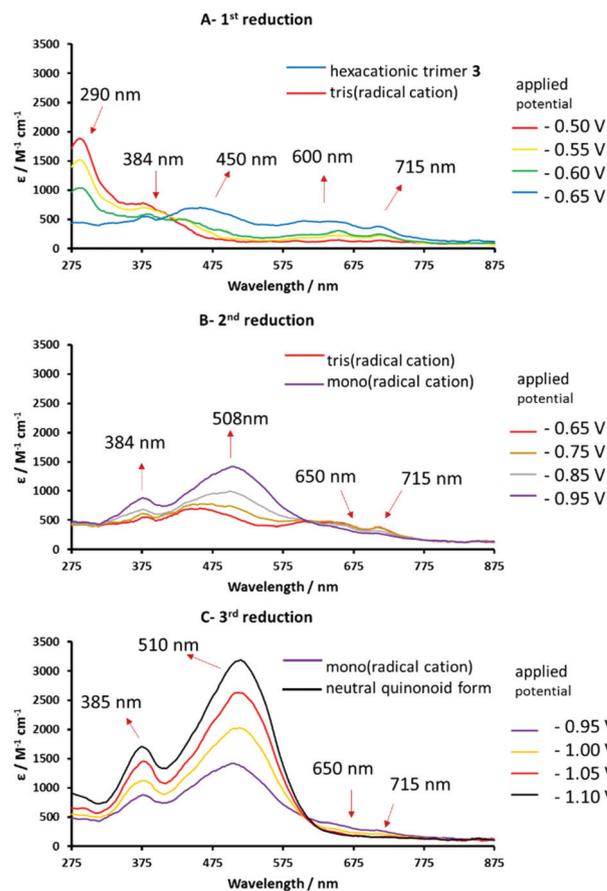
**Fig. 5** Square wave voltammogram of 0.2 mM trimer 3 on a glassy carbon disc electrode in anhydrous DMF/0.1 M TBAPF<sub>6</sub>.



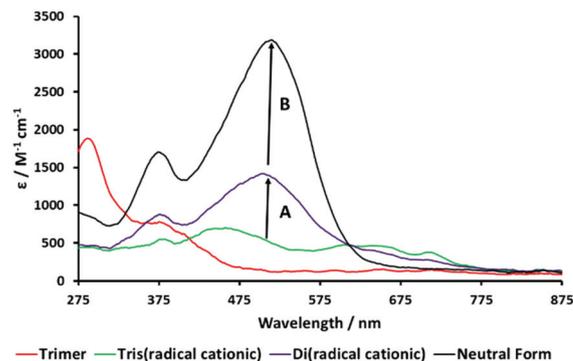
**Fig. 6** Reversible UV-vis spectral changes accompanying the stepwise 1e reduction of tetracationic dimer **2** to the corresponding bis(radical cation) (spectrum A) and the neutral quinonoid form (spectrum B). Spectra recorded in anhydrous DMF/0.1 M TABPF<sub>6</sub> using an OTTLE cell.<sup>67</sup>

which confirm the formation of fully reduced, quinoidal heterocycles. The absorption bands of the radical cation above 600 nm are diminished but still evident (Fig. 7B). This spectrum is consistent with a mixed oxidation state species containing both radical cationic and quinoidal heterocyclic species. Further lowering of the applied potential delivers the final species with intense absorption bands at 385 and 510 nm and no significant absorption bands above 600 nm (Fig. 7C), indicative of a fully reduced, neutral species.

Fig. 8 shows an overlaid plot of the UV-vis spectra of the four stable, spectroscopically distinct species in the redox cycle of **3**. The increase in intensity of the absorption band of the quinoidal bipyridinium species at 515 nm as the molecule becomes reduced from the tris(radical cation) to the mixed oxidation state system is 870 au (Arrow A). This compares to a change in intensity of 1800 au when moving from the mixed oxidation state species to the fully reduced, neutral trimer (Arrow B). These data suggest that the intermediate species, which is formed at approximately  $-0.9$  V (Table 2), contains a single quinoidal bipyridinium species compared to the three quinoidal species present in the ultimate, neutral form of **3**. Data from the SWV (Fig. 5) and the UV-vis SEC (Fig. 7 and 8) of **3**, supported by thin-layer cyclic voltammetry carried out during the latter experiment (see S14 in the ESI<sup>†</sup>), prove that the reduction of the three bipyridinium groups in the hexacationic trimer **3**, giving a tris(radical cation), occurs by the addition of three electrons at the same potential ( $-0.65$  V). However, the subsequent reduction of the tris(radical cation) to the neutral species occurs in a stepwise manner, by the

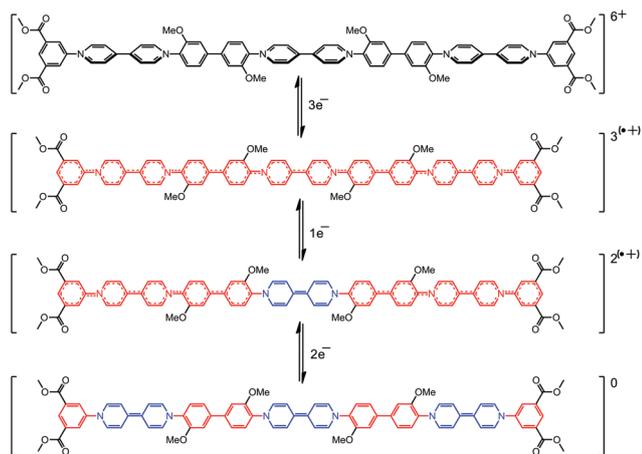


**Fig. 7** Reversible UV-vis spectral changes accompanying the stepwise reduction of hexacationic **3** to the corresponding tris(radical cation) (spectrum A) and the mono(radical cation) (spectrum B) and the neutral quinoidal form (spectrum C), recorded in anhydrous DMF/0.1 M TABPF<sub>6</sub> using an OTTLE cell.<sup>67</sup>



**Fig. 8** UV-vis spectral changes accompanying the stepwise reduction of hexacationic **3** to a neutral species. The change in intensity of the band at 515 nm associated with the formation of quinoidal bipyridinium residues occurs in the ratio 1:2 (arrows A and B respectively) during two-step reduction of the tris(radical cationic) form of **3** to the neutral species.





**Scheme 3** The structures of the stable intermediates observed during the electrochemical interconversions of trimer **3**: tris(dicationic), tris(radical cationic), di(radical cationic) and neutral, quinoidal forms (top to bottom respectively).

addition of a single electron (at  $-0.93$  V) followed by two electrons (at  $-1.05$  V). This can be accounted for if the central bipyridinium radical cation is reduced first, followed by the two terminal radical cations (Scheme 3).

## Conclusion

In this study we report an efficient synthesis of 4,4'-bipyridinium-based aromatic oligomers (**1–3**) and their spectro-electrochemical properties. Specifically, we use a controlled stepwise synthesis to produce single molecules with three 4,4'-bipyridinium residues containing, for the first time, twelve conjugated aromatic/heterocyclic rings. All the compounds undergo reversible stepwise two-electron reduction of each viologen moiety. Addition of the first electron to each viologen to form a radical cation produces a more planar structure with significantly greater electronic communication between the aromatic/heterocyclic groups along the length of each oligomer. Consequently, the tris(radical cation) of trimer **3** is reduced to the neutral, quinoidal form in two well-resolved cathodic steps.

Combined results from CV and UV-vis SEC demonstrate that, in these three oligomers, their ground states are not planar, which results in a twisted configuration between the phenyls and viologens. When going to the radical cationic state, because of the greater extent of  $\pi$  conjugation induced by their planar structure, the second  $E_{1/2}$  tends to a slightly more negative value. Spectroscopically, the UV-vis SEC result of **3** demonstrated that the apparently more electron rich, central radical cationic bipyridinium moiety is reduced to the quinoidal, neutral form *before* the two terminal bipyridinium units which appear comparatively electron poor. In support, computational studies of related conjugated oligomers and polymers have also shown that that maxima for their HOMO/LUMO coefficients are

found at the centre of the molecule.<sup>68,69</sup> If then the highest occupied spin orbital (HOSO) of the tris(radical cation) form of **3** is also localised in the centre of the molecule then the central 4,4'-bipyridinium would be reduced prior to the higher energy terminal radical cationic units. In addition, the UV-vis SEC measurements have highlighted the potential of these new molecules to find application in display technologies where dramatic and reversible colour changes are required. Work in our laboratory is continued towards integrating conjugated 4,4'-bipyridinium residues into more extended materials systems such as MOFs, mechanically interlocked molecules and high-MW polymers.

## Experimental

### General methods

Starting materials were purchased from Sigma Aldrich and Alfa Aesar and used without further purification unless otherwise stated. Anhydrous *N,N*-dimethylformamide (DMF) (Alfa Aesar, 99.8%, packaged under argon) was used as received. Tetra-*n*-butylammonium hexafluorophosphate (TBAPF<sub>6</sub>) was recrystallized twice from absolute ethanol and dried at 80 °C under vacuum overnight.

Compounds **6**<sup>61</sup> and **9**<sup>62</sup> were synthesized as described previously. All reactions were carried out in oven-dried glassware under dry nitrogen. <sup>1</sup>H and <sup>13</sup>C NMR were acquired on a Bruker DPX-400 spectrometer operating at 400 MHz and 100 MHz respectively or on a Bruker AV-700 spectrometer operating as 175 MHz for <sup>13</sup>C nuclei. Residual <sup>1</sup>H signals from the solvent were used as internal calibrants. Infrared spectra were recorded on a Perkin Elmer 17 20-X spectrometer from thin films cast from acetone, and major absorption bands are reported in wavenumbers (cm<sup>-1</sup>). For spectroelectrochemical (SEC) measurements, IR spectra were recorded on a Bruker Vertex 70v FTIR spectrometer. UV-Vis spectra were acquired on a Scinco S-3100 diode-array spectrophotometer.

Cyclic voltammetry (CV) measurements were performed on 0.2 mM solution of the compound **1**, **2** and **3** in anhydrous DMF containing 0.1 M TBAPF<sub>6</sub> as supporting electrolyte with a single-compartment three-electrode cell equipped with glassy carbon disc ( $d = 2$  mm) working, coiled platinum wire auxiliary and coiled Ag wire pseudoreference electrodes. Spectroelectrochemical measurements were carried out at 293 K, using an optically transparent thin-layer electrochemical (OTTLE) cell equipped with Pt minigrad working and auxiliary electrodes, a silver microwire pseudoreference electrode, and CaF<sub>2</sub> windows. The course of the spectroelectrochemical experiment was monitored by thin-layer cyclic voltammetry conducted with an EmStat 3 (PalmSens BV) potentiostat. The applied potentials stated in the UV-vis SEC are approximated from the CV of each compound. All the melting points were determined on a Gallenkamp melting point apparatus. Mass spectra were recorded Thermo Scientific LQT Orbitrap XL under conditions of electrospray ionisation. X-band EPR spectra were acquired on a Bruker EMX spectrometer using a TM<sub>110</sub> cylindrical mode resonator (ER 4103TM). Due to the high dielectric loss of the



solvent all samples were contained in 1 mm internal diameter quartz tubes (Wilmad 712-SQ). To allow resolution of hyperfine structure, samples were deoxygenated by a minimum of three freeze–pump–thaw cycles, and spectra were recorded with a 20  $\mu$ T modulation amplitude. A DPPH standard [ $g = 2.0036(3)$ ] was used to obtain accurate  $g$ -values.<sup>70</sup>

### Synthesis of dicationic unimer 1

1,1'-Bis(2,4-dinitrophenyl)-4,4'-bipyridinium dichloride (2.22 g, 3.96 mmol) was dissolved in a mixture of ethanol (50 mL) and water (50 mL) with excess dimethyl 5-aminoisophthalate (3.52 g, 16.83 mmol). The reaction mixture was heated to reflux and stirred vigorously for 3 days. After cooling to room temperature, the solvent was evaporated, the yellow solid was washed with THF (200 mL) and EtOAc (200 mL). Anion exchange to  $\text{PF}_6^-$  was achieved by dissolving the solid (4.60 g) in a mixture of water/methanol (8 : 2, v/v, 50 mL) and adding 10 g of  $\text{NH}_4\text{PF}_6$ . The resulting precipitate was filtered off, washed with 300 mL of water and 100 mL of ethyl acetate, and then dried under high vacuum overnight to yield **1** as a white solid (3.06 g, 93%).  $^1\text{H}$  NMR (acetone- $d_6$ , 400 MHz)  $\delta$  9.82 (d,  $J = 6.8$  Hz, 4H), 9.14 (d,  $J = 6.8$  Hz, 4H), 8.88 (t,  $J = 1.3$  Hz, 2H), 8.83 (d,  $J = 1.2$  Hz, 4H), 3.99 (s, 12H).  $^{13}\text{C}$  NMR (acetone- $d_6$ , 100 MHz)  $\delta$  164.9, 152.0, 147.7, 144.1, 133.9, 133.4, 131.0, 128.4, 53.5. IR ( $\text{cm}^{-1}$ )  $\nu = 1730$  ( $\alpha,\beta$ -unsaturated ester C=O). MS ( $m/z$ ) calc. for  $(\text{C}_{30}\text{H}_{25}\text{N}_2\text{O}_8)^+$ : 541.1605, found 541.1601; calc. for  $(\text{C}_{30}\text{H}_{26}\text{N}_2\text{O}_8)^{2+}$ : 271.0839, found 271.0840.

### Synthesis of 1-(3,5-bis(methoxycarbonyl)phenyl)-[4,4'-bipyridin]-1-ium dichloride 12

1-(2,4-Dinitrophenyl)-[4,4'-bipyridin]-1-ium chloride (**9**) (3.59 g, 10 mmol) was dissolved in ethanol (50 mL) with a large excess of dimethyl 5-aminoisophthalate (5.04 g, 24.9 mmol). The mixture was heated to reflux and stirred vigorously for 2 days. After cooling to room temperature, the solvent was evaporated and the product reprecipitated twice with ethyl acetate from methanol. The precipitate was collected by filtration, washed with THF (200 mL), and dried *in vacuo* to yield **12** as a light green solid (3.61 g, 94%) M.p. 243–246 °C (dec).  $^1\text{H}$  NMR ( $\text{D}_2\text{O}$ , 400 MHz)  $\delta$  9.29 (d,  $J = 6.8$  Hz, 2H), 8.94 (t,  $J = 1.2$  Hz, 1H), 8.81 (dd,  $J = 4.8$  and 1.6 Hz, 2H), 8.66 (d,  $J = 1.2$  Hz, 2H), 8.64 (d,  $J = 6.8$  Hz, 2H), 8.01 (dd,  $J = 1.6$  and 4.8 Hz, 2H), 4.01 (s, 6H).  $^{13}\text{C}$  NMR ( $\text{D}_2\text{O}$ , 100 MHz)  $\delta$  166.2, 155.5, 149.9, 144.9, 142.5, 142.3, 132.8, 132.81, 132.78, 129.4, 126.2, 53.3. IR ( $\text{cm}^{-1}$ )  $\nu = 1736$  (ester C=O). MS ( $m/z$ ) calc. for  $(\text{C}_{20}\text{H}_{17}\text{N}_2\text{O}_4)^+$ : 349.1183, found 349.1181.

### Synthesis of 8

1-(3,5-Bis(methoxycarbonyl)phenyl)-[4,4'-bipyridin]-1-ium dichloride (**6**) (1.35 g, 3.51 mmol) was dissolved in 8 mL of EtOH with a large excess of 1-chloro-2,4-dinitrobenzene (10.0 g, 49.4 mmol). The solvent was heated to reflux and stirred strongly for 3 days. After cooling to room temperature, the crystals were filtered off and washed with THF (2  $\times$  100 mL), further washed with EtOAc (2  $\times$  100 mL), and dried *in vacuo* to yield a light brown solid (1.41 g, 59%). M.p. 216–218 °C.  $^1\text{H}$

NMR ( $\text{D}_2\text{O}$ , 400 MHz)  $\delta$  9.53 (d,  $J = 6.8$  Hz, 2H), 9.47 (d,  $J = 7.2$  Hz, 2H), 9.42 (d,  $J = 2.4$  Hz, 2H), 8.99 (t,  $J = 1.2$  Hz, 1H), 8.96 (dd,  $J = 2.4$  and 8.8 Hz, 1H), 8.919 (d,  $J = 9.6$  Hz, 2H), 8.912 (d,  $J = 9.6$  Hz, 2H), 8.72 (d,  $J = 1.2$  Hz, 2H), 8.31 (d,  $J = 8.8$  Hz, 2H), 4.03 (s, 6H).  $^{13}\text{C}$  NMR ( $\text{D}_2\text{O}$ , 100 MHz)  $\delta$  168.7, 155.7, 153.6, 152.4, 149.3, 148.4, 145.3, 145.0, 140.8, 135.7, 135.5, 133.6, 133.2, 132.0, 130.1, 129.9, 125.3, 55.9. IR ( $\text{cm}^{-1}$ )  $\nu = 1715$  (ester, C=O), 1541 ( $\text{NO}_2$ ). MS ( $m/z$ ) calc. for  $(\text{C}_{26}\text{H}_{20}\text{N}_4\text{O}_8)^+$ : 561.1276, found 561.1266.

### Synthesis of tetracationic dimer 2

1-(3,5-Bis(methoxycarbonyl)phenyl)-1'-(2,4-dinitrophenyl)-[4,4'-bipyridin]-1,1'-dium chloride (**8**) (984 mg, 1.68 mmol) was dissolved in methanol (50 mL) with 3,3'-dimethoxybenzidine (204 mg, 0.84 mmol). The mixture was heated to reflux and stirred for 4 d. After cooling to room temperature, the solvent was evaporated, and the crude product was precipitated twice from MeOH with ethyl acetate to afford a brown solid. Anion exchange to  $\text{PF}_6^-$  was achieved by dissolving the solid in a mixture of water/methanol (10 mL) and adding  $\text{NH}_4\text{PF}_6$  (3 g). The resulting precipitate was filtered off, washed with 500 mL of water and 300 mL of ethyl acetate, and then dried under high vacuum overnight to yield **2** as a yellow solid (960 mg, 77%). M.p. 312 °C.  $^1\text{H}$  NMR (acetone- $d_6$  with 1% trifluoroacetic acid (TFA); the signal of TFA is not reported; 400 MHz)  $\delta$  9.92 (d,  $J = 6.8$  Hz, 4H), 9.64 (d,  $J = 7.2$  Hz, 4H), 9.21 (d,  $J = 6.8$  Hz, 4H), 9.12 (d,  $J = 6.8$  Hz, 4H), 8.90 (t,  $J = 1.4$  Hz, 2H), 8.87 (d,  $J = 1.2$  Hz, 4H), 8.04 (d,  $J = 8$  Hz, 2H), 7.91 (d,  $J = 1.6$  Hz, 2H), 7.69 (dd,  $J = 1.6$  and 8.4 Hz, 2H), 4.13 (s, 6H), 4.02 (s, 12H).  $^{13}\text{C}$  NMR ( $\text{CD}_3\text{CN}$  with 1% TFA; the signal of TFA is not included; 175 MHz)  $\delta$  165.3, 153.5, 152.2, 151.5, 148.4, 147.1, 145.4, 143.6, 134.3, 130.6, 128.5, 128.1, 121.6, 113.5, 27.9, 53.9. IR ( $\text{cm}^{-1}$ )  $\nu = 1722$  ( $\alpha,\beta$ -unsaturated ester C=O). MS ( $m/z$ ) calc. for  $(\text{C}_{54}\text{H}_{46}\text{N}_4\text{O}_{10})^{4+}$ : 227.5798, found 227.5798, calc. for  $(\text{C}_{54}\text{H}_{45}\text{N}_4\text{O}_{10})^{3+}$ : 303.1040, found 303.1040, calc. for  $(\text{C}_{54}\text{H}_{44}\text{N}_4\text{O}_{10})^{2+}$ : 3454.1523, found 454.1519.

### Synthesis of 1,1'-bis(4'-amino-3,3'-dimethoxy-[1,1'-biphenyl]-4-yl)-[4,4'-bipyridine]-1,1'-dium chloride 11

Zincke salt (**6**) (570 mg, 1.00 mmol) was dissolved in 50 mL of methanol/water (5 : 1, v/v) with 3,3'-dimethoxybenzidine (732 mg, 3 mmol). The reaction mixture was heated with stirring at reflux for 3 days, and then concentrated to 3 mL and treated with THF (100 mL). The crude product was precipitated twice from MeOH/EtOAc (1 : 8, v/v). The precipitate was filtered off and dried under high vacuum to afford **11** as a deep purple solid (653 mg, 95%). M.p. 280 °C (dec.).  $^1\text{H}$  NMR ( $\text{CD}_3\text{OD}$ , 400 MHz)  $\delta$  9.48 (d,  $J = 6.8$  Hz, 2H), 9.19 (d,  $J = 6.8$  Hz, 2H), 8.97 (t,  $J = 1.2$  Hz, 1H), 8.82 (d,  $J = 6.8$  Hz, 2H), 8.70 (d,  $J = 1.2$  Hz, 2H), 8.68 (d,  $J = 7.5$  Hz, 2H), 7.63 (d,  $J = 8.0$  Hz, 1H), 7.44 (s, 1H), 7.42 (d,  $J = 8.0$  Hz, 1H), 7.23 (d,  $J = 8.0$  Hz, 1H), 7.20 (s, 1H), 6.92 (d,  $J = 8.0$  Hz, 1H), 4.03 (s, 3H), 4.01 (s, 3H), 3.93 (s, 3H).  $^{13}\text{C}$  NMR ( $\text{CD}_3\text{OD}$ , 100 MHz)  $\delta$  165.7, 153.4, 152.5, 151.4, 149.9, 148.8, 148.5, 147.5, 145.3, 144.4, 137.0, 136.7, 134.6, 133.9, 131.6, 130.8, 130.5, 128.5, 128.1, 127.8, 124.8, 121.5, 120.5, 117.5, 111.9, 110.6, 57.3, 56.4, 53.6. IR ( $\text{cm}^{-1}$ )  $\nu =$



3350 (N–H, primary amine). MS ( $m/z$ ) calc. for  $(C_{38}H_{36}N_4O_4)^{2+}$ : 612.2731, found 612.2725.

### Synthesis of hexacationic trimer 3

Aromatic diamine **11** (403 mg, 0.59 mmol) was dissolved in 100 mL of ethanol/water (1 : 1, v/v) with precursor (**8**) (704 mg, 1.20 mmol) and the solution was refluxed for 3 d. The solvent was concentrated to leave approximately 5 mL, and this was added to a rapidly stirred solution of  $NH_4PF_6$  in a mixture of methanol and water (2 : 8, v/v, 30 mL). After stirring at room temperature for 1 h, the resulting precipitate was filtered off, washed with water (300 mL) and ethyl acetate (300 mL) and finally purified by column chromatography (elution gradient: acetonitrile to 1%  $NH_4PF_6$  in acetonitrile w/v). Excess salts were removed by precipitating the crude product twice from acetonitrile/water and the product was then dried under high vacuum to give a black solid, **3**, (920 mg, 72.5%). M.p. 325 °C (dec.).  $^1H$  NMR (acetone- $d_6$  with 1% TFA; the signal of TFA is not reported; 400 MHz)  $\delta$  9.92 (d,  $J = 6.8$  Hz, 4H), 9.64 (d,  $J = 7.2$  Hz, 8H), 9.22 (d,  $J = 7.2$  Hz, 4H), 9.13 (d,  $J = 6.8$  Hz, 4H), 9.10 (d,  $J = 6.8$  Hz, 4H), 8.89 (t,  $J = 1.2$  Hz, 2H), 8.86 (d,  $J = 1.6$  Hz, 4H), 8.05 (t,  $J = 7.2$  Hz, 4H), 7.91 (s, 4H), 7.82 (dd,  $J = 1.6$  and 8.0 Hz, 4H), 4.13 (s, 12H), 4.02 (s, 12H).  $^{13}C$  NMR (acetone- $d_6$  with 1% TFA; the signal of TFA is not included; 175 MHz)  $\delta$  165.0, 158.7, 158.3, 153.6, 152.2, 151.9, 148.7, 148.7, 147.7, 145.4, 144.2, 133.9, 132.2, 131.0, 128.5, 128.4, 128.2, 128.2, 121.5, 120.2, 117.4, 114.5, 113.3, 57.5, 53.5. IR ( $cm^{-1}$ )  $\nu = 1723$  ( $\alpha,\beta$ -unsaturated ester, C=O). MS ( $m/z$ ) calc. for  $(C_{78}H_{64}N_6O_{12})^{4+}$ : 319.1140, found 319.1140; calc. for  $(C_{78}H_{62}N_6O_{12})^{2+}$ : 637.2207, found 637.2210.

### Conflict of interest

The authors declare no competing financial interests.

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